



QUESTIONS FROM NCERT TEXT BOOK

Question 1. Compare the properties of electrons, protons and neutrons.

Answer:

Electrons	Protons	Neutrons
1. Negatively charged.	Positively charged	No charge
2. Mass is negligible $\left(\frac{1}{1800}\right)$ times of protons.	Mass is 1 a.m.u.	Mass is 1 a.m.u.
3. Get attracted towards +ve charge.	Get attracted towards -ve charge.	Do not get attracted, as they are neutral
4. Present outside the nucleus	Present in the nucleus	Present in the nucleus of an atom

Question 2. What are the limitations of J.J. Thomson's model of the atom?

Answer: According to J.J. Thomson's model of an atom, the electrons are embedded all over in the positively charged spheres. But experiments done by other scientists showed that protons are present only in the centre of the atom and electrons are distributed around it.

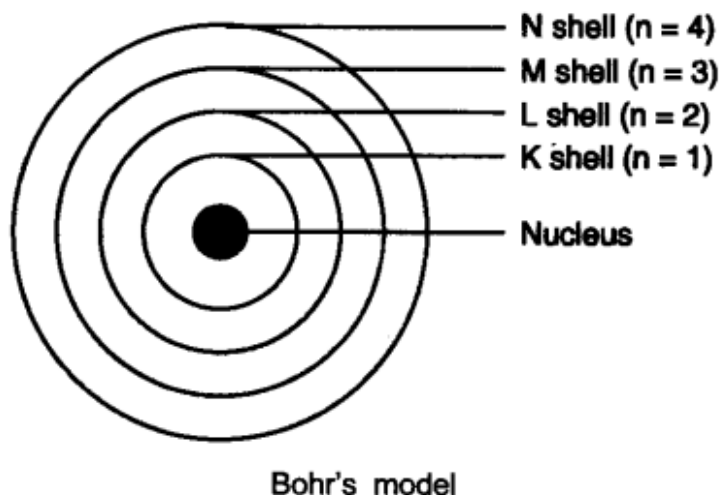
Question 3. What are the limitations of Rutherford's model of the atom?

Answer: According to Rutherford's model of an atom the electrons are revolving in a circular orbit around the nucleus. Any such particle that revolves would undergo acceleration and radiate energy. The revolving electron would lose its energy and finally fall into the nucleus, the atom would be highly unstable. But we know that atoms are quite stable.

Question 4. Describe Bohr's model of the atom.

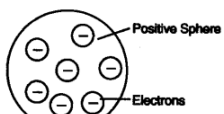
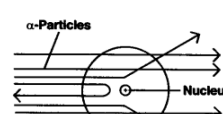

Answer: Bohr's model of the atom

- (1) Atom has nucleus in the centre.
- (2) Electrons revolve around the nucleus.
- (3) Certain special orbits known as discrete orbits of electrons, are allowed inside the atom.
- (4) While revolving in discrete orbits the electrons do not radiate energy.
- (5) These orbits or shells are called energy levels.
- (6) These orbits or shells are represented by the letters K, L, M, N or the numbers $n = 1, 2, 3, 4$



Question 5. Compare all the proposed Bohr's models of an atom given in this chapter.

Answer:

Thomson	Rutherford	Bohr
<ul style="list-style-type: none"> • Sphere of positive charge • Electrons are spread randomly all over in the sphere  <p style="text-align: center;">Thomson's Model</p> <ul style="list-style-type: none"> • Positive charge = Negative charge. • Atom is electrically neutral. 	<ul style="list-style-type: none"> • Sphere of positive charge in centre called nucleus. All mass of an atom resides in the nucleus • Electrons revolve around the nucleus in well defined orbits.  <p style="text-align: center;">Rutherford's Model</p> <ul style="list-style-type: none"> • Size of nucleus is very small as compared to size of atom. 	<ul style="list-style-type: none"> • Positive charge in centre called nucleus. • Electrons revolve in discrete orbits and do not radiate energy.  <p style="text-align: center;">Bohr's Model</p> <ul style="list-style-type: none"> • The orbits were termed as energy shells labelled as K, L, M, N or $n = 1, 2, 3, 4$ (numbered)

Question 6. Summarise the rules for writing of distribution of electrons in various shells for the first eighteen elements.

Answer: The rules for writing of distribution of electrons in various shells for the first eighteen elements are:

(i) The maximum number of electrons present in a shell is given by the formula - $2n^2$

since n = orbit number i.e., 1, 2, 3

since Maximum number of electrons in different shells are:

$$\text{K shell } n = 1 \quad 2n^2 \Rightarrow 2(1)^2 = 2$$

$$\text{L shell } n = 2 \quad 2n^2 \Rightarrow 2(2)^2 = 8$$

$$\text{M shell } n = 3 \quad 2n^2 \Rightarrow 2(3)^2 = 18$$

$$\text{N shell } n = 4 \quad 2n^2 \Rightarrow 2(4)^2 = 32$$

(ii) The maximum number of electrons that can be accommodated in the outermost orbit is 8.

(iii) Electrons are not accommodated in a given shell unless the inner shells are filled. (Shells are filled step-wise).

Question 7. Define valency by taking examples of silicon and oxygen.

Answer: Valency is the combining capacity of an atom.

Atomic number of oxygen = 8 Atomic number of silicon = 14 K L M

Electronic configuration of oxygen = 2 6 -

Electronic configuration of silicon = 2 8 4

In the atoms of oxygen the valence electrons are 6 (i.e., electrons in the outermost shell). To fill the orbit, 2 electrons are required. In the atom of silicon, the valence electrons are 4. To fill this orbit 4 electrons are required.

Hence, the combining capacity of oxygen is 2 and of silicon is 4.
 i.e., Valency of oxygen = 2
 Valency of silicon = 4

Question 8. Explain with examples:

(i) Atomic number (ii) Mass number,

(iii) Isotopes and (iv) Isobars.

Give any two uses of isotopes.

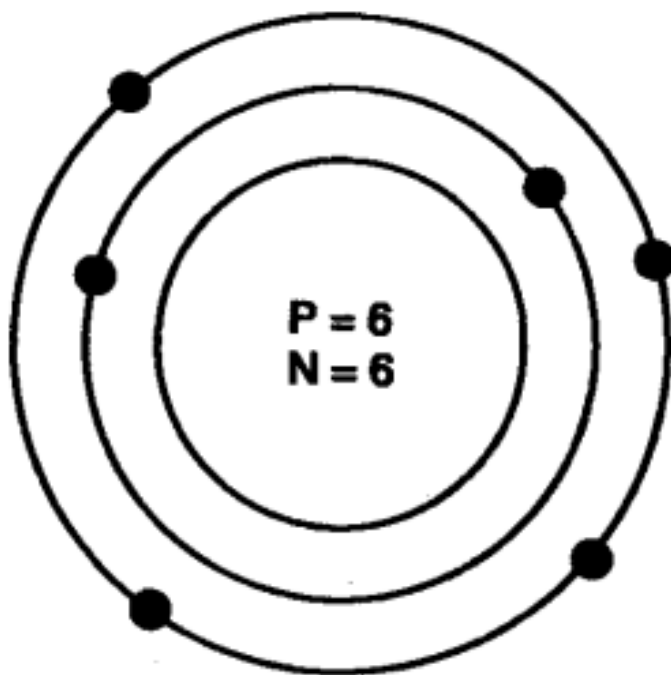
Answer: (i) Atomic number: The atomic number of an element is equal to the number of protons in the nucleus of its atom. e.g.,

Oxygen has 6 protons hence atomic no. = 6.

(ii) Mass number: The mass number of an atom is equal to the number of protons and neutrons in its nucleus.

Nucleons = number of protons + number of neutrons Example:

Protons + Neutrons = Nucleus = Mass number $6 + 6 = 12$



(iii) Isotopes: Isotopes are atoms of the same element which have different mass number but same atomic number.



(iv) Isobars: Isobars are atoms having the same mass number but different atomic numbers.



Both calcium and argon have same mass number but different atomic number.

Two uses of isotopes are:

(i) An isotope of iodine is used in the treatment of goitre.

(ii) An isotope of uranium is used as a fuel in nuclear reactors.

Question 9. Na^+ has completely filled K and L shells. Explain.

Answer: Sodium atom (Na), has atomic number = 11

Number of protons = 11

Number of electrons = 11

Electronic configuration of Na = K L M - 2 8 1

Sodium atom (Na) loses 1 electron to become stable and form Na^+ ion. Hence it has completely filled K and L shells.

Question 10. If bromine atom is available in the form of say, two isotopes $^{79}_{35}\text{Br}$ (49.7%) and $^{81}_{35}\text{Br}$ (50.3%), calculate the average atomic mass of bromine atom.

Answer:

The average atomic mass of bromine atom

$$\begin{aligned} &= \left(79 \times \frac{49.7}{100} \right) + \left(81 \times \frac{50.3}{100} \right) \\ &= 39.263 + 40.743 \\ &= 80.006 \text{ u} \end{aligned}$$

Question 11. The average atomic mass of a sample of an element X is 16.2 u. What are the percentages of isotopes $^{16}_8\text{X}$ and $^{18}_8\text{X}$ in the sample?

Answer: Let the percentage of $^{16}_8\text{X}$ be x and the percentage of $^{18}_8\text{X}$ be 100 - x.

$$\begin{aligned} \therefore \quad &\left(16 \times \frac{x}{100} \right) + \frac{18(100 - x)}{100} = 16.2 \\ &\frac{16x}{100} + \frac{1800 - 18x}{100} = 16.2 \\ \therefore \quad &\frac{16x - 18x + 1800}{100} = 16.2 \\ \therefore \quad &-2x + 1800 = 16.2 \times 100 \\ \therefore \quad &-2x = 1620 - 1800 \\ \therefore \quad &-2x = -180 \\ \therefore \quad &x = \frac{180}{2} = 90 \\ \therefore \quad &^{16}_8\text{X} = 90\% \\ \text{and} \quad &^{18}_8\text{X} = 10\% \end{aligned}$$

Question 12. If Z = 3, what would be the valency of the element?

Also, name the element.

Answer: Z = 3, (i.e., atomic number \rightarrow z)

\therefore Electronic configuration = 2, 1

Valency = 1

Name of the element is lithium.

Question 13. Composition of the nuclei of two atomic species X and Y are given as under

X - Protons = 6 and Neutrons = 6

Y - Protons = 6 and Neutrons = 8

Give the mass number of X and Y. What is the relation between the two species?

Answer: Mass number of X = Protons + Neutrons

$$= 6 + 6 = 12$$

$$\text{Mass number of Y} = \text{Protons} + \text{Neutrons} = 6 + 8 = 14$$

As the atomic number is same i.e., = 6.

[atomic number = number of protons].

Both X and Y are isotopes of same element.

Question 14. For the following statements, write T for True and F for False.

(a) J.J. Thomson proposed that the nucleus of an atom contains only nucleons.

- (b) A neutron is formed by an electron and a proton combining together. Therefore, it is neutral.
 (c) The mass of an electron is about 1/2000 times that of proton.
 (d) An isotope of iodine is used for making tincture iodine, which is used as a medicine.

Answer:

- (a) False
 (b) False
 (c) True
 (d) False

Put tick against correct choice and cross (x) against wrong choice in questions 15, 16 and 17.

Question 15. Rutherford's alpha-particle scattering experiment was responsible for the discovery of

- (a) Atomic nucleus
 (c) Proton
 (b) Electron
 (d) Neutron

Answer: (a) Atomic nucleus

Question 16. Isotopes of an element have

- (a) the same physical properties
 (c) different number of neutrons
 (b) different number of neutrons
 (d) different atomic numbers.

Answer: (c) different number of neutrons

Question 17. Number of valence electrons in Cr^{3+} ion are :

- (a) 16
 (b) 8
 (c) 17
 (d) 18

Answer: (b) 8

Question 18. Which one of the following is a correct electronic configuration of sodium?

- (a) 2, 8
 (b) 8, 2, 1
 (c) 2, 1, 8
 (d) 2, 8, 1

Answer: (d) 2, 8, 1

Question 19. Complete the following table.

Atomic Number	Mass Number	Number of Neutrons	Number of Protons	Number of Electrons	Name of the Atomic Species
9	—	10	—	—	—
16	32	—	—	—	Sulphur
—	24	—	12	—	—
—	2	—	1	—	—
—	1	0	1	0	—

Answer:

Atomic Number	Mass Number	Number of Neutrons	Number of Protons	Number of Electrons	Name of the Atomic Species
9	19	10	9	9	Fluorine
16	32	16	16	16	Sulphur
12	24	12	12	12	Magnesium
1	2	1	1	1	Hydrogen
1	1	0	1	0	Deuterium
1	1	0	1	0	Hydrogen

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