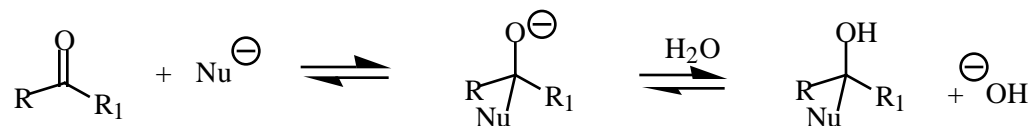


**Review:**

Base Catalyzed 1,2-addition to carbonyl groups:  $\text{H-Nu} + \text{base} \rightleftharpoons \text{Nu:}^- + \text{H-base}^+$



$\text{R}_1 = \text{H}$ , carbon group

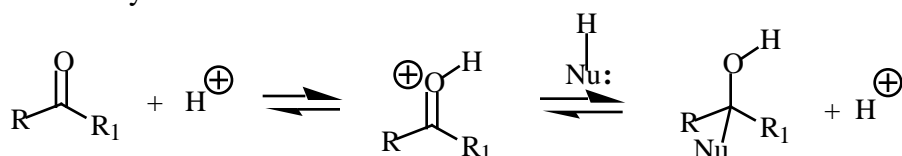
**Reversible** if  $\text{Nu}^-$  is "stabilized"

e.g.,  $\text{ROH}$ ,  $\text{R}_2\text{NH}$ ,  $\text{RO}^-$ , certain carbon anions

**Irreversible** with simple carbon anions  
[organometallic reagents]

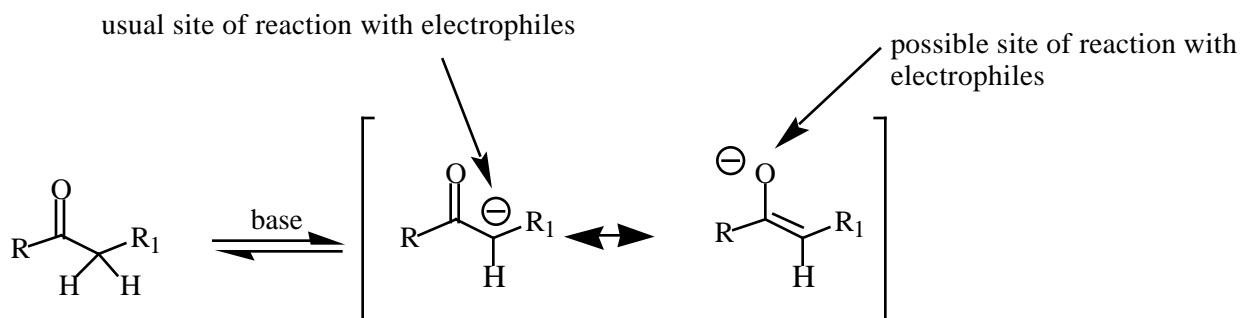
$\text{R-Li}$ ,  $\text{R-MgBr}$ ,  $\text{R-Zn-Br}$

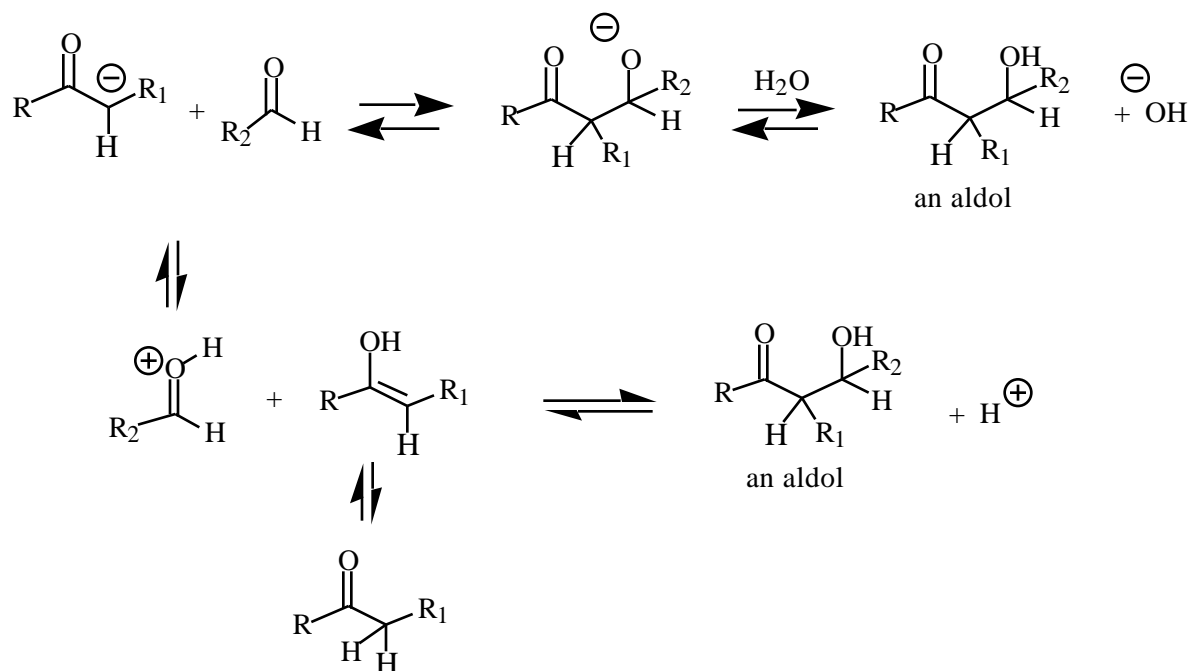
Acid catalyzed:



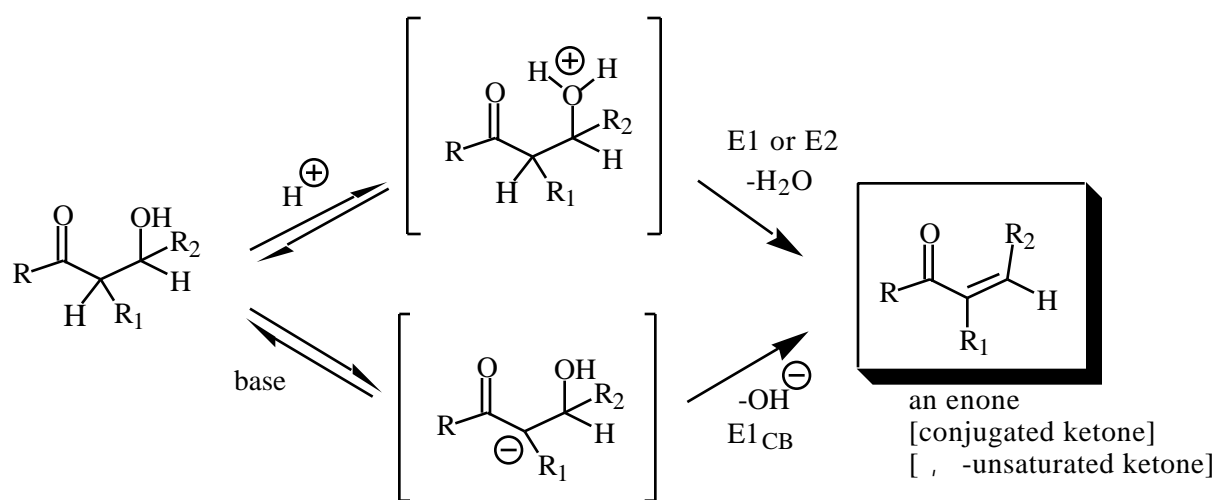
Equilibration is generally favorable ( $K > 1$ ) for aldehydes and unfavorable ( $K < 1$ ) for ketones, but important electronic and steric effects can influence  $K_{\text{eq}}$ .

Special case of nucleophile: **enolate anion** by proton abstraction from a ketone or aldehyde

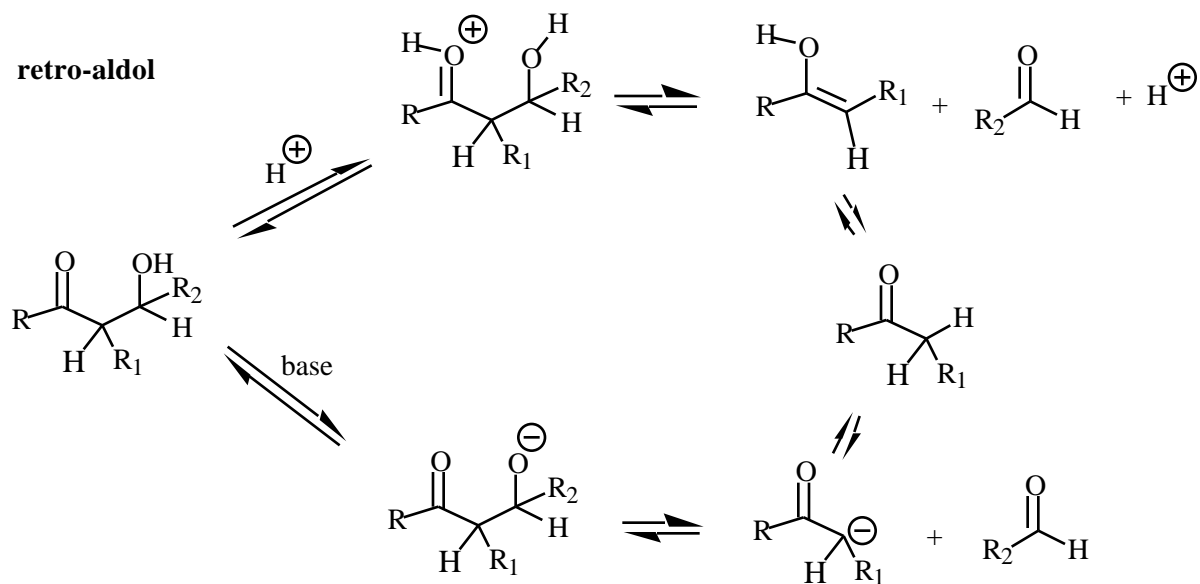




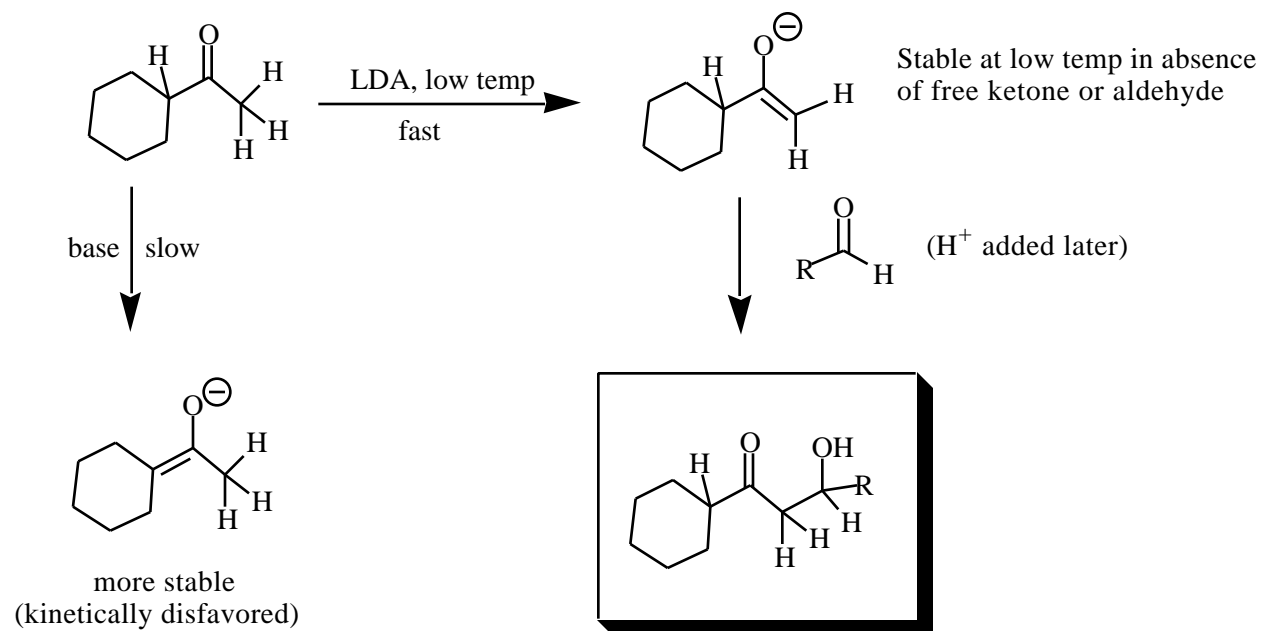
### Final Stage:



Makes the process irreversible, pulls the various equilibria steps to the right.



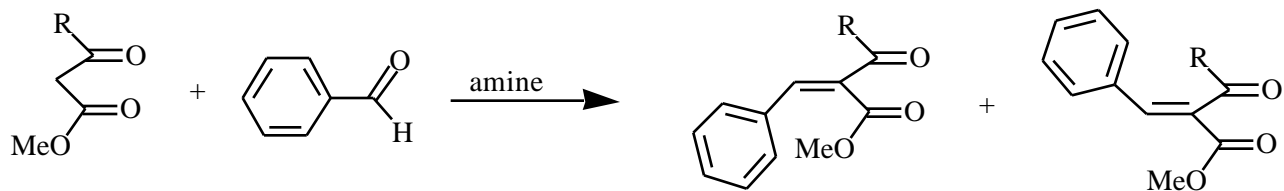
**Modern:** Selective generation of the enolate anion, add other carbonyl component.



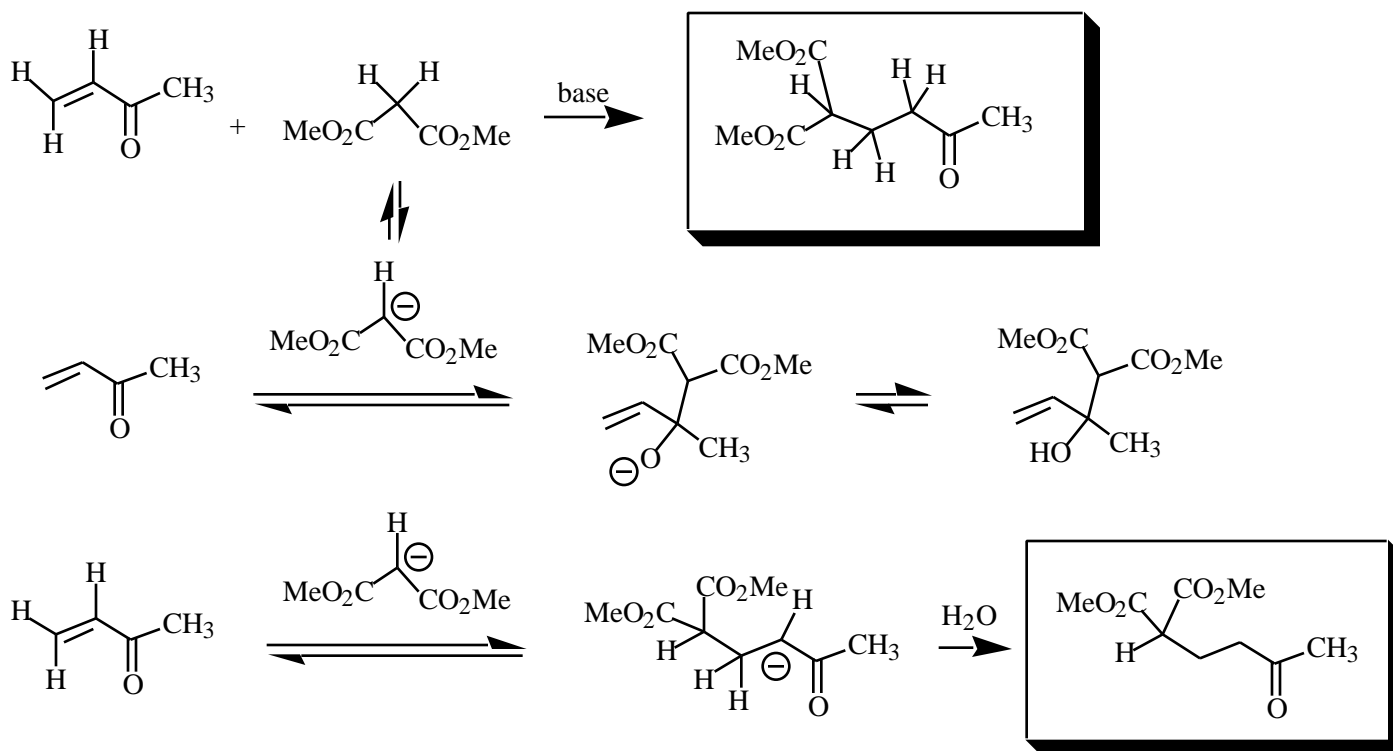
## Classics:

4

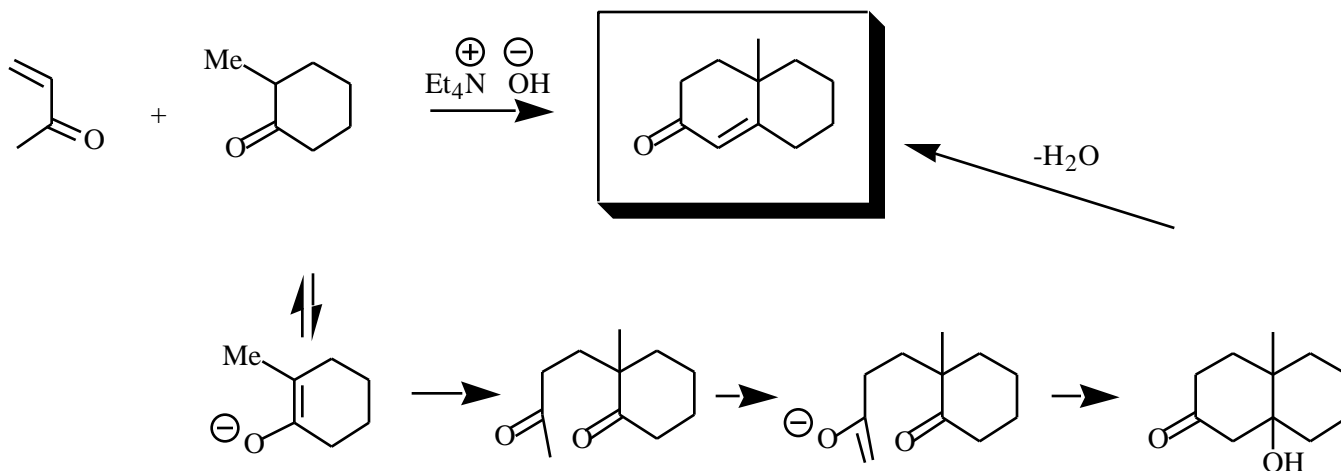
Very stabilized carbanion in one component: Knoevenagel Condensation  
elimination of water to drive the process:



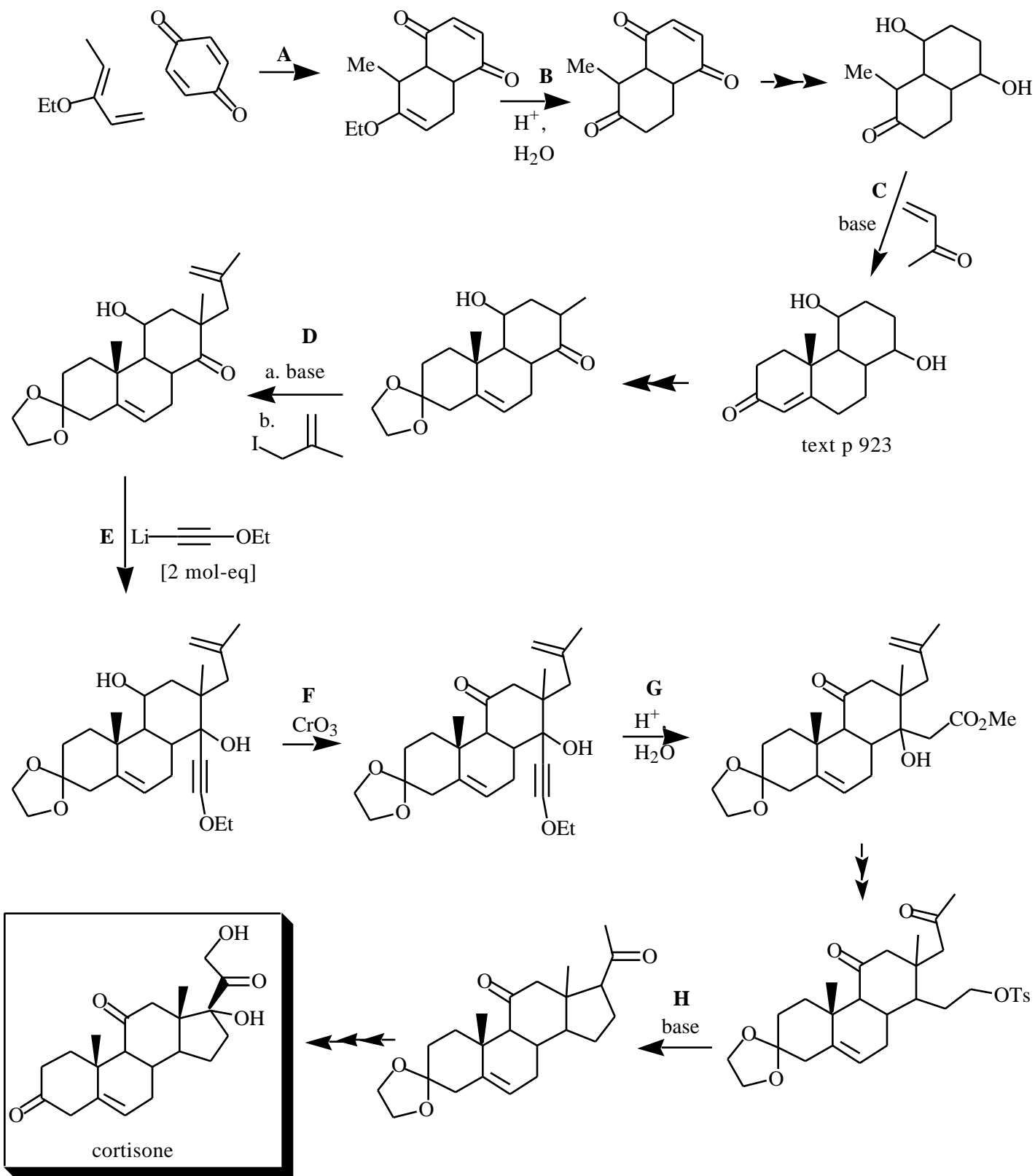
## "Conjugate Addition" The Michael Reaction



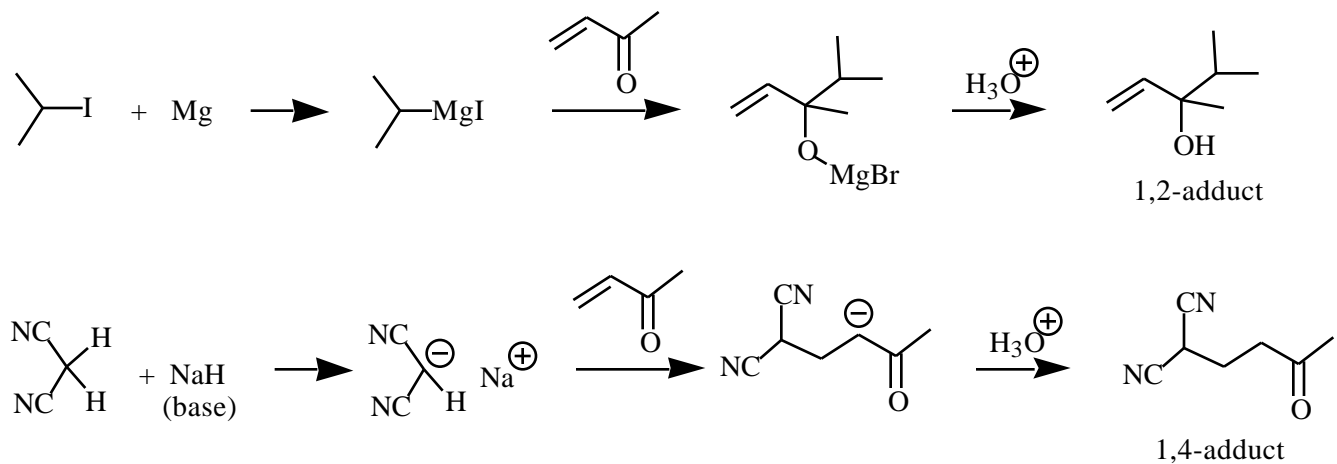
## Combination condensation reactions: Robinson Annulation



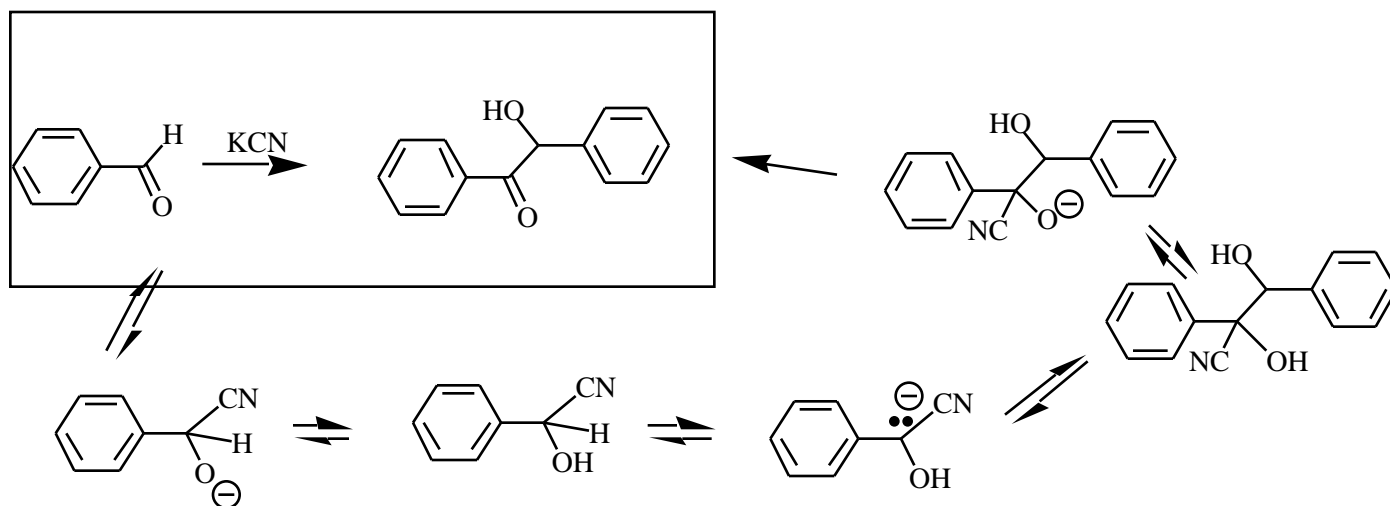
## Sarett's Cortisone synthesis:



Note: Reactive Anions tend to react "1,2" at the carbonyl carbon and not reverse

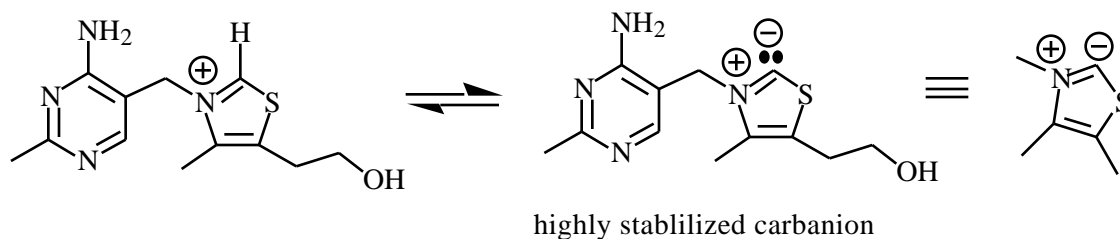


### The Benzoin reaction and Vitamin B1 [thiamine] Very old:

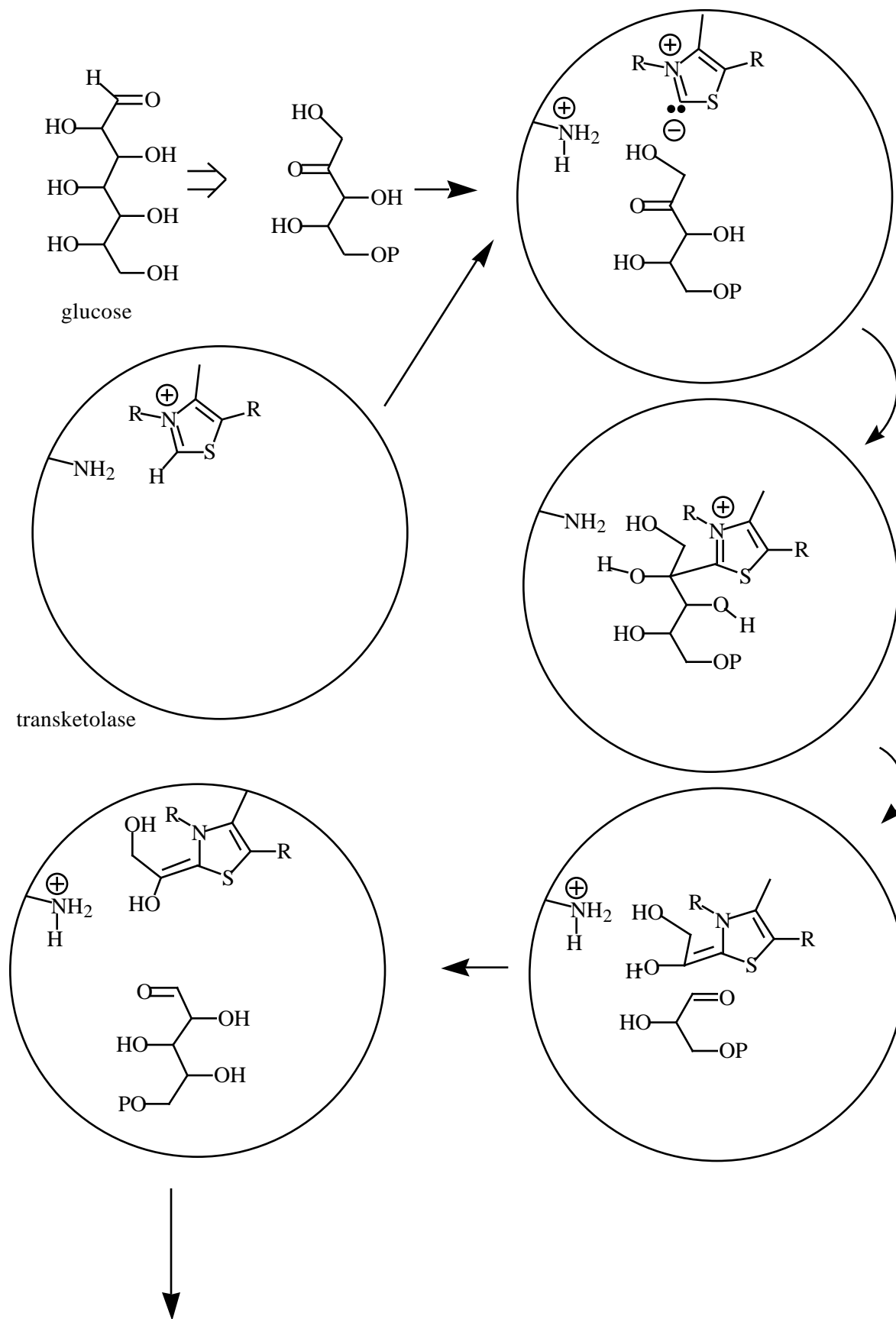


Role of CN: Add to  $\text{C}=\text{O}$ ; favor deprotonation of  $-\text{H}$   
 Easily formed nucleophile and then a good electron-withdrawing group.  
 [e.g.,  $\text{HO}^-$  does not do the same job. Why?]

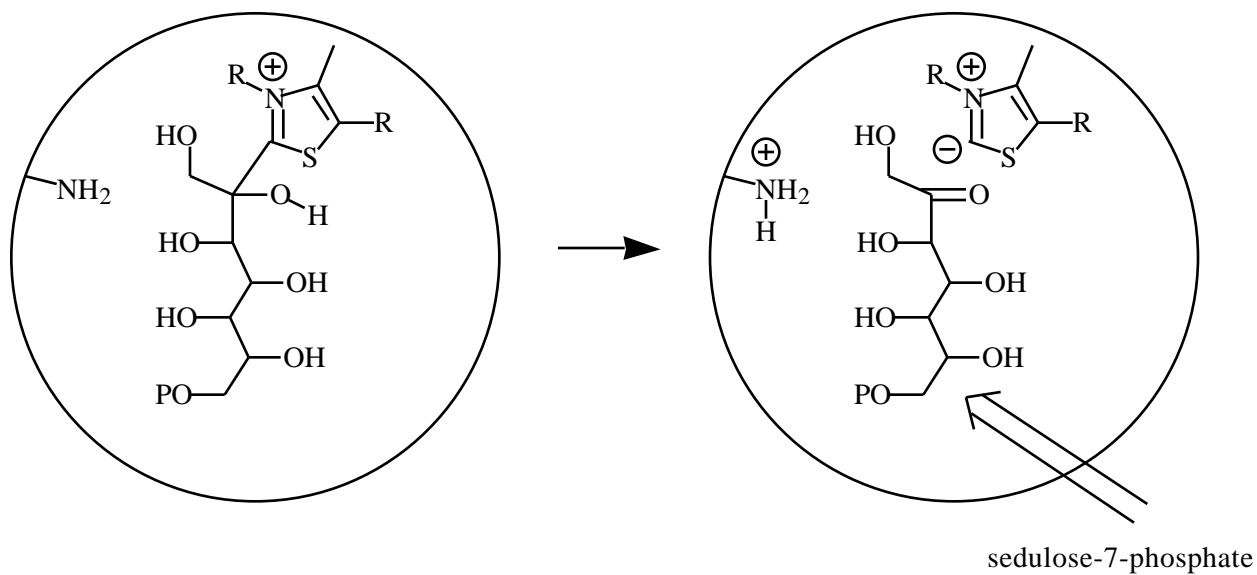
Nature's analog of CN: Thiamine-- $\text{pK}_\text{a}$  12.7











### Parallel Role of Imines/immonium ions/enamines:

