



Journal of Macromolecular Science: Part A - Chemistry: Pure and Applied Chemistry

Publication details, including instructions for
authors and subscription information:

<http://www.tandfonline.com/loi/lmsa19>

Copolymerization of 2- Hydroxypropyl Methacrylate with Alkyl Met hacrylates

G. N. Babu ^a & Anjali Deshpande ^a

^a Department of Chemistry , Indian Institute of
Technology , Powai, Bombay, 400076, India
Published online: 05 Dec 2006.

To cite this article: G. N. Babu & Anjali Deshpande (1982) Copolymerization of 2-Hydroxypropyl Methacrylate with Alkyl Met hacrylates, Journal of Macromolecular Science: Part A - Chemistry: Pure and Applied Chemistry, 17:5, 717-726, DOI: [10.1080/00222338208063268](https://doi.org/10.1080/00222338208063268)

To link to this article: <http://dx.doi.org/10.1080/00222338208063268>

PLEASE SCROLL DOWN FOR ARTICLE

Taylor & Francis makes every effort to ensure the accuracy of all the information (the "Content") contained in the publications on our platform. However, Taylor & Francis, our agents, and our licensors make no representations or warranties whatsoever as to the accuracy, completeness, or suitability for any purpose of the Content. Any opinions and views expressed in this publication are the opinions and views of the authors, and are not the views of or endorsed by Taylor & Francis. The accuracy of the Content should not be relied upon and should be independently verified with primary sources of information. Taylor and Francis shall not be liable for any losses, actions, claims, proceedings, demands, costs, expenses, damages, and other liabilities whatsoever

or howsoever caused arising directly or indirectly in connection with, in relation to or arising out of the use of the Content.

This article may be used for research, teaching, and private study purposes. Any substantial or systematic reproduction, redistribution, reselling, loan, sub-licensing, systematic supply, or distribution in any form to anyone is expressly forbidden. Terms & Conditions of access and use can be found at <http://www.tandfonline.com/page/terms-and-conditions>

Copolymerization of 2-Hydroxypropyl Methacrylate with Alkyl Methacrylates

G. N. BABU and ANJALI DESHPANDE

Department of Chemistry
Indian Institute of Technology, Powai
Bombay 400076, India

ABSTRACT

2-Hydroxypropyl methacrylate (2-HPMA) has been copolymerized with ethyl methacrylate (EMA), n-butyl methacrylate (BMA), and 2-ethylhexyl methacrylate (EHMA) in bulk at 60°C using benzoyl peroxide as initiator. The copolymer composition has been determined from the hydroxyl content. The reactivity ratios have been calculated by the Yezrielev, Brokhina, and Raskin method. For copolymerization of 2-HPMA (M_1) with EMA (M_2), the reactivity ratios are $r_1 = 1.807 \pm 0.032$ and $r_2 = 0.245 \pm 0.021$; with BMA (M_2) they are $r_1 = 2.378 \pm 0.001$ and $r_2 = 0.19 \pm 0.01$; and with EHMA the values are $r_1 = 4.370 \pm 0.048$ and $r_2 = 0.103 \pm 0.006$. Since reactivity ratios are the measure of distribution of monomer units in copolymer chain, the values obtained are compared and discussed. This enables us to choose a suitable copolymer for synthesizing thermoset acrylic polymers, which are obtained from cross-linking of hydroxy functional groups of HPMA units, for specific end-uses.

INTRODUCTION

Acrylic polymers containing hydroxy functional groups as side chains have become increasingly important in recent years for various applications such as adhesives, coatings, and binders [1]. Although linear polymers of this type are satisfactory in many ways, their thermoplastic nature affords poor resistance to strong solvents and they show a tendency to flow at high temperatures, i.e., they exhibit poor physical properties. This problem has been eliminated by reacting them with suitable cross-linking agents, depending on the end use. For instance, in 2-hydroxypropyl methacrylate/methyl methacrylate copolymer the cross-linking agents used are diethylene glycol dimethacrylate for medical use [2] and diisocyanates for adhesives [3]. The thermomechanical properties of the thermosets have been found to be dependent on the amount of HPMA copolymer [4]. To optimize the physical properties, it was necessary to control the cross-linkable units. With this objective, the copolymerization of 2-hydroxypropyl methacrylate (2-HPMA) with ethyl methacrylate (EMA), n-butyl methacrylate (BMA), and 2-ethylhexyl methacrylate (2-EHMA) was done; the reactivity ratios were determined.

EXPERIMENTAL

The monomers ethyl methacrylate (BDH), butyl methacrylate (BDH), and 2-ethylhexyl methacrylate (Fluka AG) were freed from inhibitor by washing with 10% sodium hydroxide and then with water and finally dried over anhydrous magnesium sulfate. The purified monomers were distilled under reduced pressure over 4 Å molecular sieves.

2-PMA (E. Merck) was purified by distillation at 84–85°C/5 mmHg.

Benzoyl peroxide (BDH) was purified by crystallization from methanol. A stock solution of 2% benzoyl peroxide in chloroform was prepared and stored at room temperature.

Copolymerization

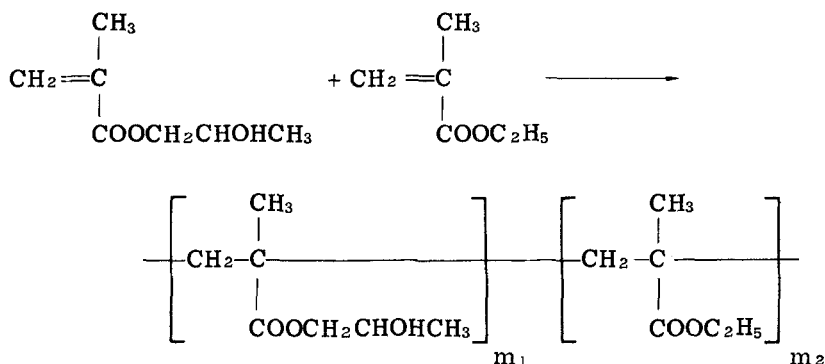
The required amounts of 2-HPMA, alkyl methacrylate, and benzoyl peroxide were charged in a Pyrex glass tube, cooled in liquid air, and sealed in a vacuum system. The sealed tubes were placed in a constant temperature bath at $60 \pm 0.1^\circ\text{C}$. Since the differential form of the copolymer equation was used, the conversion was limited to 10% in all cases. After the requisite time the contents of the tube were poured into a large amount of petroleum ether (40–60°C) and washed several times in petroleum ether. The purified polymers were dried at 50°C under vacuum to constant weight.

The composition of polymers was determined from hydroxyl content estimated by acetylation [5] with a mixture of acetic anhydride and pyridine (1:3 v/v).

RESULTS AND DISCUSSION

Reactivity Ratios in Copolymerization of 2-Hydroxypropyl Methacrylate (M_1) with Ethyl Methacrylate (M_2)

Copolymerization of 2-hydroxypropyl methacrylate with ethyl methacrylate is represented by



The reaction conditions and results are summarized in Table 1. From these data the monomer reactivity ratios were found by using the analytical method of Yezrielev, Brokhina, and Raskin (YBR method) [6], which is based on the differential form of the copolymer equation. Quite recently, Joshi [7] has reviewed various published methods of calculating r_1 and r_2 . He has stated that the linear YBR method is an outstanding method to obtain reactivity ratio values with greater accuracy than the Tidwell and Mortimer method [8] because of the symmetrical nature of the copolymer equation. The reactivity ratios for this system were found to be

$$r_1 = 1.807 \pm 0.032$$

$$r_2 = 0.245 \pm 0.021$$

Figure 1 shows the copolymer composition data as a function of monomer feed. As can be seen from the graph, the theoretical curve closely represents the experimental data. This is further reflected in the small deviations obtained for the reactivity ratios. The higher value of r_1 as compared to r_2 indicates that the copolymer containing large blocks of 2-HPMA units is interspersed by EMA units.

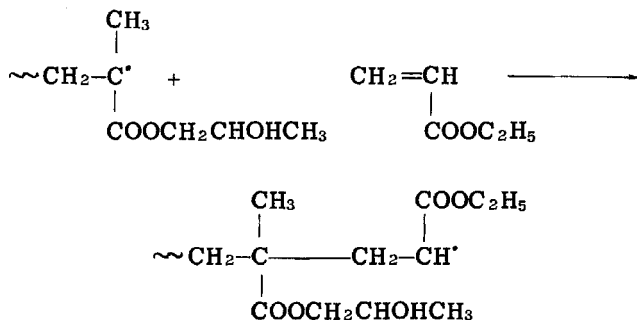
The effect of the α -methyl group on the reactivity ratios of HPMA-EMA and hydroxypropyl methacrylate-ethyl acrylate (HPMA-EA) [9]

TABLE 1. Copolymerization of 2-Hydroxypropyl Methacrylate (M_1) and Ethyl Methacrylate (M_2)^a

Initial monomer mixture		Copolymer composition			
Mole fraction of M_1	Mole fraction of M_2	Copolymer yield (% by wt of monomers)	% OH content	Mole fraction of m_1	Mole fraction of m_2
0.10	0.90	6.7	4.793	0.2599	0.7401
0.20	0.80	10.2	5.548	0.4124	0.5876
0.30	0.70	9.4	6.930	0.5295	0.4705
0.50	0.50	4.8	8.719	0.6910	0.3090
0.70	0.30	9.5	10.117	0.8259	0.1741
0.80	0.20	8.8	10.723	0.8868	0.1132

^a Polymerization conditions: Catalyst, benzoyl peroxide (0.2% by wt of monomers) temperature, $60 \pm 0.1^\circ\text{C}$.

gave an interesting observation. The $1/r_1$ values of HPMA-EMA and HPMA-EA are 0.55 and 0.08, respectively, indicating greater reactivity of ethyl methacrylate than of ethyl acrylate toward the poly(HPMA) radical. This may be due to the extra stability of the growing ethyl methacrylate radical resulting from hyperconjugation with respect to the α -methyl group in addition to conjugation. This is represented by



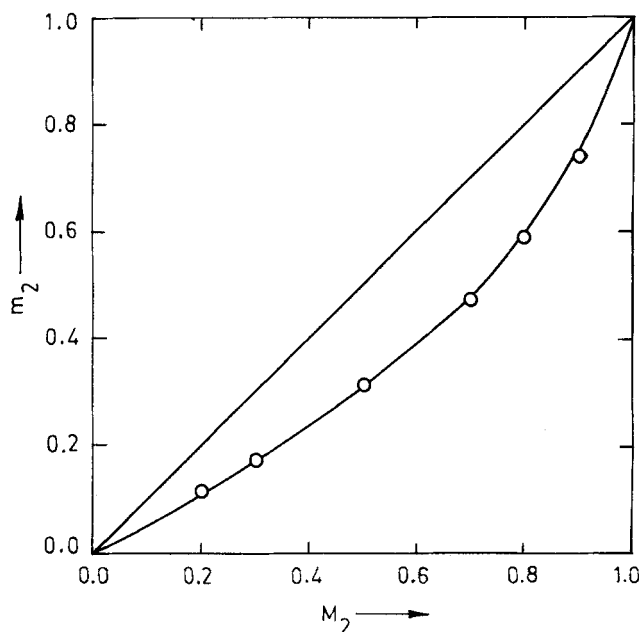


FIG. 1. Plot of mole fraction of EMA (m_2) in copolymer versus mole fraction of EMA (M_2) in monomer feed: (○) experimental values, (—) curve calculated from copolymer equation using reactivity ratios determined by the YBR method.

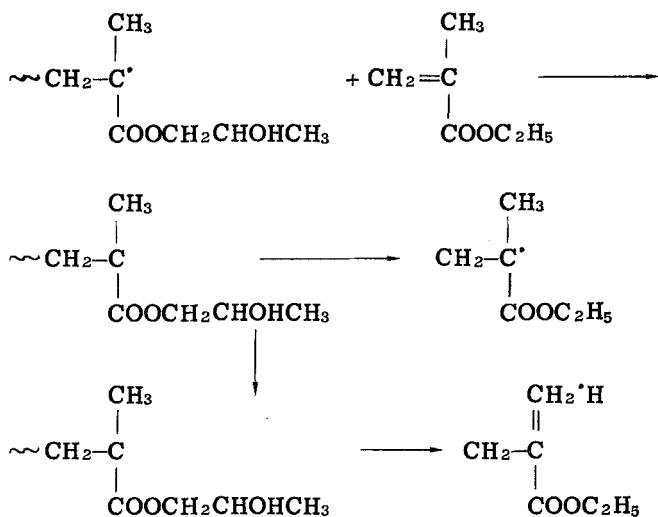


TABLE 2. Copolymerization of 2-Hydroxypropyl Methacrylate (M_1) and n-Butyl Methacrylate (M_2)^a

Initial monomer mixture		Copolymer composition			
Mole fraction of M_1	Mole fraction of M_2	Copolymer yield (% by wt of monomers)	% OH content	Mole fraction of m_1	Mole fraction of m_2
0.10	0.90	9.7	3.810	0.3197	0.6803
0.20	0.80	6.5	5.743	0.4831	0.5169
0.30	0.70	8.3	6.854	0.5772	0.4228
0.50	0.50	4.6	8.753	0.7388	0.2612
0.60	0.40	7.6	9.499	0.8025	0.1975

^a Polymerization conditions: Catalyst, benzoyl peroxide (0.2% by wt of monomers); temperature, $60 \pm 0.1^\circ\text{C}$.

Reactivity Ratios in Copolymerization of 2-Hydroxypropyl Methacrylate (M_1) with n-Butyl Methacrylate (M_2)

The conditions and results of experiments for the copolymerization of this pair of monomers are given in Table 2. The reactivity ratios calculated by using the YBR method were found to be

$$r_1 = 2.378 \pm 0.001$$

$$r_2 = 0.190 \pm 0.010$$

The copolymer composition curve, i.e., monomer feed vs copolymer feed, was calculated from r_1 and r_2 values and is shown in Fig. 2 along with experimental data. Since $r_1 \gg 1$ and $r_2 \ll 1$, the copolymer consists of larger blocks of 2-HPMA units interrupted with a few n-butyl methacrylate comonomer units.

Comparison of $1/r_1$ values of HPMA-BMA system with HPMA-butylacrylate (HPMA-BA) copolymer [9] reveals the higher reactivity of butyl methacrylate toward poly(HPMA) growing species. The $1/r_1$ values are

$$\text{HPMA-BMA} \quad 0.42$$

$$\text{HPMA-BA} \quad 0.19$$

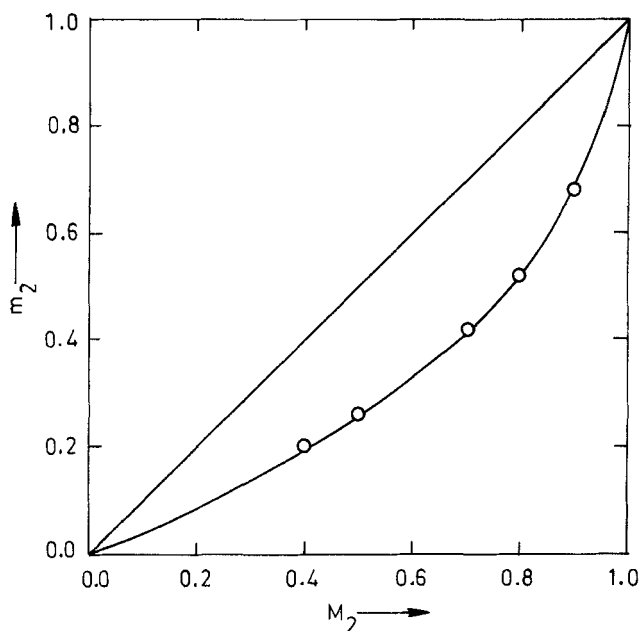


FIG. 2. Plot of mole fraction of BMA (m_2) in copolymer versus mole fraction of BMA (M_2) in monomer feed: (\circ) experimental values, (—) curve calculated from copolymer equation using reactivity ratios determined by the YBR method.

Further, the increase in reactivity of BMA as compared to BA toward poly(HPMA) radical is only twofold. On the other hand, the fact that the reactivity EMA is five times more than EA toward poly-HPMA radical clearly reveals that the former system is influenced by steric factors.

Reactivity Ratios in Copolymerization of 2-Hydroxypropyl Methacrylate (M_1) with 2-Ethylhexyl Methacrylate (M_2)

Table 3 shows the experimental conditions and results of the copolymerization of 2-HPMA. As above, the monomer reactivity ratios were found to be

$$r_1 = 4.370 \pm 0.048$$

$$r_2 = 0.103 \pm 0.006$$

See Fig. 3.

TABLE 3. Copolymerization of 2-Hydroxypropyl Methacrylate (M_1) and 2-Ethylhexyl Methacrylate (M_2)^a

Initial monomer mixture		Copolymer composition			
Mole fraction of M_1	Mole fraction of M_2	Copolymer yield (% by wt of monomers)	% OH content	Mole fraction of m_1	Mole fraction of m_2
0.10	0.90	6.0	3.379	0.4412	0.5588
0.20	0.80	9.1	4.344	0.5868	0.4132
0.30	0.70	6.4	5.004	0.6931	0.3069
0.40	0.60	6.5	5.481	0.7735	0.2265
0.50	0.50	4.6	5.816	0.8314	0.1686
0.60	0.40	10.4	6.076	0.8772	0.1228

^a Polymerization conditions: Catalyst, benzoyl peroxide (0.2% by wt of monomers); temperature, $60 \pm 0.1^\circ\text{C}$.

Because r_1 (HPMA) $\gg 1$ and r_2 (EHMA) $\ll 1$, the hydroxypropyl methacrylate is more reactive than EHMA toward both propagating species. The polymer, therefore, is richer in HPMA than in EHMA.

A perusal of the reactivity ratio data (Table 4) reveals that the reactivity of comonomer in copolymerization depends on steric factors and the polarity of the molecule. The reactivity of the comonomer is influenced by the size and the number of substituents attached to an unsaturated C=C linkage. As expected, the relative reactivity of alkyl methacrylate comonomer toward poly(HPMA) radical is of the following order: EMA > BMA > EHMA.

Although the ability of the α -methyl group to influence the polarization of an unsaturated bond is the same, the length and degree of branching of the alkyl radical exerts a certain influence on the reactivity of alkyl methacrylate in copolymerization with HPMA. The reactivity ratio of alkyl methacrylate is in accordance with the size of the alkyl group and is also higher than that of methyl methacrylate as reported by us earlier [9].

SUMMARY AND CONCLUSIONS

Based on our investigation, a different approach to obtain thermo-set acrylic polymers which can yield better mechanical properties has been reported. Thus 2-hydroxypropyl methacrylate has been

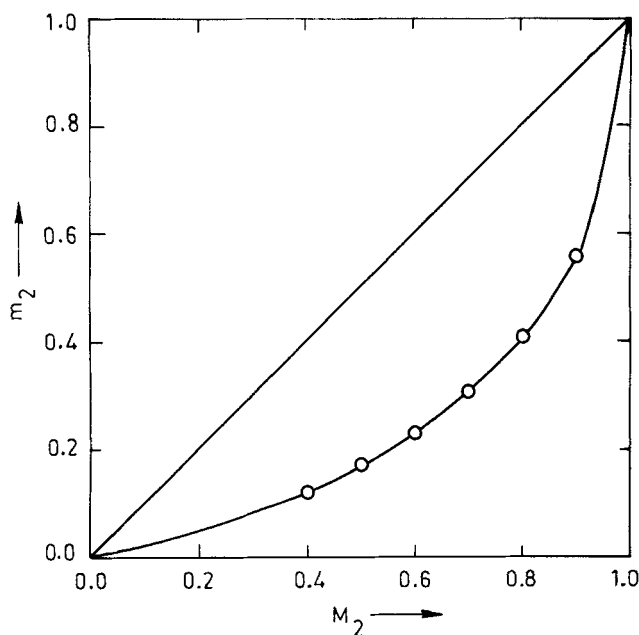


FIG. 3. Plot of mole fraction of 2-EHMA (m_2) in copolymer versus mole fraction of 2-EHMA (M_2) in monomer feed: (\circ) experimental values, (—) curve calculated from copolymer equation using reactivity ratios determined by the YBR method.

TABLE 4. Reactivity Ratios for Copolymerization of 2-Hydroxypropyl Methacrylate with Alkyl Methacrylates

Monomer M_1	Monomer M_2	$r_1 \pm \Delta r_1$	$r_2 \pm \Delta r_2$
2-Hydroxypropyl methacrylate	Ethyl methacrylate	1.807 ± 0.032	0.245 ± 0.021
	n-Butyl meth- acrylate	2.378 ± 0.001	0.19 ± 0.01
	2-Ethylhexyl methacrylate	4.370 ± 0.048	0.103 ± 0.006

copolymerized with alkyl methacrylates using free-radical initiators. Conversion was limited to less than 10% in all cases. A wide range of initial monomer feed ratio was chosen to give equal weightage to both the monomers as seen from the dependence curves.

The reactivity ratios for all the systems have been evaluated by using the YBR method. These values have been compared with the corresponding HPMA-alkyl acrylate copolymers reported earlier. Since the reactivity ratios indicate the distribution of monomer units in a polymer chain, this can serve to choose desired copolymer compositions for specific end use requirements. The copolymers selected could be used for cross-linking by using suitable cross-linking agents such as aliphatic and aromatic diisocyanates to obtain good mechanical properties. Our investigations on these aspects are under way and will be published.

REFERENCES

- [1] R. H. Yocum and E. B. Nyquist, Functional Monomers, Vol. 1 Dekker, New York, 1973, pp. 328-337.
- [2] J. D. Andrade, Med. Instrum., **1**, 110 (1973).
- [3] Henkel and Cie, British Patent 1,122,128 (1968).
- [4] G. N. Babu and A. Deshpande, To Be Published.
- [5] A. I. Vogel, Elementary Practical Organic Chemistry, Part III, Longmans, London, 1971, p. 677.
- [6] A. I. Yezrielev, E. L. Brokhina, and Y. S. Raskin, Vysokomol. Soedin., **A11**, 1670 (1969).
- [7] R. M. Joshi, J. Macromol. Sci.-Chem., **A7**, 1231 (1973).
- [8] P. W. Tidwell and G. A. Mortimer, J. Polym. Sci., **A3**, 369 (1965).
- [9] G. N. Babu, S. F. Xavioir and T. C. Goel, J. Polym. Sci., Polym. Chem. Ed., **15**, 1473 (1977).

Accepted by editor March 13, 1981

Received for publication April 6, 1981