

Outcomes

Learning

- Struggling is good. Spend more time struggling.
- If its easy, its not learning and it won't last.
- Activity: thermal expansion
- Activity: Buoyancy

Discern between the two types of charge

- Likes repel, opposites attract.
- Activity: rubbing rods with felt and silk.

Understand induced charge in conductor and insulator

- Insulators: molecular polization. Conductor: charge polarization
- Activity: rubbing single object and observe forces.

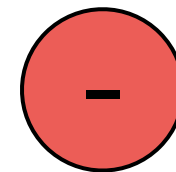
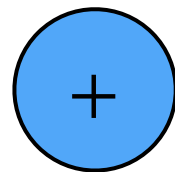
Know difference between insulator and conductor.

- Insulators: electrons tightly bound, conductors loosely bound (sea of electrons).
- Activity: Spheres connected with rods.

The electric field is a vector field that originates from positive charges and terminates at negative charges. It is a physical field that exerts a force on other charges. The electric field is a vector field that originates from positive charges and terminates at negative charges. It is a physical field that exerts a force on other charges.

Opposites attract and likes repel!

Anybody who's married knows this is true!



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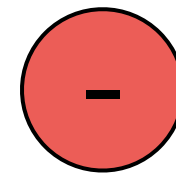
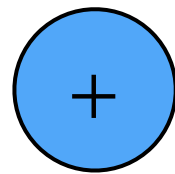
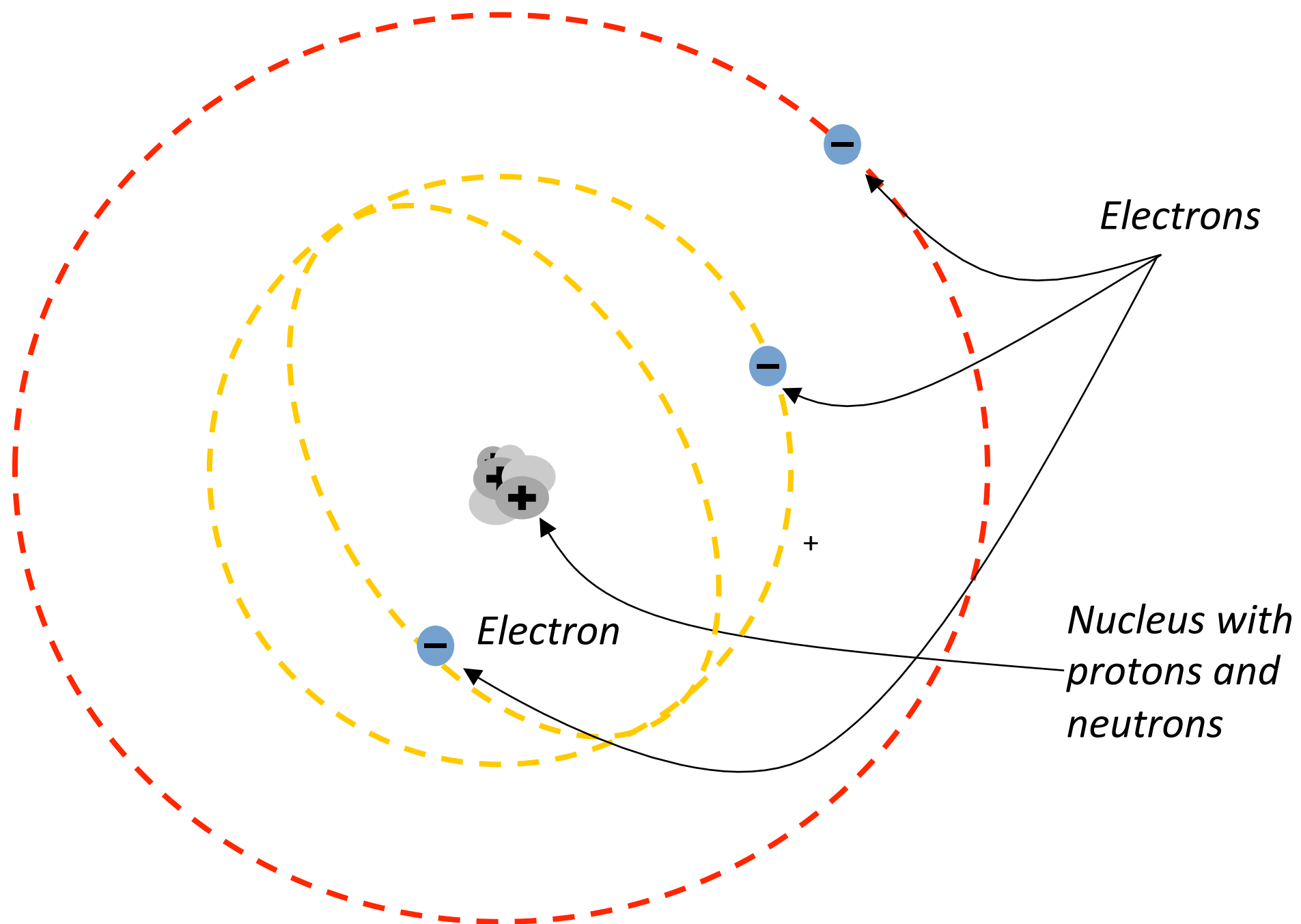


TABLE 25.1 Protons and electrons

| Particle | Mass (kg) | Charge |
|----------|------------------------|--------|
| Proton | 1.67×10^{-27} | $+e$ |
| Electron | 9.11×10^{-31} | $-e$ |



Particles not drawn to scale

Reading the periodic table

Symbol

A one- or two-letter abbreviation derived from the element's English or Latin name.

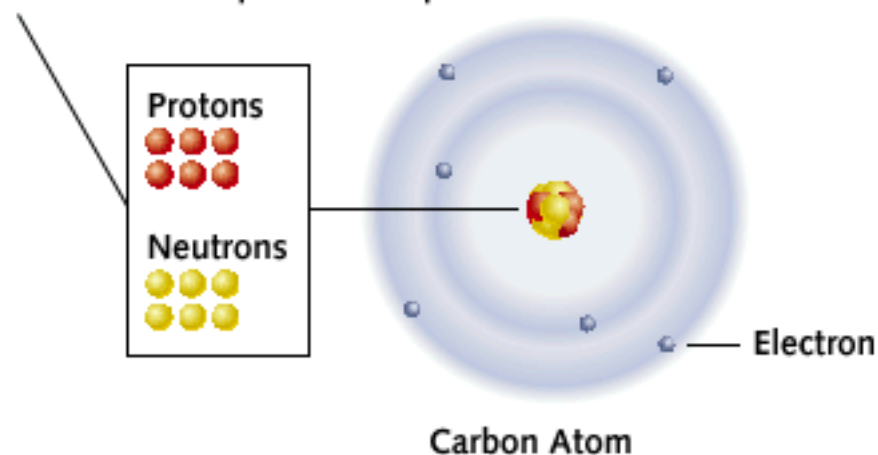
Name

Element's common name.

| |
|--------|
| 6 |
| C |
| Carbon |
| 12.011 |

Mass Number

The sum of the numbers of protons and neutrons in a specific isotope.



Reading the periodic table

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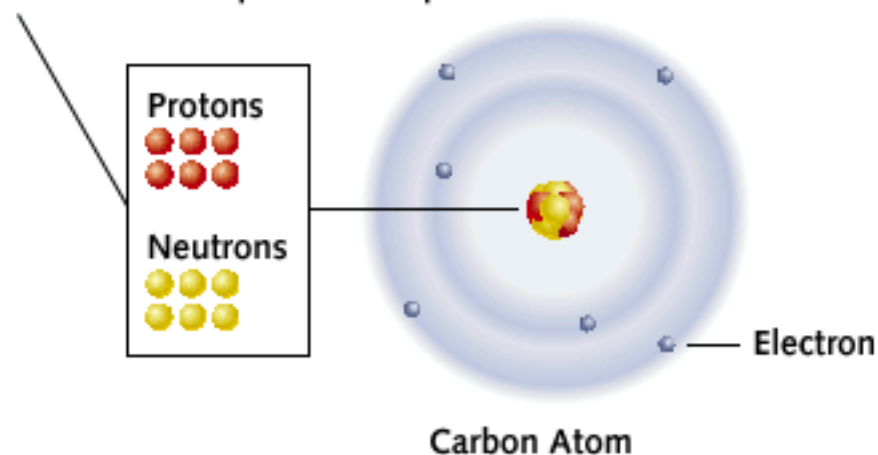
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Atomic Number

Equal to the number of protons in the nucleus, as well as the number of electrons in the electron cloud.

Mass Number

The sum of the numbers of protons and neutrons in a specific isotope.



Reading the periodic table

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Atomic Number

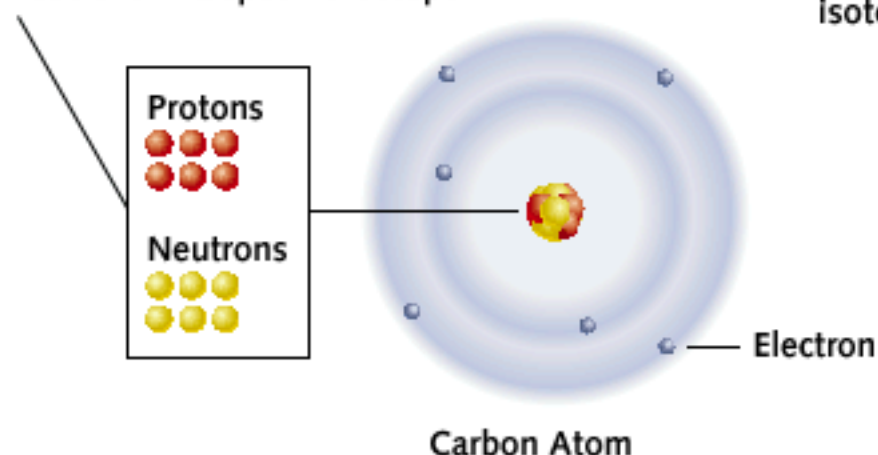
Equal to the number of protons in the nucleus, as well as the number of electrons in the electron cloud.

Atomic Mass

Weighted average of the masses of all the element's isotopes. Rounding the atomic mass to the nearest whole number yields the mass number of the most common isotope.

Mass Number

The sum of the numbers of protons and neutrons in a specific isotope.



Reading the periodic

Symbol

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Element's common name.

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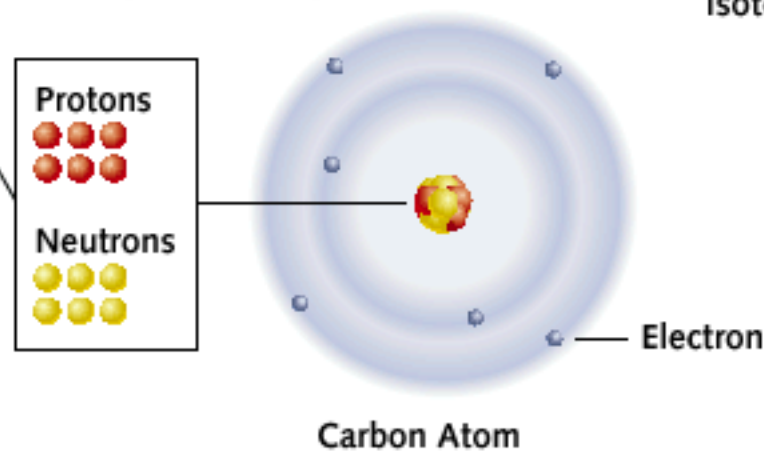
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| | | | | | |
|-------------------------------------------|----------------------------------------|---------------------------------------------|------------------------------------------|---------------------------------------------|--------------------------------------------|
| | | | | | 18 |
| | | | | | 2 He Helium 4.00 |
| | 13 | 14 | 15 | 16 | 17 |
| 5 B Boron 10.81 | 6 C Carbon 12.01 | 7 N Nitrogen 14.01 | 8 O Oxygen 16.00 | 9 F Fluorine 19.00 | 10 Ne Neon 20.18 |
| 13 Al Aluminum 26.98 | 14 Si Silicon 28.09 | 15 P Phosphorus 30.97 | 16 S Sulfur 32.06 | 17 Cl Chlorine 35.45 | 18 Ar Argon 39.95 |
| 31 Ga Gallium 69.73 | 32 Ge Germanium 72.61 | 33 As Arsenic 74.92 | 34 Se Selenium 78.09 | 35 Br Bromine 79.90 | 36 Kr Krypton 84.80 |
| 49 In Indium 114.82 | 50 Sn Tin 118.71 | 51 Sb Antimony 121.76 | 52 Te Tellurium 127.6 | 53 I Iodine 126.90 | 54 Xe Xenon 131.29 |
| 81 Tl Thallium 204.38 | 82 Pb Lead 207.20 | 83 Bi Bismuth 208.98 | 84 Po Polonium [208.98] | 85 At Astatine 209.98 | 86 Rn Radon 222.02 |
| 113 Uut Ununtrium unknown | 114 Fl Flerovium [289] | 115 Uup Ununpentium unknown | 116 Lv Livermorium [298] | 117 Uus Ununseptium unknown | 118 Uuo Ununoctium unknown |

Reading the periodic table

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Mass Number

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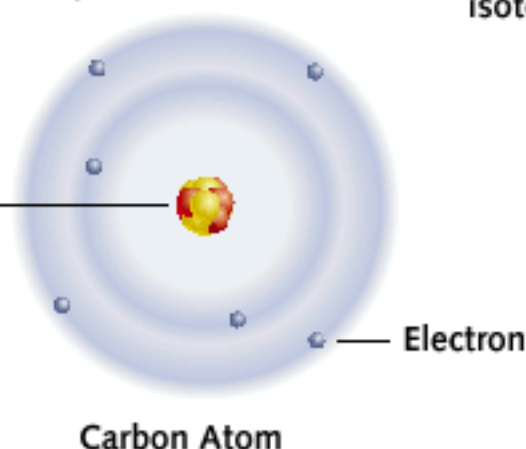
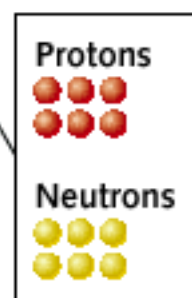
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| | |
|----------|--|
| 6 | |
| C | |
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| | | | | | |
|--------------------------------------|-------------------------------------|---------------------------------------|-----------------------------------|--------------------------------------|-----------------------------------|
| | | | | | 18 |
| | | | | | 2 He Helium 4.00 |
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| 13 Al Aluminum 26.98 | 14 Si Silicon 28.09 | 15 P Phosphorus 30.97 | 16 S Sulfur 32.06 | 17 Cl Chlorine 35.45 | 18 Ar Argon 39.95 |

| | | | | |
|-----------------------------------------------|-------------------------------------------------|-----------------------------------------------|-----------------------------------------------|----------------------------------------------|
| iron Fe 26 55.845(2) | cobalt Co 27 58.933 195(5) | nickel Ni 28 58.6934(4) | copper Cu 29 63.546(3) | zinc Zn 30 65.38(2) |
| ruthenium Ru 44 101.07(2) | rhodium Rh 45 102.905 50(2) | palladium Pd 46 106.42(1) | silver Ag 47 107.8682(2) | cadmium Cd 48 112.411(8) |
| osmium Os 76 190.23(3) | iridium Ir 77 192.217(3) | platinum Pt 78 195.084(9) | gold Au 79 196.966 569(4) | mercury Hg 80 200.59(2) |

Reading t

Symbol

A one- or two-letter abbreviation derived from the element's English or Latin name.

Name

Element's common name.

Mass Number

The sum of the numbers of protons and neutrons in a specific isotope.

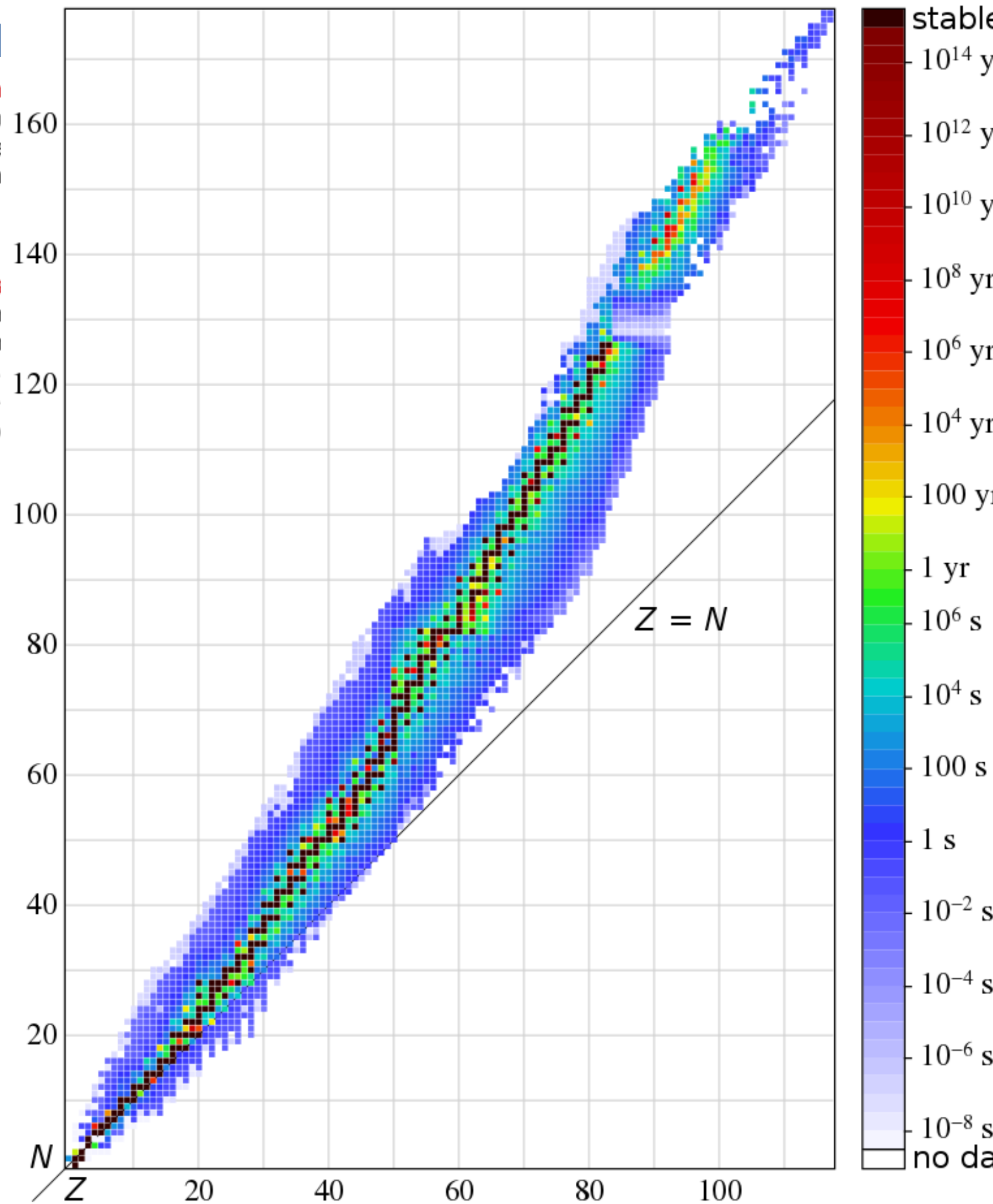
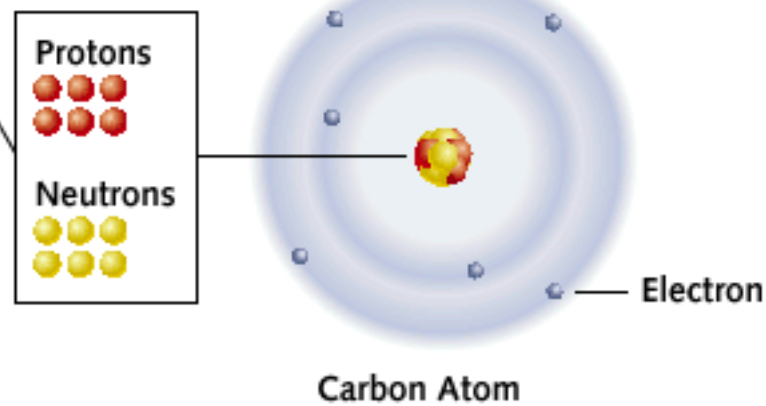
| |
|----------|
| 6 |
| C |
| Carbon |
| 12.011 |

Atomic Num

Equal to the number of protons in the nucleus, as well as the number of electrons in the neutral atom.

Atomic Mass

Weighted average of all the isotopes of the element. Rounding the nearest whole mass number gives the mass number of the most common isotope.



Reading t

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Mass Number

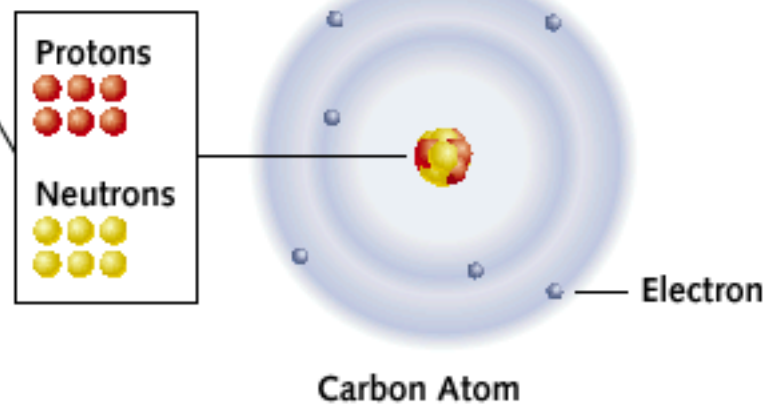
The sum of the numbers of protons and neutrons in a specific isotope.

Atomic Num

Equal to the number of protons in the nucleus, as well as the number of electrons in the atom.

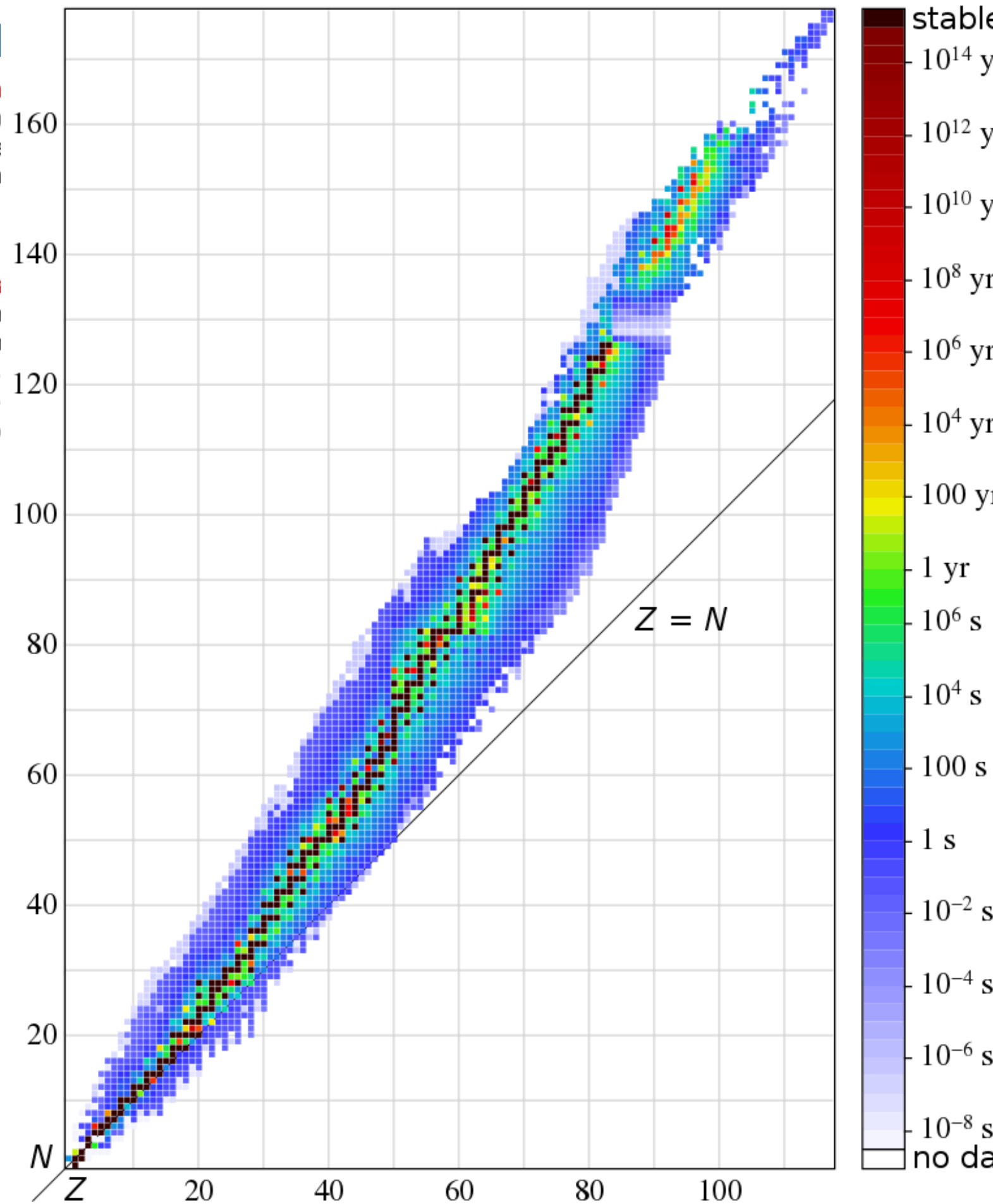
Atomic Mass

Weighted average of all the elements. Rounding the nearest whole mass number of an isotope.



12.011 amu

1 amu = 1.6605×10^{-27} kg



Reading t

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A one- or two-letter abbreviation derived from the element's English or Latin name.

Name

Element's common name.

Mass Number

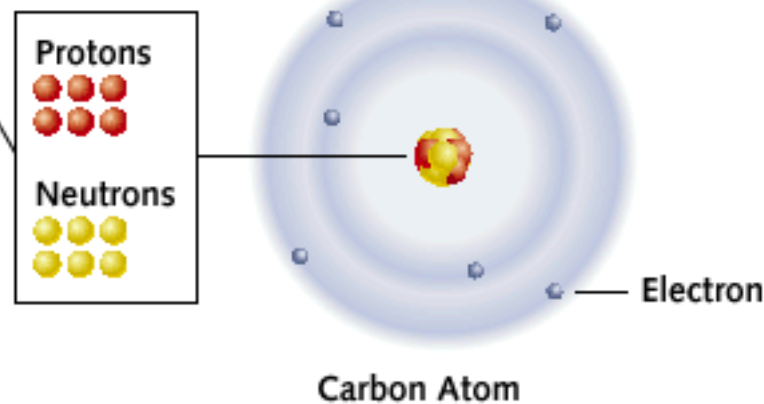
The sum of the numbers of protons and neutrons in a specific isotope.

Atomic Num

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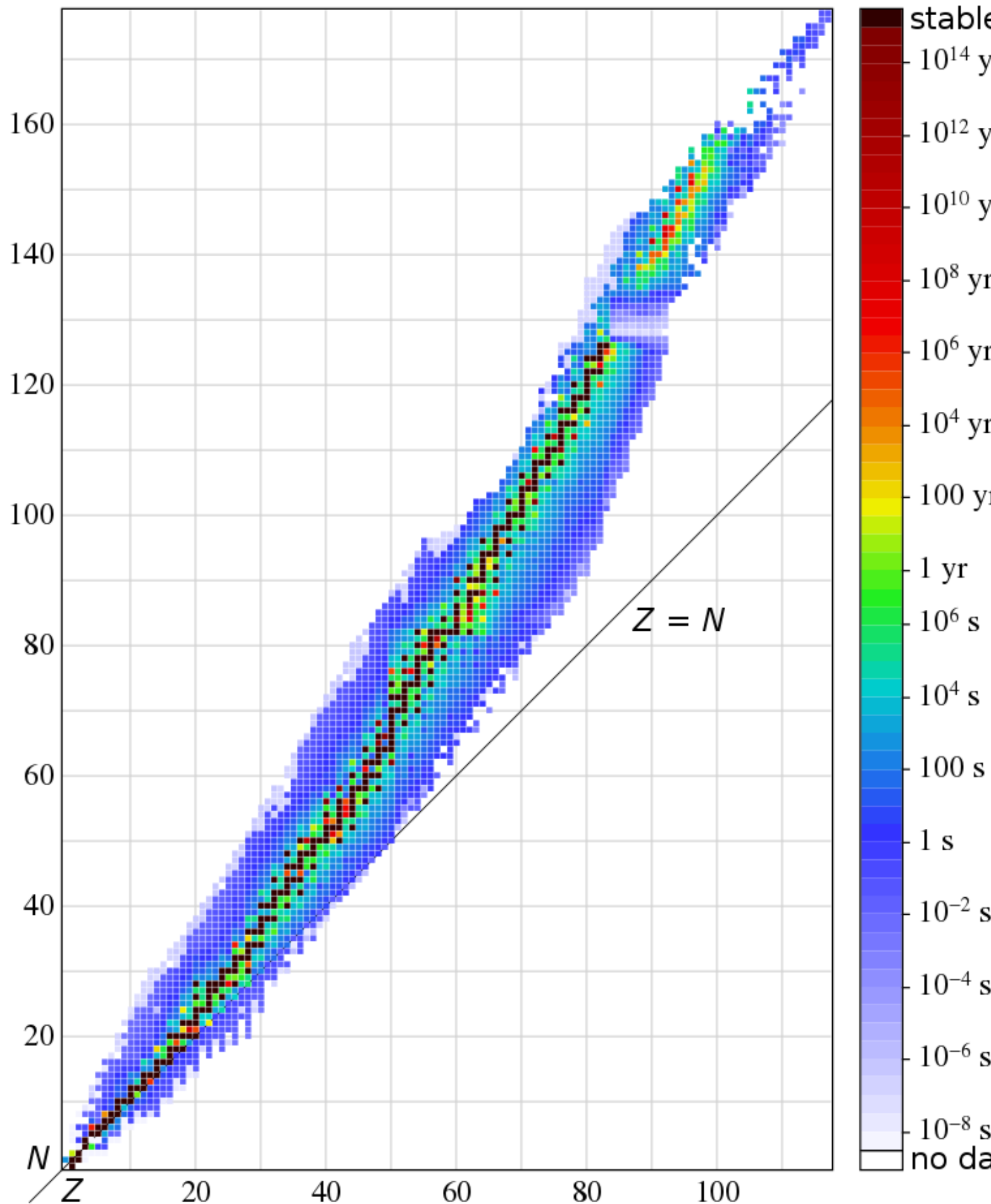
12.011 amu

$1 \text{ amu} = 1.6605 \times 10^{-27} \text{ kg}$

Or

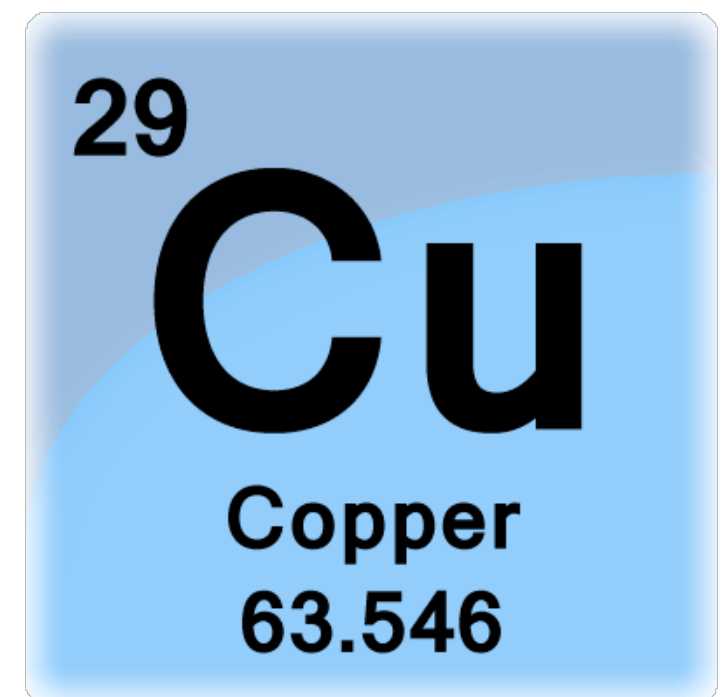
12.011 grams/mol

$1 \text{ mol} = 6.022 \times 10^{23} \text{ atoms}$



Pennies today are copper-plated zinc, but older pennies are 3.1 g of pure copper. What are the total positive charge and total negative charge in a solid copper penny that is electrically neutral?

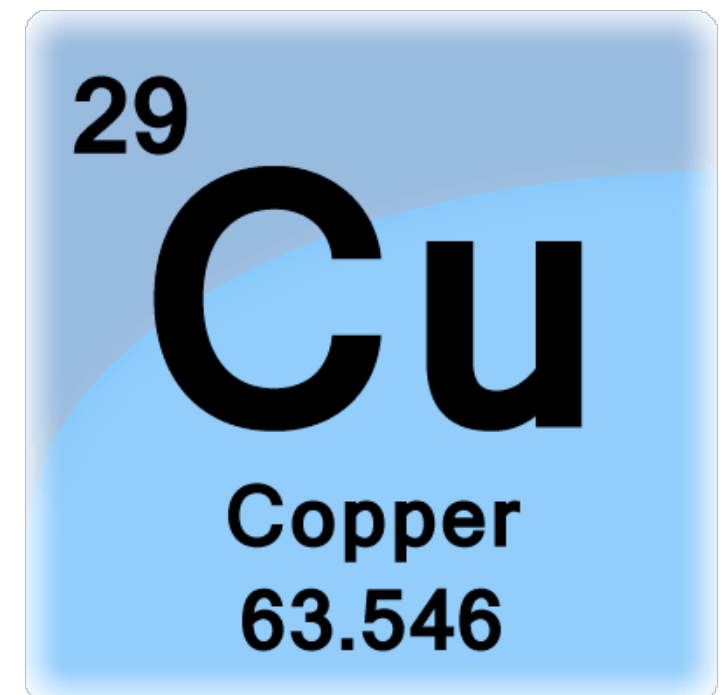
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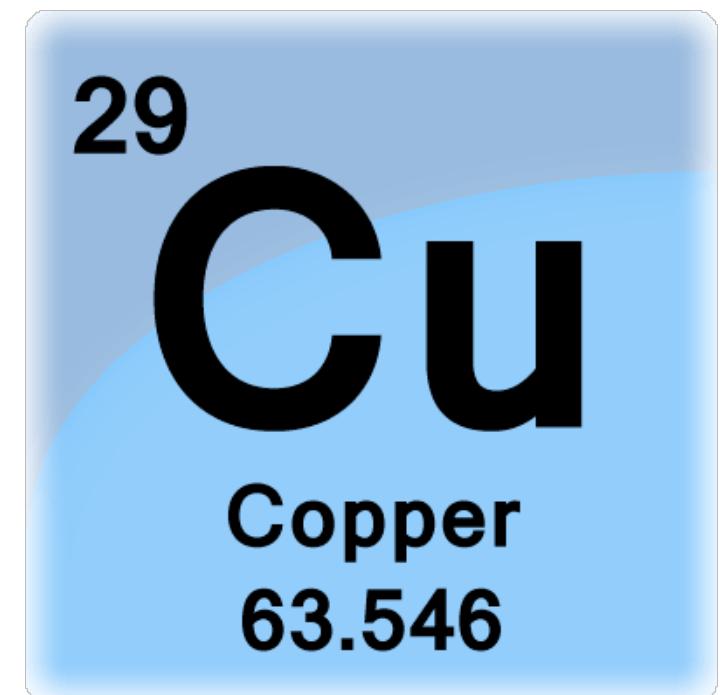
$$N_{\text{atoms}} = \frac{3.1 \text{ g}}{63.56 \times 1.6605 \times 10^{-27}} = 3.2 \times 10^{25}$$



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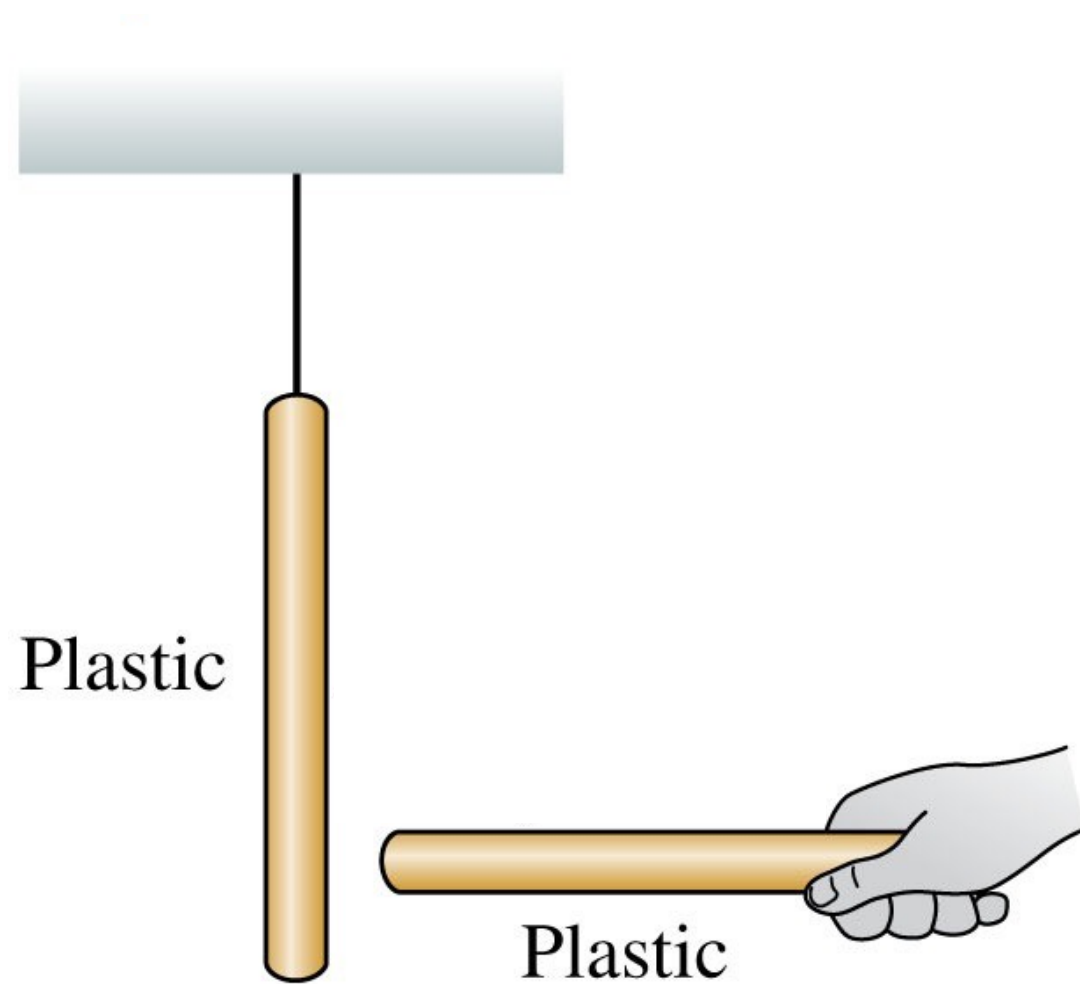
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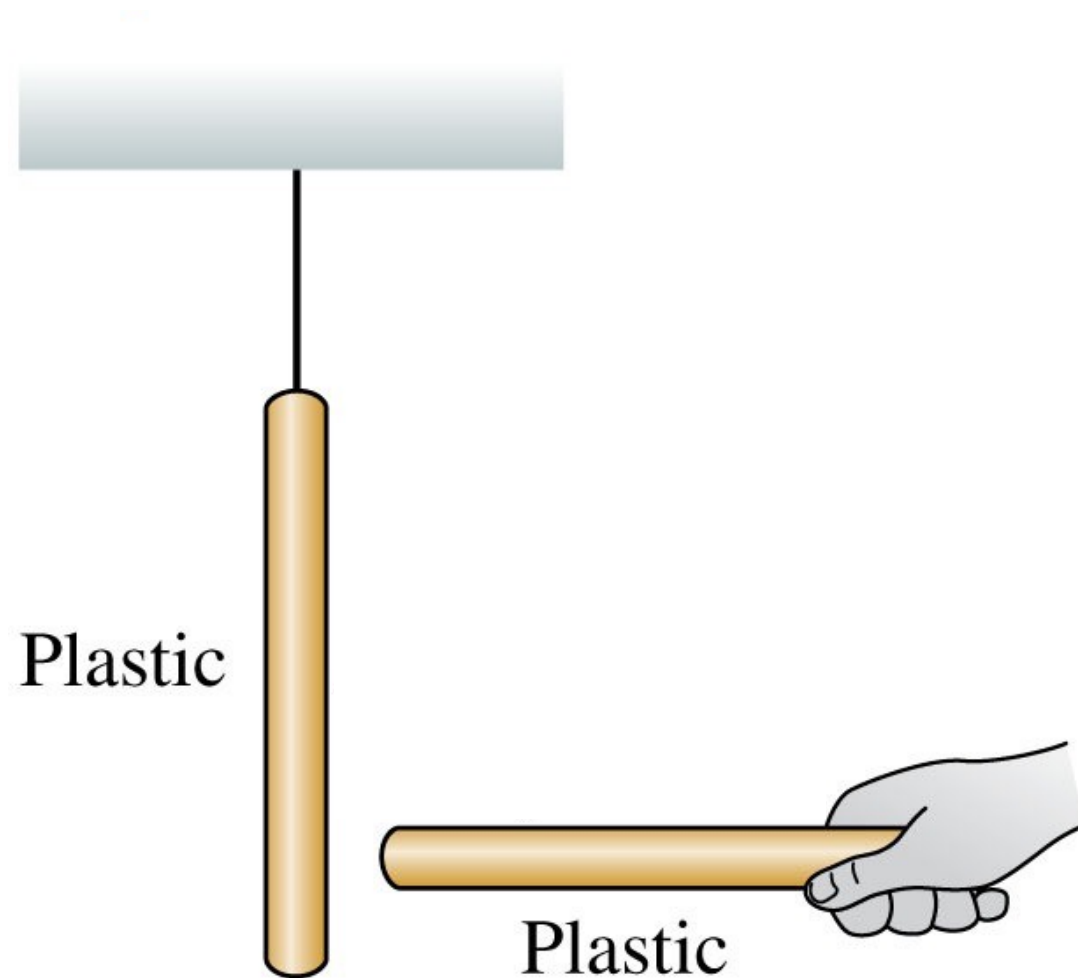


$$q = N_{\text{atoms}} \times 29 \times 1.602 \times 10^{-19} = 1.4 \times 10^6 \text{ C}$$

Do these plastic bars have charge?

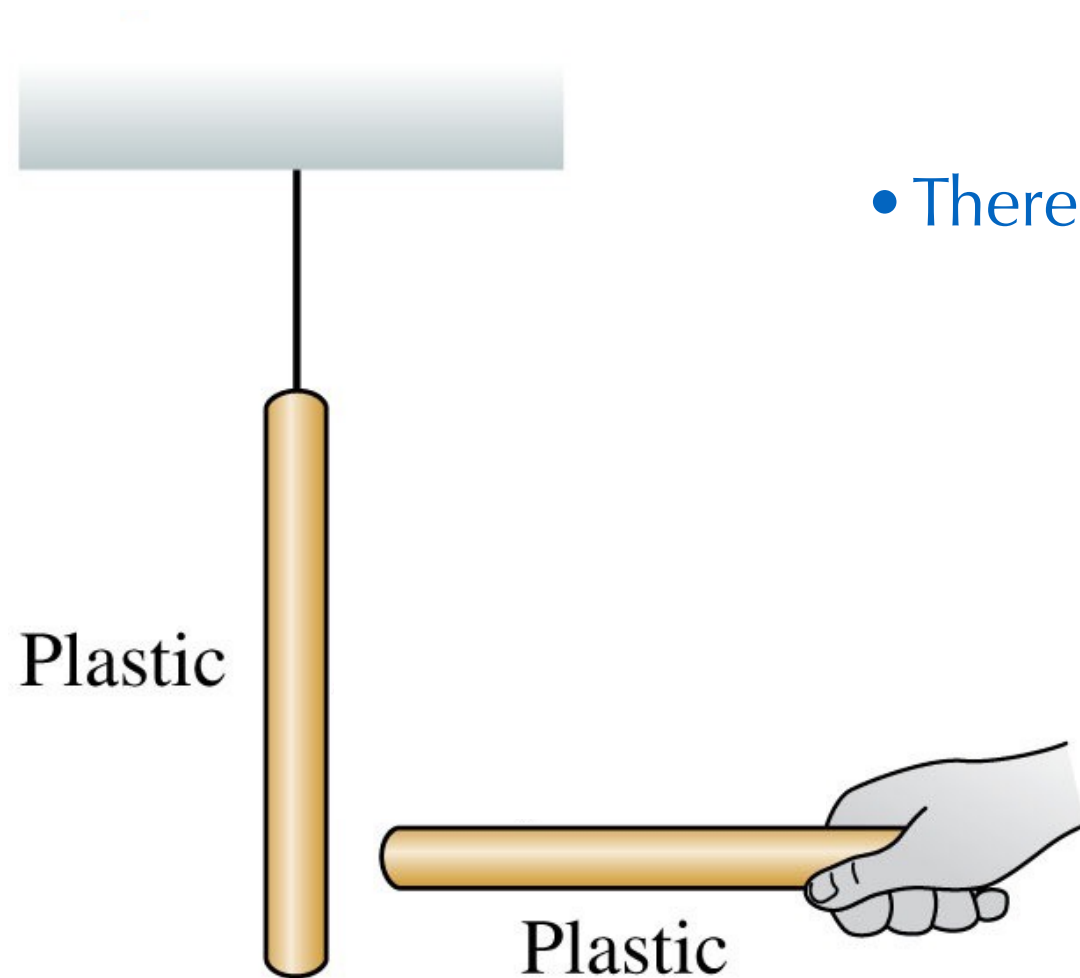


Do these plastic bars have charge?
Why don't they attract/repel one another?

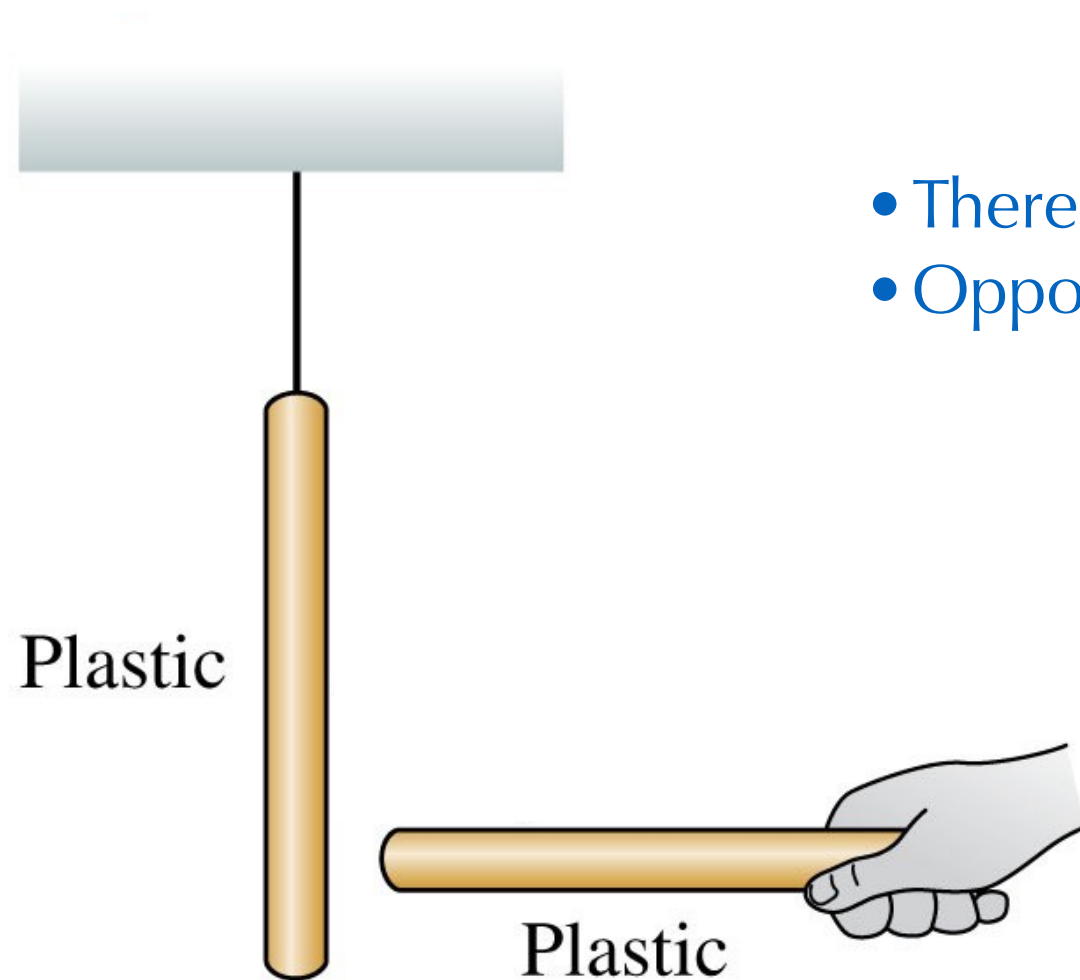


Do these plastic bars have charge?
Why don't they attract/repel one another?

- There are two fundamental charges.

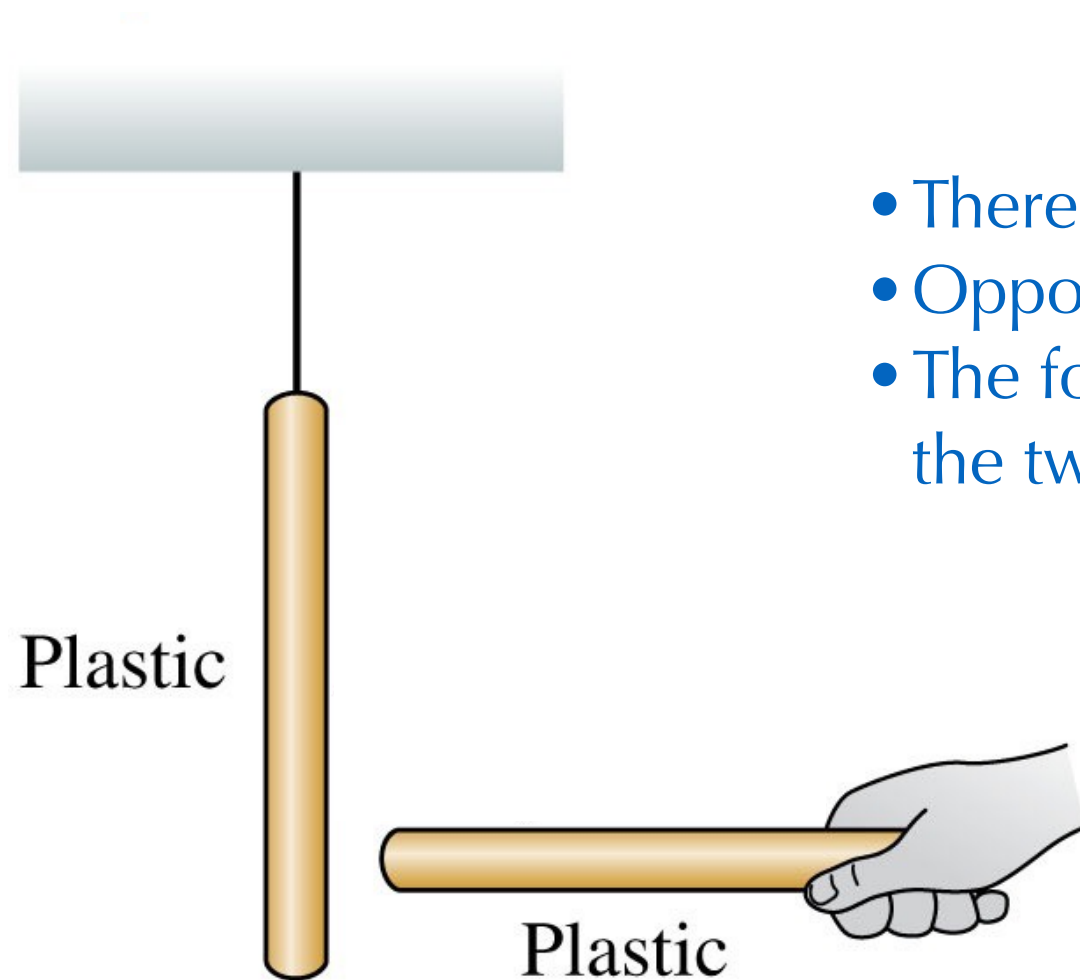


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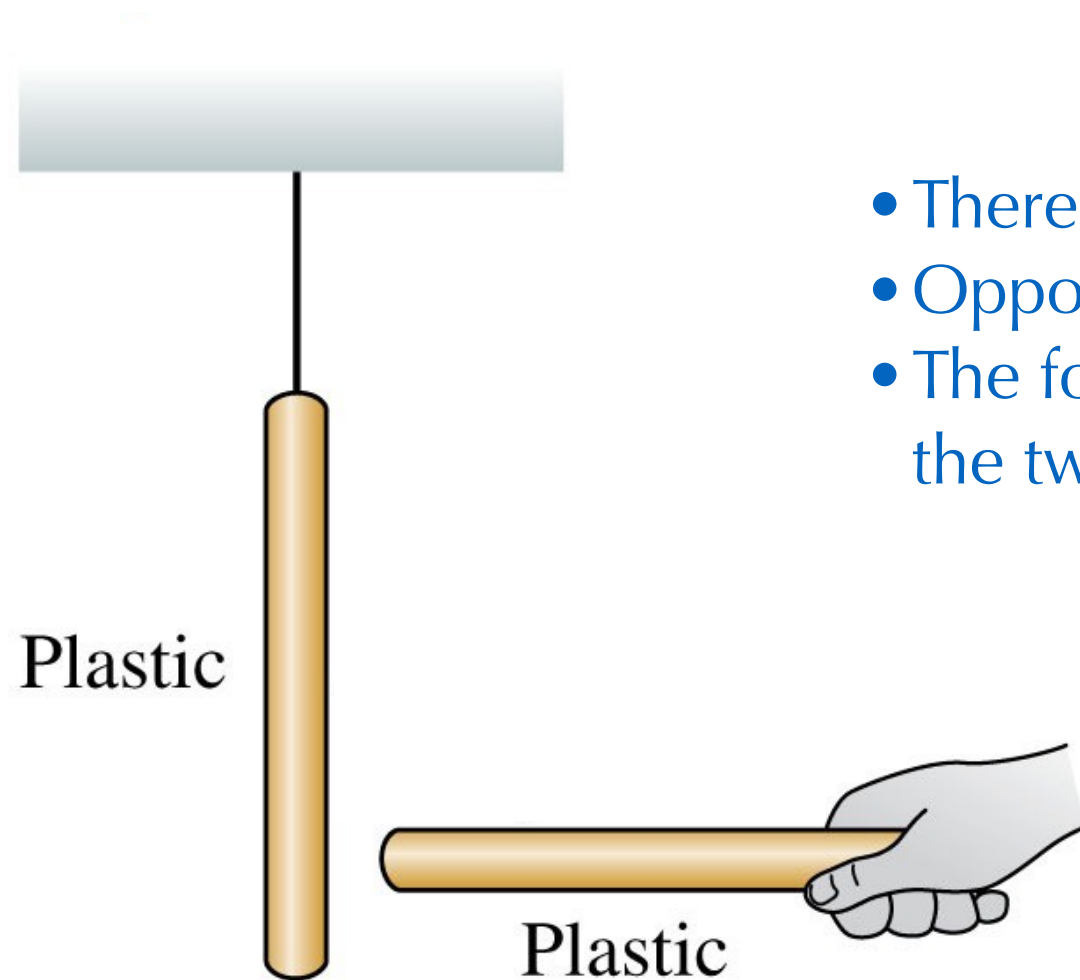
- There are two fundamental charges.
- Opposites attract, likes repel, neutrals do nothing.

Do these plastic bars have charge?
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- There are two fundamental charges.
- Opposites attract, likes repel, neutrals do nothing.
- The force seems to depend on the distance between the two objects.

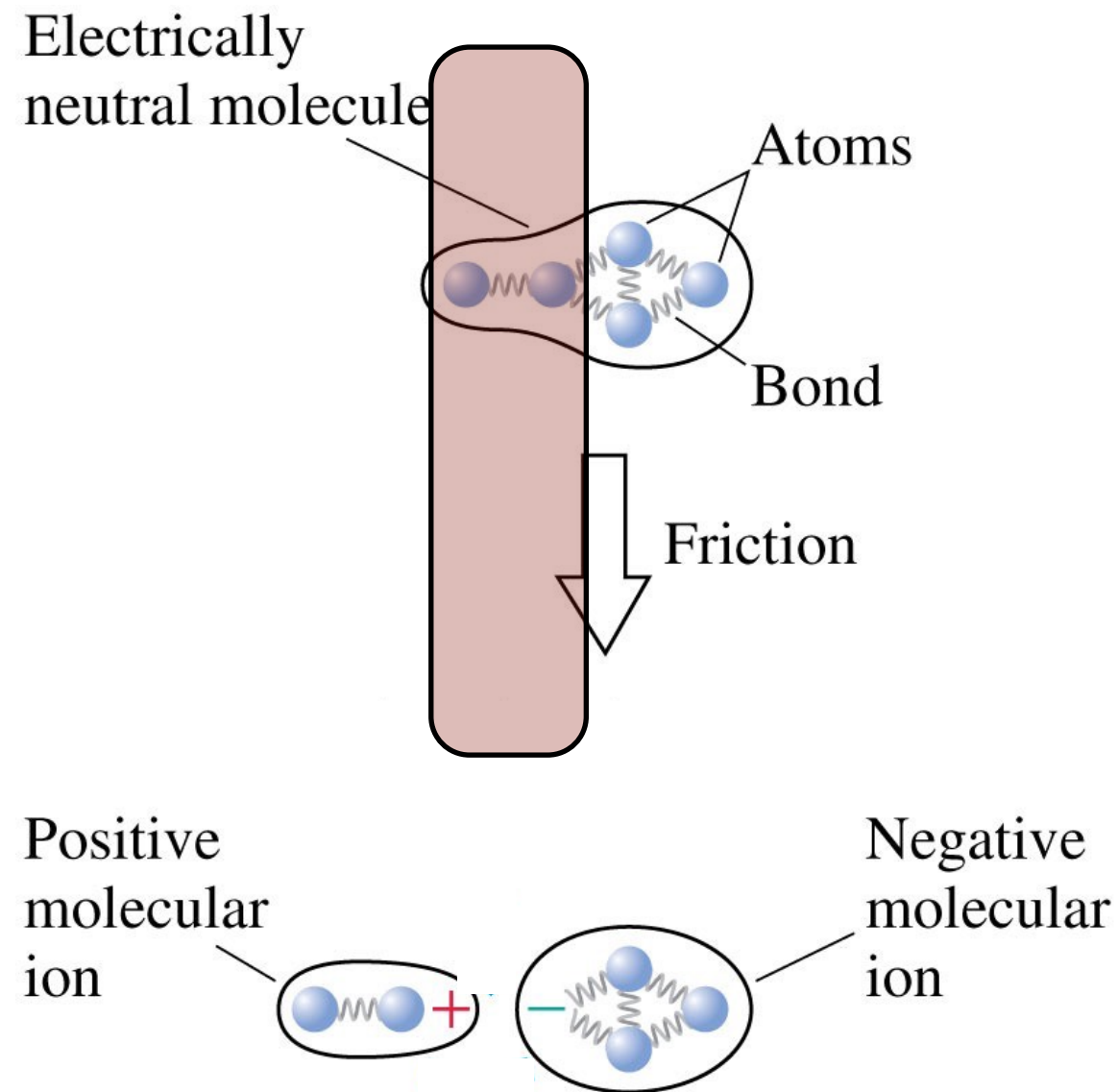
Do these plastic bars have charge?
Why don't they attract/repel one another?



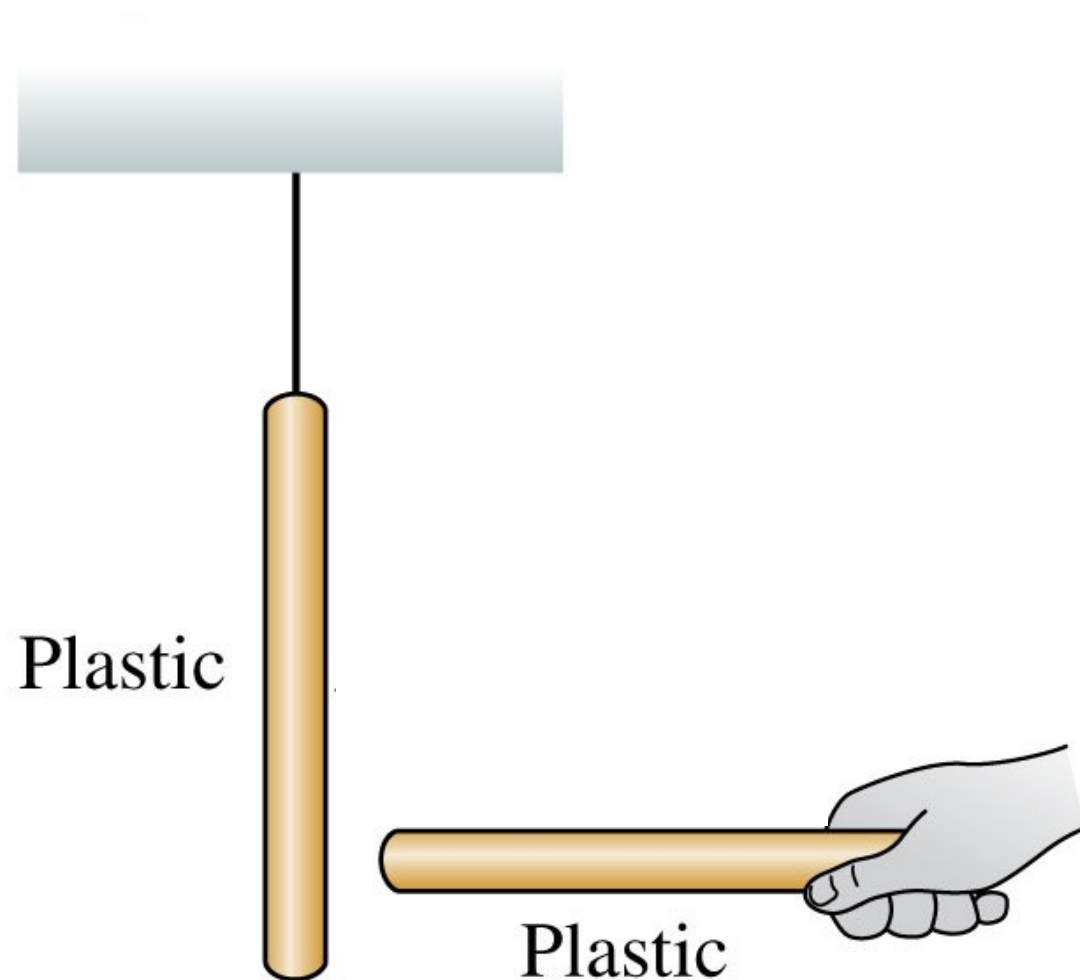
- There are two fundamental charges.
- Opposites attract, likes repel, neutrals do nothing.
- The force seems to depend on the distance between the two objects.

Outstanding question: What does rubbing do?

Triboelectric Effect

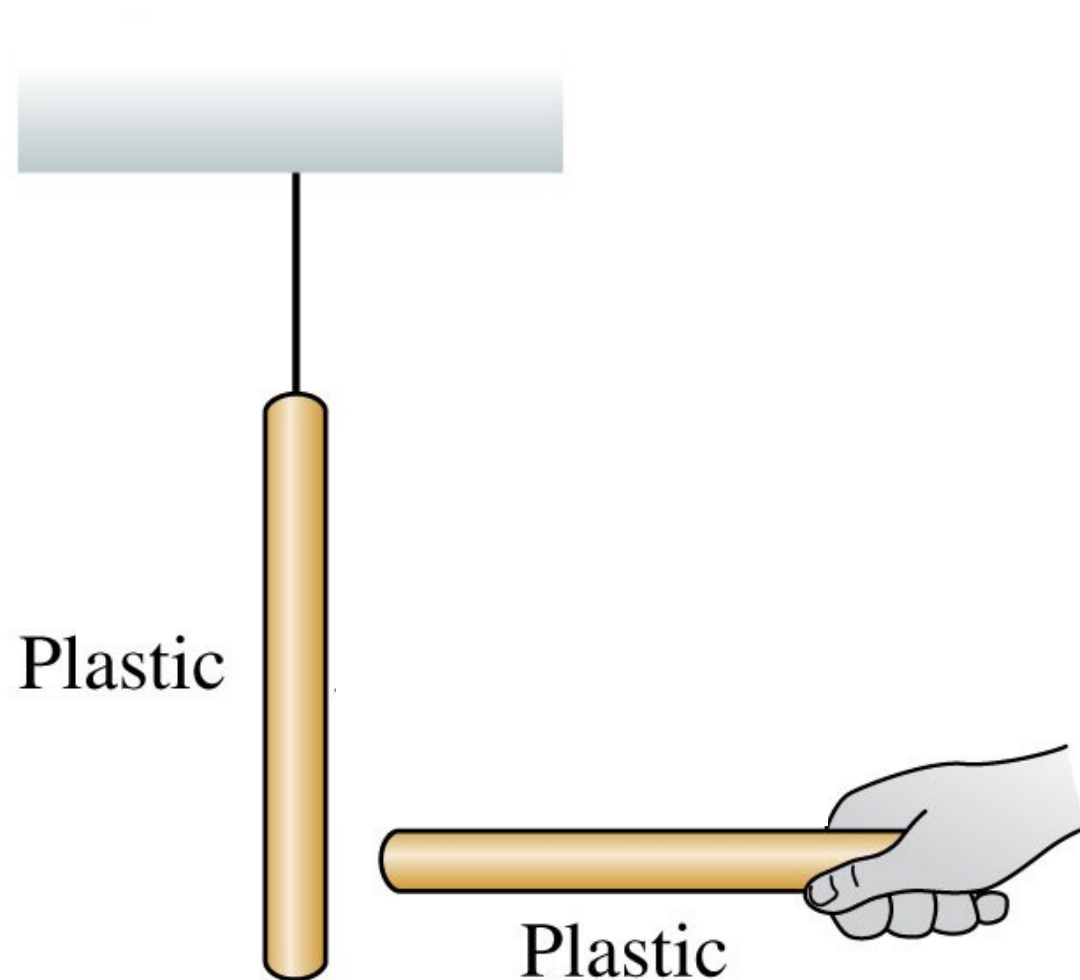


What if I only rub one of the rods?



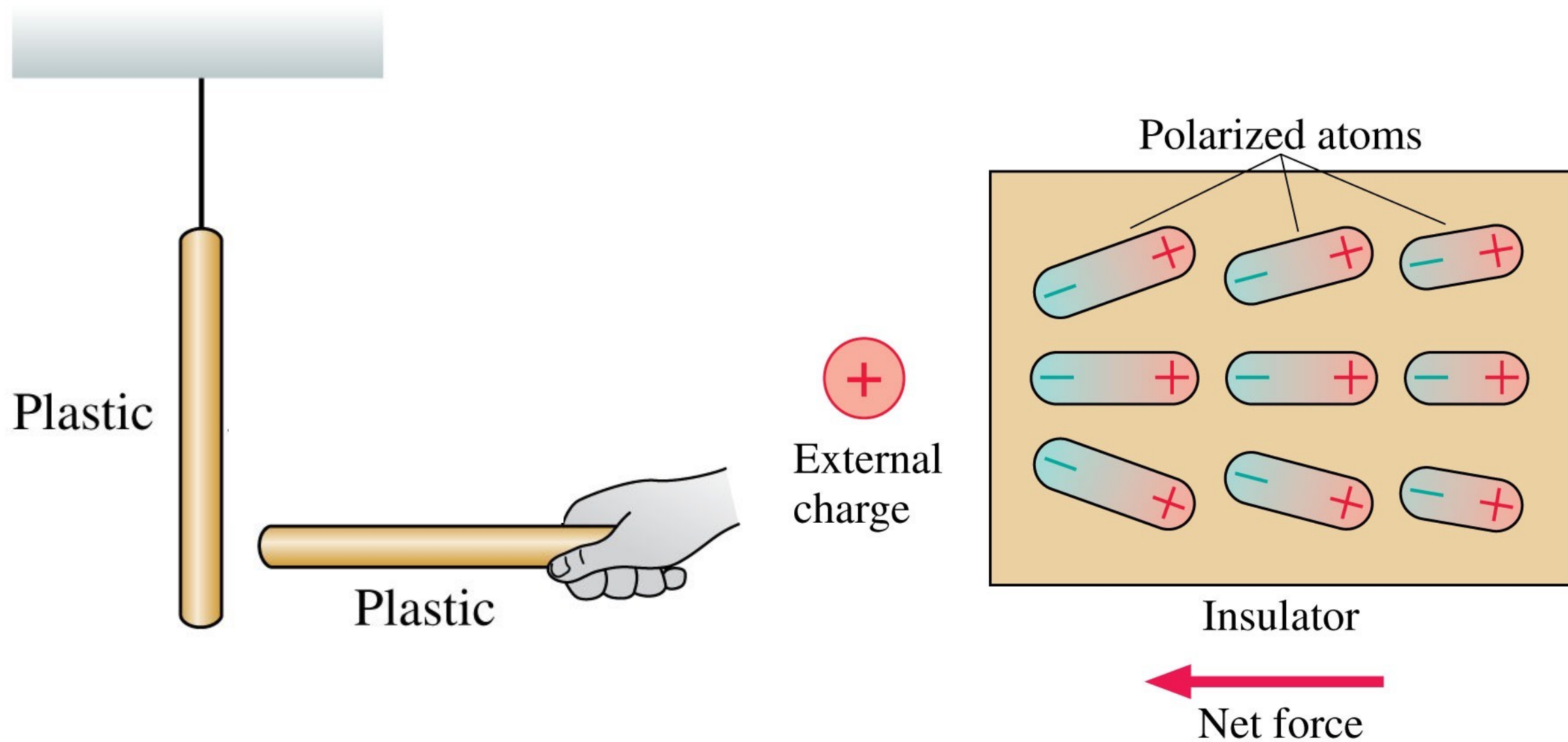
What if I only rub one of the rods?

What if I exchange the plastic rod for a glass rod?



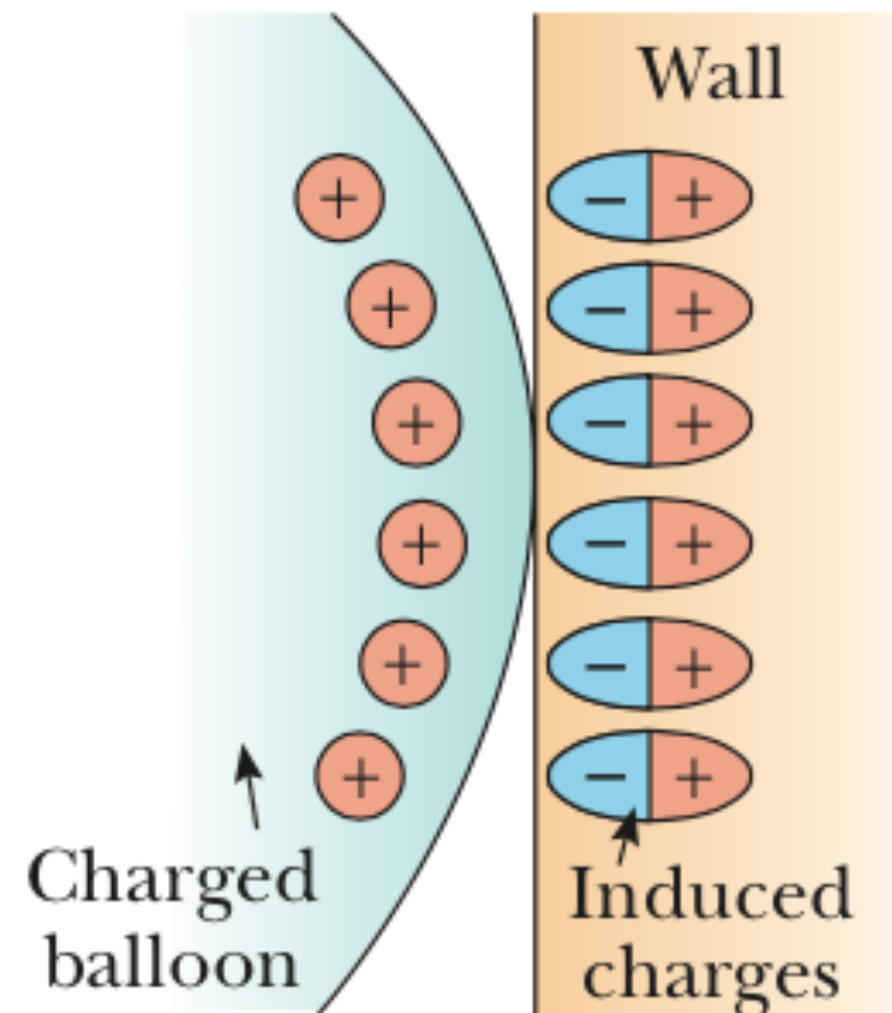
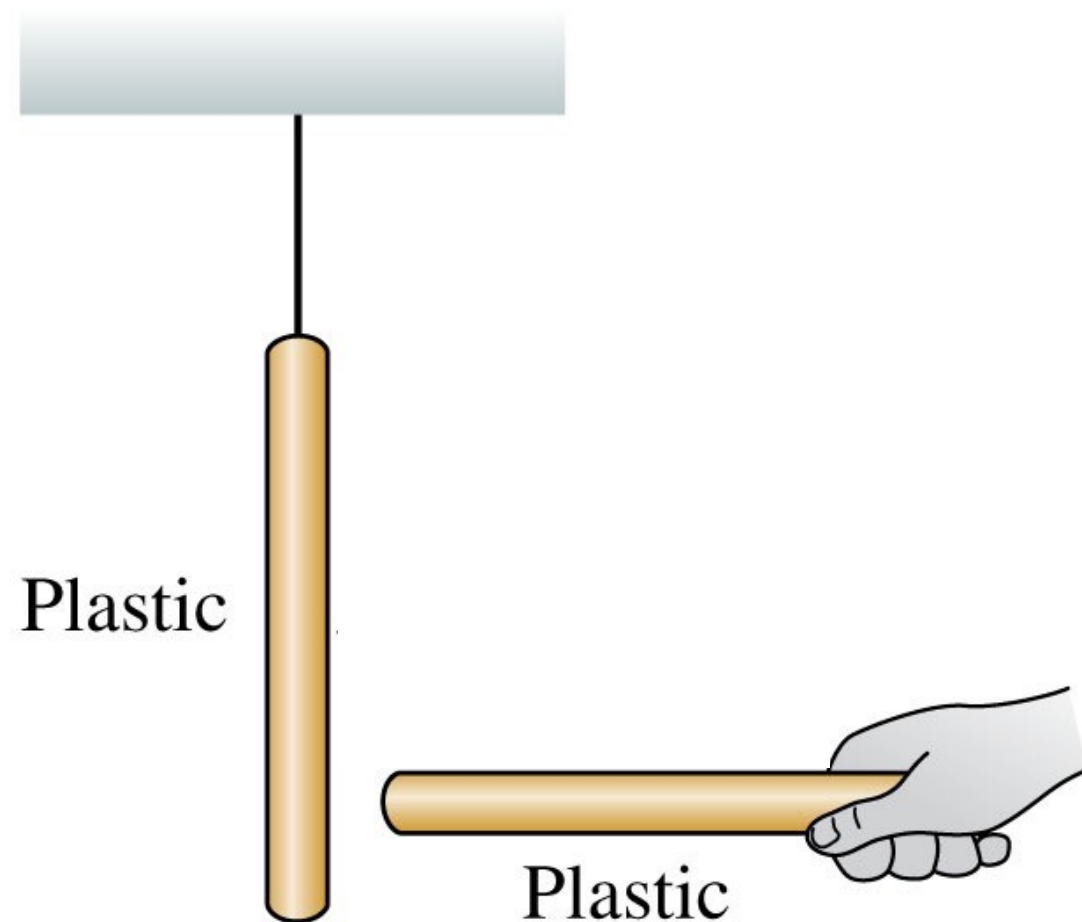
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What if I only rub one of the rods?

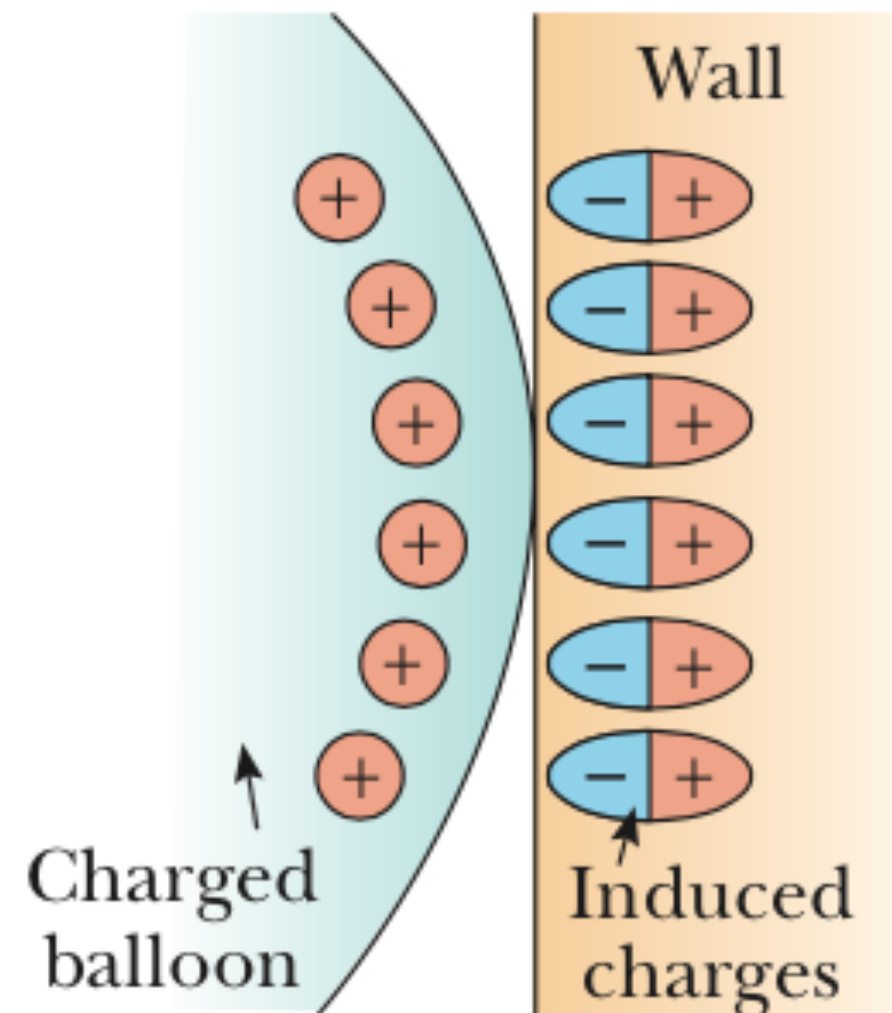
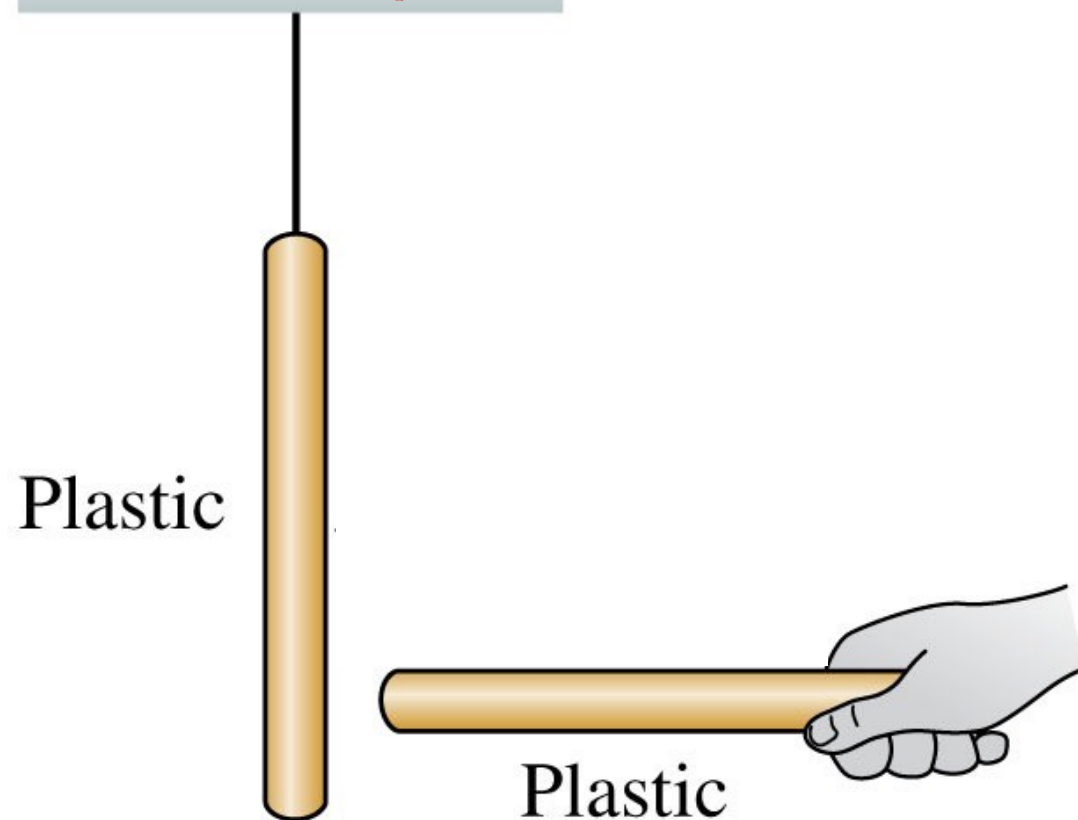
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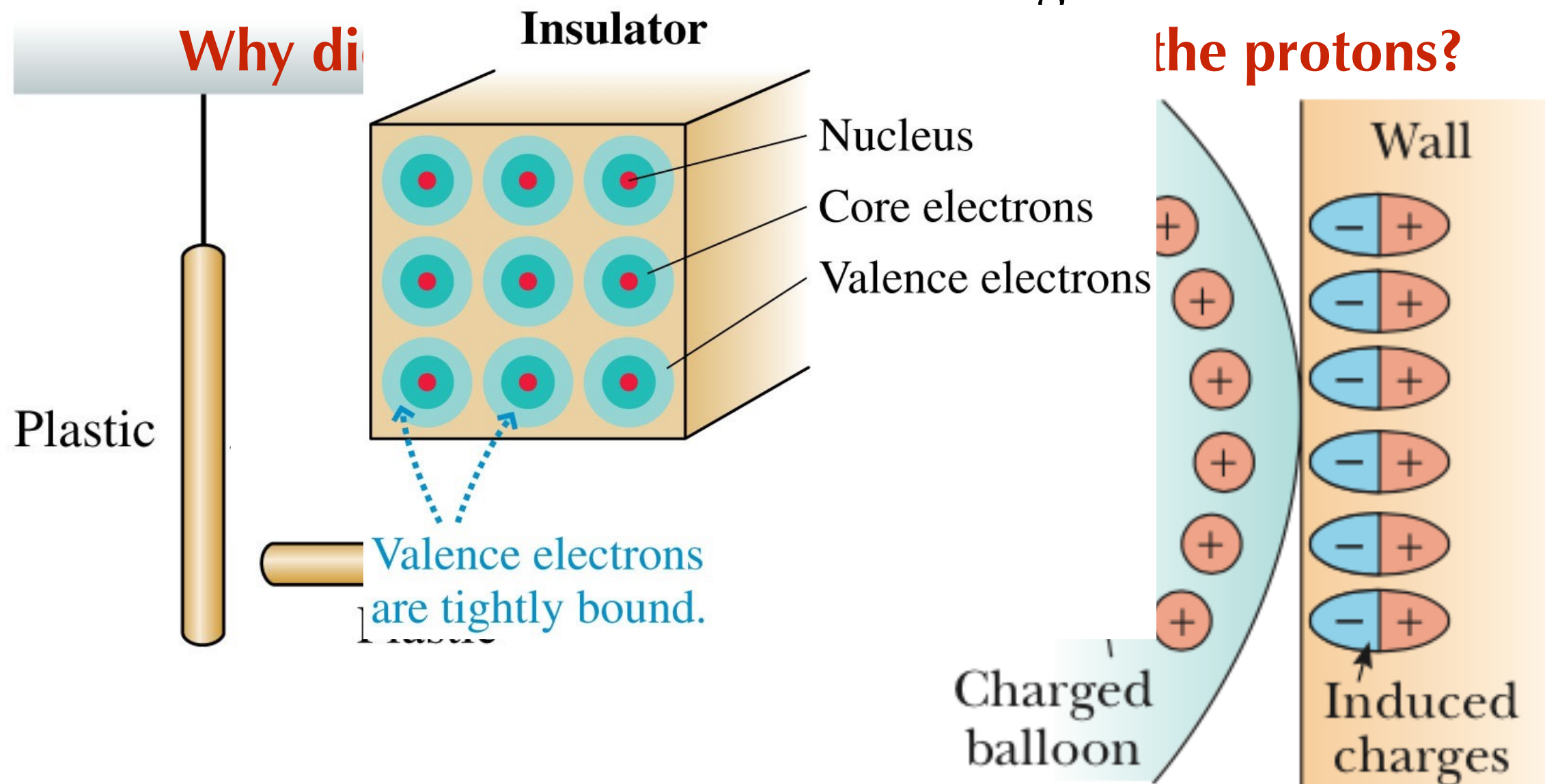
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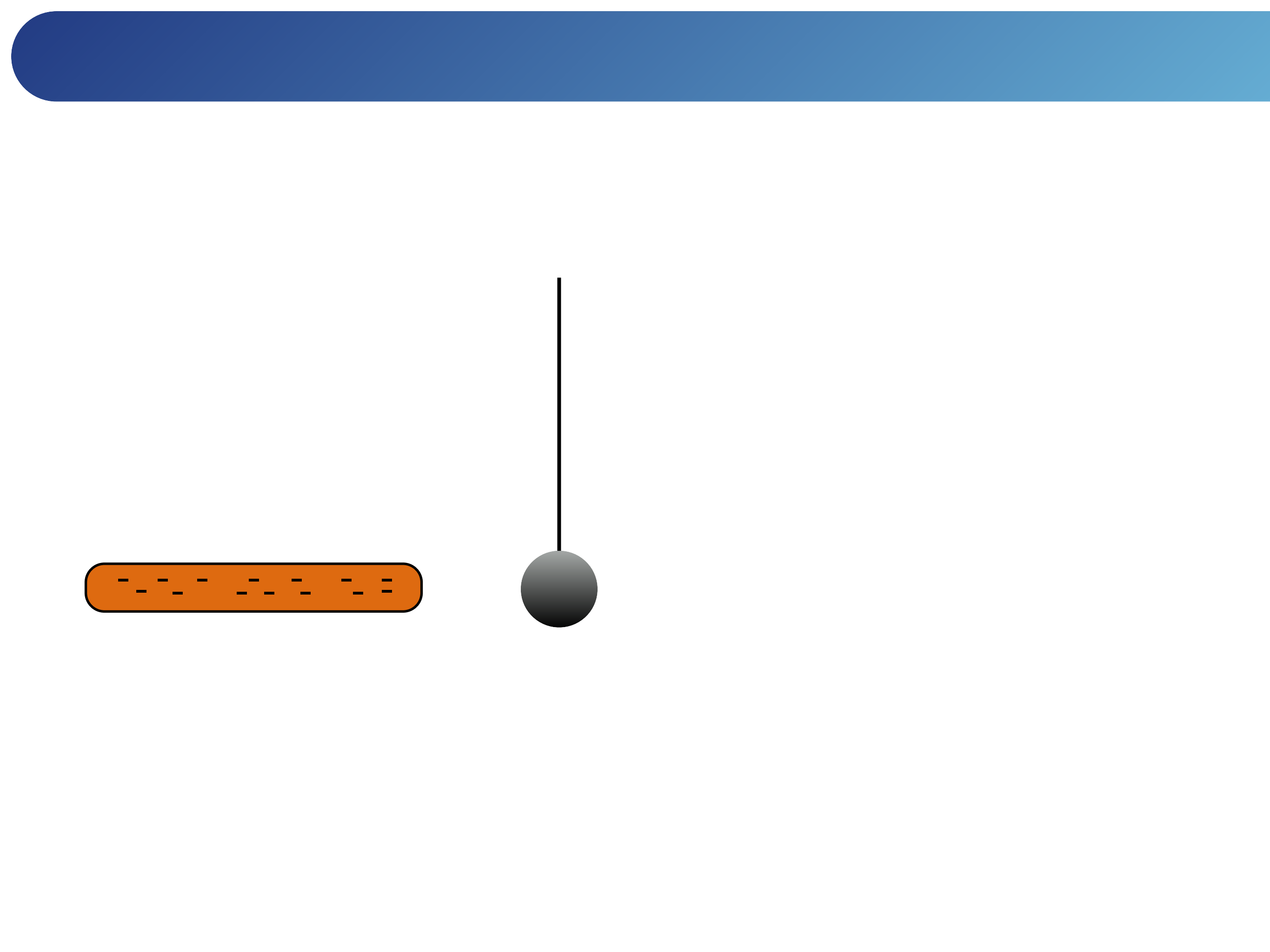
Why didn't the electrons detach from the protons?

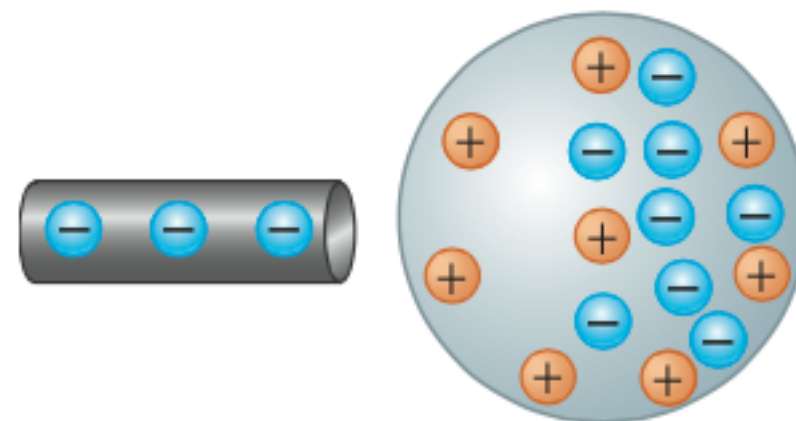
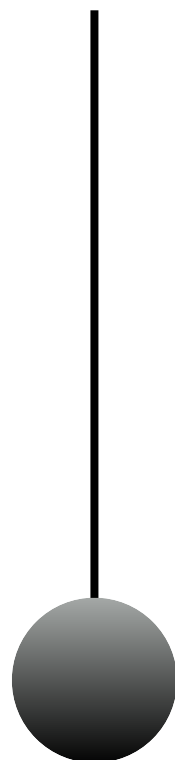
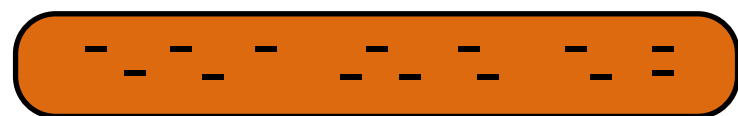


What if I only rub one of the rods?

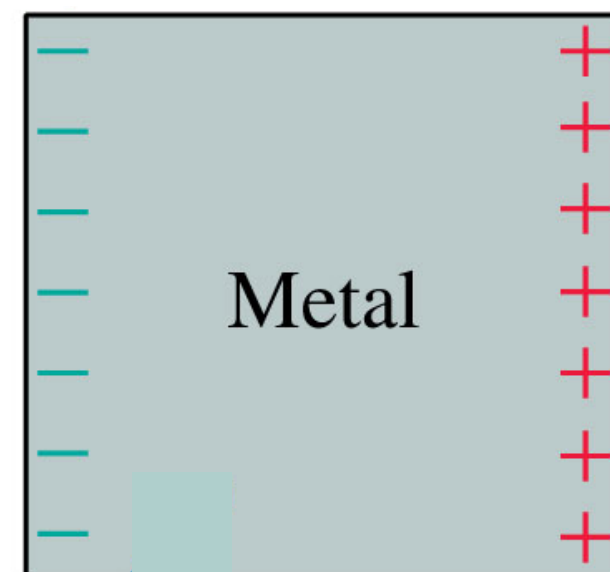
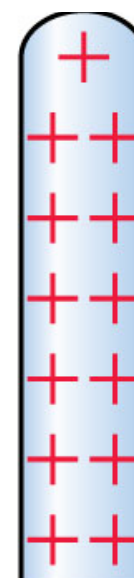
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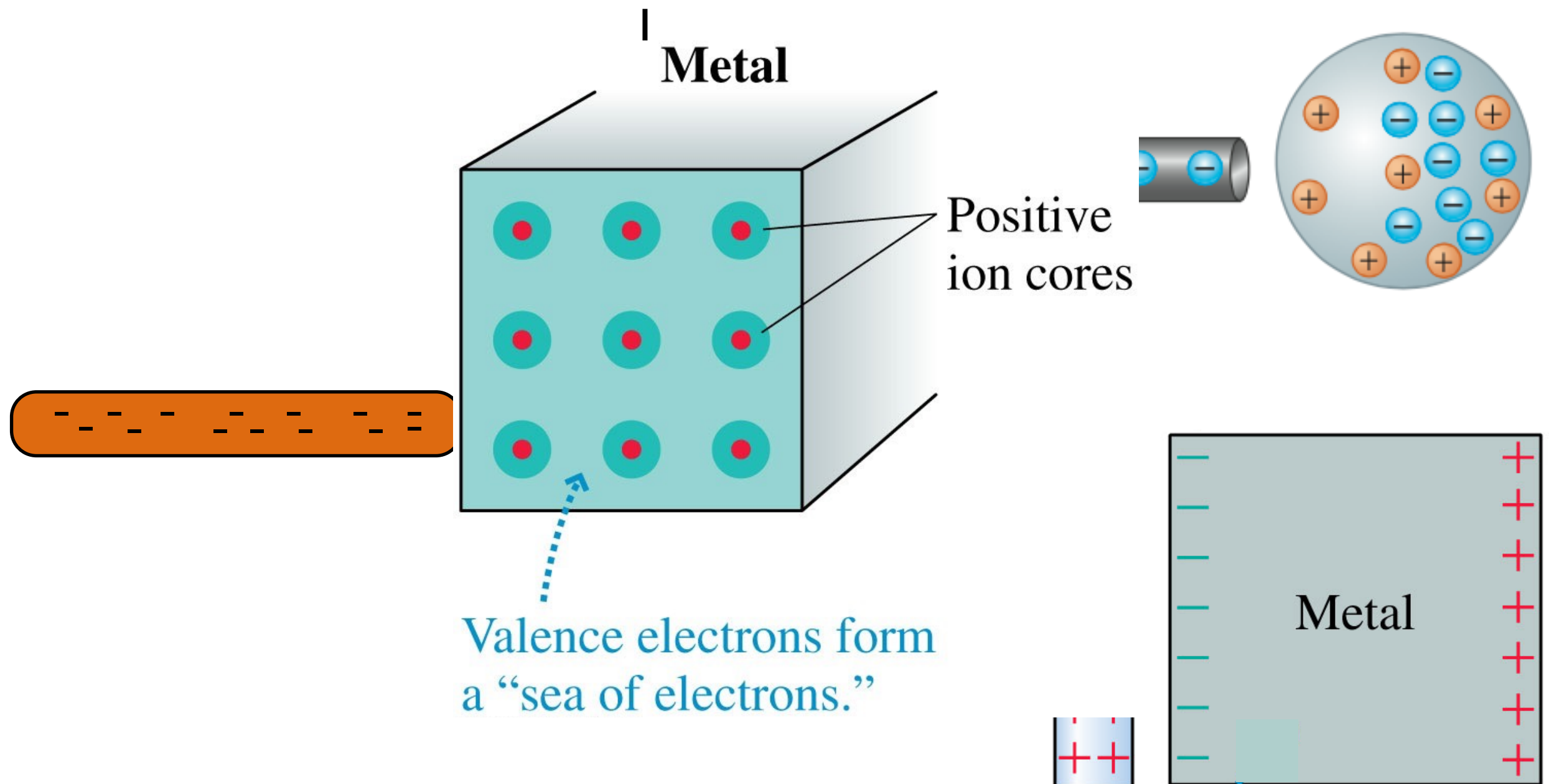






Positive
rod



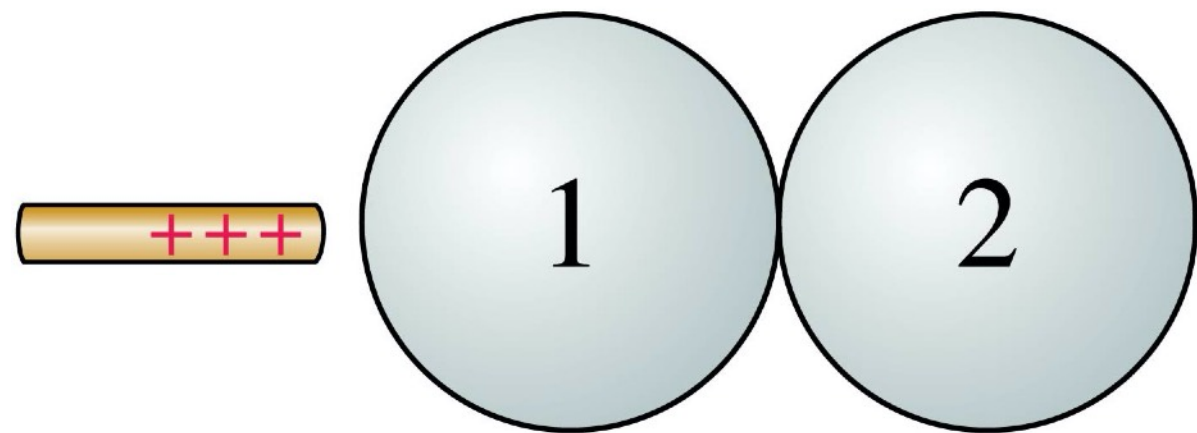


Metal spheres 1 and 2 are touching. Both are initially neutral.

Question #19

- a. The charged rod is brought near.
- b. The spheres are separated.
- c. The charged rod is then removed.

Afterward, the charges on the sphere are:

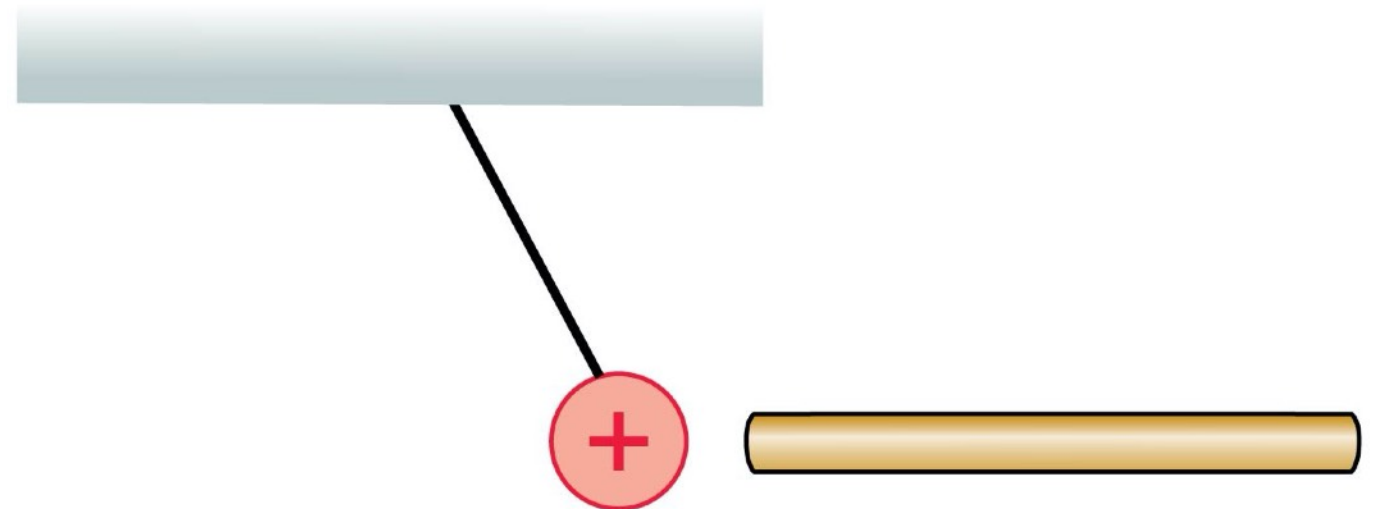


- A. Q_1 is + and Q_2 is +.
- B. Q_1 is - and Q_2 is +.
- C. Q_1 is + and Q_2 is -.
- D. Q_1 is - and Q_2 is -.
- E. Q_1 is 0 and Q_2 is 0.

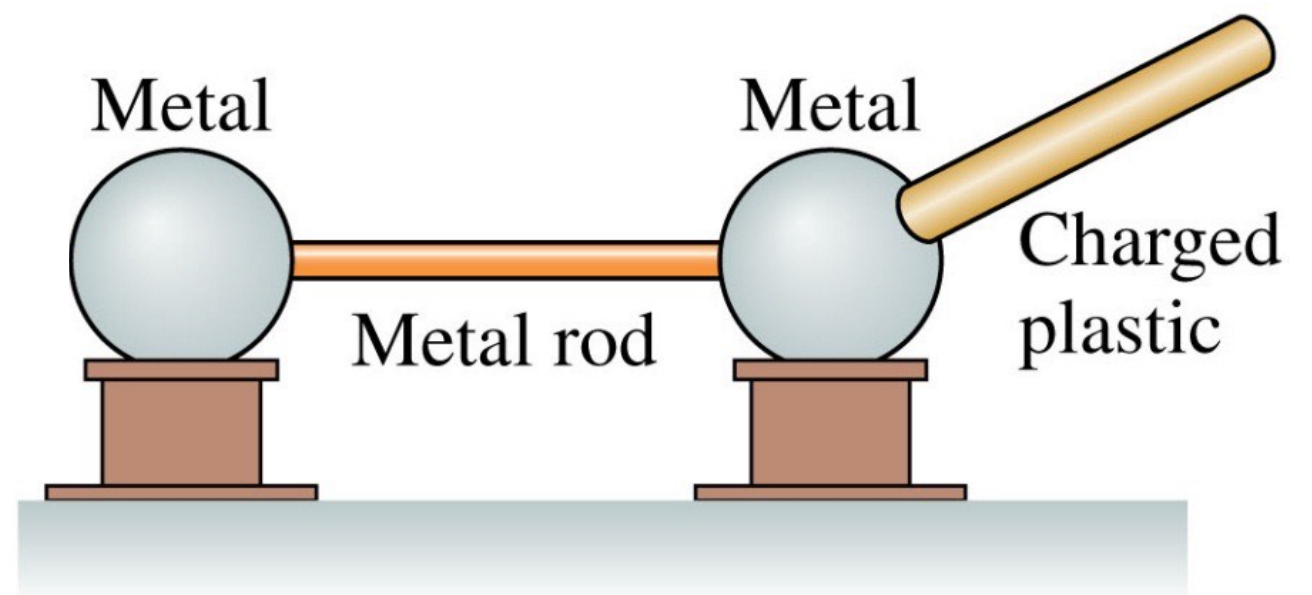
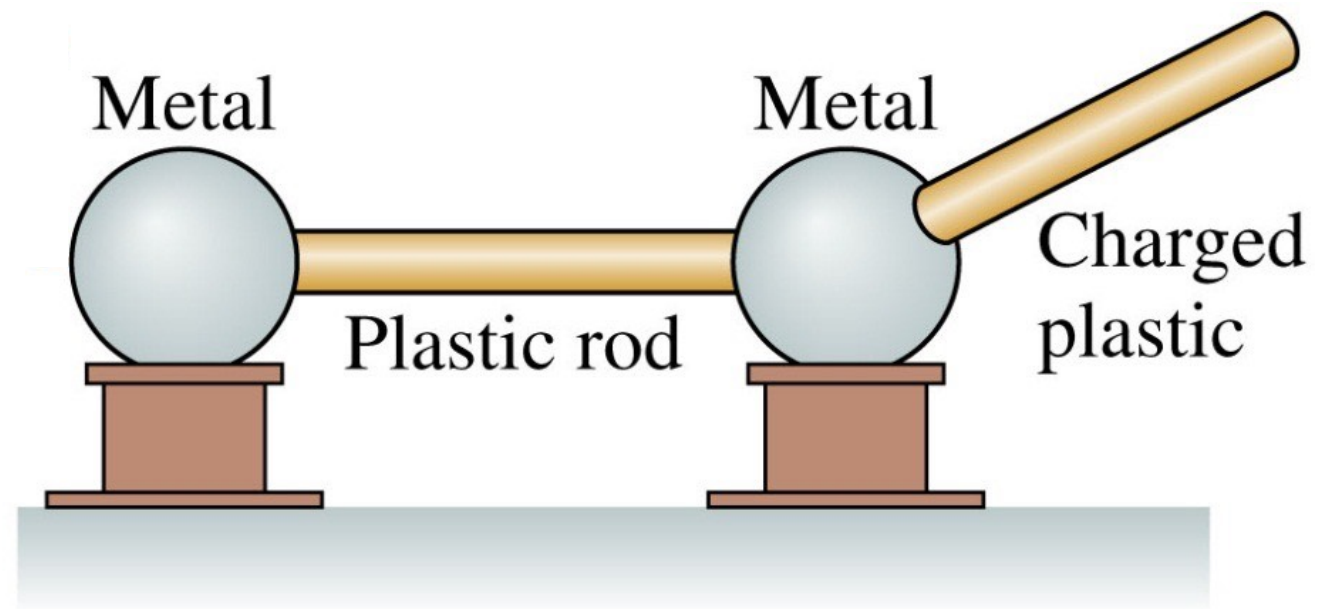
Question #20

A rod attracts a positively charged hanging ball.
The rod is

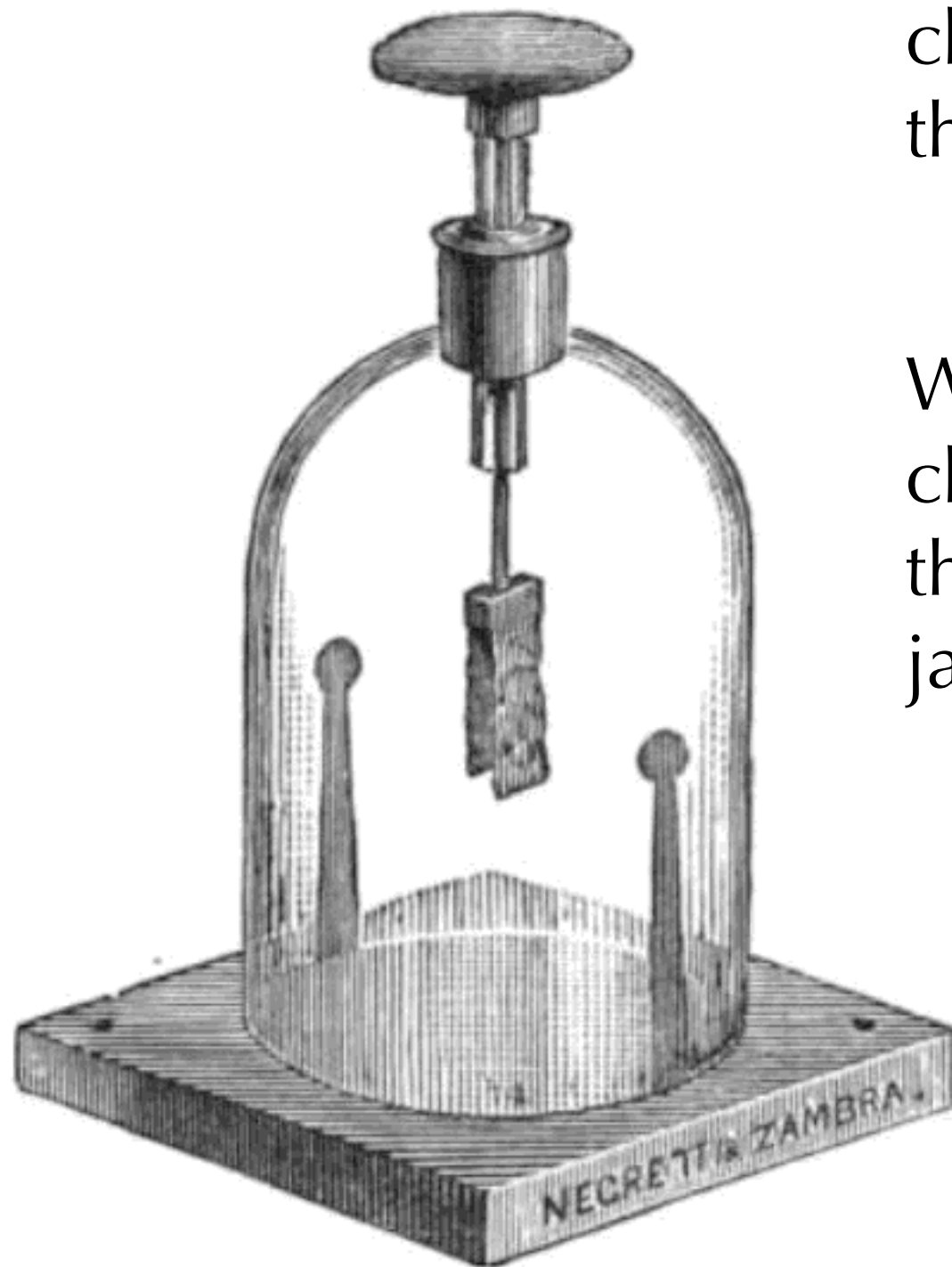
- A. Positive.
- B. Negative.
- C. Neutral.
- D. Either B or C.
- E. Either A or C.



What will be charge state of the metal spheres?



Electroscope

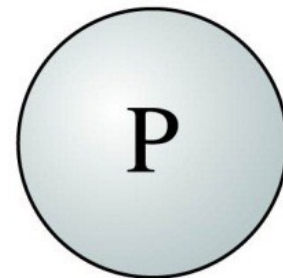


What will happen if I bring a charged object close to the top of the jar?

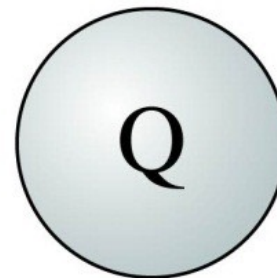
What will happen if I bring a charged object close to the top of the jar while touching the top of the jar?

Identical metal spheres are initially charged as shown. **Question #21**
Spheres P and Q are touched together and then separated.
Then spheres Q and R are touched together and separated.
Afterward the charge on
sphere R is

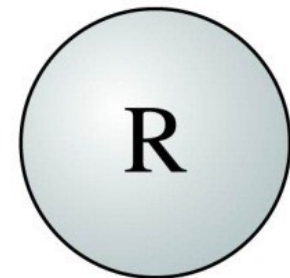
- A. 0 nC.
- B. -1 nC or less.
- C. -0.5 nC.
- D. $+0.5$ nC.
- E. $+1.0$ nC or more.



$+4$ nC



-2 nC



-1 nC

A right cylindrical shell of inner radius a , outer radius b , and height h is shown in the figure. The mass density of the object varies according to the function:

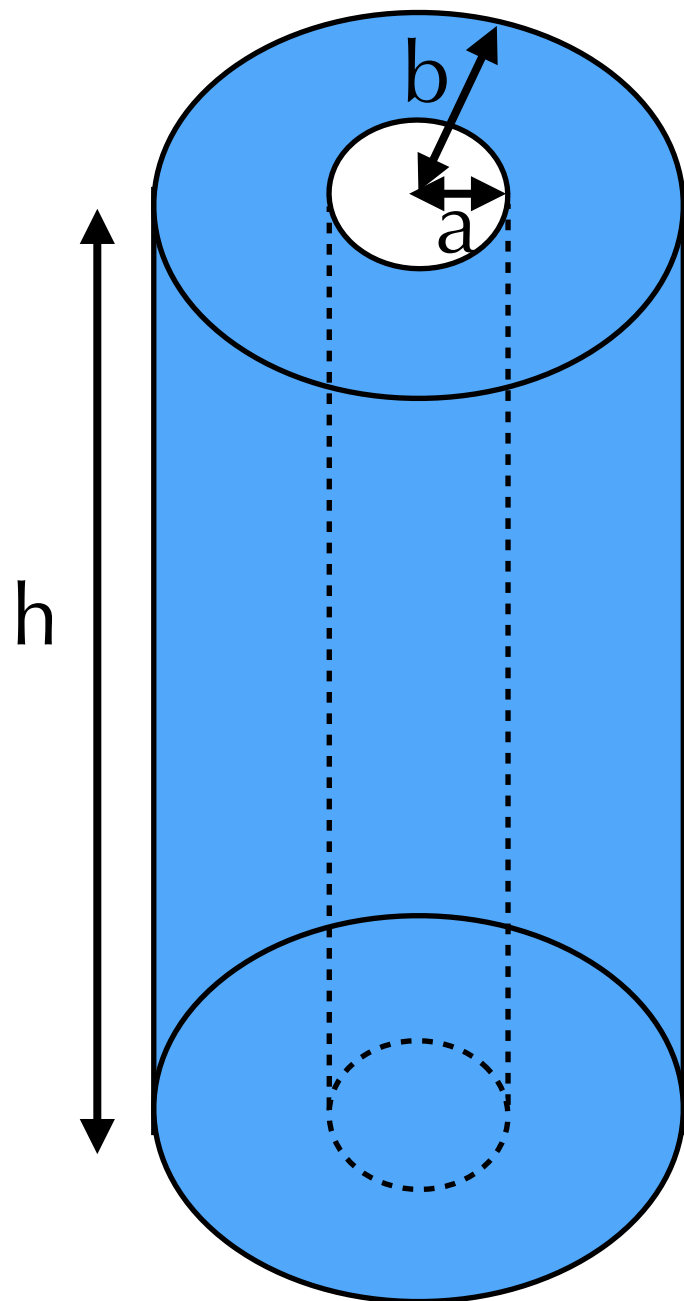
$$\sigma(\rho, \phi) = \rho^2 \sin^2 \phi$$

Find an expression for the mass of the object.

The following trig identity may be useful

$$\sin^2 u = \frac{1 - \cos 2u}{2}$$

Try solving using Mathematica



A right cylindrical shell of inner radius a , outer radius b , and height h is shown in the figure. The mass density of the object varies according to the function:

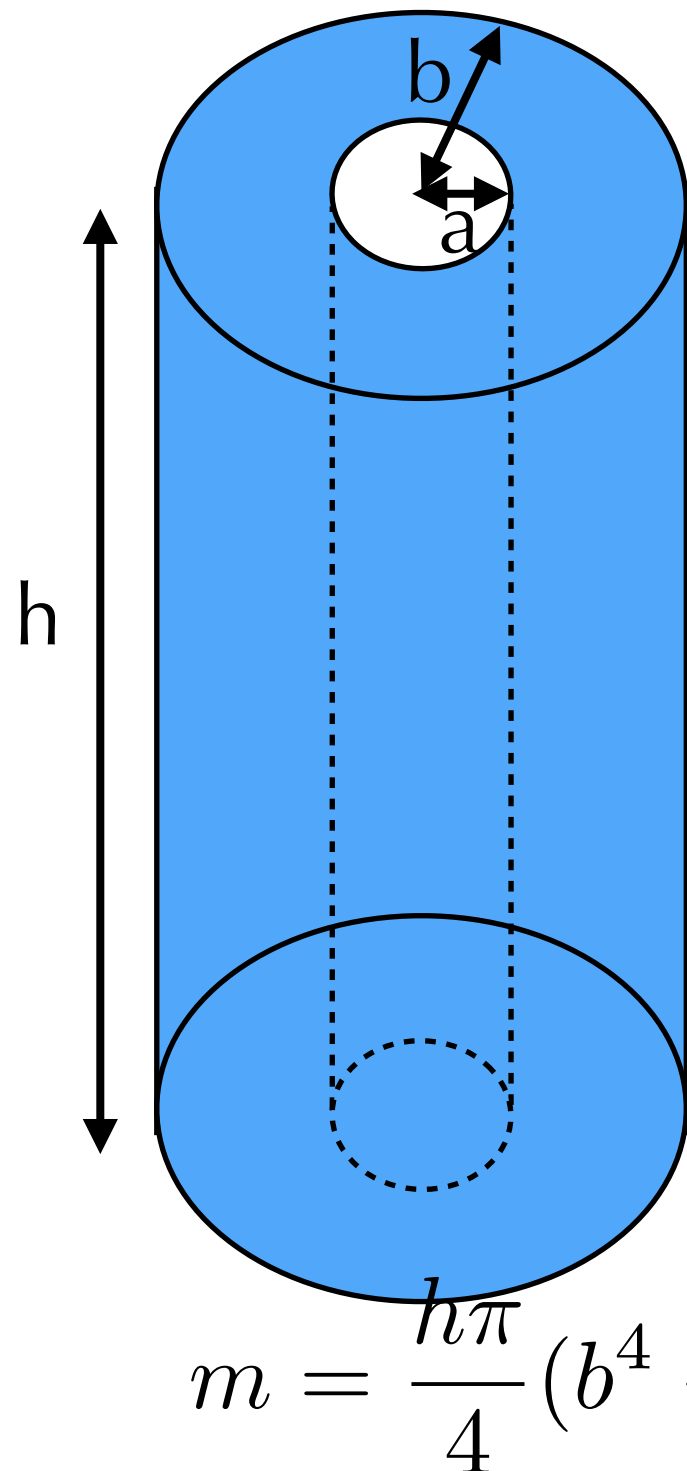
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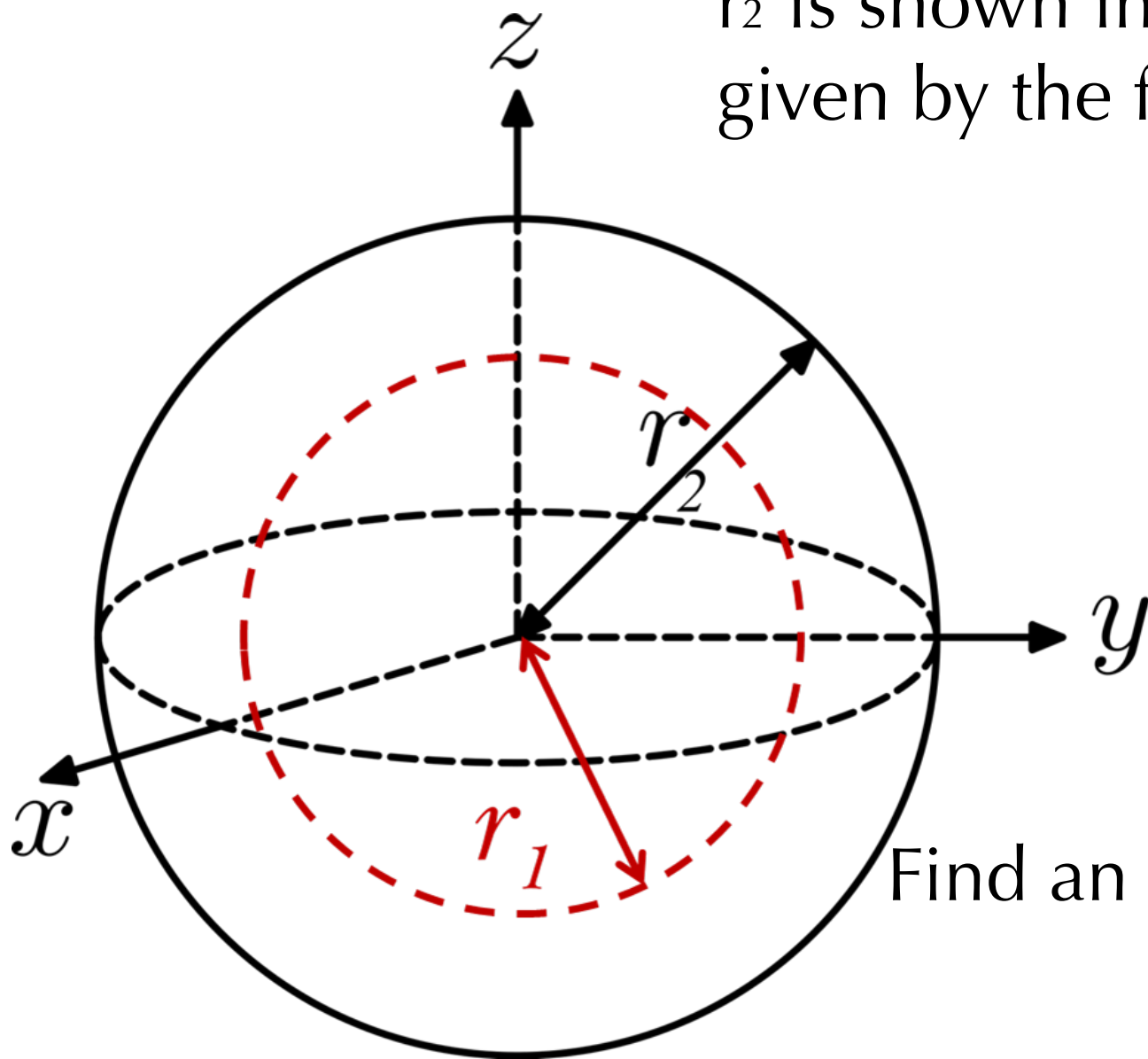
Try solving using Mathematica



$$m = \frac{h\pi}{4}(b^4 - a^4)$$

A spherical shell of inner radius r_1 and outer radius r_2 is shown in the figure. Its density varies and is given by the function:

$$\rho(r, \theta, \phi) = \sqrt{r} \sin^2 \phi$$

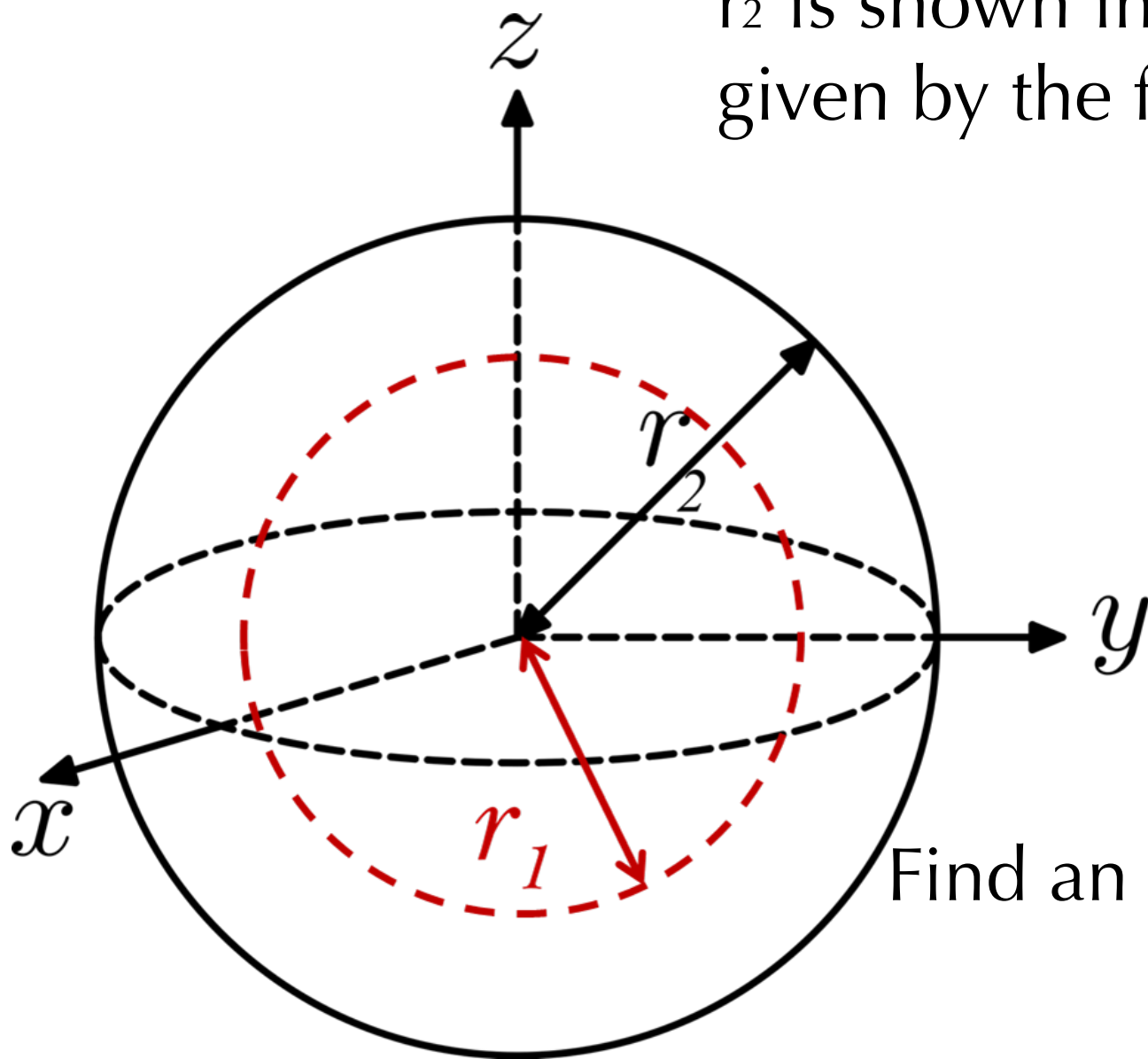


Find an expression for the mass of the object.

to ponder: what is a differential volume element in spherical coordinates.

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Find an expression for the mass of the object.

$$m = \frac{4\pi}{7} (r_2^{7/2} - r_1^{7/2})$$

to ponder: what is a differential volume element in spherical coordinates.