

The background of the slide features a stylized, painterly illustration of a city skyline. The buildings are rendered in various shades of blue and teal, with some taller structures standing out. Above the skyline, there are several fluffy, white clouds with soft yellow and orange highlights, suggesting a sunrise or sunset sky. The overall aesthetic is clean and modern.

Lecture on air quality modeling: Photostationary state of O_3 and NO_x

Learning goals for today

Learn about photostationary state between O_3 and $NO_x (= NO + NO_2)$.

- Describe the spatial distribution of NO_2 over Europe
- Derive equations for the steady state between O_3 , NO and NO_2
- Use a Jupyter Notebook to perform calculations
- Understand main features in urban measurements of NO

A short recap

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$$p \times V = N \times R \times T$$

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mole fractions (e.g. mol/mol),
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- Ways to express abundance:
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mass concentrations (e.g. $\mu\text{g}/\text{m}^3$), etc.
- Mixing ratios of main gases in the
atmosphere (not including H_2O)

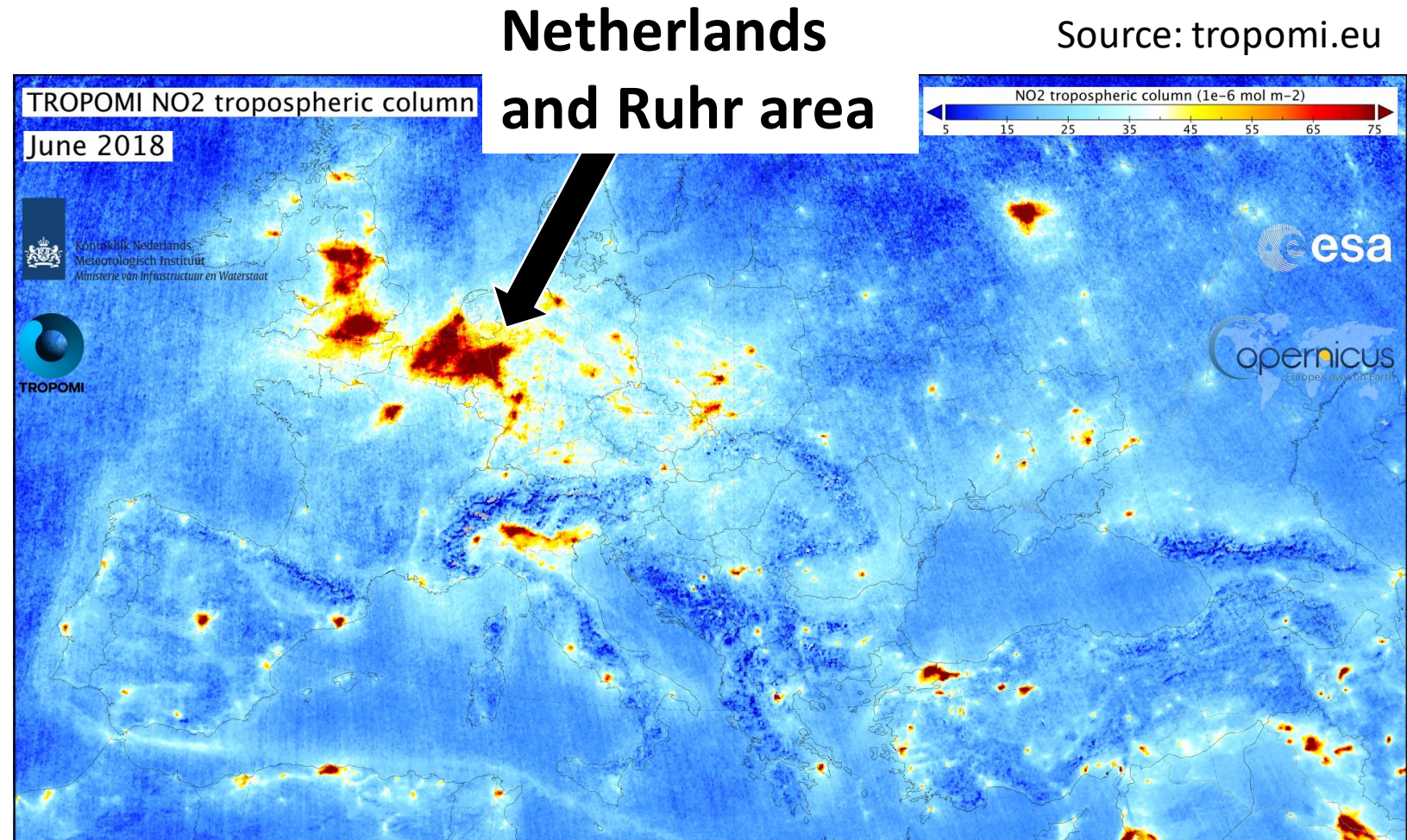
Table 1-1 Mixing ratios of gases in dry air

Gas	Mixing ratio (mol/mol)
Nitrogen (N_2)	0.78
Oxygen (O_2)	0.21
Argon (Ar)	0.0093
Carbon dioxide (CO_2)	365×10^{-6}
Neon (Ne)	18×10^{-6}
Ozone (O_3)	$0.01\text{--}10 \times 10^{-6}$

Source: Jacob (1999)

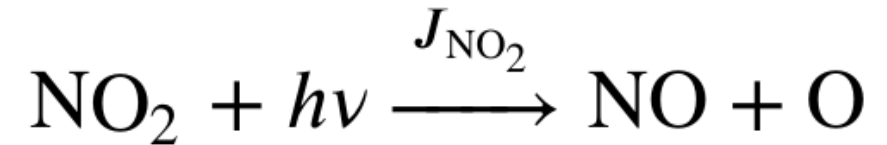
View of NO₂ from TROPOMI

- Satellites are powerful tools to study air quality
- Dutch instrument TROPOMI has a high spatial resolution
- The Netherlands is one of the hotspots for the pollutant NO₂



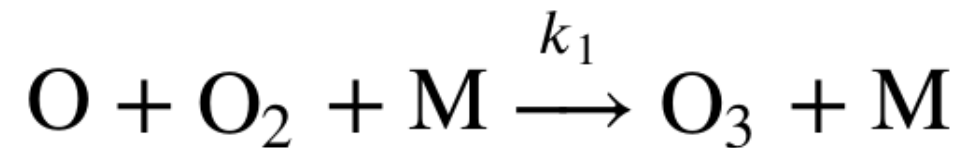
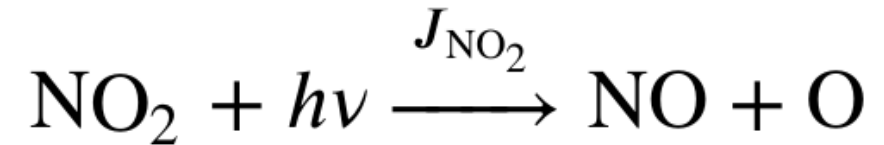
Chemical reactions

- Photolysis of NO₂ requires light and produces an oxygen radical (O)



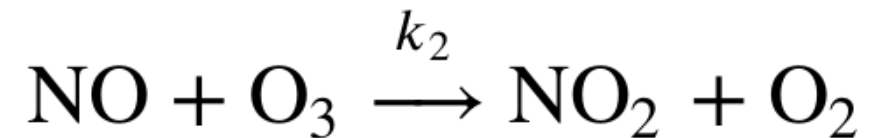
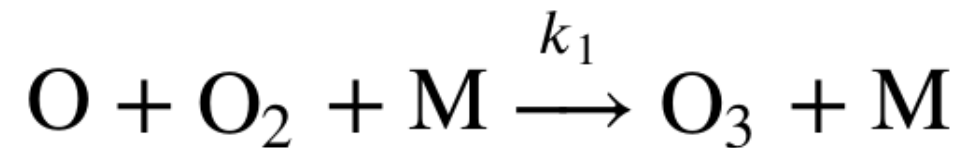
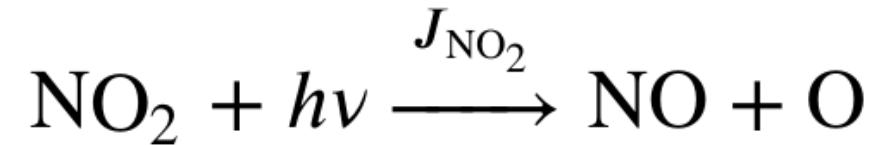
Chemical reactions

- Photolysis of NO_2 requires light and produces an oxygen radical (O)
- The oxygen radical (O) can combine with O_2 in the presence of a third body (M) to form ozone (O_3)



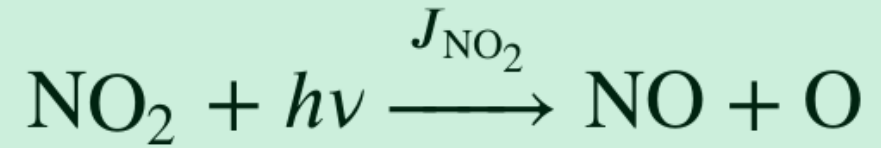
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- Finally, NO and O_3 can combine to re-create NO_2 and O_2

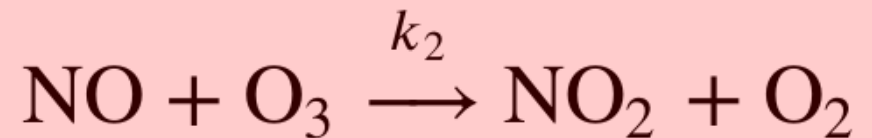
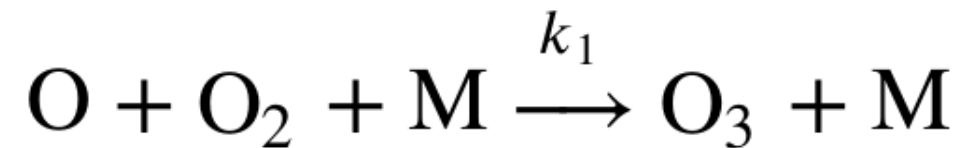


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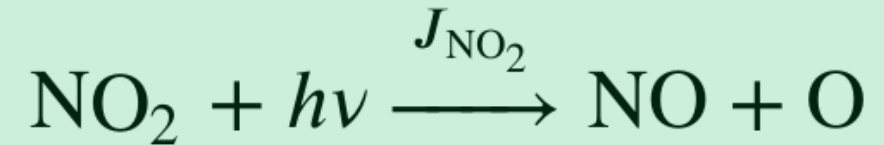
Production of NO



Destruction of NO

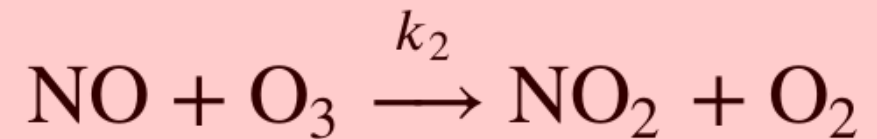
Conversion rates

$$[\text{NO}] \text{ production rate} = J_{\text{NO}_2} [\text{NO}_2]$$



Production of NO

$$[\text{NO}] \text{ destruction rate} = k_2 [\text{NO}][\text{O}_3]$$



Destruction of NO

Photostationary state

- Change in [NO] follows from the balance between production and destruction

$$\frac{d[\text{NO}]}{dt} = J_{\text{NO}_2} [\text{NO}_2] - k_2 [\text{NO}] [\text{O}_3]$$

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$J_{\text{NO}_2} \rightarrow 0$, so $[\text{NO}_2]/[\text{NO}]$ increases rapidly,
in other words NO_x will be present as NO₂

$$\frac{[\text{NO}_2]}{[\text{NO}]} = \frac{k_2 [\text{O}_3]}{J_{\text{NO}_2}}$$

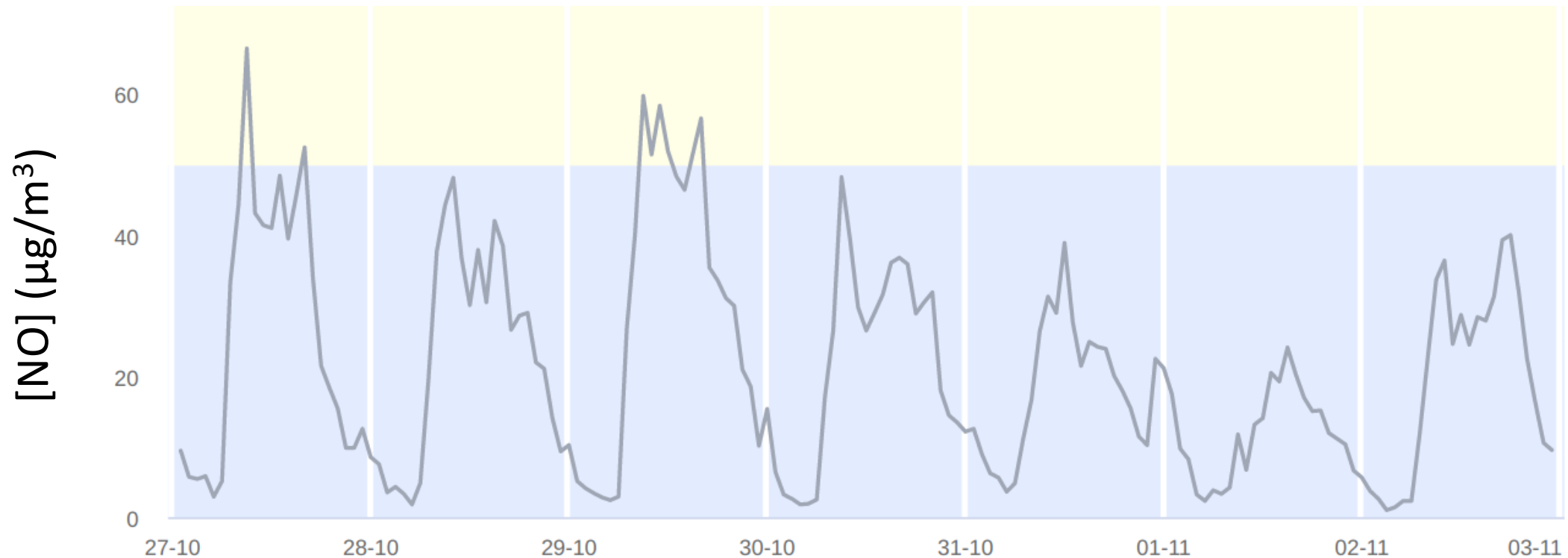
Jupyter Notebook

Source: jupyter.org



- An interactive tool to write text, code and create figures with your favorite programming language!

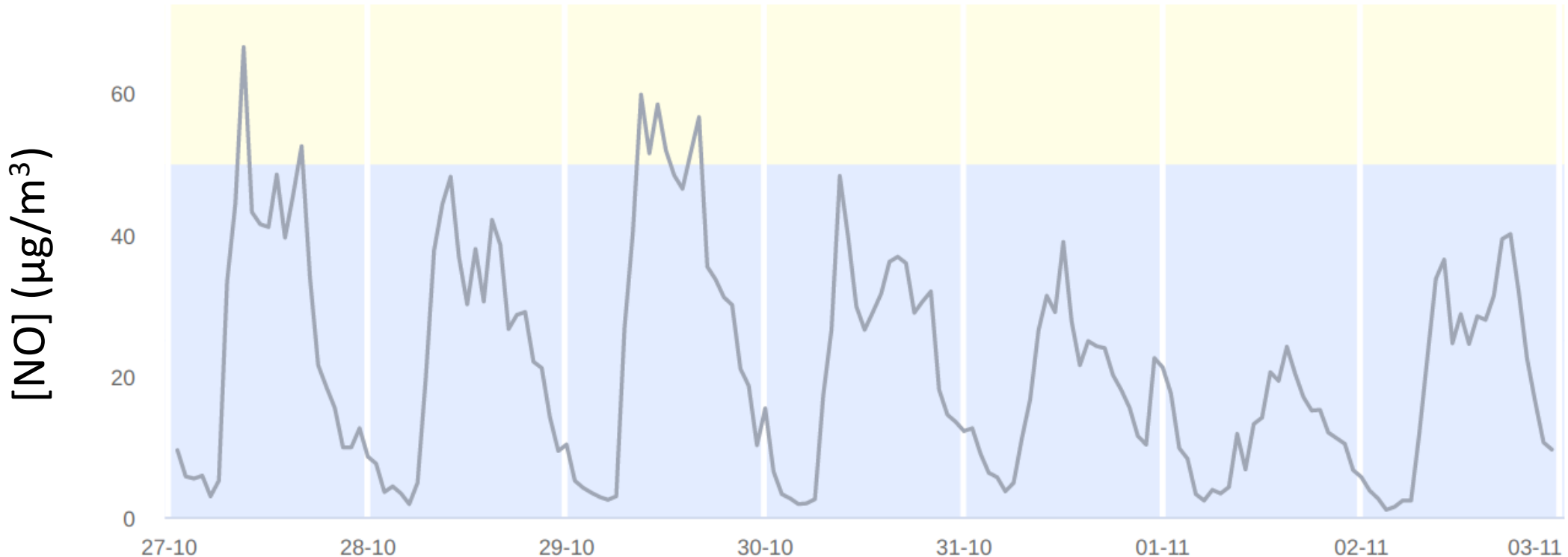
Urban air quality measurements



- Data from station Amsterdam-Haarlemmerweg for last 7 days (Luchtmeetnet.nl)

Urban air quality measurements

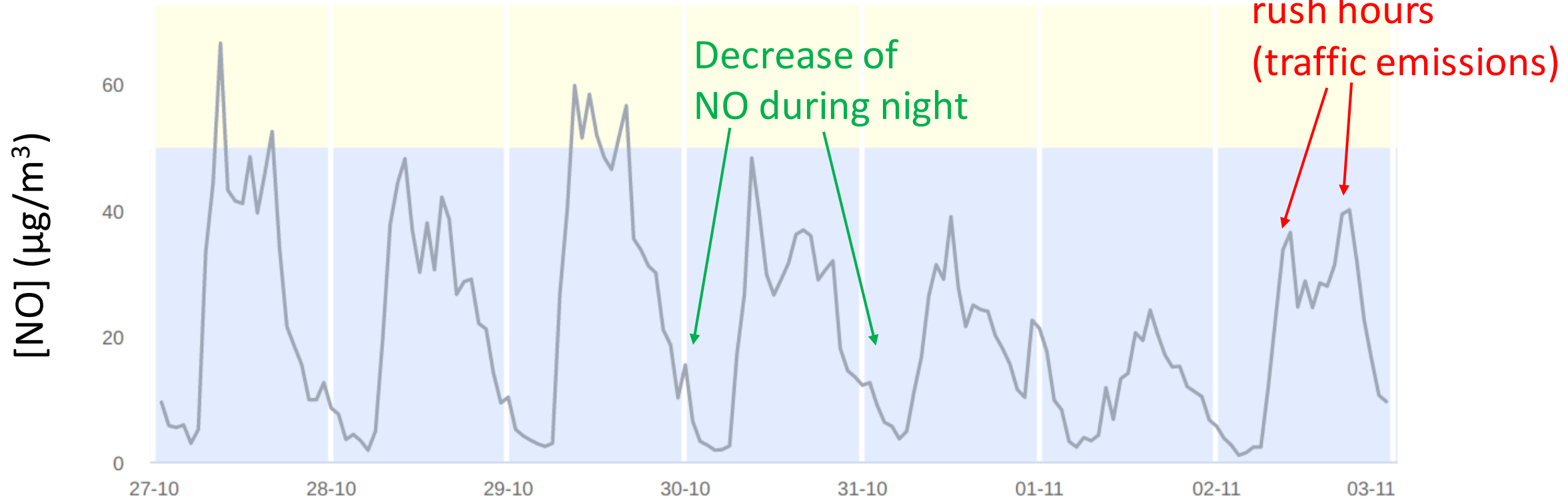
Q: Who can explain the measured signal?



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Urban air quality measurements

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Summary

Learned about photostationary state between O_3 and NO_x ($= NO + NO_2$)

- A view of NO_2 from TROPOMI: Netherlands is one of the hotspots
- Derived analytical equations for steady state between O_3 , NO and NO_2
- Worked with a Jupyter Notebook to explore these equations
- Observed peaks in NO related to traffic and decreases in NO during night in a dataset with urban measurements (Luchtmeetnet.nl).

Jupyter Notebook

Start the Jupyter Notebook in one of the following ways:

- Open the Notebook through the MyBinder link:
→ <https://mybinder.org/v2/gh/koren007/AQ/HEAD>

No additional software or account needed!

- Or: download the Notebook from GitHub and run locally:
→ <https://www.github.com/koren007/AQ>

Requires Anaconda installation (freely available) to run the Notebook



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