



Page 49:

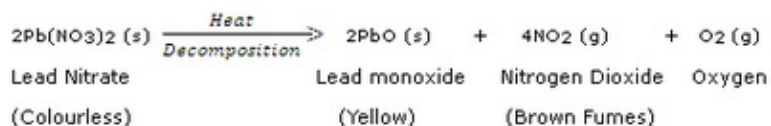
Solution 58:

- (a) Ferrous sulphate.
- (b) Decomposition reaction.
- (c)



Solution 59:

- (a) Lead nitrate.
- (b) Nitrogen dioxide.
- (c)



Solution 60:

- (a) Combination reaction.
- (b) Decomposition reaction.

Solution 61:

- (b) Copper (Cu)
 - (c) Copper sulphate (CuSO_4)
 - (d) Displacement reaction;
- $$\text{CuSO}_4 (\text{aq}) + \text{Zn} (\text{s}) \rightarrow \text{ZnSO}_4 (\text{aq}) + \text{Cu} (\text{s})$$

- (e) Brass

Solution 62:

- (a) Copper (Cu)
- (b) Hydrogen (H_2)
- (c) Copper oxide (CuO)
- (d) Water (H_2O)
- (e) $\text{CuO} + \text{H}_2 \rightarrow \text{Cu} + \text{H}_2\text{O}$
- (f) Displacement reaction (which is also a redox reaction).

Solution 63:

- (a) Silver (Ag)
- (b) Silver nitrate (AgNO_3)
- (c) Silver chloride (AgCl)
- (d) $\text{AgNO}_3 (\text{aq}) + \text{NaCl} (\text{aq}) \rightarrow \text{AgCl} (\text{s}) + \text{NaNO}_3 (\text{aq})$
- (e) Double displacement reaction.

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Solution 64:

- (a) Metal X : Copper; Metal Y : Sodium
- (b) Copper sulphate, CuSO_4
- (c) Sodium sulphate, Na_2SO_4
- (d) Barium sulphate, BaSO_4
- (e) Copper chloride, CuCl_2

Solution 65:

- (a) Copper sulphate.

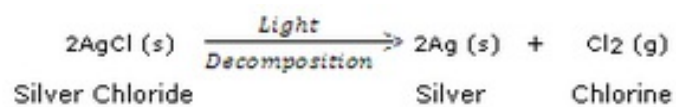
- (b) Blue colour.
- (c) Copper sulphide.
- (d) $\text{CuSO}_4 (\text{aq}) + \text{H}_2\text{S} (\text{g}) \rightarrow \text{CuS} (\text{s}) + \text{H}_2\text{SO}_4 (\text{aq})$
- (e) Double displacement reaction.

Solution 66:

- (a) Copper.
- (b) Silver nitrate.
- (c) Silver.
- (d) Copper nitrate.
- (e) Displacement reaction.

Solution 67:

- (a) Silver.
- (b) Chlorine.
- (c) Silver chloride.
- (d) Silver nitrate and Sodium chloride.
- (e) Decomposition reaction;



***** END *****