MSEG201 Homework 5

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Steady State Diffusion

This can be solved using Fick's First law of diffusion:

$$J = -D\frac{dC}{dx}$$

This can be rearranged for the concentration gradient and solved for the high-pressure concentration

$$\frac{dC}{dx} = -\frac{J}{D}$$

$$dC = C_{\text{high-pressure}} - C_{\text{low-pressure}} = +\frac{J}{D}dx$$

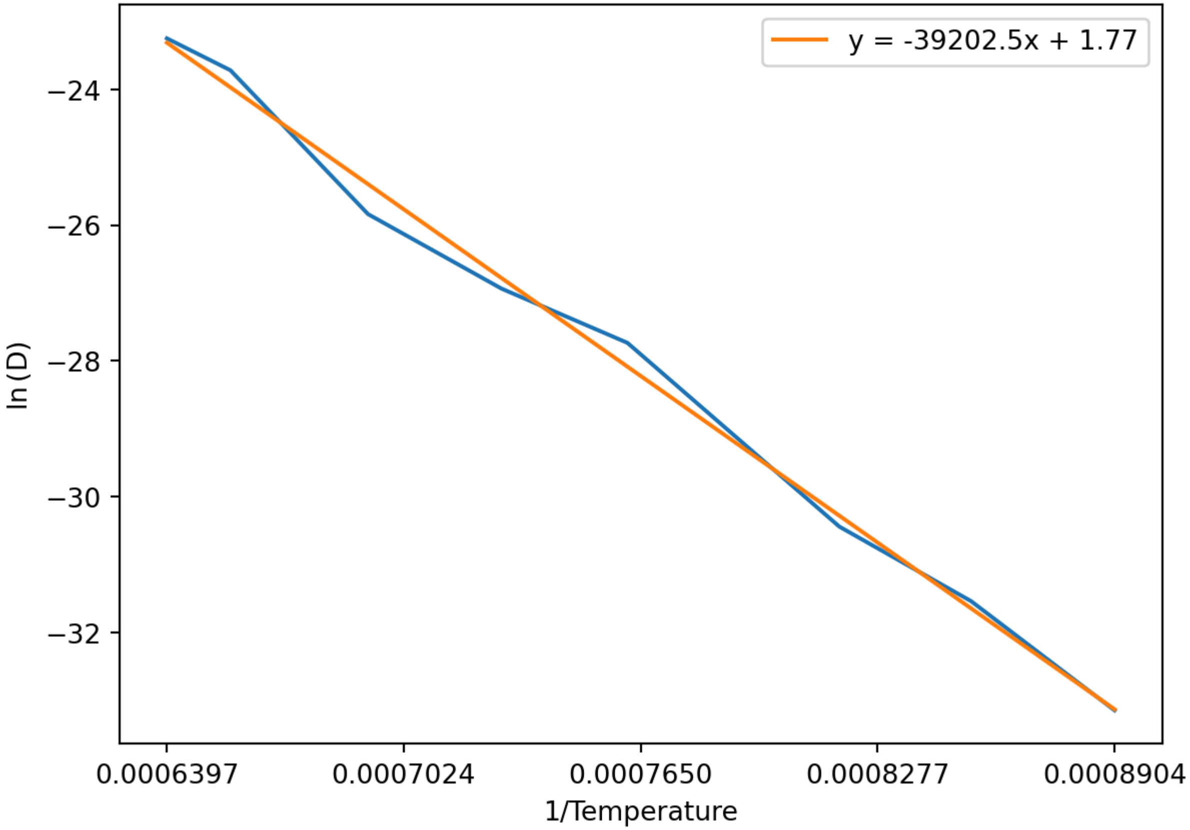
$$C_{\text{high-pressure}} = C_{\text{low-pressure}} + \frac{J}{D}dx$$

Now plugging in the given values:

$$C_{\text{high-pressure}} = 0.5\,\mathrm{kg/m^3} + (0.002\,\mathrm{m})\frac{2.48\times10^{-8}\,\mathrm{kg/m^2s}}{4.50\times10^{-11}\,\mathrm{m^2/s}} = 1.60\,\mathrm{kg/m^3}$$

Diffusion of Aluminum in Silicon

▶ Code



a)

This has effectively plotted $\ln(D) = \ln(D_0) + \frac{1}{T}\left(\frac{-Q_d}{R}\right)$. The intercept is $\ln(D_0)$, so $D_0 = \exp(b) = 58.4\,\mathrm{cm}^2/\mathrm{s}$.
The slope is $\frac{-Q_d}{R}$, so $Q_d = -mR = -(-39202)(8.6173 \times 10^{-5}) = 3.38$ eV

b)

Going back to the equation

$$\ln(D(T = 1000^{\circ}C)) = -39202.5 \times (1000 + 273.15)^{-1} + 1.77 = -29.02$$

$$D = \exp(-29.02) = 2.48 \times 10^{-13}\,\mathrm{cm}^2/\mathrm{s}$$

Doping and Diffusion

First we need to find D!

$$D = 2.14 \times 10^{-5} \exp \left[\frac{3.65}{\left(8.62 \times 10^{-5}\right)(1200 + 273.15)} \right]$$

$$D = 6.509 \times 10^{7}$$

Now rearranging the thin source solution

$$\frac{C\sqrt{\pi D t}}{B} = \exp \left[\frac{-x^2}{4Dt} \right]$$

$$x = \sqrt[2]{-4Dt \ln \left[\frac{C\sqrt{\pi D t}}{B} \right]}$$

$$x = -2 \sqrt[2]{(6.509 \times 10^{7})(1.75 \times 3600) \ln \left[\frac{3.63 \times 10^{17} \sqrt{\pi (6.509 \times 10^{7})(1.75 \times 3600)}}{5.2 \times 10^{3}} \right]}$$

Carburization

a)

It matches up with the semi-infinite bar solution since there is in initial concentration throughout the system, and the diffusion takes place on mostly the first 4mm and hopefully the gear is much thicker than 4mm.

b)

Remembering the semi-infinite bar solution and reorganizing for t

$$\frac{C(x,t)-C_0}{C_s-C_x} = 1-\text{erf}\left(\frac{x}{2\sqrt{Dt}}\right)$$

$$\text{erf}\left(\frac{x}{2\sqrt{Dt}}\right) = 0.75$$

And using wolfram alpha for the inverse erf

$$\frac{x}{2\sqrt{Dt}} = 0.8134$$

$$t = \frac{1}{D}\left[\frac{x}{2\times 0.8134}\right]^2$$

finding D

$$D = D_0\exp\left(\frac{-Q}{RT}\right) = 2.5\times 10^{-5}\exp\left(\frac{-1.48\times 10^5}{8.314\times 1373}\right) = 5.851\times 10^{-11}\,\text{m}^2/\text{s}$$

and finally plugging in

$$t = \frac{1}{5.851\times 10^{-11}\,\text{m}^2/\text{s}}\left[\frac{0.004\,\text{m}}{2\times 0.8134}\right]^2 = 103328.84\,\text{sec}$$

and into the right units

$$103328.84\,\text{sec}\times\frac{1\,\text{hour}}{3600\,\text{sec}} = 28.70\,\text{hours}$$

c)

▶ Code

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Text(0.5, 1.0, 'Temperature Profile')
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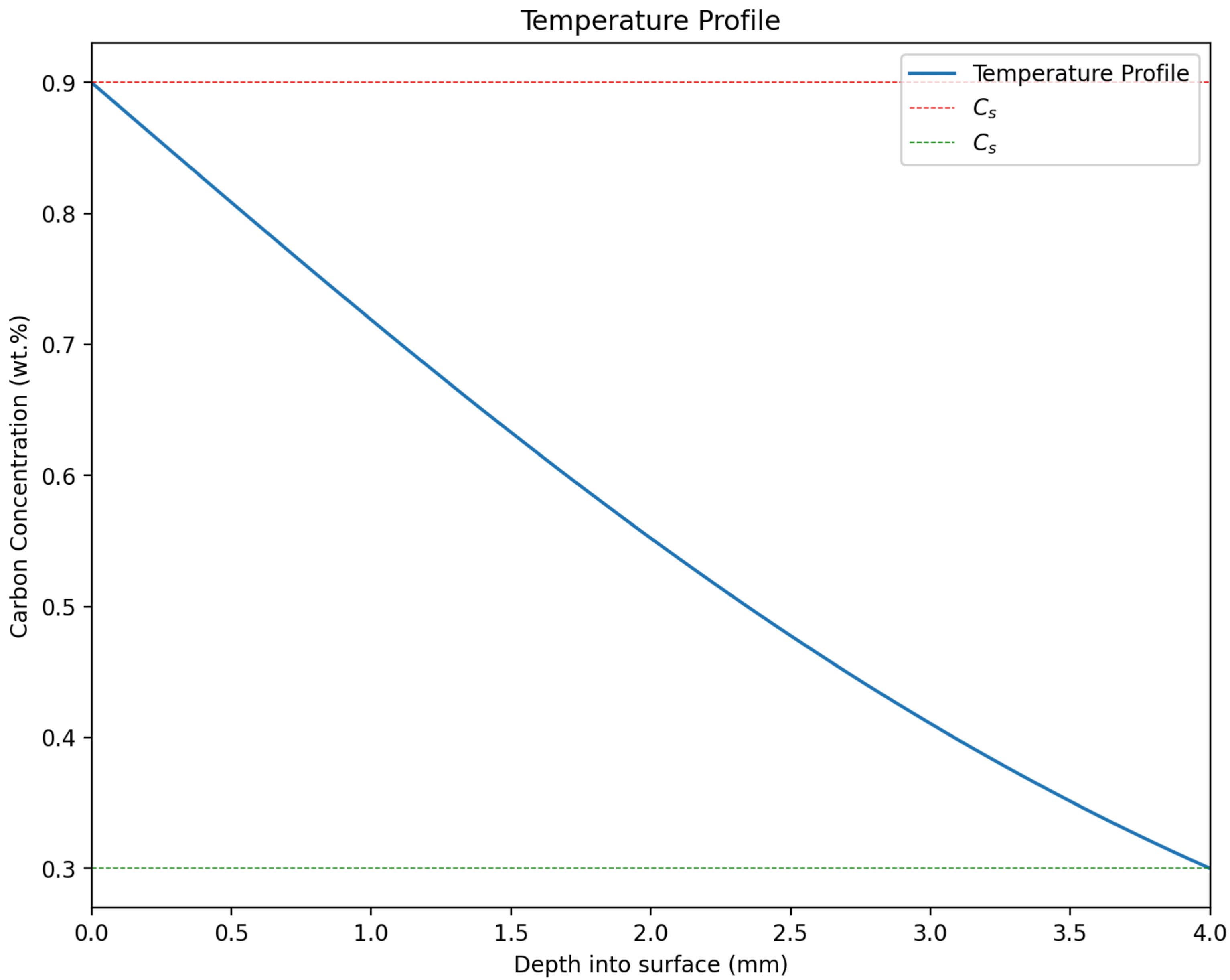

Temperature Profile
Temperature Profile
C_s
C_s
0.9
0.8
0.7
0.6
0.5
0.4
0.3
Carbon Concentration (wt.%)
0.0
0.5
1.0
1.5
2.0
2.5
3.0
3.5
4.0
Depth into surface (mm)

d)

Now D is

$$D = 2.5 \times 10^{-5} \exp\left(\frac{-1.48 \times 10^5}{8.314 \times 973}\right) = 2.834 \times 10^{-13}\,\mathrm{m^2/s}$$

so the time is

$$t = \frac{1}{\left(2.834 \times 10^{-13}\,\mathrm{m^2/s}\right)(3600)}\left[\frac{0.004\,\mathrm{m}}{2 \times 0.8134}\right]^2 = 5926\,\mathrm{hours}$$