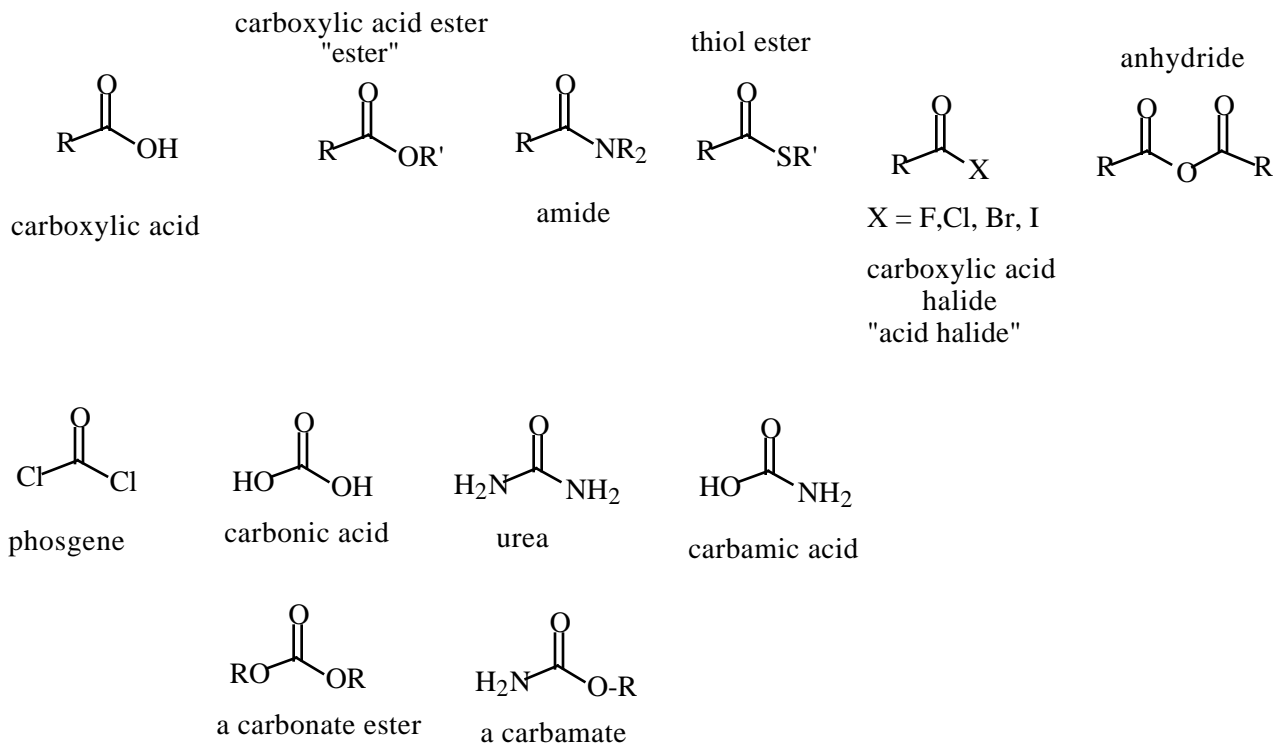


Chapter 20-- Nomenclature sec 20.1

Note bp data, p 1009 Table 20.1; Note pKa data, p 1010 Table 20.2; IR C=O stretch table 20.3

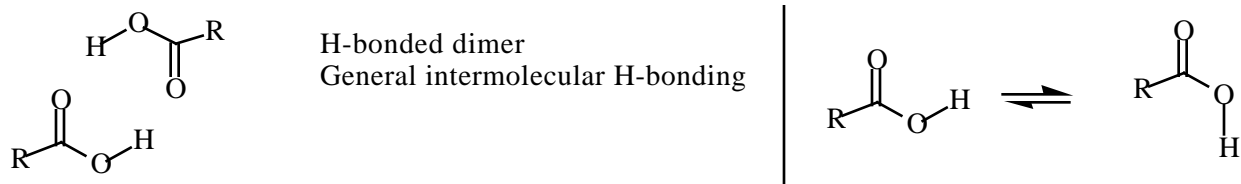
Note: NMR data Table 20.4

**DERIVATIVES OF CARBOXYLIC ACIDS:**

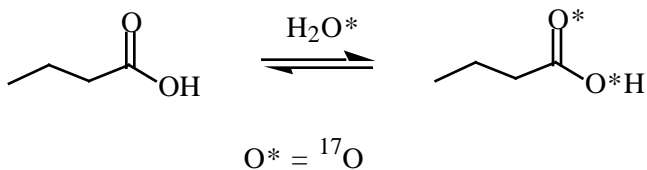
READ: Nomenclature, text 19.1

**Carboxylic Acids:****Dominant feature:** acidity and influence of structure on acidity. You know all of this stuff.

Structure:

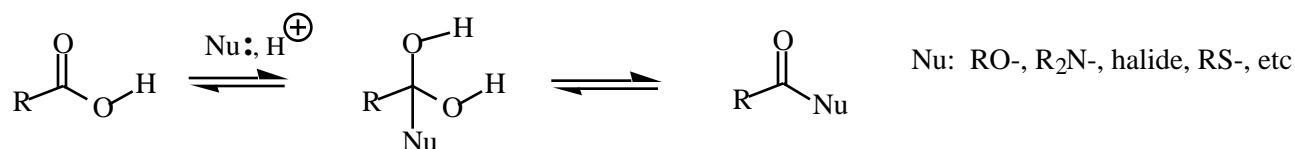
**Dominant reaction:**

Accelerated by acid and base

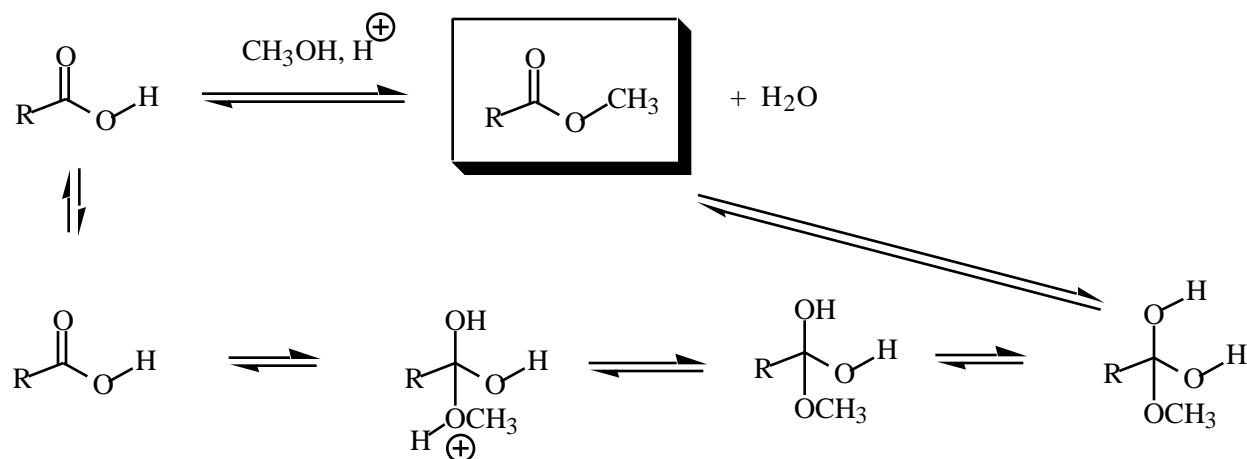


Not very interesting for carboxylic acids; not even noticeable unless the water is labeled.

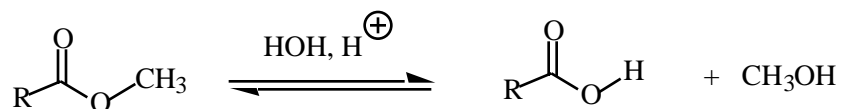
**BUT:** a major reaction with other nucleophiles. Addition/elimination



### Formation of Esters:

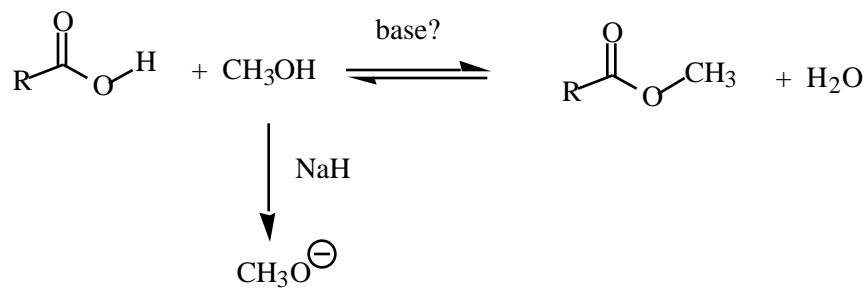


Completely reversible: also serves as a mechanism for hydrolysis of esters to carboxylic acids:

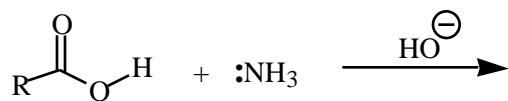
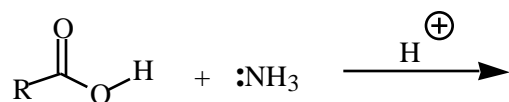
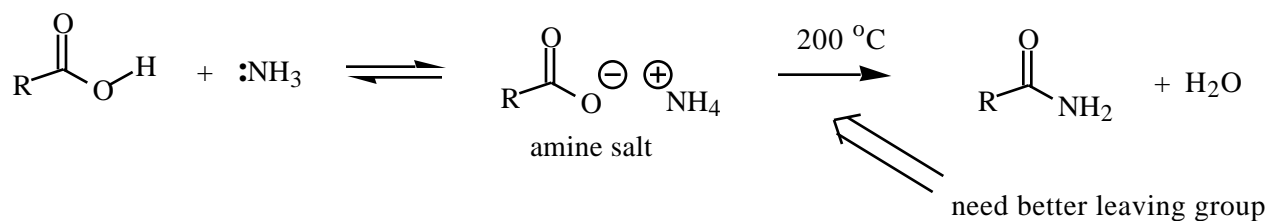


Drive the reaction by using excess of one or the other of the reactants: e.g., MeOH as solvent to make ester  
water as solvent to hydrolyze ester

What about reaction under basic conditions?

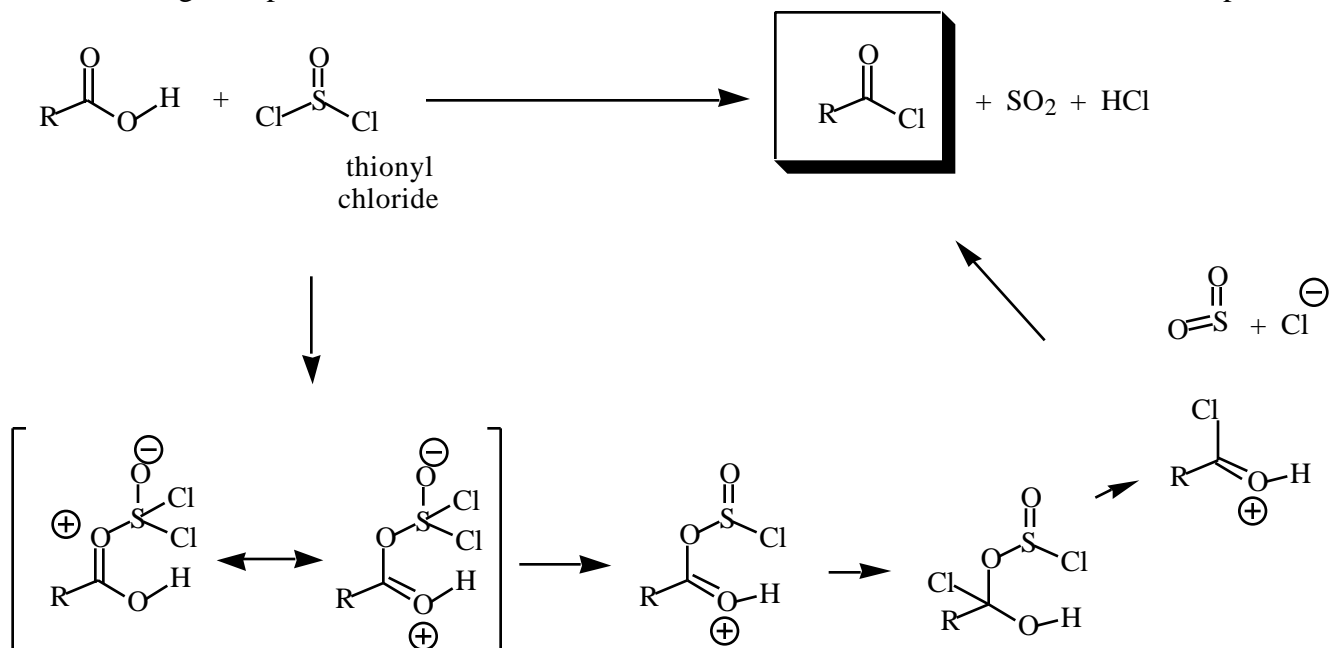


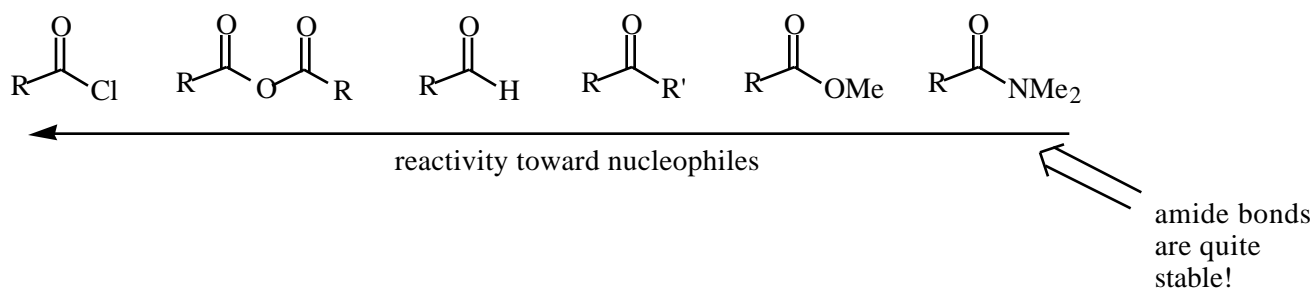
**What about amines as the nucleophile?** Amide bond formation.



Better Leaving Group?  $\text{Cl}^-$

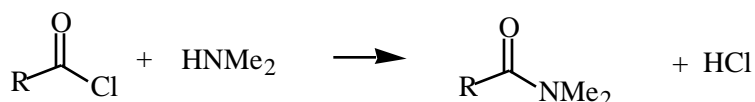
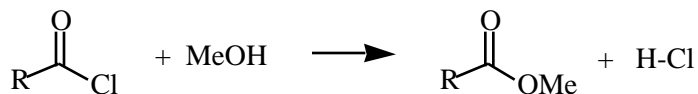
text p 978



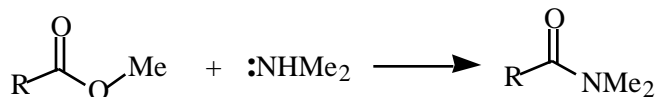


Recall correlation of IR stretching frequencies of C=O with substituent effect, via polar resonance structure (more resonance stabilization, less reactive toward nucleophiles)

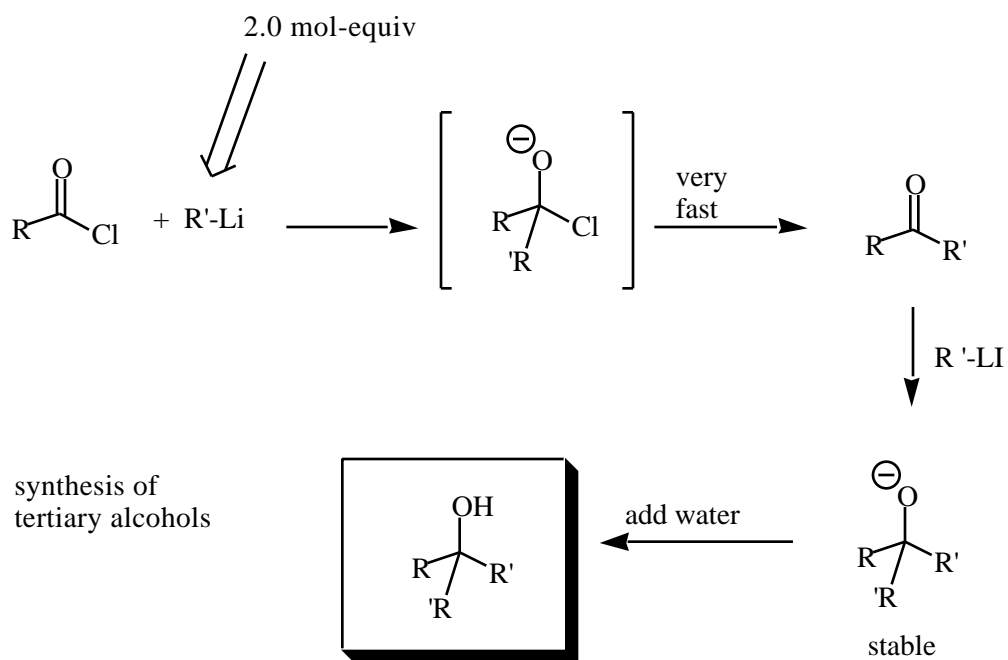
**A.** Addition of stabilized nucleophiles (O, N)



Also:

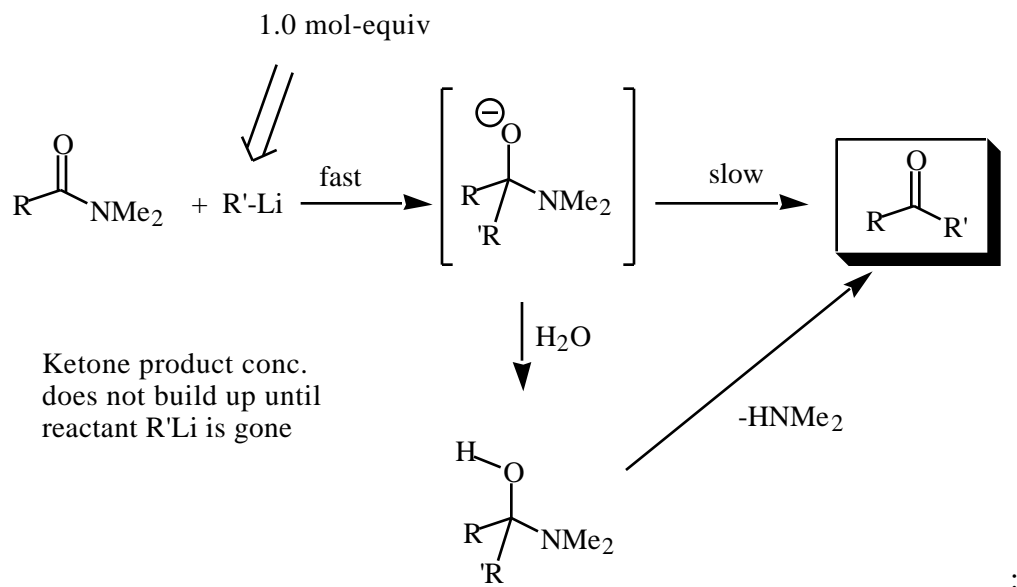


MeOH very slow

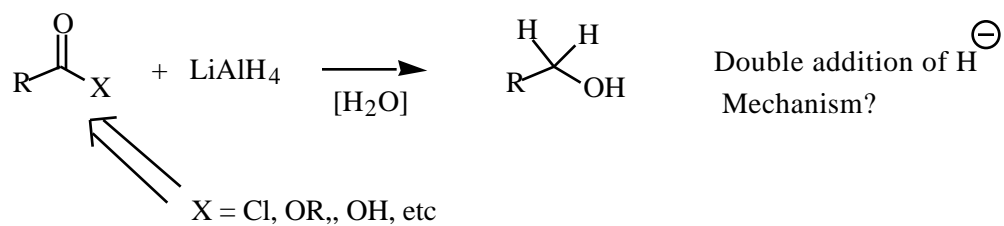


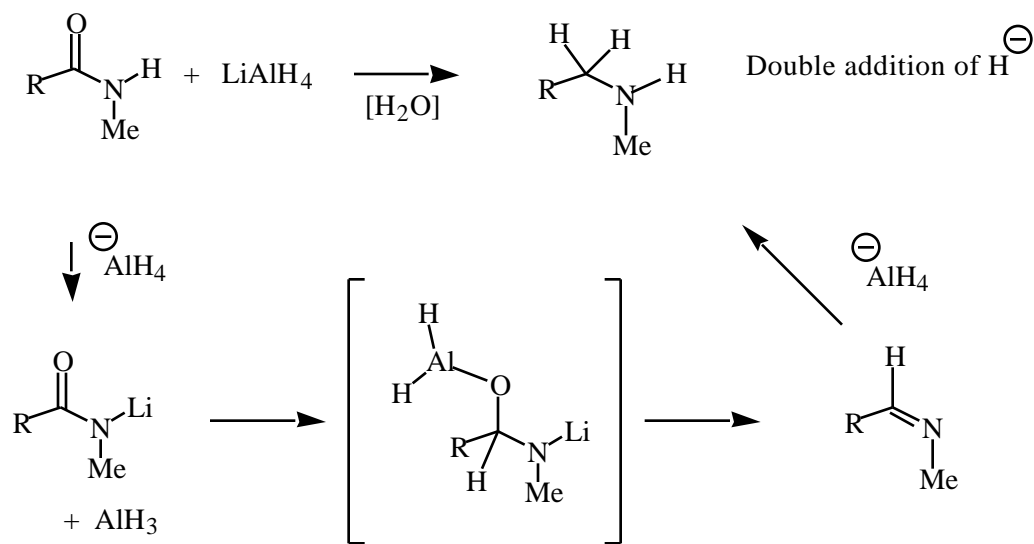
Same for esters, anhydrides.

But

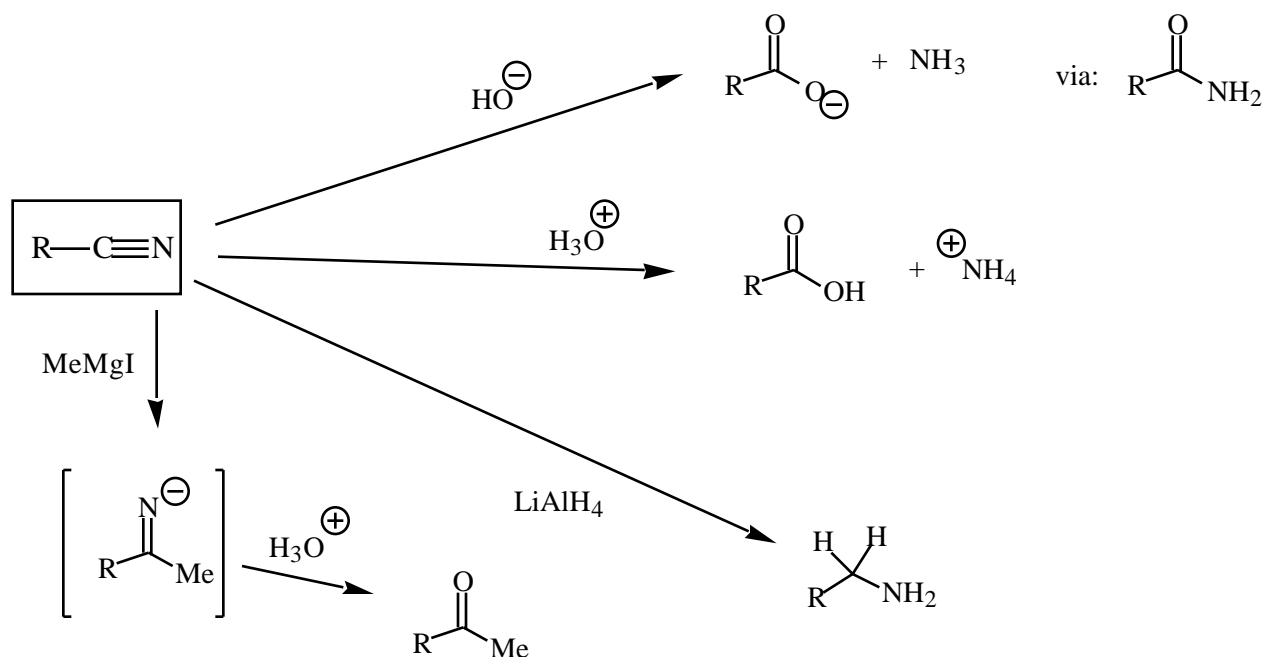


## C. Addition of metal hydrides





### Analogy with Nitriles:



**Write mechanisms:** multistep

### Enolate chemistry of esters:

