

باقی رہے ہیں



$$? \text{ mol O}_2 \text{ s } 1/2 \text{ mol } \text{C}_9\text{H}_{12}\text{O}_4 \times \frac{9 \text{ mol O}_2}{1 \text{ mol } \text{C}_9\text{H}_{12}\text{O}_4} = 4.5 \text{ mol O}_2$$

$$? \text{ CO}_2 \text{ s } 1/2 \text{ mol O}_2 \times \frac{11.1 \text{ CO}_2}{1 \text{ mol O}_2} = 5.55 \text{ CO}_2$$

$$? \text{ H}_2\text{O} \text{ s } 1/2 \text{ mol O}_2 \times \frac{11.1 \text{ H}_2\text{O}}{1 \text{ mol O}_2} = 5.55 \text{ H}_2\text{O}$$

$$? \text{ g H}_2\text{O} \text{ s } 1/2 \text{ mol } \text{C}_9\text{H}_{12}\text{O}_4 \times \frac{9 \text{ mol CO}_2}{1 \text{ mol } \text{C}_9\text{H}_{12}\text{O}_4} \times \frac{18 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 90 \text{ g H}_2\text{O}$$

$$? \text{ CO}_2 \text{ s } 1/2 \text{ mol } \text{C}_9\text{H}_{12}\text{O}_4 \times \frac{9 \text{ mol CO}_2}{1 \text{ mol } \text{C}_9\text{H}_{12}\text{O}_4} \times \frac{11.1 \text{ CO}_2}{1 \text{ mol CO}_2} = 55.5 \text{ CO}_2$$

تجزیه های دورانی :

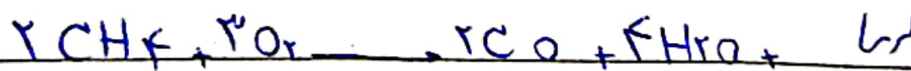
$$NH_4 = 1 \times 14 + 4 \times 1 = 18 \text{ g mol}^{-1}$$

$$? \text{ mol H}_2 = 18.2 \text{ kg NH}_4 \times \frac{1000 \text{ g NH}_4}{1 \text{ kg NH}_4} \times \frac{1 \text{ mol NH}_4}{18 \text{ g NH}_4} \times \frac{4 \text{ mol H}_2}{1 \text{ mol NH}_4} = 4050 \text{ mol H}_2$$

$$N_2 = 28 \text{ g mol}^{-1}$$

$$? \text{ g H}_2 = 4050 \text{ mol NH}_4 \times \frac{1 \text{ mol NH}_4}{18 \text{ g NH}_4} \times \frac{4 \text{ mol H}_2}{1 \text{ mol NH}_4} \times \frac{2 \text{ g H}_2}{1 \text{ mol H}_2} = 1820 \text{ g H}_2$$

$$? \text{ g N}_2 = 4050 \text{ mol NH}_4 \times \frac{1 \text{ mol NH}_4}{18 \text{ g NH}_4} \times \frac{1 \text{ mol N}_2}{1 \text{ mol NH}_4} \times \frac{28 \text{ g N}_2}{1 \text{ mol N}_2} = 6100 \text{ g N}_2$$



$$? \text{ CO}_2 = 14 \text{ g CHF}_3 \times \frac{1 \text{ mol CHF}_3}{14 \text{ g CHF}_3} \times \frac{2 \text{ mol CO}_2}{1 \text{ mol CHF}_3} \times \frac{44 \text{ g CO}_2}{1 \text{ mol CO}_2} = 44 \text{ g CO}_2$$