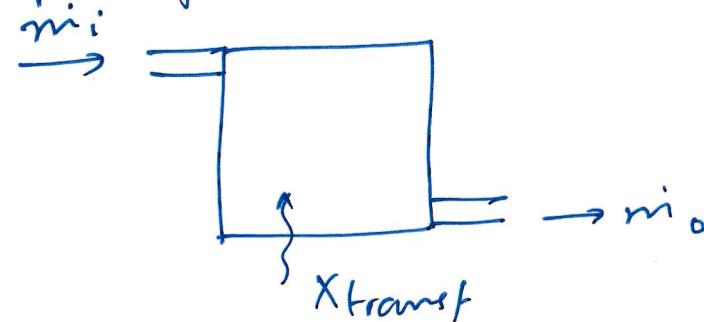


## Open Systems



$$\dot{X}_{in} - \dot{X}_{out} + \dot{X}_{transf} + \dot{X}_{gen} = \frac{\partial}{\partial t} X_{cv} \quad ①$$

Mass Balance:  $m_i - m_o + m_{transf} + m_{gen} = \frac{\partial}{\partial t} m_{cv}$

If no mass transfer at boundaries & no mass generation.

$$m_i - m_o = \frac{\partial}{\partial t} m_{cv}$$

Steady State:  $\dot{m}_i = \dot{m}_o$

Energy:  $m_i(h_i + \frac{V_i^2}{2} + g z_i) - m_o(h_o + \frac{V_o^2}{2} + g z_o) + \dot{Q} - \dot{W} = \frac{\partial}{\partial t} E_{cv}$   
(no energy generation)

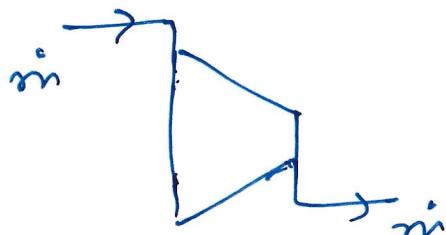
Entropy:  $m_i s_i - m_o s_o + \int \frac{\delta Q_{rev}}{T} + \sum \dot{S} = \frac{\partial S_{cv}}{\partial t}$   
↓ entropy generated

Linear momentum:  $m_i v_i - m_o v_o + \Sigma F = \frac{\partial (m_{cv} v)}{\partial t}$

Angular momentum:  $(r_i \times m_i v_i) - (r_o \times m_o v_o) + \Sigma T = \frac{\partial (r \times m_{cv} v)}{\partial t}$

②

### Compressor:



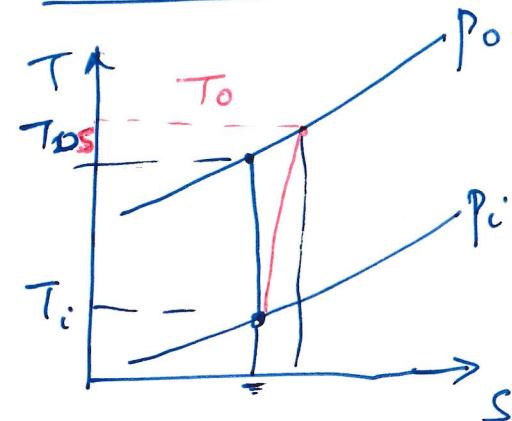
$$\dot{m}(h_i) - \dot{m}h_o + \dot{Q} - \dot{W} = 0$$

Ideal compressor is reversible adiabatic:  $\dot{Q} = 0$

$$\dot{W} = \dot{m}(h_i - h_o)$$

$$w = \frac{\dot{W}}{\dot{m}} = h_i - h_o$$

Ideal gas:  $w = C_p(T_i - T_o)$



$$\left(\frac{T_{os}}{T_i}\right) = \left(\frac{P_o}{P_i}\right)^{\frac{r-1}{r}}$$

Define isentropic efficiency

$$\eta_{is0} = \frac{T_{os} - T_i}{T_o - T_i}$$

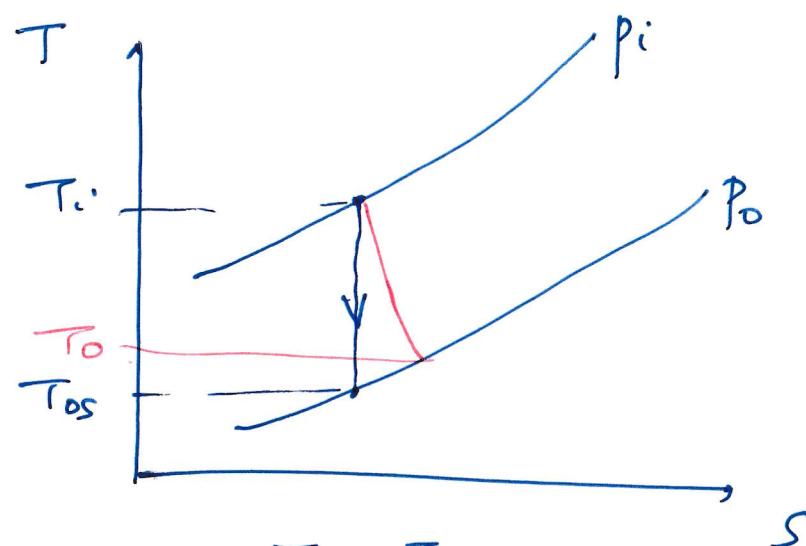
$$\Delta s = C_p \ln \frac{T_o}{T_i} - R \ln \frac{P_o}{P_i}$$

### Turbine:

$$\dot{m}h_i - \dot{m}h_o + \dot{Q} - \dot{W} = 0$$

$$\text{If } \dot{Q} = 0 \Rightarrow \dot{W} = \dot{m}(h_i - h_o)$$

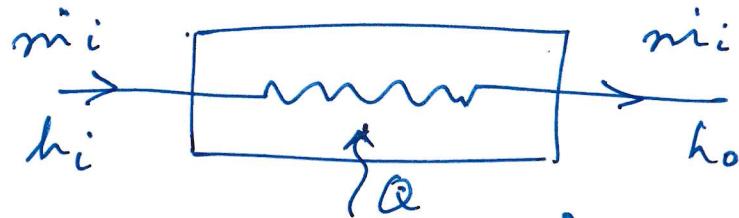
$$w = h_i - h_o.$$



$$\eta_{is0} = \frac{T_i - T_{os}}{T_i - T_o}$$

(3)

## Heat Exchanger



$$m_i h_i - m_i h_o + \dot{Q} = 0$$

$$\dot{Q} = m_i (h_o - h_i)$$

$$q = \frac{\dot{Q}}{m_i} = h_o - h_i$$

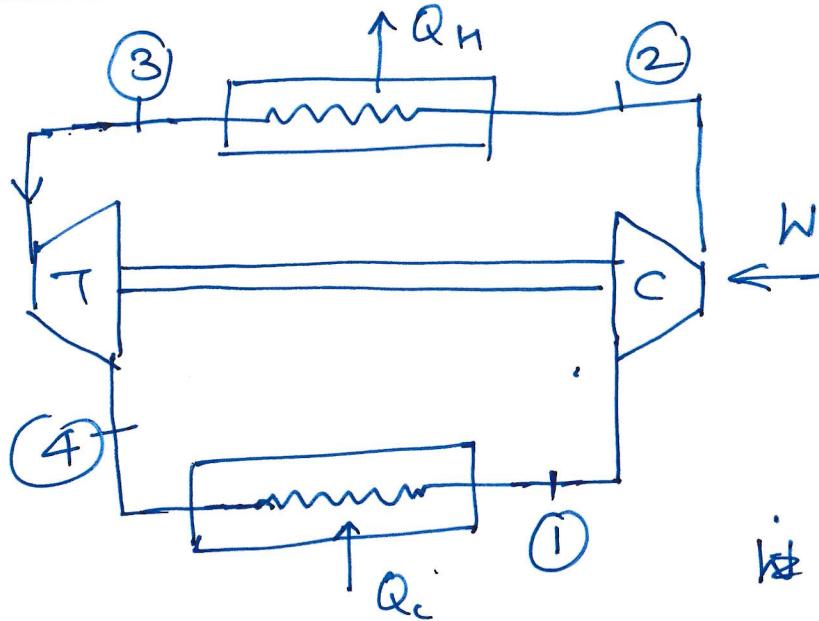
Ideal gas:  $q = \dot{m}_i c_p (\cancel{h_o - h_i})$   
 $= c_p (T_o - T_i)$

$$\Delta S = C_p \ln \frac{T_o}{T_i} - R \ln \frac{P_o}{P_i} \quad (P_i = P_o)$$

$$= C_p \ln \frac{T_o}{T_i}$$

# Reversed Brayton Refrigeration Cycle

(4)



1-2 : Isentropic compression

2-3 Constant pressure heat rejection

$3 \rightarrow 4$  : Isentropic expansion

$4 \rightarrow 1$  : Constant pressure heat addition.

$$\text{at } w_{\text{comp}} = h_2 - h_1$$

$$w_{\text{turbine}} = h_3 - h_4$$

$$w_{\text{net}} = w_{\text{comp}} - w_{\text{turbine}}$$

$$= (h_2 - h_1) - (h_3 - h_4)$$

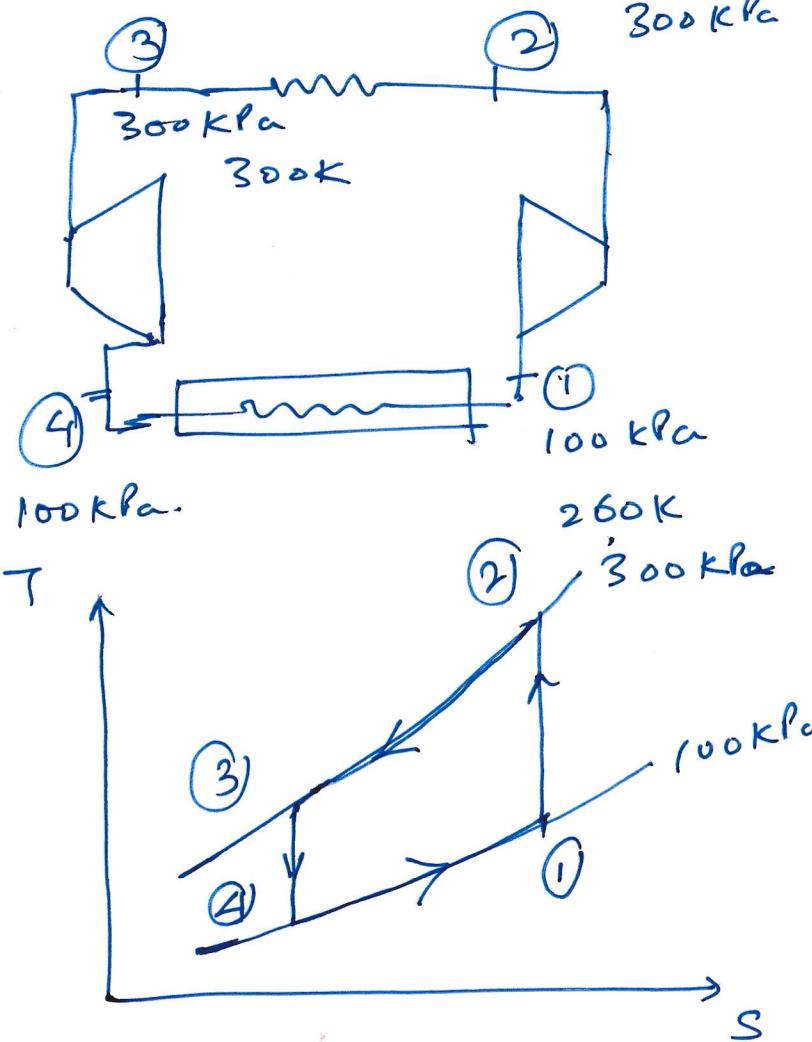
$$q_C = h_1 - h_4$$

$$\text{Coefficient of performance} \quad \text{COP} = \frac{q_C}{w_{\text{net}}} = \frac{h_1 - h_4}{(h_2 - h_1) - (h_3 - h_4)}$$

$$\gamma = 1.4$$

(3)

Air enters the compressor of a Brayton refrigeration cycle at 100 kPa, 260K and is compressed adiabatically to 300 kPa. Air enters the turbine at 300 kPa, 300 K and expands adiabatically to 100 kPa. For the cycle, determine the net work per unit mass of air flow, assuming compressor and turbine isentropic efficiencies are both 100%. Also determine the coefficient of performance. If the cooling load is 200 kW, determine the air flow required and the power required to be supplied to the compressor.



$$\dot{Q} = 200 = \dot{m}(h_1 - h_4) = \dot{m}(41) \Rightarrow \dot{m} = 4.87 \frac{\text{kg}}{\text{s}}$$

$$w_{\text{comp}} = h_2 - h_1 = Cp(T_2 - T_1)$$

$$\frac{T_2}{T_1} = \left(\frac{P_2}{P_1}\right)^{\frac{\gamma-1}{\gamma}} \Rightarrow \frac{T_2}{260} = \left(\frac{300}{100}\right)^{\frac{0.4}{1.4}} \Rightarrow T_2 = 355.9 \text{ K}$$

$$w_{\text{comp}} = (1.005)(355.9 - 260) = 96.4 \text{ kJ/kg}$$

$$w_{\text{turbine}} = h_3 - h_4 = Cp(T_3 - T_4)$$

$$\frac{T_3}{T_4} = \left(\frac{P_3}{P_4}\right)^{\frac{\gamma-1}{\gamma}} \Rightarrow \frac{T_3}{219.2} = \left(\frac{300}{100}\right)^{\frac{0.4}{1.4}}$$

$$\Rightarrow T_4 = 219.2 \text{ K}$$

$$w_{\text{turbine}} = (1.005)(300 - 219.2) = 81.2 \text{ kJ/kg}$$

$$w_{\text{net}} = w_{\text{comp}} - w_{\text{turb}} = 96.4 - 81.2 = 15.2 \frac{\text{kJ}}{\text{kg}}$$

$$\text{COP} = \frac{q_c}{w_{\text{net}}} = \frac{41}{15.2} = 2.7$$

$$q_c = h_1 - h_4 = Cp(T_1 - T_4) = (1.005)(260 - 219.2) = 41 \text{ kJ/kg}$$

$$P = \dot{m} w = (4.87)(15.2) = 74.1 \text{ kW}$$