*Empty cells denote water-soluble salts.

**Colours of salts without specifically stated are white.

Anion		Cation																
		Na^+	K^+	Mg^{2+}	Ca^{2+}	Ba^{2+}	Al^{3+}	Zn^{2+}	Fe^{2+}	Fe^{3+}	Ni^{2+}	Sn^{2+}	Pb^{2+}	Cu^+	Cu^{2+}	Hg^{2+}	Ag^+	Mn^{2+}
O^{2-}				White (Solubilities increase)			White	Yellow (hot) White (cold)	Black	Reddish brown	Green	White	Orange (hot) Yellow (cold)	Red	Black	Yellow	Black brown	Black brown
NO_3^-																		
X^-	F^-																	
	Cl^-												White				White	
	Br^-												White				Creamy	
	I^-												White				Yellow	
SO_4^{2-}					White	White							White				White	
CO_3^{2-}					White	White												
CrO_4^{2-}						Yellow							Yellow				Reddish Brown	
NaOH	Conc.			White (Solubilities increase)														
	Diluted			White			White	White	Green	Brown	Green	White	White	White	Blue	Yellow (HgO)	Black brown (Ag_2O)	White
	Excess			White			Dissolved as $Al(OH)_4^-$	Dissolved as $Zn(OH)_4^{2-}$	Green	Brown	Green	Dissolved as $Sn(OH)_4^{2-}$	Dissolved as $Pb(OH)_4^{2-}$	White	Blue	Yellow (HgO)	Black Brown (Ag_2O)	White
NH_3	Excess			White			White	Dissolved as $Zn(NH_3)_4^{2+}$	Green	Brown	Dissolved as bluish violet $Ni(NH_3)_6^{2+}$	White	White	White	Dissolved as blue $Cu(NH_3)_4^{2+}$	Yellow (HgO)	Dissolved as $Ag(NH_3)_2^+$	
S^{2-}	Acidic							White	Black	Black	Black							
	Others							White	Black	Black	Black	Brown	Black	Black	Black	Black	Black	Pale red

Special:

CdS is a yellow precipitate regardless of the environment.

Rust indicators: $Fe_3[Fe(CN)_6]_2$, $Fe_4[Fe(CN)_6]_3$, $Fe(SCN)_3(aq)$