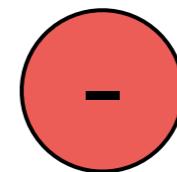
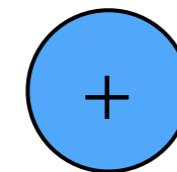


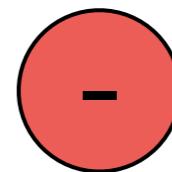
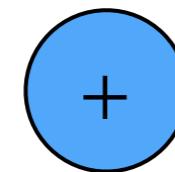
Opposites attract and likes repel!

Anybody who's married knows this is true!



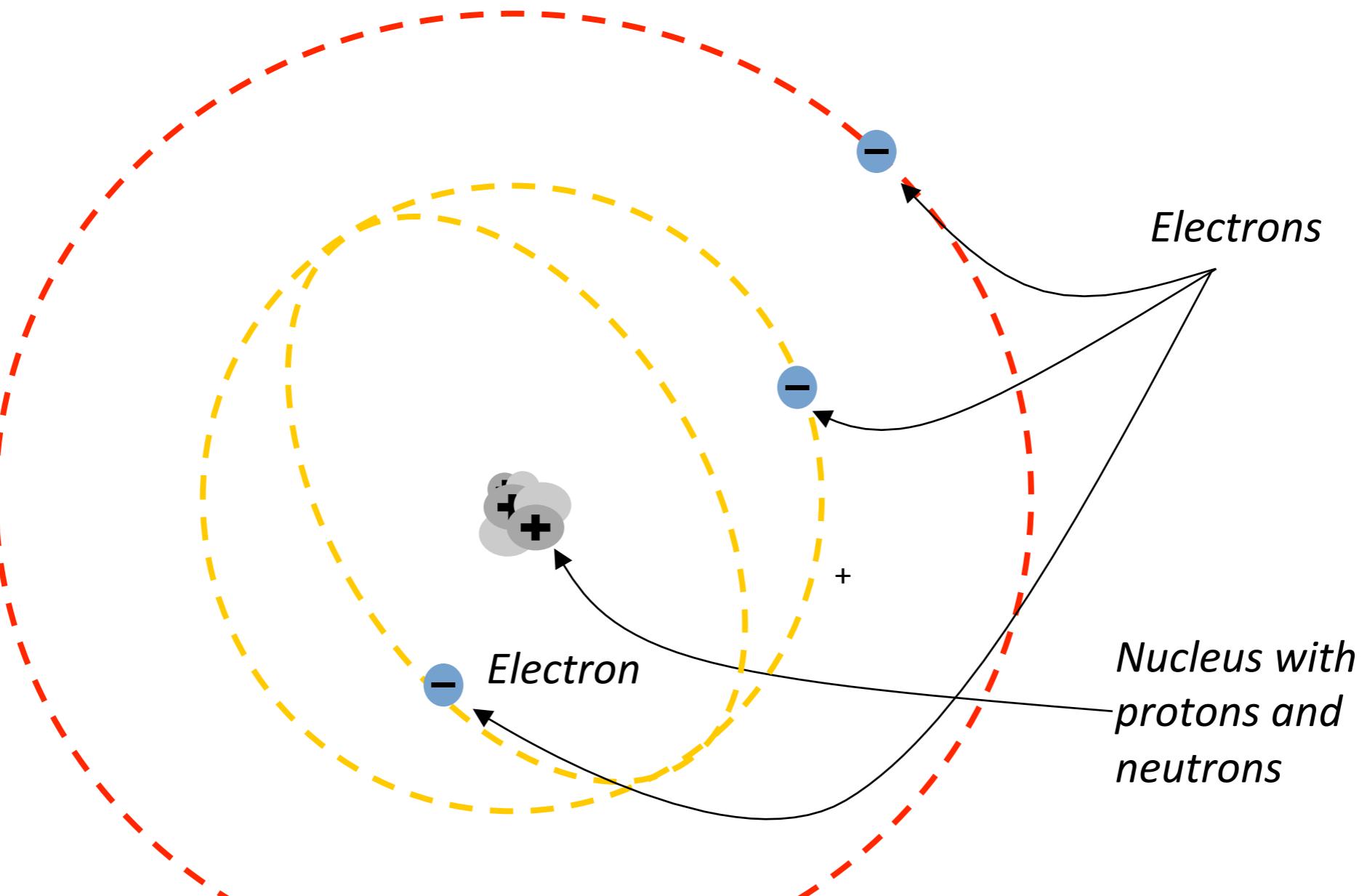
Opposites attract and likes repel!

Anybody who's married knows this is true!



**TABLE 25.1** Protons and electrons

Particle	Mass (kg)	Charge
Proton	$1.67 \times 10^{-27}$	$+e$
Electron	$9.11 \times 10^{-31}$	$-e$

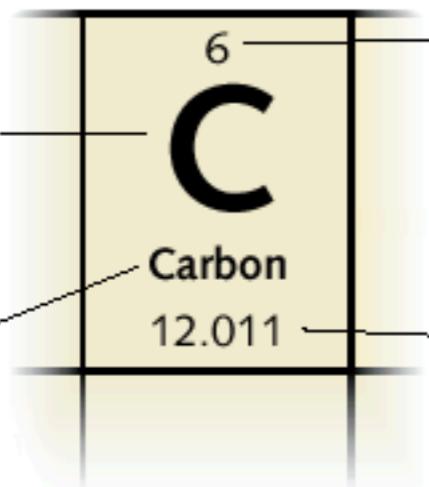


*Particles not drawn to scale*

# Reading the periodic table

## Symbol

A one- or two-letter abbreviation derived from the element's English or Latin name.

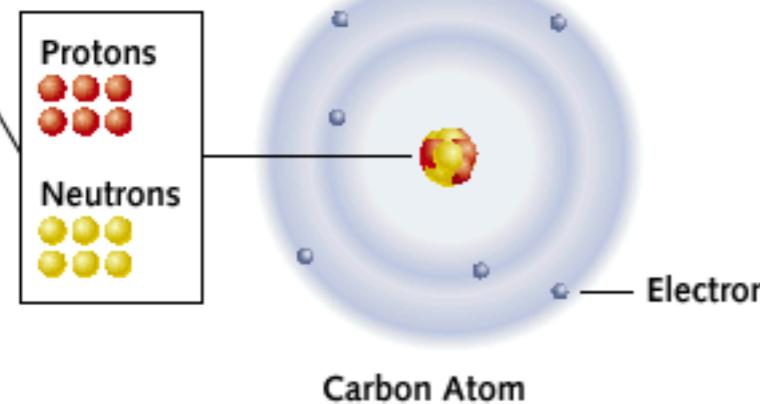


## Name

Element's common name.

## Mass Number

The sum of the numbers of protons and neutrons in a specific isotope.



# Reading the periodic table

## Symbol

A one- or two-letter abbreviation derived from the element's English or Latin name.

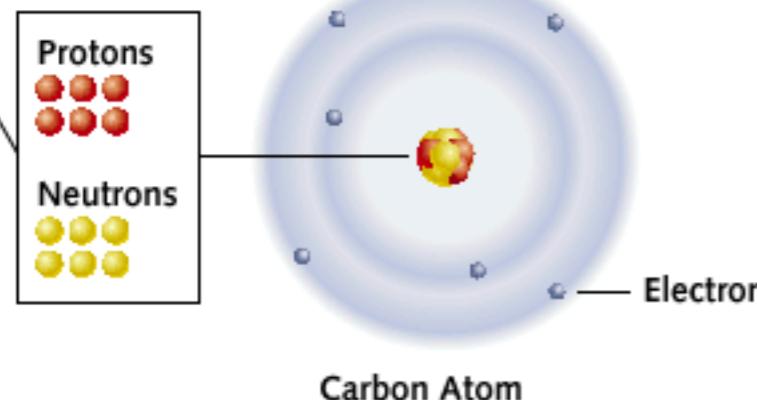
6	C	Atomic Number
Carbon		Equal to the number of protons in the nucleus, as well as the number of electrons in the electron cloud.
12.011		

## Name

Element's common name.

## Mass Number

The sum of the numbers of protons and neutrons in a specific isotope.



# Reading the periodic table

## Symbol

A one- or two-letter abbreviation derived from the element's English or Latin name.

6	C
Carbon	
12.011	

## Atomic Number

Equal to the number of protons in the nucleus, as well as the number of electrons in the electron cloud.

## Name

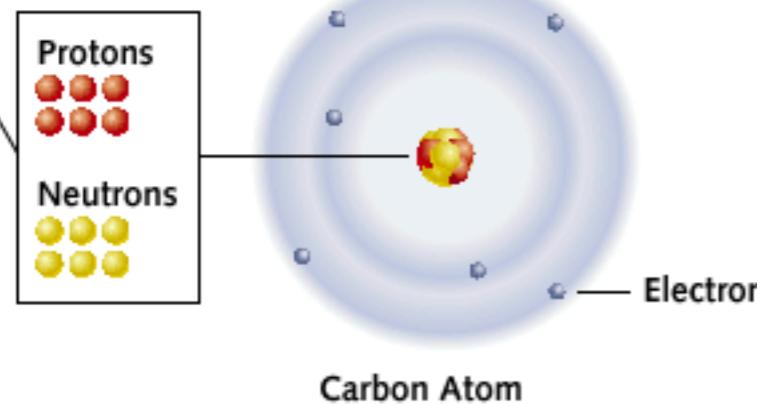
Element's common name.

## Mass Number

The sum of the numbers of protons and neutrons in a specific isotope.

## Atomic Mass

Weighted average of the masses of all the element's isotopes.  
Rounding the atomic mass to the nearest whole number yields the mass number of the most common isotope.



# Reading the periodic table

## Symbol

A one- or two-letter abbreviation derived from the element's English or Latin name.

6	C
Carbon	
12.011	

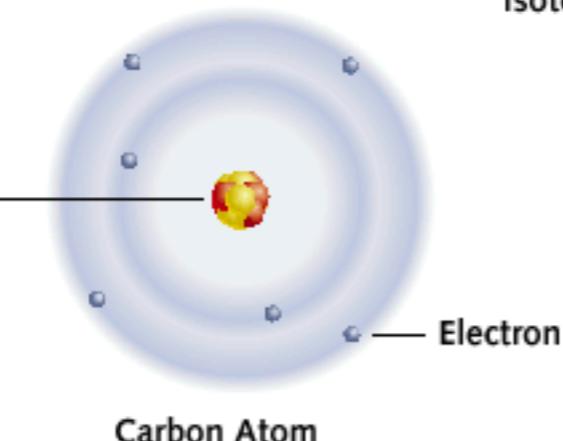
## Name

Element's common name.

## Mass Number

The sum of the numbers of protons and neutrons in a specific isotope.

Protons	
Neutrons	



## Atomic Number

Equal to the number of protons in nucleus, as well as the number of electrons in the electron cloud.

## Atomic Mass

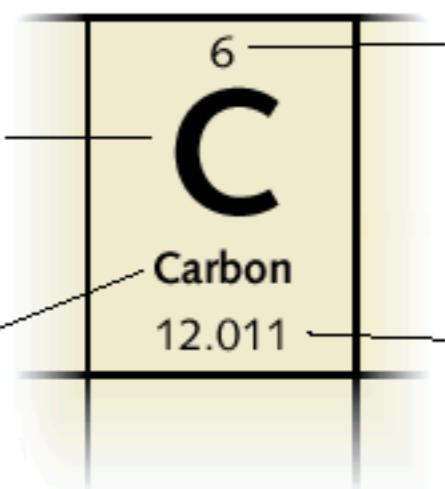
Weighted average of the masses of all the element's isotopes. Rounding the atomic mass to the nearest whole number yields the mass number of the most common isotope.

13	14	15	16	17	18
5	6	7	8	9	2
Boron 10.81	Carbon 12.01	Nitrogen 14.01	Oxygen 16.00	Fluorine 19.00	Helium 4.00
13	14	15	16	17	10
Aluminum 26.98	Silicon 28.09	Phosphorus 30.97	Sulfur 32.06	Chlorine 35.45	Neon 20.18
31	32	33	34	35	18
Gallium 69.73	Germanium 72.61	Arsenic 74.92	Selenium 78.09	Bromine 79.90	Argon 39.95
49	50	51	52	53	54
Inium 114.82	Tin 118.71	Antimony 121.76	Tellurium 127.6	Iodine 126.90	Xenon 131.29
81	82	83	84	85	86
Thallium 204.38	Lead 207.20	Bismuth 208.98	Polonium [208.98]	Astatine 209.98	Radon 222.02
113	114	115	116	117	118
Ununtrium unknown	Flerovium [289]	Ununpentium unknown	Livermorium [298]	Ununseptium unknown	Ununoctium unknown
6	67	69	70	71	

# Reading the periodic table

## Symbol

A one- or two-letter abbreviation derived from the element's English or Latin name.

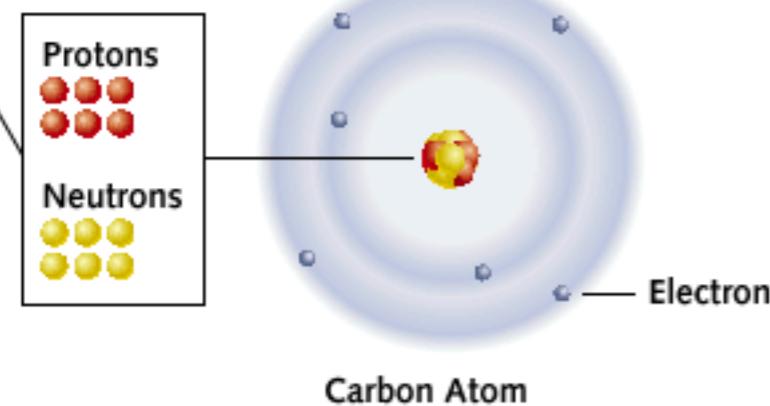


## Name

Element's common name.

## Mass Number

The sum of the numbers of protons and neutrons in a specific isotope.



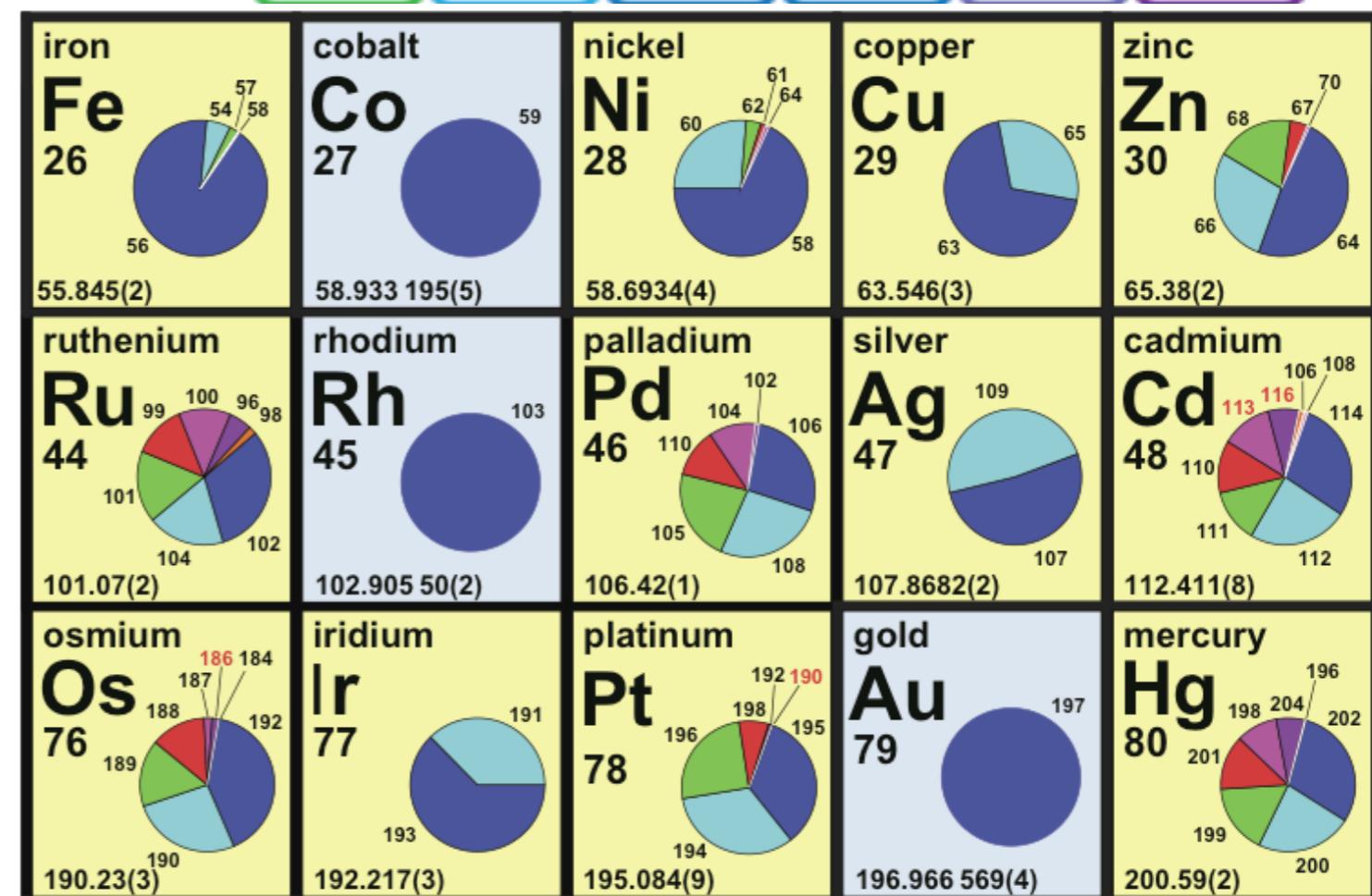
## Atomic Number

Equal to the number of protons in nucleus, as well as the number of electrons in the electron cloud.

## Atomic Mass

Weighted average of the masses of all the element's isotopes. Rounding the atomic mass to the nearest whole number yields the mass number of the most common isotope.

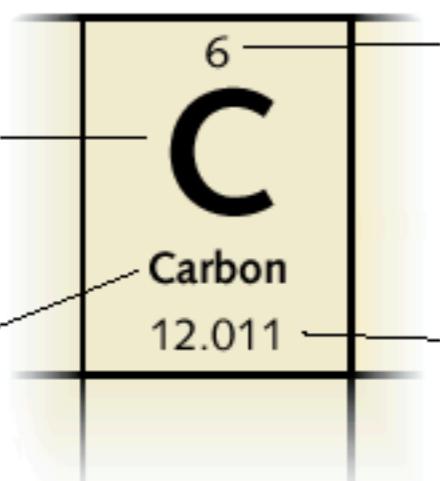
13	14	15	16	17	18
5	6	7	8	9	2
Boron	Carbon	Nitrogen	Oxygen	Fluorine	Helium
10.81	12.01	14.01	16.00	19.00	4.00
13	14	15	16	17	10
Aluminum	Silicon	Phosphorus	Sulfur	Chlorine	Neon
26.98	28.09	30.97	32.06	35.45	20.18
18					
Ar					
39.95					



# Reading the Periodic Table

## Symbol

A one- or two-letter abbreviation derived from the element's English or Latin name.



## Name

Element's common name.

## Mass Number

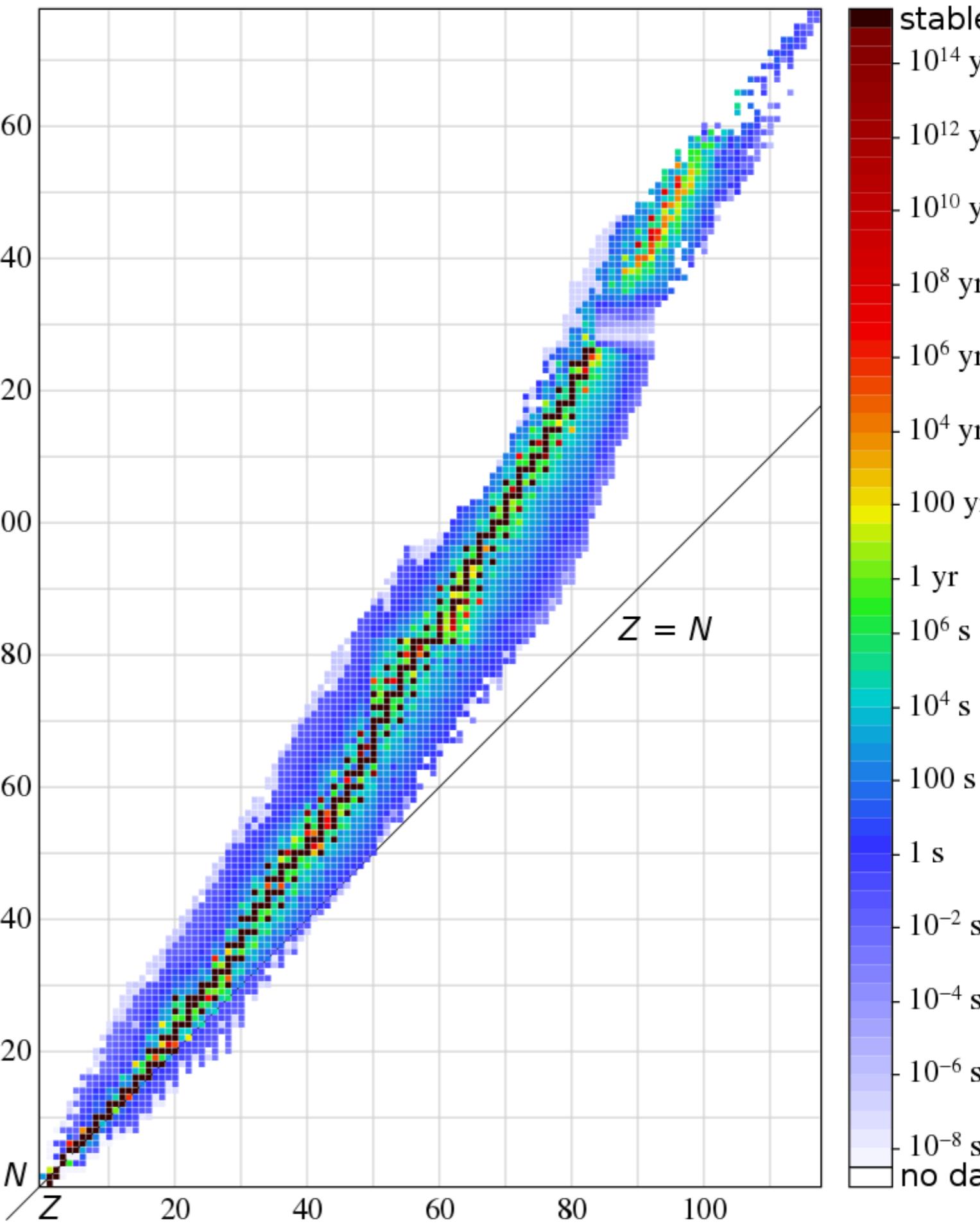
The sum of the numbers of protons and neutrons in a specific isotope.

## Atomic Number

Equal to the number of protons in the nucleus, as well as the number of electrons in the atom.

## Atomic Mass

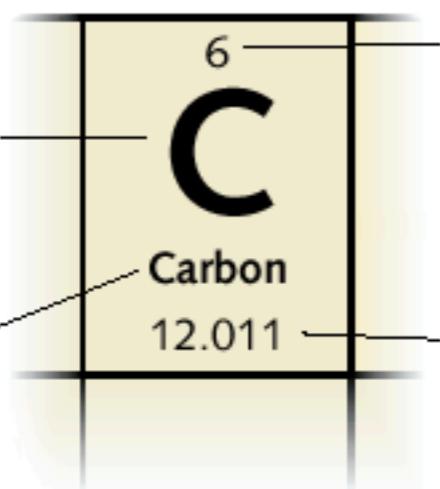
Weighted average of all the elements in the periodic table. Rounding the nearest whole number to the mass number of a specific isotope.



# Reading 1

## Symbol

A one- or two-letter abbreviation derived from the element's English or Latin name.

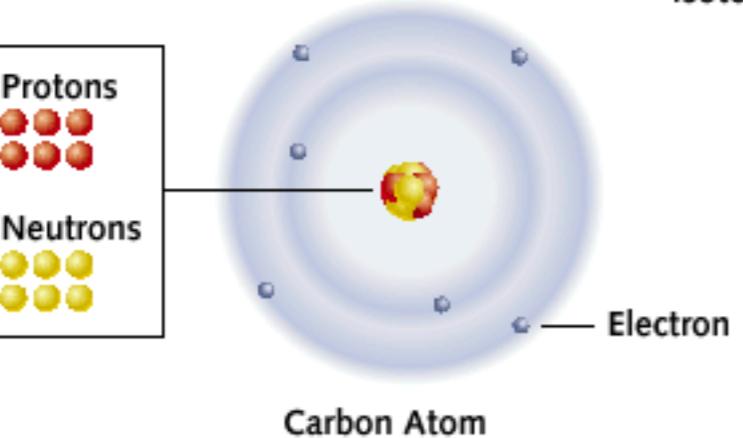


## Name

Element's common name.

## Mass Number

The sum of the numbers of protons and neutrons in a specific isotope.



12.011 amu

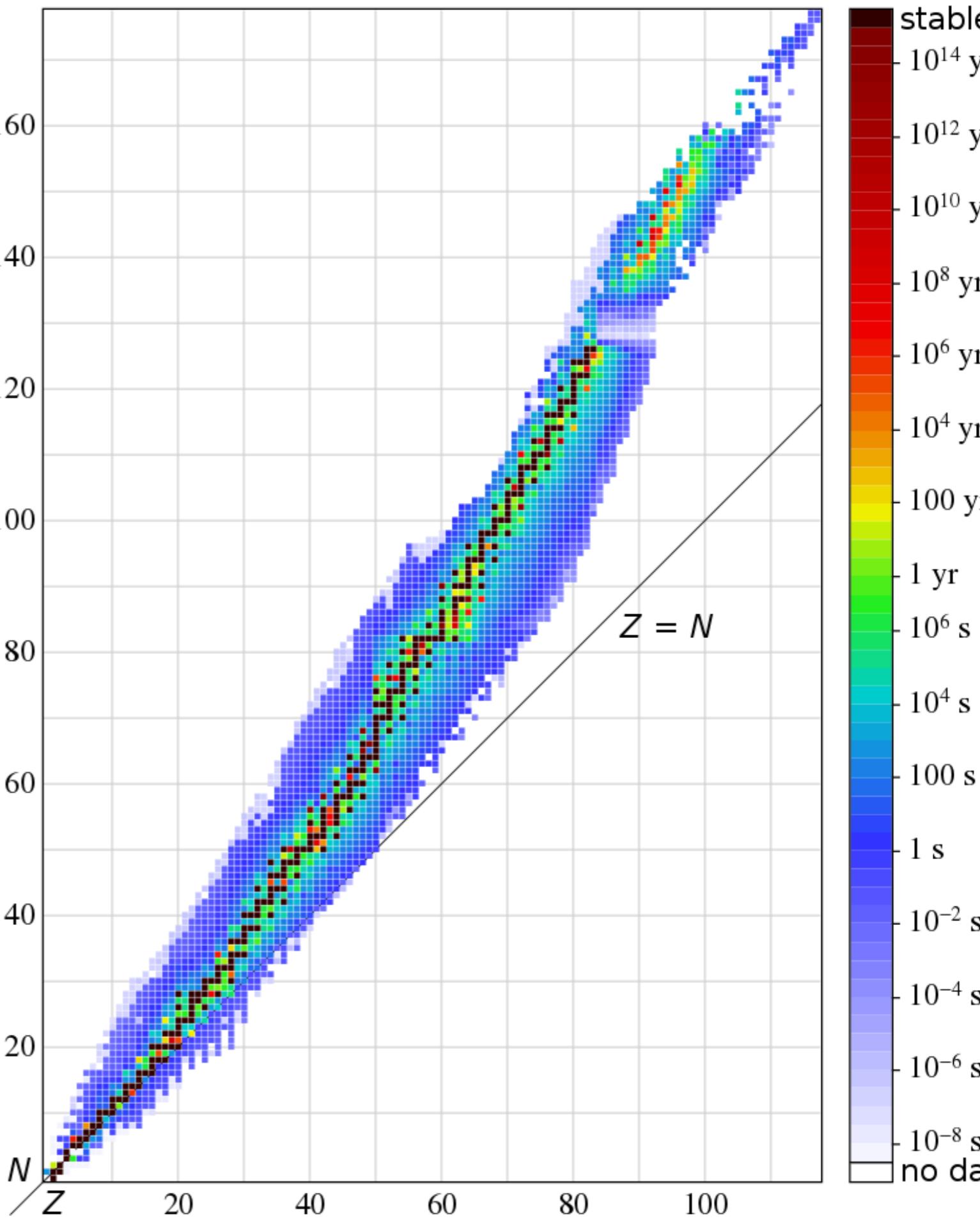
$$1 \text{ amu} = 1.6605 \times 10^{-27} \text{ kg}$$

## Atomic Number

Equal to the number of protons in the nucleus, as well as the number of electrons in the neutral atom.

## Atomic Mass

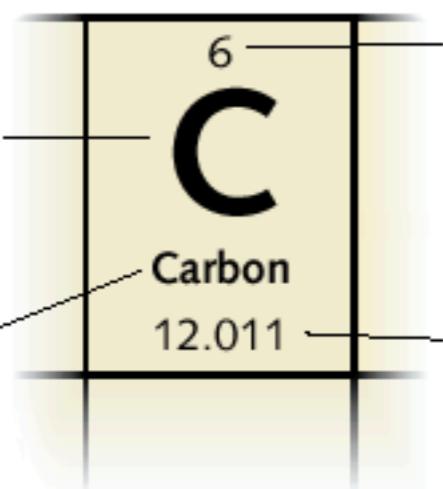
Weighted average of all the elements in the periodic table. Rounding the nearest whole number to the mass number of a specific isotope.



# Reading the Periodic Table

## Symbol

A one- or two-letter abbreviation derived from the element's English or Latin name.



## Name

Element's common name.

## Mass Number

The sum of the numbers of protons and neutrons in a specific isotope.

## Atomic Number

Equal to the number of protons in the nucleus, as well as the number of electrons in the atom.

## Atomic Mass

Weighted average mass of all the elements in nature. Rounding the mass number to the nearest whole number gives the mass number of the most abundant isotope.

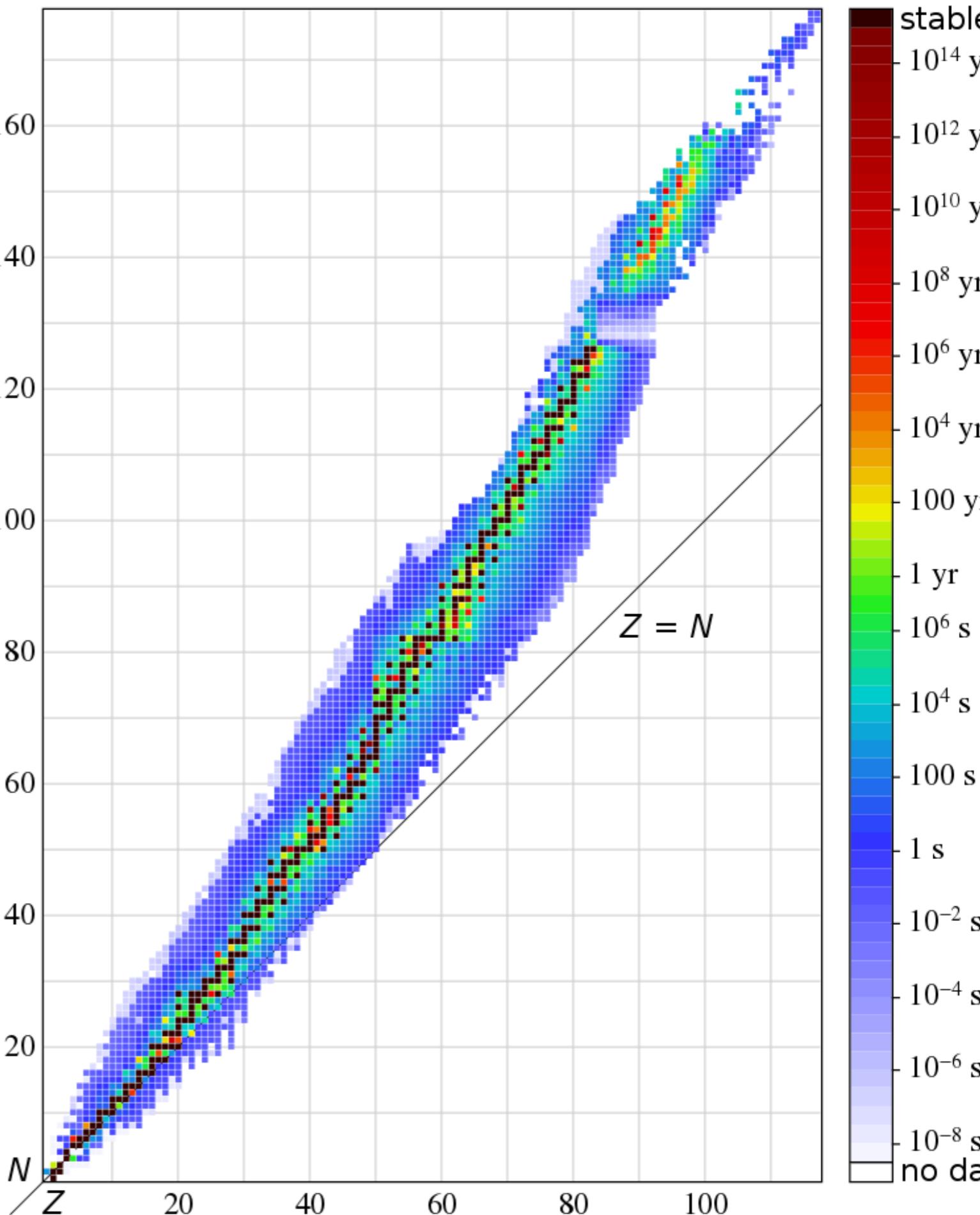
12.011 amu

$1 \text{ amu} = 1.6605 \times 10^{-27} \text{ kg}$

Or

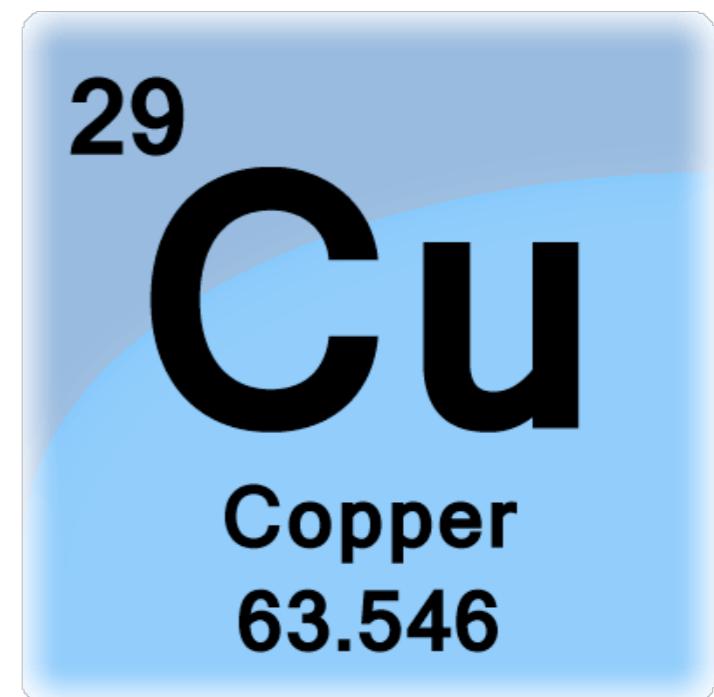
12.011 grams/mol

$1 \text{ mol} = 6.022 \times 10^{23} \text{ atoms}$



Pennies today are copper-plated zinc, but older pennies are 3.1 g of pure copper. What are the total positive charge and total negative charge in a solid copper penny that is electrically neutral?

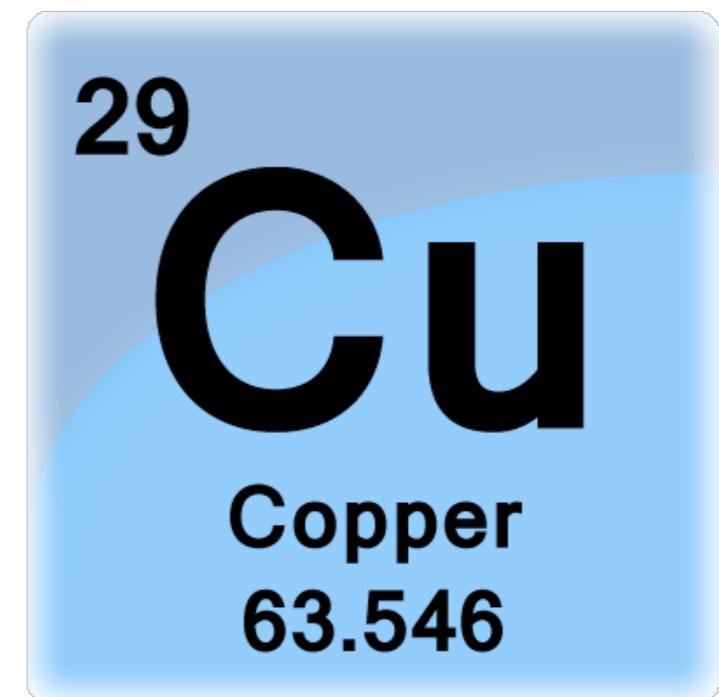
$$1 \text{ amu} = 1.6605 \times 10^{-27} \text{ kg}$$



Pennies today are copper-plated zinc, but older pennies are 3.1 g of pure copper. What are the total positive charge and total negative charge in a solid copper penny that is electrically neutral?

$$1 \text{ amu} = 1.6605 \times 10^{-27} \text{ kg}$$

$$N_{\text{atoms}} = \frac{.0031 \text{ kg}}{63.546 \times 1.6605 \times 10^{-27} \text{ kg}} = 2.9378 \times 10^{22}$$

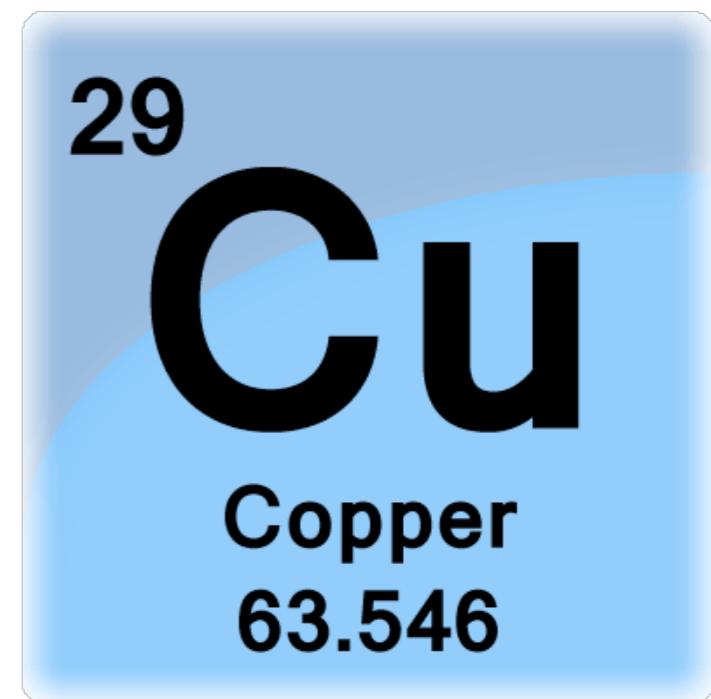


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$$1 \text{ amu} = 1.6605 \times 10^{-27} \text{ kg}$$

$$N_{\text{atoms}} = \frac{.0031 \text{ kg}}{63.546 \times 1.6605 \times 10^{-27} \text{ kg}} = 2.9378 \times 10^{22}$$

$$N_{\text{atoms}} = \frac{3.1 \text{ g/mol}}{63.546 \times 6.022 \times 10^{23} \text{ particles/mol}} = 2.9378 \times 10^{22}$$

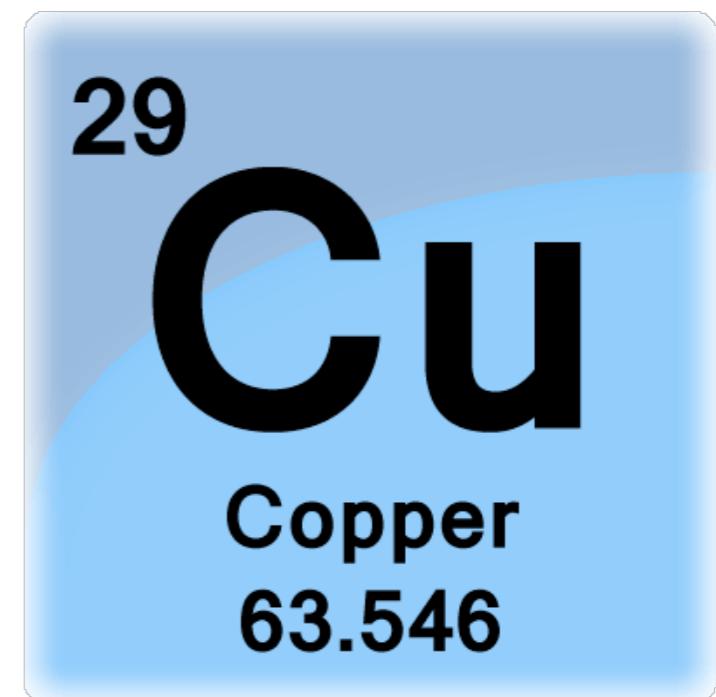


Pennies today are copper-plated zinc, but older pennies are 3.1 g of pure copper. What are the total positive charge and total negative charge in a solid copper penny that is electrically neutral?

$$1 \text{ amu} = 1.6605 \times 10^{-27} \text{ kg}$$

$$N_{\text{atoms}} = \frac{.0031 \text{ kg}}{63.546 \times 1.6605 \times 10^{-27} \text{ kg}} = 2.9378 \times 10^{22}$$

$$N_{\text{atoms}} = \frac{3.1 \text{ g/mol}}{63.546 \times 6.022 \times 10^{23} \text{ particles/mol}} = 2.9378 \times 10^{22}$$



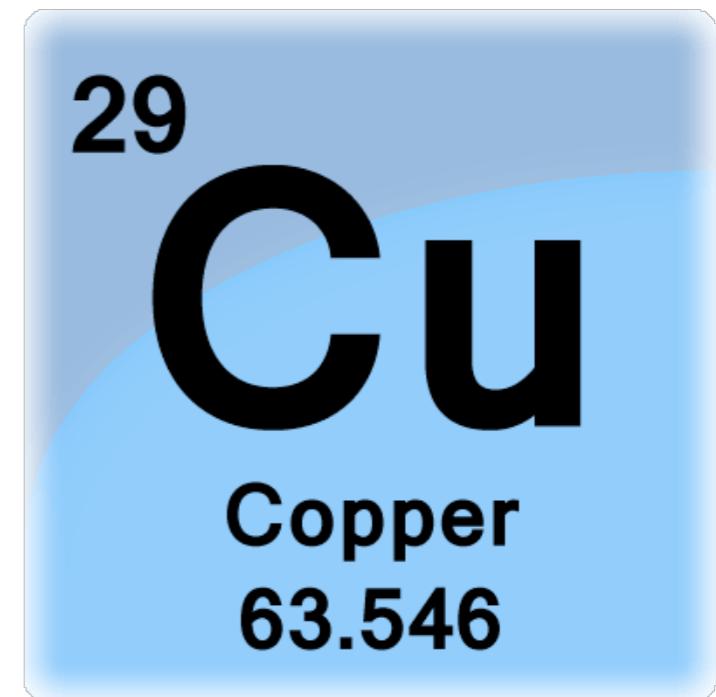
$$q = N_{\text{atoms}} \times 29 \times 1.602 \times 10^{-19} \text{ C} = 1.4 \times 10^5 \text{ C}$$

Pennies today are copper-plated zinc, but older pennies are 3.1 g of pure copper. What are the total positive charge and total negative charge in a solid copper penny that is electrically neutral?

$$1 \text{ amu} = 1.6605 \times 10^{-27} \text{ kg}$$

$$N_{\text{atoms}} = \frac{.0031 \text{ kg}}{63.546 \times 1.6605 \times 10^{-27} \text{ kg}} = 2.9378 \times 10^{22}$$

$$N_{\text{atoms}} = \frac{3.1 \text{ g/mol}}{63.546 \times 6.022 \times 10^{23} \text{ particles/mol}} = 2.9378 \times 10^{22}$$



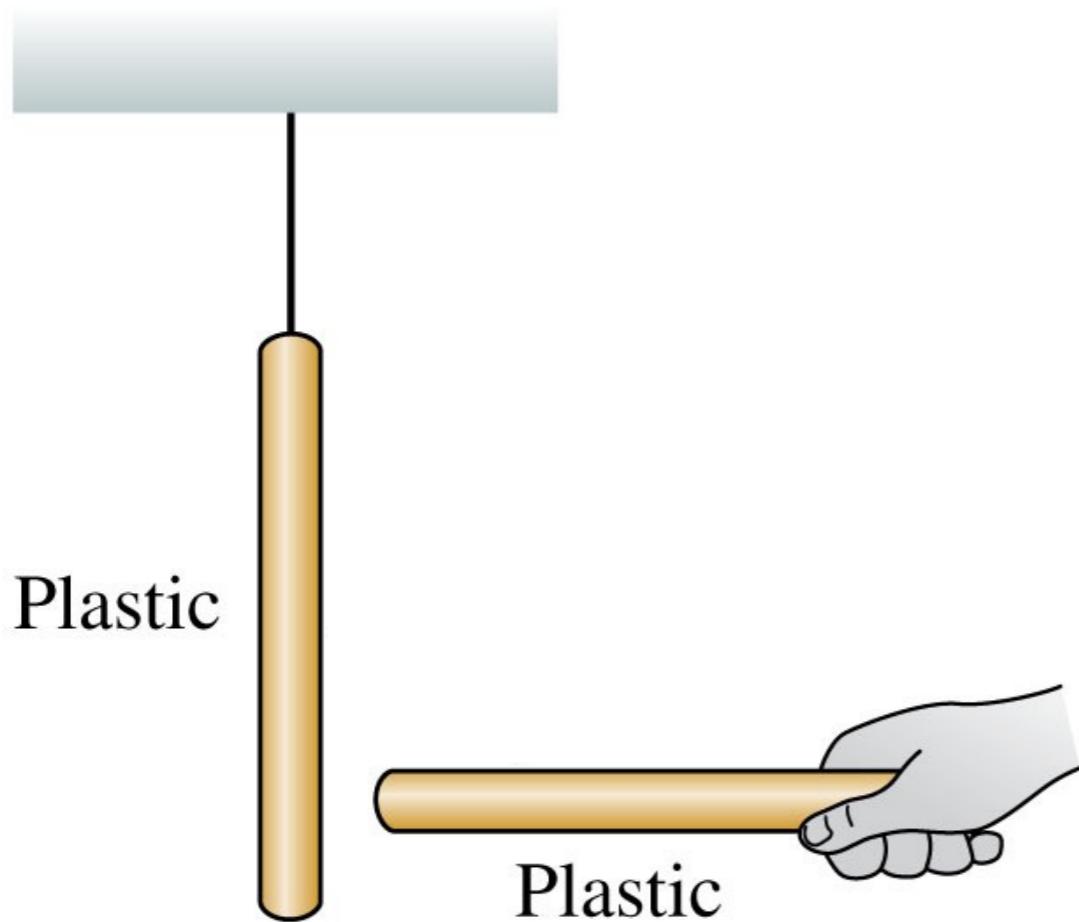
$$q = N_{\text{atoms}} \times 29 \times 1.602 \times 10^{-19} \text{ C} = 1.4 \times 10^5 \text{ C}$$

How much charge is a lot of charge?

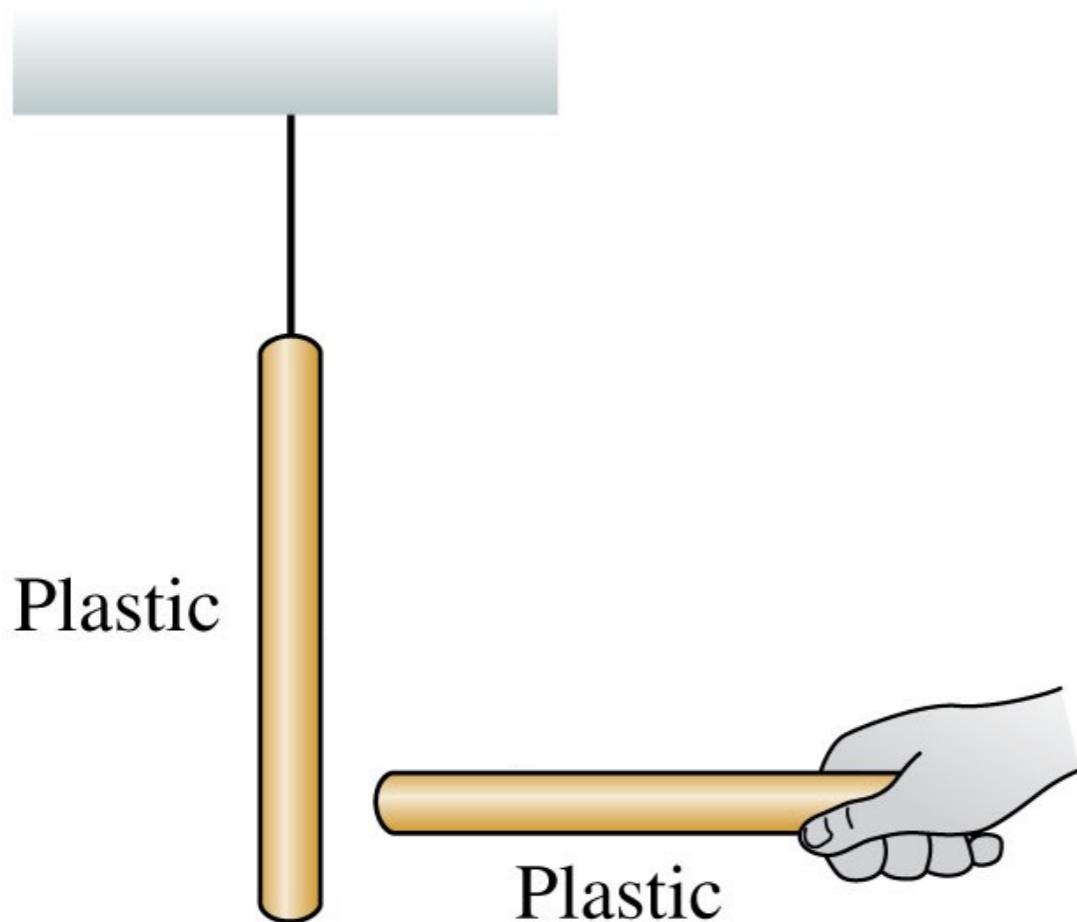
A typical lightning strike releases 10-20 C of charge.

Charge and Voltage are not the same thing.

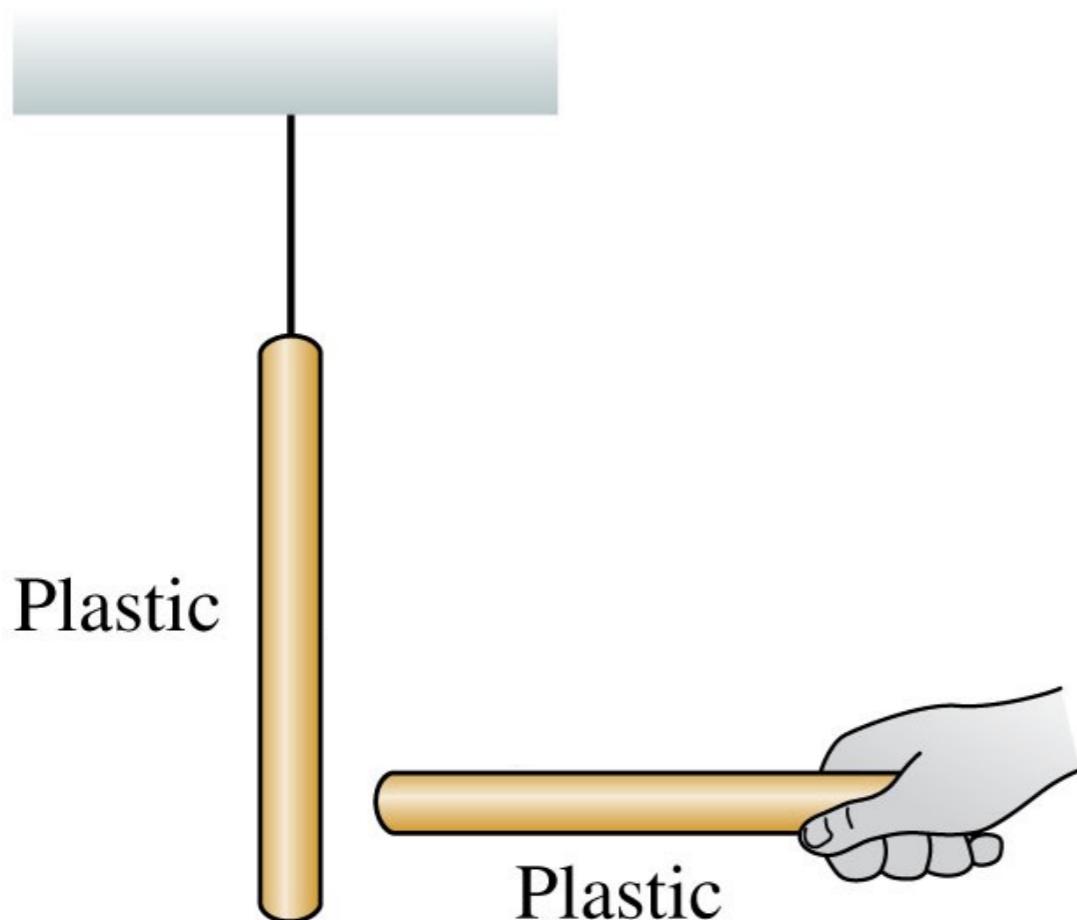
Do these plastic bars have charge?



Do these plastic bars have charge?  
Why don't they attract/repel one another?

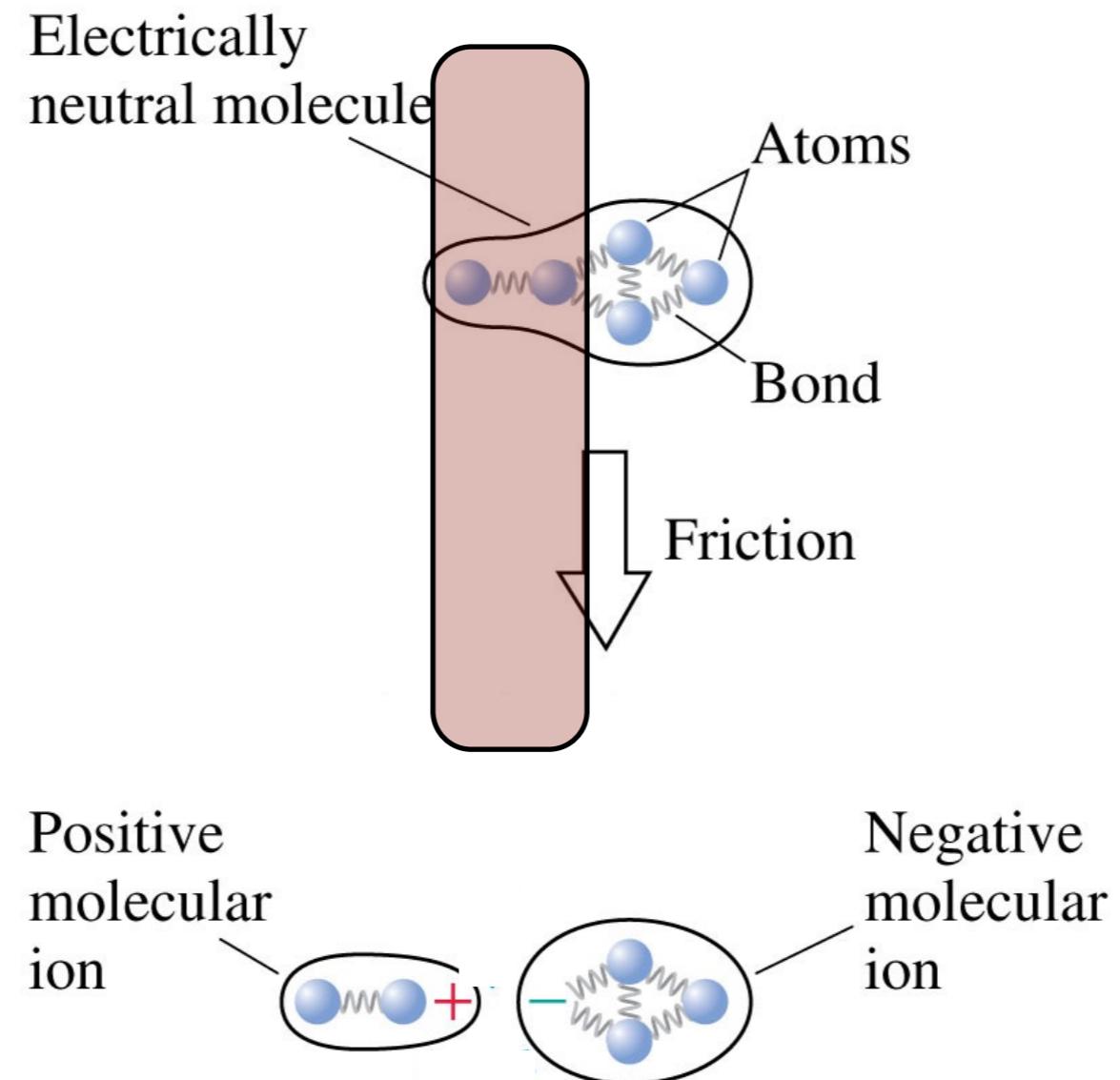


Do these plastic bars have charge?  
Why don't they attract/repel one another?

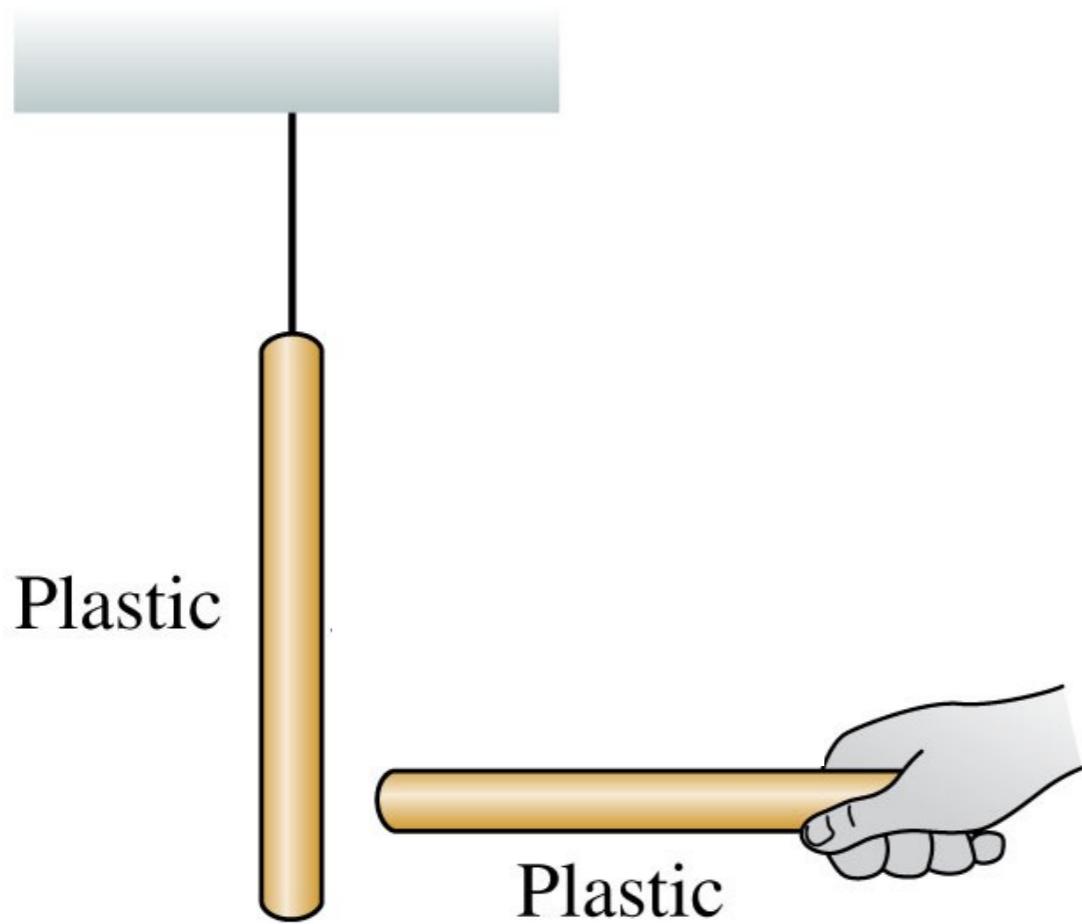


Outstanding question: What does rubbing do?

# Triboelectric Effect

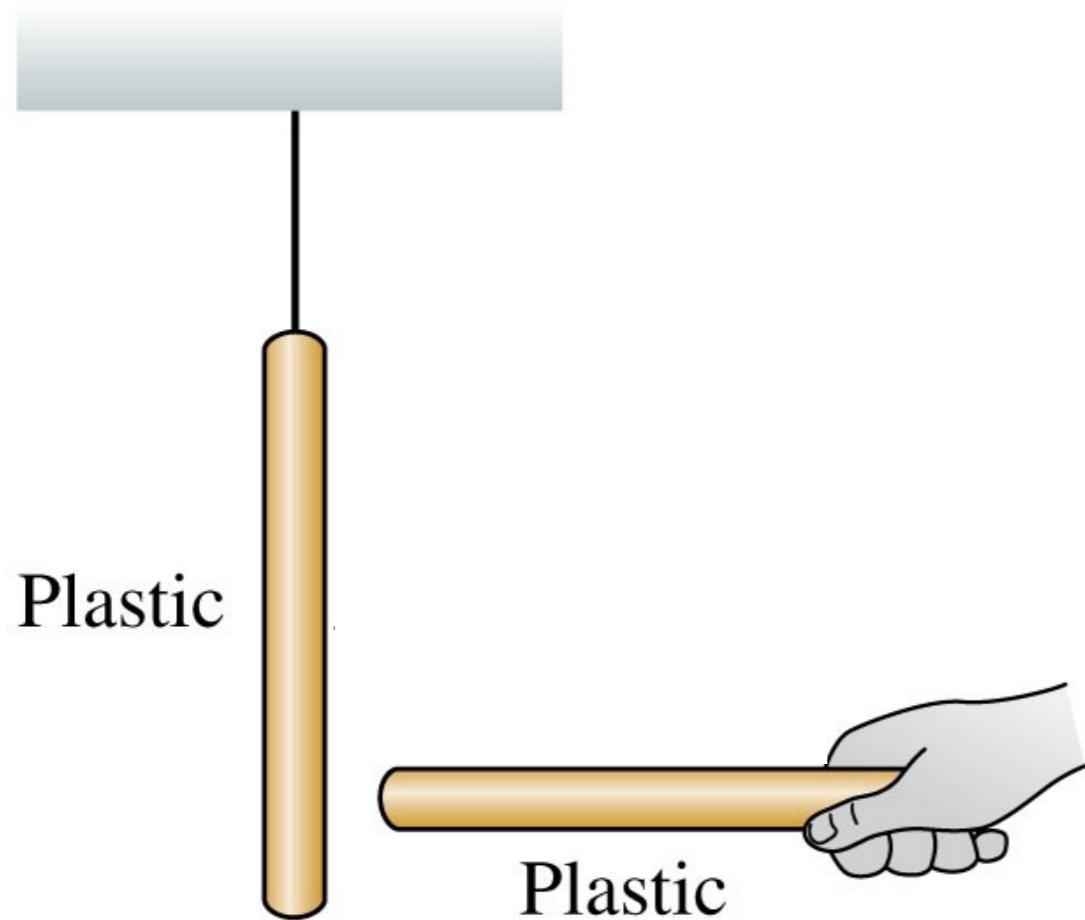


What if I only rub one of the rods?



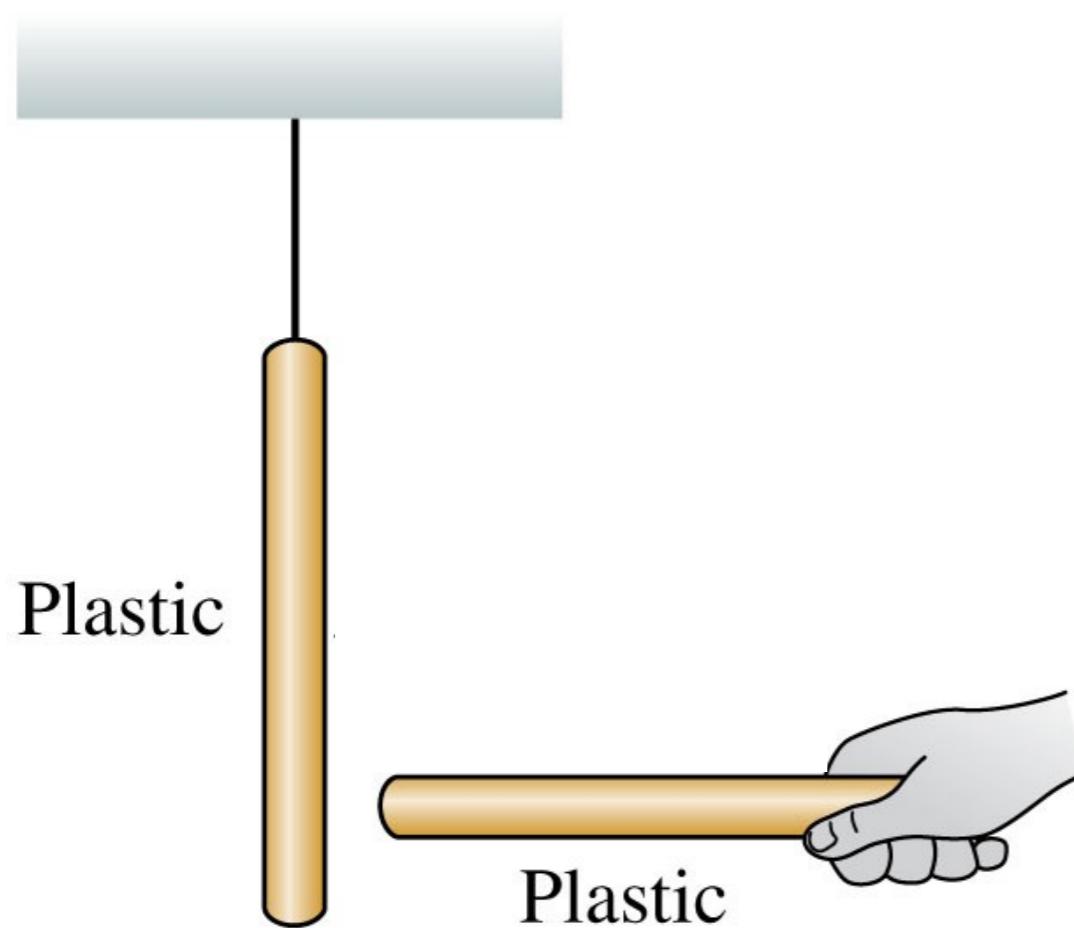
What if I only rub one of the rods?

What if I exchange the plastic rod for a glass rod?

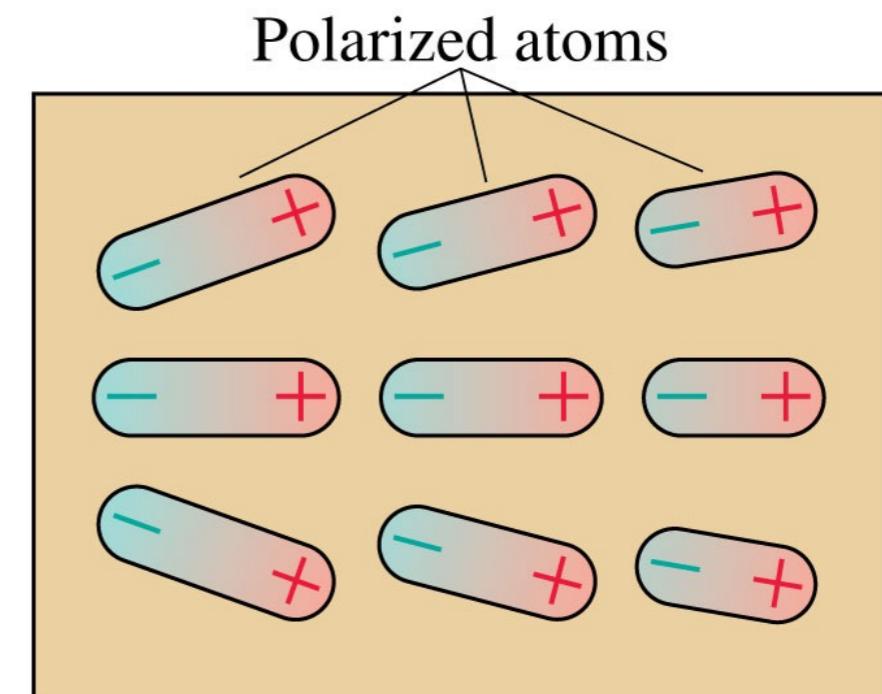


What if I only rub one of the rods?

What if I exchange the plastic rod for a glass rod?



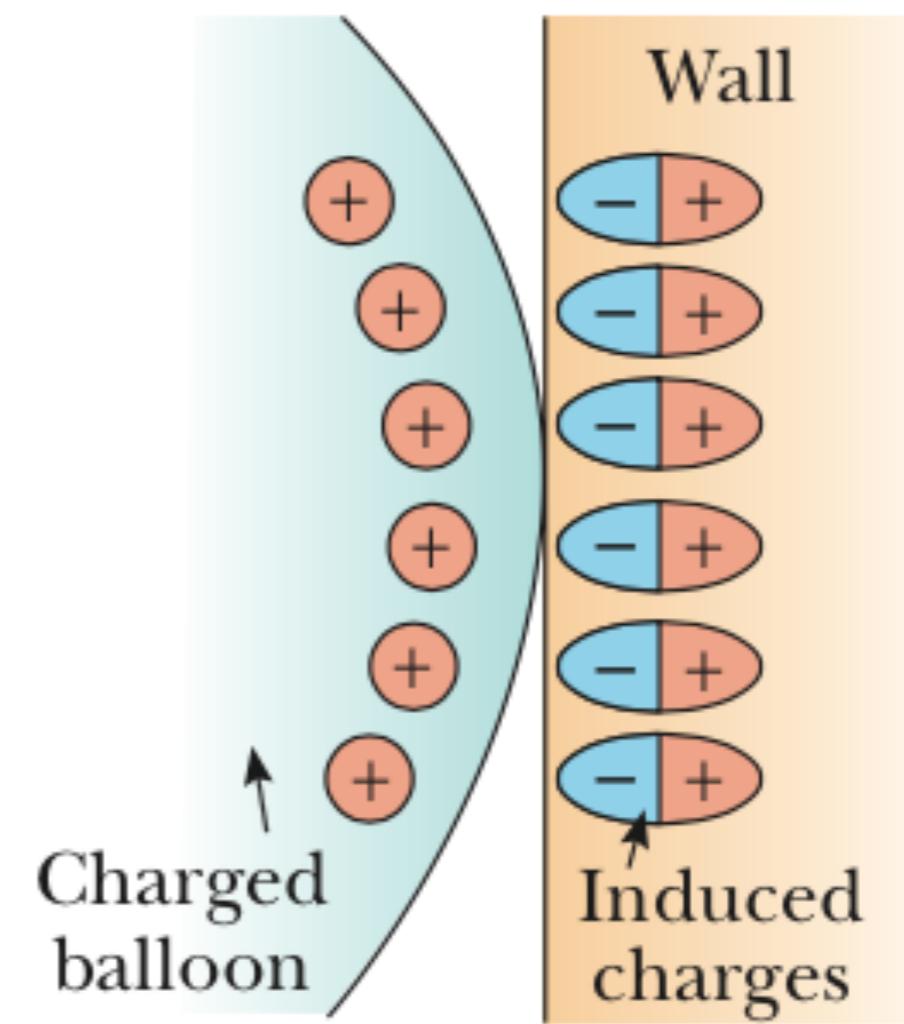
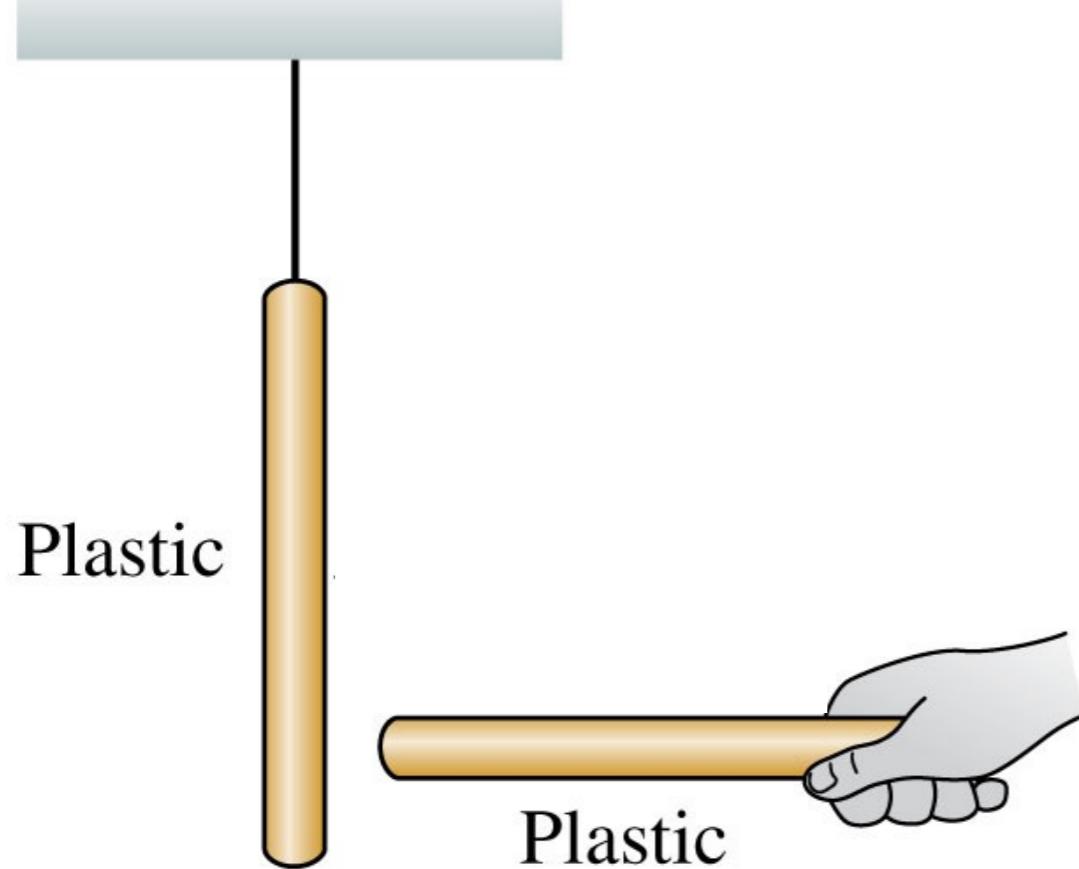
+  
External  
charge



←  
Net force

What if I only rub one of the rods?

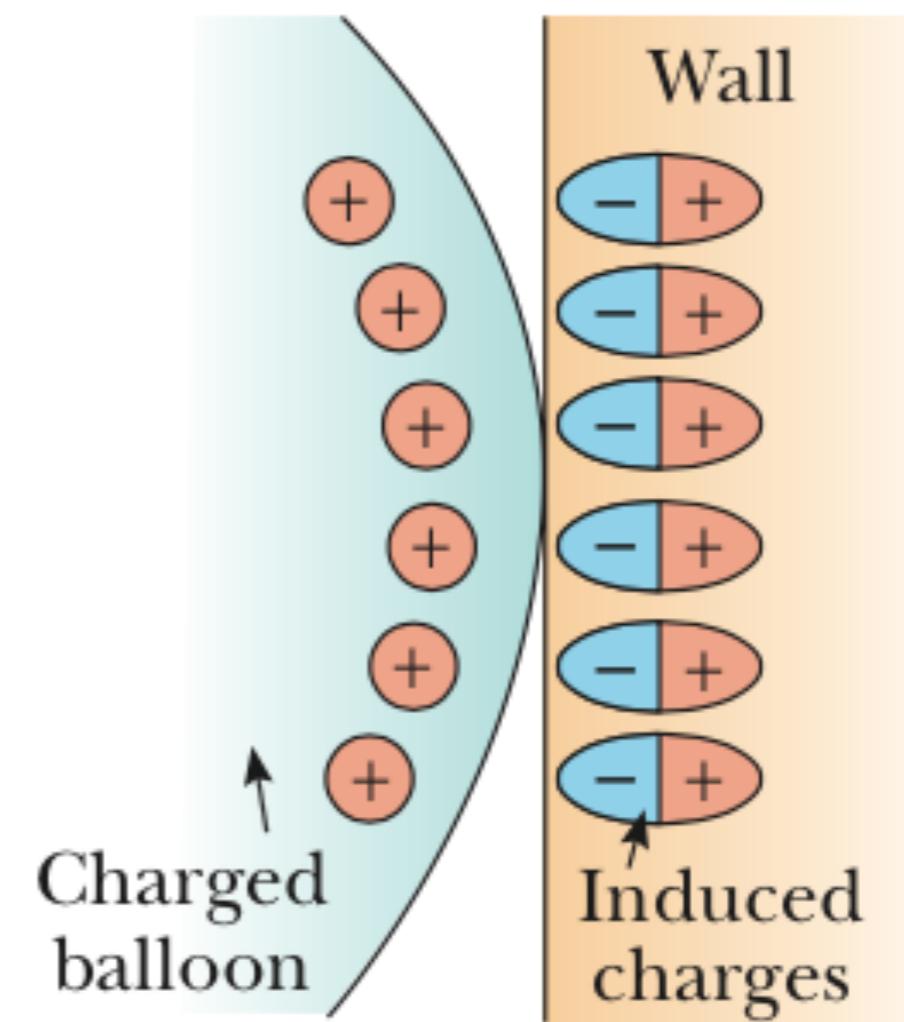
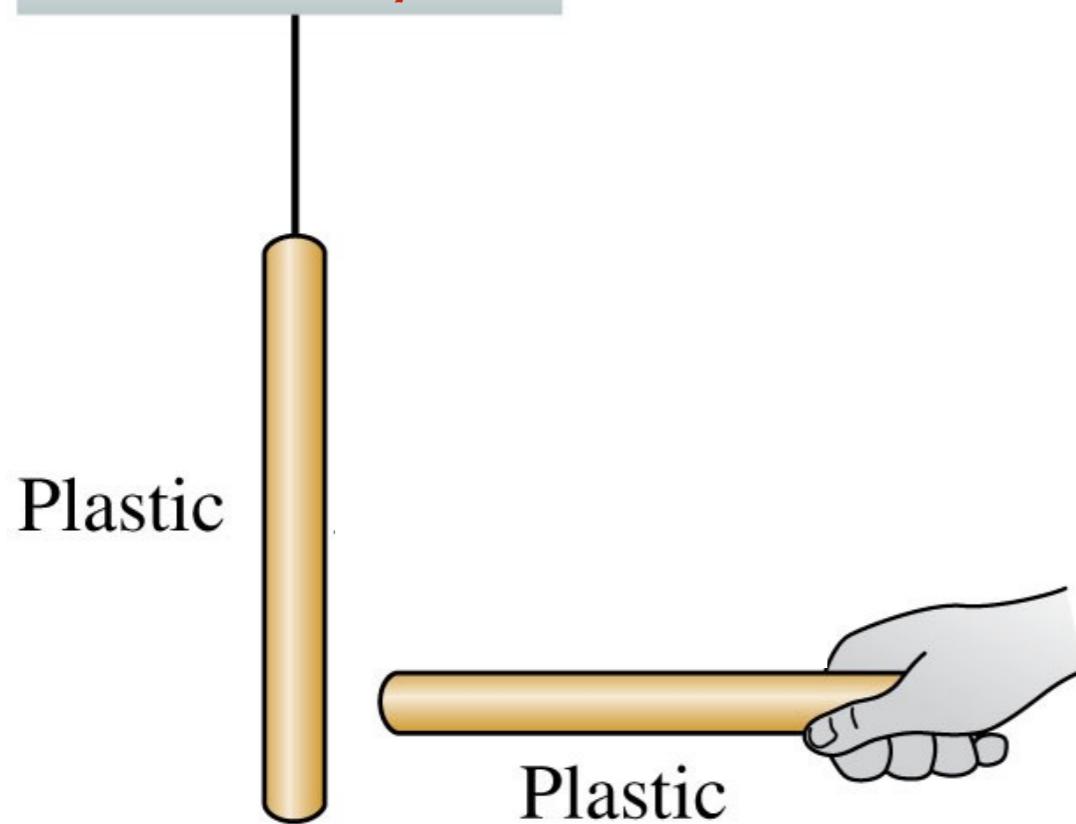
What if I exchange the plastic rod for a glass rod?



What if I only rub one of the rods?

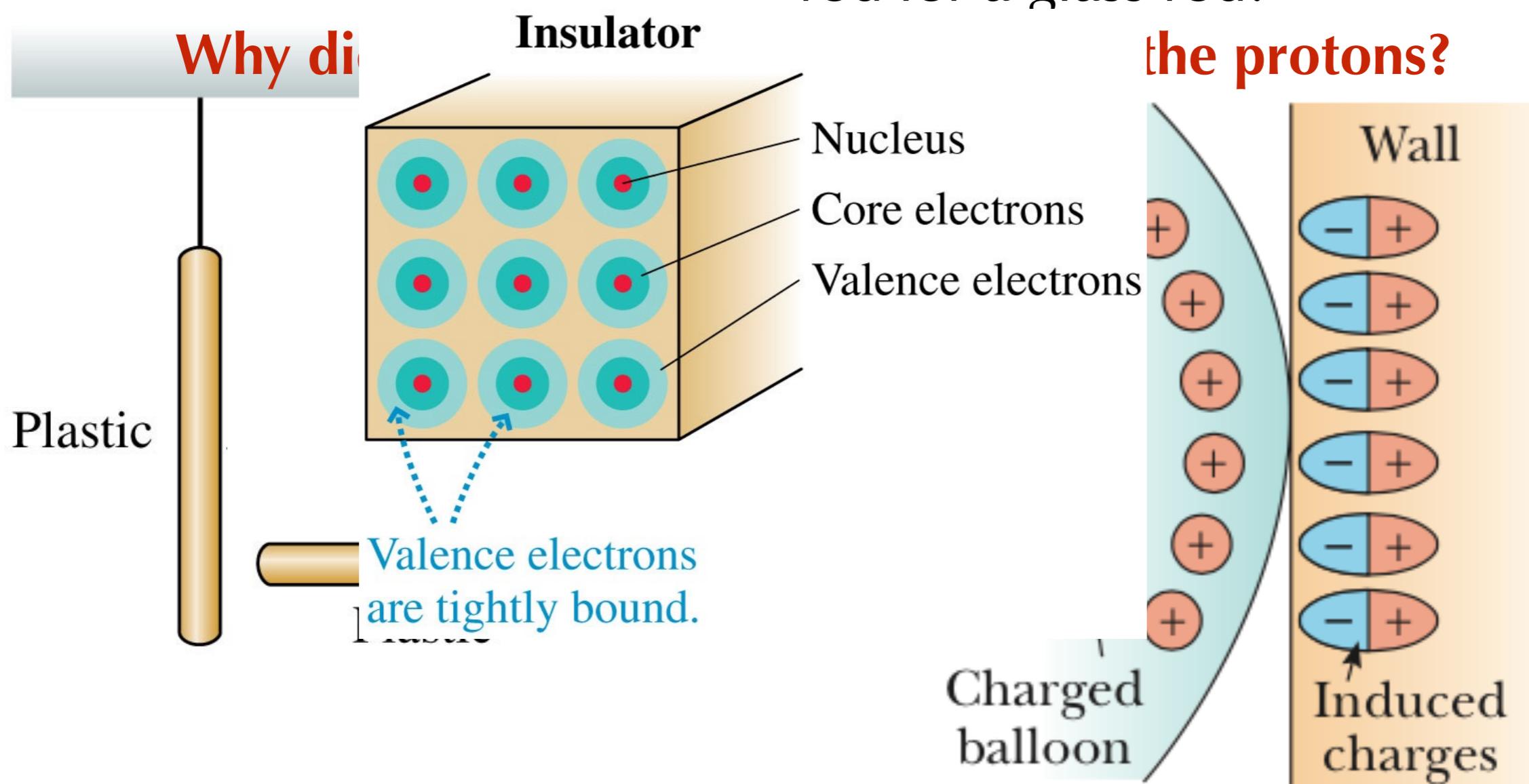
What if I exchange the plastic rod for a glass rod?

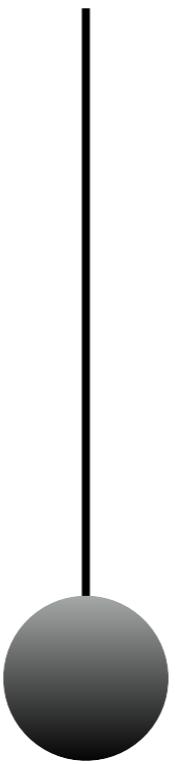
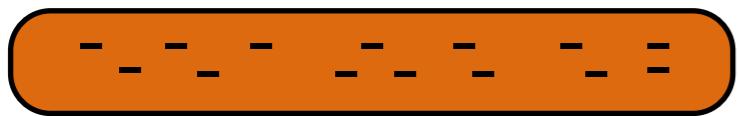
### Why didn't the electrons detach from the protons?

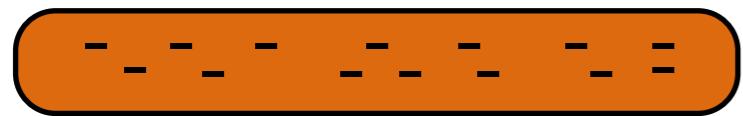


What if I only rub one of the rods?

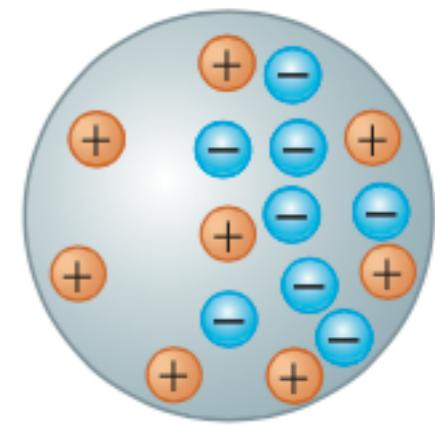
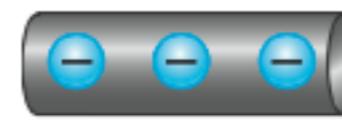
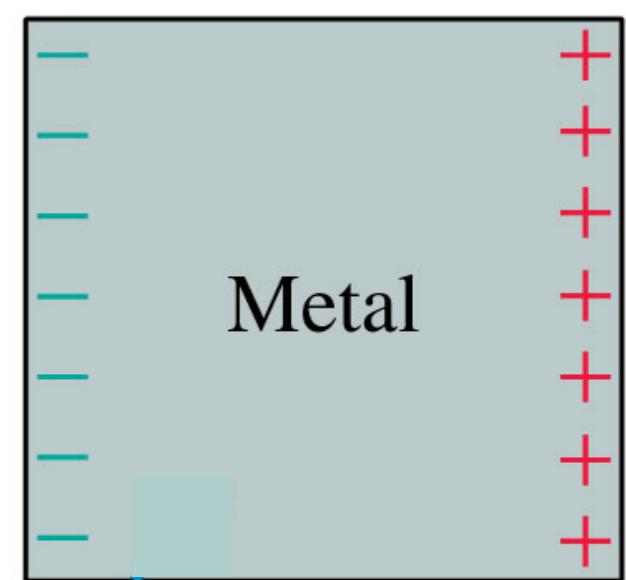
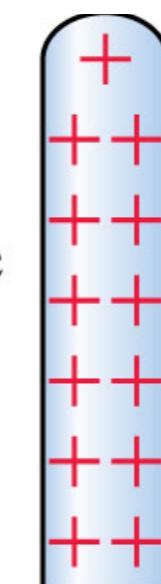
What if I exchange the plastic rod for a glass rod?

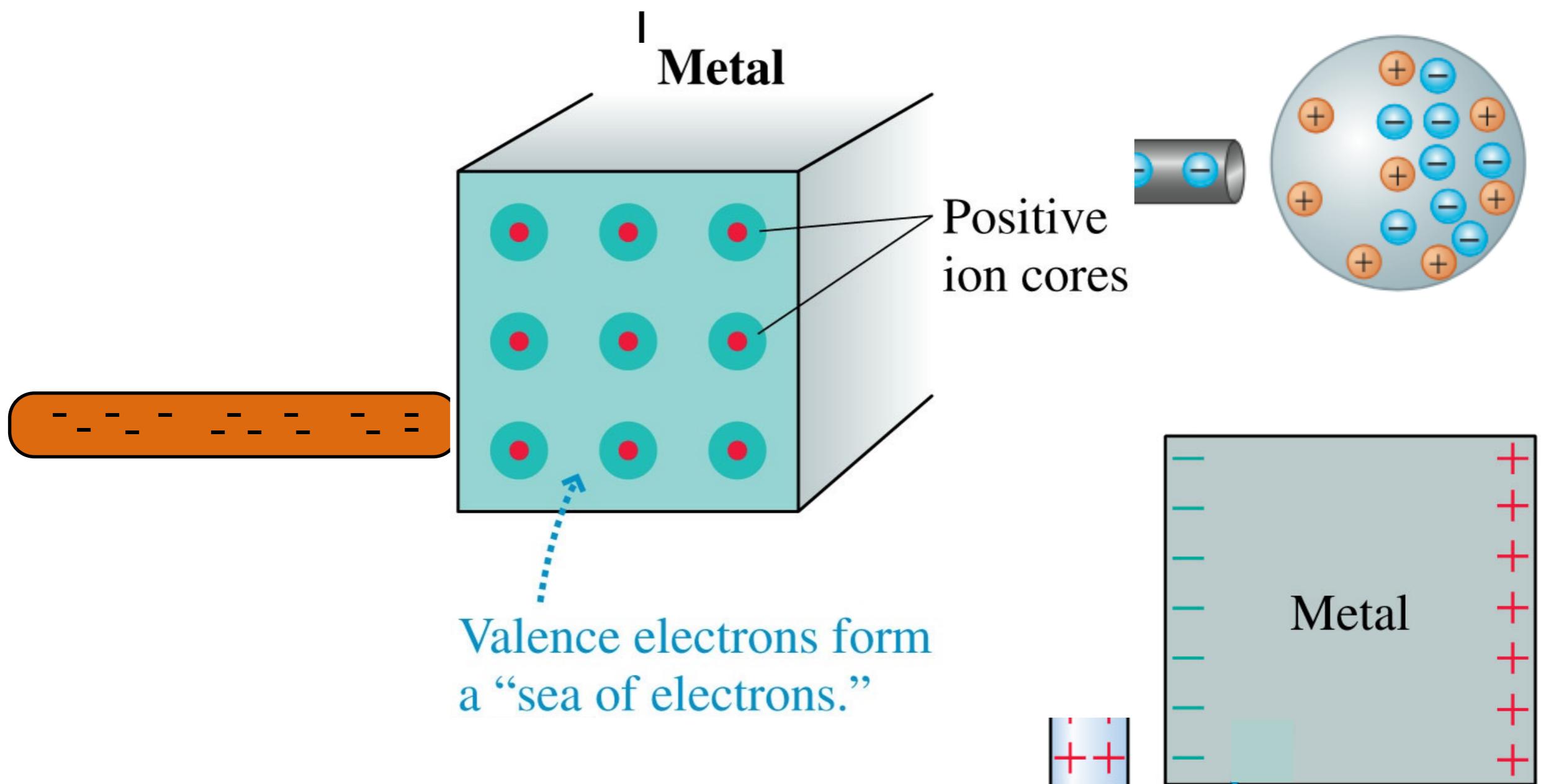






Positive  
rod



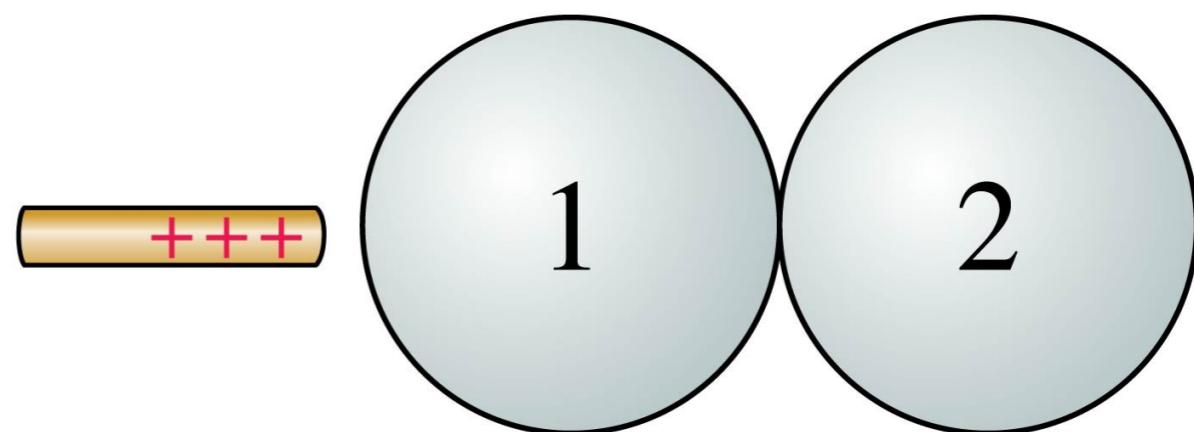


# Quiz

Metal spheres 1 and 2 are touching. Both are initially neutral.

- a. The charged rod is brought near.
- b. The spheres are separated.
- c. The charged rod is then removed.

Afterward, the charges on the sphere are:

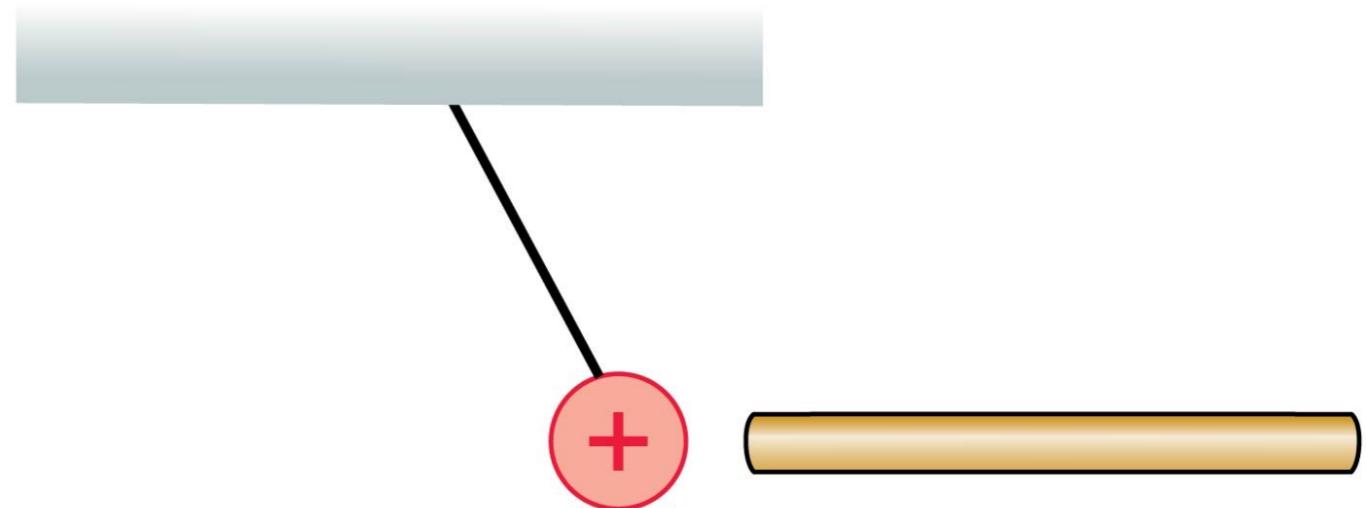


- A.  $Q_1$  is + and  $Q_2$  is +.
- B.  $Q_1$  is - and  $Q_2$  is +.
- C.  $Q_1$  is + and  $Q_2$  is -.
- D.  $Q_1$  is - and  $Q_2$  is -.
- E.  $Q_1$  is 0 and  $Q_2$  is 0.

# Quiz

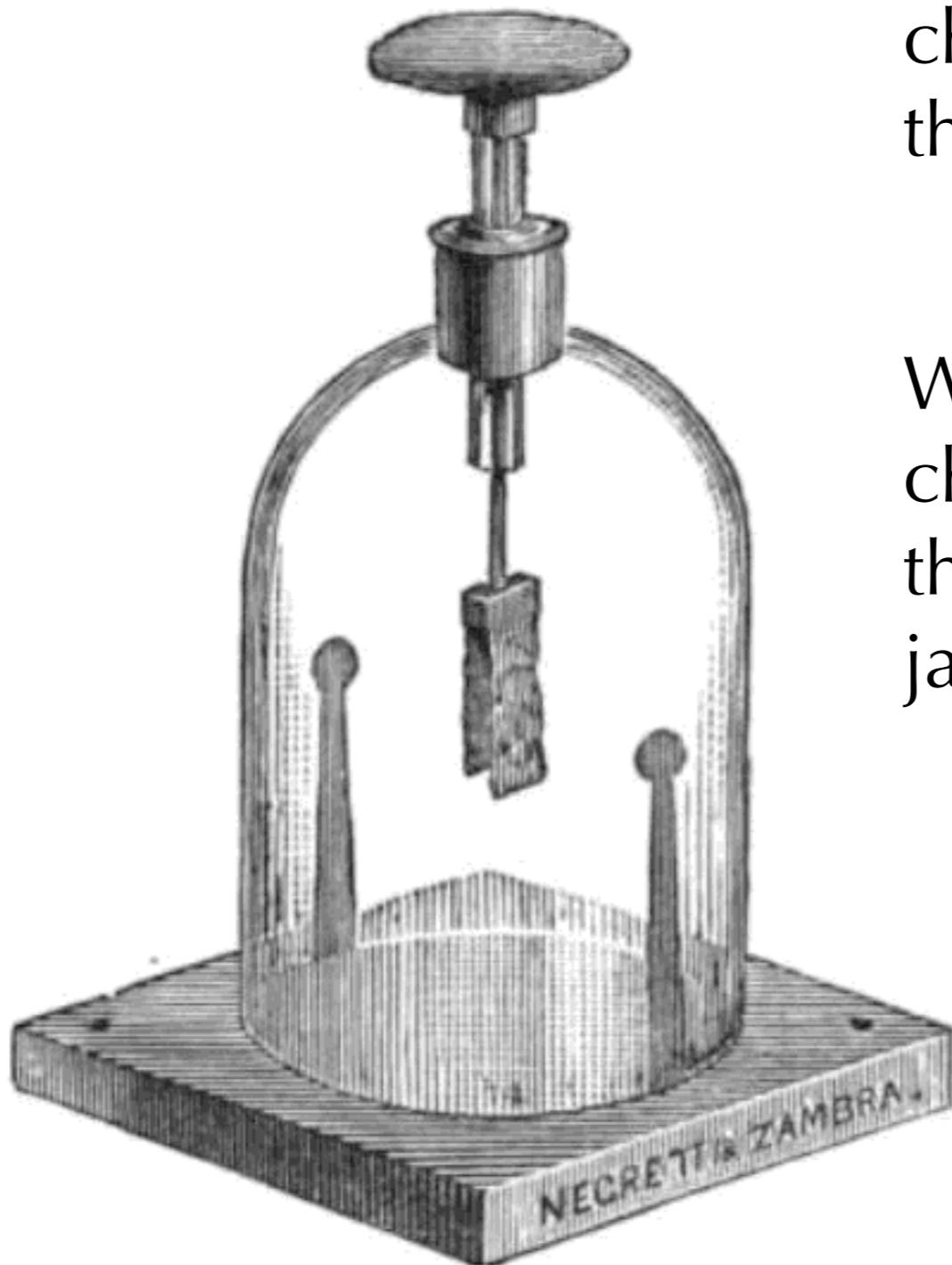
A rod attracts a positively charged hanging ball.  
The rod is

- A. Positive.
- B. Negative.
- C. Neutral.
- D. Either B or C.
- E. Either A or C.



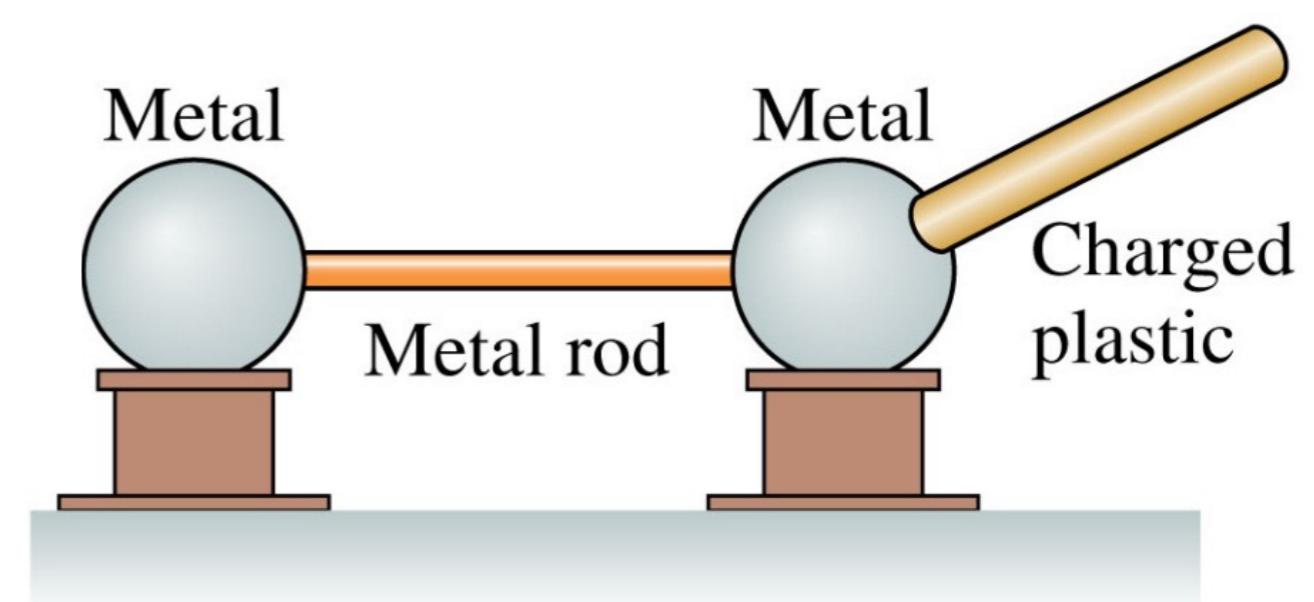
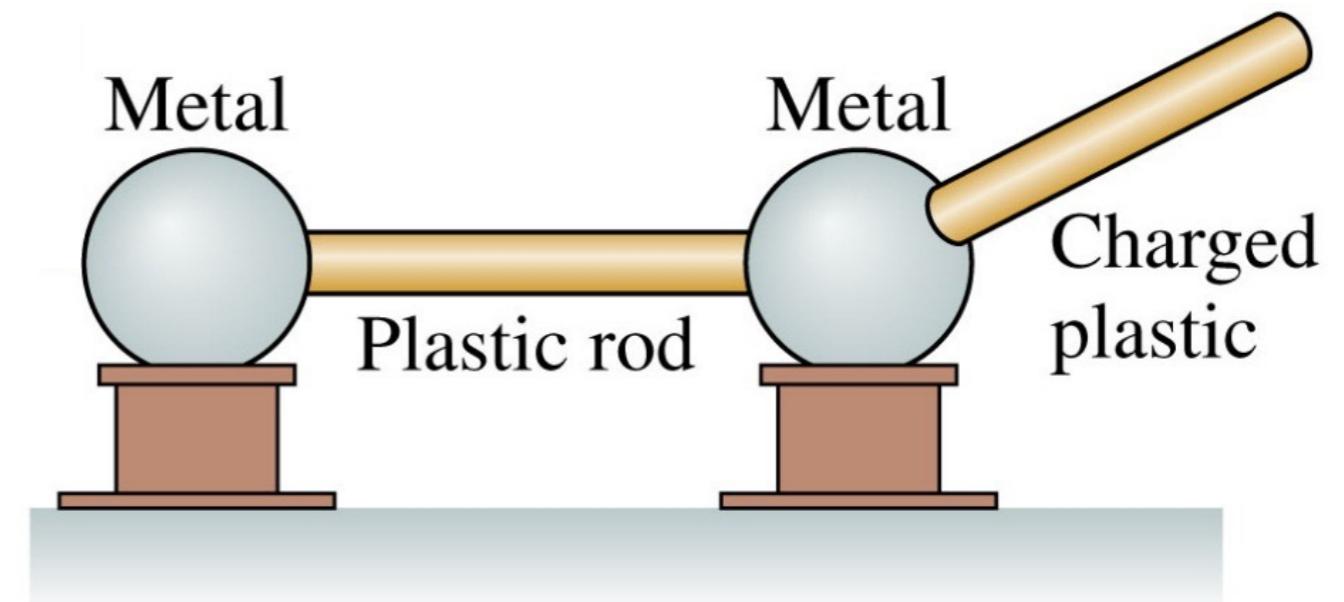
# Electroscope

What will happen if I bring a charged object close to the top of the jar?



What will happen if I bring a charged object close to the top of the jar while touching the top of the jar?

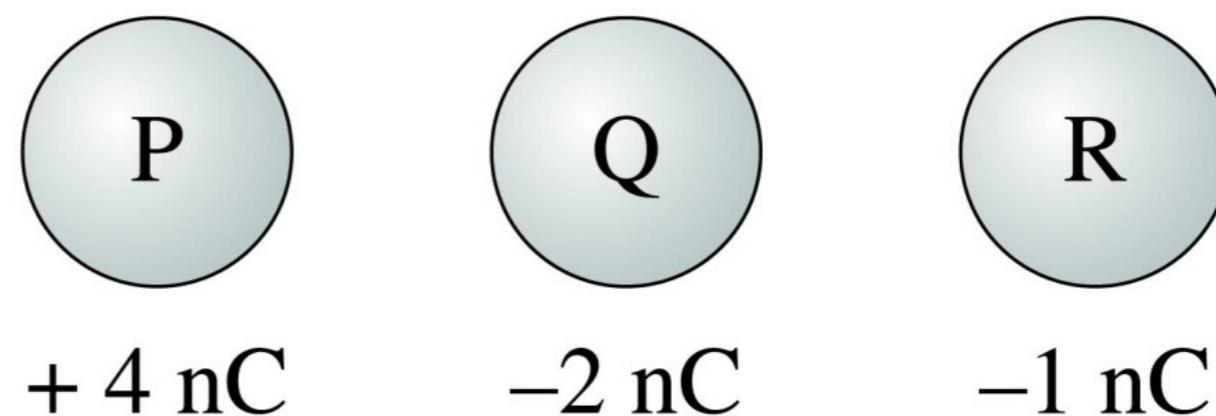
What will be charge state of the metal spheres?



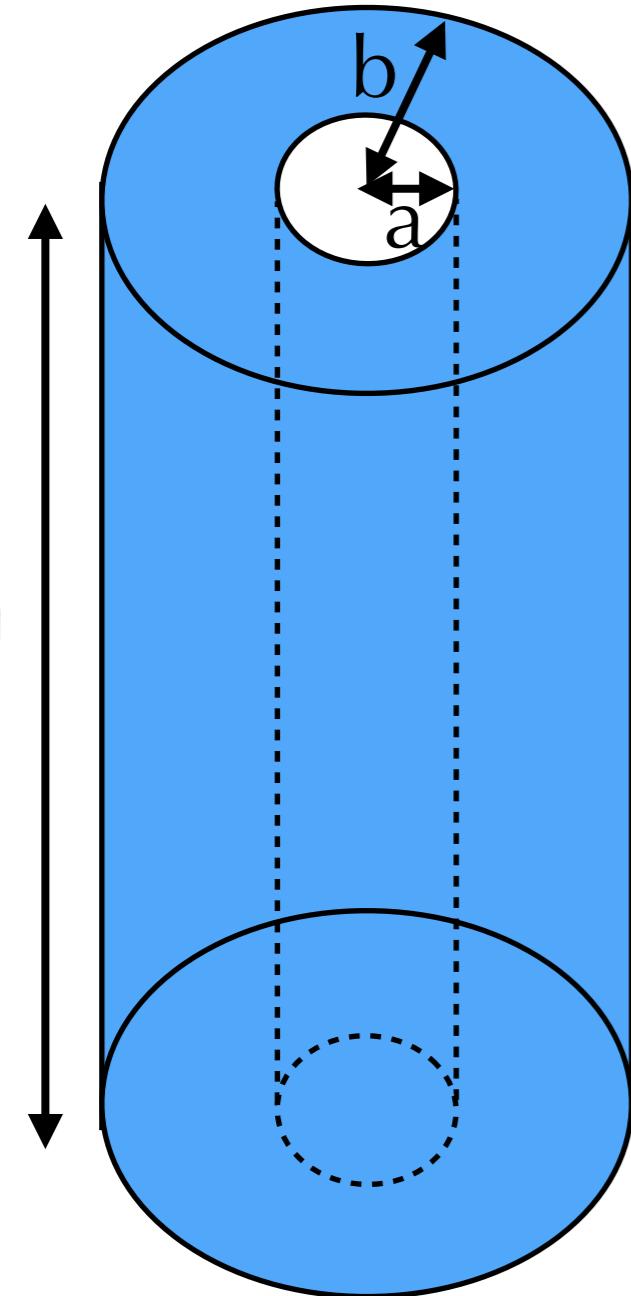
# Quiz

Identical metal spheres are initially charged as shown.  
Spheres P and Q are touched together and then separated.  
Then spheres Q and R are touched together and separated.  
Afterward the charge on  
sphere R is

- A. 0 nC.
- B. -1 nC or less.
- C. -0.5 nC.
- D. +0.5 nC.
- E. +1.0 nC or more.



A right cylindrical shell of inner radius  $a$ , outer radius  $b$ , and height  $h$  is shown in the figure. The mass density of the object varies according to the function:



$$\sigma(\rho, \phi) = \rho^2 \sin^2 \phi$$

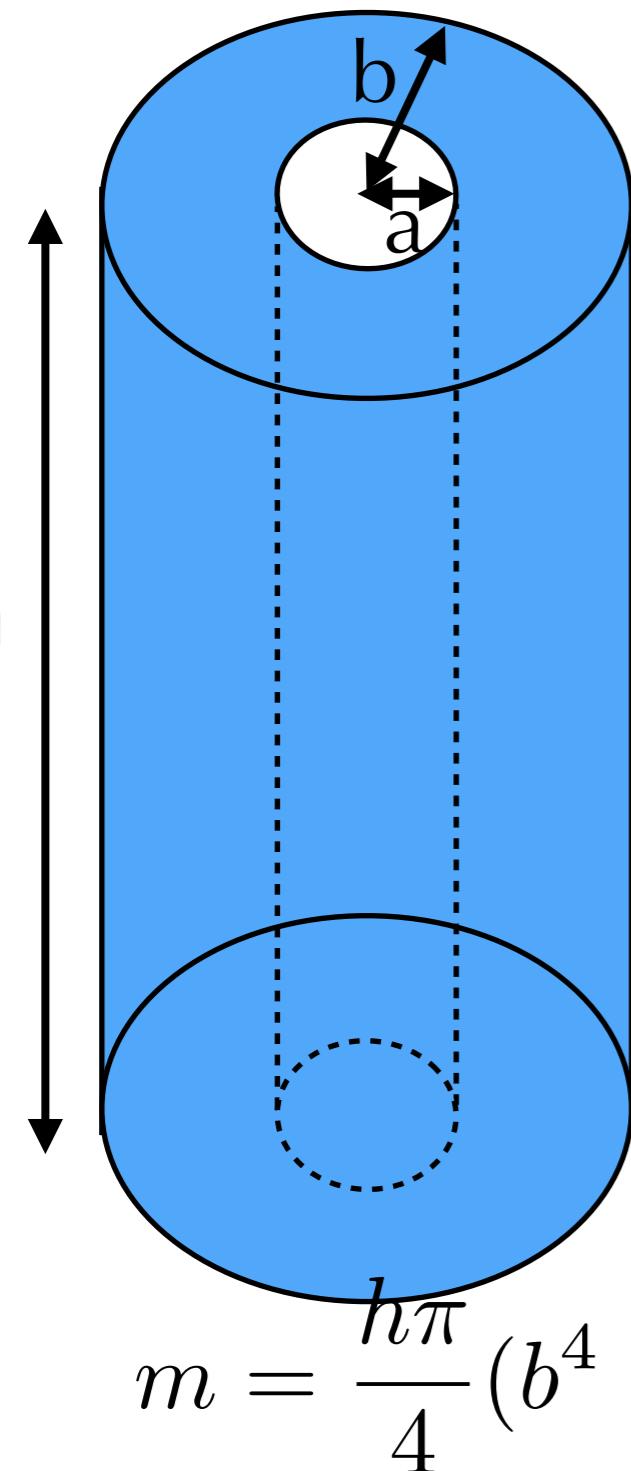
Find an expression for the mass of the object.

The following trig identity may be useful

$$\sin^2 u = \frac{1 - \cos 2u}{2}$$

Try solving using Mathematica

A right cylindrical shell of inner radius  $a$ , outer radius  $b$ , and height  $h$  is shown in the figure. The mass density of the object varies according to the function:



$$\sigma(\rho, \phi) = \rho^2 \sin^2 \phi$$

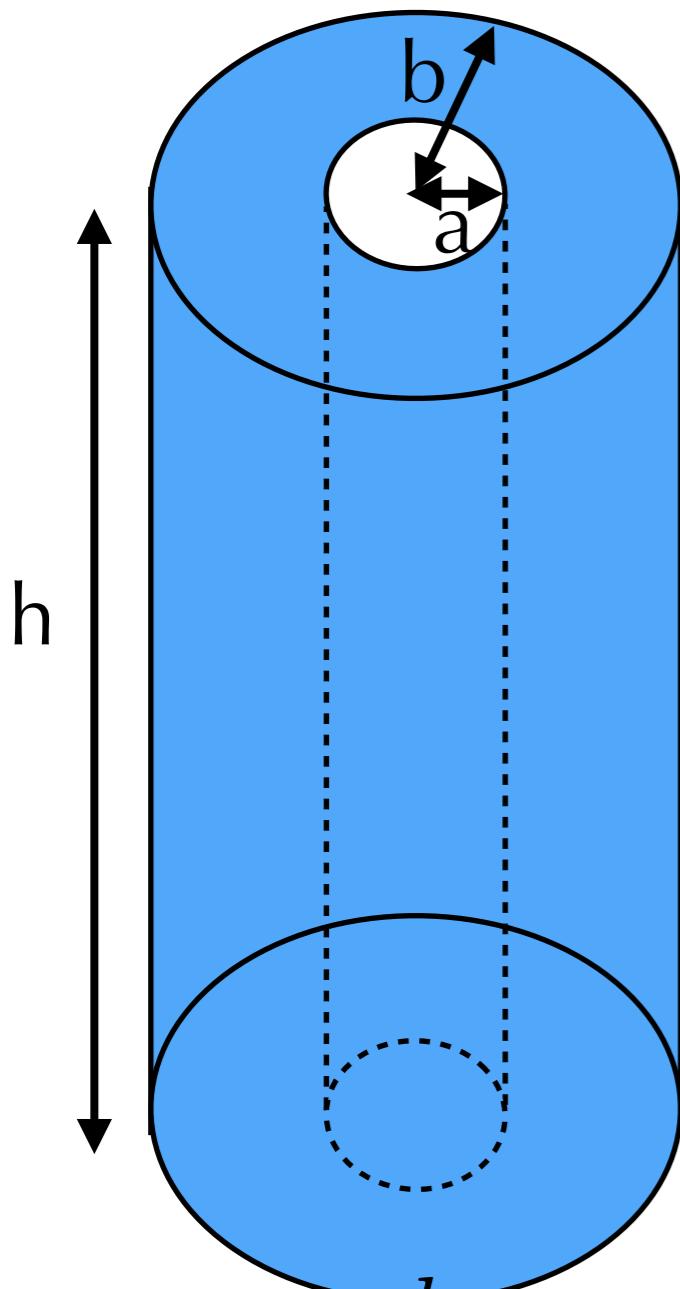
Find an expression for the mass of the object.

The following trig identity may be useful

$$\sin^2 u = \frac{1 - \cos 2u}{2}$$

Try solving using Mathematica

A right cylindrical shell of inner radius  $a$ , outer radius  $b$ , and height  $h$  is shown in the figure. The mass density of the object varies according to the function:

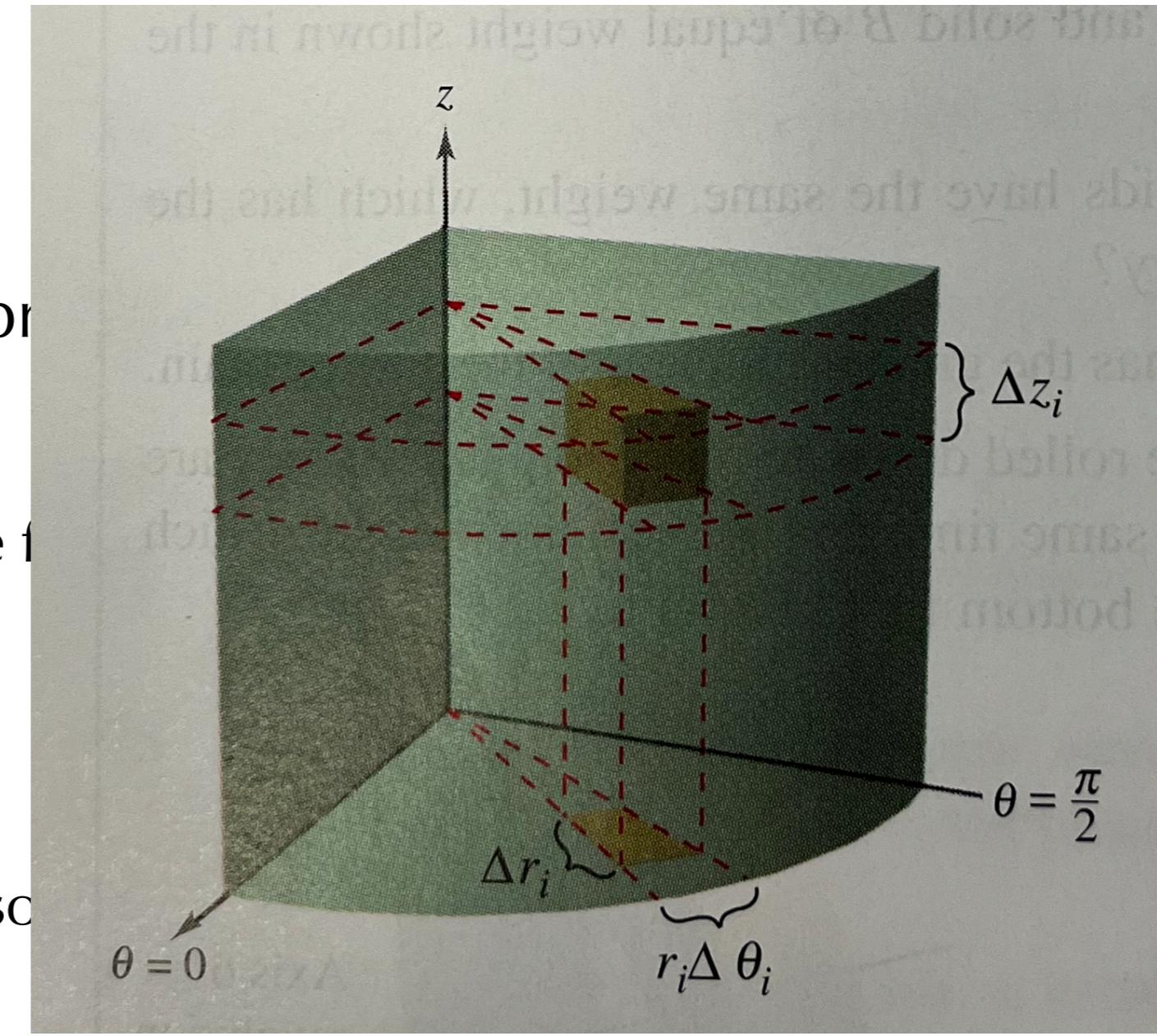


$$m = \frac{h\pi}{4} (b^4 - a^4)$$

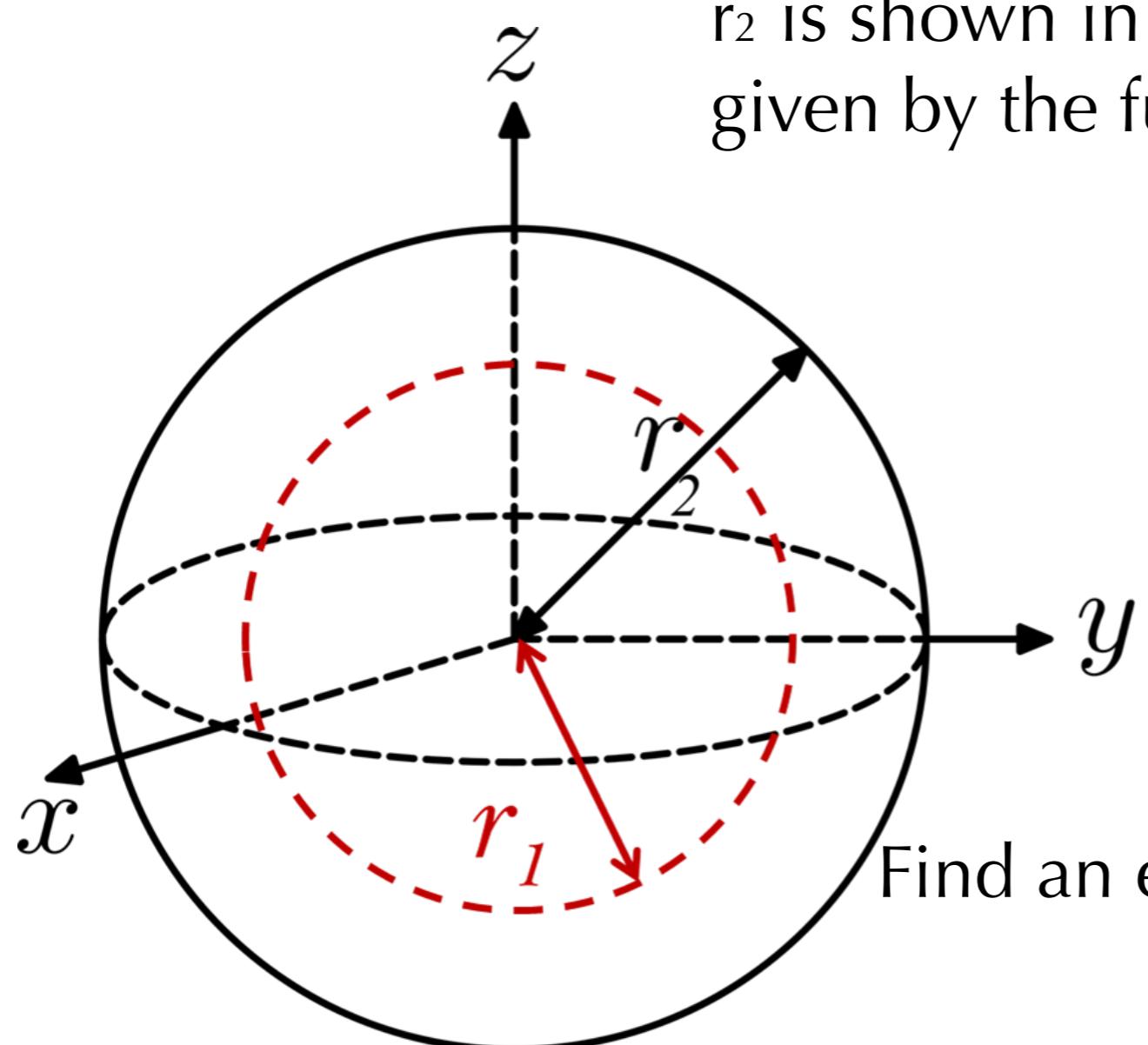
Find an expression for the mass of the shell.

The formula is

Try so



A spherical shell of inner radius  $r_1$  and outer radius  $r_2$  is shown in the figure. Its density varies and is given by the function:

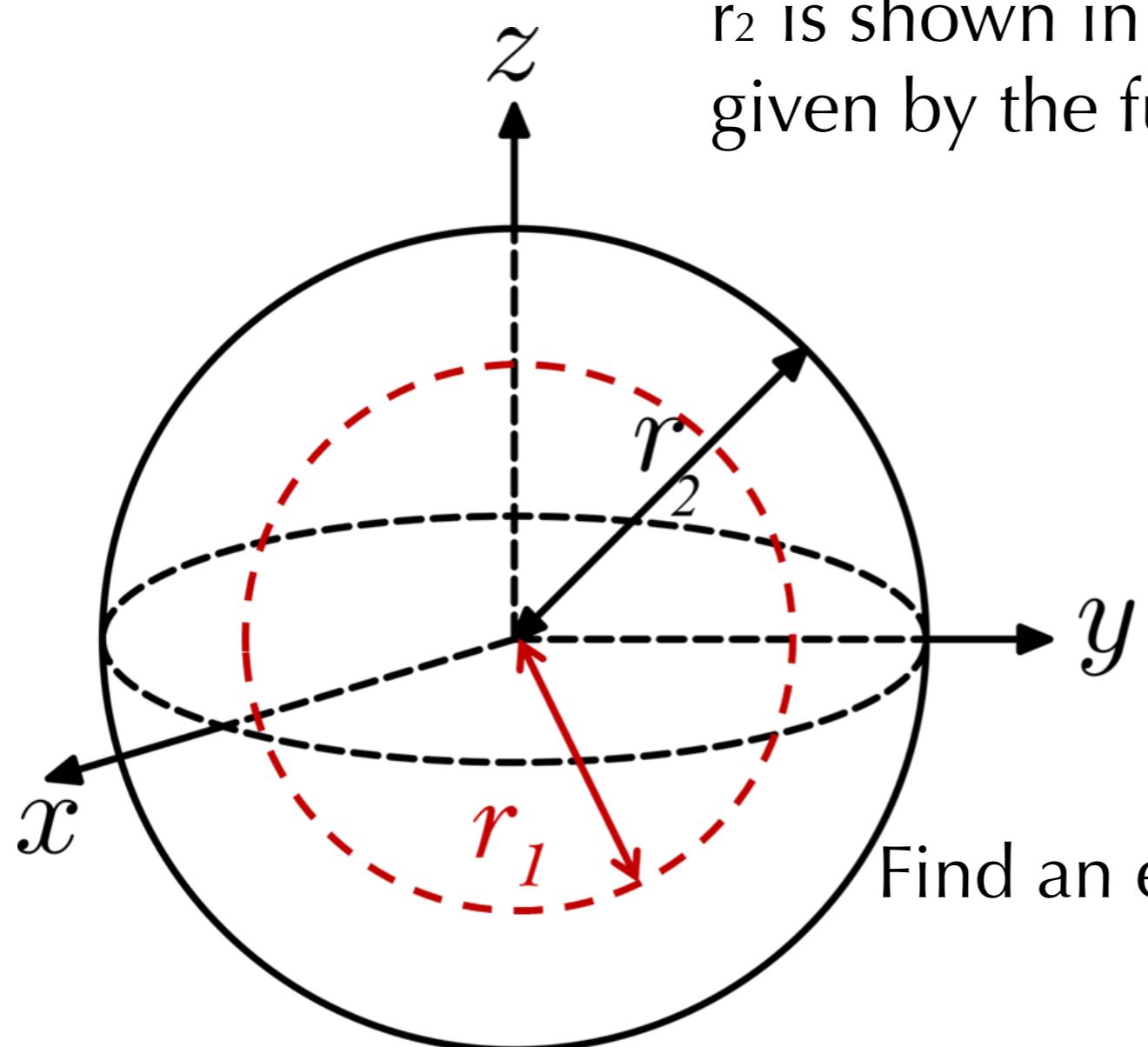


$$\rho(r, \theta, \phi) = \sqrt{r} \sin^2 \phi$$

Find an expression for the mass of the object.

to ponder: what is a differential volume element in spherical coordinates.

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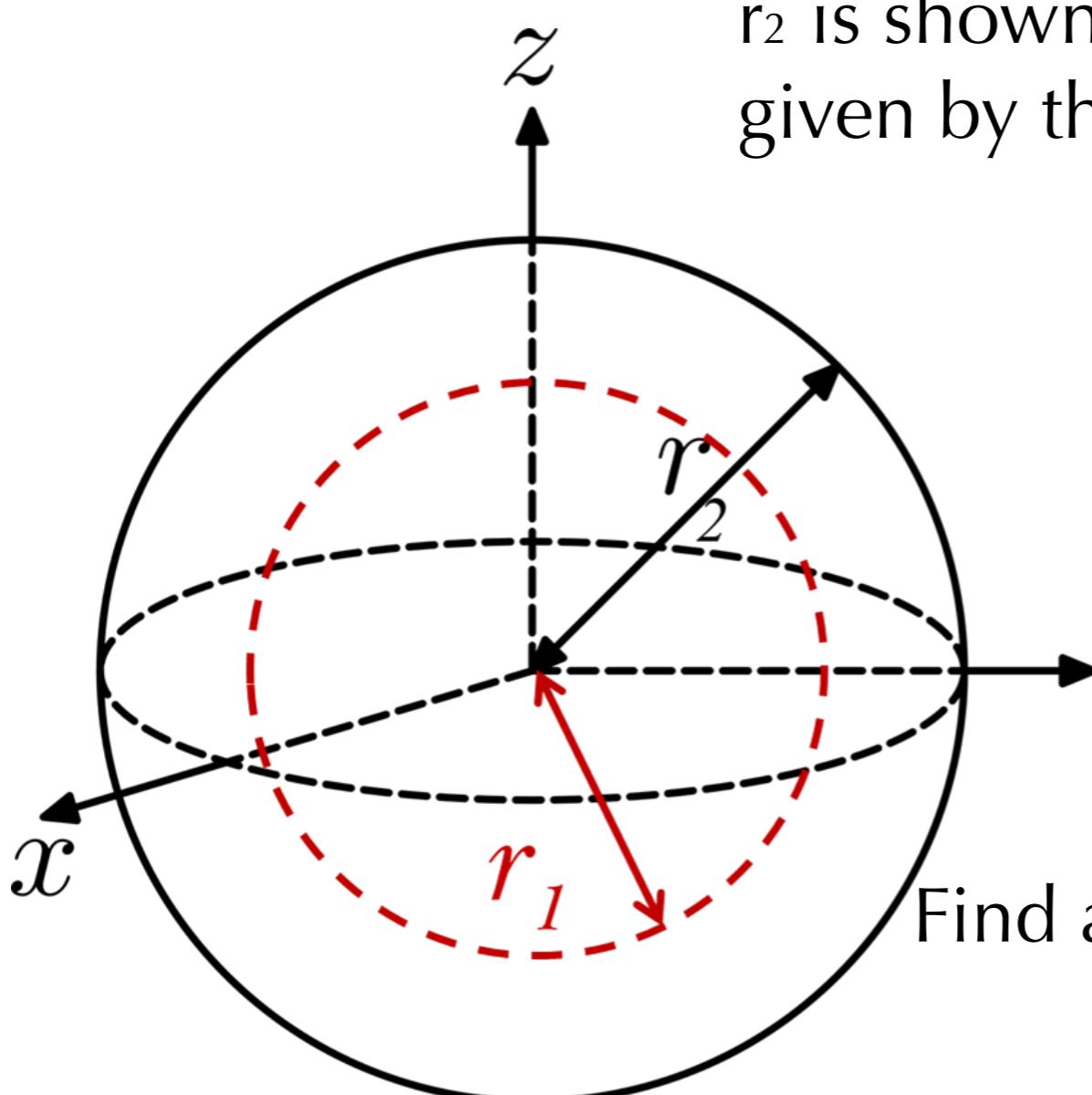
$$\rho(r, \theta, \phi) = \sqrt{r} \sin^2 \phi$$

Find an expression for the mass of the object.

$$m = \frac{4\pi}{7} (r_2^{7/2} - r_1^{7/2})$$

to ponder: what is a differential volume element in spherical coordinates.

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Find an

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to pond  
spherical

