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Solution 69

- (a) No solution will turn blue litmus to red.
- (b) Solution Q (sodium hydroxide) will turn red litmus blue.

Solution 70

A is copper sulphate pentahydrate, $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$

B is water, H_2O

C is anhydrous copper sulphate, CuSO_4

D is water, H_2O

Solution 71

X is Sodium chloride. The process is called Chlor-alkali process.

Solution 72

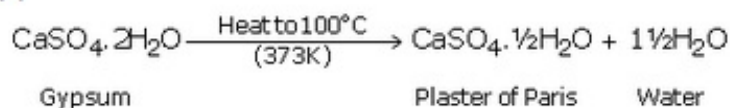
- (a) $\text{Ca}(\text{OH})_2$ and Cl_2
- (b) NaCl , NH_3 , H_2O and CO_2
- (c) $2\text{CaSO}_4 \cdot \text{H}_2\text{O}$
- (d) NaHCO_3
- (e) NaHCO_3

Solution 73

- (a) Ammonium chloride, NH_4Cl
- (b) Sodium chloride, NaCl
- (c) Sodium carbonate, Na_2CO_3

Solution 74

- (a) Plaster of Paris.
- (b)



- (c) POP is used in hospitals for setting fractured bones in the right position to ensure correct healing.

Solution 75

- (a) NH_4Cl , $(\text{NH}_4)_2\text{SO}_4$
- (b) NaCl , K_2SO_4
- (c) Na_2CO_3 , CH_3COONa

Solution 76

Bleaching powder, CaOCl_2 .



Solution 77

Salt X is like sodium carbonate, Na_2CO_3 , which is made from a strong base and a weak acid. On dissolving in water, salt X gets hydrolysed to form some strong base and some weak acid. The strong base thus formed makes the solution alkaline which turns red litmus blue.

Solution 78

Baking powder; When baking powder mixes with water, then

sodium hydrogencarbonate reacts with tartaric acid to evolve carbon dioxide gas which gets trapped in the wet dough and bubbles out slowly making the cake soft and spongy.

Solution 79

Plaster of Paris.

Solution 80

Sodium hypochlorite, NaClO ; used in making household bleaches and for bleaching fabrics.

Solution 81

(a) Gypsum - $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$

(b) Copper sulphate crystals - $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$

(c) Sodium carbonate crystals - $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$

Solution 82

(a) 5.

(b) 10.

(c) 2.

***** END *****