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Solution 58:
(a) Ferrous sulphate.
(b) Decomposition reaction.
(c)

$$2FeSO_4 \xrightarrow[Decomposition]{Heat} Fe_2O_3 + SO_2 + SO_3$$

Solution 59:
(a) Lead nitrate.
(b) Nitrogen dioxide.
(c)

$$2Pb(NO_3)_2\,(s) \xrightarrow[Decomposition]{Heat} 2PbO\,(s) + 4NO_2\,(g) + O_2\,(g)$$

Lead Nitrate	Lead monoxide	Nitrogen Dioxide	Oxygen
(Colourless)	(Yellow)	(Brown Fumes)	

Solution 60:
(a) Combination reaction.
(b) Decomposition reaction.
Solution 61:
(b) Copper (Cu)
(c) Copper sulphate ($CuSO_4$)
(d) Displacement reaction;
$CuSO_4\,(aq) + Zn\,(s) \rightarrow ZnSO_4\,(aq) + Cu\,(s)$
(e) Brass
Solution 62:
(a) Copper (Cu)
(b) Hydrogen (H_2)
(c) Copper oxide (CuO)
(d) Water (H_2O)
(e) $CuO + H_2 \rightarrow Cu + H_2O$
(f) Displacement reaction (which is also a redox reaction).
Solution 63:
(a) Silver (Ag)
(b) Silver nitrate ($AgNO_3$)
(c) Silver chloride (AgCl)
(d) $AgNO_3(aq) + NaCl\,(aq) \rightarrow AgCl\,(s) + NaNO_3\,(aq)$
(e) Double displacement reaction.

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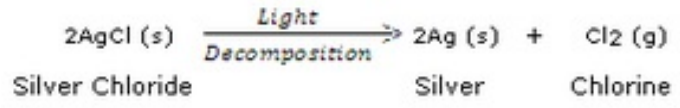
Solution 64:
(a) Metal X : Copper; Metal Y : Sodium
(b) Copper sulphate, $CuSO_4$
(c) Sodium sulphate, Na_2SO_4
(d) Barium sulphate, $BaSO_4$
(e) Copper chloride, $CuCl_2$
Solution 65:
(a) Copper sulphate.

(b) Blue colour.
(c) Copper sulphide.
(d) $CuSO_4$ (aq) + H_2S (g) → CuS (s) + H_2SO_4 (aq)
(e) Double displacement reaction.
Solution 66:
(a) Copper.
(b) Silver nitrate.
(c) Silver.
(d) Copper nitrate.
(e) Displacement reaction.
Solution 67:
(a) Silver.
(b) Chlorine.
(c) Silver chloride.
(d) Silver nitrate and Sodium chloride.
(e) Decomposition reaction;



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