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Solution 01
Neutron is not present in ordinary hydrogen atom.
Solution 02
J. J. Thomson
Solution 03
Maximum of 2 e⁻ can be accommodated in K shell of an atom.
Solution 04
Maximum of 8 e⁻ can be accommodated in L shell of an atom.
Solution 05
Maximum of 18 e⁻ can be accommodated in M shell of an atom.
Solution 06
Maximum of 32 e⁻ can be accommodated in N shell of an atom.
Solution 07
(a) Maximum of 2 e⁻ can be accommodated in innermost shell of an atom.
(b) Maximum of 8 e⁻ can be accommodated in outermost shell of an atom.
Solution 08
Three subatomic particles present in atom are electrons, protons and neutrons.
Solution 09
Electron is negatively charged particle present in atoms of all the elements.
Solution 10
J. J. Thomson discovered electron.
Solution 11
(a) e⁻
(b) p⁺
(c) n

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Solution 12
(a) False
(b) True
(c) False
Solution 13
Nucleus is the central part of an atom in which protons and neutrons are held together.
Solution 14
K, L, M, N were the letters used by Bohr to represent electron shells in an atom.
Solution 15
Protons and neutrons actually determine the mass of atom.
Solution 16
Proton is the positively charged particle present in atoms of all the elements.
Solution 17
Electronic configuration of hydrogen : 1
Solution 18
Proton is 1840 times heavier than electron.
Solution 19
Hydrogen gas produces anode rays consisting of protons in the

discharge tube experiment.
Solution 20
Nucleus was discovered by Rutherford in the alpha particle scattering experiment.
Solution 21
Positive charge on the nucleus is due to presence of protons.
Solution 22
(a) 8 electrons are present in outermost shell of Neon.
(b) 7 electrons are present in outermost shell of Chlorine.
Solution 23
(a) L shell can accommodate maximum of 8 e^-.
(b) N shell can accommodate maximum of 32 e^-.
Solution 24
(a) K shell can accommodate maximum of 2 e^-.
(b) M shell can accommodate maximum of 18 e^-.
Solution 25
(i) Chadwick discovered 'neutron'.
(ii) Thomson discovered 'electron'.
(iii) Goldstein discovered 'proton'.
Solution 26
(a) Proton has relative charge of +1.
(b) Electron has relative charge of -1.
(c) Neutron has relative charge of 0.
Solution 27
(a) Atomic number
(b) Mass number
(c) 11
(d) 23
(e) 20
(f) Nucleus
(g) Electrons
(h) Protons
(i) 8
(j) 18
(k) Neutron
(l) Negative; Positive; No
Solution 28
Electron is a negatively charged particle found in the atoms of all elements.
The relative mass of an electron is 1/1840 u.
A charge of -1 is carried by an electron.
Solution 29

Absolute mass of electron is 9×10^{-28} Kg.

Absolute charge on electron is 1.6×10^{-19} C.

Solution 30
The deflection of fast moving alpha- particles through small and large angles in Rutherford's scattering experiment is the evidence for the presence of nucleus in an atom.
Solution 31
Important information furnished about nucleus in Rutherford's alpha- particle scattering experiment is:
(i) Nucleus of an atom is positively charged.
(ii) Nucleus of an atom is very hard and dense.
(iii) Nucleus of an atom is very small as compared to the size of an atom as a whole.
Solution 32
Most of the alpha- particles passed straight through the gold foil without any deflection in Rutherford's alpha- particle scattering experiment, this shows that most of the space in an atom is empty.
Solution 33
There are equal no. of positive and negative particles in an atom, so

it is neutral as a whole.

Solution 34

(a) Proton is present in same fixed no. in the atoms of any particular element.

(b) Atomic no. is characteristic for any particular element.

********** END *********