NEET(UG)–2017 TEST PAPER WITH ANSWER & SOLUTIONS (HELD ON SUNDAY 07th MAY, 2017)

1. A spring of force constant k is cut into lengths of ratio 1 : 2 : 3. They are connected in series and the new force constant is k'. Then they are connected in parallel and force constant is k'' . Then k' : k'' is : -

(1) 1 : 9 (2) 1 : 11

(3) 1 : 14 (4) 1 : 16

Ans. (2)

Sol. Length of the spring segments $= \frac{\ell}{6}, \frac{\ell}{3}, \frac{\ell}{2}$

As we know $K \propto \frac{1}{\ell}$

so spring constants for spring segments will be

$K_1 = 6K, \quad K_2 = 3K, \quad K_3 = 2K$

so in parallel combination

$K'' = K_1 + K_2 + K_3 = 11K$

in series combination

$K' = K$ (As it will become original spring)

so $K' : K'' = 1 : 11$

2. The ratio of resolving powers of an optical microscope for two wavelengths $\lambda_1 = 4000$ Å and $\lambda_2 = 6000$ Å is :-

(1) 9 : 4 (2) 3 : 2

(3) 16 : 81 (4) 8 : 27

Ans. (2)

Sol. Resolving power $\propto \frac{1}{\lambda}$

$$\frac{RP_1}{RP_2} = \frac{\lambda_2}{\lambda_1} = \frac{6000Å}{4000Å} = \frac{3}{2}$$

3. The two nearest harmonics of a tube closed at one end and open at other end are 220 Hz and 260 Hz. What is the fundamental frequency of the system?

(1) 20 Hz (2) 30 Hz

(3) 40 Hz (4) 10 Hz

Ans. (1)

Sol. Difference between any two consecutive frequencies of COP $= \frac{2v}{4\ell} = 260 - 220 = 40$ Hz

$$\Rightarrow \frac{v}{4\ell} = 20\,\text{Hz}$$

So fundamental frequency = 20 Hz

4. Consider a drop of rain water having mass 1 g falling from a height of 1 km. It hits the ground with a speed of 50 m/s. Take 'g' constant with a value 10 m/s^2. The work done by the (i) gravitational force and the (ii) resistive force of air is :-

(1) (i) 1.25 J (ii) – 8.25 J

(2) (i) 100 J (ii) 8.75 J

(3) (i) 10 J (ii) – 8.75 J

(4) (i) – 10 J (ii) – 8.25 J

Ans. (3)

Sol. Work done by the gravity (W_g) = mgh

$= 10^{-3} \times 10 \times 10^3 = 10$ J

By work–energy theorem = $W_g + W_{res} = \Delta KE$

$$10 + W_{res} = \frac{1}{2} \times 10^{-3} \times (50)^2$$

$$W_{res} = -8.75 \text{ J}$$

5. A physical quantity of the dimensions of length that can be formed out of c, G and $\frac{e^2}{4\pi\varepsilon_0}$ is [c is velocity of light, G is universal constant of gravitation and e is charge] :-

(1) $c^2\left[G\frac{e^2}{4\pi\varepsilon_0}\right]^{1/2}$ (2) $\frac{1}{c^2}\left[\frac{e^2}{G4\pi\varepsilon_0}\right]^{1/2}$

(3) $\frac{1}{c}G\frac{e^2}{4\pi\varepsilon_0}$ (4) $\frac{1}{c^2}\left[G\frac{e^2}{4\pi\varepsilon_0}\right]^{1/2}$

Ans. (4)

Sol. $[L] = [c]^a [G]^b \left[\frac{e^2}{4\pi\varepsilon_0}\right]^c$

$[L] = [LT^{-1}]^a [M^{-1}L^3T^{-2}]^b \left[ML^3T^{-2}\right]^c$

$[L] = L^{a+3b+3c} M^{-b+c} T^{-a-2b-2c}$

$a + 3b + 3c = 1$

$-b + c = 0$

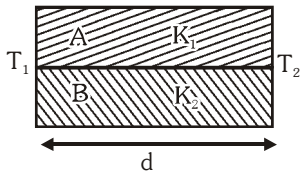
$a + 2b + 2c = 0$

On solving,

$$a = -2, \quad b = \frac{1}{2}, \quad c = \frac{1}{2}$$

$$\therefore \quad L = \frac{1}{c^2}\left[G.\frac{e^2}{4\pi\varepsilon_0}\right]^{\frac{1}{2}}$$

6. Two rods A and B of different materials are welded together as shown in figure. Their thermal conductivities are K_1 and K_2. The thermal conductivity of the composite rod will be :-



(1) $\frac{3(K_1 + K_2)}{2}$

(2) $K_1 + K_2$

(3) $2(K_1 + K_2)$

(4) $\frac{K_1 + K_2}{2}$

Ans. (4)

Sol. In parallel $\frac{1}{R_{eq}} = \frac{1}{R_1} + \frac{1}{R_2}$

$$\frac{K_{eq}(2A)}{\ell} = \frac{K_1 A}{\ell} + \frac{K_2 A}{\ell}$$

$$K_{eq} = \frac{K_1 + K_2}{2}$$

7. A capacitor is charged by a battery. The battery is removed and another identical uncharged capacitor is connected in parallel. The total electrostatic energy of resulting system :-

(1) Decreases by a factor of 2

(2) Remains the same

(3) Increases by a factor of 2

(4) Increases by a factor of 4

Ans. (1)

Sol. $U_i = \frac{1}{2}CV^2$

$$U_f = \frac{1}{2}[2C]\left[\frac{V}{2}\right]^2 = \frac{1}{2}U_i$$

Decrease by a factor of 2

8. In a common emitter transistor amplifier the audio signal voltage across the collector is 3V. The resistance of collector is 3 kΩ. If current gain is 100 and the base resistance is 2 kΩ, the voltage and power gain of the amplifier is :-

(1) 15 and 200

(2) 150 and 15000

(3) 20 and 2000

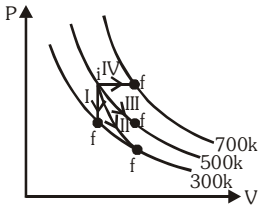
(4) 200 and 1000

Ans. (2)

Sol. $A_V = \beta\frac{R_C}{R_B} = 100 \times \frac{3k\Omega}{2k\Omega} = 150$

Power gain = $\beta A_V = 100 \times 150 = 15000$

9. Thermodynamic processes are indicated in the following diagram :



Match the following

	Column-1		**Column-2**
P.	Process I	a.	Adiabatic
Q.	Process II	b.	Isobaric
R.	Process III	c.	Isochoric
S.	Process IV	d.	Isothermal

(1) P → c, Q → a, R → d, S → b

(2) P → c, Q → d, R → b, S → a

(3) P → d, Q → b, R → a, S → c

(4) P → a, Q → c, R → d, S → b

Ans. (1)

Sol. Process (1) → volume constant → Isochroic

Process (2) → adiabatic

Process (3) → Temperature constant → Isothermal

Process (4) → Pressure constant → Isobaric

10. Suppose the charge of a proton and an electron differ slightly. One of them is – e, the other is (e + Δe). If the net of electrostatic force and gravitational force between two hydrogen atoms placed at a distance d (much greater than atomic size) apart is zero, then Δe is of the order of [Given mass of hydrogen $m_h = 1.67 \times 10^{-27}$ kg]

(1) 10^{-23} C (2) 10^{-37} C

(3) 10^{-47} C (4) 10^{-20} C

Ans. (2)

Sol. $$\frac{K \times (\Delta e)^2}{r^2} = \frac{Gm^2}{r^2}$$

$$\Delta e = m\sqrt{\frac{G}{K}} = 1.67 \times 10^{-27}\sqrt{\frac{6.67 \times 10^{-11}}{9 \times 10^9}}\ C$$

$$= 1.436 \times 10^{-37}\ C$$

11. The resistance of a wire is 'R' ohm. If it is melted and stretched to 'n' times its original length, its new resistance will be :-

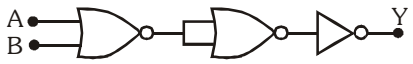
(1) $\frac{R}{n}$ (2) n^2R (3) $\frac{R}{n^2}$ (4) nR

Ans. (2)

Sol. $$R = \frac{\rho\ell}{A} = \frac{\rho\ell^2}{\text{volume}} \Rightarrow R \propto \ell^2$$

$$\Rightarrow R_2 = n^2R_1$$

12. The given electrical network is equivalent to :

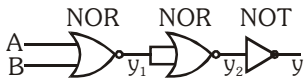


(1) OR gate (2) NOR gate

(3) NOT gate (4) AND gate

Ans. (2)

Sol.



$$y_1 = \overline{A + B}$$

$$y_2 = \overline{y_1 + y_1} = \overline{y_1} = \overline{\overline{A + B}} = A + B$$

$$y = \overline{y_2} = \overline{A + B}$$

NOR GATE

13. The de-Broglie wavelength of a neutron in thermal equilibrium with heavy water at a temperature T (Kelvin) and mass m, is :-

(1) $\frac{h}{\sqrt{3mkT}}$ (2) $\frac{2h}{\sqrt{3mkT}}$

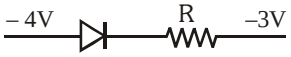
(3) $\frac{2h}{\sqrt{mkT}}$ (4) $\frac{h}{\sqrt{mkT}}$

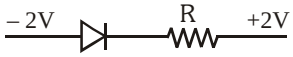
Ans. (1)

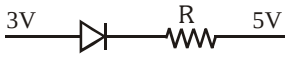
Sol. Kinetic energy of thermal neutron with equilibrium is $\frac{3}{2}KT$

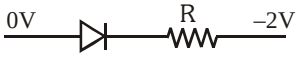
$$\lambda = \frac{h}{mv} = \frac{h}{\sqrt{2m\,K.E}} = \frac{h}{\sqrt{2m\left(\frac{3}{2}KT\right)}} = \frac{h}{\sqrt{3\,mKT}}$$

14. Which one of the following represents forward bias diode ?

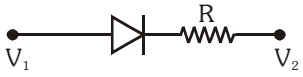
(1) 

(2) 

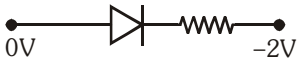
(3) 

(4) 

Ans. (4)

Sol. 

In forward bias $V_1 > V_2$

$\Rightarrow$ only 

is in forward bias

15. A long solenoid of diameter 0.1 m has 2×10^4 turns per meter. At the centre of the solenoid, a coil of 100 turns and radius 0.01 m is placed with its axis coinciding with the solenoid axis. The current in the solenoid reduces at a constant rate to 0A from 4 A in 0.05 s. If the resistance of the coil is $10\pi^2\Omega$. the total charge flowing through the coil during this time is :-

(1) 16 μC (2) 32 μC

(3) 16 π μC (4) 32 π μC

Ans. (2)

Sol. $$q = \left[\left(\frac{\Delta\phi}{\Delta t}\right)\cdot\frac{1}{R}\right]\Delta t$$

$$q = \left[\mu_0 nN\pi r^2\frac{\Delta i}{\Delta t}\right]\frac{1}{R}\Delta t$$

$$q = \left[4\pi \times 10^{-7} \times 2 \times 10^4 \times 100 \times \pi \times (10^{-2})^2 \times \left(\frac{4}{.05}\right)\right]\frac{1}{10\pi^2} \times 0.05$$

$$q = 32\ \mu C$$

16. Preeti reached the metro station and found that the escalator was not working. She walked up the stationary escalator in time t_1. On other days, if she remains stationary on the moving escalator, then the escalator takes her up in time t_2. The time taken by her to walk up on the moving escalator will be

(1) $\frac{t_1t_2}{t_2-t_1}$ (2) $\frac{t_1t_2}{t_2+t_1}$

(3) $t_1 - t_2$ (4) $\frac{t_1+t_2}{2}$

Ans. (2)

Sol. $V_1 \rightarrow$ velocity of Preeti
$V_2 \rightarrow$ velocity of escalator
$\ell \rightarrow$ distance

$$t = \frac{\ell}{V_1+V_2} = \frac{\ell}{\frac{\ell}{t_1}+\frac{\ell}{t_2}} = \frac{t_1t_2}{t_1+t_2}$$

17. Young's double slit experment is first performed in air and then in a medium other than air. It is found that 8th bright fringe in the medium lies where 5th dark fringe lies in air. The refractive index of the medium is nearly :-

(1) 1.59 (2) 1.69 (3) 1.78 (4) 1.25

Ans. (3)

Sol. $(y_8)_{\text{Bright, medium}} = (y_5)_{\text{Dark, air}}$

$$\frac{8\lambda_m D}{d} = \left(\frac{2(5)-1}{2}\right)\frac{\lambda D}{d}$$

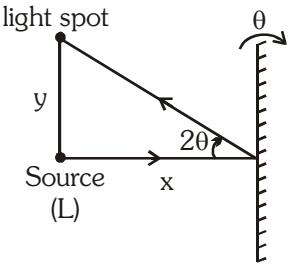
$$\frac{8\lambda}{\mu}\frac{D}{d} = \frac{9}{2}\frac{\lambda D}{d} \Rightarrow \mu = \frac{16}{9} = 1.78$$

18. A beam of light from a source L is incident normally on a plane mirror fixed at a certain distance x from the source. The beam is reflected back as a spot on a scale placed just above the source I. When the mirror is rotated through a small angle θ, the spot of the light is found to move through a distance y on the scale. The angle θ is given by :-

(1) $\frac{y}{x}$ (2) $\frac{x}{2y}$ (3) $\frac{x}{y}$ (4) $\frac{y}{2x}$

Ans. (4)

Sol.



$$2\theta = \frac{y}{x} \quad ; \quad \theta = \frac{y}{2x}$$

19. If θ_1 and θ_2 be the apparent angles of dip observed in two vertical planes at right angles to each other, then the true angle of dip θ is given by :-

(1) $\tan^2\theta = \tan^2\theta_1 + \tan^2\theta_2$

(2) $\cot^2\theta = \cot^2\theta_1 - \cot^2\theta_2$

(3) $\tan^2\theta = \tan^2\theta_1 - \tan^2\theta_2$

(4) $\cot^2\theta = \cot^2\theta_1 + \cot^2\theta_2$

Ans. (4)

Sol. $\tan\theta_1 = \frac{\tan\theta}{\cos\alpha}$

$$\Rightarrow \tan\theta_2 = \frac{\tan\theta}{\cos(90-\alpha)} = \frac{\tan\theta}{\sin\alpha}$$

$$\Rightarrow \sin^2\alpha + \cos^2\alpha = 1$$

$$\Rightarrow \cot^2\theta_2 + \cot^2\theta_1 = \cot^2\theta$$

20. Tow cars moving in opposite directions approach each other with speed of 22 m/s and 16.5 m/s respectively. The driver of the first car blows a horn having a a frequency 400 Hz. The frequency heard by the driver of the second car is [velocity of sound 340 m/s] :-

(1) 361 Hz (2) 411 Hz

(3) 448 Hz (4) 350 Hz

Ans. (3)

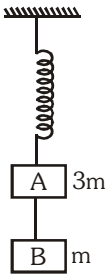
Sol. A (v_s = 22 m/s, f_0 = 400 Hz) → ← B (v_o = 16.5 m/s)

As we know for given condition

$$f_{app} = f_0\left(\frac{v + v_{observer}}{v - v_{source}}\right) = 400\left(\frac{340+16.5}{340-22}\right)$$

$$f_{app} = 448 \text{ Hz}$$

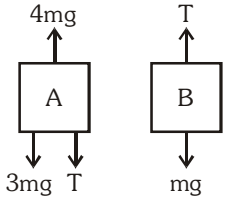
21. Two blocks A and B of masses 3 m and m respectively are connected by a massless and inextensible string. The whole system is suspended by a massless spring as shown in figure. The magnitudes of acceleration of A and B immediately after the string is cut, are respectively :-



(1) $\frac{g}{3}, g$ (2) g,g (3) $\frac{g}{3}, \frac{g}{3}$ (4) g $\frac{g}{3}$

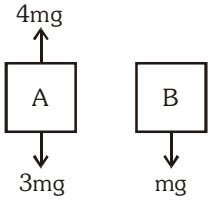
Ans. (1)

Sol. Before cutting the strip :-



$\therefore$ T = mg

After cutting the strip :-



$$a_A = \frac{4mg - 3mg}{3m} = \frac{g}{3}$$

$$a_B = \frac{mg}{m} = g$$

22. A thin prism having refracting angle 10° is made of glass of refractive index 1.42. This prism is combined with another thin prism of glass of refractive index 1.7. This combination produces dispersion without deviation. The refracting angle of second prism should be :-

(1) 6° (2) 8° (3) 10° (4) 4°

Ans. (1)

Sol. For dispersion without deviation

$\delta_1 = \delta_2$

$A_1(\mu_1 - 1) = A_2(\mu_2 - 1)$

$10(1.42 - 1) = A_2(1.7 - 1)$

$A_2 = 6°$

23. The acceleration due to gravity at a height 1 km above the earth is the same as at a depth d below the surface of earth. Then :-

(1) d = 1 km

(2) d = $\frac{3}{2}$km

(3) d = 2 km

(4) d = $\frac{1}{2}$ km

Ans. (3)

Sol. $\because$ $g_h = g_d$

$$g\left(1 - \frac{2h}{R}\right) = g\left(1 - \frac{d}{R}\right)$$

d = 2h = 2 km

24. A potentiometer is an accurate and versatile device to make electrical measurements of E.M.F. because the method involves :-

(1) Potential gradients

(2) A condition of no current flow through the galvanometer

(3) A combination of cells, galvanometer and resistances

(4) Cells

Ans. (2)

Sol. In zero deflection condition, potentiometer draws no current.

25. A spherical black body with a radius of 12 cm radiates 450 watt power at 500 K. If the radius were halved and the temperature doubled, the power radiated in watt would be :-

(1) 450 (2) 1000

(3) 1800 (4) 225

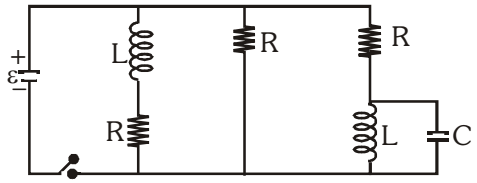
Ans. (3)

Sol. $P \propto r^2 T^4$

$$\Rightarrow \frac{P_1}{P_2} = \left(\frac{r_1}{r_2}\right)^2 \left(\frac{T_1}{T_2}\right)^4$$

P_2 = 1800 watt

26. Figure shows a circuit that contains three identical resistors with resistance R = 9.0 Ω each, two identical inductors with inductance L = 2.0 mH each, and an ideal battery with emf ε = 18 V. The current 'i' through the battery just after the switch closed is,..... :-

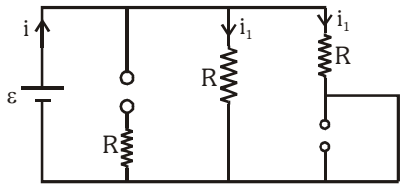


(1) 0.2 A (2) 2 A

(3) 0 ampere (4) 2 mA

Ans. (Bonus)

Sol. at t = 0



$$i_1 = \frac{\varepsilon}{R} = \frac{18}{9} = 2A$$

∴ Current through the battery is

$i = 2i_1 = 2 \times 2 = 4A$ (Bonus)

OR

According to question language :

Capacitor is not mentioned so i = 2 A

27. Radioactive material 'A' has decay constant '8 λ' and material 'B' has decay constant 'λ'. Initially they have same number of nuclei. After what time, the ratio of number of nuclei of material 'B' to that 'A' will be

$\frac{1}{e}$?

(1) $\frac{1}{7\lambda}$ (2) $\frac{1}{8\lambda}$ (3) $\frac{1}{9\lambda}$ (4) $\frac{1}{\lambda}$

Ans. (1)

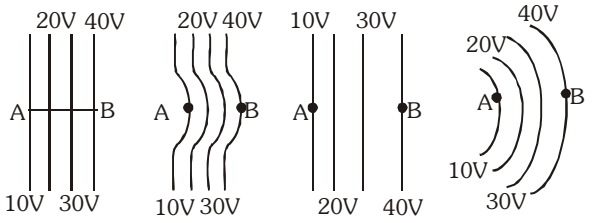
Sol. $\lambda_A = 8\lambda$, $\lambda_B = \lambda$

$$\Rightarrow N_B = \frac{N_A}{e} \Rightarrow N_0 e^{-\lambda t} = \frac{N_0 e^{-8\lambda t}}{e}$$

$$\Rightarrow -\lambda t = -8\lambda t - 1 \Rightarrow 7\lambda t = -1 \Rightarrow t = -\frac{1}{7\lambda}$$

Best answer is $t = \frac{1}{7\lambda}$

28. The diagrams below show regions of equipotentials:-



(a) (b) (c) (d)

A positive charge is moved from A to B in each diagram.

(1) In all the four cases the work done is the same

(2) Minimum work is required to move q in figure (a)

(3) Maximum work is required to move q in figure (b)

(4) Maximum work is required to move q in figure (c)

Ans. (1)

Sol. $W = q\Delta V$

as ΔV is same in all conditions, work will be same.

29. Two astronauts are floating in gravitational free space after having lost contact with their spaceship. The two will :-

(1) Move towards each other.

(2) Move away from each other.

(3) Will become stationary

(4) Keep floating at the same distance between them.

Ans. (1)

Sol. Astronauts move towards each other under mutual gravitational force.

30. The x and y coordinates of the particle at any time are $x = 5t - 2t^2$ and $y = 10t$ respectively, where x and y are in meters and t in seconds. The acceleration of the particle at t = 2s is :-

(1) 5 m/s^2 (2) – 4 m/s^2

(3) – 8 m/s^2 (4) 0

Ans. (2)

Sol. $v_x = 5 - 4t$, $v_y = 10$

$a_x = -4$, $a_y = 0$

$\vec{a} = a_x\hat{i} + a_y\hat{j}$

$\vec{a} = -4\hat{i}$ m/s^2

31. One end of string of length l is connected to a particle of mass 'm' and the other end is connected to a small peg on a smooth horizontal table. If the particle moves in circle with speed 'v' the net force on the particle (directed towards centre) will be (T represents the tension in the string) :-

(1) $T+\frac{mv^2}{l}$ (2) $T-\frac{mv^2}{l}$

(3) Zero (4) T

Ans. (4)

Sol. Net force on the particle in uniform circular motion is centripetal force, which is provided by the tension in string.

32. A particle executes linear simple harmonic motion with an amplitude of 3 cm. When the particle is at 2 cm from the mean position, the magnitude of its velocity is equal to that of its acceleration. Then its time period in seconds is :-

(1) $\frac{\sqrt{5}}{2\pi}$ (2) $\frac{4\pi}{\sqrt{5}}$ (3) $\frac{2\pi}{\sqrt{3}}$ (4) $\frac{\sqrt{5}}{\pi}$

Ans. (2)

Sol. Amplitude A = 3 cm

When particle is at x = 2 cm ,

its | velocity | = | acceleration |

$$\text{i.e., } \omega\sqrt{A^2-x^2} = \omega^2 x \Rightarrow \omega = \frac{\sqrt{A^2-x^2}}{x}$$

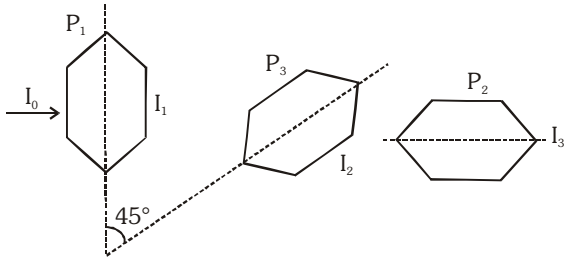
$$T = \frac{2\pi}{\omega} = 2\pi\left(\frac{2}{\sqrt{5}}\right) = \frac{4\pi}{\sqrt{5}}$$

33. Two Polaroids P_1 and P_2 are placed with their axis perpendicular to each other. Unpolarised light I_0 is incident on P_1 . A third polaroid P_3 is kept in between P_1 and P_2 such that its axis makes an angle 45° with that of P_1 . The intensity of transmitted light through P_2 is :-

(1) $\frac{I_0}{4}$ (2) $\frac{I_0}{8}$ (3) $\frac{I_0}{16}$ (4) $\frac{I_0}{2}$

Ans. (2)

Sol.



$$I_1 = \frac{I_0}{2}$$

$$I_2 = \frac{I_0}{2}\cos^2 45° = \frac{I_0}{4}$$

$$I_3 = \frac{I_0}{4}\cos^2 45° = \frac{I_0}{8}$$

34. The bulk modulus of a spherical object is 'B'. If it is subjected to uniform pressure 'p', the fractional decrease in radius is :-

(1) $\frac{B}{3p}$ (2) $\frac{3p}{B}$ (3) $\frac{p}{3B}$ (4) $\frac{p}{B}$

Ans. (3)

Sol. $B = \frac{\Delta P}{-\frac{\Delta V}{V}}$, $\frac{\Delta V}{V} = \frac{3\Delta R}{R}$

$$B = \frac{\Delta P}{\frac{-3\Delta R}{R}} \Rightarrow -\frac{\Delta R}{R} = \frac{P}{3B} \quad (\Delta P = P)$$

35. In an electromagnetic wave in free space the root mean square value of the electric field is E_{rms} = 6V/m. The peak value of the magnetic field is :-

(1) 2.83×10^{-8} T (2) 0.70×10^{-8} T

(3) 4.23×10^{-8} T (4) 1.41×10^{-8} T

Ans. (1)

Sol. $E_0 = CB_0$

$$E_{rms} = \frac{E_0}{\sqrt{2}}$$

$$\Rightarrow E_{rms}\sqrt{2} = CB_0$$

$$\Rightarrow B_0 = \frac{E_{rms}\sqrt{2}}{C} = \frac{6\times\sqrt{2}}{3\times10^8} = 2.83 \times 10^{-8} \text{ T}$$

36. A rope is wound around a hollow cylinder of mass 3 kg and radius 40 cm. What is the angular acceleration of the cylinder if the rope is pulled with a force of 30 N ?

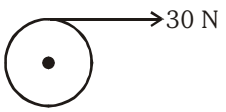
(1) 0.25 rad/s^2 (2) 25 rad/s^2

(3) 5 m/s^2 (4) 25 m/s^2

Ans. (2)

Sol. $\tau = I\alpha$

$RF = mR^2\alpha$



$$\alpha = \frac{F}{mR} = \frac{30}{3\times\frac{40}{100}} = 25 \text{ rad/s}^2$$

37. Two discs of same moment of inertia rotating about their regular axis passing through centre and perpendicular to the plane of disc with angular velocities ω_1 and ω_2 . They are brought into contact face to face coinciding the axis of rotation. The expression for loss of energy during this process is:-

(1) $\frac{1}{4}I(\omega_1-\omega_2)^2$ (2) $I(\omega_1-\omega_2)^2$

(3) $\frac{I}{8}(\omega_1-\omega_2)^2$ (4) $\frac{1}{2}I(\omega_1+\omega_2)^2$

Ans. (1)

Sol. COAM : $I\omega_1 + I\omega_2 = 2I\omega \Rightarrow \omega = \frac{\omega_1+\omega_2}{2}$

$(K.E.)_i = \frac{1}{2}I\omega_1^2 + \frac{1}{2}I\omega_2^2$

$(K.E.)_f = \frac{1}{2}\times 2I\omega^2 = I\left(\frac{\omega_1+\omega_2}{2}\right)^2$

Loss in K.E. = $(K.E.)_i - (K.E)_f = \frac{I}{4}(\omega_1-\omega_2)^2$

38. The photoelectric threshold wavelength of silver is 3250×10^{-10}m. The velocity of the electron ejected from a silver surface by ultraviolet light of wavelength 2536×10^{-10} m is :-

(Given h = 4.14×10^{-15} eVs and c = 3×10^8 ms^{-1})

(1) $\approx 0.6\times10^6$ ms^{-1} (2) $\approx 61\times10^3$ ms^{-1}

(3) $\approx 0.3\times10^6$ ms^{-1} (4) $\approx 6\times10^5$ ms^{-1}

Ans. (1 or 4)

Sol. λ_0 = 3250 Å

λ = 2536 Å

$\frac{1}{2}mv^2 = hc\left[\frac{1}{\lambda}-\frac{1}{\lambda_0}\right]$

$v = \sqrt{\frac{2hc}{m}\left[\frac{1}{\lambda}-\frac{1}{\lambda_0}\right]}$

$= \sqrt{\frac{2\times12400\times1.6\times10^{-19}}{9.1\times10^{-31}}\left[\frac{714}{2536\times3250}\right]}$

$= 0.6\times10^6$ m/s = 6×10^5 m/s

39. A 250-Turn rectangular coil of length 2.1 cm and width 1.25 cm carries a current of 85 µA and subjected to magnetic field of strength 0.85 T. Work done for rotating the coil by 180º against the torque is:-

(1) 4.55 µJ (2) 2.3 µJ

(3) 1.15 µ J (4) 9.1 µ J

Ans. (4)

Sol. Work = MB[cos θ_1 – cos θ_2]

Work = MB[cos 0 – cos 180°]

W = NiAB[1 – (–1)]

W ≃ 9.1 µJ

40. The ratio of wavelengths of the last line of Balmer series and the last line of Lyman series is :-

(1) 1 (2) 4 (3) 0.5 (4) 2

Ans. (2)

Sol. For last line of Balmer : n_1 = 2 & $n_2 = \infty$

$\frac{1}{\lambda_B} = RZ^2\left[\frac{1}{n_1^2}-\frac{1}{n_2^2}\right] = R(1)^2\left[\frac{1}{2^2}-\frac{1}{\infty^2}\right]$

$\lambda_B = \frac{4}{R}$ ...(1)

For last line of Lyman series : n_1 = 1 & $n_2 = \infty$

$\frac{1}{\lambda_L} = RZ^2\left[\frac{1}{n_1^2}-\frac{1}{n_2^2}\right] = R(1)^2\left[\frac{1}{1^2}-\frac{1}{\infty^2}\right]$

$\lambda_L = 1/R$ ...(2)

$\boxed{\frac{\lambda_B}{\lambda_L} = \frac{(4/R)}{(1/R)} = 4}$

41. A carnot engine having an efficiency of $\frac{1}{10}$ as heat engine, is used as a refrigerator. If the work done on the system is 10 J, the amount of energy absorbed from the reservoir at lower temperature is :-

(1) 90 J (2) 99 J

(3) 100 J (4) 1 J

Ans. (1)

Sol. $\beta = \frac{Q_2}{W} = \frac{1-\eta}{\eta}$

$\Rightarrow \frac{Q_2}{9} = \frac{1-0.1}{0.1}$

$\Rightarrow Q_2 = 9\times10 = 90$ J

42. A gas mixture consists of 2 moles of O_2 and 4 moles of Ar at temperature T. Neglecting all vibrational modes, the total internal energy of the system is :-

(1) 15 RT (2) 9 RT (3) 11 RT (4) 4 RT

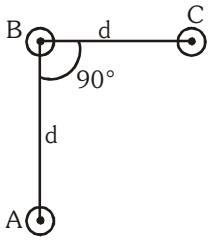
Ans. (3)

Sol. $U = \frac{f}{2}nRT$

$U_{total} = \frac{5}{2}(2)RT + \frac{3}{2}(4)RT$

$U_{total} = 11RT$

43. An arrangement of three parallel straight wires placed perpendicular to plane of paper carrying same current 'I along the same direction is shown in fig. Magnitude of force per unit length on the middle wire 'B' is given by :-



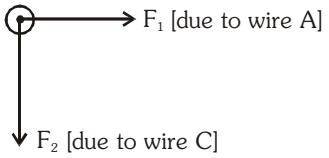
(1) $\frac{2\mu_0 i^2}{\pi d}$ (2) $\frac{\sqrt{2}\mu_0 i^2}{\pi d}$

(3) $\frac{\mu_0 i^2}{\sqrt{2}\pi d}$ (4) $\frac{\mu_0 i^2}{2\pi d}$

Ans. (3)

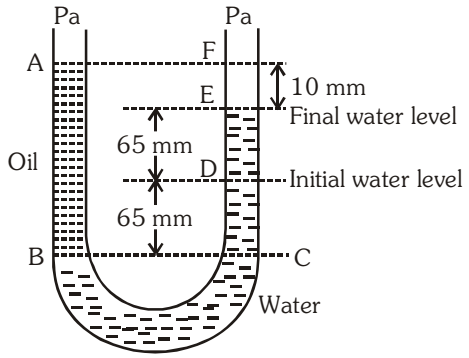
Sol. $F = \frac{\mu_0 i_1 i_2}{2\pi d}$ = force per unit length

$F_1 = \frac{(\mu_0 i)i}{2\pi d} = \frac{\mu_0 i^2}{2\pi d} = F_2$



$F_{net} = \sqrt{F_1^2 + F_2^2} = \frac{\mu_0 i^2}{\sqrt{2}\pi d}$

44. A U tube with both ends open to the atmosphere, is partially filled with water. Oil, which is immiscible with water, is poured into one side until it stands at a distance of 10 mm above the water level on the other side. Meanwhile the water rises by 65 mm from its original level (see diagram). The density of the oil is :-



(1) 425 kg m^{-3} (2) 800 kg m^{-3}

(3) 928 kg m^{-3} (4) 650 kg m^{-3}

Ans. (3)

Sol. $\rho_0 g \times 140 \times 10^{-3} = \rho_w g \times 130 \times 10^{-3}$

$\rho_0 = \frac{130}{140} \times 10^3 \approx 928 \text{ kg/m}^3$

45. Which of the following statements are **correct** ?

(a) Centre of mass of a body always coincides with the centre of gravity of the body

(b) Central of mass of a body is the point at which the total gravitational torque on the body is zero

(c) A couple on a body produce both translational and rotation motion in a body

(d) Mechanical advantage greater than one means that small effort can be used to lift a large load

(1) (a) and (b) (2) (b) and (c)

(3) (c) and (d) (4) (b) and (d)

Ans. (4)

Sol. Centre of mass may lie on centre of gravity net torque of gravitational pull is zero about centre of mass.

Mechanical advantage = $\frac{\text{Load}}{\text{Effort}} > 1$

$\Rightarrow$ Load > Effort

NEET(UG)–2017 TEST PAPER WITH ANSWER & SOLUTIONS (HELD ON SUNDAY 07th MAY, 2017)

46. Which one of the following statements is **correct**, with reference to enzymes ?
(1) Holoenzyme = Apoenzyme + Coenzyme
(2) Coenzyme = Apoenzyme + Holoenzyme
(3) Holoenzyme = Coenzyme + Co-factor
(4) Apoenzyme = Holoenzyme + Coenzyme
Ans. (1)

47. A decrease in blood pressure / volume will not cause the release of :
(1) Atrial natriuretic factor (2) Aldosterone
(3) ADH (4) Renin
Ans. (1)

48. Which cells of "Crypts of Lieberkuhn" secrete antibacterial lysozyme ?
(1) Paneth cells (2) Zymogen cells
(3) Kupffer cells (4) Argentaffin cells
Ans. (1)

49. Which of the following are not polymeric ?
(1) Proteins (2) Polysaccharides
(3) Lipids (4) Nucleic acids
Ans. (3)

50. Functional megaspore in an angiosperm develops into ?
(1) Endosperm (2) Embryo sac
(3) Embryo (4) Ovule
Ans. (2)

51. Myelin sheath is produced by :
(1) Astrocytes and Schwann cells
(2) Oligodendrocytes and Osteoclasts
(3) Osteoclasts and Astrocytes
(4) Schwann cells and Oligodendrocytes
Ans. (4)

52. Attractants and rewards are required for :
(1) Entomophily (2) Hydrophily
(3) Cleistogamy (4) Anemophily
Ans. (1)

53. Receptor sites for neurotransmitters are present on :
(1) Pre-synaptic membrane
(2) Tips of axons
(3) Post-synaptic membrane
(4) Membrane of synaptic vesicles
Ans. (3)

54. Coconut fruit is a :
(1) Berry (2) Nut
(3) Capsule (4) Drupe
Ans. (4)

55. Adult human RBCs are enucleated. Which of the following statement(s) is/are **most appropriate** explanation for this feature ?
(a) They do not need to reproduce
(b) They are somatic cells
(c) They do not metabolize
(d) All their internal space is available for oxygen transport
(1) only (a) (2) (a), (c) and (d)
(3) (b) and (c) (4) only (d)
Ans. (4)

56. Capacitation occurs in :
(1) Epididymis
(2) Vas deferens
(3) Female reproductive tract
(4) Rete testis
Ans. (3)

57. Which of the following are found in extreme saline conditions ?
(1) Eubacteria (2) Cyanobacteria
(3) Mycobacteria (4) Archaebacteria
Ans. (4)

58. Asymptote in a logistic growth curve is obtained when :
(1) K = N
(2) K > N
(3) K < N
(4) The value of 'r' approaches zero
Ans. (1)

59. Artificial selection to obtain cows yielding higher milk output represents :
(1) Directional as it pushes the mean of the character in one direction
(2) Disruptive as it splits the population into two, one yielding higher output and the other lower output
(3) Stabilizing followed by disruptive as it stabilizes the population to produce higher yielding cows
(4) Stabilizing selection as it stabilizes this character in the population
Ans. (1)

60. Select the mismatch :
(1) *Rhodospirillum* - Mycorrhiza
(2) *Anabaena* - Nitrogen fixer
(3) *Rhizobium* - Alfalfa
(4) *Frankia* - *Alnus*
Ans. (1)

61. Good vision depends on adequate intake of carotene rich food :
Select the best option from the following statements :
(a) Vitamin A derivatives are formed from carotene
(b) The photopigments are embedded in the membrane discs of the inner segment
(c) Retinal is a derivative of Vitamin A
(d) Retinal is a light absorbing part of all the visual photopigments
Options :
(1) (a), (c) and (d) (2) (a) and (c)
(3) (b), (c) and (d) (4) (a) and (b)

Ans. (1)

62. The DNA fragments separated on an agarose gel can be visualised after staining with :
(1) Acetocarmine (2) Aniline blue
(3) Ethidium bromide (4) Bromophenol blue

Ans. (3)

63. The hepatic portal vein drains blood to liver from :
(1) Stomach (2) Kidneys
(3) Intestine (4) Heart

Ans. (3)

64. The vascular cambium normally gives rise to :
(1) Primary phloem (2) Secondary xylem
(3) Periderm (4) Phelloderm

Ans. (2)

65. Thalassemia and sickle cell anemia are caused due to a problem in globin molecule synthesis. Select the correct statement :
(1) Both are due to a quantitative defect in globin chain synthesis
(2) Thalassemia is due to less synthesis of globin molecules
(3) Sickel cell anemia is due to a quantitative problem of globin molecules
(4) Both are due to a qualitative defect in globin chain synthesis

Ans. (2)

66. The genotypes of a husband and Wife are I^AI^B and $I^A i$.
Among the blood types of their children, how many different genotypes and phenotypes are possible?
(1) 3 genotypes ; 4 phenotypes
(2) 4 genotypes ; 3 phenotypes
(3) 4 genotypes ; 4 phenotypes
(4) 3 genotypes ; 3 phenotypes

Ans. (2)

67. Which of the following facilitates opening of stomatal aperture ?
(1) Decrease in turgidity of guard cells
(2) Radial orientation of cellulose microfibrils in the cell wall of guard cells
(3) Longitudinal orientation of cellulose microfibrils in the cell wall of guard cells
(4) Contraction of outer wall of guard cells

Ans. (2)

68. In Bougainvillea thorns are the modifications of :
(1) Adventitious root (2) Stem
(3) Leaf (4) Stipules

Ans. (2)

69. Which one of the following is related to Ex-situ conservation of threatened animals and plants ?
(1) Biodiversity hot spots
(2) Amazon rainforest
(3) Himalayan region
(4) Wildlife safari parks

Ans. (4)

70. Root hairs develop from the region of :
(1) Elongation (2) root cap
(3) Meristematic activity (4) Maturation

Ans. (4)

71. A disease caused by an autosomal primary non-disjunction is :
(1) Klinefelter's Syndrome (2) Turner's Syndrome
(3) Sickel Cell Anemia (4) Down's Syndrome

Ans. (4)

72. The water potential of pure water is :
(1) Less than zero
(2) More than zero but less than one
(3) More than one
(4) Zero

Ans. (4)

73. Which of the following options gives the correct sequence of events during mitosis ?
(1) Condensation → nuclear membrane disassembly → arrangement at equator → centromere division → segregation → telophase
(2) Condensation → crossing over → nuclear membrane disassembly → segregation → telophase
(3) Condensation → arrangement at equator → centromere division → segregation → telophase
(4) Condensation → nuclear membrane disassembly → crossing over → segregation → telophase

Ans. (1)

74. The process of separation and purification of expressed protein before marketing is called :
(1) Downstream processing
(2) Bioprocessing
(3) Postproduction processing
(4) Upstream processing

Ans. (1)

75. A temporary endocrine gland in the human body is :
(1) Corpus cardiacum (2) corpus luteum
(3) Corpus allatum (4) Pineal gland

Ans. (2)

76. Which of the following is made up of dead cells?
(1) Collenchyma (2) Phellem
(3) Phloem (4) Xylem parenchyma

Ans. (2)

77. An example of colonial alga is :
(1) *Volvox* (2) *Ulothrix*
(3) *Spirogyra* (4) *Chlorella*

Ans. (1)

78. Match the following sexually transmitted diseases (Column-I) with their causative agent (Column-II) and select the correct option :

Column-I		Column-II	
(a)	Gonorrhea	(i)	HIV
(b)	Syphilis	(ii)	*Neisseria*
(c)	Genital Warts	(iii)	*Treponema*
(d)	AIDS	(iv)	Human papilloma-Virus

	(a)	(b)	(c)	(d)
(1)	iii	iv	i	ii
(2)	iv	ii	iii	i
(3)	iv	iii	ii	i
(4)	ii	iii	iv	i

Ans. (4)

79. The function of copper ions in copper releasing IUD's is :
(1) They inhibit gametogenesis
(2) They make uterus unsuitable for implantation
(3) They inhibt ovulation
(4) The suppress sperm motility and fertilising capacity of sperms

Ans. (4)

80. Which of the following in sewage treatment removes suspended solids ?
(1) Secondary treatment (2) Primary treatment
(3) Sludge treatment (4) Tertiary treatment

Ans. (2)

81. An important characteristic that Hemichordates share with Chordates is :
(1) Ventral tubular nerve cord
(2) Pharynx with gill slits
(3) Pharynx without gill slits
(4) Absence of notochord

Ans. (2)

82. The final proof for DNA as the genetic material came from the experiments of :
(1) Hershey and Chase
(2) Avery, Mcleod and McCarty
(3) Hargobind Khorana
(4) Griffith

Ans. (1)

83. Among the following characters, which one was not considered by Mendel in his experiments on pea ?
(1) Trichomes – Glandular or non-glandular
(2) Seed – Green or Yellow
(3) Pod – Inflated or Constricted
(4) Stem - Tall or Dwarf

Ans. (1)

84. Plants which produce characteristic pneumatophores and show vivipary belong to :
(1) Halophytes (2) Psammophytes
(3) Hydrophytes (4) Mesophytes

Ans. (1)

85. The pivot joint between atlas and axis is a type of :
(1) Cartilaginous joint (2) Synovial joint
(3) Saddle joint (4) Fibrous joint

Ans. (2)

86. With reference to factors affecting the rate of photosynthesis, which of the following statements is not correct ?
(1) Increasing atmospheric CO_2 concentration up to 0.05% can enhance CO_2 fixation rate
(2) C_3 plants respond to higher temperatures with enhanced photosynthesis while C_4 plants have much lower temperature optimum
(3) Tomato is a greenhouse crop which can be grown in CO_2 - enriched atmosphere for higher yield
(4) Light saturation for CO_2 fixation occurs at 10% of full sunlight

Ans. (2)

87. DNA fragments are:
(1) Negatively charged
(2) Neutral
(3) Either positively or negatively charged depending on their size
(4) Positively charged

Ans. (1)

88. Which of the following components provides sticky character to the bacterial cell ?
(1) Nuclear membrane
(2) Plasma membrane
(3) Glycocalyx
(4) Cell wall

Ans. (3)

89. Which of the following options best represents the enzyme composition of pancreatic juice ?
(1) amylase, pepsin, trypsinogen, maltase
(2) peptidase, amylase, pepsin, rennin
(3) lipase, amylase, trypsinogen, procarboxypeptidase
(4) amylase, peptidase, trypsinogen, rennin

Ans. (3)

90. Which among these is the correct combination of aquatic mammals ?
(1) Dolphins, Seals, *Trygon*
(2) Whales, Dolphins, Seals
(3) *Trygon*, Whales, Seals
(4) Seals, Dolphins, Sharks

Ans. (2)

91. Fruit and leaf drop at early stages can be prevented by the application of:
(1) Ethylene (2) Auxins
(3) Gibberellic acid (4) Cytokinins

Ans. (2)

92. Select the **correct** route for the passage of sperms in male frogs:
(1) Testes → Vasa efferentia → Kidney → Seminal Vesicle → Urinogenital duct → Cloaca
(2) Testes → Vasa efferentia → Bidder's canal → Ureter → Cloaca
(3) Testes → Vasa efferentia → Kidney → Bidder's canal → Urinogenital duct → Cloaca
(4) Testes → Bidder's canal → Kidney → Vasa efferentia → Urinogenital duct → Cloaca

Ans. (3)

93. In case of a couple where the male is having a very low sperm count, which technique will be suitable for fertilisation ?
(1) Gamete intracytoplasmic fallopian transfer
(2) Artificial Insemination
(3) Intracytoplasmic sperm injection
(4) Intrauterine transfer

Ans. (2)

94. Which ecosystem has the maximum biomass ?
(1) Grassland ecosystem
(2) Pond ecosystem
(3) Lake ecosystem
(4) Forest ecosystem

Ans. (4)

95. Lungs are made up of air-filled sacs, the alveoli. They do not collapse even after forceful expiration, because of:
(1) Inspiratory Reserve Volume
(2) Tidal Volume
(3) Expiratory Reserve Volume
(4) Residual Volume

Ans. (4)

96. Presence of plants arranged into well defined vertical layers depending on their height can be seen best in:
(1) Tropical Rain Forest (2) Grassland
(3) Temperate Forest (4) Tropical Savannah

Ans. (1)

97. Which of the following statements is **correct** ?
(1) The descending limb of loop of Henle is impermeable to water.
(2) The ascending limb of loop of Henle is permeable to water.
(3) The descending limb of loop of Henle is permeable to electrolytes.
(4) The ascending limb of loop of Henle is impermeable to water.

Ans. (4)

98. Alexander Von Humbolt described for the first time:
(1) Laws of limiting factor
(2) Species area relationships
(3) Population Growth equation
(4) Ecological Biodiversity

Ans. (2)

99. Zygotic meiosis is characteristic of;
(1) *Fucus* (2) *Funaria*
(3) *Chlamydomonas* (4) *Marchantia*

Ans. (3)

100. If there are 999 bases in an RNA that codes for a protein with 333 amino acids, and the base at position 901 is deleted such that the length of the RNA becomes 998 bases, how many codons will be altered ?
(1) 11 (2) 33 (3) 333 (4) 1

Ans. (2)

101. Flowers which have single ovule in the ovary and are packed into inflorescence are usually pollinated by:

(1) Bee (2) Wind

(3) Bat (4) Water

Ans. (2)

102. Transplantation of tissues/organs fails often due to non-acceptance by the patient's body. Which type of immune-response is responsible for such rejections ?

(1) Cell - mediated immune response

(2) Hormonal immune response

(3) Physiological immune response

(4) Autoimmune response

Ans. (1)

103. Life cycle of *Ectocarpus* and *Fucus* respectively are:

(1) Diplontic, Haplodiplontic

(2) Haplodiplontic, Diplontic

(3) Haplodiplontic, Haplontic

(4) Haplontic, Diplontic

Ans. (2)

104. A gene whose expression helps to identify transformed cell is known as :

(1) Vector (2) Plasmid

(3) Structural gene (4) Selectable marker

Ans. (4)

105. A dioecious flowering plant prevents both :

(1) Autogamy and geitonogamy

(2) Geitonogamy and xenogamy

(3) Cleistogamy and xenogamy

(4) Autogamy and xenogamy

Ans. (1)

106. Which statement is wrong for Krebs' cycle ?

(1) There is one point in the cycle where FAD^+ is reduced to $FADH_2$

(2) During conversion of succinyl CoA to succinic acid, a molecule of GTP is synthesised

(3) The cycle starts with condensation of acetyl group (acetyl CoA) with pyruvic acid to yield citric acid

(4) There are three points in the cycle where NAD^+ is reduced to $NADH + H^+$

Ans. (3)

107. Phosphoenol pyruvate (PEP) is the primary CO_2 acceptor in:

(1) C_4 plants

(2) C_2 plants

(3) C_3 and C_4 plants

(4) C_3 plants

Ans. (1)

108. During DNA replication, Okazaki fragments are used to elongate:

(1) The lagging strand towards replication fork.

(2) The leading strand away from replication fork.

(3) The lagging strand away from the replication fork.

(4) The leading strand towards replication fork.

Ans. (3)

109. Which of the following RNAs should be most abundant in animal cell ?

(1) t-RNA (2) m-RNA

(3) mi-RNA (4) r-RNA

Ans. (4)

110. GnRH, a hypothalamic hormone, needed in reproduction, acts on:

(1) anterior pituitary gland and stimulates secretion of LH and FSH.

(2) posterior pituitary gland and stimulates secretion of oxytocin and FSH.

(3) posterior pituitary gland and stimulates secretion of LH and relaxin.

(4) anterior pituitary gland and stimulates secretion of LH and oxytocin.

Ans. (1)

111. What is the criterion for DN A fragments movement on agarose gel during gel electrophoresis ?

(1) The smaller the fragment size, the farther it moves

(2) Positively charged fragments move to farther end

(3) Negatively charged fragments do not move

(4) The larger the fragment size, the farther it moves

Ans. (1)

112. Hypersecretion of Growth Hormone in adults does not cause further increase in height, because:
(1) Epiphyseal plates close after adolescence.
(2) Bones loose their sensitivity to Growth Hormone in adults.
(3) Muscle fibres do not grow in size after birth.
(4) Growth Hormone becomes inactive in adults.

Ans. (1)

113. DNA replication in bacteria occurs:
(1) Within nucleolus
(2) Prior to fission
(3) Just before transcription
(4) During S phase

Ans. (2)

114. Which one from those given below is the period for Mendel's hybridization experiments ?
(1) 1840 - 1850
(2) 1857 - 1869
(3) 1870 - 1877
(4) 1856 - 1863

Ans. (4)

115. Viroids differ from viruses in having;
(1) DNA molecules without protein coat
(2) RNA molecules with protein coat
(3) RNA molecules without protein coat
(4) DNA molecules with protein coat

Ans. (3)

116. MALT constitutes about _____ percent of the lymphoid tissue in human body.
(1) 20% (2) 70%
(3) 10% (4) 50%

Ans. (4)

117. Which of the following is correctly matched for the product produced by them ?
(1) *Methanobacterium* : Lactic acid
(2) *Penicillium notatum* : Acetic acid
(3) *Sacchromyces cerevisiae* : Ethanol
(4) *Acetobacter aceti* : Antibiotics

Ans. (3)

118. Which among the following are the smallest living cells, known without a definite cell wall, pathogenic to plants as well as animals and can survive without oxygen ?
(1) *Pseudomonas* (2) *Mycoplasma*
(3) *Nostoc* (4) *Bacillus*

Ans. (2)

119. Which of the following represents order of'Horse' ?
(1) Perissodactyla (2) Caballus
(3) Ferus (4) Equidae

Ans. (1)

120. Frog's heart when taken out of the body continues to beat for sometime.
Select the best option from the following statements.
(a) Frog is a poikilotherm.
(b) Frog does not have any coronary circulation.
(c) Heart is "myogenic" in nature.
(d) Heart is autoexcitable
Options:
(1) Only(d) (2) (a) and (b)
(3) (c)and(d) (4) Only(c)

Ans. (3)

121. Homozygous purelines in cattle can be obtained by:
(1) mating of unrelated individuals of same breed.
(2) mating of individuals of different breed.
(3) mating of individuals of different species.
(4) mating of related individuals of same breed.

Ans. (4)

122. Identify the wrong statement in context of heartwood:
(1) It is highly durable
(2) It conducts water and minerals efficiently
(3) It comprises dead elements with highly lignified walls
(4) Organic compounds are deposited in it

Ans. (2)

123. Anaphase Promoting Complex (APC) is a protein degradation machinery necessary for proper mitosis of animal cells. If APC is defective in a human cell, which of the following is expected to occur ?
(1) Chromosomes will be fragmented
(2) Chromosomes will not segregate
(3) Recombination of chromosome arms will occur
(4) Chromosomes will not condense

Ans. (2)

124. Which of the following cell organelles is responsible for extracting energy from carbohydrates to form ATP ?
(1) Ribosome (2) Chloroplast
(3) Mitochondrion (4) Lysosome

Ans. (3)

125. Mycorrhizae are the example of:
(1) Amensalism (2) Antibiosis
(3) Mutualism (4) Fungistasis

Ans. (3)

126. Out of 'X' pairs of ribs in humans only 'Y' pairs are true ribs. Select the option that correctly represents values of X and Y and provides their explanation:

(1) X = 12, Y = 5 — True ribs are attached dorsally to vertebral column and sternum on the two ends.

(2) X = 24, Y = 7 — True ribs are dorsally attached to vertebral column but are free on ventral side.

(3) X = 24, Y = 12 — True ribs are dorsally attached to vertebral column but are free on ventral side.

(4) X = 12, Y = 7 — True ribs are attached dorsally to vertebral column and ventrally to the sternum.

Ans. (4)

127. In case of poriferans, the spongocoel is lined with flagellated cells called:

(1) oscula

(2) choanocytes

(3) mesenchymal cells

(4) ostia

Ans. (2)

128. Which one of the following statements is not valid for aerosols ?

(1) They alter rainfall and monsoon patterns

(2) They cause increased agricultural productivity

(3) They have negative impact on agricultural land

(4) They are harmful to human health

Ans. (2)

129. A baby boy aged two years is admitted to play school and passes through a dental check - up. The dentist observed that the boy had twenty teeth. Which teeth were absent?

(1) Canines

(2) Pre-molars

(3) Molars

(4) Incisors

Ans. (2)

130. Select the mismatch

(1) *Cycas* – Dioecious

(2) *Salvinia* – Heterosporous

(3) *Equisetum* – Homosporous

(4) *Pinus* – Dioecious

Ans. (4)

131. The morphological nature of the edible part of coconut is:

(1) Cotyledon

(2) Endosperm

(3) Pericarp

(4) Perisperm

Ans. (2)

132. Double fertilization is exhibited by :

(1) Algae

(2) Fungi

(3) Angiosperms

(4) Gymnosperms

Ans. (3)

133. Spliceosomes are not found in cells of;

(1) Fungi

(2) Animals

(3) Bacteria

(4) Plants

Ans. (3)

134. The association of histone H1 with a nucleosome indicates:

(1) DNA replication is occurring.

(2) The DNA is condensed into a Chromatin Fibre.

(3) The DNA double helix is exposed.

(4) Transcription is occurring.

Ans. (2)

135. The region of Biosphere Reserve which is legally protected and where no human activity is allowed is known as:

(1) Buffer zone

(2) Transition zone

(3) Restoration zone

(4) Core zone

Ans. (4)

NEET(UG)–2017 TEST PAPER WITH ANSWER & SOLUTIONS (HELD ON SUNDAY 07th MAY, 2017)

136. Name the gas that can readily decolourise acidified $KMnO_4$ solution :

(1) SO_2 (2) NO_2 (3) P_2O_5 (4) CO_2

Ans. (1)

Sol. $\underset{(O.A.)}{\overset{+7}{KMnO_4}} + \underset{(R.A.)}{\overset{+4}{SO_2}} \rightarrow \underset{\text{Colourless}}{MnSO_4} + H_2SO_4 + K_2SO_4$

137. Mechanism of a hypothetical reaction

$X_2 + Y_2 \rightarrow 2XY$ is given below :

(i) $X_2 \rightarrow X + X$(fast)

(ii) $X + Y_2 \rightleftharpoons XY + Y$ (slow)

(iii) $X + Y \rightarrow XY$ (fast)

The overall order of the reaction will be :

(1) 2 (2) 0 (3) 1.5 (4) 1

Ans. (3)

Sol. According to law of mass action

$r = K[X][Y_2]$ ....(1)

From fast step-1

$$K_{eq} = \frac{[X]^2}{[X_2]}$$

$$[X]^2 = K_{eq.}[X_2]$$

$$[X] = \sqrt{K_{eq.}}[X_2]^{1/2} \quad ....(2)$$

From equation (1) & (2)

$$r = K.\sqrt{K_{eq.}}[X_2]^{1/2}[Y_2]$$

$$r = K'[X_2]^{1/2}[Y_2].$$

Overall order of reaction = 1 + 0.5 = 1.5

Option (3)

138. The element Z = 114 has been discovered recently. It will belong to which of the following family/group and electronic configuration ?

(1) Carbon family, [Rn] $5f^{14}\ 6d^{10}\ 7s^2\ 7p^2$

(2) Oxygen family, [Rn] $5f^{14}\ 6d^{10}\ 7s^2\ 7p^4$

(3) Nitrogen family, [Rn] $5f^{14}\ 6d^{10}\ 7s^2\ 7p^6$

(4) Halogen family, [Rn] $5f^{14}\ 6d^{10}\ 7s^2\ 7p^5$

Ans. (1)

Sol. Z = 114 $[Rn]^{86}\ 7s^2\ 5f^{14}\ 6d^{10}\ 7p^2$

14^{th} gp. (carbon family)

139. The heating of phenyl–methyl ethers with HI produces

(1) iodobenzene (2) phenol

(3) benzene (4) ethyl chlorides

Ans. (2)

Sol. $Ph-\ddot{O}-CH_3 \xrightarrow{H^{\oplus}} Ph-\overset{\oplus}{O}(H)-CH_3 \xrightarrow[SN^2]{I^{\ominus}} Ph-OH + I-CH_3$

140. Which one is the correct order of acidity ?

(1) $CH \equiv CH > CH_3 - C \equiv CH > CH_2 = CH_2 > CH_3 - CH_3$

(2) $CH \equiv CH > CH_2 = CH_2 > CH_3 - C \equiv CH > CH_3 - CH_3$

(3) $CH_3 - CH_3 > CH_2 = CH_2 > CH_3 - C \equiv CH > CH \equiv CH$

(4) $CH_2 = CH_2 > CH_3 - CH = CH_2 > CH_3 - C \equiv CH > CH \equiv CH$

Ans. (1)

Sol. Correct order of acidic strength $\Rightarrow$

$CH \equiv CH > CH_3 - C \equiv CH > CH_2 = CH_2 > CH_3 - CH_3$

acc. to EN and Inductive effect.

141. Predict the correct intermediate and product in the following reaction :

$$H_3C-C\equiv CH \xrightarrow[HgSO_4]{H_2O,\ H_2SO_4} \underset{(A)}{\text{Intermediate}} \longrightarrow \underset{(B)}{\text{product}}$$

(1) A : $H_3C-\underset{OH}{\underset{|}{C}}=CH_2$ B : $H_3C-\underset{SO_4}{\underset{|}{C}}=CH_2$

(2) A : $H_3C-\underset{O}{\underset{\|}{C}}-CH_3$ B : $H_3C-C\equiv CH$

(3) A : $H_3C-\underset{OH}{\underset{|}{C}}=CH_2$ B : $H_3C-\underset{O}{\underset{\|}{C}}-CH_3$

(4) A : $H_3C-\underset{SO_4}{\underset{|}{C}}=CH_2$ B : $H_3C-\underset{O}{\underset{\|}{C}}-CH_3$

Ans. (3)

Sol. $CH_3 - C \equiv CH \xrightarrow[HgSO_4]{H_2O,\ H_2SO_4} CH_3 - \underset{OH}{\underset{|}{C}} = CH_2 \xrightarrow{\text{Tautomerism}} CH_3 - \underset{O}{\underset{\|}{C}} - CH_3$

142. The equilibrium constant of the following are :

$N_2 + 3H_2 \rightleftharpoons 2NH_3$ K_1

$N_2 + O_2 \rightleftharpoons 2NO$ K_2

$H_2 + \frac{1}{2}O_2 \rightarrow H_2O$ K_3

The equilibrium constant (K) of the reaction :

$2NH_3 + \frac{5}{2}O_2 \overset{K}{\rightleftharpoons} 2NO + 3H_2O$, will be :

(1) $K_2 K_3^3 / K_1$ (2) K_2K_3/K_1

(3) $K_2^3K_3 / K_1$ (4) $K_1K_3^3 / K_2$

Ans. (1)

Sol. $N_2 + 3H_2 \rightleftharpoons 2NH_3$ $K_1 \rightarrow (1)$

$N_2 + O_2 \rightleftharpoons 2NO$ $K_2 \rightarrow (2)$

$H_2 + \frac{1}{2}O_2 \longrightarrow H_2O$ $K_3 \rightarrow (3)$

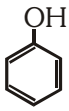
For reaction $2NH_3 + \frac{5}{2}O_2 \overset{K}{\rightleftharpoons} 2NO + 3H_2O \rightarrow$ (4)

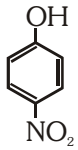
equation (4)

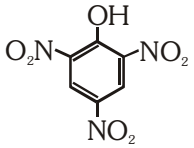
= equation(2) + 3 × equation(3) − equation(1)

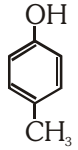
$\Rightarrow K = \frac{K_2.K_3^3}{K_1}$, Option(1)

143. Which one is the most acidic compound ?

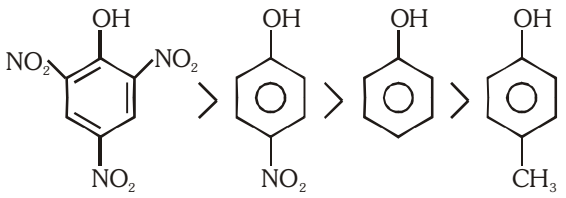
(1) 

(2) 

(3) 

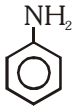
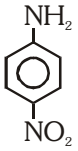
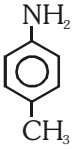
(4) 

Ans. (3)

Sol. 

More − I, − M, more acidic

144. The **correct** increasing order of basic strength for the following compounds is :

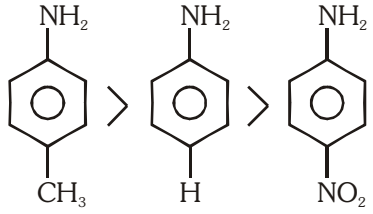
(I)  (II)  (III) 

(1) III < I < II (2) III < II < I

(3) II < I < III (4) II < III < I

Ans. (3)

Sol. Order of Basic Strength :-



electron donating

electron withdrawing

145. Ionic mobility of which of the following alkali metal ions is lowest when aqueous solution of their salts are put under an electric field ?

(1) K (2) Rb

(3) Li (4) Na

Ans. (3)

Sol. Ionic mobility $\propto \frac{1}{\text{size of hydrated ion}}$

Smaller size hydrated ion in aq. sol^n - $Rb^+(aq)$

Larger size hydrated ion in aq. sol^n - $Li^+(aq)$

Lowest ionic mobility in aq. sol^n $\rightarrow Li^+(aq)$ due to high hydration

146. The most suitable method of separation of 1 : 1 mixture of ortho and para-nitrophenols is :

(1) Chromatography (2) Crystallisation

(3) Steam distillation (4) Sublimation

Ans. (3)

Sol. The ortho and para isomers can be separated by steam distillation o-Nitrophenol is steam volatile due to intramolecular hydrogen bonding while p-nitro phenol is less volatile due to intermolecular hydrogen bonding which cause association of molecule.

147. $HgCl_2$ and I_2 both when dissolved in water containing I^- ions the pair of species formed is :

(1) HgI_2, I^- (2) HgI_4^{2-}, I_3^-

(3) Hg_2I_2, I^- (4) HgI_2, I_3^-

Ans. (2)

Sol. $HgCl_2 + 2I^- \longrightarrow HgI_2 + 2C\ell^-$

$\downarrow + 2I^-$

$[HgI_4]^{-2}$

Soluble complex

$I_2 + I^- \longrightarrow I_3^-$

water soluble

148. Mixture of chloroxylenol and terpineol acts as :

(1) antiseptic (2) antipyretic

(3) antibiotic (4) analgesic

Ans. (1)

Sol. Antiseptic (dettol)

149. An example of a sigma bonded organometallic compound is :

(1) Grignard's reagent (2) Ferrocene

(3) Cobaltocene (4) Ruthenocene

Ans. (1)

150. A first order reaction has a specific reaction rate of 10^{-2} sec^{-1}. How much time will it take for 20g of the reactant to reduce to 5 g ?

(1) 138.6 sec (2) 346.5 sec

(3) 693.0 sec (4) 238.6 sec

Ans. (1)

Sol. Half life of first order reaction $t_{1/2} = \frac{0.693}{K}$

$= \frac{0.693}{10^{-2}} = 69.3 sec$

Method-1

$20g \xrightarrow{t_{1/2}} 10g \xrightarrow{t_{1/2}} 5g$

Total time = $2t_{1/2} = 2 \times 69.3 = 138.6$ sec

Method-2

$t = \frac{2.303}{K} \log \frac{[A]_o}{[A]_t}$

$t = \frac{2.303}{10^{-2}} \log \frac{20}{5} \Rightarrow t = 138.6$ sec (Option 2)

151. Match the interhalogen compounds of column-I with the geometry in column II and assign the correct. code.

Column-I		Column-II	
(a)	XX'	(i)	T-shape
(b)	XX'_3	(ii)	Pentagonal bipyramidal
(c)	XX'_5	(iii)	Linear
(d)	XX'_7	(iv)	Square-pyramidal
		(v)	Tetrahedral

Code :

	(a)	(b)	(c)	(d)
(1)	(iii)	(i)	(iv)	(ii)
(2)	(v)	(iv)	(iii)	(ii)
(3)	(iv)	(iii)	(ii)	(i)
(4)	(iii)	(iv)	(i)	(ii)

Ans. (1)

Sol. $XX' \Rightarrow$ Linear

$XX_3' \Rightarrow$ T-shape sp^3d

$XX_5' \Rightarrow$ Square pyramidal sp^3d^2

$XX_7' \Rightarrow$ Pentagonal bipyramidal (sp^3d^3)

152. Concentration of the Ag^+ ions in a saturated solution of $Ag_2C_2O_4$ is 2.2×10^{-4} mol L^{-1} Solubility product of $Ag_2C_2O_4$ is :-

(1) 2.66×10^{-12} (2) 4.5×10^{-11}

(3) 5.3×10^{-12} (4) 2.42×10^{-8}

Ans. (3)

Sol. $Ag_2C_2O_4 \rightleftharpoons 2Ag^+ + C_2O_4^{2-}$

2.2×10^{-4}M 1.1×10^{-4}M

$K_{sp} = [Ag^+]^2[C_2O_4^{2-}]$

$= [2.2 \times 10^{-4}]^2.[1.1 \times 10^{-4}]$

$K_{sp} = 5.3 \times 10^{-12}$

153. In the electrochemical cell :-
$Zn\,|\,ZnSO_4(0.01M)\,||\,CuSO_4(1.0\,M)\,|\,Cu$, the emf of this Daniel cell is E_1. When the concentration of $ZnSO_4$ is changed to 1.0M and that of $CuSO_4$ changed to 0.01M, the emf changes to E_2. From the followings, which one is the relationship between E_1 and E_2? (Given, $\frac{RT}{F}=0.059$)

(1) $E_1 < E_2$ (2) $E_1 > E_2$
(3) $E_2 = 0 \neq E_1$ (4) $E_1 = E_2$

Ans. (2)

Sol. For cell
$Zn\,|\,ZnSO_4(0.01M)\,||\,CuSO_4(1M)\,|\,Cu$
Cell reaction $\rightarrow Zn + Cu^{+2} \longrightarrow Zn^{+2} + Cu$

$$E_1 = E^o - \frac{0.059}{2}\log\frac{Zn^{+2}}{Cu^{+2}}$$

$$E_1 = E^o - \frac{0.059}{2}\log\frac{0.01}{1}$$

$$= E^o - \frac{0.059}{2}\log\frac{1}{100} \quad ....(1)$$

For cell
$Zn\,|\,ZnSO_4(1M)\,||\,CuSO_4(0.01M)\,|\,Cu$

$$E_2 = E^o - \frac{0.059}{2}\log\frac{1}{0.01}$$

$$= E^o - \frac{0.059}{2}\log 100 \quad ...(2) \qquad \boxed{E_1 > E_2}$$

Option (2)

154. Which of the following pairs of compounds is isoelectronic and isostructural ?

(1) TeI_2, XeF_2

(2) IBr_2^-, XeF_2

(3) IF_3, XeF_2

(4) $BeCl_2, XeF_2$

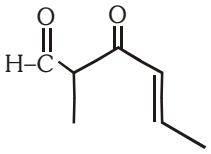
Ans. (2)

Sol. IBr_2^{-1} & Xef_2 are iso-structural

(Linear shape)

and Both C.A. consist of same no. of valence e^-s

155. The IUPAC name of the compound

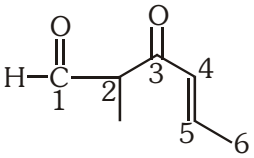


is :-

(1) 5-formylhex-2-en-3-one
(2) 5-methyl-4-oxohex-2-en-5-al
(3) 3-keto-2-methylhex-5-enal
(4) 3-keto-2-methylhex-4-enal

Ans. (4)

Sol.



3–keto–2–methylhex–4–en–1–al

156. Which one is the wrong statement ?

(1) The uncertainty principle is $\Delta E \times \Delta t \geq h/4\pi$
(2) Half filled and fully filled orbitals have greater stability due to greater exchange energy, greater symmetry and more balanced arrangement.
(3) The energy of 2s orbital is less than the energy of 2p orbital in case of Hydrogen like atoms
(4) de-Broglies's wavelength is given by $\lambda = \frac{h}{m\nu}$, where m = mass of the particle, ν = group velocity of the particle

Ans. (3)

Sol. In H-like atom energy of 2s = 2p. orbital
Incorrect statement is (3)

157. Which is the **incorrect** statement ?

(1) Density decreases in case of crystals with Schottky's defect
(2) NaCl(s) is insulator, silicon is semiconductor, silver is conductor, quartz is piezo electric crystal
(3) Frenkel defect is favoured in those ionic compounds in which sizes of cation and anions are almost equal
(4) $FeO_{0.98}$ has non stoichiometric metal deficiency defect

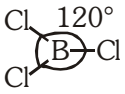
Ans. (3)

Sol. In frenkel defect the radius of cation must be very less than anion.
Incorrect statement is (3)

158. The species, having bond angles of 120° is :-

(1) ClF_3 (2) NCl_3 (3) BCl_3 (4) PH_3

Ans. (3)

Sol. $BCl_3 \Rightarrow$  Regular geometry
Hybridysation $\Rightarrow sp^2$

159. For a given reaction, $\Delta H = 35.5\ kJ\ mol^{-1}$ and $\Delta S = 83.6\ JK^{-1}mol^{-1}$. The reaction is spontaneous at : (Assume that ΔH and ΔS do not vary with tempearature)

(1) $T > 425\ K$ (2) All temperatures

(3) $T > 298\ K$ (4) $T < 425\ K$

Ans. (1)

Sol. $\Delta G = \Delta H - T\Delta S$

for equilibrium $\Delta G = 0$

$\Delta H = T\Delta S$

$$T_{eq.} = \frac{\Delta H}{\Delta S} = \frac{35.5 \times 1000}{83.6} = 425K$$

Since the reaction is endothemic it will be spontaneous at $T > 425K$. Option (1)

160. Which of the following is a sink for CO ?

(1) Micro organism present in the soil

(2) Oceans

(3) Plants

(4) Haemoglobin

Ans. (1)

Sol. Microorganism present in the soil.

161. If molality of the dilute solutions is doubled, the value of molal depression constant (K_f) will be :-

(1) halved (2) tripled

(3) unchanged (4) doubled

Ans. (3)

Sol. K_f does not depend on concentration of solution. It only depends on nature of solvent so it will be unchanged. option (3)

162. Which of the following is dependent on temperature?

(1) Molarity (2) Mole fraction

(3) Weight percentage (4) Molality

Ans. (1)

Sol. Temperature dependent unit is molarity.

163. Which one of the following statements is not correct?

(1) The value of equilibrium constant is changed in the presence of a catalyst in the reaction at equilibrium

(2) Enzymes catalyse mainly bio-chemical reactions

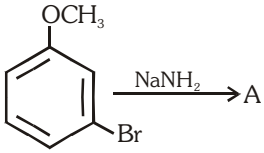
(3) Coenzymes increase the catalytic activity of enzyme

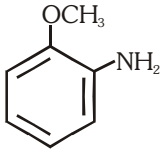
(4) Catalyst does not initiate any reaction

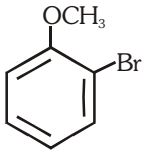
Ans. (1)

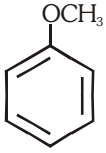
Sol. Equilibrium constant is not affected by presence of catalyst hence statement (1) is incorrect.

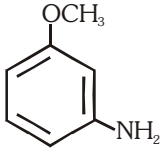
164. Identify A and predict the type of reaction



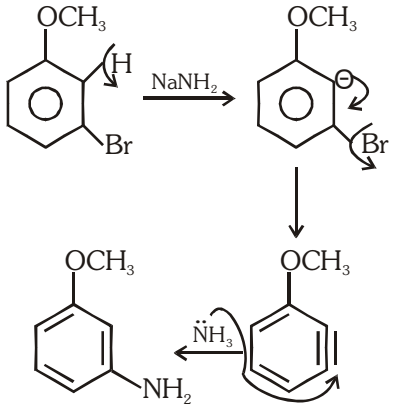
(1)  and elimination addition reaction

(2)  and cine substitution reaction

(3)  and cine substituion reaction

(4)  and substitution reaction

Ans. (4)

Sol. 

Example of substitution reaction.

165. The correct order of the stoichiometries of AgCl formed when $AgNO_3$ in excess is treated with the complexs : $CoCl_3.6NH_3$, $CoCl_3.5NH_3$, $CoCl_3.4NH_3$ respectively is :-

(1) 3 AgCl, 1 AgCl, 2 AgCl

(2) 3 AgCl, 2 AgCl, 1 AgCl

(3) 2 AgCl, 3 AgCl, 1 AgCl

(4) 1 AgCl, 3 AgCl, 2 AgCl

Ans. (2)

Sol. $[CO(NH_3)_6]Cl_3 \xrightarrow{AgNO_3}$ 3 mol AgCl

$[CO(NH_3)_5\ Cl]Cl_2 \xrightarrow{AgNO_3}$ 2 mol AgCl

$[CO(NH_3)_4\ Cl_2]Cl \xrightarrow{AgNO_3}$ 1 mol AgCl

166. The **correct** statement regarding electrophile is :-

(1) Electrophile is a negatively charged species and can form a bond by accepting a pair of electrons from another electrophile

(2) Electrophiles are generally neutral species and can form a bond by accepting a pair of electrons from a nucleophile

(3) Electrophile can be either neutral or positively charged species and can form a bond by accepting a pair of electrons from a nucleophile

(4) Electrophile is a negatively charged species and can form a bond by accepting a pair of electrons from a nucleophile

Ans. (3)

Sol. Electrophile can be either neutral or positively charged species and can form a bond by accepting a pair of electron from a nucleophile.

167. A gas is allowed to expand in a well insulated container against a constant external pressure of 2.5atm from an initial volume of 2.50 L to a final volume of 4.50L. The change in internal energy ΔU of the gas in joules will be:-

(1) –500J (2) –505J

(3) +505J (4) 1136.25J

Ans. (2)

Sol. Work done in irreversible process

$W = -P_{ext}\Delta V$

$= -2.5\ [4.5 - 2.5] = -5$ L atm

$= -5 \times 101.3J = -505J$

Since system is well insulated q = 0

By FLOT $\Delta E = q + W$

$\Delta E = W = -505$ J Option(2)

168. Which of the following reactions is appropriate for converting acetamide to methanamine ?

(1) Hoffmarnn hypobromamide reaction

(2) Stephens reaction

(3) Gabriels phthalimide synthesis

(4) Carbylamine reaction

Ans. (1)

Sol. $CH_3 - \underset{\underset{O}{\|}}{C} - NH_2 \xrightarrow{Br_2/4KOH} CH_3 - NH_2 + 2KBr + K_2CO_3$

This reaction is known as hoffmann hypobromamide reaction.

169. With respect to the conformers of ethane, which of the following statements is **true** ?

(1) Bond angle changes but bond length remains same

(2) Both bond angle and bond length change

(3) Both bond angles and bond length remains same

(4) Bond angle remains same but bond length changes

Ans. (3)

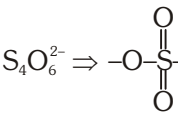
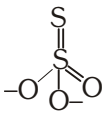
Sol. In conformation bond angle and bond length remain same.

170. In which pair of ions both the species contain S–S bond?

(1) $S_4O_6^{2-}$, $S_2O_3^{2-}$ (2) $S_2O_7^{2-}$, $S_2O_8^{2-}$

(3) $S_4O_6^{2-}$, $S_2O_7^{2-}$ (4) $S_2O_7^{2-}$, $S_2O_3^{2-}$

Ans. (1)

Sol. $S_4O_6^{2-} \Rightarrow$  | $S_2O_3^{2-} \Rightarrow$ 

171. It is because of inability of ns^2 electrons of the valence shell to participate in bonding that:-

(1) Sn^{2+} is oxidising while Pb^{4+} is reducing

(2) Sn^{2+} and Pb^{2+} are both oxidising and reducing

(3) Sn^{4+} is reducing while Pb^{4+} is oxidising

(4) Sn^{2+} is reducing while Pb^{4+} is oxidising

Ans. (In English-4, In Hindi-1)

Sol. $Sn^{+2} \longrightarrow Sn^{+4}$

(R.A) $Sn^{+2} < Sn^{+4}$ Stability order

$Pb^{+4} \longrightarrow Pb^{+2}$

(O.A) $Pb^{+2} > Pb^{+4}$ Stability order

(Inert pair effect)

172. Correct increasing order for the wavelengths of absorption in the visible region the complexes of Co^{3+} is :-

(1) $[Co(H_2O)_6]^{3+}$, $[Co(en)_3]^{3+}$, $[Co(NH_3)_6]^{3+}$

(2) $[Co(H_2O)_6]^{3+}$, $[Co(NH_3)_6]^{3+}$, $[Co(en)_3]^{3+}$

(3) $[Co(NH_3)_6]^{3+}$, $[Co(en)_3]^{3+}$, $[Co(H_2O)_6]^{3+}$

(4) $[Co(en)_3]^{3+}$, $[Co(NH_3)_6]^{3+}$, $[Co(H_2O)_6]^{3+}$

Ans. (4)

Sol. $\left[\varepsilon_a \propto \frac{1}{\lambda_a}\right]$

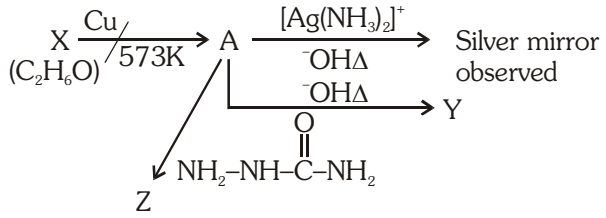
where $\varepsilon_a \Rightarrow$ absorbed energy (splitting energy)

$\lambda_a \Rightarrow$ absorbed wavelength

Presence of SFL $\Rightarrow \varepsilon_a(\uparrow)\ \lambda_a(\downarrow)$

$H_2O < NH_3 <$ en ligand strength $\uparrow$ splitting energy $\uparrow$ so absorbed $\lambda\downarrow$

173. Consider the reactions :-



Identify A, X, Y and Z

(1) A-Methoxymethane, X-Ethanol, Y-Ethanoic acid, Z-Semicarbazide.

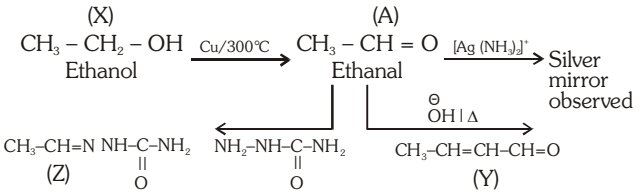
(2) A-Ethanal, X-Ethanol, Y-But-2-enal, Z-Semicarbazone

(3) A-Ethanol, X-Acetaldehyde, Y-Butanone, Z-Hydrazone

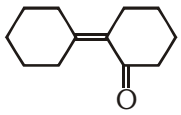
(4) A-Methoxymethane, X-Ethanoic acid, Y-Acetate ion, Z-hydrazine

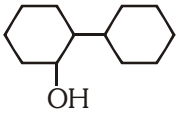
Ans. (2)

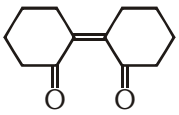
Sol.

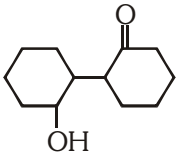


174. Of the following, which is the product formed when cyclohexanone undergoes aldol condensation followed by heating ?:-

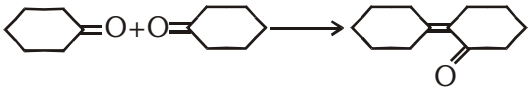
(1) 

(2) 

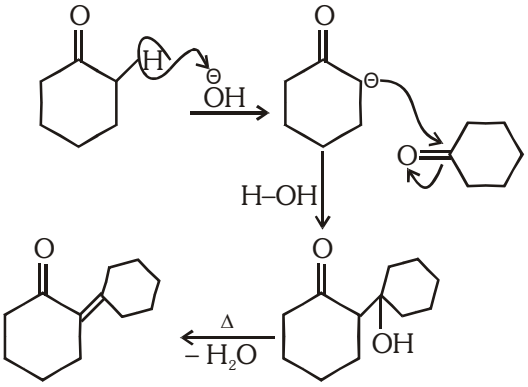
(3) 

(4) 

Ans. (1)

Sol. 

Mechanism



175. Which of the following pairs of species have the same bond order ?

(1) O_2, NO^+ (2) CN^-, CO

(3) N_2 , O_2^- (4) CO, NO

Ans. (2)

Sol. Total no. of electrons in CN^- is 14

Total no. of electrons in CO is also 14

hence B.O. of both CN^- & CO is 3

176. Extraction of gold and silver involes leaching with CN^-ion. Silver is later recovered by :-

(1) distillation

(2) zone refining

(3) displacement with Zn

(4) liquation

Ans. (3)

Sol. Mac arther forest process/cyanide process

$$Ag_2S + 4NaCN \xrightarrow{O_2} 2Na[Ag(CN)_2] + Na_2SO_4$$

$$2Na[Ag(CN)_2] \xrightarrow{Zn} Na_2[Zn(CN)_4] + Ag(\downarrow)$$

Soluble complex

Ag extracts by displacement with Zn

177. A 20 litre container at 400 K contains $CO_2(g)$ at pressure 0.4 atm and an excess of SrO (neglect the volume of solid SrO). The volume of the container is now decreased by moving the movable piston fitted in the container. The maximum volume of the container, when pressure of CO_2 attains its maximum value, will be :-

(Given that : $SrCO_3(s) \rightleftharpoons SrO(s) + CO_2(g)$, Kp = 1.6atm)

(1) 10 litre (2) 4 litre

(3) 2 litre (4) 5 litre

Ans. (4)

Sol. $SrCO_3(s) \rightleftharpoons SrO(s) + CO_2(g)$

$$K_p = P_{CO_2}$$

maximum pressure of CO_2 = 1.6 atm

$$P_1V_1 = P_2V_2$$

$0.4 \times 20 = 1.6\ V_2$

V2 = 5L option (4)

178. Pick out the correct statement with respect to $[Mn(CN)_6]^{3-}$:-

(1) It is sp^3d^2 hybridised and tetrahedral

(2) It is d^2sp^3 hybridised and octahedral

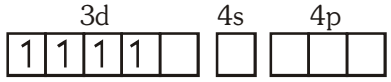
(3) It is dsp^2 hybridised and square planar

(4) It is sp^3d^2 hybridised and octahedral

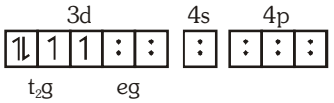
Ans. (2)

Sol. $[Mn(CN)_6]^{3-} \rightarrow$ O.S. of Mn is (+3)

C.N. = 6

$Mn^{+3} \rightarrow 3d^4 \rightarrow$ 

Presence of SFL (Pairing is possible)



$[d^2sp^3$, octahedral]

179. The reason for greater range of oxidation states in actinoids is attributed to :-

(1) actinoid contraction

(2) 5f, 6d and 7s levels having comparable energies

(3) 4f and 5d levels being close in energies

(4) the redioactive nature of actinoids

Ans. (2)

Sol. Minimum energy gap between

5f, 6d & 7s subshell. Thats why e^- exitation will be easeir.

180. Which of the following statements is not correct :-

(1) Ovalbumin is a simple food reserve in egg-white

(2) Blood proteins thrombin and fibrinogen are involved in blood clotting

(3) Denaturation makes the proteins more active

(4) Insulin maintanis sugar level in the blodd of a human body

Ans. (3)

Sol. Denaturation makes the protein more active.