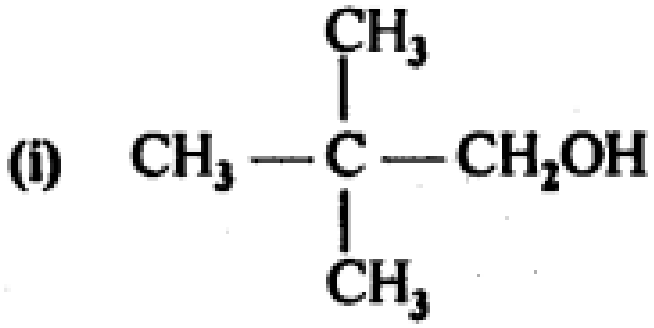
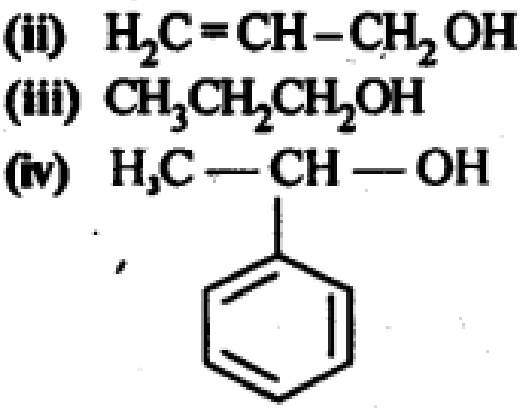
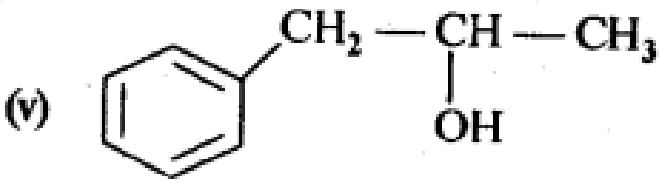
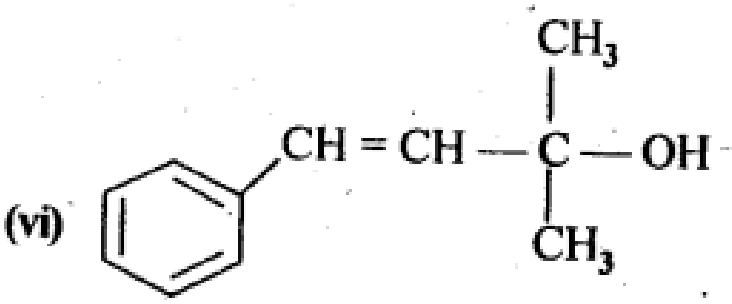
NCERT TEXTBOOK QUESTIONS SOLVED

11.1. Classify the following as primary, secondary and tertiary alcohols.



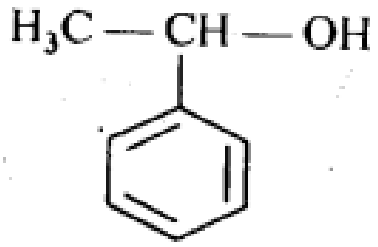






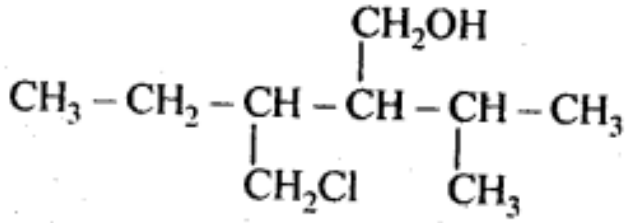
Ans:
Primary alcohols: (i), (ii), (iii)
Secondary alcohols: (iv), (v)
Tertiary alcohols: (vi)

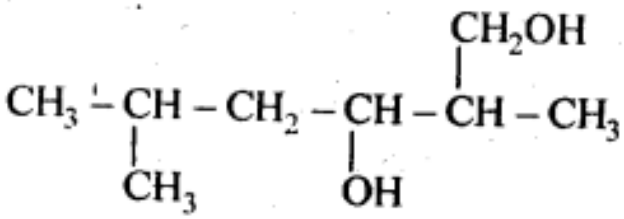
11.2. Identify allylic alcohols in the above examples.
Ans: (ii) and (iv) i.e. $H_2C=CH - CH_2OH$ and

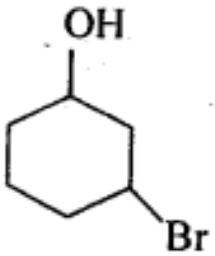


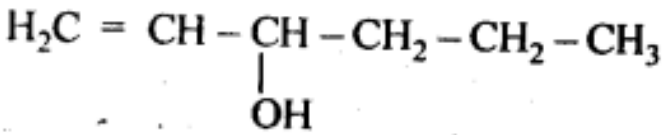
are allylic alcohols

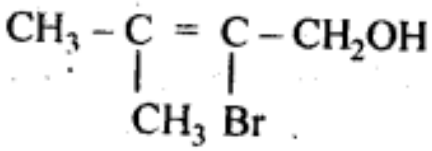
11.3. Name the following compounds according to IUPAC system.

(i) 

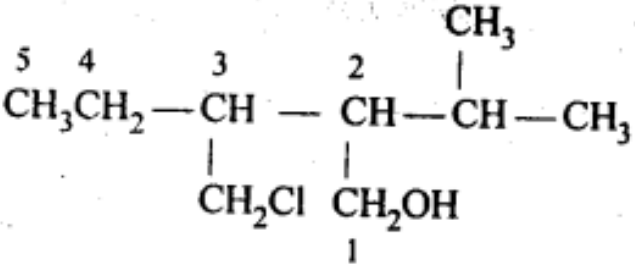
(ii) 

(iii) 

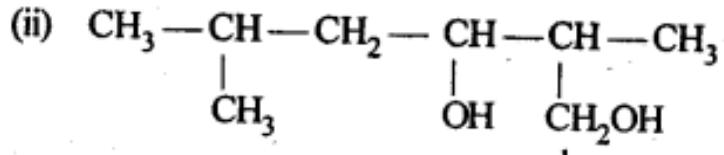
(iv) 

(v) 

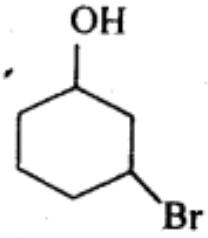
Ans:

(i) 

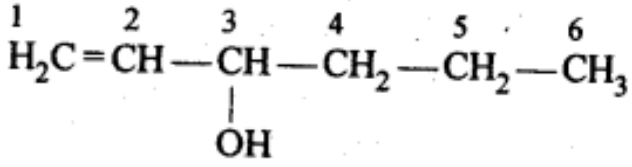
3 – Chloromethyl – 2 – isopropylpentan – 1 – ol
6 5 4 3 2

(ii) 

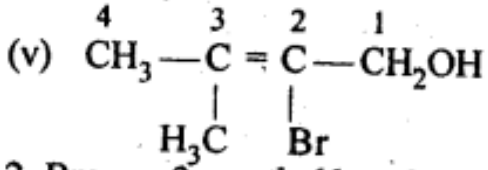
2, 5–Dimethylhexane – 1, 3 – diol

(iii) 

3 – Bromocyclohexanol

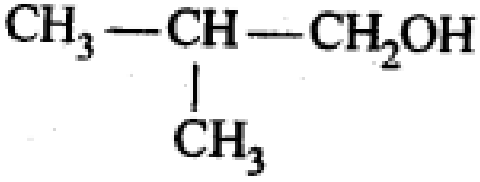
(iv) 

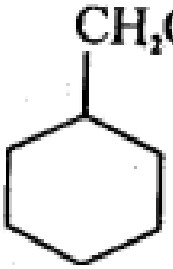
Hex–1–en–3–ol

(v) 

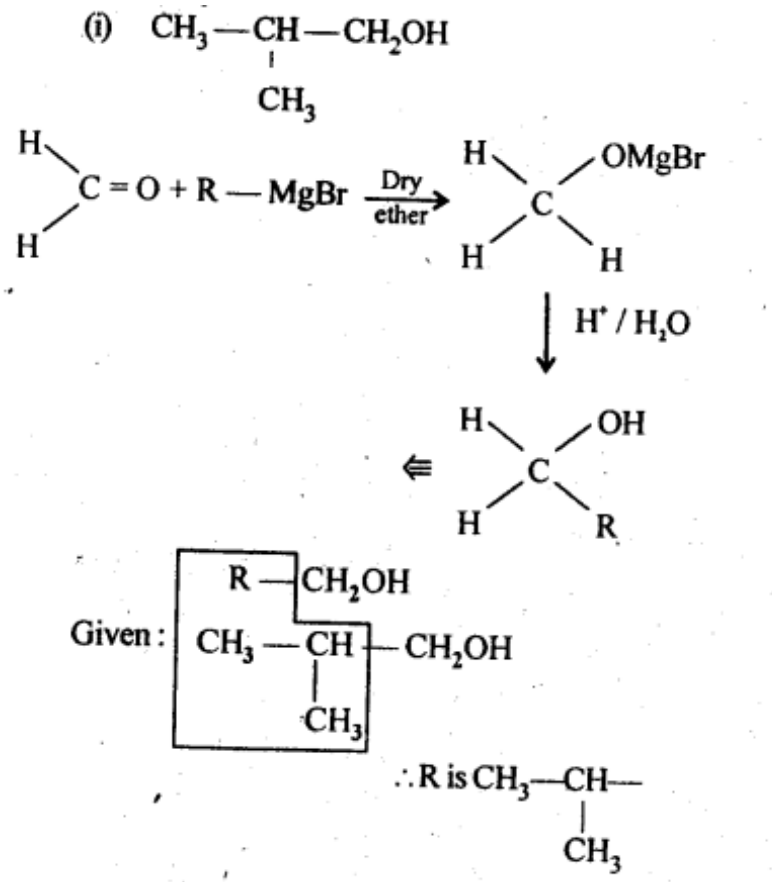
2–Bromo–3– methyl but–2–en–1–ol

11.4. Show how are the following alcohols prepared by the reaction of a suitable Grignard reagent on methanal ?

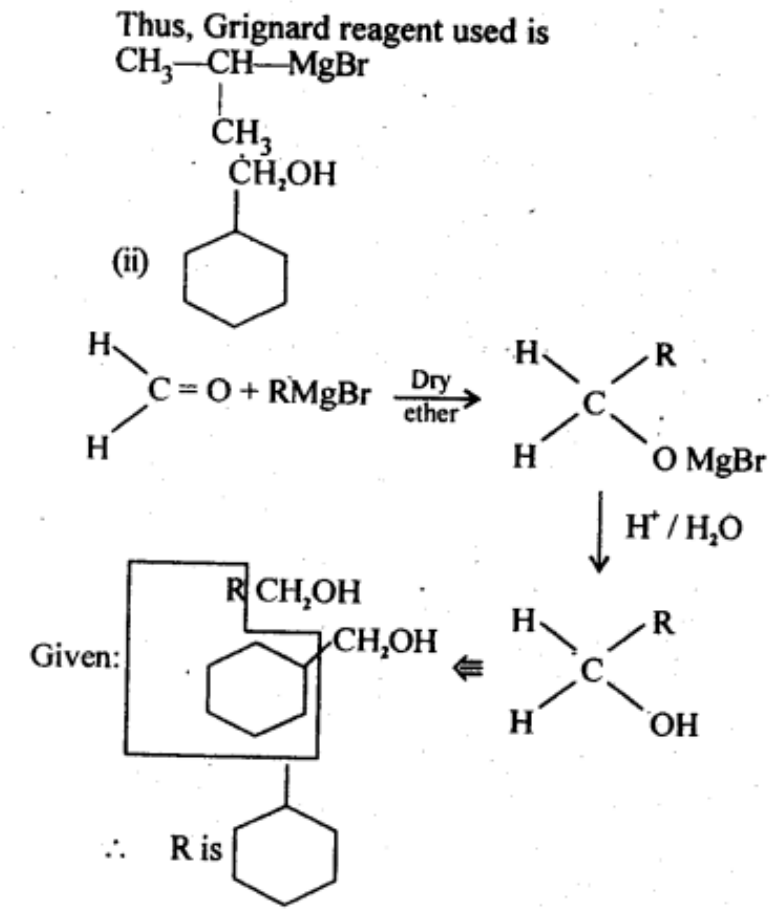
(i) 

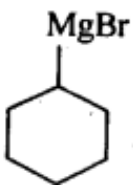
(ii) 

Ans:



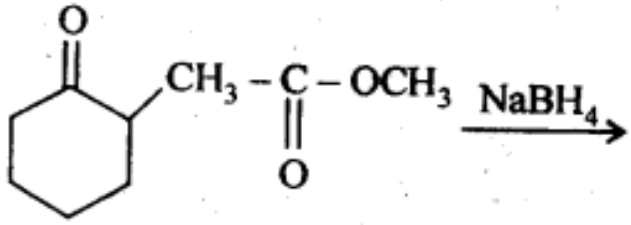
Thus, Grignard reagent used is $CH_3-CH(CH_3)-MgBr$



Thus, the Grignard reagent used is 

11.5. Write structures of the products of the following reactions:

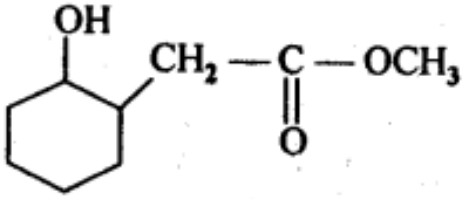
(i) $CH_3 - CH = CH_2 \xrightarrow{H_2O/H^+}$

(ii) 

(iii) $CH_3 - CH_2 - \underset{\displaystyle CH_3}{\underset{|}{CH}} - CHO \xrightarrow{NaBH_4}$

Ans:

(i) $CH_3 - \underset{\displaystyle OH}{\underset{|}{CH}} - CH_3$ **Propan-2-ol**

(ii) 

Methyl (2-hydroxy cyclohexyl) ethanoate

(iii) $CH_3 - CH_2 - \underset{\displaystyle CH_3}{\underset{|}{CH}} - CH_2OH$

2–Methylbutan–1–ol

11.6. Give structures of the products you would expect when each of the following alcohol reacts with (a) HCl-$ZnCl_2$ (b) HBrand (c) $SOCl_2$

(i) Butan-1-ol

(ii) 2-Methylbutan-2-ol

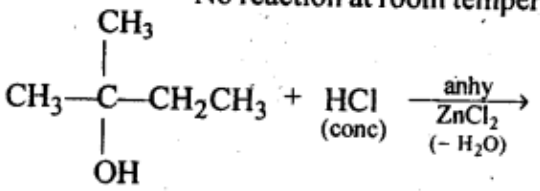
Ans:

(a) with $HCl - ZnCl_2$

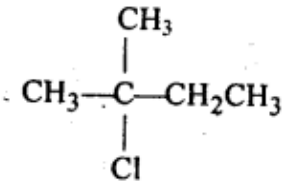
$$CH_3CH_2CH_2CH_2OH + HCl \xrightarrow[ZnCl_2]{anhy}$$

butan -1-ol (1°) (conc)

No reaction at room temperature



2-Methylbutan-2-ol (3°)

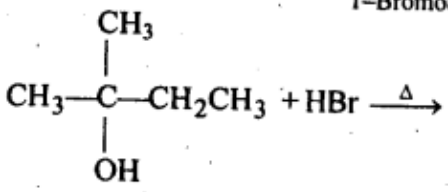


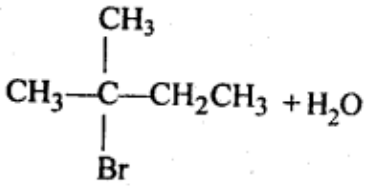
2-Chloro-2-methylbutane
(white turbidity)

(b) with HBr

$$CH_3CH_2CH_2CH_2OH + HBr \xrightarrow{\Delta} CH_3CH_2CH_2CH_2Br + H_2O$$

1–Bromobutane



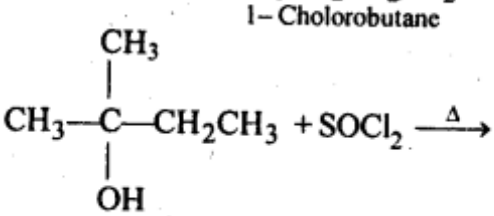


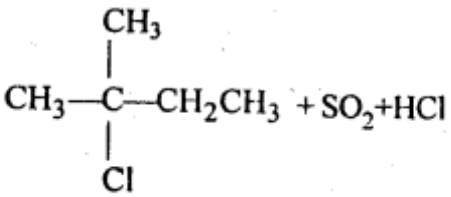
2-Bromo-2-methylbutane

(c) with $SOCl_2$

$$CH_3CH_2CH_2CH_2OH + SOCl_2 \xrightarrow{\Delta} CH_3CH_2CH_2CH_2Cl + SO_2 + HCl$$

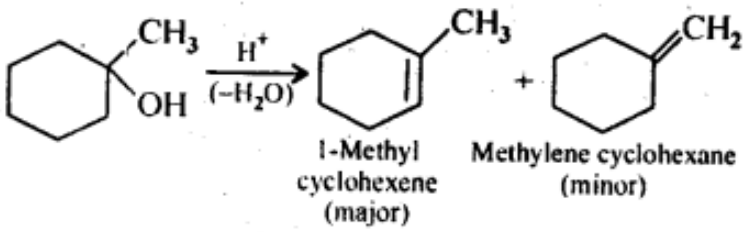
1– Cholorobutane





2-Chloro-2-methyl-butane

11.7. Predict the major product of acid catalysed dehydration of
(i) 1-nicthylcyclohcxanoland
(ii) butan-1-ol
Ans:



1-Methyl cyclohexene (major)

Methylene cyclohexane (minor)

(ii) $CH_3CH_2CH_2CH_2OH \longrightarrow$

$CH_3CH{=}CH{-}CH_3$ + $CH_3CH_2CH{=}CH_2$

But-2-ene (major)

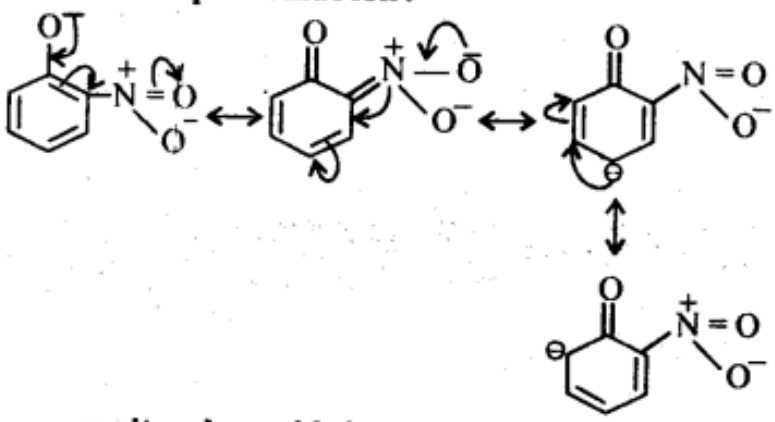
But-1-ene (minor)

11.8. Ortho and para nitrophenols are more acidic than phenol. Draw the resonance structures of the corresponding phenoxide ions.

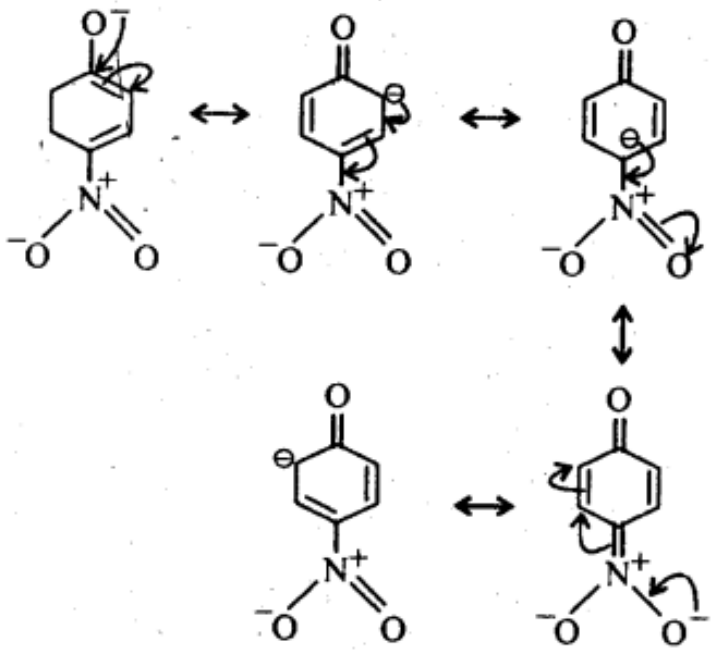
Ans:

The resonance structures of o-and p- nitrophenoxide ions and phenoxide ion are given below:

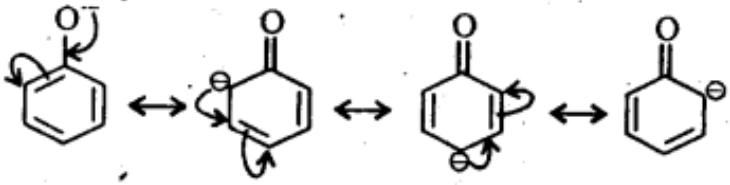
***o*-nitrophenoxide ion :**



***p*-nitrophenoxide ion :**

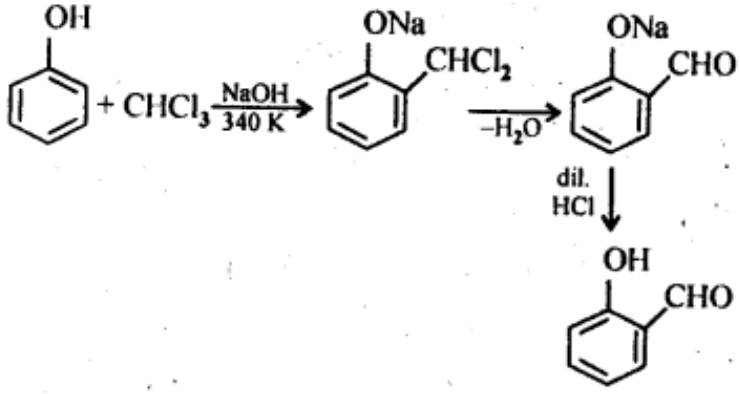


phenoxide ion :

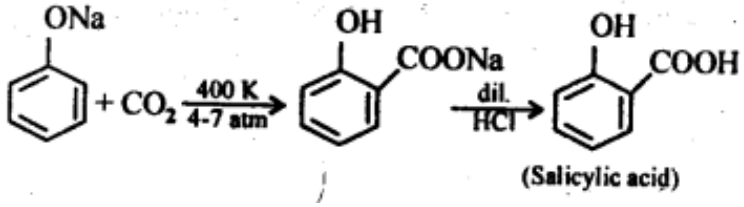


Due to –R effect of $-NO_2$ group, *o*– and *p*–nitrophenoxide are more stable than phenoxide ion. As a result, *o*– and *p*–nitrophenols are more acidic than phenol.

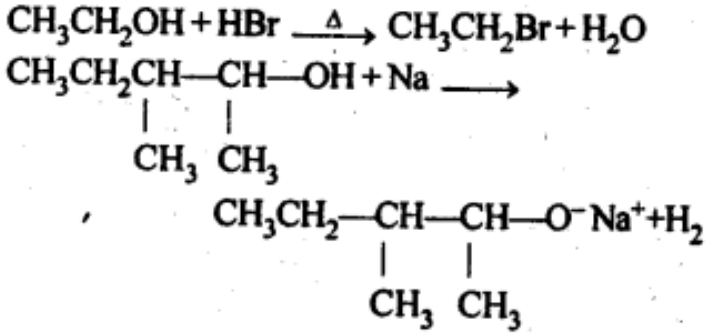
11.9 Write the equations involved in the following reactions:
(i) Reimer-Tiemann reaction
(ii) Kolbe's reaction
Ans: (i) Reimer-Tiemann reaction

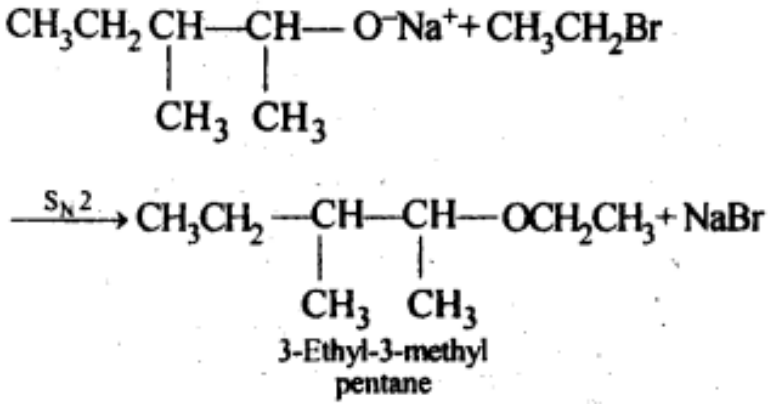


(ii) Kolbe's reaction



11.10. Write the reactions of Williamson synthesis of 2-ethoxy-3-methylpentane starting from ethanol and 3-methylpentan-2-ol.
Ans: In Williamsons's synthesis, the alkyl halide should be primary. Thus, the alkyl halide should be derived from ethanol and the alkoxide ion from 3-methylpentan-2-ol. The synthesis is as follows





3-Ethyl-3-methyl pentane

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