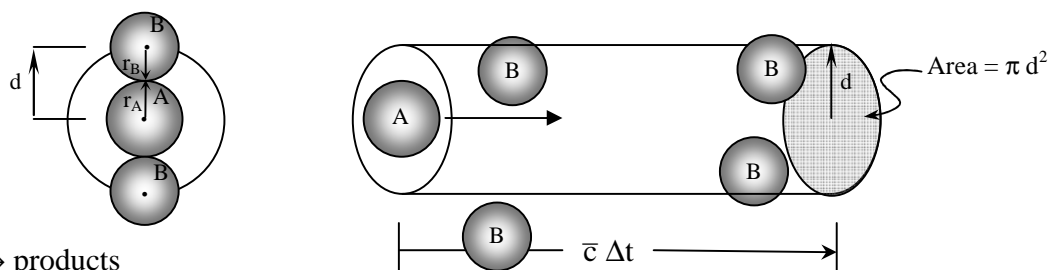


## Collision Theory



$N_A$  and  $N_B$  molecules in a container of volume  $V$  gives:  
number densities of molecules:  $\rho_A = N_A/V$  and  $\rho_B = N_B/V$

hard core collision cross section:  $\sigma = \pi (r_A + r_B)^2 = \pi d^2$   
 $d$  is the collision diameter

molecule  $A$  moving, all others fixed:  
in time  $\Delta t$ ,  $A$  sweeps out volume  $= \sigma \bar{c} \Delta t$

number of molecules inside this volume  $= \sigma \bar{c} \Delta t N_B/V$

collision frequency = number of collisions per unit time  $= \sigma \bar{c} N_B/V$

total collisions  $= N_A \sigma \bar{c} N_B/V$

total collisions per unit volume:  $= \sigma \bar{c} (N_A/V)(N_B/V) = \sigma \bar{c} \rho_A \rho_B$

$B$  molecules are moving:  $z_{AB} = \sigma \bar{c}_{\text{rel}} (N_A/V)(N_B/V)$

collisions per unit volume per unit time:  $z_{AB} = \sigma \left( \frac{8kT}{\pi\mu} \right)^{1/2} \left( \frac{N_A}{V} \right) \left( \frac{N_B}{V} \right)$  with  $\mu = \frac{m_A m_B}{m_A + m_B}$

reaction rate:  $-\frac{d(N_A/V)}{dt} = z_{AB} = \sigma \bar{c}_{\text{rel}} \left( \frac{N_A}{V} \right) \left( \frac{N_B}{V} \right)$   $[A] = \frac{1}{N_A} \left( \frac{N_A}{V} \right)$

reaction rate:  $-\frac{d(N_A/N_A V)}{dt} = z_{AB} = \sigma \bar{c}_{\text{rel}} N_A \left( \frac{N_A}{N_A V} \right) \left( \frac{N_B}{N_A V} \right)$

$-\frac{d[A]}{dt} = z_{AB} = \sigma \bar{c}_{\text{rel}} N_A [A] [B]$

$-\frac{d[A]}{dt} = k_2 [A] [B]$   $k_2 = \sigma \bar{c}_{\text{rel}} N_A = \sigma \left( \frac{8kT}{\pi\mu} \right)^{1/2} N_A = \pi d^2 \left( \frac{8kT}{\pi\mu} \right)^{1/2} (1000 \text{ L/m}^3) N_A$

$H_2 + I_2 \rightarrow 2 HI$   $r_{H_2} = 1.1 \text{ \AA}$   $r_{I_2} = 1.7 \text{ \AA}$   $\sigma = 0.246 \text{ nm}^2$   $\bar{c}_{\text{rel}} = 2060 \text{ m s}^{-1}$   
 $k_2 = \pi (1.1 \times 10^{-10} + 1.7 \times 10^{-10} \text{ m})^2 \left( \frac{8(1.381 \times 10^{-23} \text{ J K}^{-1})(400 \text{ K})}{\pi 3.33 \times 10^{-27} \text{ kg}} \right)^{1/2} (1000 \text{ L/m}^3) 6.022 \times 10^{23} \text{ mol}^{-1}$   
 $= 3.1 \times 10^{11} \text{ L mol}^{-1} \text{ s}^{-1}$