

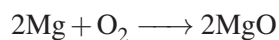
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Alumno 2

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Resumen

Este es un resumen.

§ 1. Propósito. stoichiometry of the reaction (as defined in 1.1):



1.1. Definiciones.—

Stoichiometry The relationship between the relative quantities of substances taking part in a reaction or forming a compound, typically a ratio of whole integers.

Atomic mass The mass of an atom of a chemical element expressed in atomic mass units. It is approximately equivalent to the number of protons and neutrons in the atom (the mass number) or to the average number allowing for the relative abundances of different isotopes.

§ 2. Datos experimentales. Ejemplo de una inclusión de un ambiente *tabular*

Mass of empty crucible	7,28 g
Mass of crucible and magnesium before heating	8,59 g
Mass of crucible and magnesium oxide after heating	9,46 g
Balance used	#4
Magnesium from sample bottle	#1

§ 3. Cálculo. Otro ejemplo *tabular* con mas datos y unidades en sistema internacional (SI).

Mass of magnesium metal	= 8,59 g - 7,28 g
	= 1,31 g
Mass of magnesium oxide	= 9,46 g - 7,28 g
	= 2,18 g
Mass of oxygen	= 2,18 g - 1,31 g
	= 0,87 g

§ 4. Resultados. Un ejemplo de una gráfica en formato .eps

Figura 1: Figure caption.

Un ejemplo de la inclusión de programación en matlab

```
clear all;  
close all;  
clc  
  
syms z n  
  
%%% Ejemplo de la Parte II. inciso 1)  
  
fprintf('Se muestra la transformada Z de la parte II inciso 1 ')
```

```

f1=2^n;
fz1=ztrans(f1)
pretty(fz1)

%%% Ejemplo de Parte II. inciso 2)

fprintf('Se muestra la transformada Z de la parte II inciso 2 ')

f2=(0.5)^n;
fz2=ztrans(f2)
pretty(fz2)

```

§ 5. Conclusiones y discusión. The most obvious source of experimental uncertainty is the limited precision of the balance. Other potential sources of experimental uncertainty are: the reaction might not be complete; if not enough time was allowed for total oxidation, less than complete oxidation of the magnesium might have, in part, reacted with nitrogen in the air (incorrect reaction); the magnesium oxide might have absorbed water from the air, and thus weigh “too much.” Because the result obtained is close to the accepted value it is possible that some of these experimental uncertainties have fortuitously cancelled one another.