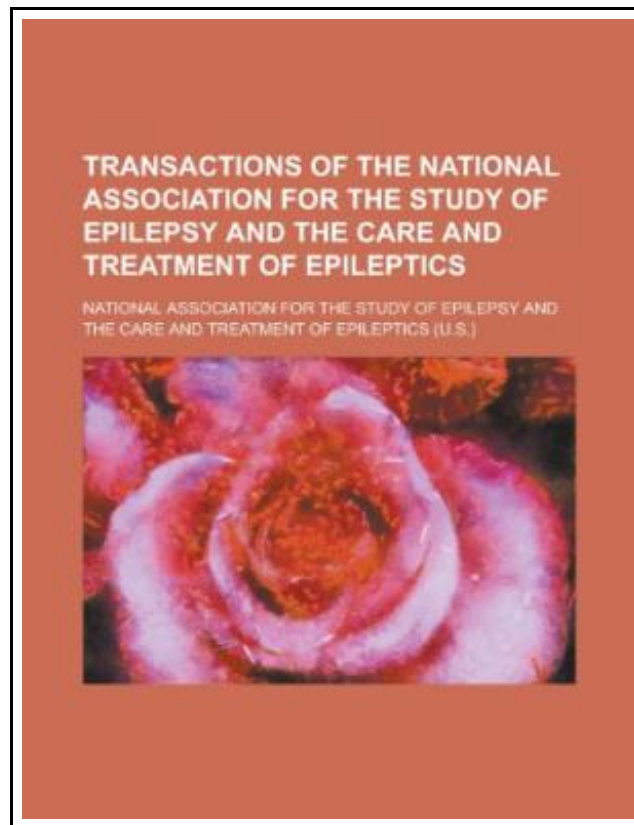


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An extremely wonderful publication with lucid and perfect reasons. It typically will not expense too much. You are going to like the way the blogger compose this publication.

(Prof. Maya Hand)

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RareBooksClub. Paperback. Book Condition: New. This item is printed on demand. Paperback. 30 pages. Original publisher: Aberdeen Proving Ground, MD : Army Research Laboratory, 2007 OCLC Number: (OCoLC)489487226 Subject: Surface active agents. Excerpt: . . . The total free energy of adsorption G can be expressed as the sum of the following free ads energy contributions: , (1) G G G G G G ads elec chem c c c s H bonding vdw G where G 10 kJ mol is the electrostatic interaction term, 25 kJ mol is due to the elec chem G G covalent bonding, 1 kJ mol chain segment is the lateral chain-chain interaction, c c c s 1 kJ mol chain segment is the interaction between hydrophobic chains and hydrophobic surface G sites, 10 kJ mol is the hydrogen bonding interaction term, and 0. 1 kJ mol is G vdw H bonding the term due to van der Waals forces. The goal in surfactant design is to tailor the distribution of functional groups to promote hydrogen bonding and electrostatic attraction to the particle surface. 5. 2 Acid Base Behavior of Functional Groups As described next, the acid base behavior of a particular functional group depends on the polarizability of the bonding, its ability to form hydrogen bonds with the surface and solvent, and the steric environment surrounding the group. 5. 2. 1 Polarizability The ionization potential is a measure of the amount of energy required to remove an electron from the molecule. The highest ionization energies are for molecules that have a symmetrical charge distribution, such as methane and ethane. As the electronegativity difference between atoms increases, the charge distribution becomes more asymmetrical, and the ionization potential decreases. Nonbonding electrons, such as those on N and O, and the bonding electrons of...



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