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| --- |
| Which among the following represent intercepted in Arrhenius plot |
| Ln A |
| -2.303/k |
| -Ea/R |
| Log10 a |
| The given equation is Arrhenius equation k = A e⁻Ea/RT, Here A is the Arrhenius factor or the frequency factor. It is also called pre – exponential factor. R is a gas constant and Ea is a activation energy measured in joule/mole. |
| In the Arrhenius plot, slope = -Ea/R and intercept = Ln A = Ln k = ln A – EA/RT SO intercept is Ln A |
| In the Arrhenius plot, slope = -Ea/R and intercept = Ln A = Ln k = ln A – EA/RT SO intercept is Ln A |
| Rate of a Chemical Reaction |