

nanomaterials

Green Synthesis of Nanomaterials

Edited by
Giovanni Benelli

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Green Synthesis of Nanomaterials

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Special Issue Editor

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About the Special Issue Editor

Giovanni Benelli serves as Senior Research Entomologist at the Department of Agriculture, Food and Environment, University of Pisa, Italy. He received an International Ph.D. in Agrarian and Veterinary Sciences from University of Pisa and Sant'Anna School of Advanced Studies. Giovanni has worked in several international institutions, including University of Hawaii at Manoa (USA) and University of Jaén (Spain). Giovanni's research focuses on insect behaviour, insect-inspired robotics, and biological control, covering agricultural pests as well as vectors of medical and veterinary importance. He has cooperated with a large number of researchers worldwide on various research projects (e.g., FP7 CoCoRo and H2020 subCULTron). Giovanni serves as Editor for several top-ranked international journals with an impact factor, including *Acta Tropica* (Editor in Chief), *Insects*, *Entomologia Generalis*, *Environmental Science and Pollution Research*, *Molecules*, *Nanomaterials*, *International Journal of Environmental Research and Public Health*, *Animals*, *Journal of King Saud University—Science*, *Symmetry*, and others. He has been awarded with various research prizes from international and national organizations, including the Odile Bain Memorial Prize 2018 (Parasites and Vectors and Boehringer Animal Health) and the Antico Fattore Prize 2016 (Accademia dei Georgofili, Firenze).

Editorial

Green Synthesis of Nanomaterials

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Nanomaterials possess stunning physical and chemical properties. They have a major role in the development of novel and effective drugs, catalysts, sensors, and pesticides. The synthesis of nanomaterials is usually achieved via chemical and physical methods, both of which require the use of extremely toxic chemicals or high-energy inputs. To move towards more environmentally friendly processes, researchers have recently focused on so-called “green synthesis”, in which microbial, animal-, and plant-borne compounds can be used both as economic forms of waste reduction and stabilizing agents to fabricate nanomaterials. Green synthesis routes have been proposed as cheap and environmentally sustainable; they can lead to the fabrication of nano-objects with controlled size and shape, two key features affecting their bioactivity.

However, real-world applications of green-fabricated nanomaterials are still largely unexplored, and the number of marketed products is limited. One may question our knowledge about their non-target toxicity and their main modes of action. Such questions also bring up issues regarding their possible fate in the environment [1,2].

In this framework, the present Special Issue includes studies by top-ranked experts on nanosynthesis and related applications. This Special Issue includes articles on relevant and pressing issues in green nanomaterial science. Most are original research articles and all highlight theoretical concepts and practical protocols of interest for real-world applications related to nanomaterials.

Recent approaches to synthesize nanomaterials have focused on the use of various natural products. For example, healthy and microwave-injured bacteria have been used to produce finely characterized palladium nanoparticles [3]. Furthermore, plant-borne products have been employed to produce interesting nanomaterials, such as chocolate extract-fabricated Au nanoparticles with good biocompatibility features [4] as well as gum kondagogu-synthesized anatase TiO₂ nanoparticles stable at high temperatures, which are relevant for the photocatalytic degradation of organic dyes [5]. Other nanomaterials produced include ZnO nanoparticles, which have been reduced and stabilized using the *Scadoxus multiflorus* leaf aqueous extract. These nanoparticles showed relevant antifungal and insecticidal activity and are highly effective against young instars (eggs and larvae) of *Aedes aegypti* (Diptera: Culicidae) [6], an important dengue and Zika virus mosquito vector [7].

Other valuable applications ranged from enzyme technology to biocatalysis; efforts to develop biocompatible antimicrobial surfaces have been studied by Aflori et al. [8], relying to the employ of poly-L-lactic acid. Silva et al. [9] shed light on polydopamine-mediated green reduction of graphene oxide whereas Fotiadou et al. [10] successfully designed lipase-hybrid nanoflowers, which have been enriched with carbon and magnetic nanomaterials, allowing biocatalytic transformations. Carbon nanochemistry is the focus of research by Kukulka et al. [11], which proposed a novel and reliable time-dependent facile synthesis of carbon solid spheres.

Additional studies are dedicated to other important nanomaterials, including silica-based nanomaterials. Because the granulometric characterization of silica nanomaterials requires harmonized protocols, Retamal Marin et al. [12] proposed a novel approach to investigate the impact of sample preparation on suspended nanosilica size.

Two research papers focused on the peculiarities of nanocrystals, proposing a robust protocol for the solvothermal synthesis of nanocrystalline CuInS₂ thin films [13] and how to synthesize photofunctional mesocrystals through a polyol-based fluoride ion slow-releasing approach [14].

The Special Issue also contains a review [15] on micro- and nanoemulsions, which covers theory and practice about their preparation as well as novel applications in the fields of entomology and parasitology. A growing number of recent papers have stressed the important advantages arising from the use of green micro- and nanoemulsions to enhance the efficacy and stability of selected bioactive compounds of natural origin. Micro- and nanoemulsions of selected natural products have been successfully proposed for the management of parasites and vectors of interest for public health (e.g., mosquitoes and ticks) as well as the control of insect and mite species of agricultural importance. In the final section, the review also highlights challenges and constraints arising from the use of green micro- and nanoemulsions to promote their commercial development for various biological and biomedical purposes [15].

Overall, as the Editor of *Nanomaterials*, I am fully aware that the present Special Issue cannot fully reflect the high diversity and creativity of new concepts and tools rapidly developing in this multidisciplinary research field. However, I am confident that this focus on the green synthesis of nanomaterials will contribute to the research interest in the field, providing our readership with a multi-faceted scenario that outlines the importance of cross-field green nanoresearch, its quick growth, as well as its wide-ranging applications.

It is also expected that the present Special Issue will encourage multidisciplinary research on green nanomaterials, broadening the range of potential practical uses. This needs to be coupled with research efforts improving large-scale synthesis and economic viability of the proposed processes, ecotoxicology insights to understand the post-application fate of green nanomaterials, and their long-term stability and impact on human health and the environment.

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Article

Characterization of Palladium Nanoparticles Produced by Healthy and Microwave-Injured Cells of *Desulfovibrio desulfuricans* and *Escherichia coli*

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Abstract: Numerous studies have focused on the bacterial synthesis of palladium nanoparticles (bio-Pd NPs), via uptake of Pd (II) ions and their enzymatically-mediated reduction to Pd (0). Cells of *Desulfovibrio desulfuricans* (obligate anaerobe) and *Escherichia coli* (facultative anaerobe, grown anaerobically) were exposed to low-dose radiofrequency (RF) radiation(microwave (MW) energy) and the biosynthesized Pd NPs were compared. Resting cells were exposed to microwave energy before Pd (II)-challenge. MW-injured Pd (II)-treated cells (and non MW-treated controls) were contacted with H₂ to promote Pd(II) reduction. By using scanning transmission electron microscopy (STEM) associated with a high-angle annular dark field (HAADF) detector and energy dispersive X-ray (EDX) spectrometry, the respective Pd NPs were compared with respect to their mean sizes, size distribution, location, composition, and structure. Differences were observed following MWinjury prior to Pd(II) exposure versus uninjured controls. With *D. desulfuricans* the bio-Pd NPs formed post-injury showed two NP populations with different sizes and morphologies. The first, mainly periplasmically-located, showed polycrystalline Pd nano-branches with different crystal orientations and sizes ranging between 20 and 30 nm. The second NPpopulation, mainly located intracellularly, comprised single crystals with sizes between 1 and 5 nm. Bio-Pd NPs were produced mainly intracellularly by injured cells of *E. coli* and comprised single crystals with a size distribution between 1 and 3 nm. The polydispersity index was reduced in the bio-Pd made by injured cells of *E. coli* and *D. desulfuricans* to 32% and 39%, respectively, of the values of uninjured controls, indicating an increase in NP homogeneity of 30–40% as a result of the prior MWinjury. The observations are discussed with respect to the different locations of Pd(II)-reducing hydrogenases in the two organisms and with respect to potential implications for the catalytic activity of the produced NPs following injury-associated altered NP patterning.

Keywords: palladium nanoparticles; microwave injured cells; microwave energy; *Escherichia coli*; *Desulfovibrio desulfuricans*

1. Introduction

Platinum group metals (PGMs) (e.g., Pd, Pt, Ru, Rh, Os, Ir) are widely used as catalysts in many different reactions to obtain valuable products with industrial applications [1]. They are of particular importance due to their unique properties (i.e., high catalytic activity, oxidation resistant properties, mechanical strength, and outstanding resistance to corrosion) [2]. PGM catalysts are used to control the emission of gaseous pollutants from automobiles. Due to this high global demand the price of PGMs has increased substantially since the mid-2000s [3], while high demand for PGMs [4,5] has also

increased the focus on recovery processes. Chemical methods offer an alternative for PGM recovery from wastes; these methods include ion exchange, solvent extraction or electrochemical recovery, but they have the challenge of using strong chemicals which are often toxic and environmentally damaging [6].

Bacterially-mediated recovery of PGMs is considered as an emerging green and cheap alternative to traditional physical and chemical approaches. Bio-derived methods can exhibit numerous advantages since bacterial species used as templates are easy to grow in large amounts and are capable of rapid metal reduction to form metallic nanoparticles that are comparably active to commercial catalysts in chemical synthesis [7].

Bacteria can interact with soluble metal species in many different ways (e.g., via enzymatic reduction, biosorption, biomineratization, etc.) [8–10]. Bacterially-mediated reduction of metals into a neo-catalyst has attracted much interest with other potential applications in, for example, fuel cells [11], decontamination of groundwater [12], and catalytic upgrading of heavy fossil and pyrolysis oils [13,14]. Some microorganisms are able to recover Pd (II) from acidic solutions similar to the conditions that are present in industrial wastes (and from actual waste leachates) and convert waste PGMs into a green neo-catalyst [15,16]. A life cycle analysis of the latter, as applied to catalytic upgrading of heavy fossil oil, showed the economic potential of this approach even before factoring in the energy (carbon) savings in refinery and mitigation of the high carbon impact and environmental damage involved in mining and metal extraction from primary ores [17].

The use of bacteria for synthesis of metallic nanoparticles (NPs) offers the advantage of NP size control via bio-patterning and the use of enzymes for the Pd reduction avoids the use of toxic chemicals as capping agents that would add to the process cost [18]. Additionally, living systems operate at ambient temperatures, making the process of synthesis of NPs economically attractive. For example, *Desulfovibrio desulfuricans*, a Gram-negative strain, has been shown to use periplasmic hydrogenases supplied with hydrogen to form Pd NPs in the periplasm [9]. NP-synthetic capability has been shown also in other Gram-negative bacteria like *Shewanella oneidensis*, *Escherichia coli*, and *Pseudomonas putida* [7,19–21] as well as Gram-positive bacteria such as *Bacillus sphaericus* and *Arthrobacter oxydans* [22,23]. With the use of modern microscopes, recent studies reported the accumulation of small intracellular Pd NPs in both bacterial types [24], as well as in cell surface layers. Although the former brings possible limitations of substrate access, the use of acetone-washed cells permeabilizes them, whereas NPs stripped of their biochemical scaffold agglomerated and lost activity [25], while partial cleaning altered the catalytic activity as the metal surface was progressively unmasked [26]. However, such processing would add to the production cost and hence this study reports the use of a supported Pd catalyst made on whole cells.

In addition to cellular location, particle size, and shape, dispersity can play an important role in catalyst reactivity in some reactions [27]. In the case of microbial synthesis of Pd NPs, some studies have shown an influential role of the biological component in the control of shape, size, and distribution of NPs and, as a consequence, their catalytic activity [7]. A possible association of the initial Pd “seeds” with intracellular phosphate structures has been postulated in cells of *Bacillus benzoevorans*, preventing the Pd NPs from agglomeration [24]. Electron donors such as formate or hydrogen used in NP fabrication can influence the sizes of the biochemically-formed Pd NPs and, with this, their electrocatalytic activity [28]. Taking into account these different factors a method of manipulating the formation of NPs to influence their size and distribution could result in a tailored catalyst for increased reaction rates and selectivity in a given reaction.

The main challenges that face the synthesis of nanoparticles are: control of the size and shape of the NPs and monodispersity. It is known that thermal factors can affect the size and uniformity of nanoparticles [29]. Localized heating can be achieved by the use of microwave radiation. Any material (but particularly water) can absorb microwave energy and this is expressed by its dielectric loss factor combined with the dielectric constant. When the microwave heats the desired material through the dielectric loss, it converts the radiation energy into thermal energy [30]. In organic synthesis this

has been shown to accelerate processes involved in homogeneous catalysis [31]. The efficiency of microwave energy for the synthesis of a variety of nanomaterials including metals, metal oxides, and bimetallic alloys has been shown [32]. The effect of microwave (MW) radiation on microorganisms has also been studied [33,34]. Some authors noted that application of radiofrequency (RF) microwave radiation (2.45 GHz) altered the activities of some enzymes expressed in *Staphylococcus aureus* resulting from some changes in the cell that could not be explained by the thermal effect [35]. More recently, Shamis et al. [34] confirmed that MW radiation on cells could result in toxic effects when the heating effect was discounted. By modulating the frequency of the MW radiation [36] different biological effects in terms of protein structures were observed, together with alterations in the routes of some biochemical reactions.

In this study MW energy was applied to cells of *E. coli* and *D. desulfuricans* before their exposure to palladium solution. Following the MW treatment, synthesis of Pd NPs was performed using molecular hydrogen as the electron donor. Characterization of size, shape, cellular localization, and atomic structure of the fabricated NPs was conducted by means of scanning transmission electron microscopy (STEM) associated with a high-angle annular dark field (HAADF) detector and energy dispersive X-ray micro analysis (EDX). The use of X-ray diffraction analysis of bulk material was largely precluded by the small nanoparticle sizes and hence poorly resolved powder patterns of the largely amorphous biomaterial [37]. The possible application of the MW treatment to moderate the synthesis of more dispersed and homogeneous Pd NPs is discussed with reference to data obtained from high-resolution electron microscopy in conjunction with image analysis.

2. Materials and Methods

2.1. Bacterial Strains and Culture Conditions

Two Gram-negative bacterial strains were used in this study, a facultatively anaerobic strain *Desulfovibrio desulfuricans* NCIMB 8307 and the facultatively anaerobic *Escherichia coli* MC4100 as described previously [19,24]. *D. desulfuricans* was grown anaerobically under oxygen-free nitrogen (OFN) in Postgate's medium C (Sigma-Aldrich) ($\text{pH } 7.5 \pm 0.2$) at 30°C (inoculated from a 24 h pre-culture, 10% v/v) in sealed anaerobic bottles [24], while *E. coli* was grown anaerobically (37°C) on nutrient broth N° 2 (Sigma-Aldrich) supplemented with 0.5% (v/v) glycerol (Sigma-Aldrich) and 0.4% (w/v) fumarate (Sigma-Aldrich) as described previously [19]. Cells were grown to mid-exponential phase and harvested (Beckman Coulter Avati J-25 Centrifuge, U.S.A) by centrifugation ($12,000 \times g$, 15 min), washed 3 times in 20 mM MOPS-NaOH buffer (pH 7), concentrated in a small volume of buffer to between 20 and 30 mg dry weight per ml and stored under OFN at 4°C until next day [38]. Cell dry weight was calculated from optical density (OD_{600}) by a previously-determined dry weight conversion factor (mg dry cells = CF \times $\text{OD}_{600} \times n$, (where n is the dilution factor)).

2.2. Microwave Irradiation of *E. coli* and *D. desulfuricans* Cells

2.2.1. Microwave Irradiation Conditions

This study was carried out using a portable commercial apparatus (CEM corporation, North Carolina, United States) (CEM Discover SP microwave digestion system; single-mode energy source; 300 W magnetron; ~3 GHz, 300 W). Vials containing cells in 6mL volume re-suspended in 20 mM buffer with concentration between 20 and 30 mg dry weight/ml were exposed in short bursts (10 s) interspersed with periods of 30 s of cooling in ice cold acetone after exposure. This process was repeated three times (total irradiation period of 30 s). During the microwave irradiation sample vials were cooled in hexane.

2.2.2. Microwave Irradiation of Resting Cells Suspended in MOPS Buffer

A 5 mL volume of concentrated cell suspension between 20 and 30 mg/mL in 20 mM MOPS-NaOH buffer (Sigma-Aldrich) pH 7 was added into a 6 mL sealed tube under OFN and treated as above. After microwave exposure, a known amount of treated cells was taken and added immediately to a new sealed tube containing Pd (II) solution (below), representing a final 5 wt% of Pd on the cells. As a control, a 6mL sealed tube under OFN of Pd(II) solution in buffer was added and exposed to microwave radiation under the same conditions as above.

2.3. Preparation of Palladium-Challenged Cells

2.3.1. Palladium Solution

For “palladization” of cells a Pd(II) solution was used: 2 Mm Na₂PdCl₄ (Sigma-Aldrich, St. Louis, Missouri, United States) pH 2 in 0.01 M HNO₃ placed in sealed tubes (final volume of 6 mL) and degassed with oxygen-free nitrogen (OFN) under vacuum prior to addition of bacteria.

2.3.2. Formation of PdNanoparticles by Control and MW-Treated Cells

Following microwave treatment, tubes with cells (and untreated controls) were allowed to stand in a water bath (30 min, 30 °C) for uptake of the Pd (II) ions in order to form nucleation sites on the biomass. Hydrogen was added as an electron donor by bubbling H₂ gas through the suspensions in the sealed bottles (15 min) which were left under H₂ (24 h) for complete reduction of palladium on the cells (confirmed by assay of residual soluble Pd (II)). Palladized cells were harvested by centrifugation (12,000× g, 15 min) and washed with distilled water twice prior to fixation (2.5% glutaraldehyde in 0.1 M cacodylate buffer (pH 7)) for examination by electron microscopy. Controls of palladium-challenged cells without MW treatment were prepared in the same way.

2.3.3. Residual Pd(II) Quantification Using the Tin(II) Chloride Method

In order to confirm complete depletion of Pd (II) ions from the solution, the spectrophotometric tin (II) chloride-based method was used as described previously [39].

2.4. High-Resolution Scanning Transmission Electron Microscopy (STEM) with HAADF (High-Angle Annular Dark Field) Detector and EDX Analysis

For STEM analysis, thin sections of palladized MW-treated and non-treated *E. coli* and *D. desulfuricans* cells were prepared according to the method described by Merroun et al. [40]. To determine the location of palladium NPs in the cells, palladized cells were examined under a high-angle annular dark field scanning transmission electron microscope (HAADF–STEM) FEI TITAN G2 80–300 at 300 KeV. For elemental analysis, EDX (energy dispersive X-ray microanalysis of specimen microareas) was used with a spot size of 4 Å and a live counting time of 50s coupled with a high-resolution STEM and HAADF detector. Element co-localizations (Pd, P, S) were enumerated by use of the Manders overlap coefficient (MOC) [41] implemented in ImageJ via the JACoP [38] and co-localization was assumed if the Manders coefficient was greater than 0.9.

2.5. Image Processing, Lattice Spacing Determination and Particle Size Analysis

The HAADF–STEM images were used to determine the average size of Pd NPs produced in different experiments as well as their distribution by means of the image processing software ImageJ (National Institutes of Health, Maryland, United States) [42]. In order to distinguish the Pd nanoparticles on/in cells from background signals and artifacts the same methodology as described by Omajali et al. [24] was used and mean particle size was calculated (mean ± SEM from at least 3 different areas of samples; total NPs analyzed was > 100). Significant differences were assigned using the two sample test of the variance at P = 0.95. The polydispersity index or coefficient of variation

was calculated from the particle size distribution dividing the standard deviation of the means by the means [43]. At least 100 particles were counted in each case using ImageJ software. The particle size distribution was estimated using Origin Pro 8. The lattice spacing was determined using ImageJ through profiling of HR-TEM images and compared against lattice spacing of bulk palladium from the database generated using Powder cell 2.400 software (2.4, Informer Technologies, Inc.).

3. Results

3.1. Microwave-Injury of Bacteria

Examination of the cells by electron microscopy post-injury showed cellular damage as compared to uninjured controls (Supplementary information Figure S1), similar to the response observed by Shamis et al. [34], i.e., shrinkage of the cytoplasmic compartment away from the wall layers with an enlarged periplasmic space. Shamis et al. [34] attributed this to the electromagnetic radiation and not the heating effect. Even though cooling was applied, it was not possible to make this distinction unequivocally using the commercial equipment in this study. Instead, a prior study [44] was carried out using purpose-built equipment that decoupled the electromagnetic and thermal effects at a comparable applied dose (Supplementary information Figure S1). This showed an identical cellular response visually to that using the commercial equipment (with cooling) in the present study and hence, as noted by Shamis et al. [34] the reported cellular response can be attributed to the effect of the MW irradiation.

3.2. Examination of the Pd Nanoparticles Produced by Native and MW-Injured Cells of *E. coli* MC4100 and *D. Desulfuricans* NCIMB 8307

Controlsof Pd(II) solution in buffer alone exposed to microwave radiation showed no removal of Pd (II) from the solution after microwave exposure usingthe tin (II) chloride method, nor the appearance of any black Pd(0), indicating that microwave irradiation had no active role in the reduction of Pd (II) under the condition tested in this work.

In order to examine and identify the Pd NPs made by native and injured cells, high-resolution HAADF-STEM with EDX was used. For both strains the Pd loading was 5 wt% (1:19 mass of Pd:dry weight of cells). Electron opaque NPs, identified as Pd by EDX (Supplementary information Figure S2) were found in the cell surface layers and within the intracellular matrices (Figure 1B,E and Figure 2B,E). In the case of untreated cells of *E. coli* (Figure 1A–C) large clusters were observed within the intracellular matrices (Figure 1C inset bottom left) and membrane (Figure 1C, main panel, arrowed) at high magnification while MW-treated cells showed apparently more dispersed intracellular NPs with few clusters (Figure 1F).

In contrast, untreated cells of *D. desulfuricans* showed a deposition of surface bound NPs in clusters (Figure 2B,C) in agreement with Omajali et al. [24]. Pd NPs located at the level of the periplasm showed inclusions in the form of nano-branches with sizes ranging from 20 to 30 nm (Figure 2C arrowed), while intracellular NPs were visible which were smaller with sizes between 1 and 5 nm (Figure 2C inset top left). Following MW treatment, and in contrast to *E. coli* (above), the cytoplasmic compartment of MW-treated *D. desulfuricans* remained contracted to reveal NPs in the outer and inner membranes (Figure 2D) with NP-deposition also in the periplasmic space (Figure 2D circled area). The main differences were in morphology of the NPs observed at the level of the surface (Figure 2F) in comparison with untreated cells (Figure 2C), where larger clusters were observed at high magnification. No major differences in number and morphology of Pd NPs were apparent visually by HR-TEM alone within the intracellular matrices in treated (Figure 2C top left) and untreated cells of *D. desulfuricans* (2F bottom right).

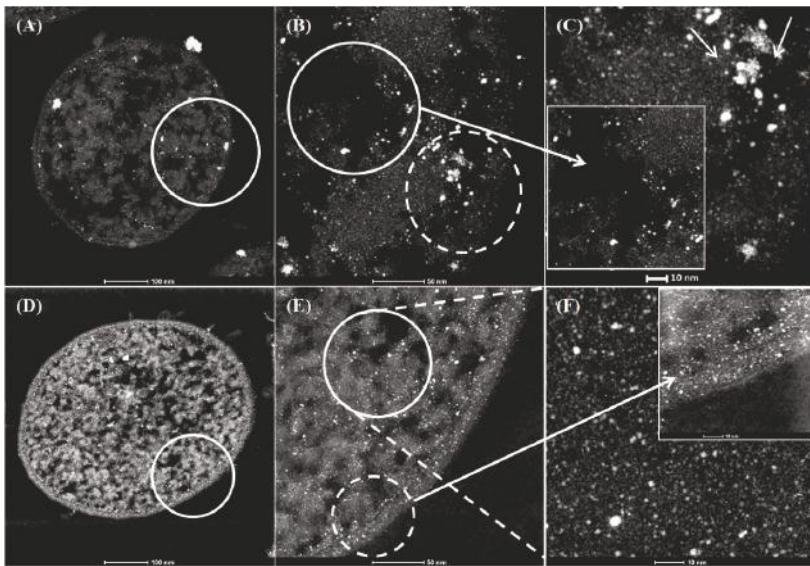


Figure 1. High-angle annular dark field scanning transmission electron microscope (HAADF-STEM) micrographs of Pd nanoparticles synthesized using 5 wt% Pd loading (1:20) on *E. coli* MC4100 from 2 mM Na_2PdCl_4 solution, in 0.01 M HNO_3 using hydrogen as an electron donor without prior microwave (MW) treatment (A,B,C) and with 30 s prior MW treatment (D,E,F).

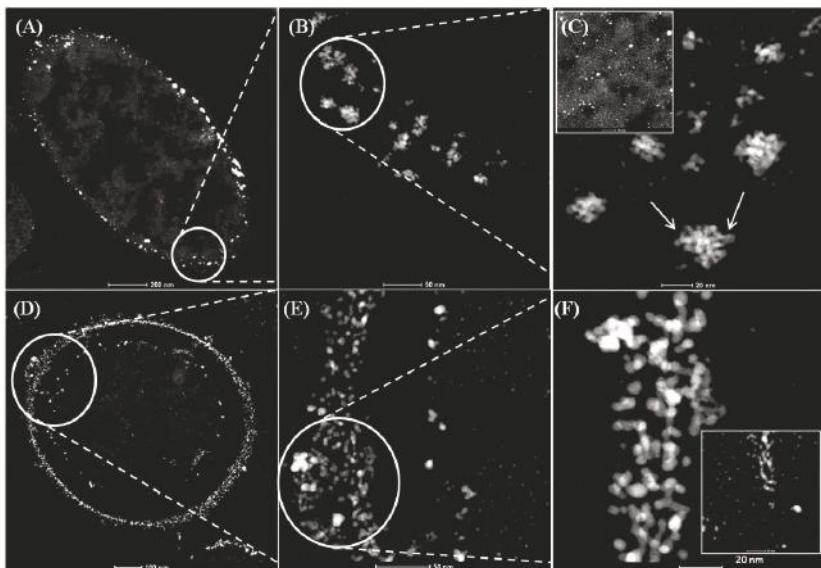


Figure 2. HAADF-STEM micrographs of Pd nanoparticles synthesized using 5 wt% Pd loading (1:20) on *D. desulfuricans* from 2 mM Na_2PdCl_4 solution, in 0.01 M HNO_3 using hydrogen as electron donor without MW treatment (A,B,C) and with 30 second MW treatment (D,E,F).

The distribution of PdNPs within the intracellular matrices, cell surface layers, and membrane was established by using elemental mapping (Figures 3 and 4). The main elements associated with Pd were phosphorus (P) and sulfur (S). Statistical analysis using ImageJ software [41] was done in order to

determine the Manders overlap coefficient to reveal any correlation in localization between Pd and S, and Pd and P in each strain and the effect of the MW treatment on any co-localizations. The Manders overlap coefficient was above 0.9 for both strains and treatments (Figures 3 and 4). According to the statistics analysis done using ImageJ, no differences in the degree of co-location for the selected elements for control and MW-treated cells were observed for either strain (Figures 3 and 4).

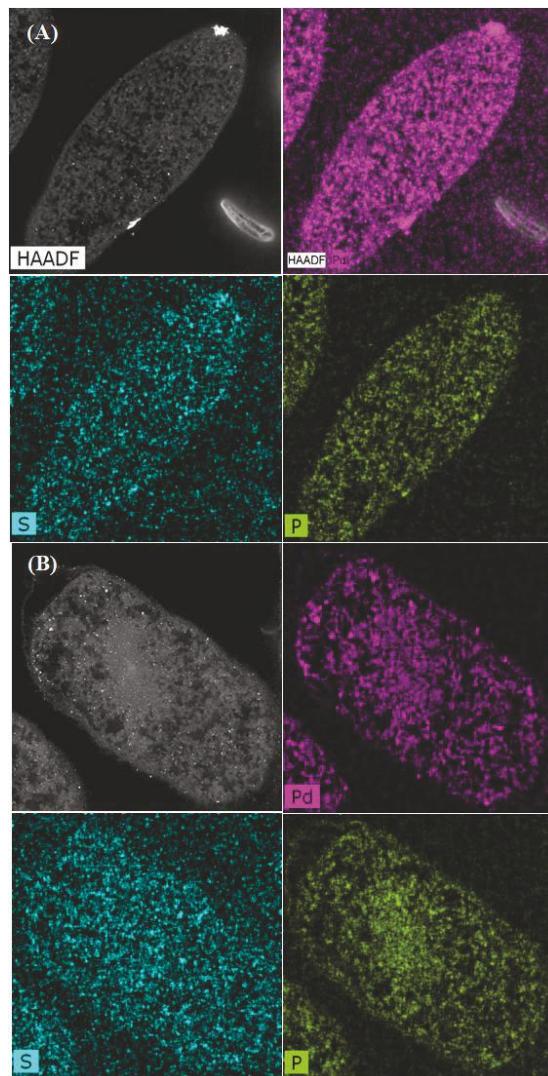


Figure 3. Elemental mapping showing distribution of Pd, P, and S in untreated cells of *E. coli* MC4100 (A) and cells treated with MW for 30 sec (B). The Manders overlap coefficients were higher than 0.9 for Pd/P and Pd/S in control and MW-treated cells.

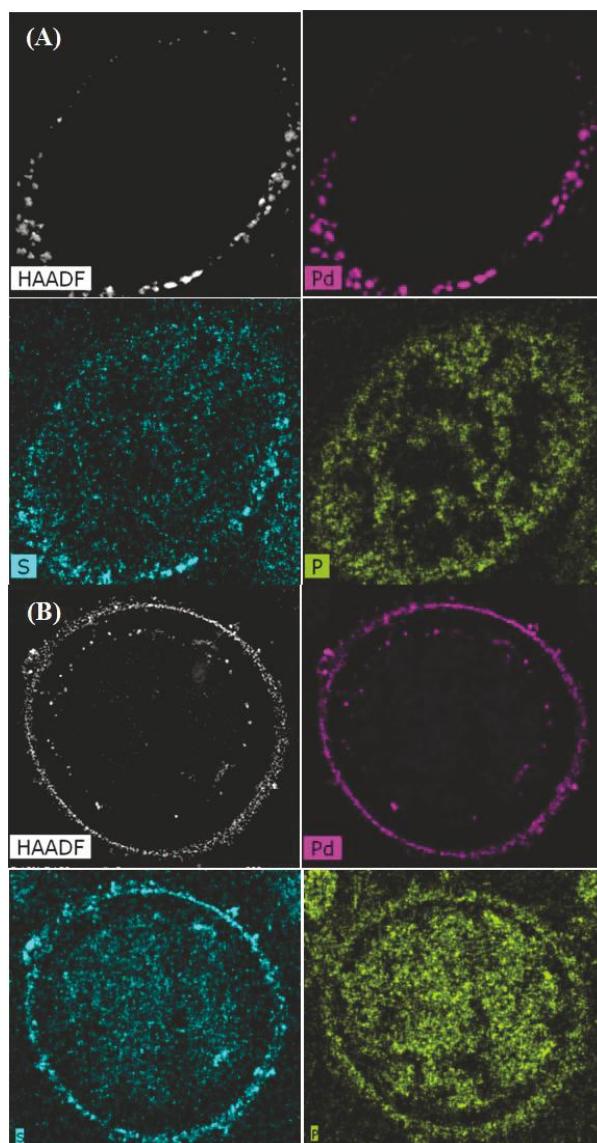


Figure 4. Elemental mapping showing distribution of Pd, P, and S in untreated cells of *D. desulfuricans* (A) and cells treated with MW for 30 sec (B). The Manders overlap coefficients were higher than 0.9 for Pd/P and Pd/S in control and MW-treated cells.

The co-location of Pd with S in *D. desulfuricans* is assumed to be PdS resulting from biogenic H₂S from residual metabolism [24] and the formation of PdS was confirmed in cell surface layers of sulfidogenic bacteria using X-ray photoelectron spectroscopy [45]. In addition, sulfur is present in the amino acids, cysteine and methionine (components of proteins), while phosphorus is ubiquitous within deoxyribonucleic acids, ribonucleic acids, phospholipids, etc., as well as phosphorylated proteins and ATP. The role of these biological components in the patterning of palladium deposition is under current consideration.

3.3. Dispersity and Size Distribution of Pd Nanoparticles

Analysis of the nanoparticles size distribution was performed using high-resolution images and ImageJ software [46]. Cells previously MW-treated and then exposed to Pd (II) followed by reduction to Pd (0) were analysed. The mean particle size of the intracellular Pd nanoparticles was 1.28 nm and 1.17 nm for the control and treated cells of *E. coli* MC4100, respectively (Figure 5A,B).

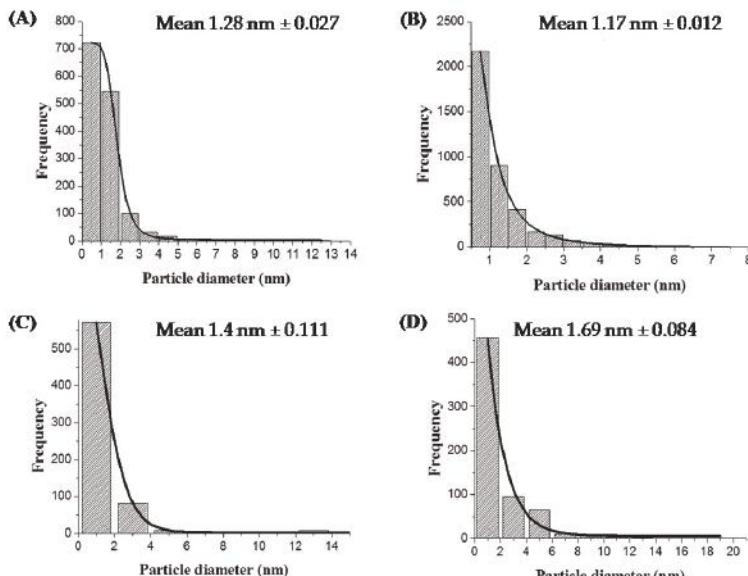


Figure 5. Size distribution and mean nanoparticle (NP) sizes of Pd nanoparticles made by *E. coli* MC4100 untreated (A) and MW-treated cells (B) and *D. desulfuricans* untreated (C) and MW-treated cells (D). The mean NP sizes (nm, mean \pm SEM) were (A): 1.28 ± 0.027 ; (B): 1.17 ± 0.012 ; (C): 1.40 ± 0.11 ; (D): 1.69 ± 0.084 .

For *D. desulfuricans* the major differences in terms of shape and size of the NPs were observed on the surface so the analyses were mainly focused on these areas. The corresponding mean particle sizes of membrane Pd NPs produced by treated cells was 1.69 nm and 1.4 nm for control cells (Figure 5C,D, respectively). These differences were significant at $P = 0.95$ (two sample test of the variance) and hence *E. coli* makes smaller intracellular NPs in response to MW irradiation whereas the NP size increases in the case of *D. desulfuricans* surface-located NPs. However since the differences were small (~10–20%), a mechanistic biological significance cannot be attributed to them at this stage and further work is required to reveal the underlying reasons which may be attributed to the mechanisms of NP deposition in the two strains (see Discussion).

Despite the small differences in mean sizes of the Pd nanoparticles, notable differences were found in their degree of dispersity. The polydispersity value of Pd nanoparticles produced by treated cells of *E. coli* was 0.55 as compared to 0.80 for untreated cells. In the case of *D. desulfuricans* the polydispersity value was 1.26 and 2.07 for MW-treated and untreated cells, respectively. The lower value of polydispersity indexes (32% and 39% lower than controls for *E. coli* and *D. desulfuricans*, respectively) for the two strains resulting from the MW injury suggests a higher degree of homogeneity of the size of nanoparticles. This is accordance with the visual appearance of the cells as noted above.

3.4. Crystallinity and Lattice Spacing of Pd Nanoparticles

No differences were observed in terms of crystallinity of the particles under these conditions (Figures 6 and 7). The most representative lattice spacing for *E. coli* was 0.204 nm and 0.213 nm consistent with the (200) facets, and 0.241 nm and 0.23 nm consistent with the (111) facets of Pd, showing a mixed-facet arrangement for both membrane and intracellular Pd nanoparticles (Figure 6A,B).

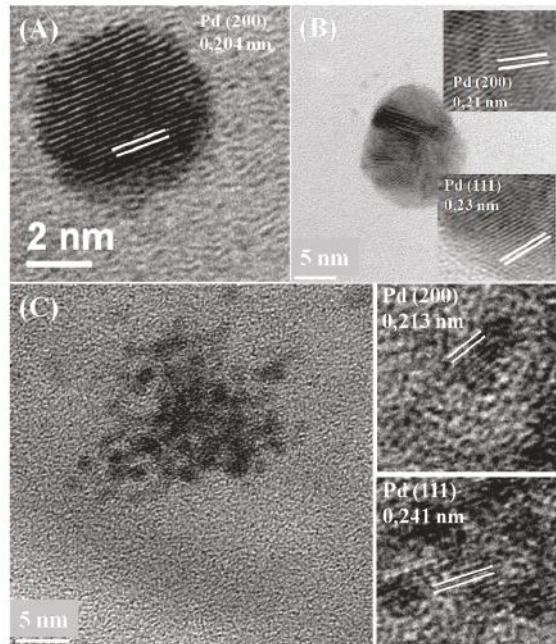


Figure 6. TEM images of intracellular Pd crystals made by *E. coli* MC41005% bio-Pd 30 seconds MW treatment before being exposed to Pd(II) (A,B) and untreated cells (C) revealing lattice spacing in crystals.

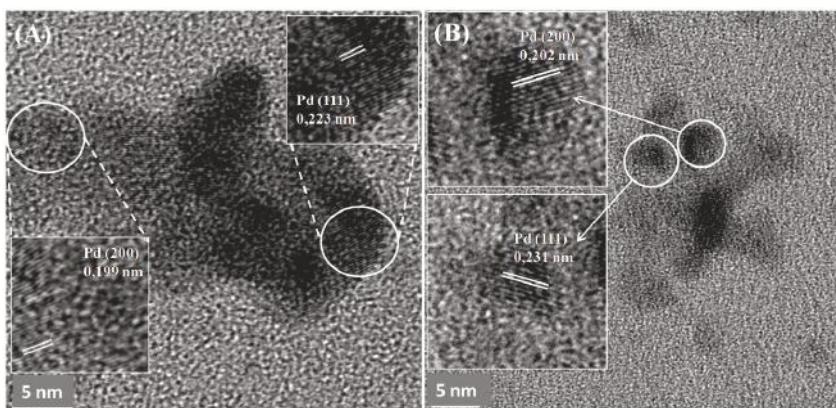


Figure 7. TEM images of membrane-bound Pd crystals made by *D. desulfuricans* 5% bio-Pd 30 seconds MW treatment before being exposed to Pd(II) (A) and untreated cells (B) revealing lattice spacing in crystals.

MW-treated cells of *D. desulfuricans* showed lattice spacing of 0.223 nm and 0.199 nm, again consistent with the (111) and (200) facets (Figure 7A) and, similarly, 0.231 and 0.202 nm for untreated cells. A similar NP structure was observed in previous studies by Omajali et al. [24] of Pd nanoparticles in *D. desulfuricans* with a mix-facet arrangement of (111) and (200) with different crystal orientations in the case of the larger clusters (Figure 7A) when made under H₂ as in the present study.

4. Discussion

This study focuses on the synthesis and characterization of palladium nanoparticles produced by two different bacterial strains that were previously injured via application of microwave energy compared to cells that had no MWexposure. The capacity of these two related genera for the synthesis of Pd NPs is well known via the activity of hydrogenases [9,19] as well as via other (unidentified) mechanisms in cells under conditions in which hydrogenases are not expressed [47]. The chemical reduction of Pd (II) under hydrogen or with formate as an electron donor using killed cells was almost negligible [9,19]. Few studies have been published using cells previously treated with MWenergy and none involve the synthesis of Pd NPs. Previous studies proposed that the application of radiofrequency (RF) energy in *E. coli* under similar conditions to those used in this work might cause electrokinetic modification of the cell surface with a destabilization of the cell membrane [34]. Another observation made by this group was that the application of MW energy resulted in disruption of the cellular membrane and, as a consequence, cytosolic fluids within the *E. coli* cells could pass out through the membrane. However, this was a temporary effect as the shape of the cells was restored within a 10 min recovery period. In the present study the Pd (II) was applied immediately after the removal of the cells from the MW apparatus within 1 min at the early stage of the 10 min “recovery window”. However, follow-up work showed that incorporating the Pd (II) at the outset, during MW exposure, gave similar results as with MW exposure during the pre-Pd (II) period [44] using the experimental setup described here as well as that of the earlier study (Supplementary information Figure S1).

In normal processes of formation of PdNPs on *E. coli*, a possible mechanism to explain the transport of Pd inside the cells was highlighted by Deplanche et al. [7], while a previous report [47] showed that Pd (II) is transported across the membrane through an unknown translocation mechanism. It is known that Ni (II) is a key component in many metalloenzymes [48,49] that are located in the cytoplasm (e.g., ureases and hydrogenases); the latter are also located in the cell membrane and, in the case of *D. desulfuricans*, in the periplasmic space. Deplanche et al. [7] suggested that, due to the similar chemistry of Ni (II)and Pd (II), the latter could be transported inside the cells through the Ni (II) “trafficking system” and deposited as NPs in the cytoplasm as a possible mechanism to forestall cytotoxicity if the Pd (II) is taken up in lieu of Ni (II), but cannot substitute for Ni (II) functionally.

Considering cells exposed to MW energy followed by exposure to Pd (II) solution, apart from the mechanisms mentioned above, additional responses could be activated in the cells that could influence the deposition of Pd inside the cells. According to Shamis et al. [34] cytosolic fluids would be extruded from the cells as a response to the MW radiation and the Pd (II) ions may have more sites becoming available for initial formation of NPs on re-absorption of extruded material along with Pd (II) ions. As a consequence, a higher number of initial binding sites would be occupied by Pd (II) due to the higher accessibility to the binding ligands originating from the cytoplasm caused by the MW. Deplanche et al. [7] confirmed a correlation between the initial uptake of Pd (II) onto cellular components and the initial formation of nuclei at many coordination sites, resulting in smaller Pd NPs per given biomass per constant amount of Pd. Once the effect of MW radiation is finished and cells are recovered in shape and membrane permeability (with re-absorption of the cytosolic fluids [34]) the resulting cells would have a greater proportion of Pd (II) ions contained intracellularly (as compared to surface-localized) than untreated cells. Once the intracellular reduction of Pd (II) into Pd (0) occurs, the association of the resulting Pd NPs with phosphate or sulfide structures would reduce NP mobility, reducing NP agglomeration. The combination of the translocation mechanisms

combined with mechanisms activated in the cells as a response to the MW energy may lead to the formation of Pd NPs with higher dispersity than native cells.

A recent study using *S. oneidensis* for the synthesis of Pd/Au nanoparticles showed how a post-treatment consisting of a doping process, carbonization of bacteria, and reduction of graphene oxide avoided NP agglomeration and, as a consequence, increased the dispersity of the nanoparticles, resulting in higher electrochemical activity than a commercial electrocatalyst [50]. Earlier work had shown co-formation of bio-Pd and reduced graphene oxide on cells of *E. coli* [51] but the catalytic activity of the *E. coli* material was not tested. In contrast, for use as a chemical catalyst the cells are not carbonized but are washed in acetone which destabilizes the membrane and lipid structures, making the intracellular Pd NPs more accessible for the reaction. Recent work showed biogenic palladium catalyst of *D. desulfuricans* cells that had been exposed to MW energy had a higher hydrogenation activity and product selectivity in the hydrogenation of 2-pentyne [44], which will be reported as a companion to the present study. Further, follow-up work showed cells of *E. coli* MW-injured under the conditions described here had a similarly increased activity as a selective oxidation catalyst when developed as Pd/Au core-shell NP structures.

By the altered deposition of the NPs observed in the biogenic Pd NPs synthesized by MW-treated cells of *D. desulfuricans* in the present work, a causal relationship may be suggested although conclusive proof is awaited. Since the deposition of the derived Pd NPs of untreated cells differs between *E. coli* and *D. desulfuricans* (Figures 1A and 2A) it may be suggested that both types of cell may have different mechanisms relevant to the synthesis of NPs in terms of shape and location via cellular responses during exposure to MW energy and during the recovery period. The process of the synthesis of Pd NPs in *Desulfovibrio* strains has been studied previously [9,24]. With respect to the derived Pd NPs synthesized by *D. desulfuricans*, with cells exposed to MW energy, the main differences were observed in the periplasm where the hydrogenases are predominantly located [52,53]. A higher porosity of the membrane may facilitate the deposition of the Pd (II) in many different nucleation sites for the initial seeds that without the increased porosity of the membrane caused by the MW energy would not be possible. This hypothesis could explain why the polydispersity index decreases when the cells were previously exposed to MW energy, indicating a higher number of NPs with homogeneous size, given the same dose of Pd (II). This effect was seen regardless of the bacterial strain used.

The relationship between the catalytic activity and the size of the NPs in these two strains has previously been shown [10]. A related study [44] to evaluate the effect of the MW treatment of resting cells (by the method shown in Supplementary information Figure S1) showed the initial rate of conversion of 2-pentyne to be increased by 50% (from 1.1 μmol per litre per second) by Pd NPs made following MW injury of the cells, with selectivity to the desired product cis-2-pentyne, being approximately doubled. One of the major goals in the optimization of the catalytic reactions using NPs as a catalyst is the NPs size control that will help to further understand the relationship between the size and location of the Pd NPs and the product selectivity for a specific reaction. A good example of size control and good dispersion of Au and Pd NPs is the S-layer protein of the Gram-positive strain *Lysinibacillus sphaericus* JG-A12 [54,55]. The monomer of the S-layer offers a good biotemplate for the formation of NPs with a regular structure, pores with identical size (1–5 nm), and good binding sites for Pd(II) such as glutamic and aspartic groups. The crystal structure of the Pd NPs synthesized by the MW-treated cells in the present study did not differ from the untreated cells, showing similar results for *D. desulfuricans* as those obtained by Omajali et al. [24] where twinned NPs were also seen (Figures 7 and 2b). Future studies will focus on comparative studies of the catalytic activity of the Pd NPs synthesized by MW-treated cells versus untreated cells that will inform the controlled synthesis of bio-Pd NPs with higher catalytic activity.

5. Conclusions

This study shows the application of microwave radiation on Gram-negative cells of *D. desulfuricans* and *E. coli* prior to the exposure of Pd (II) solution for the synthesis of PdNPs with a higher degree of

dispersity compared to cells that had no MW exposure. The response to the MW on the synthesis of the PdNPs is strain specific. The main differences of the NPs made by treated cells of *E. coli* were at the level of the cytoplasm with an increase of 32% approximately in the level of homogeneity compared to untreated cells. However, the main differences in treated cells of *D. desulfuricans* were observed at the level of the surface with an increase of 39% in the level of homogeneity of the size of the nanoparticles compared to the control.

Supplementary Materials: The following are available online at <http://www.mdpi.com/2079-4991/9/6/857/s1>, Figure S1: Effect of RF irradiation (microwaves) on cells of *E. coli* and *D. desulfuricans* in the absence of microwave treatment (A,B, respectively) and following exposure to MW in short bursts with cooling (C, D, respectively) (this study; see text). A companion study was done (E,F,G) which applied RF radiation at 2W (E),4W (F) or 8W (G) (20 min) using purpose-built equipment to decouple the thermal and electromagnetic components. The equipment was developed in house at C-Tech Innovation Ltd. via adaptation of commercial equipment. The *E* (electric) component of electromagnetic radiation produces heating in dielectric (non-conducting) materials so it is important to exclude heating from the sample by isolating the *H* (magnetic) component. This was achieved using a series LC circuit tuned with an external matching box. This generated a magnetic field inside the induction coil and an electric field in the air gap between the two conducting plates. The sample vial, placed in the centre of the induction coil, was exposed to a magnetic field for the desired time. After exposure, the final temperature of the bottle was checked and absence of heating confirmed. The samples were treated by exposure inside a solenoid coil as the electromagnetic field within the coil will be almost entirely magnetic; consequently, as the magnetic susceptibilities of the components of the sample being treated (glass, water, bacteria) are very low at this frequency there would be minimal heating of the sample, allowing long treatment times. The total dose in the two studies was similar. The commercial equipment (this study) delivered 300 W which was applied for 30 sec (9000 'units of stress'). In the purpose-built equipment the cells were treated with 2–8 W over 20 min (up to 9600 'units of stress'); the power level can be considered as a rate (power = energy per second; 8 W; 20 × 60 sec) and a dose as rate × time of the integral of rate over time. Figure S2: HAADF–STEM–EDX analysis of Pd nanoparticles (circled areas shown) with phosphorus (P) and sulfur (S) of untreated cells of *E. coli* MC4100 (A) and *D. desulfuricans* (C), 30 second MW treated cells of *E. coli* MC4100 (B) and *D. desulfuricans* (D). Silicon is from the oil diffusion pump of the column of the TEM system, copper is from the TEM grid and osmium from the staining.

Author Contributions: Biomaterials were made and characterized by J.G.B. and I.P.M.; high-resolution SEM/TEM/elemental mapping were done by J.G.B. and M.L.M.; the paper was lead-authored by J.G.B. and work was supervised by L.E.M. and M.L.M.

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Article

“Chocolate” Gold Nanoparticles—One Pot Synthesis and Biocompatibility

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Abstract: The chemical synthesis of nanoparticles can involve and generate toxic materials. Here, we present for the first time, a one pot direct route to synthesize gold nanoparticles (AuNPs) using natural cacao extract as both a reducing and stabilizing agent. The nanoparticles were characterized by UV-visible spectroscopy (UV-VIS), dynamic light scattering (DLS), and transmission electron microscopy (TEM); and have excellent biocompatibility with human primary dermal fibroblasts.

Keywords: green synthesis; cacao; non-cytotoxic

1. Introduction

For decades, nanoparticles (NP) of noble metals such as gold, silver, and platinum have captivated the researchers and the general public with their remarkable physical and chemical properties, as well as for their potent therapeutic power [1]. Gold is one of the least reactive among the noble metals [2], but its nanoparticulate forms possess unique chemical, electrical, and optical properties [3]. These properties, which are size and shape dependent, can be tuned by a variety of means, such as the synthesis route, reactants, and experimental conditions [4]. Numerous applications have benefited from the special properties of gold nanoparticles, including optics, imaging, sensing, catalysis [5,6] and biomedicine [7,8] (in particular dentistry, cancer diagnostics, and photothermal and photodynamic therapies) [9–11].

As a result of the growing interest in gold nanoparticles, numerous chemical and physical synthesis routes have been proposed during the last decades [12,13]. However, major drawbacks with some of the conventional synthesis methods include toxic, hazardous chemicals and challenging reaction parameters [14–16]. Modern synthetic trends are shifting to alternative synthetic routes to minimize the use of harmful chemicals. Several studies have reported the benefits of the biosynthesis approaches using plant extracts, and unicellular and multicellular organisms [14,17]. Although chemically complex, phytochemicals have major advantages over other biosynthesis methods, as they are generally non-toxic to mammalian cell types and to the environment [14,18]. The use of plant derivatives also reduces the possibility for the absorption of toxic chemicals on the surface of the AuNPs [19]. Various studies have demonstrated the power of phytochemicals in gold nanoparticles (AuNPs) synthesis as well as the biocompatibility of the generated AuNPs to different cell types [14,17].

Here, we report for the first time, on the potential of cacao extract as a reducing and stabilizing agent in the synthesis of AuNPs. In addition to being a popular constituent in various foods and beverages, cacao has been speculated to alleviate health disparities such as aging, inflammation, depression, cancer, and stress [20–24]. The hypothesis behind this work is that oxalic acid, which is a constituent of cacao, can reduce Au^{3+} in HAuCl_4 to metallic gold and stabilize the resultant nanoparticle colloidal solution. This hypotheses is also further substantiated by our previous work, where we reported the synthesis of silver nanoparticles (AgNPs) facilitated by cacao extract [25]. Herein, we extend this synthesis approach to prepare biocompatible ‘green’ gold nanoparticles and explore their properties. This easy single-step synthesis route was optimized and the prepared samples were characterized using UV-visible spectroscopy (UV-VIS), dynamic light scattering (DLS), and transmission electron microscopy (TEM). Finally, primary human dermal fibroblast (HDFs) cells were used to evaluate the biocompatibility of the gold nanoparticles.

2. Materials and Methods

2.1. Reagents and Chemicals

Hydrogen tetrachloroaurate (HAuCl_4), penicillin, streptomycin, and L-glutamine were bought from Pro Sci Tech, Kirwan, Australia. Cold pressed cacao powder was obtained from Forest Super Foods, Melbourne, Australia, and stored in an air-tight container (Goodguys, Adelaide, Australia). NaOH pellets, phosphate buffer saline (PBS) tablets, foetal bovine serum (FBS), nitric acid (70%), and Dulbecco’s Modified Eagle Medium (DMEM) were purchased from Sigma-Aldrich, Sydney, Australia. Hydrochloric acid (36%) was procured from Ajax Finechem Pty. Ltd., Sydney, Australia. All of the reagents were used as received. Ultra-pure MilliQ water (resistivity 18.2 Ω , Sigma-Aldrich, Sydney, Australia) was used for all of the experimental and cleaning procedures.

2.2. Synthesis of Gold Nanoparticles

The aqueous extracts of cacao were prepared by mixing a varying amount of cacao powder (Table 1) in 10 mL of ultrapure water (MilliQ system, Millipore Corp., Burlington, MA, USA) at room temperature. The extract obtained after the filtration (0.45 μm —sterile EO, Sartorius Stedim Australia Pty. Ltd., Dandenong South, Australia) of the suspension was stored for the synthesis of AuNPs. Then, the cacao extracts were mixed with aqueous solution of gold chloride (0.1 mg/mL in MilliQ water) (Table 1). The reaction mixtures were stirred continuously for 30 min at 100 °C under reflux. After 30 min, the heating source was removed, the reaction mixtures were cooled down to room temperature (25 °C), and stirring continued for 24 h. The periodic (30 min, 1 h, 2 h, 3 h, 4 h, and 24 h) monitoring of the prepared AuNPs was carried out using a UV-VIS spectrophotometer. The samples, S1, S2, S3, S4, and S5, refer to the AuNPs suspensions synthesized with 0.1, 1, 2.5, 10, and 50 mg/mL of cacao extract, respectively. The pH of the nanoparticle solutions was six.

Table 1. Concentrations of reactants used for the synthesis of different cacao-gold nanoparticles (AuNPs).

Sample	Gold Chloride (mg/mL)	Cacao (mg/mL)
S1	0.1	0.5
S2	0.1	1
S3	0.1	2.5
S4	0.1	10
S5	0.1	50

2.3. Characterization

The progress of the reaction was periodically monitored using a Cary 5 UV-VIS spectrophotometer (Varian Australia Pty. Ltd., Mulgrave, Australia) at room temperature in the wavelength range of 400–800 nm. All of AuNPs’ suspensions were diluted 2X (*v/v*) with MilliQ water prior to UV-VIS

spectral characterization. MilliQ water was used as a blank throughout the experiment. Quartz cuvettes were used for all of the measurements.

All of the samples were diluted to a suitable concentration using MilliQ water prior to DLS analysis to determine the hydrodynamic diameter of the nanoparticles. A Nicomp 380 particle size analyzer (Nicomp Particle Sizing Systems, Port Richey, FL, USA) operating at 25 °C was used for all of the DLS and zeta potential measurements. The mean hydrodynamic diameters reported are the average of the three measurements taken of the three independent nanoparticle batches (separate syntheses). Disposable plastic cuvettes were used for all of the measurements. All of the analyses were carried out at pH-6.

A 'JEOL 2100F' (Tokyo, Japan) transmission electron microscope (TEM), operated at an acceleration voltage of 200 kV, was used to determine the size and morphology of the synthesized cacao-AuNPs. Samples for the TEM analysis were prepared by depositing a small volume (10 µL) of the AuNPs solution on a carbon coated copper grid (ProSciTech, Kirwan, Australia). The grid was left to dry overnight at room temperature prior to TEM analysis. The crystal structure of the AuNPs were determined with the selected area electron diffraction (SAED) pattern obtained from TEM images.

2.4. Fibroblasts Study

Primary derived HDFs were gifted from Dr. Louise Smith, the University of South Australia. The HDFs were harvested and grown, as described elsewhere [26]. Briefly, cells were grown from frozen stocks and maintained in DMEM at 37 °C in 95% humidity and 5% CO₂. The DMEM was changed every 3–4 days. Ethics approval was approved by the Ethical Committee at the Queen Elizabeth Hospital and the University of South Australia Human Ethics Committee, described elsewhere [26].

The viability of the cacao-AuNPs treated primary human dermal fibroblast (HDFs) cells was tested using a resazurin assay based on the reduction of non-fluorescent resazurin by metabolically active living cells to form fluorescent resorufin which was quantified using a microplate reader. Cells (1×10^4 cells per well in DMEM) were seeded in 24 well plates on air plasma cleaned (5 min, 40 W, 2×10^{-1} mbar) thermanox coverslips and incubated (24 h in 95% air, 5% CO₂ at 37 °C) until they reached 50 and 80% confluence. The DMEM was supplemented with FBS, penicillin (100 IU), and streptomycin (100 lg) (Invitrogen). After incubation, the DMEM was removed and the cells (50 and 80% confluent) were briefly washed with PBS. The cells were then treated with different concentrations of AuNPs (containing 500, 250, and 125 µg/mL of Au) prepared in warm DMEM and incubated for 24 and 72 h. After each incubation period, the media was aspirated again and the cells were rinsed with PBS. A stock solution of 110 mg/mL resazurin was prepared in phosphate buffered saline and filter sterilized using a 0.2 mm filter. The stock was then diluted 1:10 in fresh warm DMEM and 600 µL of the diluted solution was added to each well. After 1 h, 200 µL of the reduced solution was transferred into a 96 well plate and the fluorescent intensity was recorded using a plate reader ($\lambda_{\text{ex}} = 544$ nm and $\lambda_{\text{em}} = 590$ nm). Fresh DMEM without any AuNPs served as the control. The media from the wells were replaced with fresh DMEM every 2–3 days during the course of the assay. The percentage of cell viability was calculated with the following equation.

$$\text{Viability (\%)} = 100 \times \text{Absorption}_{\text{test}} / \text{Absorption}_{\text{control}}$$

2.5. Statistical Analysis

All of the statistical analyses were performed using graph pad prism 6 software. All of the data were expressed as mean ± standard error mean (SEM). Statistical significance was determined using one-way ANOVA with a Dunnett's post-test. All of the experiments were performed in biological and technical triplicates on three separate days.

3. Results and Discussion

3.1. Synthesis and Characterization of AuNPs

The procedure for synthesizing gold nanoparticles with cacao extract is outlined in Figure 1. Briefly, under rigorous stirring, cacao extract was mixed with an aqueous solution of gold chloride (HAuCl_4) at 100 °C (boiling temperature). Upon the addition of the reactants, the solution became increasingly darker and changed from transparent light yellow to purple-red within 5 min. The color change was a visual indication of the formation of AuNPs (Figure 2A–E insets). The samples, S1, S2, S3, S4, and S5, refer to the AuNPs suspensions synthesized with increasing the concentration of the cacao extract by 0.5, 1, 2.5, 10, and 50 mg/mL, respectively (see Table 1).

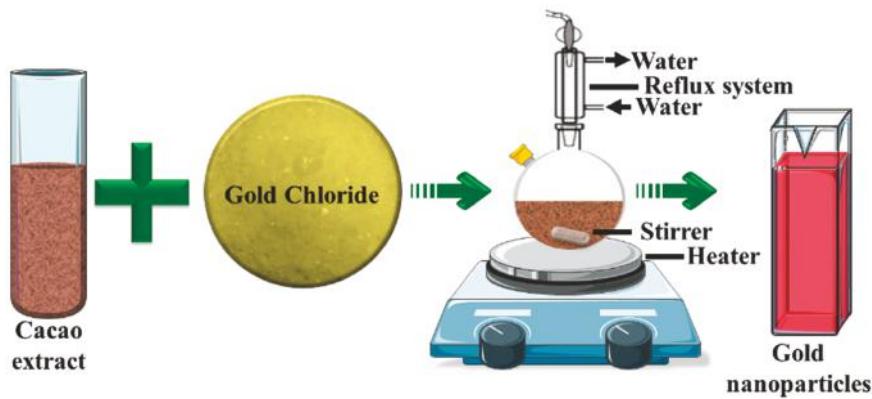


Figure 1. Schematic illustration of gold nanoparticles (AuNPs) synthesis from cacao extract.

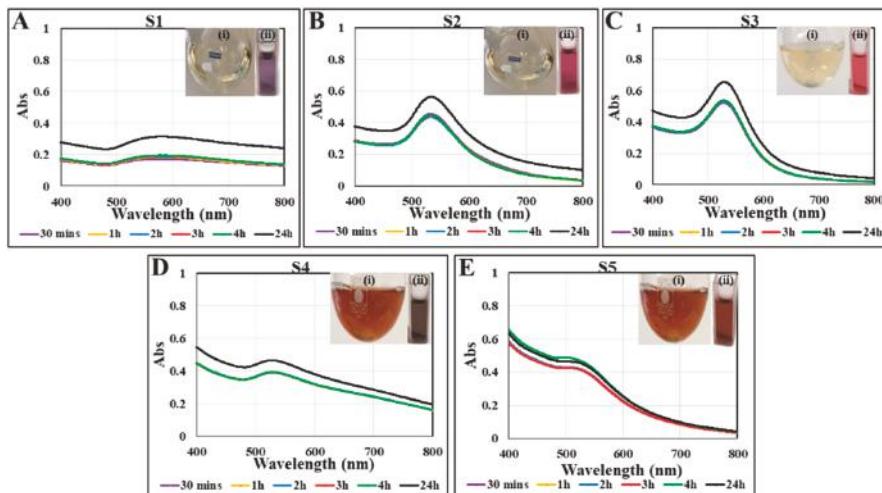
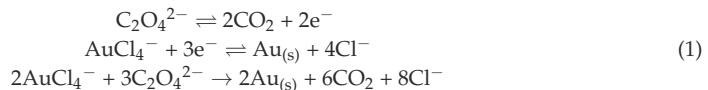


Figure 2. UV-Vis spectra of gold nanoparticles S1, (A); S2 (B); S3 (C); S4 (D); and S5 (E) obtained using cacao extract. Absorption spectra of samples recorded at 30 min, 1, 2, 3, 4, and 24 h from the initiation of AuNPs synthesis at 100 °C. Insets displaying the AuNPs solution before (i) and after (ii) synthesis.

The intense purple-red color of the reaction mixture is due to the well-known phenomenon of plasmon resonance (PR), which is the result of resonant oscillations of the semiconfined electrons in the nanoparticles with the incident photons [27]. Several distinct parameters, including amount, size, and shape of the nanoparticles, interparticle electronic interactions, and the surrounding media have an influence on the position, shape, and intensity of the PR band [28,29]. UV-visible spectrophotometry was employed to evaluate the time course of the reaction kinetics by taking measurements at intervals of 30 min, 1, 2, 3, 4, and 24 h. UV-visible profiles of the reaction mixture at these time points are shown in Figure 2. The spectra of all of these samples exhibited a maximum of absorption (λ_{max}) at around 530 nm, consistent with the plasmon resonance absorption band of AuNPs at ~510–560 nm [30], which confirmed the formation of gold nanoparticles. As an overall trend, the reaction had a fast-initial phase, much of the reduction being completed within 30 min. Clearly, there is no significant change in the absorption intensity within time points of 30 min, 1, 2, 3, and 4 h. However, a noticeable increase was observed after 24 h, which demonstrates that the reaction was slowly continuing, but was completed within this period as there was no further changes in the adsorption intensity beyond 24 h.

The UV-visible spectra pointed to some interesting trends. Initially, when the cacao concentration increased to 2.5 mg/mL (S3) (Figure 2A–C), the intensity of the plasmon resonance absorption increased, which suggests the formation of more gold nanoparticles. In the same time, the maximum of the peak shifted to the left, indicating a decrease in nanoparticles size. A further increase in the cacao concentration to 10 and 50 mg/mL (S4 and S5, Figure 2D,E) led to a decrease in the plasmon resonance absorption and a broadening of the spectra, pointing to aggregations in the system.

Oxalic acid in cacao is a natural reducing agent, which we showed to reduce silver ions into nanoparticles [25,31]. Oxalic acid exists as oxalate ions ($\text{C}_2\text{O}_4^{2-}$) in the experimental conditions ($\text{pH} > 4.3$) used for this study. The standard potential of the $\text{C}_2\text{O}_4^{2-}/\text{CO}_2$ oxido-reduction couple is $E^0_{\text{red}} = -0.49$ V, while the one of $\text{AuCl}_4^-/\text{Au(s)}$ is $E^0_{\text{red}} = 1.002$ V. The oxido-reduction reaction scheme resulting from the reduction of gold ions by oxalate is as follows:



The synthesized AuNPs were characterized for their hydrodynamic diameter and zeta potential. The results are summarized in Table 2. The hydrodynamic diameter (as determined by the DLS) of the samples S1, S2, and S3 was 54, 29, and 18 nm, respectively. These results are in good agreement with the UV-vis absorption spectra, where the PR maximum shifted to shorter wavelengths, indicating a smaller particle size. The hydrodynamic diameter of samples S4 and S5 could not be reliably determined because of aggregates, also suggested by the UV-vis spectra.

The zeta potential (ζ) provides important cues about the stabilization mechanisms in the colloidal suspension of nanoparticles. The zeta potential (Table 2) of samples S1, S2, and S3 was between -11 mV to -17 mV, and this range is known to confer incipient stability for colloids [32]. The negative charge on the surface of the nanoparticles appears to play a significant role by ensuring repulsion between the particles in the suspension. The samples were very stable and even after a month, no visible particle agglomeration was observed.

Table 2. Hydrodynamic diameter and zeta potential of green AuNPs synthesized using cacao extract.

Sample	Hydrodynamic Size (nm)	Zeta Potential (mV)
S1	54.4 ± 9.1	-11.65
S2	28.7 ± 3.4	-14.10
S3	17.9 ± 1.5	-17.52

The TEM images in Figure 3 shows that the morphology of the as-synthesized AuNPs was mostly spherical, and there was no particle aggregation when the cacao concentration was below 10 mg/mL (Figure 3A–C). The particle size analysis of samples S1, S2, S3, and S4 are shown in Figure 3E–H, respectively. The AuNPs had an average particle size of 35 ± 10 nm (S1), 20 ± 9.1 nm (S2), 10 ± 11.6 nm (S3), and 7 ± 4.2 nm (S4). The particle sizes of samples S1, S2, and S3, determined from the TEM images, are smaller than the hydrodynamic diameter measured by DLS. These variations in the particles sizes are the result of the different measurement principles used by these two methods [33]. The size distribution analysis of sample S5 could not be performed reliably because of the presence of particle aggregation. However, both methods suggest an overall trend of particle size reduction when increasing the concentration of cacao was observed.

The selected area electron diffraction (SAED) analysis confirmed that the synthesized AuNPs are crystalline in nature (Figure 3I–K). As a result of random orientation of crystal planes, concentric diffraction rings were observed in the SAED patterns of the samples S1, S2, and S3, and the reflection rings were indexed to (111), (200), and (220) planes of the face centered cubic (fcc) crystalline lattice of gold.

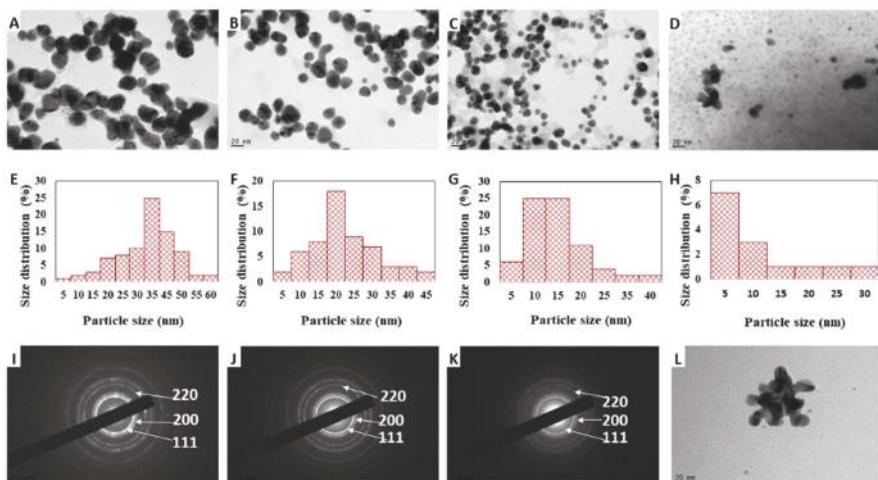


Figure 3. TEM images of the cacao-AuNPs S1 (A), S2 (B), S3 (C), S4 (D), and S5 (L). The corresponding histograms showing the particle size distribution (E, F, G, and H for S1, S2, S3, and S4, respectively). Representative SAED patterns of S1 (I), S2 (J), and S3 (K).

3.2. Viability of Human Dermal Fibroblasts after Exposure to AuNPs

Gold nanoparticles have found numerous applications in advanced medical therapies ranging from the sensing to treatment of cancers. It is thus important to evaluate the cytotoxicity of this new nanomaterial to human cells. Human dermal fibroblasts were selected for this experiment because these cells play important role in connective tissue. We tested AuNPs resulting from S1, S2, and S3 only since these preparations were free of aggregations, which is important for potential applications.

Cells having two different levels of confluence (50% and 80%) were assessed for their viability after 24 and 72 h of exposure to concentrations of cacao-AuNPs that contained 125, 250, and 500 $\mu\text{g}/\text{mL}$ of Au. Representative microphotographs of the untreated and HDFs treated with AuNPs are shown in Figure S1. After 24 h of exposure, 50% confluent HDFs showed no morphological changes in any treatment groups compared with the control (Figure S1A). At a treatment time of 72 h (Figure S1B), the cells morphology is similar to the control for all three of the AuNPs exposure concentrations. No cell shrinkage or floating cells were observed in the AuNPs treated HDFs at both time points. For the 80%

confluent cells, at 24 h of treatment, the cells are well spread out (Figure S1C) and exhibited a typical fibroblast morphology. When the treatment time was increased to 72 h, the area of cell spreading decreased when the HDFs were exposed to sample S3 (Figure S1D). As all of the AuNPs samples contained the same amount of gold, this effect could be due to the smaller size of the AuNPs in S3.

The viability of the HDFs in the culture conditions was determined using the resazurin assay. Figure 4A,B represents the influence of AuNPs on a 50% confluent HDFs after 24 and 72 h of treatment. The cells showed a greater than 90% viability for all of the AuNPs-HDFs treatments compared to the control. Interestingly, at both time points, the number of viable HDFs treated with S3 were significantly higher than the control, suggesting that S3 may have contributed in the proliferation of the HDFs as well.

When 80% confluent HDFs were treated with AuNPs for 24 and 72 h, the results were different (Figure 4C,D respectively). At 24 h, there was a significant increase in the viable cell number for S2 and S3 at all of the tested concentrations of Au, but S1 remained non-significant compared to the control. However, when the exposure time was increased to 72 h, a non-significant reduction in the cell viability was observed for S3. The degrees of freedom (DF) and probability (P) values determined from the viability assay for samples S1, S2, and S3 are provided in the Supporting Information (Table S1).

It is important to note that a variable degree of cytotoxicity against a wide range of cells has been reported for gold nanoparticles [34]. However, our data indicate that none of the cacao extract derived AuNPs samples caused any acute toxicity to HDFs. In general, cacao and its phytochemical constituents are known to be beneficial for humans [22] and to promote wound healing [21,35]. In this respect, HDFs are extremely important for controlling the wound healing process [36]. The fact that there was not an adverse cytotoxic effect observed on these cells indicates that the new cacao-AuNPs have a good biocompatibility and may be useful in the field of biomedicine. The synthesized AuNPs also have potentials as nano drug carriers. The negative surface charge and the carboxyl acid groups of the oxalic acid in these AuNPs can be used to bind and deliver other antibiotics or medically relevant drugs [37].

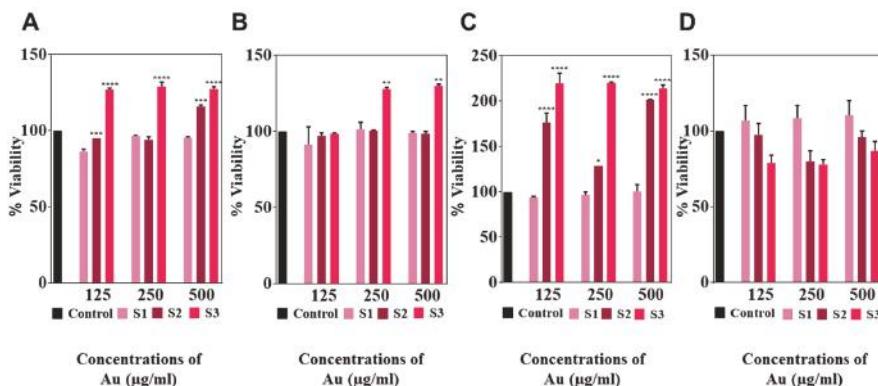


Figure 4. Determination of AuNPs generated cytotoxicity towards primary human dermal fibroblasts (HDFs). Cells were treated with different concentrations of AuNPs for 24 and 72 h. Cell viability of 50% confluent HDFs after 24 (A) and 72 h (B) exposure; and 80% confluent cells after 24 (C) and 72 h (D). The untreated cells served as controls. The results are represented as \pm standard error mean (SEM) ($n = 3$). * $p < 0.05$, ** $p < 0.01$, *** $p < 0.001$, **** $p < 0.0001$. Asterisks indicate statistical significance compared to the control.

4. Conclusions

Collectively, we developed a fast, single-step, and reproducible method for the synthesis of gold nanoparticles using the extract of cacao as a reducing and stabilizing agent. The resultant AuNPs

were mostly spherical, had a crystalline structure, and were negatively charged. We determined the experimental conditions that lead to stable colloidal suspensions, which are important for future applications. Furthermore, the size of the nanoparticles could be tuned by adjusting the concentration of the reactants. In vitro studies suggested that the cacao derived AuNPs are biocompatible, as none of the tested formulations exhibited cytotoxicity towards 50% and 80% confluent HDFs. This is important as gold nanoparticles have gained significant attention for application in fields of medical diagnostics and therapies. The toxic chemical free method for gold nanoparticles preparation developed in this work presents also opportunities in other fields, such as sensing. The surface of the nanoparticles can potentially be functionalised with desired ligands, which will provide opportunities for surface immobilization to surfaces for various applications. Another possibility, reinforced by the tunability of nanoparticles sizes, would be attachment of drugs and biomolecules to provide vehicles for delivery of cargo inside biological cells. Overall, this exciting, simple, green, and single-step new procedure for AuNPs preparation provides endless opportunities in numerous fields of research and practical application.

Supplementary Materials: The following are available online at <http://www.mdpi.com/2079-4991/8/7/496/s1>, Figure S1: Optical micrographs of HDFs cells exposed to different concentrations of cacao-AuNPs and the untreated controls. Panel I, panel II, and panel III denote the Au concentrations of 125, 250, and 500 µg/mL. (A) and (B) represent 50% confluent cells exposed to S1, S2, and S3 for 24 h and 72 h, respectively. HDFs of 80% confluency were incubated with AuNPs for 24 h (C) and 72 h (D).

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Article

Green Synthesis of High Temperature Stable Anatase Titanium Dioxide Nanoparticles Using Gum Kondagogu: Characterization and Solar Driven Photocatalytic Degradation of Organic Dye

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Abstract: The present study reports a green and sustainable method for the synthesis of titanium dioxide (TiO_2) nanoparticles (NPs) from titanium oxysulfate solution using Kondagogu gum (*Cochlospermum gossypium*), a carbohydrate polymer, as the NPs formation agent. The synthesized TiO_2 NPs were categorized by techniques such as X-Ray Diffraction (XRD), Fourier transform infrared (FTIR) spectroscopy analysis, Raman spectroscopy, scanning electron microscope- Energy-dispersive X-ray spectroscopy (SEM-EDX), Transmission electron microscopy (TEM), High-resolution transmission electron microscopy (HR-TEM), UV-visible spectroscopy, Brunauer-Emmett-Teller (BET) surface area and particle size analysis. Additionally, the photocatalytic actions of TiO_2 NPs were assessed with regard to their ability to degrade an organic dye (methylene blue) from aqueous solution in the presence of solar light. Various parameters affecting the photocatalytic activity of the TiO_2 NPs were examined, including catalyst loading, reaction time, pH value and calcination temperature of the aforementioned particles. This green synthesis method involving TiO_2 NPs explores the advantages of inexpensive and non-toxic precursors, the TiO_2 NPs themselves exhibiting excellent photocatalytic activity against dye molecules.

Keywords: titanium dioxide nanoparticles; green synthesis; gum kondagogu; methylene blue; photocatalysis

1. Introduction

Nanoparticles (metal and metal oxides) of various types have been widely employed via physical and chemical methods. Although these systems have resulted in the formation of numerous extremely diverse nanostructures, many environmental toxicity issues have emerged [1]. Metal oxides have recently been widely explored because of the huge variety of structural, material and functional properties exhibited by their nanoparticles. Transition metal oxide nanoparticles have generated great scientific interest owing to their unusual properties compared with their corresponding bulk

metals. Moreover, they have numerous industrial applications [2]. Metal oxide nanoparticles with disparate morphologies and sizes have been synthesized using different synthetic routes. These include hydrothermal, solvothermal, microwave, vapour deposition, spray pyrolysis and wet-chemical methods [3–5]. However, the usage of solvents in the chemical synthesis route adopted pose limitations for the application of NPs in medicine, pharmacy and other areas, primarily due to the toxicity caused by the solvents. Thus, there is an urgent need for the development of alternative, novel nanoparticle synthesis processes that could be exploited (at both the industrial and commercial level) in order to introduce cleaner, safer and smarter products suitable for application in communication technologies, medicine, pharmacy, agriculture and other industries. The introduction of environmentally benign approaches for designing NPs provides solutions to mounting tasks related to ecological concerns. Use of greener, safe and environmentally non-hazardous chemicals and green protocols for the synthesis of nanoparticles decreases the expense of synthesis while minimizing the utilization of harmful substances and their subsequent disposal [6].

Titanium dioxide (TiO_2) is regarded as an extremely promising metal oxide that can perform a key role in solving the global energy crunch, as well as serving to assuage environmental concerns [7,8]. Nanoparticles of TiO_2 have featured in numerous photovoltaic applications and in various sectors including cosmetic, pharmaceutical and skin care products. These versatile NPs can also confer whiteness to paints, plastics, papers, inks, food colorants and toothpaste while also finding usage in cancer treatment [9,10]. The synthesis of TiO_2 NPs, which possess controlled crystal phases and morphologies that make them eminently suitable in diverse applications requiring high performance, has proven to be fundamentally challenging for the scientific community [11].

Carbohydrate polymers isolated from trees [12], are an innovative type of potentially cost-effective and biologically favorable biomaterials that display highly particular and selective characteristics towards the design of nanostructures for unique applications. They are natural biopolymers based on plant exudates and possess interesting properties e.g., reducing, stabilizing, suspending, gelling etc. [13–15]. Tree exudates (gums) serve as ‘green’ media and harbor numerous hydroxyl, carbonyl and carboxylic functional groups. Such functional groups can act as good chemical reductants and their presence in these gum hydrocolloid materials facilitates the formation of metal or metal oxide NPs. This is achieved in two ways: either through the reduction of metal ions or by the gum molecules behaving as a stabilizing mediator (to prevent nanoparticle agglomeration). Hence, the essential criterion for designing nanomaterials with desirable properties is fulfilled [16,17].

The Kondagogu (*Cochlospermum gossypium*, KG) is a native tree growing in the forests of India and its exudates have been categorized as being substituted rhamnogalacturonans [16]. KG is a complex and acidic polysaccharide with high solution viscosity and gelation characteristics. KG contains sugars such as arabinose, rhamnose, glucose, galactose, mannose, glucuronic acid and galacturonic acid. The structural features assigned to KG are $(1\rightarrow 2)\ \beta\text{-D-Gal}\ p$, $(1\rightarrow 6)\ \beta\text{-D-Gal}\ p$, $(1\rightarrow 4)\ \beta\text{-D-Glc}\ A$, $4\text{-O-Me-}\alpha\text{-D-Glc}\ A$ and $(1\rightarrow 2)\ \alpha\text{-L-Rha}$ [18,19]. It has a high content of uronic acid and diverse functional groups (hydroxyl, acetyl, carbonyl and carboxylic) [18,20]. The morphological, physicochemical and structural analysis of this biopolymer is currently being extensively researched [18,19].

Organic dyes represent major clusters of pollutants flushed out into wastewaters from textiles and other industrialized procedures. Due to their potential injuriousness and presence in external waters, their elimination and removal has been a matter of considerable importance, since even small amounts of released dyes can discolor surface waters and impact negatively on the otherwise aesthetic and pristine surroundings [21]. The dyes altered the absorption and reflection of sunlight entering the water, which in turn affected bacterial growth and subsequently the level of biological impurities in the water [22]. Wastewater purification is discernible as one of the most serious environmental challenges of the present day. To this end, solid catalysts have been widely used in various water treatment technologies, both in processing and industry. Nanoparticles, courtesy of their small size and high

surface-to-volume ratios, display high absorbing, interacting and reactive capabilities, underscoring their value in wastewater remediation [23].

Photocatalysis is an environmentally benign process involving the conversion of light energy into chemical energy under ambient conditions. The outstanding performance of solar driven photocatalytic processes in solving environmental problems has gained much attention in recent years, effluents from the textile and paint industries being the main sources of environmental organic contaminants. Thus, it is essential that the latter is degraded and converted into harmless mineral compounds. The photocatalyst can be harnessed for environmental remediation, which includes abstraction of pesticides, fungicides and fertilizers from wastewater and the eradication of organic pollutants from the air [24]. Activation of a semiconductor photocatalyst is achieved by the absorption of a photon, which results in the transfer of an electron from the valence band to the conduction band by creation of a hole (h^+) in the valence band [25]. These photo-generated charge carriers cause redox reactions on the surface of the photocatalyst, i.e., any contaminant that is adsorbed onto the photocatalyst surface will undergo reduction or oxidation by the electron-hole pair respectively. Generally, the metal oxide photocatalyst surface acts as an active center in photocatalysis, either by the generation of OH^- radicals (by oxidation of OH^- anions) or by the generation of O_2^- radicals (by the reduction of O_2). Subsequently, these photo-generated radicals and anions react with the adsorbed organic contaminants, degrading or mineralizing them into less harmful by-products. The photocatalytic reaction can be employed to bring about the transformation of highly toxic chemicals into less noxious or non-toxic products such as CO_2 and H_2O [26,27]. Consequently, metal oxides can be employed as prime candidates for the effective photocatalytic degradation of environmental contaminants.

In this study, we have focused on the green synthesis of TiO_2 NPs and their effective application in the photocatalytic degradation of a commercially used organic dye, methylene blue (MB). The influence of various solution pH values on photocatalytic efficiency was also studied.

2. Materials and Methods

2.1. Materials

Titanium oxysulfate and methylene blue were purchased from Sigma Aldrich and HiMedia Chemicals, Mumbai, India, respectively. KG samples were obtained from the Girijan Co-operative Corporation Ltd. (GCC), Hyderabad, India. All the other chemicals and solvents used were of analytical grade.

2.2. Fabrication of TiO_2 Nanoparticles

Titanium dioxide NPs were synthesized using a typical procedure as described here. The procedure involved adding KG (50 mg) to 10 mL of titanium oxysulfate (0.1 M) and stirring vigorously (750 rpm) on a magnetic stirrer at 90–95 °C. Later, the product was centrifuged, washed and dried. The dried sample was calcined at 500 °C for 4 h, then pulverized and stored until further use.

2.3. Characterization of TiO_2 Nanoparticles

2.3.1. X-Ray Diffraction (XRD) Analysis

X-Ray Diffraction (XRD) configurations of the calcined samples were obtained with the diffraction angle range (2θ) set between 20° and 90° using a diffractometer (Rigaku Miniflex 600, Tokyo, Japan) with nickel filtered $Cu K\alpha$ ($\lambda = 1.54 \text{ \AA}$) radiation and a liquid nitrogen cooled, germanium solid state detector. The spectral plots were compared with details obtained from Joint Committee on Powder Diffraction Standards (JCPDS), data files for analytical purposes.

2.3.2. FTIR Analysis

For the Fourier transform infrared spectroscopy (FTIR) analysis, a spectrometer (Perkin-Elmer FTIR Spectrum Two, Singapore) in attenuated total reflection (ATR) mode and with the spectral range set between 4000 and 400 cm^{-1} and a resolution of 4 cm^{-1} was used.

2.3.3. Raman Spectra

A Raman microscope (NICOLET DXR, Thermo Scientific, Waltham, MA, USA), equipped with an optical microscope, was used. An argon-ion (532 nm) or helium-neon (632.8 nm) laser was used for the excitation of the Raman signal with appropriate holographic notch filters to eliminate the laser line after excitation. Spectral analysis and curve fitting were performed using GRAMS/AI 8.00 Spectroscopy software (Alfasoft GmbH, Frankfurt, Germany)

2.3.4. Scanning electron microscope- Energy-dispersive X-ray spectroscopy (SEM-EDX) Analysis

The elemental composition and morphology of TiO_2 nanoparticles were determined using a scanning electron microscope (ZEISS, Ultra/Plus, Potsdam, Germany). Energy-dispersive X-ray spectroscopy (EDX) measurements were carried out using a scanning electron microscope equipped with an EDX attachment (JSM-6390 172, JEOL, Tokyo, Japan).

2.3.5. TEM and High-resolution transmission electron microscopy (HR-TEM) Analysis

The nanoparticles and particle distributions determined were recorded by a transmission electron microscopy (TEM) (JEOL, JEM-2100, Tokyo, Japan). The samples were prepared on standard copper TEM grids covered with thin carbon foil. Drops of TiO_2 nanoparticles were dispersed in 1 mL of isopropanol using ultrasound for 10 min. and a drop of the resulting solution was gently spread onto the upper surface of the carbon covered copper TEM grid.

2.3.6. Particle Size Analysis

Nanoparticle size distributions were measured by centrifugal particle sedimentation (CPS) using the Disc Centrifuge technique (DC24000UHR, CPS Instruments Inc., Prairieville, LA, USA).

2.3.7. BET Surface Area

The surface area of the TiO_2 NPs was analyzed using the Brunauer–Emmett–Teller (BET) technique (Autosorb iQ, Quantachrome Instruments, Boynton Beach, FL, USA).

2.3.8. Thermal Stability

Thermal properties of KG and green-synthesized TiO_2 nanoparticles were studied using thermogravimetric analysis (TGA) by means of a Perkin Elmer STA 6000 Thermal Analyzer (Singapore) instrument.

2.3.9. Optical Properties

Optical properties were determined using an ultraviolet (UV)-visible spectrophotometer (Perkin Elmer Lambda 35, Singapore) over the spectral region 200–800 nm.

2.4. Photocatalytic Degradation of Methylene Blue

The photocatalytic activity of green-synthesized TiO_2 nanoparticles was scrutinized by the degradation of MB under sunlight. Photocatalytic activity in the presence of sunlight was determined under direct normal sunlight at an intensity of 100,000 Lux and the solar intensity measurement was carried out throughout the experiment at different time intervals. The degradation of dye was examined by collecting 5 mL aliquots from the reaction mixture at different intervals of time.

In addition, these aliquots were centrifuged at 7000 rpm for 15 min. Photocatalytic degradation of the dye was monitored by measuring the absorbance spectra of the supernatants using a UV–visible spectrophotometer over the wavelength range 200–800 nm.

2.4.1. Photocatalytic Studies Based on Catalyst Concentration

In order to study the effect of catalyst loading on solar light driven photocatalysis of methylene blue, different amounts (1–15 mg) of green-synthesized titanium dioxide nanoparticles calcined at 500 °C were added to 50 mL of dye solution. The reaction suspension was mixed thoroughly using a magnetic stirrer for 90 min in the presence of sunlight. The photocatalytic degradation of the dye was monitored by measuring the absorbance of the solution at regular intervals using a UV-visible spectrophotometer.

2.4.2. Photocatalytic Studies Based on Time

In a typical experiment, 50 mL of dye solution (1.0×10^{-5} M) was mixed with 10 mg of TiO₂ nanoparticles. The mixture was stirred continuously at 600 rpm in the presence of sunlight. The rate of degradation was monitored by measuring the absorbance of the solution (over the 200 to 800 nm wavelength range) with a UV-visible spectrophotometer by removing and monitoring 5 mL aliquots at defined time intervals. The process was continued for 90 min.

2.4.3. Photocatalytic Studies Based on pH

The influence of pH on solar driven photocatalytic activity of the catalyst was studied by conducting the experiment at different pH values (4–9) of dye solution. Green-synthesized titanium dioxide nanoparticles (10 mg) calcined at 500 °C were added to 50 mL of the dye solution having different pH values. The resulting suspension was thoroughly mixed using a magnetic stirrer for 90 min in the presence of sunlight and the absorbance of the solution measured with a UV-visible spectrophotometer.

2.4.4. Photocatalytic Studies Based on Temperature

The green-synthesized TiO₂ nanoparticles were calcined at different temperatures (500–900 °C) and each sample (10 mg) was added to 50 mL of dye solution. The resulting suspension was thoroughly mixed using a magnetic stirrer for 90 min, in the presence of sunlight. The absorbance of the solution was measured with a UV-visible spectrophotometer.

3. Results and Discussion

3.1. Mechanism of TiO₂ NPs Formation Via Green Synthesis

Natural tree based carbohydrate polymers—an environmentally benign medium—contain extensive numbers of hydroxyl, carbonyl and carboxylic groups which can act as good chemical reductants. The presence of these functional groups in the gum hydrocolloid material facilitates the formation of metal nanoparticles. When a metal oxide precursor is introduced into a well dissolved KG homogeneous solution or hydrogel, the polyhydroxylated macromolecules inherent in the gum matrix are able to absorb metal cations. For example, the Ti³⁺ ion could be chelated with KG by means of -OH and -COOH groups. The sequestration of cations [Mⁿ⁺] or hydroxylated cations [M(OH)]ⁿ⁺ that can undergo nucleation or growth processes is accelerated by the highly reactive hydroxyl groups present in KG. The last step in the TiO₂ nanoparticle synthesis is the calcination process at 500 °C which removes the gum template via combustion of their organic scaffold, giving rise to the formation of dispersed oxides [11,28]. While XRD and Raman spectroscopy confirmed the formation of anatase as the only distinguished mineral phase; FTIR, TGA and EDX collectively proved that no organic compounds (polysaccharides, hydroxyl, carboxylate and other oxidic groups) were present in the final nanoparticles.

3.2. XRD Analysis

The XRD pattern of TiO₂ NPs calcined at 500 °C showed a dominant peak at the 2θ value of 25.25° (Figure 1). This matches the (101) crystallographic plane of the TiO₂ anatase structure, indicating that the crystal composition is predominantly anatase. The other characteristic diffraction peaks were at the following 2θ values: 37.75, 48.02, 53.86, 54.95, 62.63, 68.88, 70.28, 75.13 and 82.60°. These values correspond to (004), (200), (105), (211), (204), (116), (220), (215) and (224) crystallographic planes of anatase, respectively (JCPDS No. 01-071-1166), thus confirming the formation of TiO₂ nanoparticles in the anatase phase [29]. The average crystallite size of the NPs [i.e., the mean size of the ordered (crystalline) domains, which may be smaller or equal to the grain size] was calculated using the Scherrer formula [$d = 0.89\lambda/\beta\cos\theta$] to be 12.58 nm. The crystallinity and higher purity of prepared TiO₂ NPs in the anatase form was validated by the presence of sharp peaks while the absence of peaks represented other crystallite forms of TiO₂ [30–32].

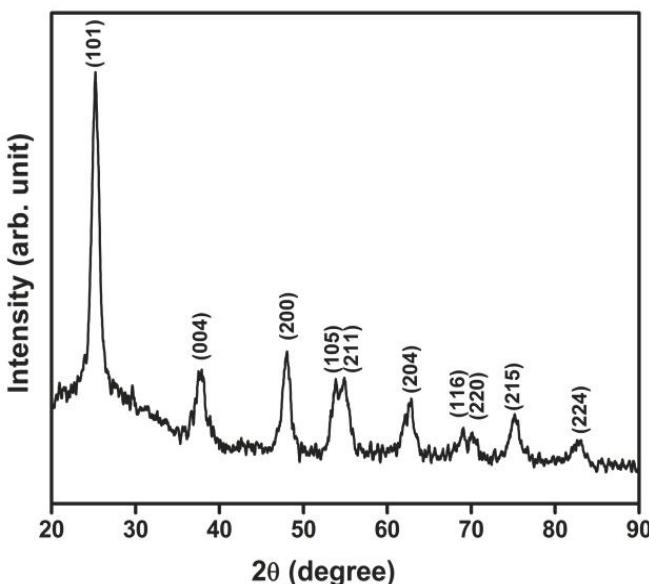


Figure 1. X-ray diffraction analysis (XRD) pattern of titanium dioxide (TiO₂) nanoparticles calcined at 500 °C.

XRD patterns of TiO₂ nanocatalysts calcined at various temperatures are shown in Figure 2. When the calcination temperatures did not exceed 700 °C, anatase was the only phase, with Joint Committee on Powder Diffraction Standards (JCPDS) No. 01-071-1166, while phase transformation of TiO₂ was observed at 800 °C. The rutile phase appeared in the 800 °C sample and became the dominant phase in the 900 °C sample with peaks at $2\theta = 27.4^\circ(110)$, $36.04^\circ(101)$, $41.2^\circ(111)$, $54.3^\circ(211)$ (JCPDS no. 01-073-2224). Formation of the rutile phase of TiO₂ is normally observed above 600 °C with a complete transformation to the rutile form occurring at 800 °C [33]. The present XRD patterns illustrate that the anatase to rutile phase transformation of the synthesized TiO₂ first took place at 800 °C and was almost completed at 900 °C, thus revealing the formation of a high temperature, stable anatase phase of TiO₂ nanocatalysts through the green synthesis method adopted.

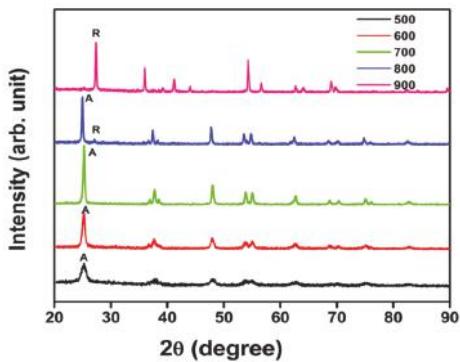


Figure 2. XRD pattern of TiO_2 nanoparticles calcined at different temperatures (R denotes rutile and A denotes anatase phases of TiO_2 nanoparticles).

While anatase and rutile are both mineral forms of titanium dioxide (possessing tetragonal crystal systems), in terms of optical activity, anatase is optically negative whereas rutile is positive. Furthermore, the luster exhibited by anatase is more strongly adamantine or metallic-adamantine than that shown by rutile.

3.3. Fourier Transform Infrared Spectroscopy Analysis

FTIR spectroscopy was employed to identify different functional groups present in KG and on the surface of the formed NPs (Figure 3). The major stretching frequencies in the spectrum for KG were observed at 3368, 1719, 1609, 1417, 1247, 1145 and 1035 cm^{-1} . The band observed at 3368 cm^{-1} suggests the presence of hydroxyl groups while those noted at 1719 cm^{-1} and 1609 cm^{-1} were ascribed to carbonyl stretching vibrations and the asymmetric stretching of carboxylate, respectively. The band seen at 1417 cm^{-1} was due to the symmetrical stretching of the carboxylate group present in the gum's uronic acid. The presence of an acetyl group was inferred by the band appearing at 1247 cm^{-1} while the bands registered at 1145 and 1035 cm^{-1} were indicative of C–O stretching vibrations of ether and alcohol groups, respectively [34]. The absence of these characteristic peaks in the FTIR spectra of TiO_2 nanoparticles promoted by KG may be a consequence of the higher purity of prepared nanoparticles in anatase crystal formations on calcination at $500\text{ }^\circ\text{C}$. The characteristic signal for TiO_2 nanoparticles observed below 1000 cm^{-1} in the FTIR spectra was due to Ti–O–Ti vibrations [35].

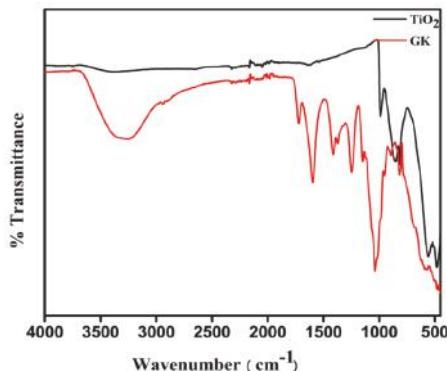


Figure 3. Fourier transform infrared spectroscopy (FTIR) spectra of Kondagogu (KG) and TiO_2 nanoparticles.

3.4. Raman Spectroscopy

To identify and quantify both the amorphous and crystalline TiO_2 phases, Raman spectroscopy was employed. Figure 4 reveals that TiO_2 anatase displays Raman bands at $639, 516, 399$ and 197 cm^{-1} , as well as a precise sharp and intense peak at 144 cm^{-1} . According to the literature, the anatase phase of TiO_2 has six Raman active modes, namely $A_{1g} + 2B_{1g} + 3E_g$, determined by group analysis $D_{4h}(\text{I4l}/\text{amd})$ [36], and known in the first-order Raman spectrum of single crystal TiO_2 at 144 cm^{-1} (E_g), 197 cm^{-1} (E_g), 399 cm^{-1} (B_{1g}), 516 cm^{-1} ($A_{1g} + B_{1g}$) and 639 cm^{-1} (E_g) [37].

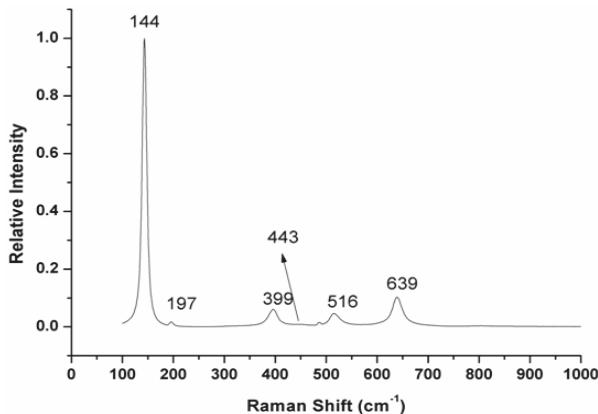


Figure 4. Raman spectrum of TiO_2 nanoparticles.

3.5. Scanning Electron Microscopy (SEM)

The SEM micrograph of TiO_2 nanoparticles is shown in Figure 5a, and the resulting nanoparticles were observed to have a spherical morphology. The chemical composition, analyzed using EDX spectra, confirmed the presence of Ti and O (Figure 5b). Since the sample was coated on a copper grid, peaks corresponding to Cu were also visible in Figure 5b.

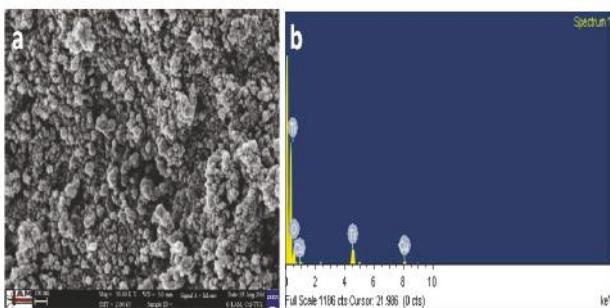


Figure 5. Typical (a) scanning electron microscopy (SEM) and (b) EDX (Energy Dispersive X-Ray) analysis of TiO_2 nanoparticles.

3.6. TEM and HR-TEM Analysis

Both particle size and morphology of TiO_2 were confirmed by TEM analysis (Figure 6a) which revealed that the particles are monodisperse and spherical in shape. The sizes of particles were in the $8\text{--}13\text{ nm}$ range and the selected area electron diffraction (SAED) pattern indicated that the TiO_2 nanoparticles possessed good crystallinity (Figure 6b). HR-TEM observations (Figure 6c) suggest that

the TiO_2 nanoparticles have a perfect lattice structure. The particle sizes determined by the TEM image were similar to the reported values obtained by applying the Scherrer equation to the XRD patterns (12.58 nm). The aforementioned equation correlates the size of sub-micrometre particles or crystallites (in a solid) to the broadening of a peak in a diffraction pattern. It is employed for the determination of the size of particles of crystals in the form of powders.

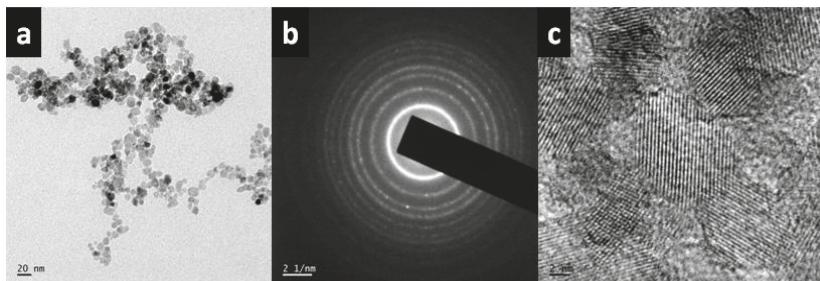


Figure 6. (a) TEM (Transmission Electron Microscopy) image, (b) SAED pattern and (c) HR-TEM (High Resolution TEM) micrograph of TiO_2 nanoparticles.

3.7. Particle Size & BET Analysis

The TiO_2 particle size distribution analyzed by CPS determined the mean size to be 11.2 ± 0.2 nm (Figure 7). Evidently, the TiO_2 nanoparticles appeared to be more stable in the current study involving green synthesis and the particle size measurements corresponded very well with the TEM analytical data (Figure 6). The specific surface area of TiO_2 was determined from the isotherms to be $42.6 \text{ m}^2/\text{g}$ based on the BET analysis.

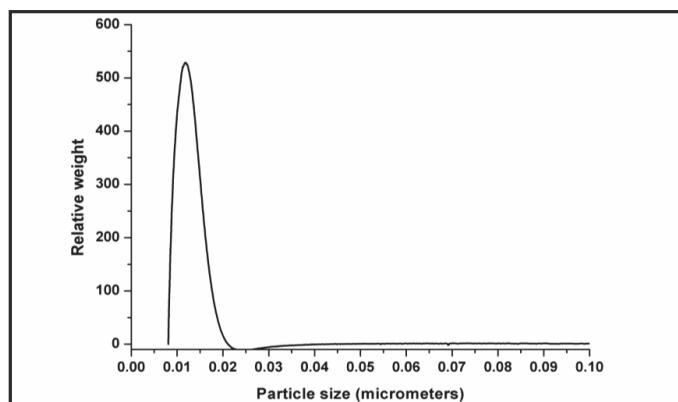


Figure 7. Particle size distributions of TiO_2 nanoparticles as determined by the centrifugal particle sedimentation (CPS) method.

3.8. Thermal Studies

Thermal properties of KG and green-synthesized TiO_2 nanoparticles were studied using thermogravimetric analysis. Depicted in Figure 8 are the TGA curves of KG and TiO_2 nanoparticles heated from 35°C to 950°C . In the case of KG, there were two major weight loss events. The first occurrence (observed between 35°C and 111°C) of approximately 17% probably represented the loss of adsorbed water - as hydrogen bonded water - from the polysaccharide structure. The second weight loss event of roughly 33% was very significant and was discerned between 232 and 309°C .

It was ascribed to decomposition of the polysaccharide. A third, far smaller weight loss instance was registered at 590 °C, possibly due to the conversion of the remaining polymer to carbon residue [20].

The TGA curve of TiO₂ (Figure 7b, red curve) did not indicate any weight loss up to 766 °C, a reflection of its high thermal stability, probably due to four hours of calcination at 500 °C. There were small linear weight reductions detected in the range of 766–911 °C, possibly caused by the loss of residual carbon from the gum matrix.

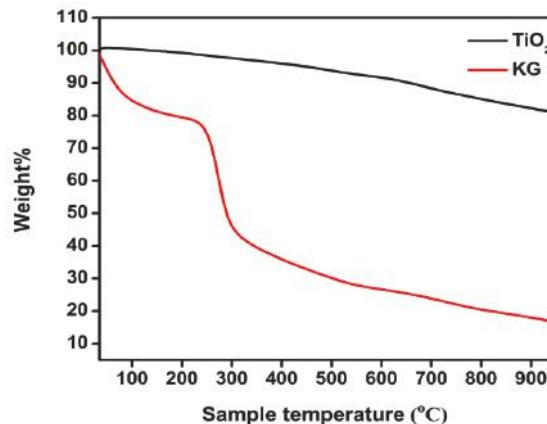


Figure 8. Thermogravimetric analysis (TGA) curves of KG and green-synthesized TiO₂ nanoparticles.

3.9. Optical Properties

The UV-visible absorption spectrum of biosynthesized TiO₂ nanoparticles is shown in Figure 8. The observed absorption spectrum matches those obtained with TiO₂ nanoparticles produced by chemical methods, displaying a broad absorption band in the UV region, up to 380 nm [38].

The UV-visible spectrum was utilized to deduce the optical absorption properties of green-synthesized TiO₂ nanoparticles (Figure 9). The band gap energy of green-synthesized TiO₂ nanomaterials was found to be 3.13 eV from the Tauc plot used to determine the optical bandgap in semiconductors as shown in Figure 10.

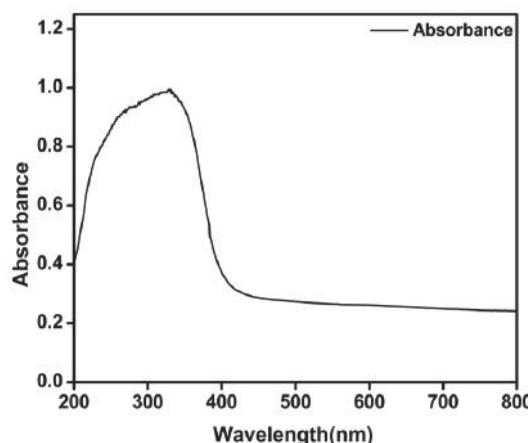


Figure 9. Ultraviolet (UV)-visible absorption spectrum of green-synthesized TiO₂ nanoparticles.

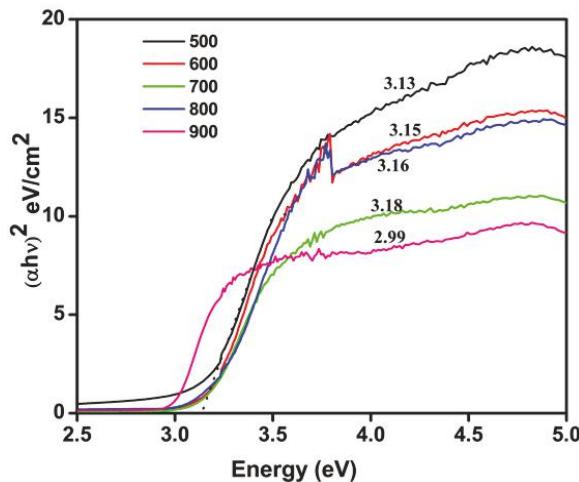


Figure 10. UV-Tauc plot of TiO_2 nanoparticles calcined at different temperatures.

After the bulk, TiO_2 had a band gap energy of 3.2 eV; the marginal reduction of 0.07 eV in this parameter being attributed to particle size dependence [39]. The absorbance spectra and the corresponding Tauc plot of TiO_2 nanoparticles calcined at different temperatures were as given in Figures 10 and 11, respectively.

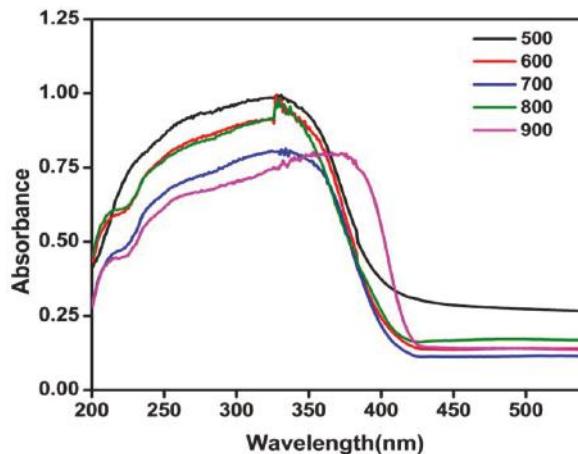


Figure 11. UV-Visible spectra of TiO_2 nanoparticles calcined at different temperatures.

The calculated band gap of TiO_2 nanoparticles, calcined at different temperatures was as shown in Table 1. It is also evident from the UV data (Table 1) that the band gap increased with the elevation of calcination temperatures from 500 to 800 °C, before decreasing gradually at 900 °C. This variation can be expected and is plausible, given that with an increase in the TiO_2 calcination temperature, the anatase phase became endowed with good crystal characteristics. Furthermore, the decrease in band gap at 900 °C was mainly attributed to the complete formation of the rutile phase.

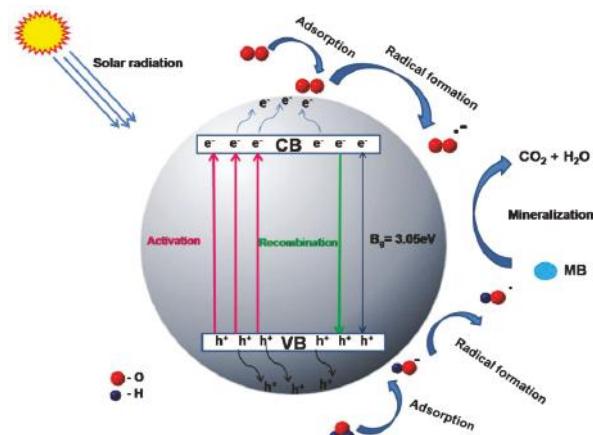
Table 1. Band gap of TiO₂ nanoparticles calcined at different temperatures.

Calcination Temperatures of TiO ₂ (°C)	Band Gap (eV)
500	3.13
600	3.15
700	3.16
800	3.18
900	3

3.10. Photocatalytic Activity

The photocatalytic activity of green-synthesized TiO₂ NPs was demonstrated by using an organic dye (methylene blue) under solar light, the dye degradation being initially identified by color change (Scheme 1). Additionally, we monitored the intensity of solar light throughout our experiment and confirmed it was in the range 100,000 Lux (see supplementary information, Figure S1). Furthermore, solar UV-radiation, as a function of time, and the UV intensity were measured using a Lux meter and UV filtering goggles. This information has been given in the supplementary document (Figures S2 and S3, respectively). The UV intensity from solar light was ascertained by calculating the difference between the intensity of solar radiation and the radiation through UV filtering goggles.

The dye displayed a distinct absorbance peak in visible light at a wavelength of 663 nm, where absorbance was at a maximum. This peak was used to monitor dye concentration in the solution over time. From the absorbance spectra corresponding to dye degradation, it was indisputable that the presence of titanium dioxide nanocatalysts resulted in a linear increase of the percentage of dye degradation with time, reflected by the decrease in absorbance. This finding indicated that when the time of irradiation was prolonged, the percent degradation increased and reached a maximum after 90 min of solar irradiation. When TiO₂ nanocatalysts, dispersed in the dye solution, were irradiated with solar radiation, photo-generated charge carriers induced redox reactions on the surface of the photocatalyst. Essentially, any contaminant adsorbed onto the photocatalyst surface, by virtue of electron-hole pair generation, will undergo either reduction or oxidation, respectively. Generally, the TiO₂ photocatalyst surface behaves as an active center in photocatalysis, either through the generation of OH radicals (by oxidation of OH⁻ anions) or by the generation of O₂⁻ radicals (via the reduction of O₂ molecules) [40]. Subsequently, these photo-generated radicals and anions react with the adsorbed organic contaminants to degrade or mineralize them into less harmful by-products, such as CO₂ and H₂O (Scheme 1).

**Scheme 1.** Photocatalysis mechanism of TiO₂.

As the irradiation time lengthened, absorption of increasingly more light energy impinging on the catalyst surface occurred. This led to the increased formation of photo-excited species, and consequently, enhanced photocatalytic activity. From this study, it was observed that the photocatalytic dye degradation process was enhanced by lengthening exposure time (Figure 12). The Beer-Lambert-Bouguer law—relating the attenuation of light to the properties of the material through which the light is travelling—was used to determine molar concentrations of the degraded dyes.

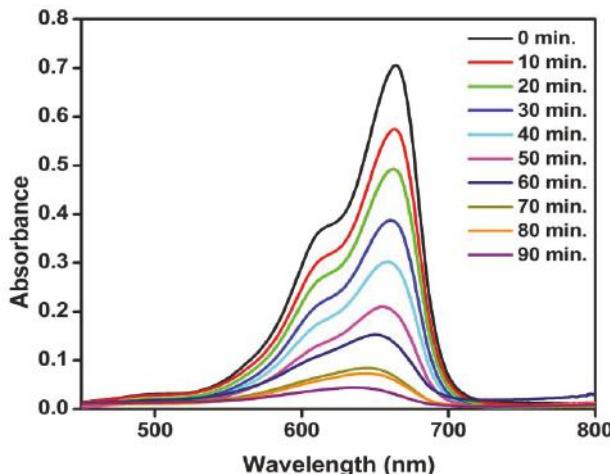


Figure 12. UV-visible spectra of methylene blue (MB) over various time intervals, in the presence of TiO_2 nanocatalysts.

No degradation of dye was discerned in the absence of solar light (supplementary information; Figure S2). The small decrease in absorbance observed was due to the insignificant adsorption of dye molecules on the catalyst surface. We have conducted a control experiment with MB on its own, as shown in supplementary information Figure S4. From this data, it is clear that TiO_2 was activated by solar light and was responsible for dye degradation. Moreover, photocatalytic degradation of dye was confirmed by the supplementary information (Figure S5, i.e., in the absence of sunlight, there is no degradation observed, confirming the major role of sunlight in the activation of the TiO_2 nanocatalyst). From Figures S4 and S5, it was confirmed that dye discoloration was entirely due to photocatalytic degradation, not adsorption.

3.10.1. Effect of Catalyst Concentration on Photocatalytic Activity of TiO_2 Nanoparticles

The effect of catalyst concentration on photocatalytic activity was tested by loading 1 to 15 mg/50 mL of TiO_2 nanocatalyst in methylene blue solution. The photocatalytic degradation of methylene blue was highly influenced by the level of catalyst loading, as evident in Figure 13 which shows that the percent degradation of the dye increased with the amount (from 1 to 15 mg/50 mL) of TiO_2 nanocatalyst loading and remained virtually constant above a certain level. This is because as the amount of catalyst increased, a greater number of active sites on the photocatalyst surface became available. Consequently, more OH radicals were produced, which facilitated their participation in the dye degradation process. However, beyond a certain limiting value of catalyst loading, the solution appeared turbid. As a result, the passage of solar radiation into the reaction mixture (required for the reaction to proceed) was obstructed, and thus, the percent degradation of the dye decreased or remained constant [41].

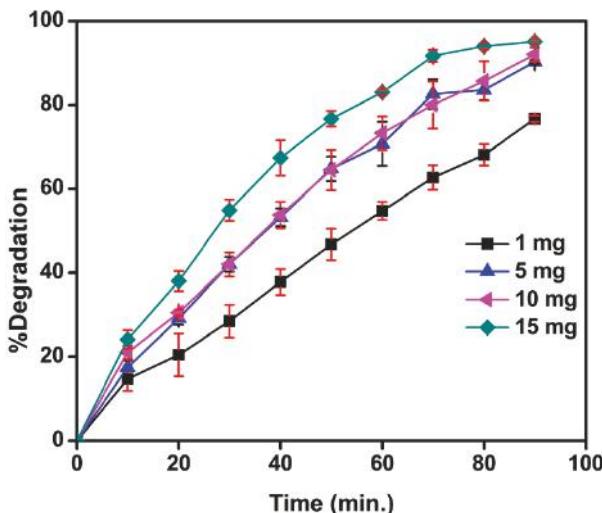
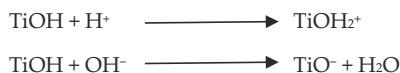


Figure 13. UV-visible spectra of MB dye (10^{-5} M) with different amounts of catalyst loading. Data obtained by degradation versus time are average values of three independent experiments (reported values = mean \pm S.D.)

3.10.2. Effect of pH on Photocatalytic Activity of TiO₂ Nanoparticles

The role of pH on the rate of photocatalytic degradation was studied over the 4–9 pH range and the results are illustrated in Figure 14. It was observed that the percent degradation increased with a rising pH, exhibiting a maximum between the pH 7–9 ranges. This pH variation may bring about changes to the surface charge on the TiO₂ nanoparticles and vary the potential associated with catalytic reactions. With the variation of potential, the extent of dye adsorption on the catalyst surface also fluctuates, culminating in the alteration of reaction velocity. Furthermore, under alkaline conditions, the surface of TiO₂ could acquire a negative charge. Since methylene blue is a cationic dye and the surface of TiO₂ nanoparticles in alkaline media attains a negative charge, the latter can be easily adsorbed onto the catalyst surface. This may lead to enhanced photocatalytic dye degradation under basic conditions [42]. The surface of TiO₂ nanoparticles, in acidic or alkaline circumstances, can be protonated or deprotonated, respectively, according to the following reactions:



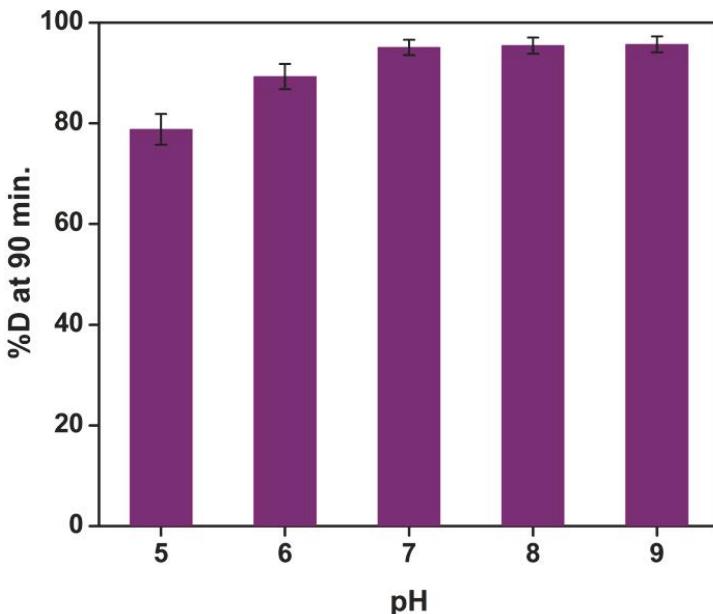


Figure 14. The photocatalytic degradation of MB dye (10^{-5} M) after treatment with photocatalyst TiO_2 (0.2 g/L) in solar light under different pH conditions (reported values are for three independent experiments; average value = mean \pm S.D.; $n = 3$).

3.10.3. Effect of Calcination Temperature on Photocatalytic Activity of TiO_2 Nanoparticles

The effect of calcination temperatures of the TiO_2 nanocatalyst on photocatalytic degradation is depicted in Figure 15. With increasing calcination temperatures of the TiO_2 nanocatalyst, the percentage degradation was found to have risen and attained a maximum of 96.85% at $700\text{ }^\circ\text{C}$. It then decreased gradually at temperatures between $800\text{ }^\circ\text{C}$ and $900\text{ }^\circ\text{C}$. This variation can be expected, given that an increase in calcination temperature of the TiO_2 generated an anatase phase possessing good crystal characteristics—an essential criterion for photocatalytic degradation. It is also apparent from the UV data (Table 1) that the band gap increased with rising calcination temperatures, from $500\text{ }^\circ\text{C}$ to $800\text{ }^\circ\text{C}$. Hence, the rate of electron-hole recombination decreased, and consequently, photocatalytic degradation increased. The decline in photocatalytic degradation at $800\text{ }^\circ\text{C}$ was mainly ascribed to the formation of the rutile phase, which existed as a mixture of both the anatase and rutile forms. Complete formation of the rutile phase transpired at $900\text{ }^\circ\text{C}$, and thus site activity fell and degradation decreased once again.

It is clear that the photocatalytic ability of titanium dioxide NPs is largely dependent on its crystalline form. Thermodynamically, the efficiency corresponding to photo-oxidation by the anatase and rutile phases should be similar. The surface recombination of photo-excited electrons and positive holes was higher in rutile than in anatase, which has a greater free (electron)-carrier mobility. Hence, a range of photoactivities was observed for each crystal form. This showed that variables such as crystal type and particulate sizes, as well as various synthesis routes (temperature, heating time etc.) largely determine its photocatalytic activity [40].

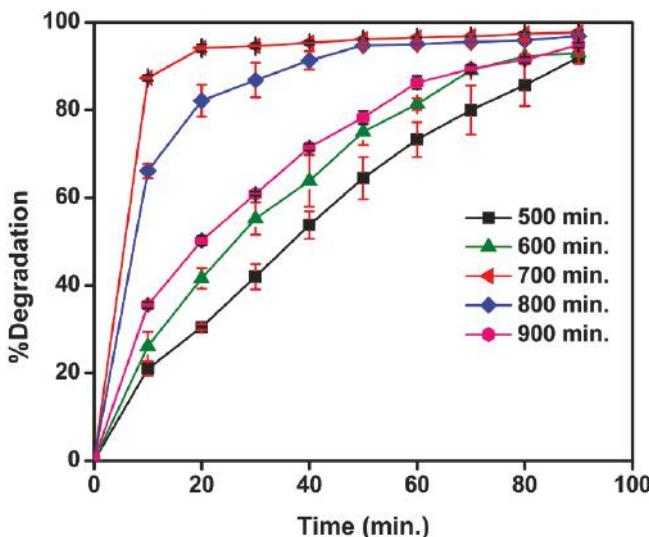


Figure 15. The photocatalytic degradation of MB dye (10^{-5} M) after treatment with photocatalyst TiO_2 (0.2 g/L) calcined at different temperatures under solar light (Values were reported as mean \pm S.D; $n = 3$).

4. Conclusions

Green nanotechnology has, in recent years, been accorded increasing importance for two reasons, namely, its contribution towards the elimination of harmful reagents and its ability to facilitate the synthesis of valuable products in a cost-effective manner. The green synthesis of TiO_2 NPs involves more compatible, eco-friendly, low cost and less time-consuming processes compared to other synthetic methods such as the sol-gel technique, which has been widely used to achieve the same ends. In the present study, titanium dioxide nanoparticles were produced using a natural hydrocolloid, gum Kondagogu (*Cochlospermum gossypium*). The crystallinity and high purity of the synthesized TiO_2 nanoparticles in the anatase form were unambiguously confirmed by the presence of sharp peaks (and the absence of unidentified peaks) in the X-Ray Diffraction patterns, Raman spectroscopic results and TEM images obtained. The absence of the original gum residue on the nanoparticles after calcination for 4 h at 500 °C was confirmed by FTIR, EDX, BET and TGA. The nanoparticles created had a mean particle size of approximately 11 nm, a value which was determined by three independent techniques: Scherrer's formula from XRD (12.58 nm); TEM (8–13 nm) and CPS (11.2 \pm 0.2 nm), the figures in parentheses being the respective particle sizes. The photocatalytic activity of green-synthesized titanium dioxide nanoparticles was evaluated by adopting methylene blue dye as a model system. It is apparent that the photocatalytic effectiveness of titanium dioxide nanoparticles is largely dependent on various factors including its crystalline form, exposure time, extent of catalyst loading and solution pH conditions. The present study has demonstrated that titanium dioxide nanoparticles (synthesized through a green route by the use of natural, renewable and eco-friendly materials) exhibit excellent photocatalytic activity towards organic dye degradation. This system can be employed in water purification and dye effluent treatment.

Supplementary Materials: The following are available online at <http://www.mdpi.com/2079-4991/8/12/1002/s1>, Figure S1: The intensity of solar light measured throughout the experiment (range 100,000 Lux); Figure S2: The % of UV intensity in solar radiation at various intervals of time; Figure S3: The solar UV intensity measurement using Lux meter and UV filtering goggles; Figure S4: UV-visible spectra of MB over various time intervals after treatment with TiO_2 nanocatalyst under dark condition; Figure S5: UV-visible spectra of MB over various time intervals in the absence of TiO_2 nanocatalyst.

Author Contributions: Formal analysis, S.W.; Funding acquisition, M.Č.; Investigation, K.S.; Methodology, C.S.; Project administration, M.Č.; Resources, R.P.; Supervision, B.G. and M.Č.; Writing—original draft, K.S.S.; Writing—review & editing, V.V.T.P.

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Conflicts of Interest: The authors declare no conflict of interest.

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Article

Environmentally-Friendly Green Approach for the Production of Zinc Oxide Nanoparticles and Their Anti-Fungal, Ovicidal, and Larvicidal Properties

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Abstract: Green synthesis of nanoparticles can be an important alternative compared to conventional physio-chemical synthesis. We utilized *Scadoxus multiflorus* leaf powder aqueous extract as a capping and stabilizing agent for the synthesis of pure zinc oxide nanoparticles (ZnO NPs). Further, the synthesized ZnO NPs were subjected to various characterization techniques. Transmission electron microscope (TEM) analysis showed an irregular spherical shape, with an average particle size of 31 ± 2 nm. Furthermore, the synthesized ZnO NPs were tested against *Aedes aegypti* larvae and eggs, giving significant LC₅₀ value of 34.04 ppm. Ovicidal activity resulted in a higher percentage mortality rate of 96.4 ± 0.24 at 120 ppm with LC₅₀ value of 32.73 ppm. Anti-fungal studies were also conducted for ZnO NPs against *Aspergillus niger* and *Aspergillus flavus*, which demonstrated a higher inhibition rate for *Aspergillus flavus* compared to *Aspergillus niger*.

Keywords: *Scadoxus multiflorus*; leaf; ZnO NPs; larvicidal; ovicidal; anti-fungal

1. Introduction

Currently, nanotechnology is a field of intense interest. The process of nanotechnology has been generally classified into three techniques: computational, wet, and dry. While the computational process deals solely with nano-sized structures, the wet process deals with components present in the cells, tissues, and membranes of living organisms. Additionally, the dry process deals with the synthesis of inorganic materials with the help of physical chemistry techniques. The major function of nanotechnology is said to be the synthesis of nanoparticles, mainly relying on the three methodologies such as physical, chemical and biological methods. Of these methodologies, biological synthesis plays a major role when compared with the two other methodologies [1–5]. Biologically-mediated synthesis is further classified into eco-friendly synthesis, which is comprised of plants and plant sources with the corresponding advantages of simplification and lower cost [6–13]. Therefore, we decided to mainly focus on the green synthesis of nanoparticles. For this method of green synthesis of nanoparticles, our research group chose *Scadoxus multiflorus* (*S. multiflorus*) leaf powder aqueous extract (SA) as a green source. This plant is also said to be one of the ancient medicinal plants of India, and belongs to the Caesalpiniaceae family. Different sources of this plant are highly recommended for various treatment purposes, such as irregular menstruation [14]. *S. multiflorus* is a bulbous plant found in most of sub-Saharan Africa which has been used as traditional medicine.

Metal oxide nanoparticles have various significant application possibilities, such as anti-microbial, cell line studies and dye degradation properties. Zinc oxide nanoparticles (ZnO NPs) have a band gap of 3.37 eV, which is relevant for various human applications [15,16].

In this manuscript, we synthesized ZnO NPs with the help of SA. Furthermore, the synthesized ZnO NPs were used to treat one of the major diseases, dengue fever, causing death in India. Dengue

is a global disease, with nearly 3 million people affected [17]. *Aedes ageypti* has been stated to be a common vector for causing dengue fever [18]. This manuscript concludes that ZnO NPs are anti-fungal agents effective against *Aspergillus flavus* (*A. flavus*) and *Aspergillus niger* (*A. niger*). Many researchers had reported on the anti-fungal activity of ZnO NPs, which proved to us that ZnO NPs could be utilized as fungicidal agents [19–23].

Overall, this manuscript describes the green synthesis of ZnO NPs using SA, and the subjection of the synthesized particles to various application studies, such as larvicidal and ovicidal activities against *Aedes ageypti* (*A. ageypti*). Furthermore, the synthesized particles, subjected to two different fungal strains, i.e., *A. flavus* and *A. niger*, were studied and are reported herein.

2. Materials and Methods

2.1. Materials and Reagents

The *S. multiflorus* leaf powder was directly procured from the local market and utilized in our research. Zinc acetate was obtained from Sigma-Aldrich (Riyadh, Saudi-Arabia). Reverse-osmosis and double-distilled water was used for the other experiments performed in this study.

2.2. Extraction of the *Scadoxus multiflorus* Leaf Powder Aqueous Extract Sample

The 30 g of procured powder material of the *S. multiflorus* leaf was immersed in 100 mL of distilled water and placed in a water bath at 60 °C for 1 h. Then, the solvent and powder layer were separated using a Buchner funnel and Whatmann filter paper. The filtrate solution of SA was collected and stored in a refrigerator to be utilized for the future synthesis of ZnO NPs.

2.3. Production of Zinc Oxide Nanoparticles

By using a pipette and mechanical stirrer, 20 mL of collected SA filtrate was added, drop by drop, to 80 mL of 1 mM of zinc acetate under stirring at room temperature (RT). Then, the resultant solutions were placed in a water bath at 60 °C for 3 h and monitored using UV-visible spectroscopy (Hitachi, Tokyo, Japan). Once the reaction mixtures confirmed the formation of ZnO NPs, the resultant solution was subjected to centrifugation at 3000 rpm for 20 min. The centrifugation processes were repeated three times with the help of distilled water to synthesize pure ZnO NPs. Once the centrifugation process was over, the supernatant was discarded, and the pellets were collected and placed in a furnace at 400 °C to obtain the desired product in powder form.

2.4. Analytical Techniques

After synthesis of the ZnO NPs, various analytical techniques, such as UV-visible spectrophotometry (Hitachi, Tokyo, Japan) were performed for the determination of the absorption maximum of the particles. The prepared material was mixed along with KBr to form pellets, to determine the Fourier-transform infrared (FTIR) spectroscopy, using a Shimadzu FTIR Spectrophotometer (Hitachi, Tokyo, Japan). The crystalline nature of the material was characterized by applying an X-ray diffractometer (XRD) (Model D8, Bruker, Germany). Transmission electron microscopy (TEM) (FEI company, Hillsboro, OR, USA) was performed to determine the morphology of the material. A particle-size histogram was developed using Image J software and the Zeta potential was determined by a Horiba nanoparticle analyzer (Horiba scientific, Kyoto, Japan), to identify the stability of the nanoparticles. Shimadzu atomic absorption spectrometry (Shimadzu, Kyoto, Japan) was used to determine Zn, with the help of a deuterium lamp.

2.5. Larvicidal and Ovicidal Properties of Synthesized Zinc Oxide Nanoparticles

Aedes ageypti (*A. ageypti*) larvae were cultured in the laboratory at RT. The third instar larvae were collected and utilized for larvicidal studies; the eggs were collected for ovicidal activity under various concentrations of ZnO NPs—15, 30, 60, and 120 ppm—which were studied and reported

using a MANOVA; LSD-DMRT Test. LC50 and LC90 values were also calculated, and identified to be statistically significant at $p < 0.05$. In this study, Neem azal, which is a commercially available insecticide, was utilized as the standard for ovicidal activity [24,25].

2.6. Antifungal Activity of Zinc Oxide Nanoparticles

The studied fungal strains, such as *A. flavus* MTCC 873 and *A. niger* MTCC 282, were procured from IMTECH (Chandigarh, India) and were then processed by using Clinical Laboratory and Standard Institute (CLSI) methods. An amount of 100 mL of PDB (Potato Dextrose Broth) was autoclaved, and *A. flavus* and *A. niger* fungal strains were inoculated into the broth. Test samples of 1 mg/mL were placed in an incubator while being stirred at 120 rpm at RT. After two weeks, the strains were collected, and the biomass of the fungi was filtered and kept for drying. This dried biomass was utilized for further studies, with carbendazim as the standard [26,27]. The mortality percentage of the fungal biomass was calculated using the formula below.

$$\frac{\text{Weight of the control} - \text{Weight of the test}}{\text{Weight of the control}} \times 100$$

3. Results and Discussion

3.1. UV–Visible Spectroscopy

The reaction mixtures of SA and zinc acetate were monitored using UV–visible spectroscopy at the wavelengths of 200 to 800 nm. From the observed results it can be inferred that the highest absorbance of 274 nm is at 90 min, which relies on the conversion of the starting material to end product, as clearly illustrated in Figure 1.

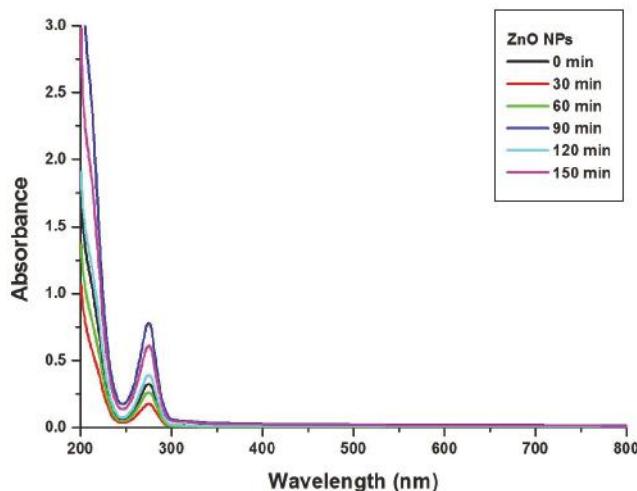


Figure 1. UV–visible spectroscopy of ZnO NPs.

3.2. FTIR Analysis of Zinc Oxide Nanoparticles

The specimens were subjected to FTIR study, as illustrated in Figure 2. Sample *S. multiflorus* leaf extract and ZnO NPs were both recorded to give the FTIR spectra. The FTIR spectrum of the *S. multiflorus* leaf extract shows peaks at 3003 and 1730 cm^{-1} , which correspond to functional groups such as C=O and C-H (stretch), present in organic molecules. These peaks completely disappear in the ZnO NPs spectrum, which clearly illustrates that the organic molecules are acting as capping

and stabilizing agents. The ZnO NPs spectrum showed a characteristic Zn–O stretching at $\sim 417\text{ cm}^{-1}$, which confirms the formation of ZnO.

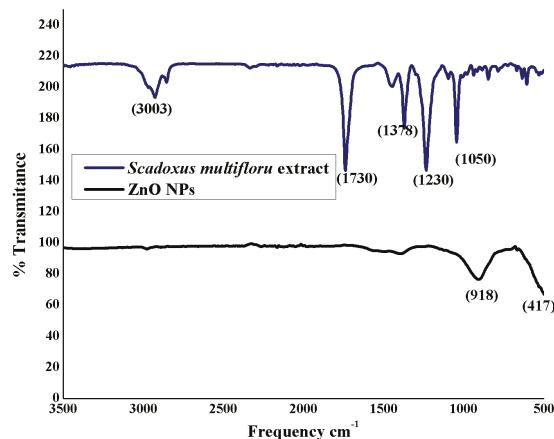


Figure 2. FTIR analysis of ZnO nanoparticles and extract.

3.3. XRD Analysis of Zinc Oxide Nanoparticles

The obtained ZnO NPs were investigated to study their crystalline nature by XRD spectroscopy. From the results it can be inferred that the synthesized ZnO NPs were synthesized in their pure phase, without any impurities. The results also confirmed the $h\ k\ l$ values of the (100), (002), (101), (102), (110), (103), (200), (112), (201), and (004) crystalline pattern. Furthermore, the crystalline structure was matched with the JCPDS data of 36-1451, and with the help θ of full-width and half-maximum data, with $d = 1.64056$ and $2\theta = 37$. Twenty-three plane crystalline data were calculated by Scherrer's formula $D = k\lambda/\beta \cos\theta$ [28]. The synthesized crystalline particles were said to be 31.8 nm in size, as illustrated in Figure 3.

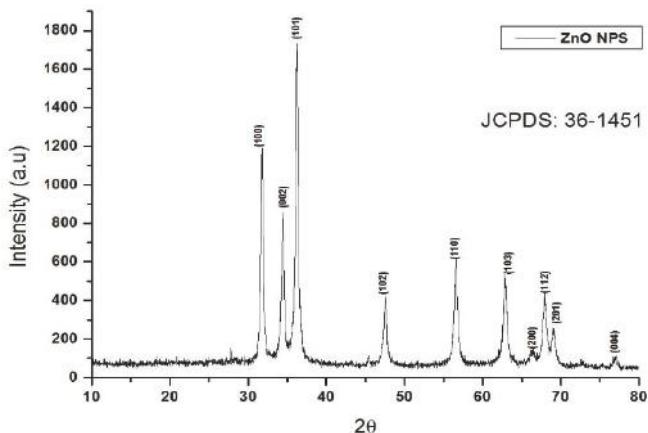


Figure 3. X-ray diffraction (XRD) analysis of the ZnO NPs.

3.4. Zinc Oxide Nanoparticles Morphological Studies

Eco-friendly synthesized ZnO NPs were identified by their morphology using transmission electron microscopy (TEM). From the observed results it can be inferred that the synthesized pure ZnO NPs show irregular, spherical-shaped particles, as illustrated in Figure 4a–c. The particles seem to be legitimately aggregated, with sizes in the range of ~100 nm. The Selected area (electron) diffraction pattern also clearly cuts the crystalline nature of eco-friendly synthesized ZnO NPs, as shown in Figure 4d. This is a typical phenomenon, taking place due to interaction of H₂O and ZnO. Due to inter-particle interactions, such as van der Waals and electrostatic or magnetic forces, the ZnO NPs in aqueous medium have a tendency to exhibit as an aggregated particle, leading to the development of soft agglomerates. Conversely, particle agglomeration is not complex, because the application purpose (i.e., larvicidal, ovicidal, and fungicidal activity) of the ZnO NPs depends upon the particle size and not on the agglomerate size.

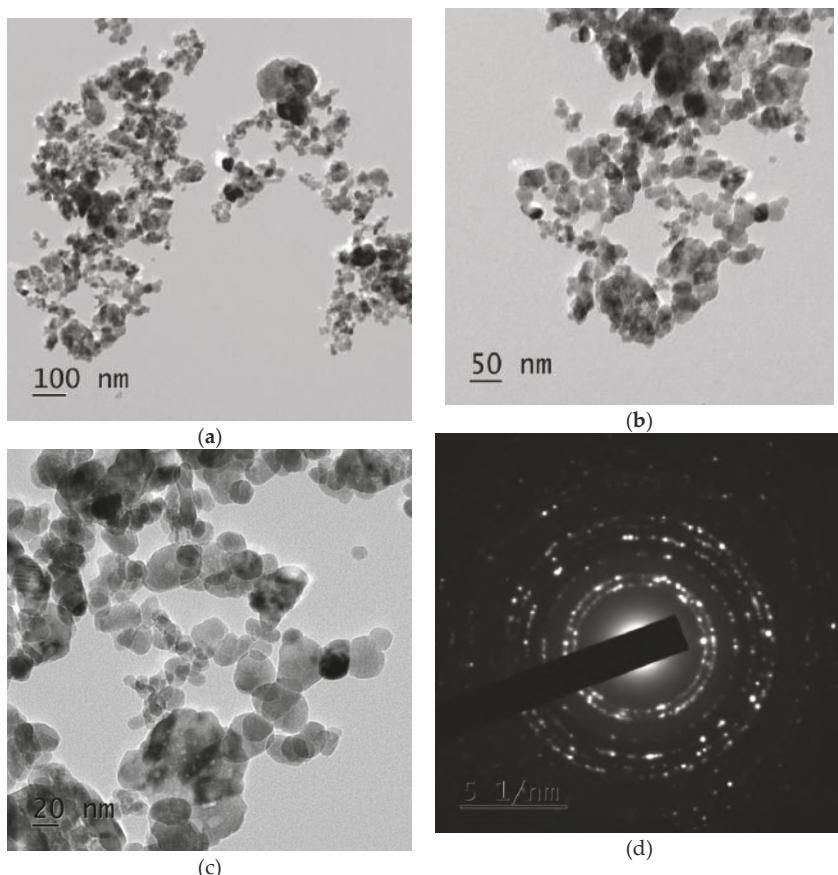
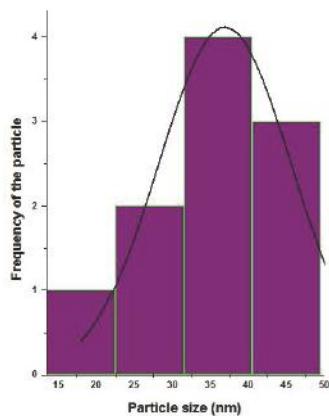


Figure 4. Cont.



(e)

Figure 4. (a–c) TEM images of ZnO NPs and (d) SAED pattern of ZnO NPs-particle size histogram (e) Particle size histogram.

3.5. Particle Size Histogram Analysis of the Zinc Oxide Nanoparticles

Our research group utilized the ImageJ software for plotting the particle size histogram. The results show that the eco-friendly synthesized ZnO NPs had an average particle size of 31 ± 2 nm, as shown in Figure 4e.

3.6. Energy Dispersive X-ray Analysis (EDAX) Spectrum of Zinc Oxide Nanoparticles

The synthesized ZnO NPs were subjected to an EDAX spectrum to quantify the mixture of metal and oxides present in the sample. The results showed that 64.12% of Zn and 35.76% of O were present on the surface area, as clearly shown in Figure 5.

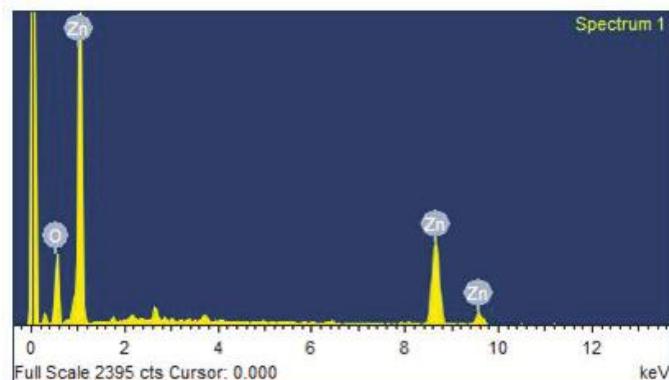


Figure 5. EDAX spectrum of ZnO NPs.

3.7. Stability of Synthesized Zinc Oxide Nanoparticles

The resultant ZnO NPs were subjected to determine the Zeta potential to test their stability, which resulted in a value of -51.8 mV, as clearly illustrated in Figure 6.

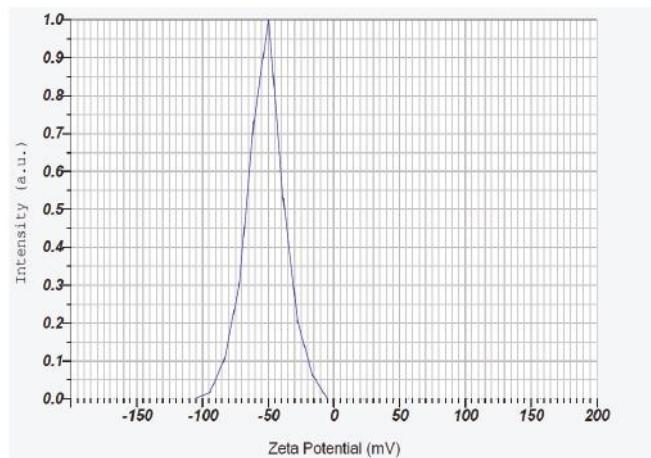


Figure 6. Zeta potential analysis of ZnO NPs.

3.8. Atomic Absorption Spectroscopy

The synthesized quantity of ZnO NPs was analyzed by atomic absorption spectroscopy (AAS) after adding the zinc acetate, with the intention of realizing the remaining concentration of zinc. AAS analysis for the nanoparticle preparing solution, performed at regular intervals of time, exhibited the formation of ZnO NPs. Initially, the standard solution of 5.02 ppm of zinc acetate was prepared and analyzed with AAS at 0 min. After adding *S. multiflorus* leaf extract and the zinc acetate, the formation of nanoparticles was observed at regular time intervals (Figure 7). The result showed a decrease in the concentration of zinc (5.02, 4.22, 3.13, 2.84, 1.87, and 0.08 ppm at 30 min intervals, respectively), indicating the conversion of zinc acetate to ZnO NPs. Additionally, in this present study, 1 gram dry weight of *S. multiflorus* leaves could synthesize 1.15 mg of ZnO NPs within 90 min. Furthermore, this is a sustainable method that does not use toxic chemicals.

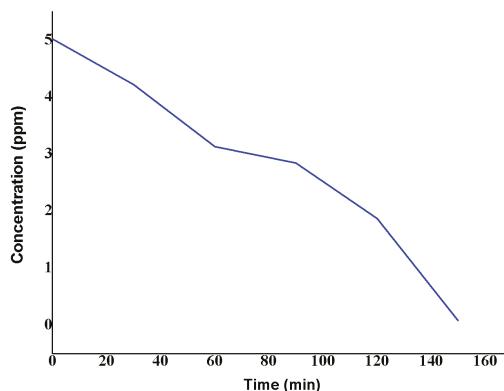


Figure 7. Atomic absorption spectroscopy analysis of zinc acetate in the nanoparticle-forming solution.

3.9. Larvicidal Activity of Zinc Oxide Nanoparticles

Dengue-causing vectors were treated with ZnO NPs at various concentrations: 15, 30, 60, and 120 ppm. The percentage mortality figures are 1.6 ± 0.4 , 28.6 ± 7.5 , 42.4 ± 2.5 , 82.2 ± 6.4 , and 98.4 ± 2.3 ,

respectively. This mortality percentage indicates a dose-dependent reaction at higher concentrations, as well as an increasing death rate. Lastly, with the help of LSD tests, we calculated LC₅₀ and LC₉₀ values with upper and lower confidence limits, as clearly illustrated in Table 1 showing significant results at $p < 0.05$ [10,28]. When compared to the literature [29], our methodologically synthesized ZnO NPs had less larvicidal activity, which may be due to the absence of a bio-organic phase on the surface of the ZnO nanoparticles. *Sargassum wightii*-mediated prepared ZnO NPs have a higher LC₅₀ value (49.22 ppm) compared to our result [30]. In another paper [31], *Ulva lactuca*-fabricated ZnO NPs were screened for larvicidal activity against *A. aegypti*, which showed an IC₅₀ value of 22.38 ppm. Our methodology provides highly crystalline, pure, and no-bio-organic-phase ZnO NPs. For the control experiment, 1.6% mortality was recorded. The LC₅₀ value for larval toxicity was 34.04 ppm.

Table 1. Larvicidal activity of synthesized ZnO NPs.

Concentration (ppm)	Mortality * (%)	LC ₅₀ (ppm)	95% Confidence Limits (ppm)		LC ₉₀ (ppm)	95% Confidence Limits (ppm)		χ^2 Value
			LCL	UCL		LCL	UCL	
Control	1.6 ± 0.4 ^a							
15	28.6 ± 7.5 ^b							
30	42.4 ± 2.5 ^c	34.04	14.82	50.32	78.06	58.75	143.75	3.189
60	82.2 ± 6.4 ^d							
120	98.4 ± 2.3 ^e							

The value represents the mean ± S.D. of five replications. * mortality of the larvae observed after 24 h of the exposure period, WHO (2005). LC₅₀: lethal concentration that causes 50% mortality; LC₉₀: lethal concentration that causes 90% mortality. LCL: lower confidence limit; UCL: upper confidence limit. Values in a column with a different superscript alphabet are significantly different at $p < 0.05$ (MANOVA; LSD-DMRT Test).

3.10. Ovicidal Activity of Zinc Oxide Nanoparticles

The eco-friendly synthesized pure form of ZnO NPs was subjected to *A. aegypti* eggs with Neem azal as a standard, with various concentrations: 15, 30, 60, and 120 ppm. The obtained results showed that the ovicidal activity relied on a dose-dependent reaction, with a higher mortality percentage of 96.4 ± 0.24 at 120 ppm. The obtained results after five replicates are depicted in Table 2 [10,28]. Our results relate to the literature [32], i.e., *Terminalia chebula* extracts against *A. aegypti*. The ovicidal activity of ZnO NPs was reported, and may be affected by diverse factors, predominantly egg age and contact period. The egg age influenced the ovicidal action of ZnO NPs. The exposure of freshly laid eggs to ZnO NPs causes higher mortality rates. Our output shows 96.4% mortality at 120 ppm, while *Terminalia chebula* (*T. chebula*) extracts exhibit only 66% mortality. The LC₅₀ value for ovicidal toxicity was 32.73 ppm.

Table 2. Ovicidal activity by green synthesized ZnO NPs.

Concentrations (ppm)	% of Mortality
15	35.5 ± 0.23
30	47.2 ± 1.21
60	63.7 ± 0.38
120	96.4 ± 0.24
Neem azal (120)	100 ± 0.00

Values represent mean ± S.D. of five replications. Different alphabets in the column are statistically significant at $p < 0.05$. (MANOVA; LSD-DMRT Test). Eggs in the control groups were not sprayed with phytochemicals. LC₅₀—32.73 ppm; LCL—24.20 ppm; UCL—44.27 ppm.

ZnO NPs were screened for ovicidal activity against which showed an IC₅₀ value of 32.73 ppm. Concerning the mechanisms of action of nanoparticles, Volker et al. noted that nanoparticles can affect various physiological parameters in treated organisms, both in vitro and in vivo. The results of in vitro

assays showed dose-dependent cell death with oxidative stress as the main likely toxicity pathway. In addition, silver nanoparticles may affect cellular enzymes by interference with free thiol groups and mimicry of endogenous ions. The nanoparticles affect the physiological process of the target organism [33]. On the other hand, strictly limited specific studies have been carried out to elucidate the precise mechanisms of the action of metal nanoparticles on insect pests and vectors [24,25]. However, in the present study, effort has been made to find the mechanism behind the mortality of the mosquitos. The scientific findings have been claimed that, the death of the mosquito may be due the absorption of the nanoparticle into the system and might affect the epithelial cell/ midgut or cortex [25]. It has been predicted that, when the ZnO nanoparticles were absorbed they gets accumulated in the midgut which leads to the shrinkage abdomen leads to the alteration of the mosquitos system. Alternatively, ZnO may affect the functions of other parts such as thorax and midgut, as well as other effects namely lateral hair loss, deformation in gills as well as brushes. Due to these damages in the system it might be the fact the mosquitoes could not undergo respiration hence forth leads to death.

3.11. Zinc Oxide Nanoparticles as Fungicides

Green synthesized pure ZnO NPs were subjected to two fungal pathogens: *A. flavus* MTCC 873 and *A. niger* MTCC 282. The ZnO NPs played a prominent role against *A. flavus*, with 75% inhibition at 500 ppm and 76% inhibition at 1000 ppm, while *A. niger* resulted in 57% and 63% inhibition, respectively [27], as clearly illustrated in Figure 8. The results were compared with the reported work [34]. The prepared ZnO NPs are active only at higher concentrations. Therefore, there is not much activity against *A. flavus* and *A. niger*. Many reports are available on ZnO NPs and their biological activity. These reports clearly state that smaller-sized nanoparticles (NPs) will have higher activity [35]. ZnO NPs might be toxic to some strains, but they are considered essential nutrients. The second reason for the antibacterial activity is that when the Zn²⁺ released by ZnO comes into contact with the cell membranes of the microbe, the cell membranes with negative charge and Zn²⁺ with positive charge mutually attract, and the Zn²⁺ penetrates into the cell membrane and reacts with sulphydryl groups inside the cell membrane. As a result, the activity of synthetase in the microbe becomes so damaged that the cells lose the ability of growth through cell division, which leads to the death of the microbe (Figure 9) [36,37].

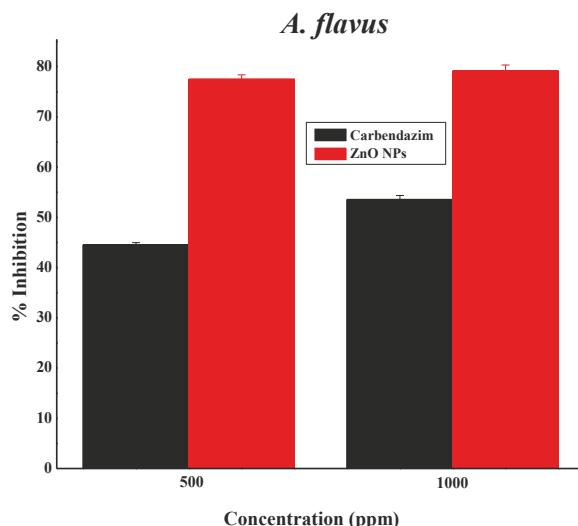


Figure 8. Cont.

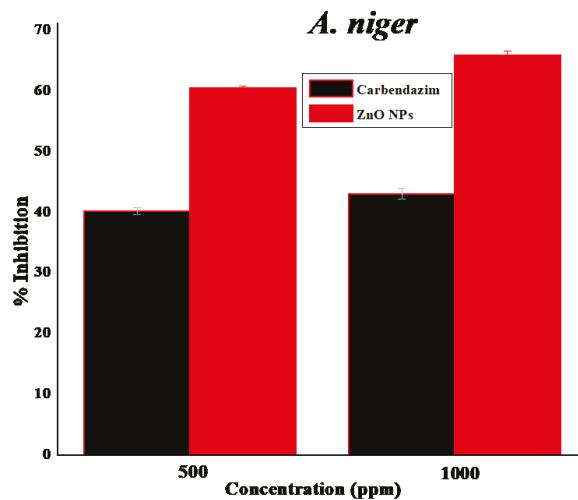


Figure 8. ZnO NPs' anti-fungal activity against *A. flavus* and *A. niger*.

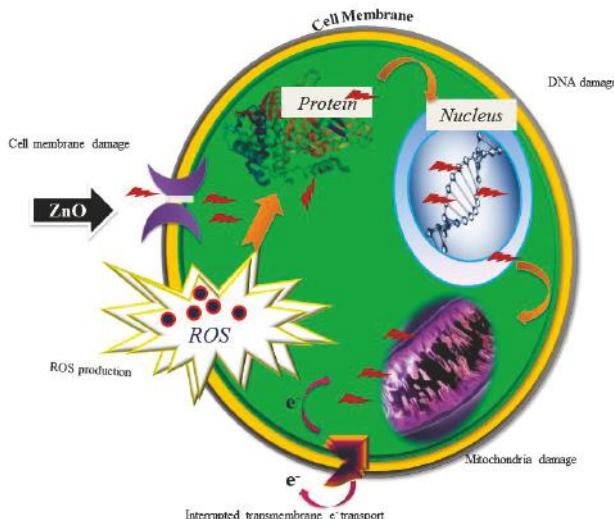


Figure 9. Mode of action of ZnO NPs on microbes.

4. Conclusions

In this manuscript, we proposed a simple process of ZnO NP synthesis by using SA. The results on the synthesized ZnO NPs showed they were irregular, spherical in shape, with an average particle size of 31 ± 2 nm. In addition, the material portrayed promising activity for dengue fever treatment by acting against *A. aegypti*, showing a significant difference at $p < 0.05$. Ovicidal activity was dose-dependent, with an increasing mortality rate at 120 ppm. The activity of the nanoparticles against two fungal pathogens resulted in a higher inhibition rate of *A. flavus* compared to *A. niger*. Moreover, they exhibited effective larvicidal properties against tested fungi and insects. Hence, this study

concludes that *S. multiflorus* mediated ZnO NPs may be used as effective control tools against mosquito larval populations and have potential applications in the pharmaceutical and biomedical field.

Author Contributions: N.A.A.-D. and M.V.A. designed the experiment. N.A.A.-D. and M.V.A performed the laboratory work. N.A.A.-D. and M.V.A. analyzed the results and drafted the manuscript.

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Article

Eco-Friendly Method for Tailoring Biocompatible and Antimicrobial Surfaces of Poly-L-Lactic Acid

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Abstract: In this study, a facile, eco-friendly route, in two steps, for obtaining of poly-L-lactic acid/chitosan-silver nanoparticles scaffolds under quiescent conditions was presented. The method consists of plasma treatment and then wet chemical treatment of poly-L-lactic acid (PLLA) films in a chitosan based-silver nanoparticles solution (Cs/AgNp). The changes of the physical and chemical surface proprieties were studied using scanning electron microscopy (SEM), small angle X-Ray scattering (SAXS), Fourier transform infrared spectroscopy (FTIR) and profilometry methods. A certain combination of plasma treatment and chitosan-based silver nanoparticles solution increased the biocompatibility of PLLA films in combination with cell line seeding as well as the antimicrobial activity for gram-positive and gram-negative bacteria. The sample that demonstrated from Energy Dispersive Spectroscopy (EDAX) to have the highest amount of nitrogen and the smallest amount of Ag, proved to have the highest value for cell viability, demonstrating better biocompatibility and very good antimicrobial proprieties.

Keywords: chitosan; poly-L-lactic acid; plasma; silver nanoparticles; antimicrobial

1. Introduction

Nowadays, there is a great demand for obtaining new environmentally-friendly and cost-effective materials to replace plastic products [1,2]. One of the most promising polymers is poly-L-lactic acid (PLLA); however, it has limited practical applications due to its low thermal stability and inherent brittle nature. Reinforcements with different substances prove to be a powerful tool in designing clean, eco-friendly materials for several applications [2]. PLLA is widely used in tissue engineering due to its slow degradation rate (almost 6 months to 1 year for complete degradation) [3,4] and cost-effective and effortless large scale production [5]. In this context, a fairly new area is represented by nanocomposites—the reinforcing material having the dimensions in nanometric scale [6]. Noble metal nanoparticles are of increased interest because of their potential applications in novel technologies due to their different properties compared to bulk metals [7,8]. Recently, silver and gold nanoparticles biosynthesis under eco-friendly conditions by using plant extracts, bio-organisms, proteins, and polysaccharides have gained an increased interest from material science researchers [9,10]. In this context, chitosan is one of the most used biopolymers for such approach mainly due to unique physicochemical properties in the presence of largely free amino and hydroxyl groups [11]. Chitosan is also recommended by proprieties like biocompatibility, non-toxicity, bioadhesivity, biodegradability, safety, and the promotion of drug absorption [12]. PLLA has been blended with chitosan to improve its wettability [13] or its tensile strength [14]. Some authors used chitosan as both capping and reducing agent for the incorporation of silver nanoparticles into the polymer matrices [15–17]. Due to the

ion–dipole intermolecular forces, chitosan stops the aggregation of silver in clusters at the macroscopic level and has a crucial role in stabilization of the formed nanoparticles [18].

Surface treatment procedures, such as plasma discharges in different kind of gases, can modify the physico-chemical proprieties at a scale of only a few atoms layers thick, without changing the bulk material properties [19,20]. Within the last decade, plasma technique has been applied to improve PLLA surface hydrophilicity, roughness and morphology and to proliferate the selective interaction between polymer surface and proteins. Some authors observed that melt extruded PLLA sheets treated with different kind of inert gas discharges did not affect PLLA biodegradation rate in soil [21]. Other authors used plasma treatment in reactive gases for enhancing cell (human skin fibroblast) adhesion on PLLA by obtaining reactive amine groups and further to immobilize collagen through polar and hydrogen bonding interactions at the treated film surfaces [22,23]. By increasing plasma treatment time, an increase of PLLA film degradation takes place [20]. It is well known that some issues related to non-permanent surface modification can occur, making polymers films unsuitable for certain applications in medicine or package food industry [24].

In this study, a facile, eco-friendly route, in two steps, for obtaining of PLLA/Chitosan-silver nanoparticles nanocomposite scaffolds under quiescent conditions was presented. This scaffold proved to be suitable for cell (preosteoblastic cell line MC3T3-E1 established from mouse C57BL/6 calvaria) attachment and proliferation and to have very good biocide proprieties.

2. Material and Methods

2.1. Materials

Chitosan flakes (molecular weight 50,000–190,000 Da based on viscosity, 75–85% deacetylated), silver nitrate and acetic acid were purchased from Sigma-Aldrich Chemie GmbH, Steinheim, Germany. Solutions were prepared with MilliQ water. Poly-L-lactic acid films of 50 µm thick were purchased from Goodfellow Cambridge Ltd, Huntingdon, UK and were cut in square shapes of 5 × 5 cm size.

Culture media and solutions: Alpha Minimum Essential Medium (α MEM, with ribonucleosides, deoxyribonucleosides, 2 mM L-glutamine and 1 mM sodium pyruvate, without ascorbic acid GIBCO, Custom Product, Catalog No. A1049001); Bovine fetal serum (BFS); Penicillin/Streptomycin/Neomycin solution (P/S/N) for cell culture; Phosphate Buffered saline (PBS) for cell culture; 3-(4,5-Dimethyl-2-thiazolyl)-2,5-diphenyl-2H-tetrazolium bromide (MTT), solution in PBS (5 mg/mL) were purchased from Sigma-Aldrich Chemie GmbH, Steinheim, Germany.

Cells: Preosteoblasts of MC3T3-E1 line, subclone 4 (passage 21) (purchased from Sigma-Aldrich Chemie GmbH, Steinheim, Germany) were thawed and multiplied in culture flasks with a surface of 75 cm^2 , in culture media α MEM, without ascorbic acid, supplemented with 10% BFS and 1% mixture of antibiotic. The initial density of cells culture was 2000 cells/ cm^2 culture surface.

2.2. The Two Step Method

2.2.1. In Situ Formation Silver Nanoparticles on Chitosan

Chitosan in a concentration of 6.87 mg/mL was dissolved in 1% glacial acetic acid solution. To avoid nanoparticle sedimentation due to the poor solubility of chitosan, the mixture was kept two days until a clear solution was obtained (Figure 1 Step 1) [25,26]. The solution was prepared by adding 5 mL of 9 mM AgNO₃ to 10 mL chitosan solution under stirring for 30 min at room temperature. The mixture was transferred to glass tubes and it was kept at 90 °C for 6 h in a temperature controlled bath. After synthesis, the AgNPs colloids were cooled at room temperature for the removal of the majority of the unreduced Ag⁺ to avoid the toxicity of the solution [27,28]. Then, the CS/AgNP solution was stored at room temperature and in dark glass tubes.

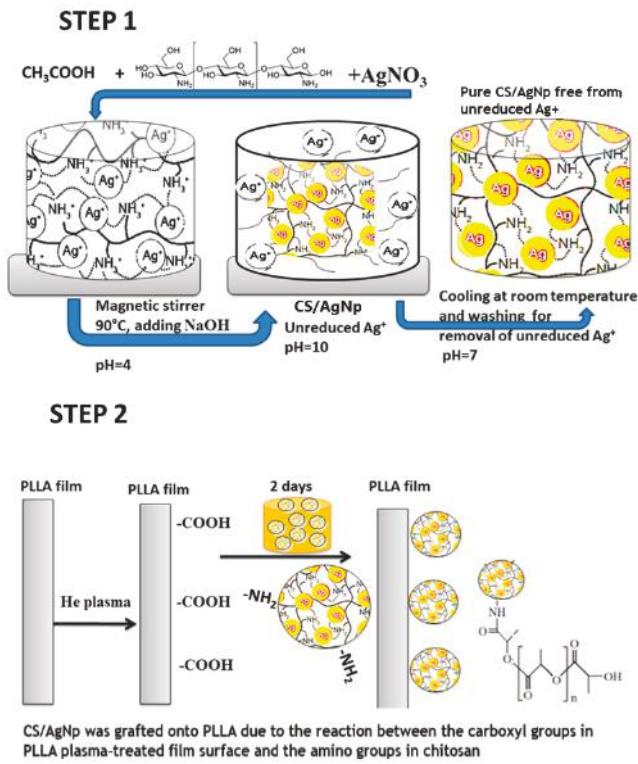


Figure 1. Scheme of two steps experimental method.

2.2.2. Combined Nonconventional and Conventional Treatments of PLLA Films

Briefly, the plasma treatments of PLLA films (Figure 1, Step 2) were performed in an EMITECH RF plasma device. Different times of treatments were performed to obtain a certain concentration of functional groups at the surface layer, without affecting the polymer bulk. The polymer was introduced in a gas vessel containing helium (He) at a pressure $p = 5 \times 10^{-2}$ mbar. Right after plasma treatments, to avoid the aging effect, the polymer was immediately immersed in the CS/AgNp solution at room temperature, in a dark place. After 2 days, the samples were rinsed with MilliQ water. Different input RF powers and times were performed for plasma treatments; however, for the present work, the following samples were selected: neat PLLA (P0) film, PLLA films treated in plasma for 4 min at 30 W, immersed in CS/AgNp (P1) and PLLA films treated in plasma for 10 min at 30 W, immersed in CS/AgNp (P2). For a lower time, no significant changes in the surfaces were observed, while for higher times the surface of PLLA film starts the degradation process, by becoming opaque and brittle.

2.3. Characterization Methods

A LUMOS Microscope Fourier Transform Infrared (FTIR) spectrophotometer (Bruker Optik GmbH, Ettlingen, Germany), equipped with an Attenuated Total Reflection (ATR) device. ATR-FTIR was used to acquire spectra in the range $600\text{--}4000\text{ cm}^{-1}$.

The scanning electron micrographs of PLLA samples were registered with a Quanta 200 microscope at an accelerating voltage of 15 kV and with an Energy Dispersive Spectroscopy (EDAX) system of elemental analysis (FEI Company, Brno, Czech Republic).

Transmission electron microscopy (TEM) images of CS/Ag Np solution were obtained on a HT1700 (Hitachi High-Technologies Corporation - Hitachi High-Tech, Tokyo, Japan) microscope using an acceleration voltage of 120 kV.

A Nanostar U-system (Bruker AXS GmbH, Karlsruhe, Germany) equipped with a Vantec detector and an X-ray micro source was used to perform small-angle X-ray scattering measurements (SAXS). The sample-to-detector distance was 107 cm and the wavelength of the incident X-ray beam was $\lambda = 1.54 \text{ \AA}$ (Cu K α). The solutions were loaded in capillary tubes and together with the film samples were measured under vacuum at a constant temperature, 25 °C for 10,000 s. Data analysis was performed by the model fitting approach using the DIFFRAC^{plus} NanoFit.

Particle sizes from Cs/AgNp solution were measured using a Zetasizer instrument (Zetasizer Nano ZS, Malvern Instruments Ltd, Malvern, Worcestershire, UK) using UV (ultra violet) Grade cuvette after treatment in an ultrasonic water bath (Model FB11012, Fisherbrand, Loughborough, UK) for 30 min to break up any aggregates present. All measurements were performed in triplicate.

A sample of the CS/AgNP solution was diluted with distilled water (1:10 sample:water) and analyzed by spectrophotometry Ultraviolet-visible (UV-vis) produced by Barloworld Scientific Ltd, Dunmow, Essex, UK.

Films roughness were determined using a Tencor Alpha-Step D-500 stylus profiler (KLA Tencor Corporation, Milpitas, CA, USA) both before and following the immersion in the Cs/AgNP for the samples at 1000 μm scan length, and 100 $\mu\text{m}/\text{s}$ scan speed. The arithmetic average of the absolute values of the profile heights over the evaluation length Ra was measured by applying a stylus force of 2.3 mg, and a long-range cutoff filter of 25 μm .

Static contact angle measurements using the sessile-drop method were performed on a CAM-101 (KSV Instruments Ltd., Helsinki, Finland) system equipped with a video camera, liquid dispenser and drop-shape analysis software (KSV CAM Optical Contact Angle and Pendant Drop Surface Tension Software, version 3.99, KSV Instruments Ltd., Helsinki, Finland). Liquid drops (double distilled water or ethylene glycol) of $\sim 1 \mu\text{L}$ were placed at room temperature, with a Hamilton syringe, on the polymer surface. For each drop 10 photos were recorded at an interval of 0.016 s. To obtain a statistical result, three different surface regions were selected for each liquid.

2.4. Culture Media Preparation and MTT Test

The samples were cut in fragments of $5 \times 5 \text{ mm}$ size and were decontaminated by immersion in a sterile solution of 70% ethyl alcohol for 20 min. Then the samples were rinsed three times in sterile PBS and were pre balanced in complete culture media for 24 h at 37 °C.

The MTT test was performed by a direct contact method, by using as test product the samples prepared as described above in 24-well culture plates populated with preosteoblasts MC3T3-E1 line, subclone 4. The initial cell population density was 1×10^4 cells/well, in 0.5 mL α MEM. The sample contact with the cells was made after 48 h after culture initiation, to obtain a cell monolayer semi-confluent. One piece of material with the size of $5 \times 5 \text{ mm}$ was then placed in each well over the cell culture for 72 h, at 37 °C, humidity 95% and 5% CO₂. Each sample was tested in triplicate material and the results were compared with the ones obtained for control samples, without testing material. The MTT test with 3-(4,5-dimethylthiazol-2-yl)-2,5-diphenyltetrazoliumbromid was carried out according to techniques from the literature [27,28]. The principle of the method is based on the reduction of yellow MTT compound, in a violet-colored product (formazan), as a result of mitochondrial dehydrogenase activity of viable cells. To achieve the MTT assay, the culture medium and the fragment of the material from each well was removed, the cells MTT solution was added in α MEM without BFS. After 3 h of incubation solution, the formazan absorbance was measured using a spectrophotometer plate reader (Tecan) at a wavelength of 570 nm. By reporting formazan absorbance from the wells with experimental samples to the control ones, the percent cell viability corresponding to the incubated culture with one of the tested materials was calculated.

2.5. Antimicrobial Activity

The polymers were cut into squares (1 cm side), plated in 24-well plates, sterilized with 70% ethanol (1) and washed with PBS (3 times). The test polymers were then incubated with 1 mL of bacterial suspension (*S. aureus* strains, *P. aeruginosa*, 0.5 McFarland turbidity). The plates were placed in the incubator at 37 °C for 2 h. Thereafter, the polymers were removed from the 24-well plates using a sterile forceps and were washed three times with PBS to remove the non-adherent bacteria. The films were then placed in tubes with 1 mL of PBS and vortexed for 120 s to remove all solutions from the adhering bacteria. Then, the solution was serially diluted in PBS, cultured on nutrient agar and the number of colonies forming units per ml (UFC/mL) was calculated.

3. Results and Discussions

In this work, chitosan was used as a mild reducing agent in silver nanoparticle synthesis. Protonized chitosan with NH₃⁺ functional groups was obtained by reaction of chitosan with H⁺ from the acetic acid solution. At the same time, the positions of Ag⁺ were fixed by coordination to the functional groups of chitosan, which simultaneously act as a stabilizing agent. The biopolymer behaves as a template or matrix which prevents the nanoparticles agglomeration.

The effect of surface modification experiments can be permanent (in the case of covalent attachment of functional groups) or non-permanent (non-covalent attachment). In the case of PLLA plasma treatment, the advantage is due to the improvement of surface wettability and cell affinity, while the disadvantage is due to surface rearrangement to minimize the interfacial energy, which affects the effectiveness of the surface modification, making the effect of plasma treatment non-permanent [29,30]. Another disadvantage of plasma treatment is degradation of PLLA in certain conditions. The two-step method described in our paper overcomes those disadvantages by immersing the PLLA plasma treated film in the chitosan-based AgNp solution immediately after the treatment, without any delay. In this way, the surface rearrangement does not have time to take place and the newly-introduced functional groups are efficient in permanent attachment of AgNP to the treated polymer surface. On the other hand, the input plasma parameters (the treatment time) were tailored to obtain no degradation of the PLLA film, the samples obtained in optimal conditions being selected to be presented in this paper. At lower values of the plasma treatment, no important changes in the film surface were noticed, while for higher values of the treatment, the PLLA film becomes brittle and opaque, the bulk properties being seriously affected.

The synthesis of AgNp in CS was demonstrated by UV-vis spectroscopy. The concentration of the chitosan-capped silver nanoparticles was approximately 0.11 ng/mL. The presence of surface plasmon resonance in the UV-vis characteristic optical spectrum indicates the presence of silver nanoparticles of certain particle size [31]. The UV-vis results (Figure 2) show a typical silver absorption peak at 415 nm which is in the reported range of silver and silver oxide nanoparticles [32–34]. The symmetry of the nanoparticles can be determined by the number of surface plasmon resonance peaks. If only one peak is observed in the UV-vis spectra, spherical silver nanoparticles were synthesized. The absorption peak is relatively narrow; thus, this method revealed a small size distribution of the particles in solution and the absence of unreduced positively charged ions, which is consistent with the data observed in the SEM and TEM images.

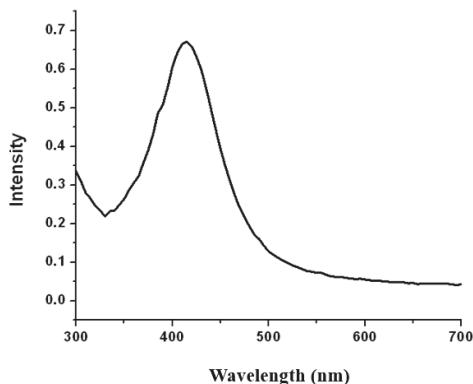


Figure 2. Ultraviolet-visible (UV-vis) spectrum of CS/AgNp solution.

3.1. TEM Images and Particle Size Results of CS/AgNp

The representative TEM images of AgNPs show that the particles were randomly distributed in the solution and due to the presence of the polymer, no agglomeration was detected (Figure 3a). The particles in Cs/Ag Np solution have a spherical shape (Figure 3b) and an average particle size of about 30 nm (Figure 3c) was obtained, in good concordance with the particle distribution obtained by dynamic light scattering method (Figure 3a). Other authors obtained similar results [35]. The contrast of TEM micrographs is correlated with the nature and size of the particles. The main discussion is between organic and inorganic particles or different organic-inorganic composites. Soft materials are predominantly composed of low number atoms, such as C, O and N. These elements, compared to heavy metals, exhibit a low level of electron-optical contrast. Thus, TEM micrographs of CS/AgNPs show AgNP formation—which appears as dark areas due to the high electron density of Ag. The contrast difference for CS/AgNP is because by drying the nanoparticle suspension, chitosan remains at the surface of the particles, and thus gives rise to areas of low contrast compared to AgNPs.

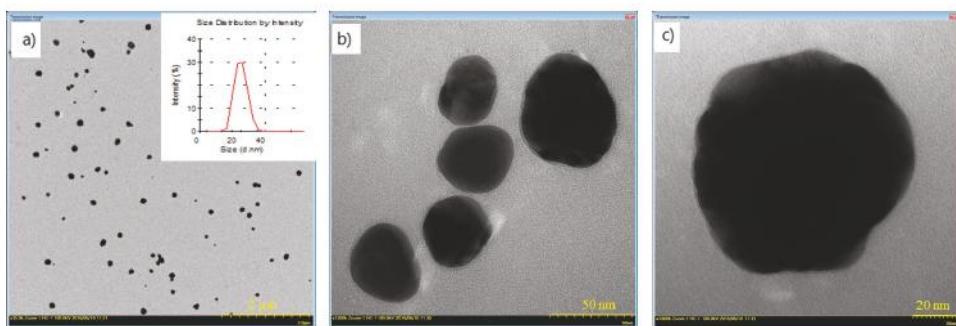


Figure 3. Transmission electron microscopy (TEM) images of CS/AgNp solution and Ag particle size distribution for scale: (a) 2 μ m; (b) 50 nm; (c) 20 nm.

3.2. SAXS Results

The confirmation of particle size as obtained from TEM images was further authenticated by the SAXS analysis performed for all studied samples and the results are illustrated in Figure 4. The SAXS plots on a double logarithmic scale of the pristine chitosan and CS/AgNp solution presented in Figure 4a) demonstrate no important modification in chitosan morphology after the reduction of silver. The change in the slope of the CS/AgNP SAXS pattern is due to the presence of the silver nanoparticles

in the system, and by applying the spherical model using DIFFRAC^{plus} NanoFit, a value of 35 nm was obtained, in good concordance with TEM results.

The structural changes in the PLLA samples after applying the two-step treatment is demonstrated by the Kratky Plot obtained from SAXS measurements (Figure 4b). It is well known that plasma treatments are accompanied by the heating of the polymer due to the interaction with high-energy plasma particles with the material surface. On the other hand, the crystal modification of PLLA is easily obtained from the melt [36]; therefore, the increase in the intensity and the shift at smaller angles of the PLLA peak in the Kratky plots is a clear indication that the process of PLLA crystallization takes place (Figure 4b). The appearance of the second SAXS peak as the treatment time in plasma increased suggests the formation of regular aligned lamellar structures in the polymer matrix due to the plasma treatment observed by other authors [37]. The average long period of a lamella L can be estimated from the maximum of the peak (q_{\max}) in the Kratky plots (Figure 4b) according to Bragg's law ($L = 2\pi/q_{\max}$). The thickness increased with the increase of plasma treatment time and lamellar structures grew in size from 0.25 nm for P1 to 0.27 nm to P2.

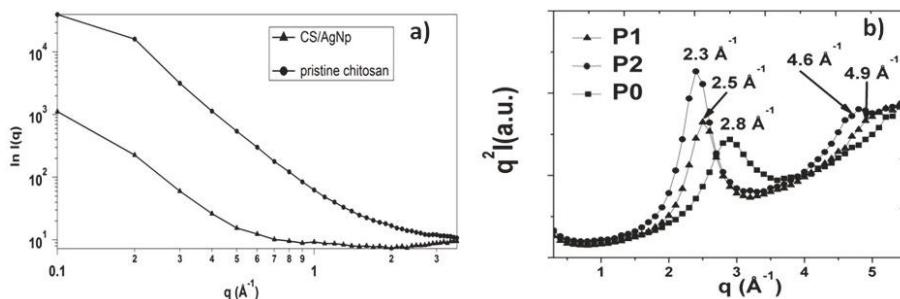


Figure 4. SAXS patterns: (a) double-logarithmic plot for pristine chitosan and CS/AgNp solution; (b) Kratky plot for pristine and treated PLLA samples.

3.3. FTIR Results

FTIR measurements were carried out to elucidate the interactions that take place in the reduction process of silver nitrate in the presence of chitosan.

In the pristine chitosan spectrum from Figure 5a, the broad absorption peak in the $2250\text{--}3800\text{ cm}^{-1}$ region is attributed to symmetric and asymmetric vibrations of CH_2 ($2250\text{--}3050\text{ cm}^{-1}$), and vibrations of O-H , N-H and intermolecular hydrogen bonds of polysaccharides ($3050\text{--}3800\text{ cm}^{-1}$). The peaks at 1588 and 1642 cm^{-1} were assigned to amino ($=\text{NH}_2$), amide I (C=O) and C=O of O-C-O-R groups in the chitosan structure, respectively. The peaks at 1060 , 1075 and 1176 cm^{-1} are the characteristic absorptions due to C-O vibrations in the C-O-C band [38]. The formation of silver/chitosan nanoparticles was confirmed by Fourier transform infrared spectroscopy. As shown in Figure 5b, the spectra of the CS/AgNP exhibited a few differences from the chitosan. The peak intensities in the range 1000 cm^{-1} and 1350 cm^{-1} due to C-N stretching and bending decreased because of the reduction of silver in chitosan. The absence of 1588 cm^{-1} peak that exists in pristine chitosan and the appearance of additional peaks at 1707 and 1744 cm^{-1} (Figure 5b), corresponding to carbonyl stretch vibrations in ketones, aldehydes and carboxylic, indicate that the silver is bound to the functional groups of chitosan. The formation of chitosan-silver nanoparticles was achieved after the reduction of silver ion through the amino group. The presence of these functional groups on the surface of the synthesized silver nanoparticles and the disappearance of the NH_2 double spike peak indicates that the polymer successfully capped the nanoparticles and the polymer network restricts the diffusion of Ag^+ . Moreover, the reduction of the silver ions is coupled to the oxidation of the hydroxyl groups in chitosan molecular and/or its hydrolyzates acids [39]. The band at 3434 cm^{-1} assigned to the overlap between the O-H stretching vibration and the N-H stretching vibration of the biopolymer moieties,

shifted to 3421 cm^{-1} due to co-ordination bond between the silver and electron rich groups [40,41]. After silver binding during reduction of silver nitrate with chitosan, the molecule weight was heavier and the vibration intensity of the N-H bond decreased, suggesting the attachment of silver to nitrogen atoms from chitosan [42].

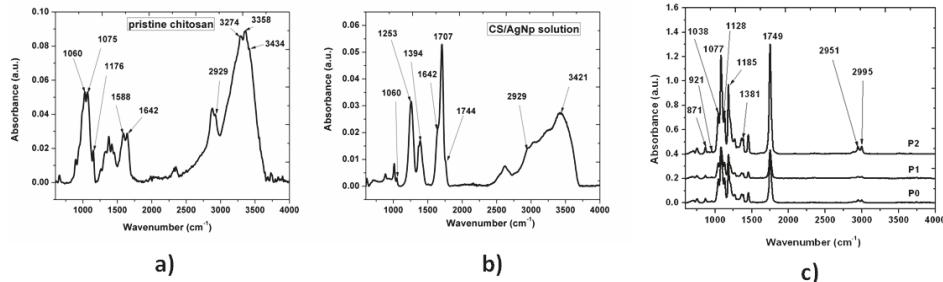


Figure 5. Fourier Transform Infrared (FTIR) results for (a) pristine chitosan; (b) CS/AgNp solution; (c) pristine and treated PLLA samples.

Figure 5c shows the IR spectra of the PLLA samples. The 1185 cm^{-1} and 1077 cm^{-1} bands were assigned to C—O—C asymmetric and symmetric stretching, while the peak at 1749 cm^{-1} was attributed to the stretching of C=O. The C—CH₃ stretching caused the peak at 1038 cm^{-1} and the C—H (of CH₃ groups) rocking mode was present in the spectrum at 1128 cm^{-1} . An increase in the degree of PLLA crystallinity due to the heating of the polymer during the plasma treatments can be observed, which is in good concordance with SAXS measurements. There was an increase of the 1749 cm^{-1} and of 1381 cm^{-1} band intensities. A shift with 3 cm^{-1} to lower wavenumbers of those bands can be observed. Those bands were assigned to the carboxylic groups and demonstrated an increase of those functional groups quantity after plasma treatments and the presence of silver ions at the polymer surface [43,44]. There was an increase of the 1749 cm^{-1} and of 1381 cm^{-1} band intensities assigned to the carboxylic groups, which demonstrate an increase of those functional groups quantity after plasma treatments and the presence of silver ions at the polymer surface. A shift with 3 cm^{-1} to lower wavenumbers of those bands can be observed, which revealed the reaction between the carboxyl groups in PLLA and the amino groups in CS; PLLA was grafted onto the backbone of CS.

3.4. Surface Roughness and Wettability

The polymer surface was affected by the plasma treatment due to the breaking of chemical bonds, heating, degradation, etc. [26]. Because all of these processes may significantly change the structure and morphology of the polymer film, the surface roughness and contact angle measurements were performed.

Hydrophobicity and hydrophilicity proprieties of polymer surface are key factors in further cell adhesion. Surface functional groups and the surface roughness of the material are very important in determining surface wettability.

To determine the wettability of neat and treated PLLA surfaces, water contact angle measurement (two liquids method) was used to provide the information on the wetting properties (Table 1). The values of the static contact angle (θ_w for water and θ_{EG} for ethylene glycol) can be used to estimate the wettability and surface tension of a solid surface. Based on these measurements, some parameters such as surface free energy (γ_{SV}), solid–liquid interfacial tension (γ_{SL}), or work of adhesion (W) were calculated using Owens–Wendt–Rabel and Kaelbe methods [45–47] and the results are listed in Table 1. The polar and the dispersive components of surface free energy were also listed in Table 1, to evaluate the surface modifications after the two step treatment.

Table 1. Surface roughness and wettability results.

	Roughness		Contact Angle Measurements Parameters						
	Ra (nm)	θ_w	θ_{EG}	W_w	W_{EG}	γ_{SV}^b	γ_{SV}^d	γ_{SV}	γ_{SL}
P0	43.103	79.9	52.19	86.19	77.42	8.05	23.93	31.98	18.58
P1	86.448	51.54	34	118.07	87.79	38.93	9.61	48.55	3.27
P2	63.831	70.11	42.22	97.56	83.54	13.96	22.39	36.36	11.60

The roughness parameters measured with profiler revealed that the increase of treatment time causes an increase of the surface roughness parameters (Table 1).

Both treatments caused a significant increase in the surface roughness and a decrease in the static water contact angle compared to the untreated polymer. The lowest water contact angles and the highest value for surface roughness were obtained for P1 treatment (Figure 6).

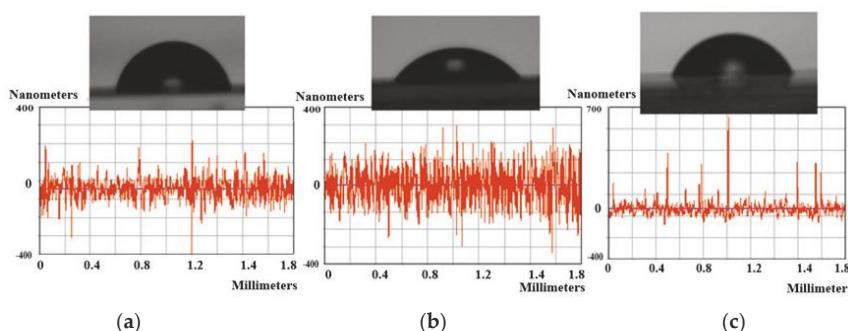


Figure 6. Surface modification results: top figures—contact angle images; bottom figures—profiler roughness measurements of samples: (a) P0; (b) P1; (c) P2.

The values listed in Table 1 revealed that both plasma treatments significantly increased the surface energy, mainly due to the increase in its polar component. Furthermore, the lowest value of the water contact angle was achieved at the largest values of the polar component of the surface energy.

3.5. SEM and EDAX Results

The morphological changes in the PLLA surface after the two-step treatments are presented in Figure 7. From Figure 7a, pristine PLLA film has a smooth surface without any irregularities. The PLLA films surfaces, after the combined plasma-wet chemical treatment, have patterns of different size and shape due to the surface interactions with different reactive species formed in plasma and due to the presence of AgNp and chitosan (Figure 7b,c, respectively). EDAX measurements (Table 2) show the presence of the AgNp on the surfaces of the P1 and P2 films after the treatments. The presence of nitrogen in the treated samples supports the presence of chitosan at the polymer surface.

It can also be observed that the highest amount of nitrogen was obtained in the P1 sample, while the high amount of Ag was obtained in the P2 sample. The carbon concentration was slightly increased with the increase of the plasma treatment time as a result of the polymer chain destruction. The low concentration of Ag nanoparticles assures low toxicity at the surface of the polymer and avoids the agglomeration of nanoparticles. Moreover, the presence of N at the treated samples surfaces demonstrates the presence of chitosan.

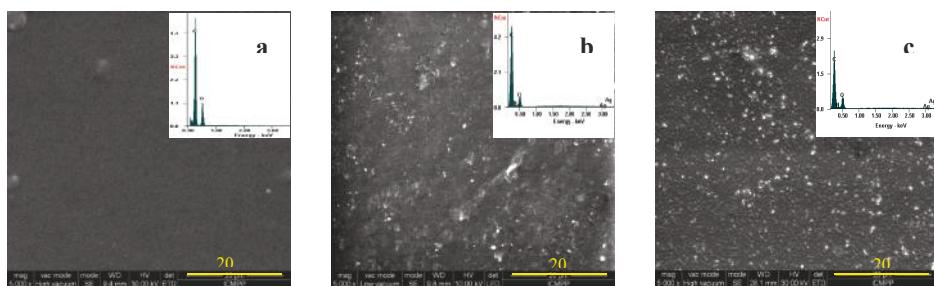


Figure 7. SEM images and Energy Dispersive Spectroscopy (EDAX) spectra of: (a) P0; (b) P1; (c) P2.

P1 and P2 samples were treated in plasma at different conditions and then introduced in the same solution of CS/AgNPs. After plasma treatment, functional groups of different concentrations (depending on the input plasma parameters) were present at the PLLA surface. Those groups are responsible for the presence of chitosan and silver at the PLLA treated surfaces, as presented in FTIR section. From Table 2, the modified N/C ratio on the surface was 0.19 for P1 and 0.17 for P2, while Ag/C ratio was 0.006 for P1 and 0.011 for P2. A possible mechanism responsible for the higher amount of Ag and the smaller amount of N in P2 sample compared to P1 is: un-reduced silver from CS/AgNPs solution can easily direct bond to the PLLA surface containing functional groups, without the aid of chitosan. On the other hand, if the larger chitosan molecule does not find sufficient functional groups with which to form bonds at the PLLA surface, it is removed by washing after treatment.

Table 2. EDAX results of PLLA studied samples.

Element (%)	CK	OK	NK	AgL
P0	70.49	29.51	-	-
P1	71.73	14.36	13.46	00.45
P2	72.43	14.61	12.14	00.82

3.6. Proliferation and Morphology of the MC3T3-E1 Cells on the PLLA Samples

Clinical biomaterial applications require good biocompatibility of the material. The biocompatibility of neat PLLA and two step treated PLLA films can be primarily evaluated by utilizing MC3T3-E1 cell lines, to be used for applications such as bone tissue engineering.

Fluorescent staining was used to study cell density and morphology after culturing MC3T3-E1 cells on P0, P1 and P2 films for 48 h and 72 h, as shown in Figure 8. For plasma-treated samples P2 and P0, spherical and round cells can be observed. The growth of cells on the surface of P1 films was better and more rapid than that on the P0 and P2 samples (Figure 8h) after culturing for 72 h. Moreover, the cells incubated on P1 show a higher degree of fibroblast cell adhesion and proliferation and a well-preserved morphology, which was flat and fully spread.

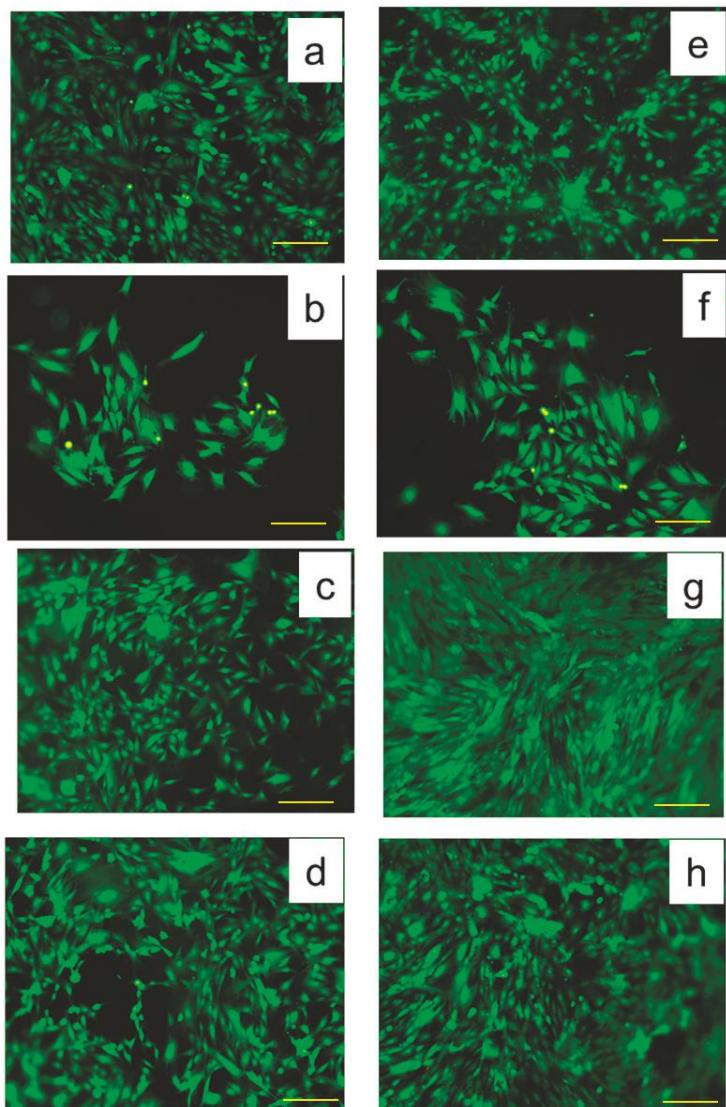


Figure 8. Cell culture (cell line MC3T3-E1) after 48 h for (a–d) and after 72 h (e–h); control (a,e), P0 (b,f), P1 (c,g), P2 (d,h). Scale: 200 μ m.

As a measure of unsaturated bond energy resulting from dangling bonds of surface material [42,43] surface energy is an important chemical cue on polymer surfaces. A polymer in contact with biological fluids has a surface energy which influences cell activities, such as serum protein adsorption and cell attachment. It was found that more fibroblasts can adhere and spread widely on the more hydrophilic polymer surface. The improvement or the suppression of cell adhesion at the polymer surface is in good concordance with high or respectively low surface free energy values [44,45]. In the case of equal surface free energy, a higher value for the polar component will induce a higher degree of cell adhesion and proliferation on the surface. The surface energy of the

PLLA surface was tailored by using plasma treatment and the surface free energy and the polar component of the P1 film was higher (Table 1) than that of the P0 and P2 films.

Two different cell morphologies were found on the studied polymer surfaces (Figure 8): (1) elongated cells well spread into polymer surface and (2) rounded cells, which are attached but have not begun to spread. Both types of morphologies are seen in different proportions on studied samples. The majority (>80%) of cells grown on PLLA films are elongated. On each of the P0, P1 and P2 substrates, up to 20% of cells were rounded. The highest density of elongated cells was found at P1 treatment after 72 h.

Figure 9 shows the MTT assay results of the MC3T3-E1 cell lines on P0, P1 and P2 films on 48 h and 72 h. The number of cells in each group increased with culture time on all of the tested groups. The MC3T3-E1 cells cultured on all samples have similar proliferation on the first day compared with the control. From the first day, the viable cell numbers on P1 were higher than those on P0 and P2.

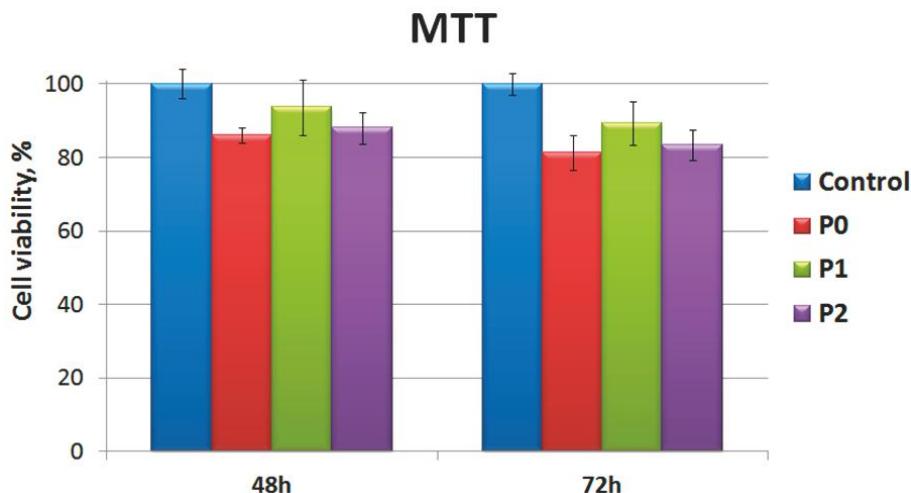


Figure 9. MTT assay results of MC3T3-E1 cells on the P0, P1 and P2.

The PLLA/Chitosan based-silver nanocomposite scaffolds appeared to be in vitro biocompatible and noncytotoxic to cells (Figure 9). The higher density of cells on P1 samples can be attributed to a higher concentration of N and a lower Ag concentration at film surfaces (Table 2 EDAX measurements). The shape, dimensions and low concentration of the silver ions prove to be nontoxic for the MC3T3-E1 cells culture. Moreover, the highest concentration of N and the highest polar component obtained for P1 samples assures better biocompatibility of P1 samples even compared to the neat PLLA.

3.7. Antimicrobial Activity

Strains adherence of *S. aureus* and *P. aeruginosa* (ATCC, clinical isolates) to polylactic acid films were studied. Figure 10 demonstrates the very good antimicrobial proprieties of P1 and P2 samples compared to the untreated P0 sample. The antimicrobial behavior of P1 sample is more pronounced in the case of *P. aeruginosa* compared to *S. aureus* and can be explained considering that both chitosan and AgNp have a bactericidal effect.

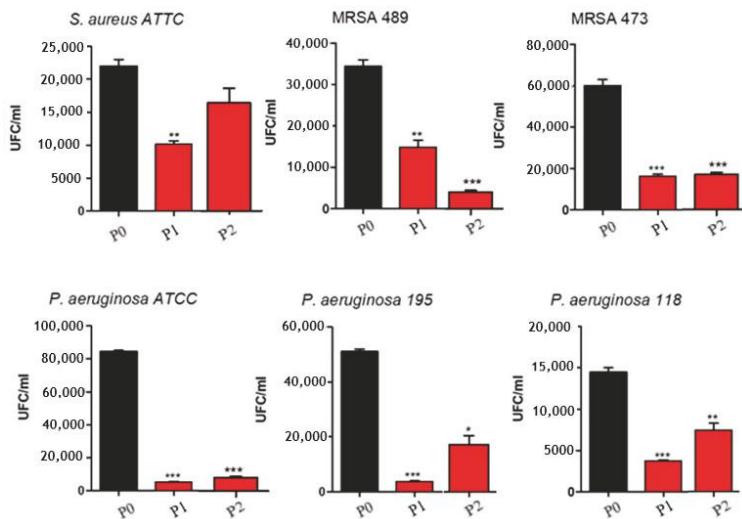


Figure 10. Strains adherence of *S. aureus* and *P. aeruginosa* (ATCC, clinical isolates) to polylactic acid films; *, ** and *** correspond to the magnitude of antimicrobial activity.

One of the proposed mechanisms of Ag action was based on AgNp capability to easily enter into the bacterial cell and form a less dense region in the center of the bacteria, causing the cell death by interacting with thiol containing enzymes [46,47]. Another proposed mechanism involves disruption of DNA/RNA caused by Ag reaction with the weak acid groups in the genetic material, such as phosphate [48].

On the other hand, chitosan has antimicrobial behavior and the main mechanism is based on the electrostatic interaction between positively charged chitosan groups and negatively charged sites on microbial cell [49]. In concordance with other authors [50], Figure 10 demonstrated that chitosan has stronger influence on Gram negative than on Gram-positive strains because the cell wall of *P. aeruginosa* (Gram-negative) has a thickness of 7–8 nm while the wall of *S. aureus* (Gram-positive) is around 20–80 nm [51].

4. Conclusions

In this paper, an environmentally-friendly synthesis of metallic nanoparticles in the presence of chitosan was performed. Silver ions underwent coordination and reduction thanks to the presence of numerous amino and hydroxyl groups in the chitosan chains. Bounding of silver nanoparticles of 30 nm average diameter to the polymer functional groups ensured a long-term stability and prevented their agglomeration. FTIR data pointed out the possible interactions of the hydroxyl or amino groups of chitosan and the carboxyl groups of PLLA. The silver nanoparticles were successfully adsorbed on PLLA films exposed to plasma treatments, by simply immersing the treated films in the chitosan solution containing silver nanoparticles. In this way, chitosan was used to fix silver nanoparticles on PLLA films surfaces. This is a time saving, inexpensive and eco-friendly synthesis that minimizes the use of toxic chemicals and does not produce toxic waste.

The biopolymer-based nanocomposite scaffolds with bioactive inorganic phases are of high interest due to their biocompatibility in combination with preosteoblastic cell line MC3T3-E1 (established from mouse C57BL/6 calvaria) seeding. The sample, which has demonstrated from EDAX to have the highest amount of nitrogen and the smallest amount of Ag, proved to have the highest value for cell viability. Moreover, it demonstrated better biocompatibility and very good antimicrobial activity against gram-negative and gram-positive bacteria. The effective component for

the biocompatibility seemed to be both PLLA and chitosan, while for the antimicrobial property both chitosan and silver were responsible.

The described two-step method is a promising technology for obtaining: poly(L-lactic acid) for tissue engineering applications like bone regeneration. In this direction, the new approach of biopolymer-polysaccharides based composite enables the scaffold surface to mimic complex local biological functions. To target clinical and medical applications, the need for additional investigations in the biological system is imperative.

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Article

Elucidating the Chemistry behind the Reduction of Graphene Oxide Using a Green Approach with Polydopamine

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Abstract: A new approach using X-ray photoelectron spectroscopy (XPS) was employed to give insight into the reduction of graphene oxide (GO) using a green approach with polydopamine (PDA). In this approach, the number of carbon atoms bonded to OH and to nitrogen in PDA is considered and compared to the total intensity of the signal resulting from OH groups in polydopamine-reduced graphene oxide (PDA-GO) to show the reduction. For this purpose, GO and PDA-GO with different times of reduction were prepared and characterized by Raman Spectroscopy and XPS. The PDA layer was removed to prepare reduced graphene oxide (RGO) and the effect of all chemical treatments on the thermal and electrical properties of the materials was studied. The results show that the complete reduction of the OH groups in GO occurred after 180 min of reaction. It was also concluded that Raman spectroscopy is not well suited to determine if the reduction and restoration of the sp^2 structure occurred. Moreover, a significant change in the thermal stability was not observed with the chemical treatments. Finally, the electrical powder conductivity decreased after reduction with PDA, increasing again after its removal.

Keywords: graphene oxide; reduced graphene oxide; X-ray photoelectron spectroscopy; Raman spectroscopy; electrical conductivity; functionalization

1. Introduction

Graphene, a 2D monolayer of sp^2 -hybridized carbon atoms arranged in a hexagonal lattice with a carbon–carbon bond length of 0.142 nm, has been extensively studied since it was first isolated in 2004 by Novoselov et al. [1,2]. It has a great potential for several applications due to its Young's modulus of 1 TPa, intrinsic strength of 130 GPa [3], room temperature (RT) electron mobility of 250,000 $cm^2 V^{-1} s^{-1}$ [4], and optical transmittance of 97.7% [5]. The promising application areas for graphene are photonics, optoelectronics, energy generation and storage, sensors for gas detection, reinforcement of composite materials, and biomedical areas, particularly in biosensing, drug and gene delivery, and tissue engineering [4].

Several techniques have been reported to produce graphene such as mechanical and electrochemical exfoliation of graphite, chemical vapor deposition, plasma enhanced chemical vapor deposition, thermal decomposition on silicon carbide (SiC), and reduction of graphene oxide (GO) [6]. Mechanical exfoliation of graphite can be either performed through atomic force microscope (AFM) probe techniques or adhesive tape exfoliation and results in high quality graphene. However, the high production cost makes this technique only feasible for research purposes [6,7]. Chemical vapor deposition is the most used technique for large-scale production of single or few-layer graphene [6,8]. Chemical vapor deposition is an expensive process due to the large energy consumption and the necessity of removing the substrate. Besides, controlling the grain size and the number of graphene layers produced is still a

challenge [8]. Graphene produced by this technique has high quality when the processing parameters are properly controlled. The main drawbacks are the high cost of SiC wafers and the high temperatures involved in the process (around 1200 °C) [4,8].

Reduction of GO appears to be another viable route to produce single-layer graphene [9]. GO can be prepared from graphite through various methods such as those reported by Brodie, Staudenmeier and Hummers [10], either following their original protocol or introducing some variations. The oxidation process of graphite introduces hydroxyl, epoxy, carbonyl, and carboxyl groups in the hexagonal lattice of graphene [8]. Then, the reduction of GO can be performed, for instance, through a chemical approach to remove the oxygen-containing functional groups and restore the conjugated graphene structure [11]. Compared to others, this process allows the production of large quantities of graphene at a low cost since no special equipment or high temperatures are needed and the starting materials, graphite and chemical reductants, such as hydrazine and sodium borohydride, are usually cost-effective [12].

The major drawback of the traditional routes for chemical reduction of GO is the use of toxic and hazardous chemicals both to living organisms and to the environment [12]. Thus, special care with the handling of these chemicals must be taken and, at an industrial scale, the remediation of the hazardous wastes generated might result in a substantial increase of the production costs. In addition, if toxic residues are still present in the final material, applications in the biological and biomedical fields are unsuitable [13].

Regarding the reduction of graphene oxide, an improvement in the use of green chemistry [14] has been observed and several potential environmental-friendly chemicals have been studied. Among these, vitamins [15–17], saccharides [18], amino acids [19–21], organic acids [22–25], microorganisms [26,27], proteins and peptides [28,29], hormones [30], urea [31], and plant extracts [32–35] have been tested as reducing agents for GO. Dopamine, a nature-based, commercially available, and inexpensive reagent, was first employed for this purpose in 2010 by Xu et al. [36]. Nevertheless, a substantive interpretation of data showing the successful reduction of GO at a molecular lever was not provided. In the presence of the oxygen functional groups in GO and at a weak alkaline pH, dopamine self-polymerizes to form polydopamine (PDA) with the catechol groups undergoing oxidation to form quinone groups. Thus, the polymerization of dopamine at the surface of GO is accompanied by the reduction of the last with dopamine acting both as a reducing and functionalization agent [36]. The presence of PDA on the surface of reduced graphene oxide (RGO) allows the preparation of more stable dispersions, compared with RGO prepared using other reducing agents, which can be a crucial factor for further processing through liquid assisted techniques. Furthermore, PDA strongly adheres to a wide range of substrates, which makes it a good material for applications in functional coatings [37]. Since 2010, several works have been published on the production of PDA-GO envisaging applications such as water purification [38], anion and proton exchange membrane fuel cells [39], functional coatings, and biomedical applications, for instance cancer treatment [40], drug delivery [41], antibacterial materials [42], biosensing [43], and tissue engineering [44]. However, these works focus mainly on the use of PDA-GO as a platform for anchoring of nanoparticles or covalent grafting of other molecules and on the characterization of the materials prepared envisaging the final application and lack deep insight into the mechanism of the GO reduction by PDA.

Considering the limited knowledge on this topic, in this work, we propose a new approach using X-Ray photoelectron spectroscopy (XPS) to supply evidence for the molecular reduction of GO by PDA. This reduction process is a good example of the use of green chemistry to replace traditional methods since no toxic solvents are used and PDA is a natural and renewable raw material, and only a little amount of waste is generated. Our investigation contributes to the field of conductive composite material development based on the strategy to use inexpensive and easily available graphite as basic raw material. In addition, the removal of the PDA to obtain RGO is studied. For this purpose, GO, PDA-GO, and RGO were prepared and characterized by Raman spectroscopy, XPS, thermogravimetric analysis (TGA), and electrical powder conductivity measurements. Thus, the effect of the chemical

treatments in the thermal and electrical properties of GO and PDA-GO was studied and a better understanding of the chemistry behind the green reduction of GO with PDA is presented.

2. Materials and Methods

2.1. Materials

Natural graphite flakes (99%; –325 mesh), tris(hydroxymethyl)aminomethane (tris base, ≥99.8%), dopamine hydrochloride (DA, 98%), and sodium hydroxide (NaOH, >97%) were purchased from Sigma Aldrich® (Munich, Germany). Sulfuric acid (H_2SO_4 , 98%), potassium permanganate (KMnO_4 , ≥99%), and hydrochloric acid (HCl, 10% v/v) were received from VRW® (Darmstadt, Germany). Hydrogen peroxide (H_2O_2 , 30% w/v) was purchased from Merck (Darmstadt, Germany). Distilled water (DW) was used in all chemical treatments.

2.2. Synthesis of Graphene Oxide, Polydopamine-Reduced Graphene Oxide, and Reduced Graphene Oxide

Graphene oxide (GO) powders were prepared through a modified Hummers' method according to our previous publication [45]. The oxidation step was performed using H_2SO_4 and KMnO_4 (graphite: KMnO_4 = 1:1). DW, H_2O_2 , and HCl were used for the purification step.

For the synthesis of PDA-GO, first a tris base solution in DW (0.1 M) was prepared and degassed by N_2 bubbling during 20 min. Then, 40 mg of GO were placed in a round bottomed flask and degassed for 10 min with N_2 . After that, 20 mL of tris base were added, and the suspension was magnetically stirred for 15 min at 60 °C. After stirring, DA was added to the GO suspension (GO:DA = 1 w/w) and the suspension was degassed by N_2 bubbling during 10 min. The reactions between GO and DA took place at 60 °C for the times indicated in Table 1. After the reaction, the suspension was vacuum filtered with a nylon filter membrane (0.45 µm pore size, Whatman, Kent, UK). Finally, the powders were collected and dried overnight in vacuum at 60 °C.

Table 1. Time of reaction between dopamine hydrochloride (DA) and graphene oxide (GO) and respective sample description.

Time of Reaction (min)	Sample Description (PDA-GO_reduction Time in min)
30	PDA-GO_30
60	PDA-GO_60
90	PDA-GO_90
120	PDA-GO_120
150	PDA-GO_150
180	PDA-GO_180
210	PDA-GO_210

Reduced graphene oxide (RGO) was prepared from PDA-GO. First, 20 mg of PDA-GO_30, PDA-GO_90, and PDA-GO_180 were stirred each in a 40 mL of NaOH solution (5 M), under N_2 flow for 6 h. Then, the suspension was vacuum filtered with a nylon filter membrane (0.45 µm pore size, Whatman, Kent, UK), and the powders collected and dried at 60 °C under vacuum, overnight, to obtain RGO_30, RGO_90, and RGO_180.

2.3. Modeling and Characterization of GO, PDA-GO, and RGO

2.3.1. Ab-Initio Calculations

Ab-initio calculations were performed to get information about an optimized geometry of GO by minimizing the corresponding Hartree–Fock energy applying the software package GAMESS (freeware version 2017-09-30R2, Iowa State University, Ames, IA, USA) [46] using the basis set STO-6G. For this purpose, first, a graphene-like structure consisting of four rows of heptacene was considered. Then, as a first oxidation step, the formation of external hydroxyl and carboxyl groups was simulated following

the production of oxirane groups in the plane of the graphene-like molecule. Finally, the oxirane rings opening and conversion into hydroxyl groups, due to the acidic medium where the reaction takes place, was calculated.

2.3.2. Raman Spectroscopy

Raman Spectroscopy was performed on a Raman microscope alpha300R (WITec, Ulm, Germany) using a laser excitation wavelength of 532 nm, laser power of 1 mW, a spectral resolution of 6 cm^{-1} , and integration time of 0.5 s. Two hundred scans were accumulated to record each spectrum. For this purpose, GO, PDA-GO, and RGO aqueous suspensions were prepared and sprayed on a glass slide positioned on a heating plate for fast water evaporation and deposition of the graphene products to be analyzed. The I_D/I_G or $I_{D+\text{PDA}}/I_{G+\text{PDA}}$ ratios were calculated considering the area under curve of the bands, which were estimated with a mean error of about 10%.

2.3.3. X-Ray Photoelectron Spectroscopy (XPS)

XPS studies were carried out by means of an Axis Ultra photoelectron spectrometer (Kratos Analytical, Manchester, UK). The spectrometer was equipped with a monochromatic Al K α ($\hbar\nu = 1486.6\text{ eV}$), X-ray source of 300 W at 15 kV. The kinetic energy of photoelectrons was determined with hemispheric analyzer set to pass energy of 160 eV for wide-scan spectra and 20 eV for high-resolution spectra. For the C 1s region, the maximum information depth of the XPS method is about 8 nm [47,48]. Employing Scotch double-sided adhesive tape (3M Company, Maplewood, MN, USA), the powdery samples were prepared as thick films on a sample holder. During all measurements, electrostatic charging of the sample was avoided by means of a low-energy electron source working in combination with a magnetic immersion lens. Later, all recorded peaks were shifted by the same value that was necessary to set the component peak G_r showing the sp²-hybridized carbon atoms of the graphite-like lattice ($-\text{C}=\text{C}- \leftrightarrow \text{=C-C=}$) to 283.99 eV [49]. In the case of considerable amounts of saturated hydrocarbons (PDA-GO samples), their corresponding component peaks in the C 1s spectrum was used as a reference with a binding energy of 285.00 eV [50]. Quantitative elemental compositions were determined from peak areas using experimentally determined sensitivity factors and the spectrometer transmission function. The shapes of the high-resolution element spectra were used to analyze the different binding states of the elements. For this purpose, the high-resolution element spectra were deconvoluted into component peaks (Kratos spectra deconvolution software, software version 2.2.9, Kratos Analytical Ltd., Manchester, UK), in which their binding energy values (BE), height, full width at half maximum and the Gaussian–Lorentzian ratios were free parameters.

2.3.4. Thermogravimetric Analysis (TGA)

TGA measurements were carried out on a TGA Q 5000 (TA Instruments Inc., New Castle, DE, USA) under N₂ atmosphere between 40 to 800 °C and heating rate of 10 K min⁻¹. Prior to the measurements, the powders, except graphite, were dried at 100 °C for 12 min using the same equipment.

2.3.5. Electrical Powder Conductivity

The electrical conductivity of the powders was measured using the equipment shown in Figure 1, which was developed and constructed at the Leibniz Institute of Polymer Research Dresden. The equipment consists of a transparent cylinder of 40 mm length with a capillary hole with diameter of 5 mm, which is mounted on a gold electrode on its bottom. After a certain amount of powder (~25 mg) was filled inside the hole, the upper movable cylindrical gold electrode with the same diameter compresses the material stepwise up to a pressure of 30 MPa using a stepper motor. The resistance was measured between the two gold electrodes using a Keithley 2001 electrometer (Tektronix, Köln, Germany) and the conductivity was calculated [51]. At least three measurements were performed to get mean values and standard deviations. In addition, based on the weighed powder mass and the volume of the sample, given by the geometrical conditions of the cylinder and the position of the

stepper motor, the bulk density and its development at different pressures can be determined. This also gives a measure of the compressibility of the powder.



Figure 1. Equipment used for electrical powder conductivity measurements (photo by Dr. Wolfgang Jenschke, Leibniz Institute of Polymer Research Dresden, Dresden, Germany).

3. Results and Discussion

3.1. The Oxidation of Graphene Stacks

Graphene is an allotrope form of carbon with a 2D honeycomb structure. According to the sp^2 orbital hybridization of all its carbon atoms, it can be considered as an infinitely large polycyclic aromatic molecule. While the electrons in the s , p_x and p_y orbitals form σ -bonds, the p_z electrons are involved in conjugated π -bonds hybridizing to π -band and π^* -bands. The half-filled π -band covering the whole molecule permits free-moving electrons, which are responsible for the graphene's electrical conductivity. In fact, the structure of a real graphene sample must be limited. Figure 2a shows a graphene-like model molecule consisting of four rows of heptacene. The degree of oxidation of all these graphene-like carbon atoms can be theoretically given to an average number of -0.307 . To understand the oxidation of graphene, in a model based on the Lefr-Klinowski model [52,53], the edges of a graphene sample were decorated with hydrogen atoms to avoid a geometric disorder and keep the aromaticity and electrical conductivity of the molecule [54,55].

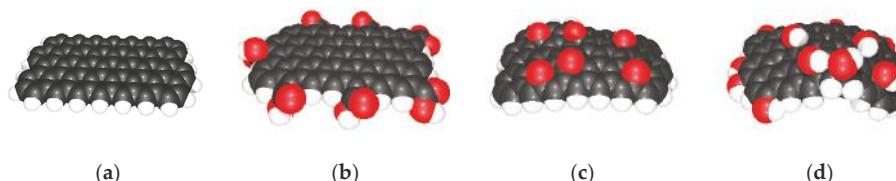


Figure 2. Results of ab-initio calculations to model the geometries of graphene-like and GO-like molecules: (a) graphene-like structure consisting of four rows of heptacene; (b) oxidized graphene carrying sterically demanding carboxylic groups on the edge; (c) oxidative attack on the carbon atoms in the molecular plane formed oxiran groups instable in acidic media; and (d) GO-like molecule decorated with hydroxyl groups on the edge and in the molecular plane.

As can be seen in Figure 2a, the optimized hydrogen-decorated graphene-like model molecule is flat and planar. With increasing degree of oxidation (exchange of 7 external H atoms with 7 carboxyl groups increases the average oxidation number of carbon atoms to 0.047), the original honeycomb

structure was disturbed. At first, external hydroxyl and carboxyl groups (probably also other carbonyl groups, such as quinone-like groups, which were not studied here) slightly deformed the planarity of the graphene-like molecule (Figure 2b). The oxidative attack on carbon atoms in the plane of the graphene-like molecule may produce oxirane groups (Figure 2c). However, due to the strongly acidic medium where the oxidation of graphene took place, the oxiran rings were opened immediately and converted into hydroxyl groups, as shown in Figure 2d. Regardless of whether hydroxyl or oxiran groups were formed in the molecular plane, the former flat plane bended up and needed more space. This effect is often used to separate single layers of graphene from their stacks. The combination of hydroxyl groups with the replacement of edge-standing hydrogens by carboxyl groups increases the average degree of oxidation of carbon atoms to 0.176.

3.2. Raman Spectroscopy

Raman Spectroscopy was performed to evaluate the level of “disorder” of the sp^2 hybridized structure of the materials prepared. The Raman spectra, presented in Figure 3, show the three major bands characteristic of sp^2 carbon materials. The D band, near 1350 cm^{-1} , is related to the presence of structural defects in the hexagonal sp^2 carbon lattice of graphene and to edge effects [56,57]. The G band, at approximately 1580 cm^{-1} , is related to the in-plane vibration of the sp^2 carbon atoms [56]. The band near 2700 cm^{-1} , the 2D band, originates on a second-order Raman scattering process and its shape, width, and position is related to the number of layers for n -layer graphene. Ferrari et al. reported that an increase in the number of layers originates a broader 2D band shifted to higher Raman shifts. [57,58]. The relative signal intensity of the D band to the G band (I_D/I_G) provides information about the level of “disorder” in terms of covalent modification of the graphene structure [57,59]. In pristine graphite, the 2D band consists of two components and appears at $\approx 2720\text{ cm}^{-1}$ while graphene presents a single sharp peak centered at a Raman shift lower than 2700 cm^{-1} [59].

Regarding PDA, there are two bands at about 1358 and 1588 cm^{-1} [45] that are assigned to the stretching vibration and deformation of chatecol groups [60]. Thus, the D and G typical for the graphene derivatives overlap with the PDA and a proper assignment of these bands is not possible. As long as PDA is present at the GO surface, contributions of these bands have to be considered. Thus, particularly the band at 1350 – 1358 cm^{-1} presents an overlap of the signals coming from the structural defects of GO, functional groups on the surface, and the amount of PDA in the material. Nevertheless, the ratios between the intensities at about 1350 and 1580 cm^{-1} were calculated and referred to as $I_{D+\text{PDA}}/I_{G+\text{PDA}}$ for the PDA containing samples.

The Raman spectra in Figure 3a show an increase of the I_D/I_G ratio when graphite was chemically converted in GO, from 0.2 to 0.5. This is a consequence of an increase of the content of structural defects caused by the introduction of oxygen-containing functional groups during the oxidation process or by the decrease in the flake size when GO was exposed to sonication [61]. Moreover, the downshift to 2717 cm^{-1} after oxidation indicates a reduction in the number of graphene stacks.

After reduction with PDA, a further increase in $I_{D+\text{PDA}}/I_{G+\text{PDA}}$ was observed, as shown in Table 2. Raman spectroscopy is a standard method to evaluate the molecular structure of graphene and graphene derivatives. However, considering the overlapping bands of GO and PDA, in this case, the method is not well suited to determine if the reduction and restoration of the sp^2 structure occurred since the typical I_D/I_G ratios also have a contribution of the PDA signal. Nevertheless, the position of the 2D band shows a further reduction on the number of graphene stacks on the PDA-GO materials.

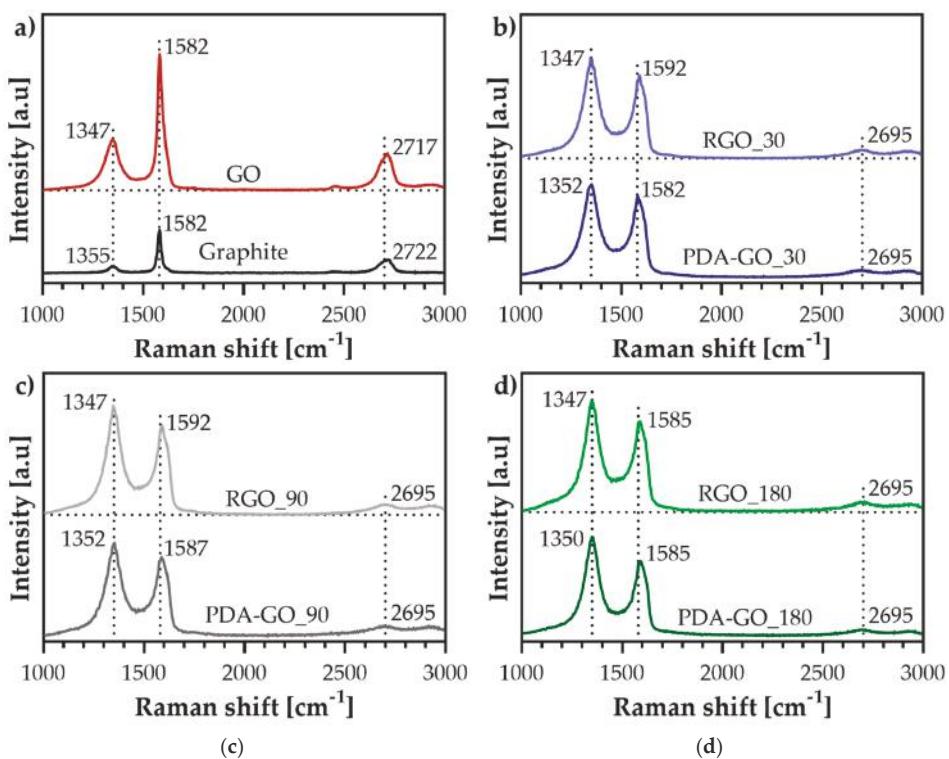


Figure 3. Raman spectra and position of the D , G , and 2D bands of: (a) graphite and GO; (b) PDA-GO_30 and RGO_30; (c) PDA-GO_90 and RGO_90; and (d) PDA-GO_180 and RGO_180.

Table 2. I_D/I_G or I_{D+PDA}/I_{G+PDA} of the graphene derivatives.

Material	I_D/I_G	I_{D+PDA}/I_{G+PDA}
Graphite	0.2	—
GO	0.5	—
PDA-GO_30	—	1.2
PDA-GO_90	—	1.2
PDA-GO_180	—	1.2
RGO_30	—	1.2
RGO_90	—	1.2
RGO_180	—	1.2

The evaluation of the structure of RGO should be possible since no contribution from PDA peaks is expected. However, no significant changes were found between the spectra of RGO and the corresponding PDA-GO, which might indicate that the PDA was not completely removed after the treatment with NaOH. Thus, additional experiments are required to study the molecular structure of these materials by Raman Spectroscopy and to optimize the removal of the PDA layer. In addition, different methods, such as XPS, are required to properly characterize the molecular structure of PDA-GO.

3.3. X-Ray Photoelectron Spectroscopy (XPS)

3.3.1. XPS Spectra of Graphene Stacks and GO

The XPS studies corroborate the information obtained by Raman spectroscopy. The shape of the C 1s spectrum of graphite in Figure 4a is very characteristic for carbonaceous substances consisting of sp^2 hybridized lattices. Due to the numerous excited states, the spectrum noticeably tailed on the high energy side. The main component peak *Gr* found at 283.99 eV results from photoelectrons that escaped from sp^2 -hybridized carbon atoms ($-C=C-\leftrightarrow=C-C-$). Photoelectrons removed from molecules with excited electron states appear as wide shake-up peaks (gray lines in Figure 4). At 285.69 eV, a further small component peak *C* was observed. This component peak shows the presence of $\underline{C}-O$ bonds on the sample surface (detailed XPS data are presented in Tables S1–S18 of the Supplementary Materials).

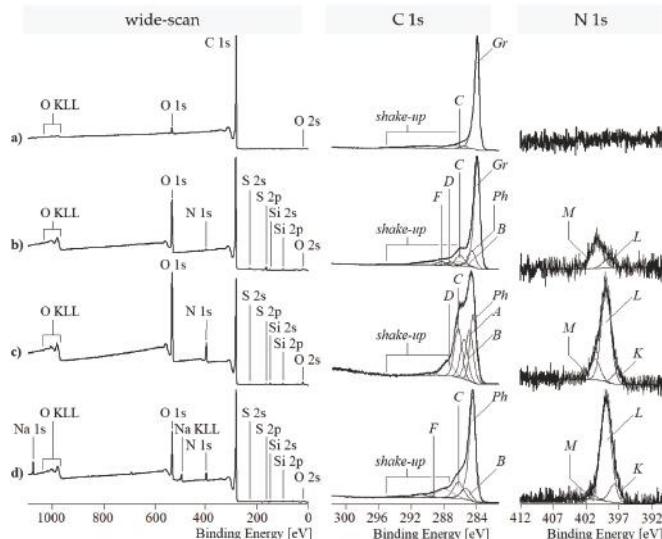


Figure 4. Wide-scan (left column), high-resolution C 1s (middle column) and N 1s (right column) XPS spectra recorded from: (a) graphene stacks; (b) GO; (c) PDA-GO_30; and (d) RGO_30.

The oxidation of graphite to GO significantly increases the relative oxygen content from 0.028 to 0.166 (Figure 4b). This increase results from the presence of numerous oxygen-containing functional groups on the sample surface. However, the C 1s spectrum (Figure 4b, middle column) shows that the oxidation reaction was gently performed. It seemed to be beneficial to minimize the destruction of the conjugated π -electron system by breaking σ -bonds during the formation of carboxylic acid groups. On the other hand, the planar graphene sheets should be sufficiently oxidized to bend them and expand the former stacks. The predominant preservation of the π -electron system becomes clear by the appearance of the intense component peak *Gr* and the shake-up peaks. It also seems necessary to introduce a component peak *Ph* (284.75 eV), indicating the presence of sp^2 -hybridized carbon atoms that are not involved in the highly conjugated π -system of the original graphene lattice [50]. Oxygen-containing functional groups were analyzed as component peaks *C* (286.07 eV), *D* (287.29 eV), and *F* (288.36 eV). Component peak *C* results from photoelectrons that escaped from phenolic $\underline{C}-OH$ groups. The binding energy value found for component peak *C* is also characteristic for ether groups ($\underline{C}-O-\underline{C}-$). However, as mentioned above, it is unlikely that oxirane groups are stable on the graphene surface. Component peak *D* shows carbonyl carbon atoms of quinone-like ($\underline{C}=O$) groups. Photoelectrons from the carbonyl carbon atoms of carboxylic acid groups ($O=C-OH$) and their corresponding carboxylates ($O=C-O^\ominus\leftrightarrow O-C=O$) were observed as the small component peak *F*. The presence of nitrogen ([N]:[C] = 0.009)

required the introduction of an additional small component peak *B* (at 285.72 eV) showing $\underline{\text{C}}-\text{N}$ bonds. These bonds can be constituents of amino and/or amide groups. In the case of the presence of amide groups, photoelectrons of the carbonyl carbon atoms ($\text{O}=\underline{\text{C}}-\text{NH}-\text{C}$) contributes to component peak *D* while component peak *B* presents the amine-sided carbon atoms ($\text{O}=\text{C}-\text{NH}-\underline{\text{C}}$).

3.3.2. XPS Spectra of PDA-GO

The application of dopamine and its subsequent polymerization on the GO surface significantly increased the relative nitrogen content from $[\text{N}]:[\text{C}]$ of the samples from 0.009 to ca. 0.090. The variation of the time for the polymerization reaction from 30 to 210 min does not correlate with the relative nitrogen content on the GO surfaces and thus with the amount of PDA deposited there. As can be seen in Figure 4c, the PDA layer strongly changes the shape of the C 1s spectra. Carbon atoms bonded in the phenyl rings of the PDA's catechol groups having no heteroatoms as binding partner and sp^2 -hybridized carbon atoms from the graphite-like lattice of the substrate materials were identified as component peak *Ph* (ca. 284.49 eV). Component peak *A* (285.00 eV) results from the presence of carbon atoms in the sp^3 -hybrid state of saturated hydrocarbons. Some of these carbon atoms are constituents of the PDA layer, but the presence of component peak *A* can be also considered as a first hint of reduced carbon species on the sample surface. Carbon–nitrogen bonds were assigned as component peak *B* (ca. 285.72 eV). The intensities of the component peaks *B* equal the twice of the $[\text{N}]:[\text{C}]$ ratios determined from the corresponding wide-scan spectra. According to the structural formula of the PDA in Figure 5, the number of carbon atoms carrying phenolic OH groups ($\underline{\text{C}}-\text{OH}$) should equal the number of carbon atoms bonded to nitrogen.

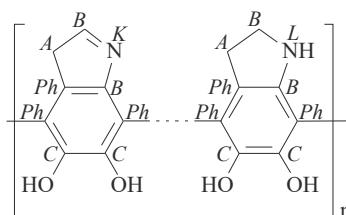


Figure 5. Characteristic cutouts from the chemical structure of PDA. Italic letters denote the assignment of the carbon and nitrogen atoms to the component peaks in the C 1s and N 1s high-resolution spectra.

However, the C 1s spectra recorded from the samples that reacted for 30 and 90 min showed component peaks *C* with higher intensities than their corresponding component peaks *B*, as shown in Figure 6a. It is assumed that the excess ($[\text{C}'] = [\text{C}] - [\text{B}]$) of the intensities of the component peaks *C* result from the contribution of $\underline{\text{C}}-\text{OH}$ groups from the GO substrate material. Quinone-like groups cannot be safely detected because component peak *D* is overlapped by intense shake-up peaks.

After longer polymerization times, such as 180 and 210 min, the excess component peaks *C'* disappears completely (Figure 6b). Obviously, the reduction of the GO by the oxidative polymerization of the adsorbed dopamine molecules requires longer periods of reaction time. Figure 6b (right column) shows that the high-resolution N 1s spectra recorded from the PDA-GO samples is deconvoluted into the three component peaks *K* (ca. 398.69 eV), *L* (ca. 400.1 eV) and *M* (ca. 401.68 eV). Corresponding to the intensities of the component peaks *B* (in the C 1s spectra), which are the twice of the $[\text{N}]:[\text{C}]$ ratios and the binding energy values found, component peak *L* shows cyclic secondary amino groups ($\text{C}-\underline{\text{NH}}-\text{C}$), such as the pyrrolidine structures in Figure 5 on the right side [50,62]. The binding energy values of the component peaks *K* are unusually small for organically bonded nitrogen. The observation of such low binding energy values indicates a high electron density at the nitrogen atom, which is characteristic for the cyclic imide nitrogen atoms ($\text{C}-\underline{\text{N}}=\text{C}$) exemplary shown in Figure 5 (left). Protonated nitrogen species ($\text{C}-\underline{\text{N}}^+\text{H}$) reflecting the protonation/deprotonation equilibrium of Brønsted basic nitrogen species are observed as component peak *M*.

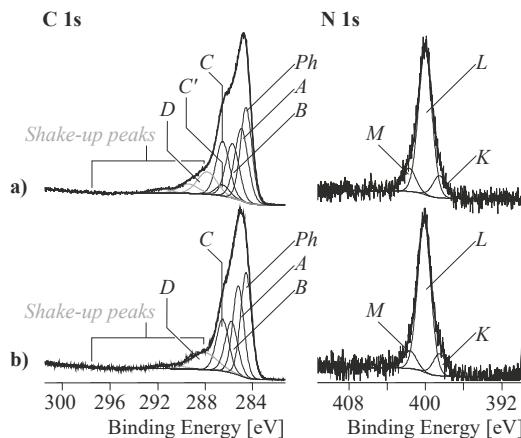


Figure 6. High-resolution C 1s (left) and N 1s (right) XPS spectra recorded from: (a) PDA-GO_90; and (b) PDA-GO_180.

3.3.3. XPS Spectra of RGO

NaOH was employed to hydrolyze and remove the PDA layer covering the carbonaceous substrate materials. However, the wide-scan spectra shows the presence of considerable amounts of nitrogen ($[N]:[C] \approx 0.04$) after treating the samples with NaOH (Figure 4d, left column). These findings can be considered as a hint that PDA was not completely removed from the substrate materials. Nevertheless, the corresponding C 1s spectra in Figure 4d exemplary shows the C 1s spectrum of sample RGO_30 with clearly intensity-reduced shoulders in the region of 286.5 eV and tailings on the high energy sides, which were previously observed in the C 1s spectrum of the graphene reference sample (Figure 4a, middle column). Corresponding to the findings of the PDA-coated samples, component peak B (285.34 eV) appears with an intensity that is twice the $[N]:[C]$ ratio found in the wide-scan spectrum of sample RGO_30. An equal number of photoelectrons escapes from the PDA's catecholic $\underline{C}-\text{OH}$ groups and contribute to component peak C at 286.29 eV. The excess of component peak C ($[C'] = 0.061$) has to be assigned to the $\underline{C}-\text{O}$ groups remaining on the surface of the substrate material. Probably, component peak D disappears completely while small traces of carboxylic acid groups ($\text{O}=\underline{C}-\text{OH}$) and/or carboxylic ester groups ($\text{O}=\underline{C}-\text{O}-\text{C}$) were observed as component peak F (289.36 eV). Latter groups could be formed by esterification of the carboxylic acid groups with some of the PDA's catechol groups. Photoelectrons of the corresponding alcohol-sited carbon atoms ($\text{O}=\underline{C}-\text{O}-\underline{C}$) contribute to component peak C. As discussed above, the N 1s spectrum was deconvoluted into the three component peaks K, L and M.

In summary, in the case of organic materials, it is often very difficult to determine or estimate the contribution of surface contaminations to the spectra. This is especially true for very complex-shaped high-resolution spectra, which were recorded for the samples studied here. C 1s spectra recorded from the graphene stacks, GO, and RGO_30 provide no information about the degrees of contaminations. A component peak showing the presence of surface contaminations arise at the same peak position. A separation was not possible. If they are present and can be separated [63] the photoelectrons of their carbon atoms would contribute to component peak observable at about 285 eV.

As shown in Figure 4, we found slightly different binding energy values for the component peaks C. The reason of these differences was the different chemical character of the $\text{C}-\text{OH}$ groups. In the case of the PDA-GO samples (PDA-GO_30, PDA-GO_90, and PDA-GO_180), $\text{C}-\text{OH}$ groups of the di-phenolic catechol units mainly contributed to the component peaks C (about 286.47 eV). After their removal with NaOH (sample RGO_30), a few residual di-phenolic catechol units and numerous OH groups from the not fully reduced substrate material contributed to component peak C. The binding

energy value was slightly lowered (286.29 eV). The phenolic OH groups on the GO, which were more intensively involved in the delocalized p-electron system of the (more or less disturbed by oxidation) graphite-like lattice, showed a binding energy value of 286.07 eV. The component peak C resulting from photoelectrons of traces of oxygen-carrying functional groups found on the surfaces of the graphene stacks had a binding energy value of 285.69 eV, which is significantly lower as usually expected for C–OH bonds. However, here we have to take into account that these few phenol groups are embedded in a largely undisturbed p-system, and - in contrast to the catechol rings - their dissociations ($\text{C}-\text{OH} + \text{H}_2\text{O}$ to $\text{C}-\text{O}^- + \text{H}_3\text{O}^+$) increasing the electron density at the carbon atoms are not hindered by negative charges in their immediate molecular neighborhood.

In all cases, the shapes of the component peaks were the result of convolutions of a Gaussian normal distribution and Cauchy–Lorentz distribution. With exception of the distribution of the component peak *Gr* and the component peak *Ph* in the RGO_30 sample, all other component peaks had the same shapes as suggested by the reviewer. High number of excited states in the well-ordered graphite-like lattice led to a tailing of the photoelectron distribution at the high-energy side of the component peak. Hence, it seemed to be necessary to adapt the line shape of the component peaks *Gr* by an increased asymmetry. In the case of the component peaks *Ph* in sample RGO_30, we have to consider that this component peak *Ph* summarized photoelectrons from the p-conjugated carbon atoms of the remaining PDA molecules and the carbon atoms of the more or less disturbed graphite-like structures of the substrate.

Furthermore, the XPS spectra recorded from the PDA-GO samples clearly show the reduction of the GO. As can be seen in the high-resolution C 1s spectrum, the reduction is mainly due to the decrease in the $\text{C}-\text{OH}$ groups of the former GO substrate. The majority of the $\text{C}-\text{OH}$ groups found in that spectrum are constituents of the PDA, which remain on the sample surface after washing with NaOH. The PDA wrapping the carbonaceous material can be used as stable anchor layer for subsequent modification reactions to functionalize and compatibilize carbon-based nanomaterials.

3.4. Thermogravimetric Analysis (TGA)

TGA was performed to evaluate the thermal stability of graphite, GO, PDA-GO, and RGO. It was expected that the amount of PDA on the GO surfaces could be evaluated from the TGA results. As shown in Figure 7, PDA has the lowest thermal stability, starting thermal degradation around 180 °C due to the partial decomposition of the main chain in three different stages with a maximum at around 260 °C [64]. The weight loss observed for graphite is very low and mainly due to the release of adsorbed water and residual oxidation [61]. Regarding GO, Figure 8 shows three loss stages. The first weight loss stage of ca. 1.7% occurring between 40 and 120 °C was attributed to the evaporation of adsorbed water [65]. The second weight loss (5.7%) has its maximum at about 220 °C and might be due to the decomposition of oxygen-containing functional groups ($\text{C}-\text{OH}$), whereas the main weight loss (8.8%) occurs at about 400 °C and is related to the loss of more stable oxygen functionalities such as carboxyl groups. The total weight losses of graphite, GO, and PDA at 800 °C are 0.7%, 19.1%, and 86.3%, respectively.

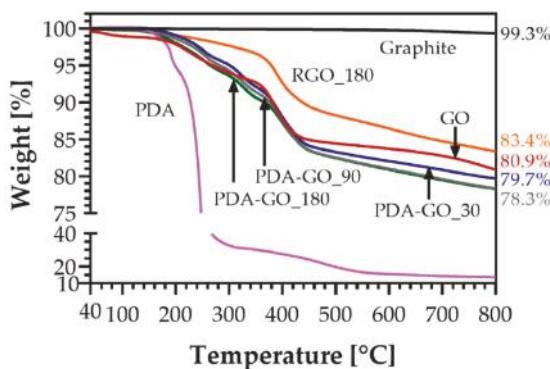


Figure 7. Thermogravimetric analysis (TGA) curves of graphite, GO, PDA, PDA-GO_30, PDA-GO_90, PDA-GO_180, and RGO_180.

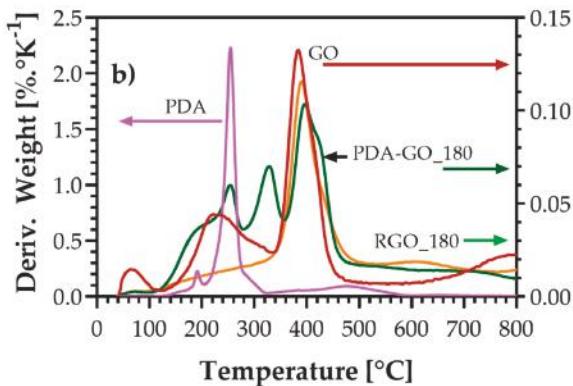


Figure 8. Derivative weight loss of GO, PDA, PDA-GO_180, and RGO_180.

The TGA curves of the PDA-GO materials show significant thermal degradation starting at around 180 °C due to the decomposition of PDA and the removal of non-reduced oxygen-containing functional groups from GO. The different polymerization times only slightly affected the thermal stability of the PDA-GO materials, which underwent a total weight loss of between 20.3% (PDA-GO_30) and 21.7% (PDA-GO_90 and PDA-GO_180) at 800 °C, as shown in Figure 7. However, a decrease in the thermal stability in comparison with GO was observed, possibly related to the increase of the oxygen content as determined by XPS due to the presence of PDA. Regarding RGO_180, the curve in Figure 8 shows only a weight loss stage with a maximum at around 400 °C, relative to the decomposition of stable oxygen-containing functional groups which were not reduced by PDA as determined by XPS. No peaks from PDA contribution were found throughout the temperature range analyzed; however, the small weight loss that occurs below 300 °C might be attributed to the decomposition of the residual PDA remaining after NaOH treatment. A total weight loss of 16.6% was observed, which reflects an improvement of the thermal stability after PDA removal. As calculated by XPS, RGO_180 possesses the lowest amount of oxygen species, which is reflected by its highest thermal stability among all prepared materials.

3.5. Electrical Powder Conductivity

The electrical powder conductivity of the graphene derivatives was measured to evaluate the effect of the reduction process on the restoration of the electrical properties. The electrical powder conductivity generally increases with the pressure applied during the measurements. As shown in

Figure 9, the starting material graphite has the highest electrical conductivity among all materials with values between 68.8 and 93.4 S/cm. As expected, after oxidation, there was a remarkable decrease in the conductivity of about 18% to 16.6 S/cm (at 30 MPa). This value of GO is higher than values reported previously when using the same measurement equipment but differently oxidized GO materials [66,67]. The reduction of GO with PDA resulted in a substantial decrease in the electrical powder conductivity, which scales with the reaction time between GO and DA and is most significant in PDA-GO_180 with a decrease to 1.9 S/cm at 30 MPa. Although the reduction process allowed the removal of oxygen-containing functionalities and potentially a restoration of the sp^2 -hybridized lattice of graphene, the electrical insulating nature of the PDA layer might contribute to this effect. By this, a decrease in the contact area between conductive graphene areas and an increase in contact resistance between the partially PDA covered conductive graphene occur, which justifies the decrease in the electrical conductivity. The conductivity of PDA-GO_180 increases with the pressure until 20 MPa, remaining constant afterwards. Regarding PDA-GO_30 and PDA-GO_90, similar electrical powder conductivity values were measured for all pressures considered. For PDA-GO_30, there was an increase of 48% from 1.8 to 3.7 S/cm while for PDA-GO_90 an increase of 44% from 1.6 to 3.6 S/cm was observed when increasing the pressure from 5 to 30 MPa.

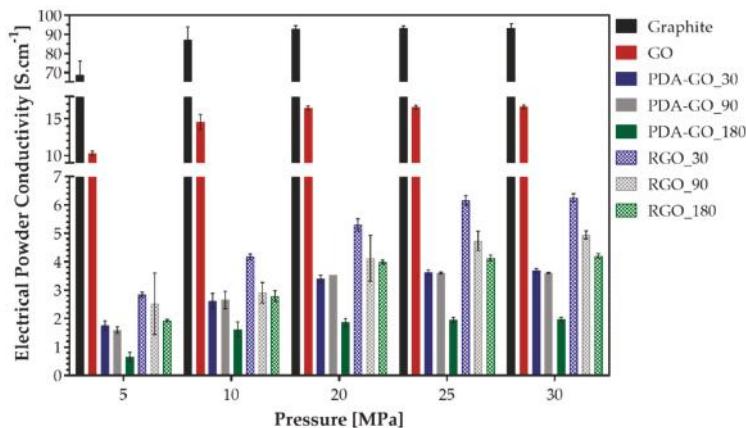


Figure 9. Electrical conductivity of the powdery materials at different pressures.

After the removal of PDA with NaOH, the electrical conductivity increases. When the PDA layer is removed, leaving only a small residue, the resultant RGO possesses a structure more suited for electrical conduction. Following the trend of decreasing conductivity with reaction time between GO and PDA, the samples after PDA removal also show this dependency. However, despite RGO₁₈₀ has the lowest conductivity among the RGO samples, the difference between PDA-GO and RGO is most pronounced in this sample which shows an increase of two times after the treatment with NaOH. The values achieved are higher than those reported for differently reduced graphite oxides measured using the same equipment in Ref. [66]. However, even after reduction and PDA removal, the desired higher electrical conductivity than that of GO was not obtained, possibly due to the presence of residual PDA, a more defected sp^2 carbon structure, or different compressibility of the powders. The representation of the electrical powder conductivity as a function of the bulk density of the various materials in Figure 10 confirms the significant differences between them. Graphite and GO have the highest bulk density values at the applied pressures, which illustrates a higher packing density and better contact between the graphene flakes inside the powder materials. Graphite has a much higher compressibility (especially at pressures between 5 and 20 MPa) than GO, which has the lowest increase in conductivity with pressure, showing the lowest compressibility of all samples. The PDA-GO and RGO powders show lower bulk densities at the initial pressure of 5 MPa and stronger density changes

at increasing pressure than graphite and GO. The increase in conductivity with pressure is slightly lower for the PDA_GO than for the RGO samples, demonstrating a slightly higher compressibility of the RGO. This may be a hint that the powder particles in the RGO samples are more flexible and less stiff. The dependencies in Figure 10 illustrate that the materials are in different packing states during the measurements, which affects the contact surfaces and the contact resistance between the material powder particles and flakes in the measured sample volume. If only very small amounts of PDA remain on the surface of graphene flakes, especially at locations where the flakes are in close contact during measurement, this may result in high contact resistance and reduced conductivity through the sample. The comparison of the electrical powder conductivity of all materials at the same density, selected here at 1.85 g cm^{-3} , however, confirms the sequence described above at different measuring pressures. Graphite has the highest conductivity, followed by GO, while PDA_GO has a lower conductivity than RGO.

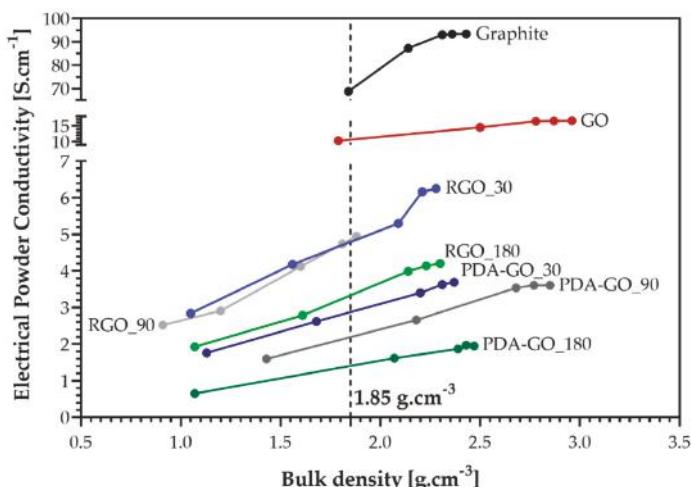


Figure 10. Dependency of electrical conductivity of the powdery materials on bulk density (density increases with pressure).

Nevertheless, despite the lower electrical powder conductivity, both PDA-GO and RGO materials possess higher conductivity values than reported for other examples in the literature [66] and show a better dispersibility in several solvents than GO, which is crucial when, for example, these materials are to be used in solvent assisted techniques to prepare conductive films.

4. Conclusions

In this work, we propose a new approach to investigate the molecular reduction state of GO by PDA and the removal of the PDA, using NaOH, to obtain RGO. It was shown that Raman spectroscopy is not well suited to determine the reduction and restoration of the sp^2 structure. However, a first hint for the presence of PDA, even after the treatment with NaOH, can be obtained using this method since no significant differences were found between the PDA-GO and RGO spectra. The reduction of GO by PDA was proven by XPS through a new approach that considers the number of carbon atoms bonded to OH and to nitrogen in PDA and compares it to the total intensity of the signal resulting from OH groups in PDA-GO to finally determine that the reduction occurs. In addition, it was shown that there was no complete removal of the PDA layer with NaOH, corroborating the Raman spectroscopy results. Regarding the thermal analysis, it was observed that the presence of PDA in PDA-GO results in a decrease in the thermal stability. However, after PDA removal, the thermal stability improved and revealed to be higher than in GO, which agrees with the XPS studies that showed RGO_180 possesses

the lowest amount of unstable oxygen-containing species. In addition, the graphene derivatives prepared in the present work possess considerably higher electrical powder conductivity values than those reported in the literature, even if the desired higher electrical conductivity than that of GO was not obtained. The small proportion of PDA remaining on the material surface can be used for subsequent functionalization, which often plays the key role for the intended application.

To summarize, in the present work, deep insight into the chemistry of PDA-GO and RGO was given. The green reduction of GO by PDA proved to be a way to replace typical reduction methods that involve toxic, corrosive, and hazardous solvents and chemicals. In the future, further studies are needed to better understand the reduction of GO by such green approaches and make the processes reproducible and scalable.

Supplementary Materials: The following are available online at <http://www.mdpi.com/2079-4991/9/6/902/s1>, Table S1: Elemental compositions investigated by XPS, Table S2: Wide-scan spectrum of the sample recorded from the graphene stacks, Table S3: C 1s spectrum of the sample recorded from the graphene stacks, Table S4: Wide-scan spectrum of the GO sample, Table S5: C 1s spectrum of the GO sample, Table S6: N 1s spectrum of the GO sample, Table S7: Wide-scan spectrum of the PDA-GO_30 sample, Table S8: C 1s spectrum of the PDA-GO_30 sample, Table S9: N 1s spectrum of the PDA-GO_30 sample, Table S10: Wide-scan spectrum of the RGO_30 sample, Table S11: C 1s spectrum of the RGO_30 sample, Table S12: N 1s spectrum of the RGO_30 sample, Table S13: Wide-scan spectrum of the PDA-GO_90 sample, Table S14: C 1s spectrum of the PDA-GO_90 sample, Table S15: N 1s spectrum of the PDA-GO_90 sample, Table S16: Wide-scan spectrum of the PDA-GO_180 sample, Table S17: C 1s spectrum of the PDA-GO_180 sample, Table S18: N 1s spectrum of the PDA-GO_180 sample.

Author Contributions: Sample preparation, electrical powder conductivity measurements, and data treatment and interpretation of Raman spectroscopy, TGA, and electrical powder conductivity measurements were performed by C.S. as well as writing. Ab initio calculations were performed by P.F. Conceptualization, reviewing, editing, project administration, and supervision were realized by C.Z. and P.P. who also interpreted the powder conductivity measurements. XPS was performed by F.S. including data treatment, interpretation and discussion.

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Conflicts of Interest: The authors declare no conflict of interest.

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Article

Development of Effective Lipase-Hybrid Nanoflowers Enriched with Carbon and Magnetic Nanomaterials for Biocatalytic Transformations

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Abstract: In the present study, hybrid nanoflowers (HNFs) based on copper (II) or manganese (II) ions were prepared by a simple method and used as nanosupports for the development of effective nanobiocatalysts through the immobilization of lipase B from *Pseudozyma antarctica*. The hybrid nanobiocatalysts were characterized by various techniques including scanning electron microscopy (SEM), energy dispersion spectroscopy (EDS), X-ray diffraction (XRD), Raman spectroscopy, and Fourier transform infrared spectroscopy (FTIR). The effect of the addition of carbon-based nanomaterials, namely graphene oxide and carbon nanotubes, as well as magnetic nanoparticles such as maghemite, on the structure, catalytic activity, and operational stability of the hybrid nanobiocatalysts was also investigated. In all cases, the addition of nanomaterials during the preparation of HNFs increased the catalytic activity and the operational stability of the immobilized biocatalyst. Lipase-based magnetic nanoflowers were effectively applied for the synthesis of tyrosol esters in non-aqueous media, such as organic solvents, ionic liquids, and environmental friendly deep eutectic solvents. In such media, the immobilized lipase preserved almost 100% of its initial activity after eight successive catalytic cycles, indicating that these hybrid magnetic nanoflowers can be applied for the development of efficient nanobiocatalytic systems.

Keywords: hybrid nanoflowers; lipase; magnetic nanomaterials; biocatalysis; enzyme immobilization

1. Introduction

Over the last decades, the immobilization of enzymes onto nanostructured supports has been extensively used and has facilitated their applications, owing to their easy handling and operational stability, as well as facile recovery and reusability of the biocatalysts, leading to more efficient bioprocesses [1,2]. Various nanostructured composite materials with extensive active surface areas and desirable pore sizes, such as nanoporous supports, nanofibers, nanoparticles, and carbon-based nanomaterials (e.g., nanotubes and graphene) have been proven to be effective in manipulating the nanoscale environment of biomolecules [3–5] and, as a consequence, their biological function and stability.

Organic-inorganic hybrid nanomaterials (nanoflowers) are a recently developed group of nanoparticles that schematically resemble plant flowers in a nanoscale range [6]. Hybrid nanoflowers

(HNFs) have attracted a lot of interest over the last years as host platforms for immobilizing enzymes, owing to the higher surface-to-volume ratio compared to spherical nanoparticles, as well as to their simple, eco-friendly, and cost-effective synthesis [7,8]. Nanoflowers containing various enzymes have been usually prepared as enzyme-Cu₃(PO₄)₂·3H₂O hybrids by combining copper sulfate (CuSO₄) with enzymes in phosphate-buffer saline (PBS). Different HNFs mainly based on copper (II) and calcium (II) ions have been used to form complexes with enzymes and other proteins [9–11]. Moreover, the development of protein-embedded HNFs based on other metal ions, such as zinc (II), cobalt (II), and iron (II), was recently reported [12–14]. The formation of HNFs comprises the following steps: the nucleation and formation of primary crystals, the growth of these aggregates, and the complete formulation of nanoflowers [8]. During the first step, protein molecules form complexes with metal ions, primarily through the coordination between nitrogen atoms of the amide groups present in the protein backbone and the metal ion. These complexes provide sites for nucleation. The intramolecular interactions between the metal ion and the protein promote the anisotropic growth of nano-petals (step 2) and, consequently, the formation of a flower-like structure in which proteins serve as the glue that binds the petals together (step 3). The formation of these enzyme-embedded HNFs do not require harsh conditions and toxic reactants for their self-assembly; thus, the immobilization procedure is facilitated with biomolecules in a one-step process. Moreover, the incorporated enzyme is subjected to minor manipulation in comparison with other conventional immobilization procedures, thus retaining its biocatalytic activity [8].

The selection of the enzyme and the metal ions—as well as the pH, the temperature, and the incubation time—plays an essential role for the configuration and the catalytic efficiency of the enzyme-containing nanoflowers [8,15]. A variety of enzymes with biotechnological interest have been encapsulated in HNFs and successfully applied in dye decolorization [11,16], the production of esters [17], the detection of phenols or glucose [9,18], the degradation of pollutants [19], and the development of biosensors [20,21] in which two or more enzymes were successfully encapsulated in the same nanoflower.

The enhanced activity of enzymes that is observed in various HNFs is mainly attributed to their high surface area, which decreases mass-transfer limitations, along with the specific interactions of the enzyme molecules and metal ions [8,22,23]. However, the biocatalytic activity and stability of some HNFs are reduced by the interactions between metal ions and proteins [24]. Recently, it was proposed that the incorporation of surfactants [25], biopolymers such as chitosan [26], and carbon-based nanomaterials [27,28] could enhance the catalytic properties as well as the mechanical strength of enzyme-containing nanoflowers, leading to stable nanohybrids.

Herein, we describe the preparation of novel hybrid nanoflowers consisting of copper (II) or manganese (II) ions, combined with magnetic nanoparticles and carbon-based nanomaterials, and we investigate their use as versatile host platforms for the development of sufficient systems for the immobilization of enzymes. The addition of carbon-based nanomaterials, namely graphene oxide and multi-walled carbon nanotubes, in the preparation of nanoflowers is expected to provide high surface area and extraordinary mechanical properties, whereas the incorporation of magnetic nanoparticles, such as maghemite, allows the easy and quick separation of the nanoflowers by the application of an external magnetic force. The use of these novel HNFs as host platforms for the immobilization of lipase B from *Pseudozyma antarctica*, an enzyme with numerous biotechnological applications, was investigated. The novel nanobiocatalysts were characterized by scanning electron microscopy (SEM), energy dispersion spectroscopy (EDS), X-ray diffraction (XRD), Raman spectroscopy, and Fourier transform infrared spectroscopy (FTIR), while the effect of the composition of nanoflowers on the catalytic activity, thermal activity, and operational stability of the immobilized enzyme was investigated. Moreover, the ability of the lipase-based nanoflowers to catalyze the synthesis of lipophilic derivatives of phenolic antioxidants, such as tyrosol, in non-aqueous media, as well as in environmental-friendly ionic solvents, was also investigated.

2. Materials and Methods

2.1. Materials

Lipase B from *Pseudozyma antarctica* (formerly *Candida antarctica*, CaLB) was purchased from Novozymes A/S (Bagsværd, Denmark) and was utilized without further purification. 4-nitrophenyl butyrate (*p*-NPB), 4-nitrophenol (*p*-NP), copper (II) sulfate pentahydrate, manganese (II) sulfate, tyrosol, and dimethyl sulfoxide were obtained from Sigma–Aldrich (St. Louis, MO, USA). Vinyl butyrate was obtained from Fluka. The ionic liquid (IL) 1-Butyl-3-methylimidazolium hexafluorophosphate ([BMIM]PF₆) with a purity of 97.0% was purchased from Sigma–Aldrich (St. Louis, MO, USA). Choline chloride (ChCl) and urea (U) were obtained from Sigma–Aldrich (St. Louis, MO, USA) and used for the preparation of deep eutectic solvents (DES), according to a previous work [17]. All organic solvents used were of analytical grade.

2.2. Preparation of CaLB Nanoflowers

The CaLB hybrid nanoflowers were prepared according to Ge et al. [8]. Typically, 0.42 mL of CuSO₄ or MnSO₄ aqueous solutions (120 mM) were added to 50 mL of phosphate buffer saline (PBS 1X, pH 7.4), which contained 0.4 mg mL⁻¹ CaLB. The mixtures were placed for incubation at 25 °C for 3 days. The nanoflower precipitates were separated by centrifugation at 4000 rpm for 10 min, washed three times with distilled water, and dried under vacuum over silica gel at room temperature. Nanoflowers were stored at 4 °C until used. The prepared copper- and manganese-based samples are labeled Cu₃(PO₄)₂ and Mn₃(PO₄)₂, respectively.

For the preparation of nanomaterials-modified CaLB nanoflowers, a similar approach was followed. Graphene oxide (GO), oxidized multi-walled carbon nanotubes (CNTs), and maghemite nanoparticles (γ -Fe₂O₃) were synthesized as reported elsewhere [29–31]. Briefly, 5 mg of GO and 3 mg of oxidized CNTs or γ -Fe₂O₃ nanoparticles were added in 49 mL of PBS and sonicated for 20 min. After the dispersion of the nanomaterials, 1 mL of CaLB solution and 0.42 mL of CuSO₄ or MnSO₄ aqueous solutions (120 mM) were added into the mixture. The next steps were the same as those described previously. Nanoflowers containing only GO or CNTs were also prepared. The prepared modified copper-based samples are labeled GO-Cu₃(PO₄)₂, CNTs-Cu₃(PO₄)₂, GO/CNTs-Cu₃(PO₄)₂, and GO/Fe₂O₃-Cu₃(PO₄)₂, and the prepared modified manganese-based samples are labeled GO-Mn₃(PO₄)₂, CNTs-Mn₃(PO₄)₂, GO/CNTs-Mn₃(PO₄)₂, and GO/Fe₂O₃-Mn₃(PO₄)₂.

2.3. Characterization of CaLB Nanoflowers

SEM images were acquired from a JEOL JSM-5600 microscope (JEOL Ltd., Tokyo, Japan) with 10 and 25 kV accelerating voltage. Moreover, the surface morphologies of the samples were determined by field emission scanning electron microscopy (FE-SEM) using a SEM Zeiss Gemini 500 (Oberkochen, Germany). Prior to SEM analysis, the nanoflowers were placed in double-sided carbon tape and sputter-coated with gold-platinum. Phase elemental distribution was studied with SEM/EDS (JEOL JSM-6510 LV equipped with an X-Act EDS-detector by Oxford Instruments, Abingdon, Oxfordshire, UK).

The XRD patterns of all CaLB-HNFs were collected on a D8 Advance Bruker diffractometer with Cu K α radiation (40 kV, 40 mA) and a secondary-beam.

Raman spectroscopy was used to confirm the presence of the carbon nanomaterials in the nanomaterials-modified CaLB-HNFs. The Raman spectra were recorded with the Labram HR system by HORIBA Scientific (HORIBA, Paris, France). The 514.5 green line of an air cooled Ar-Ion Laser was employed for the Raman excitation using a confocal aperture of 100. The laser power at the focal plane of the x100 objective was circa 0.8 mW. Spectral treatment included only a linear baseline subtraction.

FTIR was utilized to confirm the successful immobilization of CaLB in the nanoflower structure. The spectra were recorded in the range of 400–4000 cm⁻¹ using a FTIR-8400 infrared spectrometer (Shimadzu, Tokyo, Japan) equipped with a deuterated triglycine sulfate (DTGS) detector. For each sample, a total of 64 scans were averaged, using a 2 cm⁻¹ resolution. The samples were prepared

using KBr pellets containing a circa 2 wt% sample. The similarity of FTIR spectra in the Amide I region ($1600\text{--}1700\text{ cm}^{-1}$) was quantified by calculation of the correlation coefficient, r , using the following equation:

$$r = \frac{\sum xy_i}{\sqrt{\sum x_i^2 \sum y_i^2}}, \quad (1)$$

where x and y represent the spectral absorbance values of the reference and sample spectra, respectively, at the i th frequency position [32]. For identical spectra, the r value is equal to 1.0, while spectra that have differences will show lower values.

2.4. Determination of Encapsulation Yield

The amount of the immobilized CaLB was determined by calculating the protein concentration present in the supernatant after the immobilization procedure using the Bradford assay [33]. Enzyme encapsulation was estimated as the difference between the initial amount of the enzyme and the amount of the enzyme in the supernatant after immobilization.

2.5. Activity of CaLB Nanoflowers

The activity of CaLB-HNFs was determined by the hydrolysis of *p*-NPB. Specifically, 0.5 mg of CaLB nanoflowers was added into 2 mL of phosphate buffer (50 mM, pH 7.5). The reaction was initiated with the addition of 20 μL of a 50 mM *p*-NPB solution (dissolved in DMSO), and the mixture was incubated for up to 10 min at 40 °C, 650 rpm. The 4-Nitrophenol (*p*-NP) release was monitored at 405 nm. The activity was estimated by measuring the concentration of *p*-NP using a standard curve. In this work, one unit of lipase activity was defined as the specific quantity of CaLB nanoflowers required to hydrolyze 1 μmol of *p*-NPB per reaction minute. Blank measurements without any enzyme were also incubated with the substrate for ten minutes, and their absorbance was measured where no catalytic activity was observed.

2.6. Stability of CaLB Nanoflowers

The thermal stability of free CaLB and CaLB-HNFs was tested at 60 °C for up to 24 h in phosphate buffer (50 mM, pH 7.5). In order to determine the remaining activity of CaLB nanoflowers, aliquots were taken at predetermined interval times for measuring the remaining lipase activity. The remaining hydrolyzing activity was estimated as described before, monitoring the increase in the absorbance of *p*-NP.

2.7. Transesterification of Tyrosol Catalyzed by CaLB Nanoflowers

The performance of CaLB-HNFs was tested on their ability to synthesize tyrosol esters. Typically, tyrosol (20 mM), vinyl butyrate (100 mM), and 4 mg mL^{-1} of CaLB-HNFs were added in various organic solvents and ionic liquids. The reaction mixtures were incubated for 72 h under stirring at 50 °C. Synthesis reactions were repeated twice, while experiments without nanoflowers were also conducted, and any decrease in the amount of tyrosol was observed for the selected solvents. The concentration of tyrosol in the reaction mixtures was quantified by high performance liquid chromatography (HPLC), equipped with a μBondapack C18 reverse phase column (particle size 10 μm , length 300 mm, diameter 3.9 mm) and a diode array UV detector. The elution was carried out with 40% water (containing 0.1% acetic acid) in methanol at a flow rate of 1 mL min^{-1} for 30 min. Tyrosol and its ester derivative were detected at 280 nm, while the column temperature was set at 35 °C. The conversion yield of the enzymatic transesterification was based on the decrease in the concentration of tyrosol, which was calculated using a tyrosol standard curve.

2.8. Reusability of CaLB Nanoflowers

The reusability of CaLB-HNFs was tested with respect to *p*-NPB hydrolysis for nine consecutive cycles. After each catalytic cycle, the samples were recovered by centrifugation at 1000 rpm for 2 min and excessively rinsed out three times with phosphate buffer (50 mM, pH 7.5). In the case of GO/Fe₂O₃-based hybrid nanoflowers, an external magnetic field was applied after each cycle and between washing procedures. The relative activity (%) was defined as the ratio of the remaining activity to the activity of the first cycle.

Magnetic hybrid nanoflowers (GO/Fe₂O₃-based HNFs) were tested for their reusability on the transesterification of tyrosol in *tert*-butyl-methylether. Tyrosol (20 mM), vinyl butyrate (100 mM), and 4 mg mL⁻¹ of GO/Fe₂O₃ CaLB-HNFs were added in 1 mL *tert*-butyl-methylether, and the reaction mixture was incubated for 72 h under stirring at 50 °C. The nanobiocatalytic system was separated from the reaction solution by an external magnetic field and washed twice with 1 mL of *tert*-butyl-methylether. The modified nanoflowers were applied to a new reaction solution and tested as described before for eight successive cycles.

3. Results and Discussion

3.1. Morphological and Structural Characterization of CaLB Nanoflowers

In the present work, HNFs based on copper (II) or manganese (II) ions were prepared by a simple method and used as nanosupports for the encapsulation of lipase B from *Pseudozyma antarctica* (CaLB). The effect of the enrichment of HNFs with graphene oxide sheets, oxidized multi-walled carbon nanotubes, and γ-Fe₂O₃ nanoparticles on the morphological, structural, and catalytic properties of HNFs was investigated.

SEM images of unmodified Cu₃(PO₄)₂ CaLB-HNFs revealed a high quality nanoflower formation with diameters in the range of 15–30 μm (Figure 1a). Moreover, SEM images of unmodified Mn₃(PO₄)₂ CaLB-HNFs (Figure 1b) displayed a flower-like structure, though this structure was not as clear as in the case of Cu₃(PO₄)₂. The nanomaterials-modified CaLB-HNFs, either with GO, CNTs, or both carbon structures together, exhibited different structures and formations, as indicated in Figure 1c–j, while a more detailed analysis can be found in the Supplementary material (Figure S1). It is noteworthy to add that the presence of carbon nanostructures (either GO or CNTs) in manganese-based nanoflowers facilitated the formation of nanoflowers in the final structures compared to the unmodified one (Figure 1g–j). Moreover, the combination of GO and CNTs resulted in the growth of clear crystals forming particular porous flower-like structures.

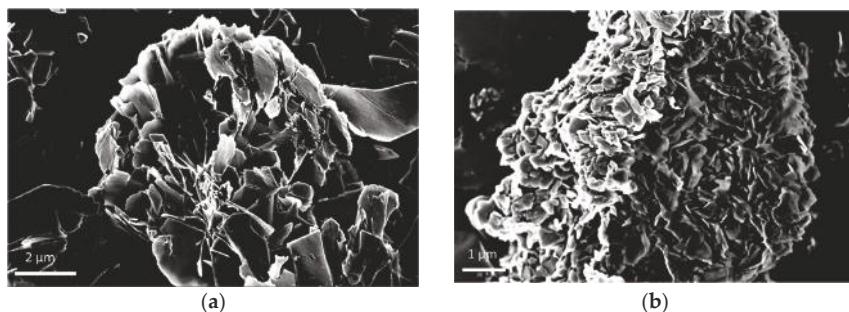


Figure 1. Cont.

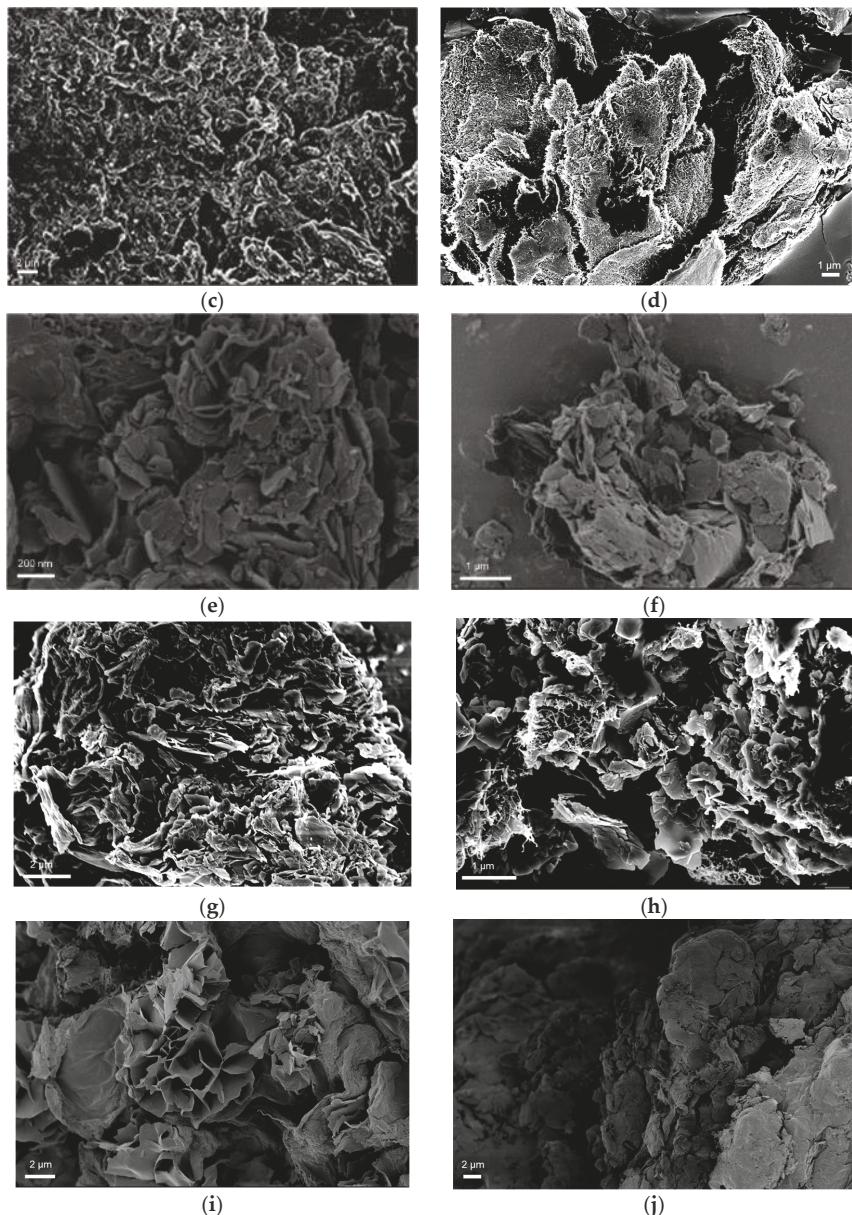


Figure 1. SEM images of: (a) unmodified $\text{Cu}_3(\text{PO}_4)_2$ CaLB-HNFs; (b) unmodified $\text{Mn}_3(\text{PO}_4)_2$ CaLB-HNFs; (c) GO- $\text{Cu}_3(\text{PO}_4)_2$ CaLB-HNFs; (d) CNTs- $\text{Cu}_3(\text{PO}_4)_2$ CaLB-HNFs; (e) GO/CNTs- $\text{Cu}_3(\text{PO}_4)_2$ CaLB-HNFs; (f) GO/ Fe_2O_3 - $\text{Cu}_3(\text{PO}_4)_2$ CaLB-HNFs; (g) GO- $\text{Mn}_3(\text{PO}_4)_2$ CaLB-HNFs; (h) CNTs- $\text{Mn}_3(\text{PO}_4)_2$ CaLB-HNFs; (i) GO/CNTs- $\text{Mn}_3(\text{PO}_4)_2$ CaLB-HNFs; and (j) GO/ Fe_2O_3 - $\text{Mn}_3(\text{PO}_4)_2$ CaLB-HNFs.

Modified GO/ Fe_2O_3 -based HNFs were further elementally analyzed using energy dispersion spectroscopy (EDS) (Figure S2). The peaks of carbon (C) and oxygen (O) were attributed to CaLB

and the incorporated nanomaterials, while the presence of nitrogen (N) and sulfur (S) confirmed the successful encapsulation of the enzyme in the nanoflower structure. The appearance of copper (Cu) (Figure S2a), manganese (Mn) (Figure S2b) and phosphate (P) peaks indicated the successful formation of the nanoflowers. Sodium (Na) and chloride (Cl) peaks appeared due to the utilized preparation buffer. Moreover, the confirmation of the presence of iron nanoparticles (Fe) in the hybrid nanoflower structures was also observed—the atomic percentages of iron were 4.34 and 5.42% in the cases of GO/Fe₂O₃-Cu₃(PO₄)₂ and GO/Fe₂O₃-Mn₃(PO₄)₂ HNFs, respectively (Table S1).

X-ray diffraction (XRD) was used to characterize the unmodified and the nanomaterials-modified CaLB-HNFs, and the XRD patterns are provided in the Supplementary material (Figure S3). For the copper-based CaLB-HNFs, the XRD patterns represented peaks for the Cu₃(PO₄)₂·3H₂O (JCPDS 00-022-0548) phase, while for the manganese-based CaLB-HNFs, the phase of manganese changed from Mn₃(PO₄)₂ for the unmodified nanoflowers to Mn₂P₂O₇ for the nanomaterials-modified HNFs [34].

The presence of the carbon-based nanostructures in the CaLB-HNFs was confirmed with Raman spectroscopy. The Raman spectra of the unmodified Cu₃(PO₄)₂ CaLB-HNFs and Mn₃(PO₄)₂ CaLB-HNFs, as well as the modified HNFs with GO, CNTs, and γ-Fe₂O₃, are presented in Figure 2. The spectrum of the unmodified Cu₃(PO₄)₂ CaLB-HNFs presented several vibrational modes; the most pronounced were located at 645 cm⁻¹, 927 cm⁻¹, and 1147 cm⁻¹ and can be attributed to the antisymmetric bending of the PO₄³⁻ ion, the symmetric stretching vibrations of PO₄³⁻ ion, and the antisymmetric stretching vibrations of the PO₄³⁻ ion, respectively [35]. The unmodified Mn₃(PO₄)₂ CaLB-HNFs presented a strong vibrational mode at 958 cm⁻¹ that can be ascribed to the symmetric stretching mode of the PO₄³⁻ ion [36].

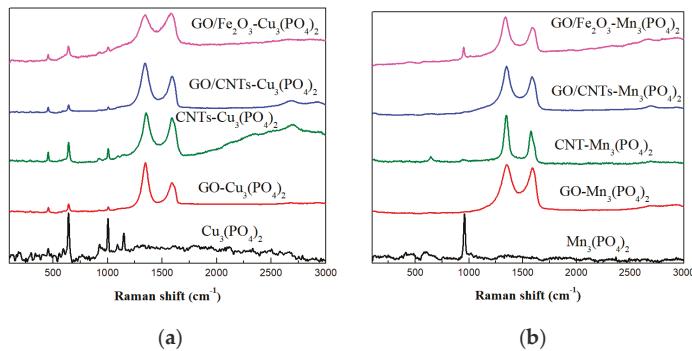


Figure 2. Raman spectra of: (a) Cu₃(PO₄)₂-based CaLB-HNFs and (b) Mn₃(PO₄)₂-based CaLB-HNFs.

The preparation of Cu₃(PO₄)₂- and Mn₃(PO₄)₂-based CaLB-HNFs with GO in both led to the appearance of the characteristic of carbon-based materials vibrational modes, D and G, located at circa 1346 cm⁻¹ and 1590 cm⁻¹, respectively [37], while a weak asymmetric 2D vibrational mode was also present in both cases at around 2685 cm⁻¹ [38]. Similarly, the preparation of both ion-based CaLB-HNFs with CNTs led to the appearance of the characteristic D and G vibrational modes located at circa 1350 cm⁻¹ and 1587 cm⁻¹, respectively, while the 2D vibrational mode of the CNTs was located at circa 2700 cm⁻¹. The more intense D and G vibrational modes were also present when both GO and CNTs were added in the CaLB-HNFs, establishing the successful incorporation of the carbonaceous nanomaterials into the HNFs.

To confirm the successful immobilization of CaLB in the 3D nanostructures, all nanobiocatalytic systems were characterized by FTIR spectroscopy by recording the spectra in the range 400 cm⁻¹ to 4000 cm⁻¹. As seen in Figure 3, peaks at the region 950 to 1060 cm⁻¹ were associated with the asymmetric stretching vibrations of PO₄³⁻, while peaks at the region 550 cm⁻¹ to 650 cm⁻¹ arose from the bending vibrations of bridging phosphate groups, such as O-P-O [39,40]. The presence of CaLB in

the nanoflower structure was confirmed by the peak at 1648 cm^{-1} , which arises from the stretching vibrations of C = O of the peptide chain of the enzyme and corresponds to the Amide I band [41,42].

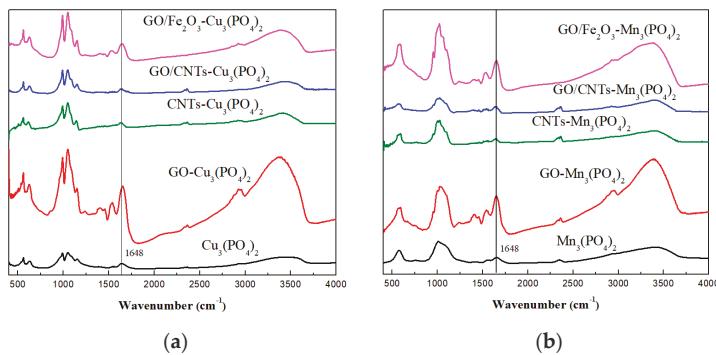


Figure 3. FTIR spectra of: (a) $\text{Cu}_3(\text{PO}_4)_2$ -based CaLB-HNFs and (b) $\text{Mn}_3(\text{PO}_4)_2$ -based CaLB-HNFs.

To better assess the dissimilarities among the spectra of the CaLB-HNFs, we compared the correlation coefficients (r) in the Amide I region ($1600\text{--}1700\text{ cm}^{-1}$), according to previously published work [43–45]. As seen from Table 1, for most of the CaLB-HNFs, r was close to 1.0, indicating that CaLB was able to maintain its native secondary structure upon encapsulation in the nanoflower structure. GO has formed a cage-like structure in which lipase was encapsulated, preserving the secondary structure of the enzyme [27]. In contrast, when CaLB-HNFs were prepared with CNTs, especially when combined with GO, the changes in the r value were more pronounced. This result could be attributed to conformational changes occurred during the encapsulation of CaLB in the nanoflower structure. The disorder of the natural conformation of CaLB, may arise from the over-crowded enzyme molecules within the strongly packed GO/CNTs HNF structure.

Table 1. Correlation coefficient (r) between the FTIR spectra of CaLB-HNFs.

Nanoflower	r	Nanoflower	r
$\text{Cu}_3(\text{PO}_4)_2$	0.976	$\text{Mn}_3(\text{PO}_4)_2$	0.982
GO- $\text{Cu}_3(\text{PO}_4)_2$	0.991	GO- $\text{Mn}_3(\text{PO}_4)_2$	0.998
CNTs- $\text{Cu}_3(\text{PO}_4)_2$	0.899	CNTs- $\text{Mn}_3(\text{PO}_4)_2$	0.837
GO/CNTs- $\text{Cu}_3(\text{PO}_4)_2$	0.879	GO/CNTs- $\text{Mn}_3(\text{PO}_4)_2$	0.801
GO/ Fe_2O_3 - $\text{Cu}_3(\text{PO}_4)_2$	0.991	GO/ Fe_2O_3 - $\text{Mn}_3(\text{PO}_4)_2$	0.998

3.2. Biocatalytic Characterization of CaLB Nanoflowers

The encapsulation yield and specific hydrolytic activity of all CaLB-HNFs are presented in Table 2. The protein loading for unmodified $\text{Cu}_3(\text{PO}_4)_2$ and $\text{Mn}_3(\text{PO}_4)_2$ CaLB-HNFs were 57.6% and 49.0%, respectively, while their specific activity was calculated at 8.3 and 96.7 U g^{-1} , respectively, pointing out that the kind of the metal ion significantly affects the hydrolyzing ability of the immobilized lipase. It has been previously proposed that enzymes provide different binding sites for metal ions, and, as such, nucleation sites are formed in different enzyme regions, affecting the 3D structure and activity of the immobilized biocatalysts [15]. Moreover, in the case of $\text{Cu}_3(\text{PO}_4)_2$ CaLB-HNFs, lipase could have been embedded deep inside the flower-like structure, preventing the active sites of the CaLB from interacting with the substrate and thus leading to low catalytic activity, due to steric hindrance phenomena [28,46].

Table 2. Encapsulation yield and specific hydrolytic activity of various CaLB-HNFs.

Nanoflower	Encapsulation Yield (%)	Specific Activity (U g ⁻¹ Immobilized CaLB)
Cu ₃ (PO ₄) ₂	57.6 ± 3.1	13.1 ± 0.5
GO-Cu ₃ (PO ₄) ₂	70.5 ± 1.7	174.4 ± 0.7
CNTs-Cu ₃ (PO ₄) ₂	57.5 ± 2.1	189.0 ± 3.9
GO/CNTs-Cu ₃ (PO ₄) ₂	61.6 ± 1.5	167.0 ± 1.7
GO/Fe ₂ O ₃ -Cu ₃ (PO ₄) ₂	59.0 ± 2.4	197.1 ± 2.5
Mn ₃ (PO ₄) ₂	49.0 ± 1.7	161.2 ± 2.6
GO-Mn ₃ (PO ₄) ₂	67.1 ± 3.6	284.7 ± 5.2
CNTs-Mn ₃ (PO ₄) ₂	57.6 ± 1.2	175.6 ± 4.0
GO/CNTs-Mn ₃ (PO ₄) ₂	65.9 ± 2.5	168.7 ± 1.0
GO/Fe ₂ O ₃ -Mn ₃ (PO ₄) ₂	60.9 ± 2.7	175.9 ± 1.9

In order to provide more binding sites for the formation of CaLB-HNFs, different carbon-based and magnetic nanomaterials were added to the hybrid nanostructures during the preparation procedure. All HNFs enriched with carbon or magnetic nanomaterials exhibited higher encapsulation yields than those without nanomaterials, regardless of the metal ion type. The highest encapsulation yields were observed when GO was used as an additive. For instance, the encapsulation efficiency reached up to 70.5 and 67.1% in the case of GO-Cu₃(PO₄)₂ and GO-Mn₃(PO₄)₂ CaLB-NHFs, respectively. Similar results have also been reported by Li and co-workers when GO was added in the formation of laccase-based nanoflowers [27]. CNTs also seem to affect the immobilization efficiency of CaLB, as is consistent with previous work [28]. The large surface area of GO and CNTs seems to increase the available binding sites and thus promote enzyme adsorption, in addition to stabilizing the 3D structure of the nanoflower. Moreover, the presence of oxygen-containing groups in the surface of these nanomaterials may result in the formation of electrostatic interactions between those functional groups and the copper cations, stabilizing the nucleation step.

The modification of CaLB-HNFs with carbon or magnetic nanomaterials enhanced the specific hydrolytic activity of the immobilized enzyme. In the case of manganese-based nanoflowers, the specific activity of the enzyme was increased up to around two-times in the presence of nanomaterials. The beneficial effect of the use of nanomaterials was more pronounced in the case of copper-based nanoflowers. More specifically, all the nanomaterials significantly outperformed in terms of activity the unmodified Cu₃(PO₄)₂ nanoflowers. GO sheets, CNTs, and γ-Fe₂O₃ nanoparticles, due to randomly distributed oxygen-containing groups on their surface, interact with positively charged metals and amino groups on the enzyme's surface, leading to more stable and active flower-like structures [28]. Such interactions could lead to a more active conformation [47,48]. Compared to each individual nanomaterial, the GO/CNTs hybrid system was not as beneficial as expected, maybe due to the uniform dispersion of lipase within the nanoflower structure or stereochemical hindrance. Moreover, CaLB immobilized on GO/CNTs nanoflowers presented the highest conformational changes (as previously discussed, Table 1), which could result in lower catalytic activity. It is important to mention that the preparation of HNFs containing both carbon nanomaterials and magnetic nanoparticles has not been previously reported. GO/Fe₂O₃-based HNFs reached high encapsulation yields, while GO/Fe₂O₃-Cu₃(PO₄)₂ CaLB-HNFs exhibited one of the highest catalytic activities among all nanoflowers.

The thermal stability of the CaLB-HNFs was also investigated. The remaining hydrolytic activity was estimated after incubation of nanoflowers for up to 24 h in phosphate buffer at 60 °C, and is presented in Figure 4.

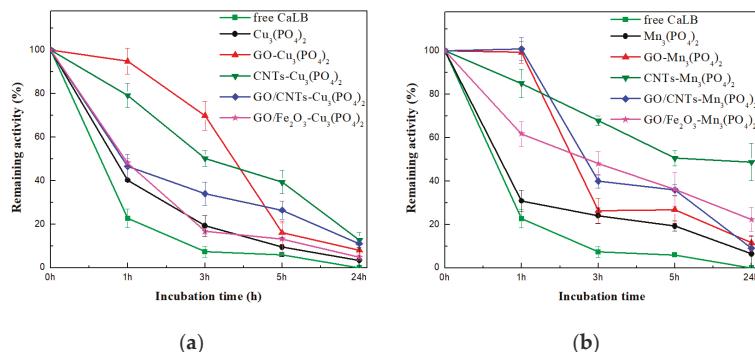


Figure 4. Thermal stability of: (a) $\text{Cu}_3(\text{PO}_4)_2$ -based CaLB-HNFs and (b) $\text{Mn}_3(\text{PO}_4)_2$ -based CaLB-HNFs at 60 °C. The 100% percentage corresponds to the activity at t = 0 min.

As seen in Figure 4, the use of hybrid nanoflowers as supports for the immobilization of CaLB increased the thermal stability of the immobilized enzyme. Specifically, the catalytic activity of free CaLB decreased to 20% after the first hour of incubation, while unmodified CaLB- $\text{Cu}_3(\text{PO}_4)_2$ and $\text{Mn}_3(\text{PO}_4)_2$ CaLB-HNFs retained up to 40% and 31% of their initial activity, respectively. Moreover, after 5 h of incubation, free CaLB was totally inactivated, while the immobilized lipase on unmodified HNFs retained up to 19.2% of their activity, indicating that the nanoflower 3D structure can protect the active conformation of the enzyme, thus enhancing its stability [49]. Similar observations have also been reported for lipase from the porcine pancreas [50]. The thermal stability of the immobilized CaLB was further improved when HNFs containing carbon and $\gamma\text{-Fe}_2\text{O}_3$ nanomaterials were used as immobilization supports. This observation could be attributed to the protective effect these nanomaterials offer on the stability of protein molecules [51,52]. Amongst the nanomaterials, CNTs stabilized the immobilized enzyme the most for both copper- and manganese-based nanoflowers (12.7% and 49% of enzyme activity, respectively, was retained after 24 h of incubation). Their high surface area, as well as the fact that CNTs are distributed within the petals of the flower-like structure, enables lipase to maintain its stability [27,28]. The conformational changes previously described (Table 1) may lead to a more rigid folding of lipase and thus enhance its stability [51]. Furthermore, in comparing the two inorganic components, it is clear that manganese HNFs exhibited higher remaining activity than copper HNFs, underlining the correlation of the different interactions developed between nanomaterials and each metal ion.

One of the major drawbacks of using soluble enzymes in large-scale reactions is reusability, due to their incapability of maintaining their stability under harsh conditions, and the difficulty of removal from the reaction system, due to their high solubility. Therefore, the immobilization of enzymes enhances their stability and enables their separation and use in successive cycles, making them an asset for industrial applications. In the present study, the operational stability of the CaLB-HNFs was investigated for the hydrolysis of *p*-NPB, and the results are presented in Figure 5. As seen, unmodified $\text{Cu}_3(\text{PO}_4)_2$ and $\text{Mn}_3(\text{PO}_4)_2$ CaLB-HNFs were almost deactivated after the fifth biocatalytic cycle. It is possible that the disruption of the non-covalent bonds between the organic and inorganic parts of the nanoflowers in the aqueous environment accelerated the enzyme leaching or gradual degradation of the flower-like morphology, leading to low enzymatic activity, which is in agreement with that recently reported [40].

In the case of nanomaterials-modified HNFs, the operational stability of CaLB was notably increased. Immobilized CaLB on nanomaterials-based HNFs could be efficiently used for nine consecutive cycles for the hydrolysis of *p*-NPB. The residual activity of nanomaterials-modified $\text{Cu}_3(\text{PO}_4)_2$ and $\text{Mn}_3(\text{PO}_4)_2$ CaLB-HNFs retained up to 83% even after nine catalytic cycles. These results infer that the presence of nanomaterials in the nanoflowers protects the enzyme configuration,

thus enhancing its stability for successive hydrolysis cycles. Similar to the thermal stability studies presented above, CNTs-modified CaLB-HNFs offered the most beneficial impact on the operational stability of CaLB, indicating that the incorporation of CNTs inside the nanoflower structure enables the adoption of a more rigid conformation of CaLB, stabilizing it against repeatable usage [53]. It is interesting to note that, although manganese-based HNFs presented higher thermal stability than copper-based HNFs (as previously discussed), their operational stability was lower compared to copper-based HNFs. It is possible that the enzyme leaching from the manganese-based HNFs is higher compared to copper-based HNFs, resulting in a higher loss of the residual activity of the enzyme.

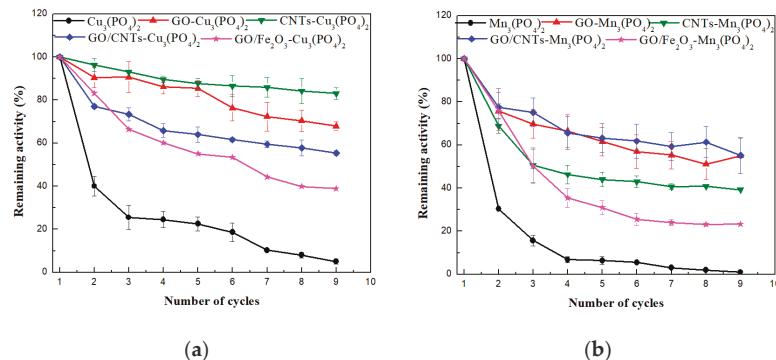


Figure 5. Operational stability of: (a) Cu₃(PO₄)₂-based CaLB-HNFs and (b) Mn₃(PO₄)₂-based CaLB-HNFs. The 100% percentage corresponds to the lipase hydrolytic activity at the first catalytic cycle.

3.3. Transesterification of Tyrosol by CaLB Nanoflowers in Non-Aqueous Media

The prepared nanomaterials-modified CaLB-HNFs were used for the synthesis of tyrosol esters in non-aqueous media. Tyrosol is a natural phenolic antioxidant derived from various plants, such as olive and green tea. This abundant product has been associated with many health-related benefits as well as plenty of industrial applications [54,55]. An increase of tyrosol lipophilicity is suggested to improve its antioxidant activity [56,57]. Thus, the enzymatic lipophilization of tyrosol may be of great interest. For this reason, GO/Fe₂O₃-Cu₃(PO₄)₂ and GO/Fe₂O₃-Mn₃(PO₄)₂ CaLB-HNFs were used as biocatalysts for the transesterification of tyrosol with vinyl butyrate (Figure S4) in a variety of organic solvents, as well as in environmentally friendly ionic and deep eutectic solvents; the results are presented in Table 3. GO/CNTs-based HNFs were also used for the transesterification of tyrosol; the results are presented in Table S2.

Table 3. Conversion yields for the enzymatic transesterification of tyrosol with vinyl butyrate in non-aqueous media catalyzed by GO/Fe₂O₃-Cu₃(PO₄)₂ and GO/Fe₂O₃-Mn₃(PO₄)₂ CaLB HNFs.

Reaction Medium	Conversion Yield (%)	
	GO/Fe ₂ O ₃ -Cu ₃ (PO ₄) ₂ CaLB-HNFs	GO/Fe ₂ O ₃ -Mn ₃ (PO ₄) ₂ CaLB-HNFs
n-Hexane	99.6 ± 0.4	100.0 ± 0.3
Acetonitrile	80.3 ± 0.3	80.7 ± 0.8
2-Methyl-2-butanol	30.2 ± 1.1	52.6 ± 1.2
tert-Butyl-methylether	98.9 ± 0.5	99.7 ± 0.6
tert-Butanol	23.2 ± 0.2	22.5 ± 0.4
[BMIM][PF ₆]	13.6 ± 1.6	20.0 ± 4.7
ChCl:U	33.2 ± 2.8	26.7 ± 4.6

As seen in Table 3, both CaLB-HNFs were able to catalyze the transesterification of tyrosol, achieving high conversion yields in most of the non-aqueous solvents. It has been recently proposed that the hydrophobic surface of the hybrid nanoflowers benefits synthetic reactions in non-aqueous solvents by promoting the oriented delivery of the substrates near the hydrophobic surface of nanoflowers and, thus, to the active sites of the enzyme [49]. Conversion yields of transesterification seem to strongly depend on the nature of the organic solvent, namely its polarity and viscosity. More specifically, the nanoflower-catalyzed reactions in non-polar solvents, e.g., n-hexane and *tert*-butyl-methylether, exhibited high conversion yields up to 100%. Moreover, the reaction rate of the transesterification reaction catalyzed by GO/Fe₂O₃-Mn₃(PO₄)₂ CaLB-HNFs in hexane and *tert*-butyl-methylether was up to 73-fold higher in comparison with that in other media (Table S3). Solvents with low polarity enable enzymes to preserve the essential water molecules bound on their surface in order to maintain their natural conformation and be fully functional [58,59]. On the other hand, the use of more hydrophilic solvents with higher affinity to interact with water molecules [60], such as 2-methyl-2-butanol and *tert*-butanol, led to a decrease of the conversion yield of the transesterification reaction.

Both HNFs were able to catalyze the transesterification of tyrosol in eco-friendly alternatives of organic solvents, such as ionic liquids and deep eutectic solvents ([BMIM][BF₆] and ChCl:U, respectively). Those green solvents have been widely employed for enzymatic biotransformations, as they present high chemical and thermal stability, low vapour pressure, low toxicity, and the ability to enhance the catalytic performance of the enzymes [17,55,61]. As seen in Table 3, the conversion yield is decreased in the ionic liquid compared to organic solvents, which could be attributed to the low dispersability of the nanoflowers in these media. Moreover, the high viscosity of [BMIM][BF₆] (381 cP at 25 °C)[62] and ChCl:U (1200 mPa s at 25 °C)[63] could lead to mass-transfer limitations, restricting the biocatalytic activity of immobilized CaLB [64,65].

The use of magnetic nanobiocatalysts could facilitate the separation from the reaction solution through the application of an external magnetic field and, thus, the reuse of the biocatalyst [52]. Considering this aspect, magnetic CaLB-HNFs were applied in consecutive reaction cycles for tyrosol transesterification. Figure 6 presents the remaining catalytic activity of the GO/Fe₂O₃-Mn₃(PO₄)₂ CaLB-HNFs for successive catalytic cycles. As seen, these magnetic nanoflowers presented excellent operational stability after eight consecutive reaction cycles (576 h of total operation) without any loss of biocatalytic activity, making this hybrid nanobiocatalyst one of the most robust nanobiocatalysts reported until now for similar reaction processes. This enhanced operational stability, in comparison with the one described above for the hydrolysis of *p*-NPB (Figure 5), could be attributed to the fact that non-polar organic solvents do not remove protein-bound water that is crucial for maintaining protein structure and function, leading to a more rigid conformation of the immobilized biocatalyst [66,67].

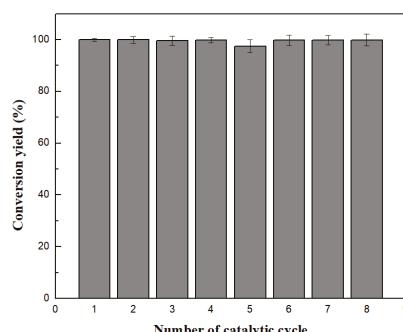


Figure 6. Operational stability of GO/Fe₂O₃-Mn₃(PO₄)₂ CaLB-HNFs after eight reaction cycles for the enzymatic transesterification of tyrosol with vinyl butyrate in *tert*-butyl-methylether. Each reaction was carried out for 72 h at 50 °C.

4. Conclusions

Herein, we report the preparation and characterization of novel hybrid nanoflowers comprised of copper (II) or manganese (II) ions combined with magnetic nanoparticles and carbon-based nanomaterials. These nanoflowers can be effectively used as versatile host platforms for the immobilization of an industrially relevant enzyme (CaLB) through biomimetic mineralization. The metal ion and the nature of the nanomaterials affect the structural and catalytic characteristics of the immobilized lipase in different manners. The nanomaterials-modified hybrid nanoflowers presented an excellent catalytic performance in the production of tyrosol esters in different organic solvents and environmental-friendly ionic solvents. Furthermore, CaLB-magnetic HNFs (combining GO and maghemite nanoparticles) exhibited remarkable operational stability for the tyrosol transesterification reaction, as the nanobiocatalyst retained almost its entire catalytic activity even after eight successive reaction cycles, indicating that these bio-nanoconjugates could potentially be used as efficient tools for heterogeneous biocatalytic transformations in large-scale applications.

Supplementary Materials: The following are available online at <http://www.mdpi.com/2079-4991/9/6/808/s1>, Figure S1: SEM images of: (A1, A2) $\text{Cu}_3(\text{PO}_4)_2$ CaLB-HNFs; (B1-B3) GO- $\text{Cu}_3(\text{PO}_4)_2$ CaLB-HNFs; (C1-C3) CNTs- $\text{Cu}_3(\text{PO}_4)_2$ CaLB-HNFs; (D1-D3) GO/CNTs- $\text{Cu}_3(\text{PO}_4)_2$ CaLB-HNFs; (E1, E2) GO/ Fe_2O_3 - $\text{Cu}_3(\text{PO}_4)_2$ CaLB-HNFs; (F1-F3) $\text{Mn}_3(\text{PO}_4)_2$ CaLB-HNFs; (G1-G3) GO- $\text{Mn}_3(\text{PO}_4)_2$ CaLB-HNFs; (H1,H2) CNTs- $\text{Mn}_3(\text{PO}_4)_2$ CaLB-HNFs; (I1-I3) GO/CNTs- $\text{Mn}_3(\text{PO}_4)_2$ CaLB-HNFs; (J1, J2) GO/ Fe_2O_3 - $\text{Mn}_3(\text{PO}_4)_2$ CaLB-HNFs, Figure S2: EDS spectra of: (a) GO/ Fe_2O_3 - $\text{Cu}_3(\text{PO}_4)_2$ CaLB-HNFs; (b) GO/ Fe_2O_3 - $\text{Mn}_3(\text{PO}_4)_2$ CaLB-HNFs, Table S1: Elemental analysis of GO/ Fe_2O_3 -based CaLB-HNFs by EDS, Figure S3: XRD patterns of: (a) $\text{Cu}_3(\text{PO}_4)_2$ -based CaLB-HNFs; (b) $\text{Mn}_3(\text{PO}_4)_2$ -CaLB-HNFs, Figure S4: Transesterification of tyrosol with vinyl butyrate catalyzed by CaLB, Table S2: Conversion yields for the enzymatic transesterification of tyrosol with vinyl butyrate in non aqueous media, by GO/CNTs- $\text{Cu}_3(\text{PO}_4)_2$ CaLB-HNFs, Table S3: Reaction rates (mM h^{-1}) of tyrosol transesterification catalyzed by GO/ Fe_2O_3 - $\text{Mn}_3(\text{PO}_4)_2$ CaLB-HNFs in non-aqueous media.

Author Contributions: H.S. contributed to the overall design of the experiments, results interpretation, and manuscript writing. D.G., A.A., A.P., and E.P.G. contributed to results interpretation and reviewed the manuscript. R.F. performed the synthesis and experiments with nanoflowers, results interpretation, and manuscript writing. M.P. contributed to results interpretation and manuscript writing. M.A.H., A.E., K.S., and D.M. performed the synthesis of nanomaterials. S.E.M., E.D.S., and X.R.D. contributed to the characterization of nanoflowers and the results analysis. K.T. performed the Raman characterization of nanoflowers, contributed to results analysis, and reviewed the manuscript.

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Conflicts of Interest: The authors declare no conflict of interest.

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Article

From Hollow to Solid Carbon Spheres: Time-Dependent Facile Synthesis

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Abstract: Here, we report a facile route for obtaining carbon spheres with fully tunable shell thickness. Using a hard template in chemical vapor deposition (CVD), hollow carbon spheres, solid carbon spheres, and intermediate structures can be obtained with optimized process time. The resulting carbon spheres with particle diameters of ~400 nm, as well as a controllable shell thickness from 0 to 70 nm, had high Brunauer–Emmett–Teller (BET) specific surface area (up to $344.8 \text{ m}^2 \cdot \text{g}^{-1}$) and pore volume (up to $0.248 \text{ cm}^3 \cdot \text{g}^{-1}$). The sphere formation mechanism is also proposed. This simple and reproducible technique can deliver carbon materials for various applications, e.g., energy storage and conversion, adsorption, catalytic, biomedical, and environmental applications.

Keywords: mesoporous materials; carbon spheres; hollow carbon spheres; solid carbon spheres; CVD process; time dependence

1. Introduction

Carbon materials have been extremely popular for decades. Among them, carbon spheres attracted huge scientific attention [1]. Their history began in the 1990s. Firstly, solid and hollow carbon spheres were obtained from the thermal decomposition of methane [2–4] or camphor vapors [5] in the presence of a metal catalyst precursor. Currently, a number of techniques for preparing carbon spheres are applied, such as hydrothermal carbonization (HTC) [6], arc-discharge [7], and laser ablation [8]. However, the chemical vapor deposition (CVD) process [9] is the most common method for carbon sphere fabrication. Various precursors and templates were used to prepare carbon spheres using the CVD method. The templates were classified into hard and soft templates. Soft templates are precursors inducing the self-assembly of the final product, such as hexadecyltrimethylammonium bromide (CTAB). Hard templates are particles acting as solid cores with carbon replicas. Soft templates are much easier to be eliminated from the products. However, hard templates allow better control over the fabricated structures. Recently, aluminosilicate templates such as halloysite are becoming more and more popular. They can be used to obtain tubular structures [10]. However, to induce the formation of spherical carbon materials, silica spheres are the most commonly used templates. This is mainly due to a wide range of available diameters from 5 nm to several hundreds of nanometers, and the ease of the removal procedure from the final product [11–13]. Obtained hollow carbon spheres have various structural properties which result in high specific surface area, large pore volume, low density, thermal conductivity, and electrical conductivity, as well as good chemical and mechanical stability [14]. Additional unique properties of carbon spheres are related to low density, excellent reactivity, high compressive strength, thermal insulation, and large cavity space [15,16]. There are many reports describing the use of carbon spheres in energy storage and conversion, adsorption, catalytic, biomedical, and environmental applications. In the energy storage field, they are used as active materials of electrodes in supercapacitors [17–19] and lithium-ion batteries [20–22]. Furthermore,

carbon spheres can act as catalysts in various reactions [23,24], and they are used for CO₂ capture [25,26]. Moreover, they also serve well for biomedical applications [27–29], for example in drug delivery [27].

So far, the control of the structure of carbon spheres is related to two main structural characteristics: (i) the shell thickness [30,31] and (ii) the order of the pores in the shell. The proposed carbon sphere can exhibit ordered [13,32] and disordered [33] pores. It is well known that the nanoparticle shape has a crucial role in its applications, such as drug delivery [34], filler for polymeric matrices [35,36], and deacidification [37,38]. In our contribution, solid silica nanoparticles serve as a hard template in CVD-grown carbon spheres. Furthermore, optimizing the process time of the procedure allowed the growth of carbon spheres with empty cores, solid carbon spheres, and intermediate carbon structures. We also propose a mechanism for sphere formation, taking into consideration the mesoporous silica particles (m-SiO₂) being utilized as the template. Schematically, the change in morphology of the spheres prepared using CVD for different times, from hollow to solid carbon spheres, is shown in Figure 1.

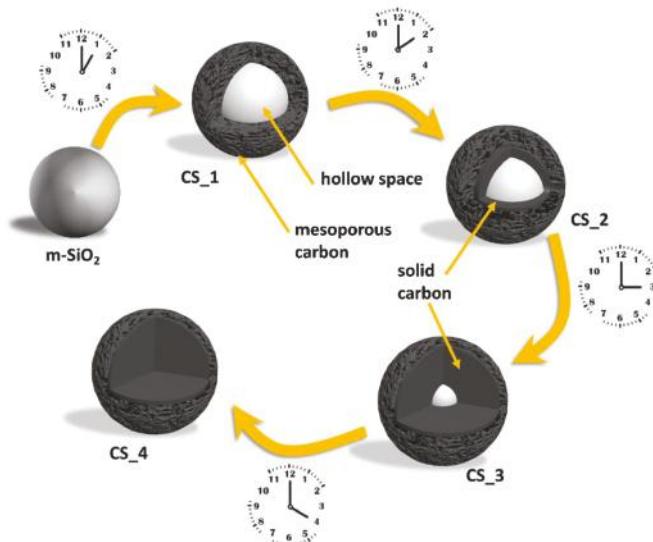


Figure 1. Scheme presenting the change in morphology of resulting carbon spheres after 1 h (CS_1), 2 h (CS_2), 3 h (CS_3) and 4 h (CS_4) of chemical vapor deposition (CVD).

2. Materials and Methods

2.1. Synthesis of Mesoporous Silica Spheres

Mesoporous silica spheres were prepared as templates for further processing. Briefly, the surfactant, hexadecyltrimethylammonium bromide (CTAB, MERCK, Darmstadt, Germany; 900 mg), was added to a mixture of ethanol (EtOH, MERCK, Darmstadt, Germany; 180 mL), distilled water (240 mL), and ammonia (MERCK, Darmstadt, Germany; 25 wt.%, 3.3 mL), before being sonicated to obtain a homogeneous solution, and stirred vigorously for 30 min. Next, the silica precursor, tetraethyl orthosilicate (TEOS, MERCK, Darmstadt, Germany; 1.2 mL), was added to the reaction mixture and subsequently stirred at room temperature overnight. Finally, the product was centrifuged and dried [39].

2.2. Synthesis of Carbon Spheres with Different Shell Thickness (from Hollow to Solid Carbon Spheres)

The as-prepared m-SiO₂ template was used to prepare the carbon spheres using the CVD method. The m-SiO₂ template in an alumina boat was placed into a tube furnace in the presence of argon and ethylene at flow rates of 100 sccm and 30 sccm, respectively. The temperature was raised to 800 °C. Processes with different carbonization times (1, 2, 3, and 4 h) were performed. Afterward, the resulting spheres (m-SiO₂_CS) were washed with hydrofluoric acid to remove the silica and obtain the final product—carbon spheres with different morphology.

2.3. Characterization

The morphology of the samples was examined with a transmission electron microscope (TEM; Tecnai F30, Thermo Fisher Scientific, Waltham, MA, USA) and a scanning electron microscope (SEM; VEGA3 TESCAN, Brno, Czech Republic; high voltage (HV): 30 kV, working distance (WD): 5.25 mm). X-ray diffraction (XRD) patterns were carried out using an X'Pert Philips Diffractometer (X'Pert PRO Philips diffractometer, Co. Ka radiation, Almelo, Holland) with a Cu lamp ($K\alpha_1 = 1.54056 \text{ \AA}$) to investigate the crystal composition of the samples. Thermogravimetric analysis (TGA) was carried out on 10-mg samples using a DTA-Q600 SDT TA Instrument (TA Instrument, New Castle, DE, USA) at a heating rate of 5 °C/min from room temperature to 900 °C in air. Raman spectra were determined using an inVia Raman Microscope (Renishaw, New Mills Wotton-under-Edge, UK) with an excitation wavelength of 785 nm. N₂ adsorption/desorption isotherms were obtained using a Quadrosorb SI (Quantachrome Instruments, Boynton Beach, FL, USA). Specific surface area was calculated according to the Brunauer–Emmett–Teller (BET) method, and pore size distribution was determined using the density functional theory (DFT) method.

3. Results and Discussion

Scanning electron microscopy was used to reveal the evolution of morphology and structure of the obtained materials, i.e., the silica template and all carbon spheres. As illustrated in Figure 2a,b, a uniform and spherical shape with a diameter of ~400 nm, as well as a smooth surface, can clearly be observed in the silica template sample. After the CVD process, carbon materials kept the initial spherical shape of the silica template and exhibited a similar diameter of ~400 nm. It can clearly be observed that, with an increase in synthesis time, the morphology of the carbon spheres did not change significantly (Figure 2c–f).

Representative TEM images of silica spheres and carbon spheres are presented in Figure 3. Silica spheres used as a template for the carbon spheres had diameter of ~400 nm (Figure 3a,b), which was confirmed by SEM analysis [34]. The carbon spheres after removal of the silica template had porous shells with thicknesses changing with time. After 1 h (CS_1) of CVD, the shell thickness was ~70 nm (Figure 3c,d). After 2 h (CS_2), the shell diameter decreased to ~30 nm (Figure 3g,h). The spheres produced after 3 h of CVD (CS_3) had a shell thickness ~20 nm (Figure 3i,j). When the process took 4 h, the sample did not exhibit the presence of a shell. Its diameter was the same as the silica template. Moreover, the core of the sphere was no longer hollow. Solid carbon spheres were produced under these experimental conditions. Interestingly, with the increase in process time, the carbon atoms increasingly diffused from the external shell into the spheres.

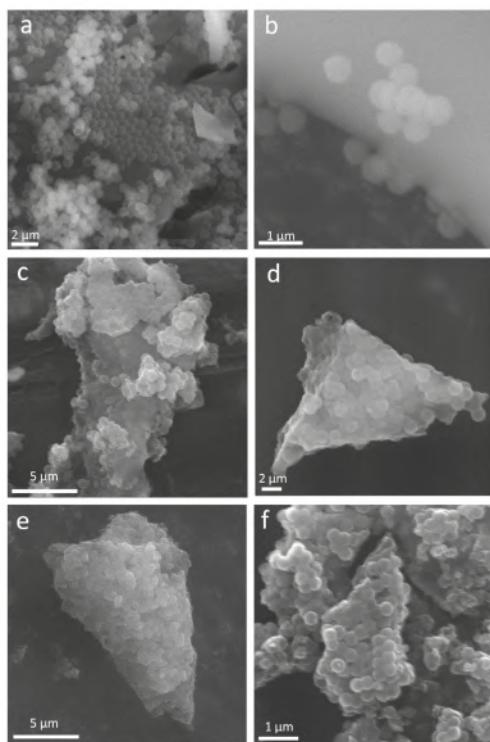


Figure 2. SEM images of silica template (a,b) and carbon spheres after 1 h (c), 2 h (d), 3 h (e), and 4 h (f) of synthesis.

The graphitization degree of the carbon spheres was determined using XRD. Figure 4a displays the XRD patterns, in which the two peaks at $\sim 25^\circ$ and $\sim 43^\circ$ can be assigned to typical graphitic (002) and (100) planes, respectively. The broadening of the two peaks suggests a low graphitization degree and the possible presence of amorphous carbon. However, upon extending the CVD process time, sharper peaks with greater intensity were detected. This means that the degree of graphitization of the final material increased.

The bonding, order, and crystallinity of the materials were studied using Raman spectroscopy (Figure 4b). The presence of disordered graphitic materials was suggested by the two Raman modes. The peak at 1604 cm^{-1} (G band) corresponds to the E_{2g} mode of hexagonal graphite, and it is related to the vibration of the sp₂-hybridized carbon atoms in the graphite layer. This implies that the carbon spheres were composed of graphitic carbon, which is consistent with the TEM and XRD results. The D band at approximately 1312 cm^{-1} is associated with the vibration of carbon atoms with dangling bonds in the plane with termination by disordered graphite. The D band had higher intensity than the G band, which suggests that the obtained carbon spheres had several defects, and they consisted mostly of amorphous carbon [40].

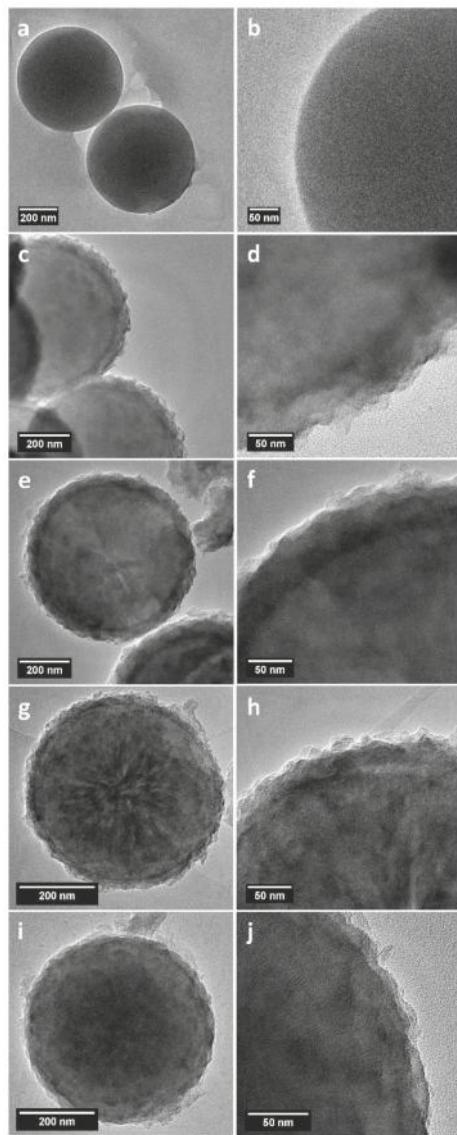


Figure 3. TEM images of silica template (**a,b**) and carbon spheres after 1 h (**c,d**), 2 h (**e,f**), 3 h (**g,h**), and 4 h (**i,j**) of synthesis.

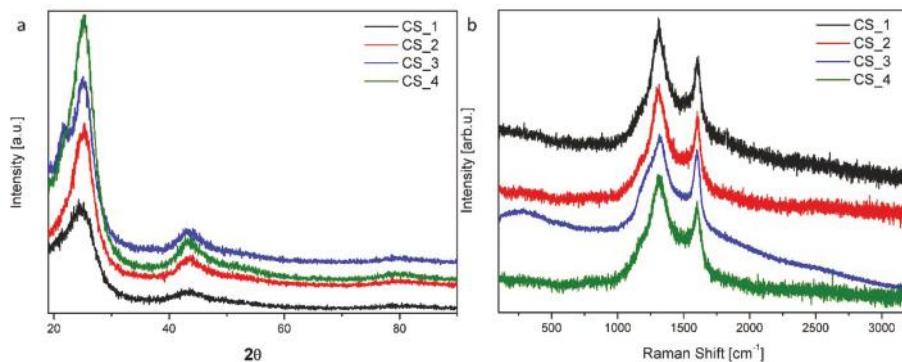


Figure 4. X-ray diffraction (XRD) patterns (a) and Raman spectra (b) of carbon spheres with different times of synthesis.

The porosity of the synthesized samples was tested using N_2 adsorption/desorption experiments. The textural properties are listed in Table 1. Type IV isotherms with H4 hysteresis loops were observed in all samples, which are typical of mesoporous materials (Figure 5). The position of the P/P_0 inflection points is associated with the range of mesopore size, and the slope degree of the steps indicate the uniformity of mesopore size. There were capillary condensation steps at P/P_0 of 0.4–1.0, ascribed to mesopores in the samples [41]. The pore size distribution curves show the existence of uniform mesopores below 4 nm. These mesopores should be located in the shell of carbon spheres. Upon increasing the time of synthesis, the BET specific surface area of carbon spheres increased gradually to $344.8 \text{ m}^2 \cdot \text{g}^{-1}$, as well as the total pore volume to $0.248 \text{ cm}^3 \cdot \text{g}^{-1}$, in CS_2. A further increase in the process time caused a decrease in BET specific surface area and total pore volume, which may be related to the blocking of mesoporous channels by the diffused carbon. When the carbon spheres were filled with carbon, the diffusion ceased and the shell disappeared; the BET specific surface area and total pore size distribution were again enhanced. The control of the texture parameters of carbon spheres, such as specific surface area, total pore volume, and pore size, is important in adsorption and catalytic applications.

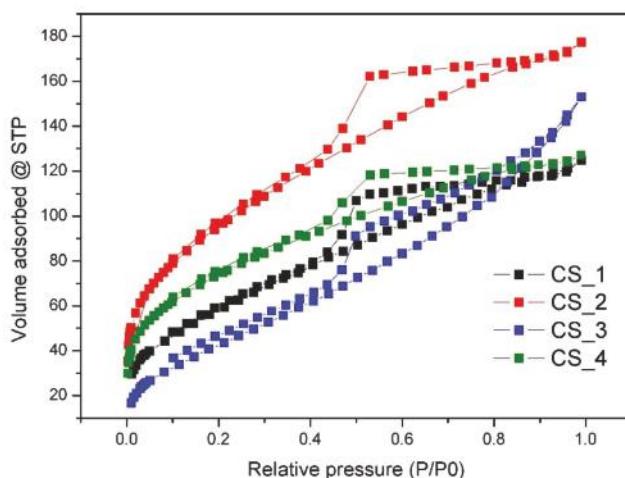


Figure 5. N_2 sorption isotherms of carbon spheres with different times of synthesis.

Table 1. Texture parameters of carbon spheres with different times of synthesis. SBET—Brunauer–Emmett–Teller surface area; Vtotal—total pore volume; m-SiO₂—mesoporous silica; CS_1–4—carbon spheres with 1–4 h of chemical vapor deposition.

Sample	SBET ^a (m ² ·g ⁻¹)	Vtotal ^b (cm ³ ·g ⁻¹)	Pore Size ^c (nm)	Error ^d
m-SiO ₂	224.2	0.210	2.313	0.391%
CS_1	218.2	0.174	3.969	0.609%
CS_2	344.8	0.248	3.167	0.970%
CS_3	169.8	0.212	2.245	0.544%
CS_4	263.5	0.177	1.178	0.621%

^a Determined using the multipoint BET method. ^b Calculated from density functional theory (DFT) method for cumulative pore volume. ^c Determined by the DFT pore diameter mode. ^d Fitting error from the DFT method summary.

The TGA measurements in Figure 6 provide information on the carbon content and quality of the structure in the carbon spheres. It is known that carbon with a better crystalline structure decomposes at higher temperature. For example, carbon nanomaterial with well-ordered sp₂ hybridization starts decomposing above 600 °C [42–44], while amorphous carbon initiates its decomposition at a lower temperature, ca. 500 °C and below [45,46]. Figure 6a shows that our carbon spheres were oxidized above 415 °C, which means that our carbon spheres consisted mostly of amorphous carbon, which is in full agreement with the Raman and XRD spectra. As the temperature increased further, weight decreased rapidly until all of carbon spheres were exhausted at approximately 730 °C. The ash contents of the samples after combustion were 54.3% (*w/w*) for m-SiO₂_CS_1, 50.6% for m-SiO₂_CS_2, 31.5% for m-SiO₂_CS_3, and 28.5% for m-SiO₂_CS_4, which indicates that the contribution of silica decreased. A simultaneous increase in synthesis time and the amount of carbon in the samples was observed. After removing the silica core, the TGA curves changed slightly (Figure 6b). All samples consisted only of carbon; thus, they were completely burnt at high temperatures of up to 800 °C, proving the high purity of the final samples. However, the temperature of total combustion was different, and it shifted with the time of the CVD process. CS_1 burned completely at about 650 °C, while CS_4 burned completely at about 755 °C. Full thermal decomposition parameters calculated from the differential thermogravimetry (DTG) curves are listed in Table 2. The start (T_{start}) and end (T_{end}) temperatures of the peak, as well as the temperature at which the peak had its maximum (T_{max}), are specified. Both T_{max} and T_{end} shifted to higher temperature values with an increase in CVD process time. This indicates that the carbon crystallinity was enhanced when the CVD process was longer. This is in good agreement with the increasing intensity of the peak at 25° in the XRD measurements shown earlier.

Detailed microscopic analysis allowed a proposal of the growth mechanism. In the initial phase of synthesis, carbon atoms adsorbed onto the surface of SiO₂, forming a porous shell. Therefore, the diameter of the spheres after 1 h of CVD process was larger than that of the pristine template. With an increase in CVD time, carbon atoms started diffusing from the external part of the shell to the core, and the shell thickness decreased. However, more carbon atoms were deposited onto the interior of the template. Gradually, the shell disappeared and, at some point, the mesoporous channels of silica were filled with diffused carbon. When the carbon spheres were filled with carbon, the diffusion slowed down and the shell disappeared, resulting in the reduction in BET specific surface area and pore size distribution of the carbon spheres. This effect was clearly observed in the case of CS_4.

Table 2. Thermal decomposition parameters of carbon spheres with different times of synthesis from differential thermogravimetry (DTG) curves in Figure 6b.

Sample	T _{start} (°C)	T _{max} (°C)	T _{end} (°C)
CS_1	110	580	650
CS_2	220	680	715
CS_3	450	705	735
CS_4	450	725	755

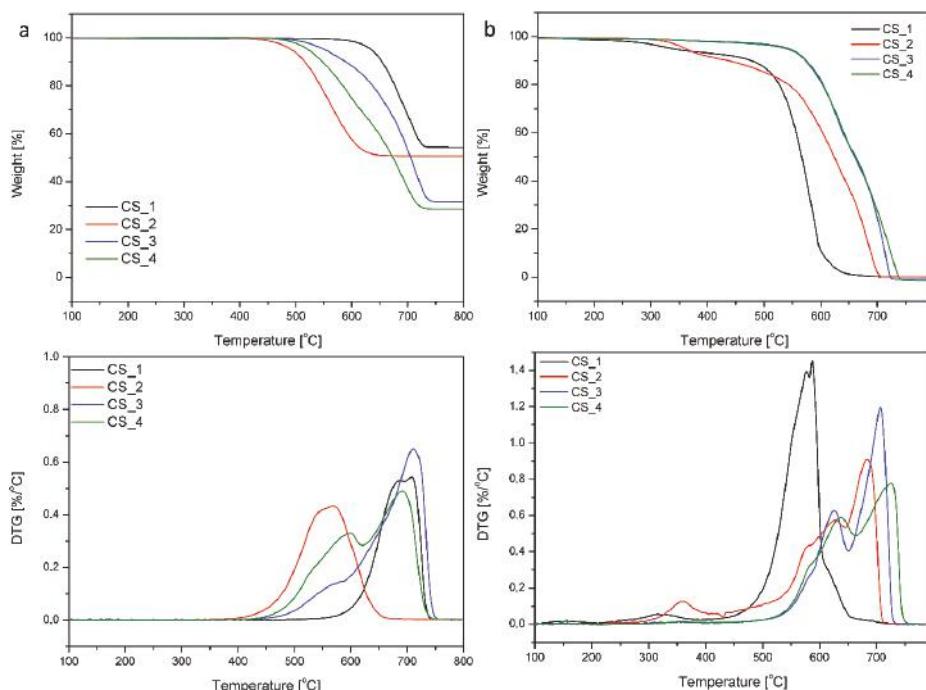


Figure 6. Thermogravimetry analysis (TGA) and differential thermogravimetry (DTG) curves of carbon spheres with different times of synthesis. Before (a) and after (b) removing silica template.

4. Conclusions

In this work, we successfully fabricated solid and hollow mesoporous carbon spheres with controllable shell thickness, as well as high specific surface area and pore volume, using the CVD method. Specific surface areas and pore volume distributions of the carbon spheres could be tuned with the CVD process time. Optimizing the process time of the procedure allowed the growth of carbon spheres with empty cores (1 h CVD), solid carbon spheres (4 h CVD), and intermediate carbon structures (2 or 3 h CVD). The mechanism of sphere formation was also proposed. It is believed that this facile synthesis route can be a way of preparing carbon spheres with different morphology (from hollow to solid), which can be tested for various applications, such as energy storage and conversion, adsorption, catalytic, biomedical, and environmental applications.

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Article

Effects of Sample Preparation on Particle Size Distributions of Different Types of Silica in Suspensions

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Abstract: The granulometric characterization of synthetic amorphous silica (SAS) nanomaterials (NMs) still demands harmonized standard operation procedures. SAS is produced as either precipitated, fumed (pyrogenic), gel and colloidal SAS and these qualities differ, among others, with respect to their state of aggregation and aggregate strength. The reproducible production of suspensions from SAS, e.g., for biological testing purposes, demands a reasonable amount of dispersing energy. Using materials representative for each of the types of SAS, we employed ultrasonic dispersing (USD) at energy densities of 8–1440 J/mL and measured resulting particle sizes by dynamic light scattering and laser diffraction. In this energy range, USD had no significant impact on particle size distributions of colloidal and gel SAS, but clearly decreased the particle size of precipitated and fumed SAS. For high energy densities, we observed a considerable contamination of SAS suspensions with metal particles caused by abrasion of the sonotrode's tip. To avoid this problem, the energy density was limited to 270 J/mL and remaining coarse particles were removed with size-selective filtration. The ultrasonic dispersion of SAS at medium levels of energy density is suggested as a reasonable compromise to produce SAS suspensions for toxicological in vitro testing.

Keywords: nanomaterials (NMs); nanostructured; synthetic amorphous silica (SAS); ultrasonic dispersing (USD); energy density; sample preparation; in vitro testing

1. Introduction

Modification of physico-chemical properties of nanomaterials (NMs) or nanostructured materials allows the control and variation of design, development and improvement of new products. Synthetic amorphous silica (SAS) comprise an important group of NMs, which are added to industrial as well as consumer products such as cosmetic or foods [1–7] within which they serve e.g., as stabilizers, thickeners, pigments, flow enhancing agents, or UV absorbers [3,8–10]. Based on some concern regarding possible health impacts and safety risks of NMs, legal authorities request the toxicological analysis of SAS NMs by means of in vivo and in vitro studies [11–15].

Generally, ultrasonic dispersing or separation has been used for sample preparation of nanomaterials for safety assessment [16]. These studies can support the optimization of nanosynthesis or nano-applications for the sake of a “Green Synthesis of Nanomaterials”. One example for such

green nano-application of amorphous silica in entomology and parasitology as a nanopesticide has been considered safe for humans because of the specific mechanisms of action [17].

An important aspect of exposure and toxicological analyses is the characterization of NMs with respect to particle size. Most SAS occur in an aggregated state with particle sizes ranging from nanometer-sized primary particles to micrometer-sized aggregates or agglomerates [18–20]. However, the sample preparation for a specific *in vitro* test should consider the particle size-distribution which is of relevance for a given exposure pathway [21–24]. For example, inhalation of particles into the human respiratory tract leads to a fractionation of particles: larger agglomerates are deposited in the nasopharyngeal region (5–30 μm), small agglomerates are partially deposited in the tracheobronchial region (1–5 μm), and only small ($<1 \mu\text{m}$) and nano-sized materials (1–100 nm) may penetrate into the alveolar region of the lung [25–30]. Thus, toxicity testing of NMs using *in vitro* lung models demands the preparation of properly suspended samples under defined conditions.

In general, studies on environmental and health risk assessment focus on transport and deposition of NMs in real-life exposure scenarios. Both processes are governed by the mobility of aggregates and agglomerates, for which reason the size of aggregates and agglomerates needs to be measured. This is different from the nanomaterial definition recommendation of the European Commission, which is based on number-weighted distribution of the minimum size of isolated particles or constituent particles within aggregates and agglomerates [31,32].

The analysis of nanomaterials (NMs) and nanostructured materials requires standard operation procedures (SOPs) for the preparation of suspension samples to ensure defined granulometric states [33–35]. Therefore, the preparation and analysis of nanomaterial suspensions needs a high degree of standardization with respect to primary sample preparation (stock suspension), secondary sample preparation, conditioning (e.g., adjusting suspension composition or concentration), sample splitting and finally measurement/interpretation. All these steps need to be considered for the characterization of liquid-suspended powders and for the comparison of different SOPs in view of their reproducibility. Although wetting and low energy dispersion of SAS powders in the suspension are substantial components of (primary) sample preparation, further ultrasonic dispersing (USD) is needed as it is the most versatile method to disintegrate large particle agglomerates into small particle aggregates or primary particles. At the same time, stabilization and homogenization of the dispersed particles are necessary.

USD is a rather intense type of dispersing, which relies on the hydrodynamic stress caused by collapsing cavitation bubbles. As USD can be performed with different types of equipment (e.g., ultrasonic bath, high-power probe sonicator, or cup-horn sonication) methods ensuring its reproducibility are mandatory. Several studies have shown that the energy density (measured in J/mL) serves as a well-suited parameter for obtaining a largely identical degree of USD of nanomaterials among different laboratories [23,36,37]. Of note, the application of different ultrasonic dispersing methods (e.g., variation of the sonotrode geometry) and parameter settings (e.g., sonication time and vibration amplitude) requires the application of the energy density concept, applicable to ultrasonication, rotor-stator systems [37,38] or high-pressure dispersing [39–41]. In all these cases, the ultrasonic dispersing energy density can be used as a main parameter for comparing available and new sample preparation protocols or SOPs.

To achieve a reproducible particle size analysis, it is necessary to use adequate sample preparation techniques. The USD facilitates the disintegration of submicron agglomerates, which are therefore of special relevance for the preparation of stable and homogeneous distribution of particles in the suspension and contribute to achieve a stable dispersion. The particles are under interaction of different dispersion forces which control their random dispersion in the sample volume. The energy density has been used in Table 1 to compare several studies, which are developed for the application to nanostructured materials).

Table 1. Published protocols of ultrasonic dispersing (USD) specifically designed for application to nanostructured materials; characteristic parameters including the (range of) inserted energy density.

Protocol	Sample Volume	Dispersing Time	(Calorimetric) Energy Density
Tantra 2016 [42]; Pradhan 2016 [43]	6 mL	16 min	1176 J/mL
Rasmussen et al., 2013 [35]	15 mL	10 min	500–400 J/mL
	10 mL	16 min	2500 J/mL
Taurozzi et al., 2012 [44]	50 mL	5 min	300 J/mL
Jensen et al., 2011 [33]	6 mL	16 min	3140 J/mL
Bihari et al., 2008 [45]	1 mL	1 min	420 J/mL
Mandzy et al., 2005 [46]	-	Time frames (2 h)	5700 J/mL
Pohl et al., 2005 [37]	10–42 mL	17–630 s	400–30,000 J/mL
Pohl et al., 2004 [47]	3–6 mL	-	100–2000 J/mL

Table 1 shows that the energy densities used in several recent studies to disperse nanomaterials differ by more than two orders of magnitude. This raises the question, as to which extent they influence the results of particle size measurement. Previous studies on different grades of NMs such as SiO₂, Al₂O₃, or TiO₂ have shown that even with a comparatively high energy density of up to 5 kJ/mL a maximum dispersion cannot be achieved for all materials [20]. However, as administration of such energy densities requires extensive cooling of the samples and prolonged periods of ultrasonic treatment, we were seeking for a reasonable compromise to achieve an acceptable dispersion of nanomaterials.

In this paper, we examine the effect of USD energy on the dispersion of SAS and characterize resulting particle size distributions (PSD). Despite their identical chemical composition, SAS products show considerable variations with respect to the synthetic routes, particle morphology, and product properties. The synthesis of silica is realized either in aqueous solution based on sodium silicate solution or in gaseous phase from SiCl₄ [48,49]. The types of silica originating from silica synthesis processes in aqueous solution are silica gel (SG), precipitated silica (PS) and colloidal silica (CS). Fumed silica (FS), also referred to as pyrogenic silica, is synthesized from gaseous phase. SAS products are nanostructured NMs [46] as they consist of aggregates and agglomerates of nanosized constituent particles (FS, PS and SG, cf. [50]) or well-dispersed nano-objects (CS). Accordingly, the preparation of suspensions of FS, PS and SG requires defined dispersion procedures for their use, e.g., in toxicity studies, and characterization, whereas this is not really necessary for colloidal silica [19,20]. Furthermore, the different types of silica have a characteristic morphology due to their varied synthetic processes. This is an important issue that needs to be considered for comparison and data interpretation (e.g., reproducibility, effectiveness).

2. Materials and Methods

2.1. Materials

This study analyzed SAS products, representative for FS, PS, SG and CS, respectively. While FS, PS, SG were provided as untreated hydrophilic powders, CS was provided as an aqueous suspension. Important physico-chemical properties are summarized in Table 2.

Table 2. SAS properties.

SAS Type Internal Code	Fumed Silica F-3	Precipitated Silica P-2	Silica Gel G-1	Colloidal Silica C-1
BET ¹ (m ² /g)	300	440	700	200 ²
solid content for suspensions (wt.-%)	-	-	-	40
pH ³	5	6.5	4.4	9.7
electric conductivity (μS/cm) at 25 °C	4	160	55	471.6

¹ BET: Surface measured according to Brunauer, Emmet and Teller [51,52]. ² measured from freeze dried material.

³ suspended in ultra-pure water (1 wt.-%, 25 °C).

Figure 1a–c show SEM (JEOL Ltd, Tokyo, Japan) images of typical aggregates of FS, PS and SG [50]. In contrast, CS (Figure 1d) contains isolated spherical nanoparticles, which have gathered into an agglomerate-like structure upon drying on the TEM grid.

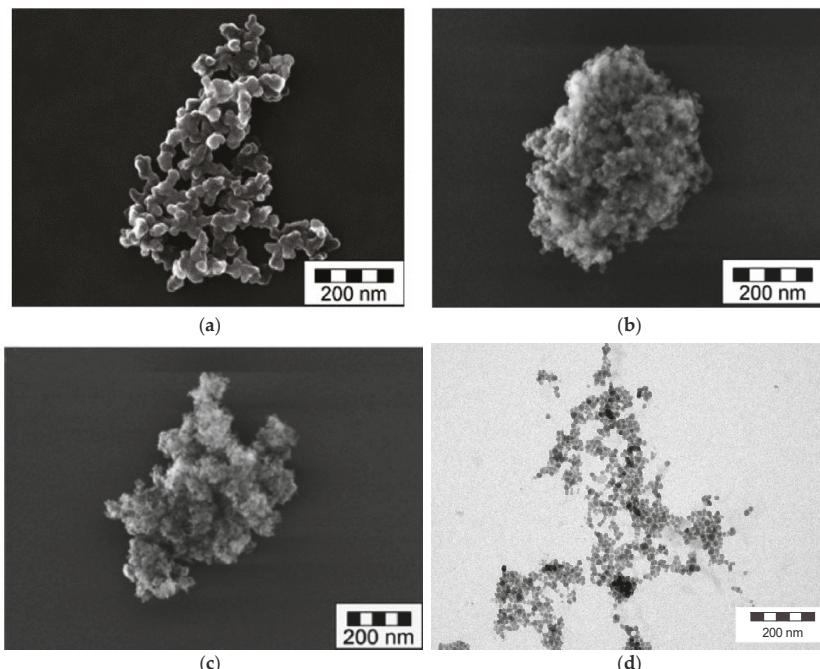


Figure 1. SEM and TEM images of different silica types. (a) fumed (pyrogenic) silica (opened fractal-like aggregates), (b) precipitated silica (compact fractal-like aggregates), (c) silica gel (compact and microporous fractal-like aggregates) and (d) colloidal silica (isolated spherical nanoparticles or small aggregates, here dried on TEM grid to opened agglomerates).

2.2. Instruments and Procedures for Sample Preparation

To prepare stock suspensions of SAS 1 wt.-% of silica powders (FS, PS, SG) were dispersed in 100 mL de-ionized water (18.3 MΩcm, 0.2 µm filtered). To avoid re-agglomeration, the particles were placed in a liquid environment that ensured high surface charges. The pH value of prepared silica suspension is far from the isoelectric point of silica (e.g., pH 1.8–2.5) and it has a low electric conductivity (see Table 2) [53–55]. Powders were dispersed by different treatments: Firstly, by means of a paddle stirrer (PS) (model RW 11 basic, IKA, Staufen, Germany), which administered the lowest input of mechanical energy into the suspension and which was used for the homogenization of SAS suspensions. Of note, the geometric size of the paddle stirrer and the sample beaker as well as stirring velocity ensured hydrodynamic equivalence to the “paddle apparatus” specified in Ph. Eur. 5.7. (2006) [56]. Secondly, by means of a turbulent shear rotor stator (RS) (Ultra-Turrax T25, IKA) which was used to achieve advanced dispersion and to investigate changes of PSD of different SAS types upon progressive dispersion energy. The RS provides shear forces, which cause shear stress on particle agglomerates. Thirdly, by means of immersion horns (three different instruments, see Table 3) which were used in most experiments [33–35]. The sonotrode or horn is in direct contact with the suspension and the dispersion effect is associated to cavitation, which occurs in highly intensive sound fields. Furthermore, the cavitation causes the formation of vapor cavities in a liquid (bubbles) which steadily grow to a critical size, at which they turn instable and implode [20]. This implosion produces

high temperature and rapid micro-jets, which exert mechanical stress to the particles close to the formed bubble [57]. This mechanical stress leads to the fragmentation or at the least to the erosion upon the direct contact with the imploding cavities [58,59].

The USD equipment consisted of three different instruments equipped with different sonotrodes (a few mm up to a few cm) (see Table 3), which were operated at frequencies in the range of 20 kHz to 100 kHz. Instruments had a nominal power consumption of a few Watt to approximately 1 kW [18]. The calorimetric energy input was measured at different dispersing instruments. Most of them work at frequencies in the range of 20 kHz. In pilot experiments, various tip diameters of the sonotrodes were selected according to the geometry of the beaker and the sample volume required by the sample preparation protocols.

Table 3. Technical data of ultrasonic dispersion instruments operated at approximately 20 kHz.

Model	Vibra-Cell 72412 ¹	UDS751 ²	SONIFIER 450D ³
Code	V	T	B
company	Sonics and Materials	Topas GmbH	Branson Ultrasonics
normal capacity (W)	600	200	400
tip diameter (\varnothing , mm)	13 19	3 7 14	5 13
amplitude (%)	0–100	0–100	10–100

¹ Vibra-Cell 72412 (Sonics & Materials, Newtown, CT, United States). ² UDS751 (Topas GmbH, Dresden, SN, Germany).

³ SONIFIER 450D (Branson Ultrasonic Corporation, Danbury, CT, United States).

For the particle size analysis of all silica types in this study one USD equipment was selected, the generator Vibra-Cell 72412 (Sonics and Materials; 20 kHz, nominal power: 600 W together with a 19 mm tip diameter solid probe. The tip of the probe was replaced for each series of dispersing experiments. USD was performed at maximum amplitude (100%) in a pulsed mode (2s:2s) with the probe being uniformly immersed in the sample. The same type of glass beaker was employed for all samples; the beakers were placed in cooled water during the USD. Even though, samples were steadily heated-up with ongoing USD, for which reason the USD was interrupted after a maximum of 4 min to cool down the complete the sample and the ultrasonic probe. This procedure ensured that the sample temperature stayed below 33 °C [20].

2.3. Instruments for Particle Size Analysis

USD leads to deagglomeration and disintegration of aggregates and the corresponding change in the granulometric state was quantified by laser diffraction (LD) and dynamic light scattering (DLS). However, these standard analytical techniques are based on mathematical models, which are not perfectly applicable to the examined particle systems, e.g., because they assume spherical particles (e.g., Stokes-Einstein relation and Mie theory for DLS), or because they do not cover the whole size range of broad distributions (e.g., Fraunhofer diffraction theory for LD) [18–20,60,61]. To avoid these technical limitations and for a better interpretation of measured data it is advantageous to apply both techniques to well-adapted dilutions of the same sample.

LD measurements were carried out with a HELOS KR (Sympatec, Clausthal-Zellerfeld, Germany) for angular ranges below 35° (i.e., forward scattering). Within this study an angular range of 0.1° to 9° (measurement range R3) was used, which is sensitive for particles of 0.5 μm to 175 μm, but which is insensitive to nanoparticles ($x \leq 100$ nm). These measurements were rather insensitive to small, weakly scattering particles $\ll 1$ μm.

To quantify particle sizes in sub-micrometer range, DLS measurements were conducted. The employed instrumentation, HPPS (Malvern, UK), bases on backscattered (173°) and sideward scattered light, respectively. Measured DLS signals (i.e., autocorrelation functions) were analyzed with inversion procedures, which compute complete size distributions, and cumulant analysis. The latter yields a polydispersity index, PDI, and a characteristic mean particle size, x_{cum} , which is the harmonic

mean of the intensity-weighted size distribution. The samples are filled in closed cuvettes (4 mL), which are placed in the temperature-controlled sample holder at least 15 min before the measurements.

2.4. Estimation of the Calorimetric Energy Input

The effectivity of the ultrasonication, in comparison with other dispersion procedures regarding size reduction, demands the necessity to evaluate the applied acoustic energy per unit suspension volume (E_V) [18,20,23,37,46]. The calculation of ultrasonic dispersion energy cannot be calculated directly. There are two ways to estimate the inserted acoustic energy. One way is the estimation from electrical energy consumption ($E_{V,\text{el}}$) and the other possibility is through the generated heat after implosion of bubbles ($E_{V,\text{cal}}$). The electrical energy consumption depends on many factors: transformation of energy of ultrasonic dispersions instruments (e.g., normal capacity, range of frequencies, types and probe diameter) that must be considered for comparison and validation. Furthermore, another important point to consider is the acoustic reflections inside the probe depending on, for example, sample volume, density and the propagation velocity of sound of liquid (acoustic impedance) [62].

This study uses the calorimetric energy input as decisive parameter for the effect of USD. Therefore, the USD devices needed to be calibrated regarding the calorimetric power input for the setup and settings (sample volume, horn diameter; nominal ultrasonic amplitude) employed in dispersing the SAS samples. The calibration comprises the evaluation of temperature increase by ultrasonication of a defined volume of de-ionized water. In this study, USD was conducted at 100 mL suspension sample placed in a 150 mL cylindrical borosilicate beaker. Please note that other studies worked with higher volumes (e.g., 500 mL, [23,34]) or smaller ones (e.g., 6 mL [33,42,43]). A further difference to other studies is that the beaker was placed in insulating foam, to minimize heat exchange with environment. The whole setup, including the ultrasonic probe as well as a thermometer with short response-time, was allowed to thermally equilibrate. Then ultrasonication was started and the temperature within the beaker was recorded as function of time.

The calorimetric energy input into a suspension sample by ultrasonication can be calculated from the heat production rate P_{cal} which is a function of the dispersion time t_{disp} . The former parameter is valid for defined conditions of ultrasonication, i.e., for defined sample properties and USD settings. It can be determined by specifically designed experiments that measure the temperature increase when sonicating the particle-free dispersion medium (e.g., water). Furthermore, it is necessary to consider the mass and specific heat of the beaker to have a correct determination of heat production and to evaluate the initial slope of temperature increase, which needs to be less than 4 Kelvin ($\Delta T \leq 4$ K). Furthermore, the following proposed equation to estimate the calorimetric energy input assumes that the sonotrode has a zero heat capacity and that there is no heat exchange with the environment:

$$P_{\text{cal}} = (m_w c_{p,w} + m_b c_{p,b}) \cdot \Delta T / t_{\text{disp}} \quad (1)$$

where m_w denotes the mass of liquid, m_b denotes the mass of beaker, $c_{p,w}$ specific heat of water, $c_{p,b}$ specific heat of beaker, ΔT the temperature increase and t the dispersion time [63,64]. The calorimetric energy density can be determined as follows:

$$E_{V,\text{cal}} = \frac{P_{\text{cal}} \cdot t_{\text{disp}}}{V} = (m_w c_{p,w} + m_b c_{p,b}) \cdot \Delta T / V \quad (2)$$

with V being the suspension volume.

The energy density is considered to be a most important process parameter when dispersing suspensions and emulsions. Frequently, a power-law relationship can be established between average particle size x and energy density $E_{V,\text{cal}}$ (e.g., [47]):

$$\bar{x} \propto E_{V,\text{cal}}^{-b} \quad (3)$$

where “average particle size” can be any characteristic distribution parameter (e.g., median size or arithmetic mean) referring to a defined type of quantity, in which the size distribution is weighted (e.g., number, volume or scattering intensity). The exponent b describes the material’s dispersibility under the specified conditions.

Finally, there are some assumptions to estimate values for calorimetric calibration of ultrasonic instruments. The first assumption is the uniform temperature of water and beaker, second—zero heat capacity of ultrasonic probe and third—no heat transfers out of the system beaker-water.

3. Results and Discussion

3.1. Calorimetric Calibration of Probe Sonication

While the ultrasonic wave propagates through the dispersion medium, its energy is absorbed and converted into heat [65,66]. Figure 2 exemplifies this progressive heating during USD for a non-insulated beaker. The images give an impression of how the heat generated upon ultrasonication is transferred from the sonotrode into the surrounding medium and associated masses. Therefore, it is important to consider some issues such as sample volume, or isolating foam to achieve a correct calorimetric calibration of probe sonication, as explained in the proposed protocol (see Section 2.4).

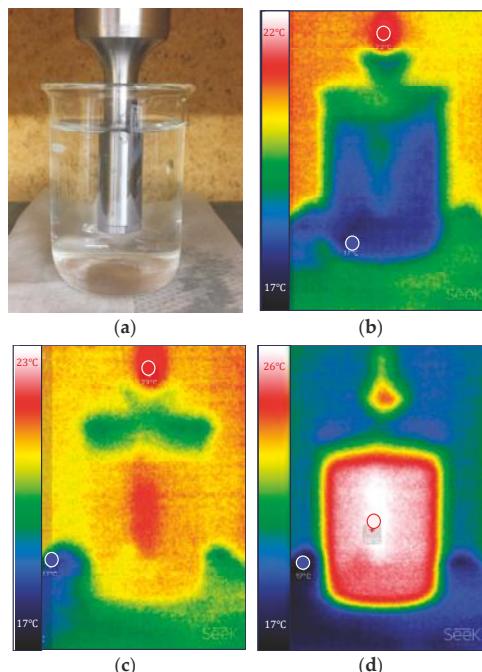


Figure 2. Progressive heating of the sonotrode, sample fluid, and beaker in the course of USD. Images were captured with a thermal imaging camera (Seek Thermal) after different dispersing periods; (a,b): 0 s, (c): 30 s, (d): 60 s. Temperature (in degree Celsius) is shown on a pseudo-color scale whose range was automatically adjusted to the temperature peak (T_{\min}/T_{\max}) of the measurement (i.e., (a) 17 °C/22 °C, (b) 17 °C/23 °C, (c) 17 °C/26 °C).

The calorimetric calibration was calculated for different ultrasonic dispersing instruments (according to the protocol in Section 2.4) to compare their heat production after setting of different parameters (see Section 2.2). The calorimetric measurements delivered the results in Figure 3.

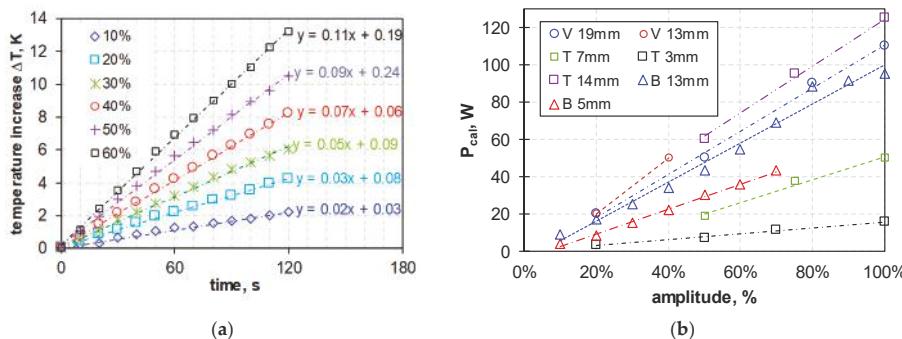


Figure 3. Calorimetric calibration curves of ultrasonic dispersing instruments. (a) temperature increases (ΔT) in Kelvin (K) over time, obtained with a 13 mm tip diameter sonotrode (Branson 450D) and increasing amplitudes (10–60% from maximum, as indicated in the graph). (b) Heat production (P_{cal}, W) of the different dispersing instruments and sonotrode geometries Vibra-Cell 72412 (V), Topas UDS751 (T) and Branson SONIFIER 450F (B) (outlined in Table 3) as a function of increasing amplitude (in % from maximum).

3.2. Sample Preparation by Probe Sonication

3.2.1. Impact of USD on Particle Size Distribution of SAS

Previous studies have shown that the PSD of, e.g., FS may be highly polydisperse [20,54] and covers a wide range from a few nanometers up to several micrometers [18–20]. As outlined above (Section 2.3), the appropriate granulometric analysis of such samples requires a combination of LD and DLS. This section addresses the effectiveness of ultrasonic dispersing on particle size of silica types measured with both techniques.

Suspensions of the SAS types were prepared by a combination of dispersing procedures and defined calorimetric energy densities (EV, J/mL). Depending on the instrumentation (see Section 2.3) dispersing energies were weak from propeller stirrer (PS), moderate from rotor-stator (RS), or intense from ultrasonication (US, different dispersing energies) and the selected calorimetric energy densities' values are based on the calorimetric calibration of probe sonication (see Figure 3). The calorimetric energy density of the US treatment ranged from 8 J/mL through 18 J/mL and 270 J/mL to a maximum energy input of 1440 J/mL. This stepwise increase of energy input allows for a comprehensive characterization of SAS with respect to particle size and morphology.

Figures 4–6 show the LD measurements of the PSD results of the transformed distribution density of upon increasing $E_{V,\text{cal}}$ for PS, FS, and SG (see Section 2.1). The transformed distribution density represents in accordance with ISO 9276-1:1998 provide the differential size distribution on a log scaled abscissa. Areas under the curve represent the volume portions of the size classes [67]. The result of a long-term sedimentation of silica suspensions (after five months) is shown in parallel.

Precipitated silica (PS, 400 m²/g, 1 wt.-%) shows a clear tendency of deagglomeration by increasing dispersion energy (Figure 4a). Especially the presence of coarse, micrometer-sized agglomerates was diminished, and this effect started mostly at 18 J/mL. In line with these results the sedimentation profiles of ultrasonicated PS suspensions show an increasing degree of opacity upon 270 J/mL and 1440 J/mL (Figure 4b), suggesting the presence of slowly settling, light scattering particles in the sub-micrometer range. However, although the zone of opacity was wider upon 1440 J/mL, the volume of the white matter at the bottom was similar.

Largely similar to PS was fumed (pyrogenic) silica (FS, 300 m²/g, 1 wt.-%) deagglomerated by increasing USD energy (Figure 5a). The deagglomeration of large particles was achieved already with a moderate dispersion (RS) and increased further upon administration of USD energy. Interestingly,

the mono-modal PSD of FS became bimodal upon 18, 270, and 1440 J/mL, such that a fine and a large fraction could be distinguished (Figure 5a). Of note, the large fraction generated by 270 J/mL and 1440 J/mL comprised larger particles as compared to the fraction induced by 18 J/mL. This suggests that elevated energy levels at least in part can provoke an agglomeration of FS. The effect was most obvious at an USD energy of 270 J/mL. These results confirm previous studies, where the coarser particles appeared at energy density levels above 171 J/mL [20]. The sedimentation profiles of FS at higher dispersion energies showed a similar degree of opacity in the supernatant, but an increased volume of the white sediment. This suggests that high USD energy leads to the formation of larger particles from the same mass of FS (see Figure 5b).

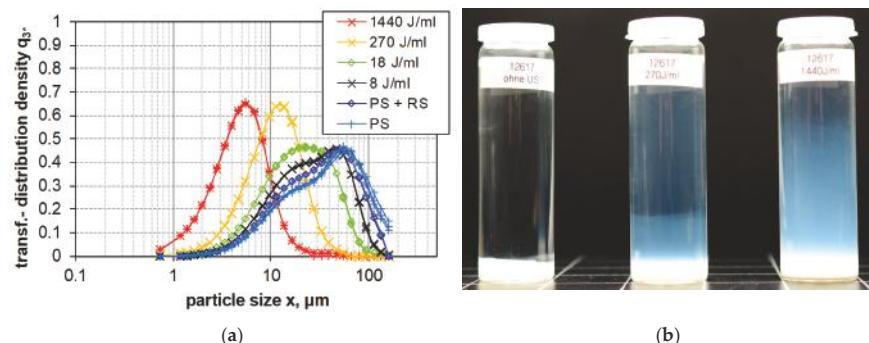


Figure 4. Particle size distribution of PS ($440 \text{ m}^2/\text{g}$) dispersed by different procedures. (a) Size distribution measured by laser diffraction spectroscopy plotted as the transformed distribution density (q^{3*}). Weak (PS), moderate (RS) and intense (US) dispersion with increasing ultrasonic energies (indicated in the diagram) were employed. (b) Silica suspensions after five months (left 1 J/mL, middle 270 J/mL, and right 1440 J/mL).

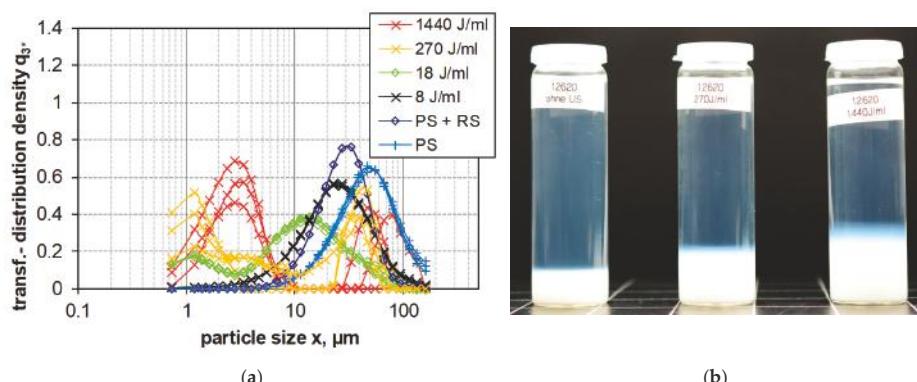


Figure 5. Particle size distribution of FS ($300 \text{ m}^2/\text{g}$) dispersed by different procedures. (a) Size distribution measured by laser diffraction spectroscopy plotted as the transformed distribution density (q^{3*}). Weak (PS), moderate (RS) and intense (US) dispersion with increasing ultrasonic energies (indicated in the diagram) were employed. (b) Silica suspensions after five months (left 1 J/mL, middle 270 J/mL, and right 1440 J/mL).

As can be seen in Figure 6a, the PSD of silica gel (SG, $700 \text{ m}^2/\text{g}$, 1 wt.-%) remained unchanged upon increasing energy density. SG consists of compact (dense) and microporous fractal-like aggregates (see Figure 1), which appear to be insensitive to high USD energies and undergo a rapid sedimentation of particles (by median $x_{50,3} = 6 \mu\text{m}$), irrespective of USD treatment (see Figure 6b), although both

ultrasonic treatments led to a similar degree of opacity of the supernatant which may be indicative of smaller particles (not measurable by LD).

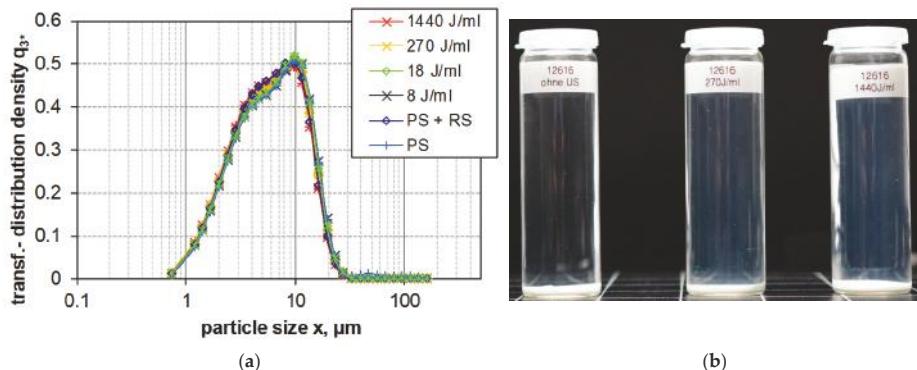


Figure 6. Particle size distribution of SG (BET: $700 \text{ m}^2/\text{g}$) dispersed by different procedures. (a) Size distribution measured by laser diffraction spectroscopy plotted as the transformed distribution density (q_3^*). Weak (PS), moderate (RS) and intense (US) dispersion with increasing ultrasonic energies (indicated in the diagram) were employed. (b) Silica suspensions after five months (left 1 J/mL, middle 270 J/mL, and right 1440 J/mL).

The effect of increasing dispersion energy on particle size of PS, SG and FS, as measured by LD, is compared in Figure 7, which shows the trend analysis for the $x_{50,3}$ and $x_{99,3}$ quantiles of the volume weighted size distribution. Whereas the particle size of SG is not altered by increased dispersion energy, particle size of PS is constantly lowered. In the case of FS, only the $x_{50,3}$ value reflects the decrease in particle size, whereas the $x_{99,3}$ is inconsistent and shows an upwards trend demonstrating the coarsening or reagglomeration upon high USD energy. Figure 8 shows a comparison the stability of intensely dispersed silica types at 1440 J/mL after long-term sedimentation (5 months). The different sediment degrees of silica types support the LD results and provide a subjective information about sedimentation velocity in the gravitational field of the earth.

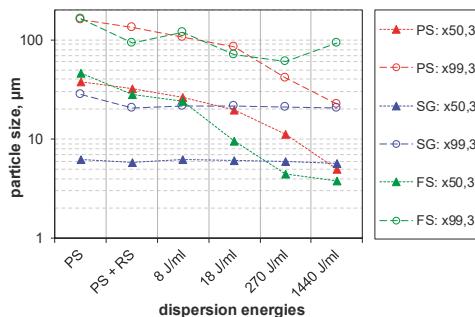


Figure 7. Particle size distribution of silica suspension after administration of increasing dispersion energy as measured by laser diffraction. Curves show the trends for the $x_{50,3}$ and $x_{99,3}$ quantiles of the volume weighted size distribution. USD energy density of ultrasonic treatment is indicated in J/mL.

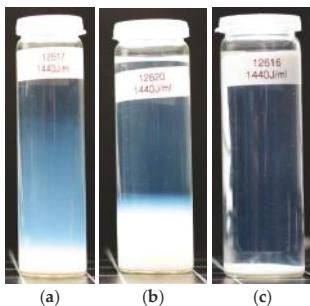


Figure 8. Intensely dispersed SAS samples (USD, $E_{V,\text{cal}}$: 1440 J/mL) after long-term sedimentation (5 months): (a) PS; (b) FS; (c) SG.

Since LD is not sensitive for silica particles smaller than 1 μm [68], the characterization of sub-micrometer particles (1 nm–10 μm) was carried out with DLS. Data of three silica types (FS, SG, PS) was expressed as intensity-weighted size distribution, using the characteristic values mean size (x_{cum}) and the polydispersity index (PDI) obtained by cumulant analysis. Figure 9a compares the as result calculated logarithmic normal distribution (LND) and shows the impact of 270 J/mL and 1440 J/mL on particle size. In Figure 9b the Intensity-weighted transformed distribution density functions are shown. While the particle size of SG and FS remains nearly unchanged in the lower size range, the long-term sedimentation e cumulative size distribution curve of PS is shifted leftwards, indicating that particle size had shifted to submicron and nano range (<1 μm). Fumed silica (FS) shows a minimal tendency to increase the size of submicron particles upon high energy density.

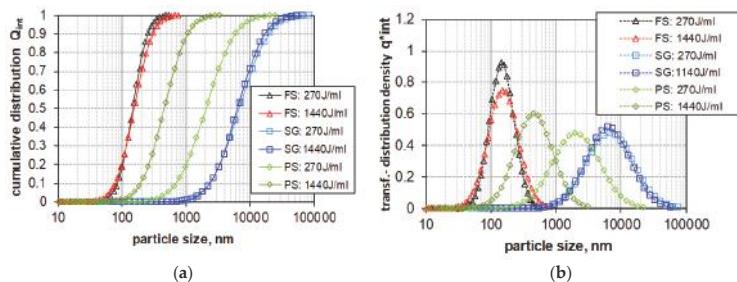


Figure 9. Particle size distribution (PSD) of amorphous silica suspensions dispersed by two different ultrasonic dispersion energy densities (270 and 1440 J/mL). PSD was measured by dynamic light scattering (DLS). FS: fumed silica, SG: gel silica, PS: precipitated silica. (a) Intensity-weighted sum functions for different energy density values. (b) Intensity-weighted transformed distribution density functions for different energy density values. Logarithmic normal distribution.

Figure 10a shows the granulometric state of colloidal silica (CS) over the full range of dispersion energies as used in Figure 7. While stirring (PS, PS + RS) had no effect on particle size, ultrasonic treatment surprisingly led to a larger and broader PSD indicated by an increase in hydrodynamic size and PDI. The effect started at a low energy density of 8 J/mL and was found to be strongly augmented upon higher USD energies. We found that wear particles from the sonotrode's tip, the larger of which appeared as a sediment at the bottom of the vial (Figure 10b), made a major contribution to this effect. Due to the strong light scattering properties of such metal particles (compared to the small and weakly scattering CS particles), even low amounts of wear particles contaminate the light optic measurements. If this increase of the mean particles size would be caused

by agglomerated colloidal silica particles, they would be visible as a sediment layer after 5 months. In the case of silica gel scattering, intensity of the micrometer particles hides the contamination signals during measurement whereas, after settling, the contamination is embedded in the silica sediment.

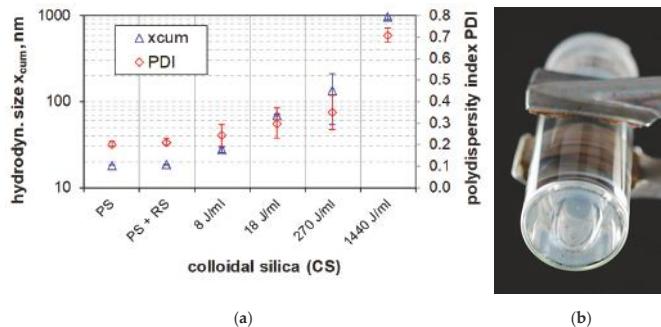


Figure 10. Particle size distribution of colloidal silica dispersed by different procedures measured by DLS. (a) Mean particle size (x_{cum}) and corresponding polydispersity index (PDI) as determined with cumulant analysis. (b) Bottom view of a vial with an intensely dispersed colloidal silica sample ($E_{V,cal}$: 1440 J/mL) taken after 5 months of gravitational settling.

3.2.2. Sample Contamination with Probe Sonication

Figure 11a,b show the effect of high dispersion energy on the sonotrode's tip if delivered over prolonged period. Wear particles ablated from the tip contaminate suspension and can be observed as a black sediment (i.e., coarse titanium particles) and/or as a well as a grey discolouration of the suspension (Figure 11c). Figure 12 shows a SEM picture of sonotrode abrasion particles collected from the bottom of silica suspension. As shown in a previous study, the abrasion of the ultrasonic probe and sample contamination occurs in the moment of ultrasonication; the number of particles increases linearly with time [20]. Furthermore, it was shown that in suspension of pyrogenic silica (PS, 1 wt.-%) at a dispersion energy >171 J/mL sonotrode wear particles contribute to the PSD and interfere with the sample analysis by LD [20]. Nevertheless, this widespread sonicator type is superior to other indirect sonicator types (e.g., ultrasonic bath, cup horn) [36,69] due to its high effectiveness of USD and with regard to the best possible disintegration of agglomerates and aggregates in a short time [20]. Therefore, a restriction of the USD energy appears reasonable. To remove larger particles at lower dispersion energy, e.g., from PS suspensions, we developed a dispersion protocol combining stirring, USD and filtration steps (see Section 3.3).

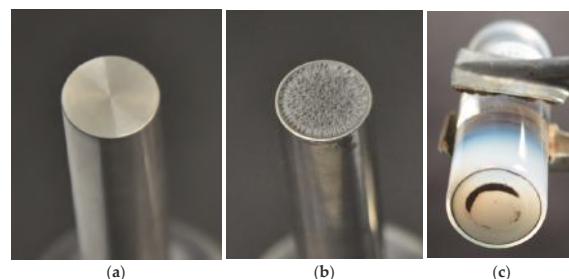


Figure 11. Comparison of new and used sonotrode with sonotrode abrasion as a consequence of long ultrasonic dispersion time (a,b); sonotrode abrasion sediment on the bottom of a precipitated silica suspension sample after high USD energy (c) (i.e., $E_{V,cal}$: 1440 J/mL).

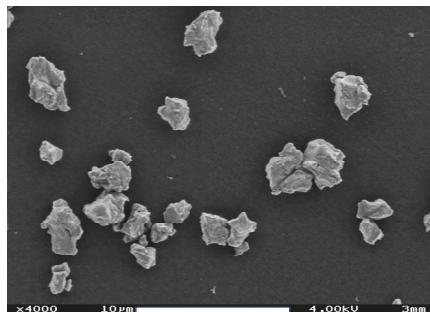


Figure 12. Scanning electron microscope image of wear particles from the sonotrode tip. Abrasion particles were collected from the sediment of a silica suspension sample after high USD energy (i.e., $E_{V,\text{cal}}$: 1440 J/mL).

3.3. Sample Preparation with Size-Selective Filtration

Testing the in vitro toxicity of nanomaterials requires that the size distribution of particles in cell culture media is well defined. With respect to inhalation exposure, which may be tested by the alveolar macrophage assay [24], larger non-respirable coarse particles need to be removed so that a mass-per-volume- or surface-per-volume dose metrics can be applied [15]. Ideally, particle size distribution should reflect inhalable fractions with aerodynamic diameters smaller than 4 μm . However, as outlined above for paddle stirring (PS), this would require high ultrasonic energy and bears the risk of metal particle contamination (see Figures 11 and 12). To circumvent this risk, a size classification by controlled filtration (with 100% fines penetration) was developed, using a commercially available nylon gaze with a pore size of nominally 5 μm (Bückmann, Germany). Figure 13 shows the grade efficiency function $T(x)$ demonstrating that glass spheres below 7 μm can freely permeate the filter, whereas spheres larger than 15 μm were retained.

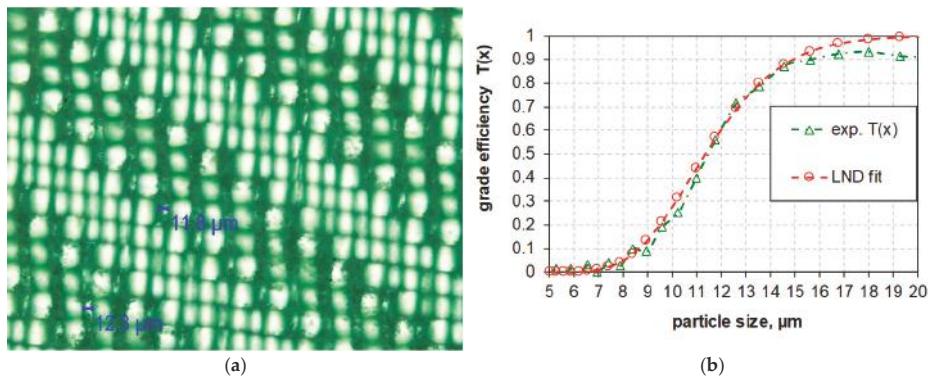


Figure 13. (a) Light microscope image of polymer gaze for size-selective suspension filtration and (b) $T(x)$: grade efficiency function after gaze test filtration with glass spheres.

The theoretical volume weighted cumulative distribution function of the filtrate $Q_{3,f}(x)$ was derived from the experimentally determined grade efficiency function $T(x)$ and the feed distribution $Q_{3,i}(x)$:

$$\Delta Q_{3,f}(x) = (1 - T(x)) \cdot \left(\frac{\Delta Q_{3,i}(x)}{1 - \sum T(x) \cdot \Delta Q_{3,i}(x)} \right) \quad (4)$$

In the next step ultrasonication and filtration were combined to prepare a suspension suitable for *in vitro* testing. The method is shown exemplarily for the PS used in this study: the powder was suspended (1 mg/mL) in de-ionized water by means of a magnet stirrer (700 rpm, 10 min). Thereafter the PS sample was filtrated by gaze filter with a grade efficiency curve shown in Figure 13b and a cut-off size of 11.5 μm . The filtrate of the silica sample was then dispersed with $E_{V,\text{cal}}$: 18 J/mL and $E_{V,\text{cal}}$: 270 J/mL. Results were expressed in Figure 14a as cumulative particle volume curves (Q3, green curve) and compared to the effects of progressive dispersion energy on the size distribution of non-filtered PS (red curve).

Figure 14a shows the advantage of combined moderate ultrasonication by 270 J/mL with filtration (green line) to remove particles larger than 11 μm in comparison with the high ultrasonic energy result (red line), including possible sample pollution. LD results (Figure 14b) show that low energies (i.e., weak dispersion (PS), moderate dispersion (RS) without ultrasonication) leave a considerable amount of micrometer-sized agglomerates in the suspension, whereas ultrasonic dispersion with 8, 18, 270 and 1440 J/mL progressively reduced micrometer-sized agglomerates.

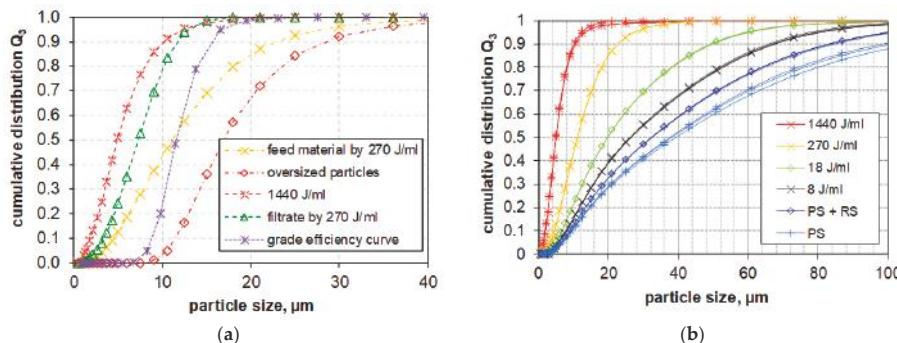


Figure 14. (a) filter grade efficiency curve and calculated results for penetrated (filtrate) and retained particle size distributions from feed size distribution (yellow) in comparison to not filtered but with 1440 J/mL dispersed sample (b) Evolution of PSD during ultrasonication of precipitated silica ($440 \text{ m}^2/\text{g}$); cumulative distribution functions measured by LD.

4. Conclusions

The effect of dispersion energy on particle size distribution of nanomaterial suspensions depends not only on a defined dispersion procedure (e.g., dispersion time, sample volume) but also on the silica types (e.g., morphology).

Ultrasonic dispersion energy density is a main parameter for comparability of sample preparation protocols. Sonication is limited by sample pollution with wear particles from the probe. Therefore, upper limit dispersion energy density values must be determined. In the case of silica it is recommended to apply dispersion energy density only up to 300 J/mL.

The resulting particle size distributions strongly depend on the type of silica. Fumed SAS reach PSDs in the submicron range even at low values of ultrasonic energy density; continued sonication leads to a steady, yet slight size reduction. Gel and colloidal SAS are hardly or even adversely affected by increasing ultrasonic dispersion energies. The PSDs of precipitated SAS strongly depends on the increasing ultrasonic dispersion energy, changing constantly to smaller sizes.

Additional size-selective filtration can remove the large and settling particles without the risk of sample contamination by too high ultrasonic energy dispersion. A combination with the above-described ultrasonic dispersion provides a general SOP for the preparation of well-defined suspensions of SAS nanoparticles for *in vitro* toxicological tests.

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Article

Facile and Robust Solvothermal Synthesis of Nanocrystalline CuInS₂ Thin Films

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Abstract: This work demonstrates that the solvothermal synthesis of nanocrystalline CuInS₂ thin films using the amino acid L-cysteine as sulfur source is facile and robust against variation of reaction time and temperature. Synthesis was carried out in a reaction time range of 3–48 h (at 150 °C) and a reaction temperature range of 100–190 °C (for 18 h). It was found that at least a time of 6 h and a temperature of 140 °C is needed to produce pure nanocrystalline CuInS₂ thin films as proven by X-ray and electron diffraction, high-resolution transmission electron microscopy, and energy-dispersive X-ray spectroscopy. Using UV-vis spectroscopy, a good absorption behavior as well as direct band gaps between 1.46 and 1.55 eV have been determined for all grown films. Only for a reaction time of 3 h and temperatures below 140 °C CuInS₂ is not formed. This is attributed to the formation of metal ion complexes with L-cysteine and the overall slow assembly of CuInS₂. This study reveals that the reaction parameters can be chosen relatively free; the reaction is completely nontoxic and precursors and solvents are rather cheap, which makes this synthesis route interesting for industrial up scaling.

Keywords: solvothermal synthesis; CuInS₂; TEM

1. Introduction

Due to the growing energy needs of our society, the scarcity of fossil fuels, and threatening greenhouse effect, research on materials that offer appropriate functionalities to overcome these problems is desperately needed; this is therefore a very active research field. Possible applications involve the generation of electricity via solar energy, the production of alternative fuels like hydrogen, and the decomposition of contaminants; but involves also research on how to store the produced energy [1–7]. For example, the photosynthesis of plants is mimicked to split water by light [8,9] or to convert CO₂ into less-harmful compounds [10]. To keep the costs economic, the synthesis of green energy materials should also be a green synthesis. Such a synthesis should be feasible without the need for expensive precursors, high pressures, or temperatures (i.e., high energy input), using a route that tolerates deviations in temperature and time, ideally accomplished in only one synthesis step, and should also avoid toxic chemicals during preparation, [11,12]. The use of biomolecules as precursors in chemical reactions and the formation of nanomaterials for diverse applications has been actively investigated in recent times [4,6,12–17].

Copper indium disulfide, CuInS₂, is a material suitable for diverse solar-driven applications [18–20]. It offers a band gap of 1.5 eV for the bulk, a high-absorption coefficient ($\alpha = 10^5 \text{ cm}^{-1}$) [21] and can be used to convert sunlight into electricity or as a photocatalyst. CuInS₂ can be fabricated with various techniques which, in most cases, require high temperatures, high pressure, and clean precursor metals, but also with wet-chemical approaches [22–25]. We chose the solvothermal route to prepare CuInS₂, as it can be considered a green synthesis route—it only uses simple solvents and metal salts, while achieving

a wide variety of nanostructures at low-reaction temperatures [26,27]. Furthermore, the solvothermal growth allows for the direct, one-pot CuInS₂ deposition on a suitable substrate—growing (thin) films in-situ without the need to deposit synthesized material afterwards [28].

Peng et al. [28] developed a solvothermal synthesis strategy for growing CuInS₂ films directly on fluorine-doped tin oxide (FTO) using simple salts as precursors. This synthesis strategy has been used and slightly modified in our group to prepare CuInS₂ thin films as well as microspheres [29,30]. However, this synthesis route involves the carcinogenic substance thioacetamide as sulfur source [31]. Therefore, we recently changed the sulfur source to the natural amino acid L-cysteine to achieve a complete non-toxic, green synthesis pathway towards CuInS₂ films on FTO substrates [32]. There, we varied the concentration and ratio of the used precursor salts to investigate the influence on sample morphology and properties, while keeping the reaction conditions constant (150 °C, 18 h). In short, at high sulfur ratios, an additional nanoflake layer of In₂S₃ on top of a compact CuInS₂ film was observed [32]. On reviewing literature, it becomes clear that it is possible to synthesize CuInS₂ thin films in a wide variety of reaction conditions and with many possible precursor salts and solvents; e.g., Peng et al. [28] used CuSO₄, InCl₃ and thioacetamide in ethanol at 160 °C for 12 h to produce pure CuInS₂. Wochnik et al. [29] were also able to synthesize pure tetragonal CuInS₂ films but at 150 °C for 24 h using the same precursors and solvent. Furthermore, it is possible to produce pure and stoichiometric CuInS₂ thin films with CuCl₂, In(NO₃)₃, thiourea, CTAB and oxalic acid in ethanol at 200 °C for 24 h as demonstrated by Xia et al. [33]. The synthesis from Cu₂O, In(OH)₃, thioacetic acid and ammonia in ethanol at 150 °C for 6 h is also possible, as shown by Liu et al. [34]. Solvothermal synthesis of nanostructured CuInS₂ using L-cysteine as a sulfur source has already been published in literature [35,36]. Liu et al. [35] reported about the formation of CuInS₂ using CuCl₂, InCl₃ and L-cystine in 1:1 ethylene diamine: water. They kept the autoclave at 200 °C for 12 h, resulting in microspheres and nanoparticles in the tetragonal Chalcopyrite structure without any visible impurities or side products. The composition of their CuInS₂ samples was also in a stoichiometric range. Wen et al. [36] synthesized CuInS₂ microspheres out of CuCl₂, InCl₂ and L-cysteine in *N,N*-dimethylformamide (DMF) as a solvent, also at 200 °C for 12 h. Their product displayed the tetragonal Chalcopyrite modification as well.

The results in literature indicate that pure, crystalline CuInS₂ nanostructures can be fabricated within a relative large reaction window; however, systematic studies, where the reaction temperature and time are varied in a broad range while keeping all other parameters constant, are rare. Additionally, many of the already existing synthesis routes involve toxic substances as raw materials.

In the present work, we fill that gap and focus on the influence of the reaction temperature and time on the L-cysteine-assisted solvothermal growth of nanocrystalline CuInS₂ films. We show that the solvothermal synthesis of CuInS₂ using L-cysteine as sulfur source is not only non-toxic but also extremely robust over a large temperature range from 140 °C to 190 °C as well as less critical on large time variations from 6 to 48 h. Thus, this synthesis pathway is very interesting for possible industrial utilization.

2. Materials and Methods

2.1. Synthesis of CuInS₂ Films

Chemicals were used as-purchased from Sigma-Aldrich (Sigma-Aldrich Chemie GmbH, Munich, Germany) without further purification. The FTO glass substrates (Sigma-Aldrich) were cut into pieces of 15 mm × 20 mm × 2 mm, cleaned ultrasonically in dilute nitric acid, double-distilled water, acetone and ethanol for 5 min each prior to synthesis. The films were grown with our recently reported synthesis strategy using L-cysteine as sulfur source [32], which is based on the method published by Peng et al. [28] and our group [29] where thioacetamide was used.

The procedure is as follows: CuSO₄·5H₂O (0.2 mol, 0.050 g) and InCl₃ (0.2 mol, 0.044 g) were weighed out directly into a Teflon liner (20 mL capacity) and dissolved in 10 mL ethanol. The mixture was stirred for 10 min after which L-cysteine (0.5 mol, 0.061 g) was added. After stirring for another

5 min, a piece of FTO was placed inside the Teflon liner, conducting side facing down, the stainless-steel autoclave was sealed and put into an electric oven. There it was kept for 3, 6, 9, 12, 15, 18, 21, 24 and 48 h at a temperature of 150 °C and for 18 h at temperatures of 100, 120, 140, 150, 160, 180 and 190 °C. That means that for a variation of the reaction time, the reaction temperature was kept constant at 150 °C and for a variation of the reaction temperature the reaction time was always 18 h. The ratio between the precursors was kept at Cu:In:S 1:1:2.5. The film grown at this concentration/ratio and at 150 °C for 18 h has been published before [32] and will be referred to as film_S (standard reaction conditions). The other films are named according to their variation in time or temperature as film_time or film_temperature.

2.2. Characterization

To investigate the crystal structure of the synthesized CuInS₂ films on a global scale, X-ray diffraction (XRD) was used. To minimize the contribution from the FTO substrate, the measurements were performed under grazing incidence geometry with an incident angle of $\alpha = 2^\circ$ in a Seifert THETA/2THETA X-ray diffractometer. The diffractometer was equipped with a Co source ($\lambda K\alpha = 1.79 \text{ \AA}$), polycapillary beam optics and an energy dispersive point detector. The 2θ values ranging from 10° to 140° were measured with a step size of 0.05°/s and a count time of 30 s/step. The X-ray generator was operated at 40 kV and 30 mA. Literature data were used to identify the obtained phases. To calculate the average crystallite size, the Scherrer equation [37] was applied, fitting the most intense CuInS₂ peaks (112) and (204) with a Gaussian function.

The morphology of the CuInS₂ films was evaluated using scanning electron microscopy (SEM). For this purpose, a ZEISS Merlin, operated at 5.0 kV and a probe current of 2.0 nA, was used. Imaging was performed using the attached InLens® ZEISS standard detector. To analyze the chemical composition, energy-dispersive X-ray (EDX) spectroscopy using the XFlash detector 6130 was done with an acceleration voltage of 20.0 kV and a probe current of 4.0 nA. Quantification was done using the Cliff-Lorimer equation. The intensities of the element-specific X-ray lines were determined by using Gaussian functions. The k-factors were calculated using the Bruker software. The results were normalized to Cu. In the case of thin films on substrates, the spectrum can also contain signals from the substrate as a result of the large interaction volume when using high acceleration voltages in SEM. For example, the In L line from CuInS₂ and the Sn L line from the substrate FTO (SnO₂:F) overlap and complicate the quantification of In (compare also our recent publication) [38]. Nevertheless, EDX measurements have been performed in the SEM at 20 kV acceleration voltage but In is not considered for the analysis and only the ratio between Cu and S is given (stoichiometric ratio Cu:S for CuInS₂ should be 1:2).

The film thicknesses were measured by focused ion beam (FIB) sectioning on a FEI Helios Nanolab 600. Cuts were performed at sample areas coated with conductive silver paint to avoid destruction of the film surface.

To conduct UV-vis measurements of the CuInS₂ films a Perkin Elmer Lambda 800 in transmission mode has been used. Spectral range was from 260 nm to 900 nm with a step size of 1 nm. From the UV-vis data band gaps were calculated using the Tauc method for direct band gap semiconductors [39]. The energy was plotted vs. (energy · absorption)² and the first linear slope was fitted and the intersection with the x-axis calculated.

For in-depth characterization of the films (scanning), transmission electron microscopy ((S)TEM) was used. Measurements were performed on a FEI Titan Themis 300 (S)TEM at 300 kV acceleration voltage. The (S)TEM is equipped with a Cs probe corrector, a Gatan Quantum ERS energy filter, and a Super X-EDX detector from Bruker. Electron diffraction data, calibrated with the help of a Si standard, were evaluated by comparing the results to literature data. (S)TEM scratch samples have been prepared to avoid an influence of the sample preparation on the crystallinity and composition of the investigated films (compare a recent publication of our group) [38]. As mentioned for EDX in SEM, quantification was done using the Cliff-Lorimer equation with the help of the Bruker software and normalizing the results relative to Cu. In STEM mode, several EDX maps have been recorded

(\approx 6 maps per sample) and quantification of the Cu:In:S ratio was done on \approx 10 areas of each map with each area \approx 100 nm² and calculating the average value.

3. Results

3.1. Reaction Time

Top-view secondary electron SEM images of CuInS₂ films, solvothermally synthesized for different reaction times at 150 °C, are shown in Figure 1 and in the Supporting Information (Figure S1). At first sight, all the films show a very similar surface topology. Only film_3 h, synthesized with the shortest reaction time of only 3 h at 150 °C, seems to be composed of individual agglomerates that grow on the FTO surface. Between these small agglomerates, the substrate is still visible, indicating an incomplete coverage after 3 h of solvothermal reaction (see Figure 1). With increasing reaction time, the agglomerates seem to grow laterally until they cover the underlying FTO substrate completely and form a compact CuInS₂ layer (for cross-sectional views see later images and the supporting information, Figure S2). After the compact layer has formed, more CuInS₂ agglomerates deposit on top of it. For some of the films, small nanoflakes growing out of the agglomerates can be observed (exemplarily marked in Figure 1). Nevertheless, large changes in the surface morphology of the films cannot be seen. From 6 h reaction time, the FTO substrate is not visible anymore in top-view and no cracks or delamination of the film is observed in cross-sectional SEM micrographs (see Figure S2 in the supporting information and also Figure 4a,b), implying a good homogeneity and adhesion of the films to the substrate.

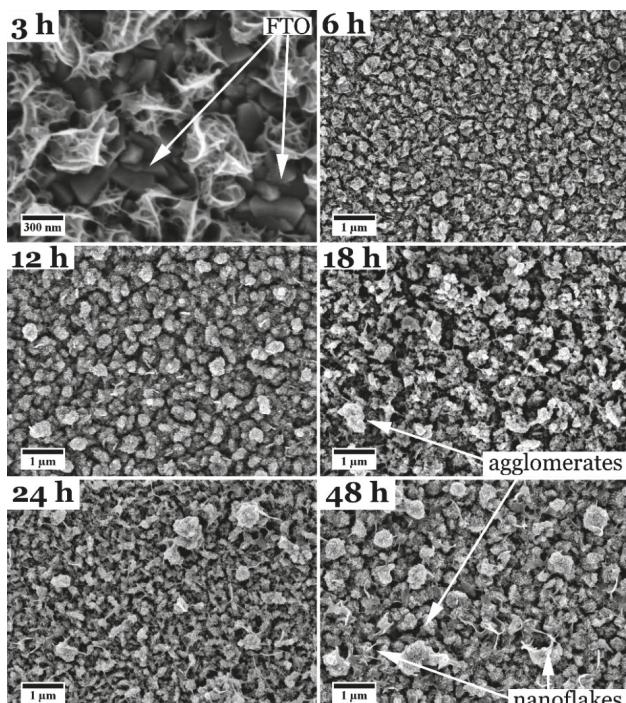


Figure 1. Top-view SEM images of CuInS₂ films synthesized with L-cysteine for different reaction times at 150 °C. The time varied between 3 h and 48 h. Please note a different scale bar for a reaction time of 3 h (film_3 h).

Measurement of the film thicknesses of the films via FIB is difficult, as the films consist of a compact CuInS₂ layer and outgrowing agglomerates. This problem is demonstrated exemplarily with the help of a FIB cross sectional cut, shown in Figure S2 (can also be seen in Figure 4a,b for a (S)TEM cross sectional sample). However, for all films, the thickness of the compact layer was determined to be in a relatively small range around 400 nm and is therefore in the same size regime.

Figure 2 shows exemplary XRD pattern of the films grown for 3, 12 and 24 h at 150 °C. The XRD pattern of the films synthesized for the other reaction times and of the pure FTO substrate are shown in the supporting information (Figures S3 and S6). Besides the strong reflections of the substrate FTO, (marked with * in Figure 2) all films show some more distinct signals. These reflections can be indexed according to tetragonal CuInS₂ in its Chalcopyrite modification (Figure 2, marked with # and compared to literature data [40]) for the films synthesized with a reaction time of at least 6 h. However, Cu^I and Cu^{II} sulfides possess very similar *d*-values as CuInS₂ and can therefore not be excluded by XRD data alone [41–43]. Additionally, amorphous phases could be present. Only film_3 h displays different reflections compared to the other films, which possess pure CuInS₂. These reflections (marked with °) can be assigned to cubic CuCl [44] and orthorhombic InS [45]. Due to the EDX measurements, described below, the presence of InS can be excluded as there is nearly no In detectable in the film as described later.

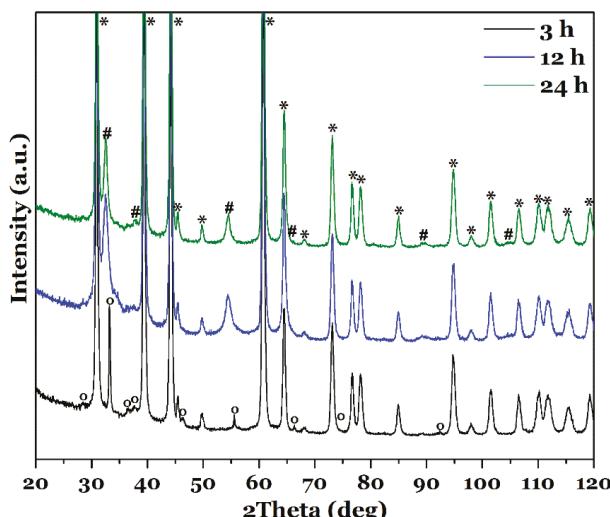


Figure 2. XRD pattern of CuInS₂ films synthesized with L-cysteine for different reaction times, 3, 12 and 24 h, at 150 °C. Signals stemming from FTO are marked with *, the ones originating from CuInS₂ with #, reflections from CuCl with °. The ° reflections could also stem from InS.

When comparing the sharp FTO reflections with the ones from tetragonal CuInS₂ it becomes clear that the latter is rather broad, indicating a small crystal and/or domain size in the films. Applying the Scherrer equation [37] to the most intense reflections (112) and (204) of CuInS₂ allows to estimate crystal sizes that are on average 9.0 ± 1.0 nm (Table 1). For film_3 h, in comparison, the distinct signals are relatively sharp, and using the (111) and (220) reflections of CuCl gives a crystal size of ≈ 39 nm for this phase. This value is three times higher than the largest calculated crystallite size of the CuInS₂ films.

Table 1. Summary of crystal size, determined with XRD, normalized elemental composition (Cu:S for SEM, Cu:In:S for STEM measurements) and band gap of the CuInS₂ films synthesized with L-cysteine for different reaction times at 150 °C. Values of film_3 h are not included in the average value in the last line.

	Crystal Size XRD (nm)	EDX SEM Cu:S (Normalized)	EDX TEM Cu:In:S (Normalized)	Band Gap UV-vis (eV)
film_3 h	39 ± 4	—	Cu:Cl 1.0:0.8 Cu:In:S 1.0:0.1:0.2	1.55
film_6 h	6.6 ± 1.4	1.0:1.3	—	1.54
film_9 h	11.0 ± 0.6	1.0:1.6	—	1.46
film_12 h	7.9 ± 1.0	1.0:2.1	—	1.54
film_15 h	8.5 ± 0.6	1.0:2.4	—	1.50
film_S [32]	9.4 ± 1.0	1.0:2.5	1.0:1.0:2.1	1.47
film_21 h	8.9 ± 1.0	1.0:2.3	—	1.51
film_24 h	11.0 ± 0.5	1.0:1.7	—	1.53
film_48 h	8.7 ± 1.0	1.0:1.9	1.0:1.0:2.2	1.44
Ø	9.0 ± 1.0	1.0:2.0 ± 0.4	—	1.50 ± 0.04

EDX measurements in SEM, as mentioned before, show not only signals from Cu, In and S, but also from the substrate (Sn, Si, O from FTO and glass) due to the large interaction volume. This is discussed in more detail in our previous publication [38]. The overlap between the In L and Sn L line make the quantification of In difficult. This is the reason why for SEM EDX measurements only the Cu:S ratio is given and In is not included. Furthermore, all the SEM EDX spectra show also signals from the conductive coating (Au, Pd) and Cl from the InCl₃ precursor.

For film_3 h, a very high amount of Cu and Cl is measured with a lower amount of S and nearly no signal from In. This proves the existence of CuCl as observed from the XRD pattern (see Figure 2). However, an InS phase, which could also explain the reflections in the XRD pattern, is not present due to the very low In amount. Possible other phases might be, e.g., a strongly distorted CuS [41]. This will be described in more detail later. As can be seen (Table 1), the sulfur ratio detected for the films is varying, but always close to the stoichiometric Cu:S ratio of CuInS₂ of 1:2, except for reaction times of 3 h and 6 h. Higher amounts of S can be attributed to e.g., incomplete cleaning of the synthesized films with water and therefore remaining precursors/amorphous side products on the film surface. To quantify also the In amount via EDX (S)TEM measurements were performed, which are described later.

All films, synthesized with reaction times from 3 to 48 h, show a similar, strong absorption behavior over the whole visible spectrum, as exemplarily shown for film_3 h, film_12 h and film_24 h in Figure 3 (the UV-vis spectra from the other samples can be found in the supporting information, Figure S4). The absorption is influenced by the film thickness and also by light scattering on film structures or on the interface to the substrate. Since the film thicknesses of our films varies slightly but stays in the same size regime (\approx 400 nm), this should not influence the absorption drastically. However, our films possess a rough surface structure (compare Figure 1) with larger agglomerates on top, which can scatter the light and ‘increase’ the absorption. Additionally, in areas with less agglomerates, the absorption is lower due to a smaller effective film thickness. Furthermore, crystal structure and composition can influence the total absorption of the measured films.

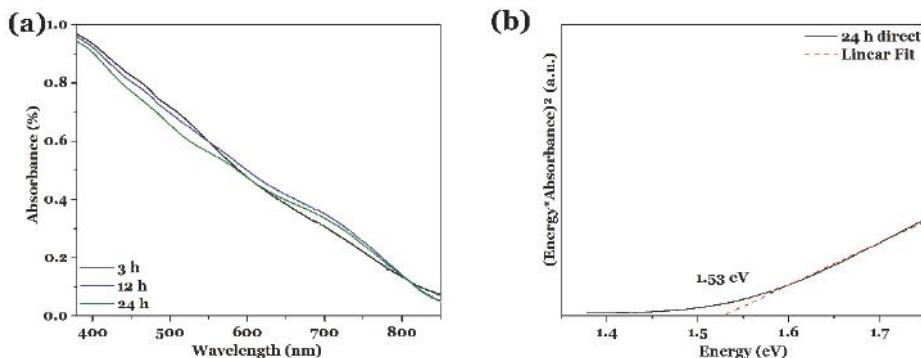


Figure 3. (a) UV-vis spectra of CuInS₂ thin films on FTO substrate, synthesized solvothermally with L-cysteine for 3, 12 and 24 h at 150 °C. (b) Exemplary Tauc plot for direct semiconductors for film_24 h, indicating a band gap of 1.53 eV.

The optical band gaps were calculated from the UV-vis data and the values are summarized in Table 1. In Figure 3b, an exemplary Tauc plot for film_24 h is shown. The smallest band gap of 1.44 eV is observed for film_48 h and the largest one with a value of 1.55 eV for film_3 h. All these values are close to the reported band gaps for CuInS₂ bulk material and nanostructures [21,46]. Variations in the band gap values can also be affected by the chemical composition, the film thickness and structure, as well as defects in the crystal structure, as mentioned above for the absorption behavior. Due to the solvothermal synthesis of our films and the small crystal size, a large number of defects, i.e., grain boundaries, are present, which could also lead to sub-band gap excitations [47–49]. Although the band gaps seem to vary (compare Table 1) no correlation can be drawn between the band gap value and, e.g., the crystallite size or chemical composition. The average band gap for the films grown for 6 to 48 h is 1.50 ± 0.04 eV. The standard deviation is relatively small, so that the band gaps can be considered as similar for all reaction times. This will be discussed in more detail later.

To investigate and compare the films grown at different reaction times in more detail, (S)TEM investigations have been performed. For the STEM EDX data, the In quantification is not problematic because the influence of the substrate FTO (and therefore the signal from Sn) can be neglected. Out of high-resolution HR TEM images, the crystal size and structure was extracted and compared to the one obtained by XRD. The results of the (S)TEM measurements on film_S [32,38] and film_48 h are shown in Figure 4, a TEM image and according diffraction pattern taken from film_3 h is shown in the supporting information, Figure S5, but also discussed in the following.

Figure 4a,b show the cross-sectional view of a focused-ion beam prepared lamella in STEM, demonstrating the vertical structure of the CuInS₂ film grown for 18 h at 150 °C. The film consists of a compact layer close to the FTO substrate with larger agglomerates growing on top of the film [32,38]. As all the other films, grown for different reaction times, display very similar topologies in SEM, it can be concluded that they also display very similar vertical structures. The HR TEM image and electron diffraction pattern in Figure 4c,d, also taken from film_S, prove the good crystallinity of the nanoparticle film and the tetragonal CuInS₂ modification [32,38]. The crystal size of the nanoparticles determined from the HR TEM images is 5.3 ± 2.2 nm and therefore smaller than the one determined with XRD (9.4 nm, compare Table 1). EDX quantification for film_S resulted in Cu 24 ± 2 at %, In 25 ± 2 at % and S 51 ± 3 at %, giving a Cu:In:S ratio of 1.0:1.0:2.1 [32]. This ratio is close to the stoichiometric value.

As for film_S, the HR TEM image and electron diffraction pattern for film_48 h reveal a good crystallinity of the solvothermally synthesized CuInS₂ film in the tetragonal Chalcopyrite modification (Figure 4e,f). Again, the film is composed of many agglomerated nanoparticles with grain sizes of

5.4 ± 2.3 nm, smaller than the value calculated from the XRD data (Table 1). The reason why the determination of the crystallite size leads to different values in XRD and TEM will be discussed later. EDX measurements on various areas gave Cu 24 ± 2 at %, In 25 ± 1 at % and S 52 ± 1 at %, resulting in a Cu:In:S ratio of 1.0:1.0:2.2. This is also very close to a stoichiometric composition of CuInS₂.

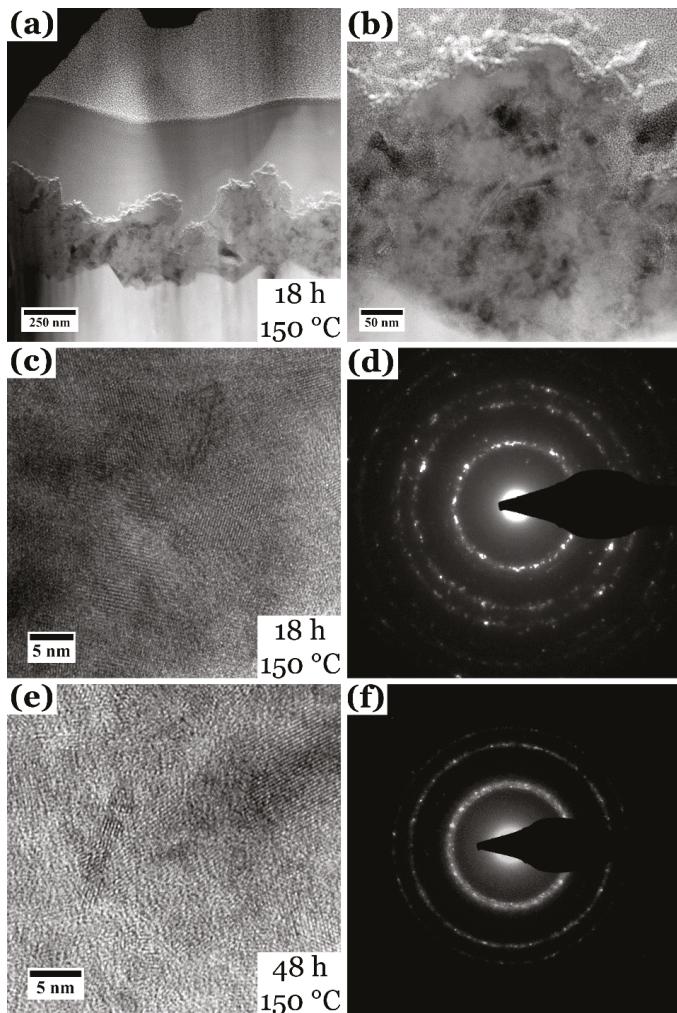


Figure 4. (a,b) cross sectional HAADF STEM images of a lamella prepared from film_S [32,38], displaying the vertical structure of the film. (c,d) HR TEM image and according electron diffraction pattern of film_S, and (e,f) from film_48 h.

Only the film grown for 3 h showed single agglomerates of nanoparticles on the FTO substrate (see Figure 1) and strong reflexes of CuCl in the XRD data. HR TEM and electron diffraction pattern (Figure S5a,b) confirm that these nanoparticles are crystalline. The crystallite sizes determined from HR TEM images resulted in a minimum value of ≈ 4 nm and a maximum crystal size of 36 nm; the latter was also obtained out of the XRD spectrum. The electron diffraction pattern can be indexed according to cubic CuCl [44] and orthorhombic InS [45]. However, due to very similar d -values of other

possible products, e.g., In_2S_3 or CuS , and the possibility of lattice distortion of these phases caused by intercalation of impurity atoms, it is very difficult to determine the unambiguous phases. EDX measurements and quantification have been performed for Cu, In, S and also Cl. The quantification led to 48 ± 5 at % Cu, 2 ± 2 at % In, 10 ± 8 at % S and 39 ± 6 at % Cl, which gives a ratio of Cu:Cl of 1:0.8. This is close to the stoichiometric Cu:Cl ratio of 1:1 for CuCl . The existence of an InS phase in film_3 h can be excluded because of the very low Indium amount.

3.2. Reaction Temperature

Figure 5 shows an overview over the CuInS_2 samples solvothermally grown with L-cysteine at different reaction temperatures for 18 h. Film_100 °C is shown in the supporting information (Figure S6). A reaction temperature of 100 °C seems to be not sufficient to grow a film. Only the pure FTO substrate is found in SEM and XRD (Figure S6). Except for film_100 °C and film_120 °C, all the films have the same appearance when studied in top-view as the films synthesized at 150 °C with different reaction times (see Figure 1). The films consist of a compact nanograined film with agglomerations of nanoparticles on top, which vary in size and density. Film_120 °C, on the other hand, looks different. The film seems to consist of nanoparticles, too, but with a smoother shape and has a white appearance on the FTO substrate when inspected by eye, while all other films are brownish. For film_180 °C, although similar to the other CuInS_2 films, a higher number of nanoflakes, which grow out of the film, and agglomerates can be observed. This can also be seen in a more distinct manner for other films (compare Figure 1). The film with the highest reaction temperature (190 °C) displays very large, round-shaped agglomerates. All films again show no cracks or delamination from the substrate.

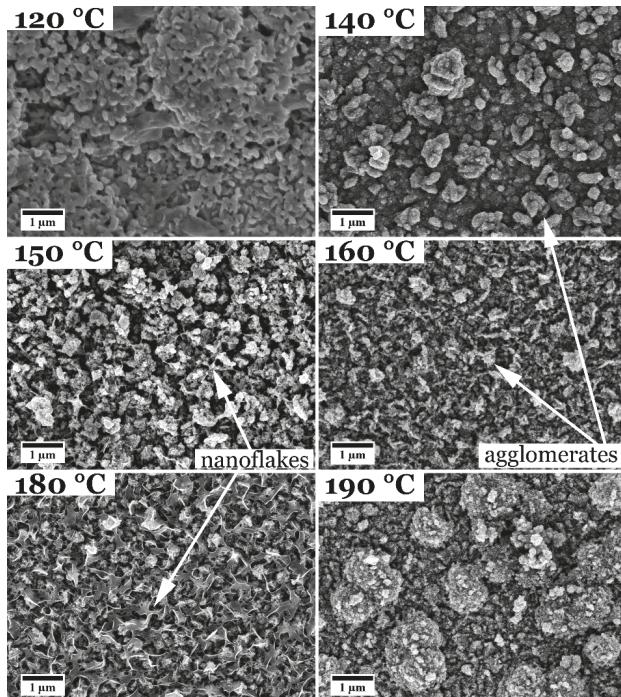


Figure 5. Top-view SEM images of CuInS_2 films synthesized with L-cysteine at different reaction temperatures for 18 h. The temperature for the films shown was varied between 120 and 190 °C.

The film thickness of these films has also been investigated by FIB cross sectional cuts. Again, the film thickness of all films is around 400 nm. As mentioned before, the measurement of the film thickness is difficult because of the structure of the film (Figure S2 and HAADF STEM image in Figure 4a).

XRD patterns obtained from the films synthesized at different reaction temperatures for 18 h are shown in Figure 6, exemplarily for 120, 160 and 190 °C, and in the supporting information, Figures S6 and S7, for the other temperatures. For film_100 °C, as already observed in the SEM, only the pure FTO substrate leads to signals in the XRD pattern. All peaks are in accordance to literature data [50]. This implies that a reaction temperature of 100 °C is not sufficient for the growth of CuInS₂. In addition, the XRD pattern of the sample grown at 120 °C (Figure 6, black) does not correspond to CuInS₂. The pattern shows a lot of reflections, which cannot be indexed unambiguously by one crystalline phase. This means that several different crystalline species are formed for this reaction conditions. It might also be that the film consists of not-reacted precursor salts or preliminary formed complexes. However, with further increase of the reaction temperature, CuInS₂ in tetragonal Chalcopyrite modification is formed. As mentioned before, the presence of side products cannot be excluded fully because of very similar *d*-values of possible compounds, e.g., copper sulfides. Amorphous phases cannot be excluded as well. However, within the detection limit, a pure CuInS₂ phase is formed for all films.

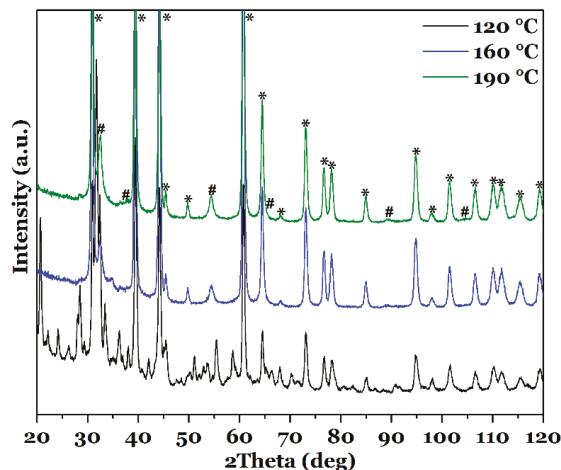


Figure 6. XRD pattern of CuInS₂ films synthesized with L-cysteine at different reaction temperatures, 120, 160 and 190 °C, for 18 h. Signals stemming from FTO are marked with *, the ones originating from CuInS₂ with #.

The CuInS₂ peaks are again rather broad and crystal sizes between 7.9 nm for film_140 °C and 11.4 nm for film_180 °C were estimated from the XRD data. On average, the crystal/domain sizes lay also in the range of 10 nm as observed for the CuInS₂ films grown for different reaction times at 150 °C (compare Table 2).

Table 2. Summary of crystal size, determined with XRD, normalized elemental composition (Cu:S for SEM, Cu:In:S for STEM measurements) and band gap of the CuInS₂ films synthesized with L-cysteine at different reaction temperatures for 18 h. Values of film_120 °C are not included in the average value in the last row.

	Crystal Size XRD (nm)	EDX SEM Cu:S (Normalized)	EDX TEM Cu:In:S (Normalized)	Band Gap UV-Vis (eV)
film_120 °C	—	1.0:3.9	—	1.51
film_140 °C	7.9 ± 0.5	1.0:1.9	1.0:1.1:2.2	1.52
film_S [32]	9.4 ± 0.6	1.0:2.5	1.0:1.0:2.1	1.47
film_160 °C	9.2 ± 1.0	1.0:1.7	—	1.54
film_180 °C	11.4 ± 1.0	1.0:1.6	—	1.55
film_190 °C	9.8 ± 0.6	1.0:1.4	1.0:1.0:1.9	1.54
Ø	9.5 ± 1.1	1.0:1.8 ± 0.4	—	1.52 ± 0.03

Analogous to the study of different reaction times, also for the films grown at different reaction temperature, EDX measurements have been performed in the SEM but only the Cu:S ratio is given.

As expected, film_100 °C does not give any signal of Cu, In or S in the EDX spectrum, as it is only the plain FTO substrate and no film has grown on top. Film_120 °C shows only a very little amount of copper (Cu:S 1.0:3.9) but large amounts of from Sn and O (FTO substrate). From film_140 °C on the films show a nearly stoichiometric ratio between Cu and S. An increase of the reaction temperature to 180 °C gives a ratio of Cu to S of 1.0:1.6 (Table 2). The decreased amount of sulfur is even more pronounced for a reaction temperature of 190 °C (Cu:S of 1.0:1.4). However, as the XRD pattern in Figure 6 only shows signals from CuInS₂ in the tetragonal Chalcopyrite modification, the reduced S amount might be caused by e.g., amorphous side products, which have not been rinsed away.

All the films show a good absorption behavior over the whole visible spectrum (see Figure 7a for film_140 °C, film_160 °C and film_190 °C, the other UV-vis spectra are shown in the supporting information, Figure S8). As described before, fluctuations in the total absorption can be caused by variations of the sample surface, relative thickness, and scattering of light. Because all the films show a similar crystal size (determined with XRD) and film thickness, the absorption is in the same order of magnitude.

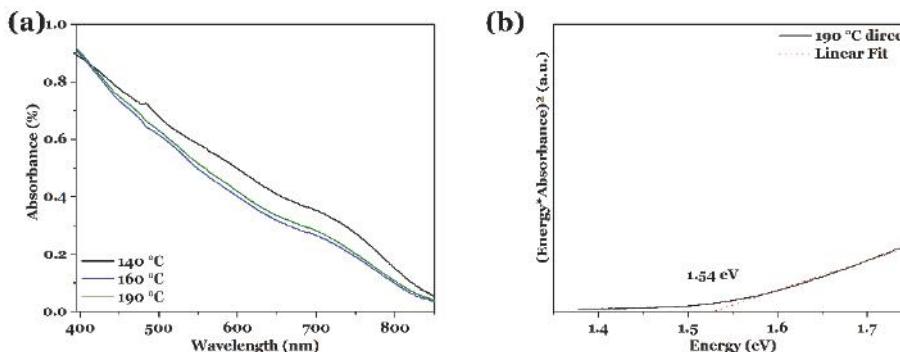


Figure 7. (a) UV-vis spectra of CuInS₂ thin films on FTO substrate, synthesized solvothermally with L-cysteine at different reaction temperatures for 18 h. (b) Exemplary Tauc plot for direct semiconductors for film_190 °C, indicating a band gap of 1.54 eV.

Figure 7b shows one exemplary Tauc plot to determine the band gap for film_190 °C; the other calculated band gaps are summarized in Table 2. The smallest band gap is found for film_S (1.47 eV) [32], while the largest one is given for a reaction temperature of 180 °C, with a value of 1.55 eV.

All the other band gap values lay in between these values and are very close to band gaps reported in literature for CuInS₂ [21,46]. Again, variations in the band gap can be induced by the chemical composition, the film thickness, and structure, as mentioned already above. The average band gap gives 1.52 ± 0.03 eV, which also displays a small standard deviation and therefore the band gaps are in the same size regime. The impact of these fluctuations in the band gap values will be discussed later.

TEM measurements on film_140 °C show a good crystallinity of the film (Figure 8a,b) and small crystallites with an average size of 5.7 ± 2.0 nm (XRD 7.9 nm, compare Table 2). The size of the crystals is in the same magnitude as for film_S [32] and film_48 h (compare also Figure 4 and Table 1); the deviation between TEM and XRD grain size determination will be discussed later. The electron diffraction pattern in (b) also proves the tetragonal Chalcopyrite modification of CuInS₂. STEM EDX quantification gives Cu 23 ± 1 at %, In 25 ± 1 at % and S 52 ± 1 at %, resulting in a Cu:In:S ratio of 1.0:1.1:2.2, which is very close to a stoichiometric CuInS₂ ratio.

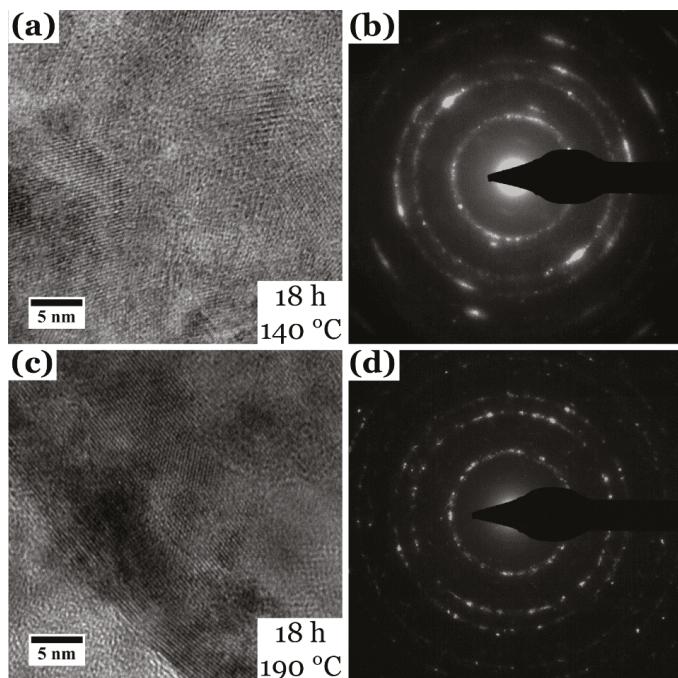


Figure 8. (a) HR TEM image and (b) corresponding electron diffraction pattern of film_140 °C. (c) HR TEM image and (d) corresponding electron diffraction pattern of film_190 °C.

The (S)TEM investigations on film_190 °C (Figure 8c,d) give similar results. The film is also well crystallized and displays the tetragonal CuInS₂ modification. Crystallite size was determined to be 10 ± 5.0 nm, which is larger than for the lower temperatures but close to the crystal size determined with XRD (≈ 9.8 nm). Additionally, the reflections in the diffraction pattern of film_190 °C appear much sharper than for film_140 °C (Figure 8b vs. Figure 8d), hinting also at a larger crystallite size. The quantitative EDX measurements in STEM mode show Cu 26 ± 1 at %, In 25 ± 1 at % and S 49 ± 1 at %, resulting in a Cu:In:S ratio of 1.0:1.0:1.9. This is also very close to stoichiometry of CuInS₂.

4. Discussion

When comparing the top-view SEM images of the samples synthesized with strongly varying reaction time and temperature, they all appear very similar. Most films are composed of a dense film with outgrowing agglomerates, also built up of nanoparticles, on top. This can also be seen in cross-sectional SEM and STEM images. The film thickness of the different films is roughly in the same size regime of ≈ 400 nm. This can be explained by the use of the same precursor concentrations for all the films, which was found to be the dominating parameter in controlling the film thickness [32]. When all available precursors are consumed, neither time nor temperature have an impact on the thickness of the grown CuInS₂ film. Only changing the number of precursor molecules would change the film thickness. However, another possible explanation for this fact is the development of a thermodynamic equilibrium between grown film and free CuInS₂ nuclei. Only adding more precursor would shift the equilibrium to an increase in film thickness.

For all synthesis conditions, a pure tetragonal Chalcopyrite phase of CuInS₂ can be observed in the XRD data, except for film_3 h, film_100 °C and film_120 °C. Applying the Scherrer equation [37] to the most intense peaks of the pattern gave an average crystal/domain size of ≈ 10 nm. Crystal sizes determined with TEM for film_S [32], film_48 h and film_140 °C give, by contrast, smaller crystal sizes of ≈ 5 nm with the exception of film_190 °C, where both methods led to similar crystal sizes (≈ 10 nm). This might be related to the fact that the films are buildup of two areas—a dense, nanograined film and overlying agglomerates of nanoparticles. Due to the global nature of XRD, the calculation of the crystallite size takes into account both parts of the film structure, resulting in an overall higher crystal size. In contrast, TEM allows the determination of the crystal size at a very local scale and the here used scratch TEM samples are most likely representing mainly the agglomerates, as they can be removed from the substrate easier. As a consequence, we conclude that our CuInS₂ thin films display larger crystal sizes in the more-dense film close to the substrate (measured with XRD), while the agglomerates on top of this film are formed of smaller crystallites (measured with TEM). Only for film_190 °C, the crystal size determined with TEM is close to the XRD measurements, which can mean that higher temperatures lead to larger crystal sizes. Nevertheless, all CuInS₂ thin film samples display a good crystallinity and the tetragonal Chalcopyrite modification of CuInS₂. Along with EDX measurements, it can be concluded that starting from a reaction time of 6 h and a reaction temperature of 140 °C, pure CuInS₂ thin films without any impurities or side products can be grown solvothermally with the help of L-cysteine. The reason for the indeterminable amount of side products like CuS, Cu₂S or In₂S₃ is the use of a high enough L-cysteine to Cu precursor ratio of 2.5:1 [32,51] as well as the complex formation between Cu, In and L-cysteine (and cystine) [35,36,52,53] as will be described below.

All the films show a very good absorption behavior in the visible regime and their band gaps, calculated with the help of the Tauc method, lie in the range between 1.46 and 1.55 eV, which is very close to the band gap for the bulk CuInS₂ [21] and has also been reported for CuInS₂ nanostructures [46]. As obvious from Tables 1 and 2 and mentioned before, the band gap values of the different films are varying and show no direct relation to the crystal size and/or chemical composition. An average value calculated from the band gaps of film_6 h to film_48 h as well as film_140 °C to film_190 °C results in 1.52 ± 0.04 eV. The small standard deviations show that the band gap variations are small and can be caused by e.g., slight variations in the chemical composition or the local crystallinity, which can be caused e.g., by amorphous side products, etc. Summarized, the absorption properties and band gap values, as well as the good crystallinity, make the CuInS₂ thin films a very interesting material for solar driven applications. Changes of the band gaps in the here observed regime should not have any influence on use in applications. A calculation of the theoretical efficiencies based on the band gap values show that, considering single-junction solar cells, a high efficiency of $\approx 30\%$ can be reached in a band gap range from ≈ 1.0 – 1.6 eV [54,55]. This means that for a solar-driven application, the observed band gap values are excellent. Equally important is the ability to absorb most of the sun light, which is true for CuInS₂.

Solvothermal reaction—and chemical reactions in general—are governed by many factors, but most important are the thermodynamic ones, pressure, and temperature, followed by reaction time. In this work, we varied the reaction time at constant temperature, and the reaction temperature for constant time intervals. In the case of a reaction time of 18 h but increasing reaction temperature, the pressure inside the reaction vessel should increase accordingly. However, the consideration in the case of a fixed reaction temperature but changing time is more complex. In general, for an increase of reaction time at fixed temperature, no significant increase in pressure is to be expected—except for the evolution of gaseous products during the reaction, which is not assumed for the CuInS₂ reaction as presented here. Nevertheless, for very short reaction times, the temperature inside the autoclave is lower than targeted and a therefore lower pressure is expected.

For film_3 h, mainly a cubic CuCl phase could be observed in the XRD and electron diffraction pattern. The film itself is very thin and the FTO substrate is visible between the agglomerates. EDX measurements in STEM mode reveal mainly Cu and Cl with only little indium and sulfur contents in the film. This shows that a reaction time of 3 h at 150 °C is not sufficient to form CuInS₂. The formation of CuCl instead can be explained as follows: Although a reaction temperature of 150 °C is chosen, the autoclave might not reach this temperature during 3 h of reaction time. For a short reaction time (i.e., lower temperature, lower pressure inside the autoclave) the decomposition of L-cysteine and cystine proceeds with a kinetically low rate, leading to (1) the metal ions In³⁺ and Cu⁺ still stabilized in complexes and (2) no S²⁻ ions are available. Instead, a lot of free Cl⁻ ions are accessible due to the solvation of the InCl₃ precursor salt. A recrystallization of InCl₃ is very unlikely because of an enthalpy of formation of 537 kJ/mol [56], which exceeds the one of In₂S₃ of −346 kJ/mol [56]. However, as mentioned before, due to the low temperature/short reaction time, no sulfur ions are available. Therefore, only CuCl with an enthalpy of formation of −137 kJ/mol [56] is thermodynamically likely to grow.

For film_100 °C, only the pure FTO substrate was obtained. It is obvious that a reaction temperature of 100 °C, although kept for 18 h, and an according low pressure inside the autoclave, is not sufficient to grow CuInS₂ or any other crystalline compound on the substrate. The formed Cu⁺ and In³⁺ complexes with L-cysteine and cystine are still stable at this temperature and a thermal decomposition is not taking place. Therefore, no S²⁻ and metal ions are released and ready to react. Additionally, the overall solubility of the precursor salts is not promoted at low temperature and pressure, and only a little amount is available in the solution.

When increasing the reaction temperature to 120 °C, a film forms, consisting of nanoparticles, probably embedded in an amorphous matrix. The XRD data show a lot of signals, arising from crystalline compounds, which means that the reaction conditions are sufficient to form crystalline compounds. However, an unambiguously assignment of all emerging signals is not achievable. Possible compounds are L-cystine, L-cysteine, their Cu and In complexes, Cu sulfides like Cu₂S, Cu_{1.8}S or CuS, indium sulfides, e.g., In₂S₃ or modifications of these phases. It is also possible that unchanged precursor materials like InCl₃ or CuSO₄·5H₂O are still present for this reaction conditions. As observed for film_3 h, the formation of CuCl is also possible.

The reaction mechanism of CuInS₂ out of CuSO₄·5H₂O, InCl₃ and L-cysteine can be described as follows: as known, L-cysteine is oxidized to cystine while Cu²⁺ is reduced to Cu⁺ [32,51,53]. This Cu⁺ and In³⁺ can be coordinated and stabilized by the chelating agents L-cysteine and cysteine [57]. The release of the metal ions is therefore very slow. As can be learned from a reaction at 100 °C for 18 h, at the chosen precursor concentrations (Cu 0.2 mol, In 0.2 mol, L-cysteine 0.5 mol, precursor ratio Cu:In:S 1:1:2.5) no free metal or sulfur ions exist in the solution. However, reaching a certain temperature/pressure (>120 °C), the organic molecules L-cysteine and L-cystine begin to decompose and release S²⁻. When the decomposition of the sulfur source starts and S²⁻ is released, (InS₂)[−] can be formed [35,36,52], but also copper sulfides are possible side products. From previous investigations [32] with higher L-cysteine contents (Cu:In:S precursor ratio 1:1:4), it is known that during a reaction at 150 °C for 18 h, a compact CuInS₂ bottom layer with an outgrowing In₂S₃ top layer is formed.

Accordingly, for a short reaction time (3 h) at 150 °C, In₂S₃ should be formed, which is not the case as described above. Here, the reason is a thermodynamically favored formation and stabilization of Cu⁺ in a chloride as no sulfur ions are present. Additionally, this gives rise to the conclusion that the combination between (InS₂)⁻ and Cu⁺ to form CuInS₂ is the main bottleneck in the reaction and rather slow, which hints at a higher importance of the reaction time compared to the reaction temperature. However, as obvious from film_100 °C and film_120 °C, the reaction temperature is important in the formation of CuInS₂. If the temperature of the reaction is not high enough, no reaction will occur (film_100 °C) or a mixture of compounds will grow (film_120 °C). However, temperatures above 120 °C allow the formation of pure CuInS₂ without obvious side products, as described above.

Our results are different from Kharkwal et al. [52] They synthesized CuInS₂ nanoparticles solvothermally using CuCl, InCl₃ and thiourea at 150 °C for different reaction times (2 to 48 h). They observed small nanoparticles of ≈5 nm for 2 h reaction time and larger nanoparticles (up to ≈27 nm) for 48 h. All nanoparticles showed a pure tetragonal CuInS₂ and with an increase in the particle size, the band gap decreased. Furthermore, with increasing reaction time, their nanoparticles evolved from agglomerates to flower-shaped structures. Kharkwal et al. ascribed the phase purity of their nanoparticles to the formation of Cu and In complexes with thiourea as we ascribe them to the formation of L-cysteine complexes. However, in our case, the crystallite size does not change with reaction time and a reaction time of 3 h is not sufficient to produce CuInS₂. This can be related to the different sulfur and copper precursors. In addition, in our case, the Cu²⁺ of the precursor has to be reduced to Cu⁺ before it is usable in the synthesis of CuInS₂, while Kharkwal et al. [52] directly used a Cu⁺ salt.

Zhuang et al. [58] prepared CuInS₂ thin films on FTO substrates as well. They used CuCl₂, In(NO₃)₂, thiourea, oxalic acid and hexadecyl trimethyl ammonium bromide (CTAB) in ethanol at 200 °C and reaction time was varied (1, 4, 8 and 20 h). For a low thiourea concentration, they observed an increase in film thickness with increasing reaction time (saturation after 8 h of reaction) and an evolution in the surface topology. For a higher concentration, the films are always composed of a close packed microsphere layer, while only the diameter and packing density changed with time. All the samples displayed the tetragonal Chalcopyrite modification of CuInS₂. In our synthesis, the influence of the reaction time on film thickness or morphology is neglectable.

To summarize the comparison with literature, it seems that our synthesis procedure using L-cysteine as sulfur source is a very robust route, as it yields CuInS₂ films on FTO substrates with excellent properties over a wide reaction temperature and time range.

5. Conclusions

To conclude, we were able to demonstrate that the green synthesis of pure, nanocrystalline CuInS₂ thin films on FTO substrates via a solvothermal, non-toxic L-cysteine assisted synthesis approach is possible and stable over a wide reaction time and temperature range. Our systematic study showed that the synthesis of CuInS₂ films with a thickness of ≈400 nm, a good crystallinity and band gaps in the range of 1.46 to 1.55 eV is independent of the exact reaction temperature and time, as long as the reaction temperature is above 140 °C and the reaction time longer than 6 h. All the results can be explained by the presented, refined reaction mechanism and the stability of initially formed precursor complexes of Cu and In complexes with L-cysteine and cystine, which provides the basis for the advantage of using L-cysteine compared to other sulfur sources. The facile and robust solvothermal synthesis of CuInS₂ using L-cysteine is therefore suggested as a possible route for up-scaling. Due to their excellent properties, the films are also viewed as possible candidates for solar driven applications like solar cells or water splitting.

Supplementary Materials: The following are available online at <http://www.mdpi.com/2079-4991/8/6/405/s1>. Figure S1, Top-view SEM images of CuInS₂ films synthesized with L-cysteine for different reaction times at 150 °C. Shown are film_9 h, film_15 h and film_21 h. Figure S2, SE image of an exemplary FIB cross section for the film thickness determination of film_140 °C is shown, displaying large fluctuations in the film thickness due to the

agglomerates. Figure S3, XRD pattern of CuInS₂ films synthesized with L-cysteine for different reaction times, 6, 9, 15, 18, 21 and 48 h, at 150 °C. Signals stemming from FTO are marked with *, the ones originating from CuInS₂ with #. Figure S4, UV-vis spectra of CuInS₂ thin films on FTO substrate, synthesized solvothermally with L-cysteine grown for different reaction times at 150 °C. Figure S5, (a) HR TEM image and (b) electron diffraction pattern of a CuInS₂ film synthesized with L-cysteine for 3 h at 150 °C (film_3 h). Figure S6, (a) SEM image and (b) XRD pattern of film_100 °C. Only pure FTO can be observed in SEM and XRD. Figure S7, XRD pattern of CuInS₂ films synthesized with L-cysteine at different reaction temperatures, 100, 140, 150, and 180 °C, for 18 h. Signals stemming from FTO are marked with *, the ones originating from CuInS₂ with #. Figure S8, UV-vis spectra of CuInS₂ thin films on FTO substrate, synthesized solvothermally with L-cysteine grown for 18 h at different reaction temperatures.

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Communication

A Polyol-Mediated Fluoride Ions Slow-Releasing Strategy for the Phase-Controlled Synthesis of Photofunctional Mesocrystals

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Abstract: There are only a few inorganic compounds that have evoked as much interest as sodium yttrium fluoride (NaYF_4). Its extensive applications in various fields, including transparent displays, luminescence coding, data storage, as well as biological imaging, demand the precise tuning of the crystal phase. Controlling the emergence of the desired α -phase has so far remained a formidable challenge, especially via a simple procedure. Herein, we represented a polyol-assisted fluoride ions slow-release strategy for the rational control of pure cubic phase NaYF_4 mesocrystals. The combination of fluorine-containing ionic liquid as a fluoride source and the existence of a polyalcohol as the reactive medium ensure the formation of uniform α -phase mesocrystallines in spite of a higher temperature and/or higher doping level.

Keywords: polyol-assisted fluoride ions slow-release strategy; NaYF_4 mesocrystals; crystallographic phase control

1. Introduction

Since inorganic micro/nanocrystals usually exist in various forms or phases, the phase transformation from kinetically stable ones to thermally stable ones is a normal phenomenon [1–4]. The intrinsic properties of a micro/nanomaterial are largely determined by its unique crystal structure [5,6]. Hence, controlling the phase formation is essential for both scientific interests and extended applications. As a typical example, sodium yttrium fluoride (NaYF_4) owns two polymorphs under ambient condition, i.e., the cubic (α -) and hexagonal (β -) phase, which is a commonly used matrix lattice for up-conversion luminescence. The former is a high-temperature metastable phase, while the latter remains thermodynamically stable [7,8]. The past decades have witnessed much exploration of its controlled synthesis and up-/down-conversion luminescent properties [1,4,9–34]. Compared with considerable work on $\alpha \rightarrow \beta$ phase transformation [1,4,9–19], the fabrication of $\alpha\text{-NaYF}_4$ as well as the investigation involving the $\beta \rightarrow \alpha$ transformation process have been neglected [26,30,33,34]. So far, some strategies have been developed to fabricate α -phase NaYF_4 nano-/micro-crystal, such as a liquid–solid–solution (LSS) procedure [21], polyol method [22], two-phase interfacial route [23], microwave-assisted ionic liquid (IL)-based technique [24], modified solvothermal approach [25], and self-sacrificing template multiple-step route [26,27]. Furthermore, introducing Mn^{2+} ($r = 81 \text{ pm}$) with a smaller size than Y^{3+} ($r = 89 \text{ pm}$) into an NaYF_4 host can dominate, forming pure α -phase NaYF_4 nanoparticles [28]. However, α -phase NaYF_4 inevitably transforms into the hexagonal ones due to its thermodynamic instability. Additionally, the cubic NaYF_4 nanoparticles are usually formed preferentially in the solution system of non-equilibrium reactions [20]. As a consequence, rationally

controlling α -NaYF₄ and simultaneously avoiding the generation of β -phase or a mixture of α and β phases remain formidable challenges, especially via a simple procedure [28,29,34].

On the other hand, the above-mentioned progress focused on NaYF₄ micro/nanocrystals instead of mesocrystals. Mesocrystals are three-dimensional (3D) order nanoparticles superstructures with unique properties and various potential applications as functional materials [35–38]. Nevertheless, the range of known mesocrystallines remains quite limited, in which few investigations to fluorine-containing compound mesocrystallines are available [39–45]. More recently, our group fabricated yttrium hydroxide fluoride mesocrystalline, as well as its Eu³⁺ doped analogue, by means of an additive-free hydrothermal procedure, which involved the reaction of Y(NO₃)₃, NaF, and NaOH aqueous solution without any organic additives [39]. Furthermore, we explored the preparation of rare-earths trifluoride mesocrystals by a solvothermal route involving IL 1-butyl-3-methylimidazolium hexafluorophosphate (BmimPF₆) as the fluorine source in the presence of 1,4-butanediol [40]. However, no effort has been made to reveal the phase control related to rare-earths fluoride mesocrystallines. Herein, we present a facile, one-pot route called a polyol-mediated fluoride slow-releasing strategy for the rational control of pure phase α -NaYF₄ mesocrystals. In spite of a higher temperature or/and higher doping level, cubic phase can be maintained.

2. Experimental Procedure

2.1. Chemicals and Materials

Analytical grade rare earth chlorides and/or nitrates (yttrium chloride hexahydrate, gadolinium chloride hexahydrate, ytterbium nitrate pentahydrate, and erbium nitrate pentahydrate, 99.9%) were provided by Aladdin Industrial Inc. Shanghai, China. NaNO₃ (99.0%), 2,2'-oxydiethanol (99.0%, diethylene glycol, abbreviated as DEG), 1,2-ethanediol (99.0%), and ethanol (99.8%) were obtained from Sinopharm Chemical Reagent Company, Shanghai, China. 1-Butyl-3-methylimidazolium hexafluorophosphate (BmimPF₆, 99%) was purchased from Aldamas-beta Co., Shanghai, China. All of the reagents and solvents were directly used without further treatment.

2.2. Synthesis

NaYF₄:Yb³⁺,Er³⁺(20/2 mol%) (abbreviated as NYF:Yb³⁺,Er³⁺ hereafter) and Gd³⁺ tri-doped NYF:Yb³⁺,Er³⁺(20/2 mol%) nanocrystals (NCs) were synthesized via a polyol-mediated solvothermal procedure. Here, we took the synthesis of NYF:Yb³⁺,Er³⁺ (20/2 mol%) as an example. The starting chemicals including NaNO₃, yttrium chloride hexahydrate, ytterbium nitrate pentahydrate, and erbium nitrate pentahydrate in the stoichiometric ratio were well mixed with 1,2-ethanediol (or DEG) under stirring, to form solution. Thereafter, the solution was slowly added into a 25-mL polytetrafluoroethylene (PTFE) vial containing a proper amount of BmimPF₆ under vigorous stirring. The autoclave was sealed after vigorous stirring at room temperature for around 15 min, and then heated at 120 °C for 24 h. The final products were collected by centrifugation, and then washed sequentially using ethanol and H₂O three times. After drying at 70 °C under dynamic vacuum for 24 h, an NYF:Yb³⁺,Er³⁺ sample was obtained. The synthetic procedure of Gd³⁺ tri-doped NYF:Yb³⁺,Er³⁺ (20/2 mol%) NCs was the same as that which was used to fabricate NYF:Yb³⁺,Er³⁺, except that the stoichiometric amount of gadolinium chloride hexahydrate was also added to 1,2-ethanediol (or DEG).

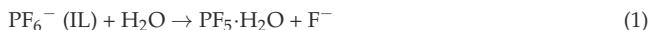
2.3. Characterization

The crystal structure and phase analysis were determined via X-ray diffraction (XRD) using a Bruker D8 Advanced X-ray diffractometer (Ni filtered, Cu K α radiation, 40 kV and 40 mA) (Bruker, Billerica, MA, USA). The morphology the products were recorded on a transmission electron microscope (TEM, JEM-2010, JOEL Ltd., Tokyo, Japan) and a Hitachi S4800 field-emission scanning electron microscope (FE-SEM) (Hitachi Ltd., Tokyo, Japan). The selected area electron diffraction (SAED) pattern were characterized by the above-mentioned TEM (JEM-2010). An up-conversion

fluorescence spectrum was obtained on an Edinburgh Instrument FLS920 phosphorimeter (Edinburgh Instruments Ltd., Livingston, UK) with a 980-nm laser diode Module (K98D08M-30mW, Changchun, China) as the excitation light source. The above-mentioned measurements were performed at room temperature from powder samples.

3. Results and Discussion

NaYF_4 -based mesocrystals were fabricated via solvothermal treatment of Na^+ , Y^{3+} , and BmimPF_6 in the presence of viscous polyol-like diethylcol or 1,2-ethanediol. Apart from serving as solvent and complexant (i.e., bonding with Na^+ and Y^{3+}), polyol also acts as a stabilizer that limits particle growth and suppresses the $\alpha \rightarrow \beta$ phase transition of NaYF_4 [22,46]. BmimPF_6 was chosen as a task-specific fluorine source (the reason why it was chosen as the fluorine source is given in the Supplementary Materials). The required fluoride anion (F^-) was provided by BmimPF_6 as a result of its slow decomposition and hydrolysis [23,45,47]. Even without additional water, BmimPF_6 can hydrolyze with the aid of the trace water and hydration water molecules from yttrium chloride hexahydrate [45,48]. During the treatment procedure, PF_6^- can slowly hydrolyze and then produce F^- through slowly increasing the temperature [45], as revealed in Equation (1):



Therefore, this procedure was defined as a fluoride slow-release strategy, which involved fluoride releasing from BmimPF_6 with the assistance of polyol [49].

Powder XRD patterns of a Yb^{3+} -Er $^{3+}$ co-doped and pure NaYF_4 submicrocube in the case of DEG as the reaction medium are illustrated in Figures 1a and 2a, respectively. All of the diffraction peaks matched the α -phase NaYF_4 crystals (PDF No.77-2042), and no impurities were found. The sharp and narrow diffraction peaks revealed the highly crystallinity of these submicrocubes despite treatment at relatively low temperature (120 °C).

As exhibited in Figure 1b–d, all of the $\text{NYF}: \text{Yb}^{3+}, \text{Er}^{3+}$ submicrocrystals show cubic shapes and edge lengths of about 120 nm. Both FE-SEM and TEM photos illustrated their novel microstructure features, which are built from many nanoparticles and exhibited rough surfaces. A few nanoparticles were found attached on its surface (Figure 1e). Especially, the SAED pattern (Figure 1f) of a single NaYF_4 cube shows sharp and periodic spots, revealing its noticeable single crystal-like feature. According to Cölfen et al. [35–40], the regular-shaped NaYF_4 cubes actually belong to typical mesocrystals. The combination of a coarse surface pattern and the attachment of nanoparticles reveal that these mesocrystallines resulted from the self-assembling of nanoparticle subunits rather than the classic crystalline growth [18,35–40].

When DEG was replaced by 1,2-ethanediol, the product can also be indexed as pure-phase cubic NaYF_4 crystal (Figure 2b). All of these results indicated that IL BmimPF_6 in the presence of polyol also acts as a crystal-phase manipulator during the formation of NaYF_4 [18].

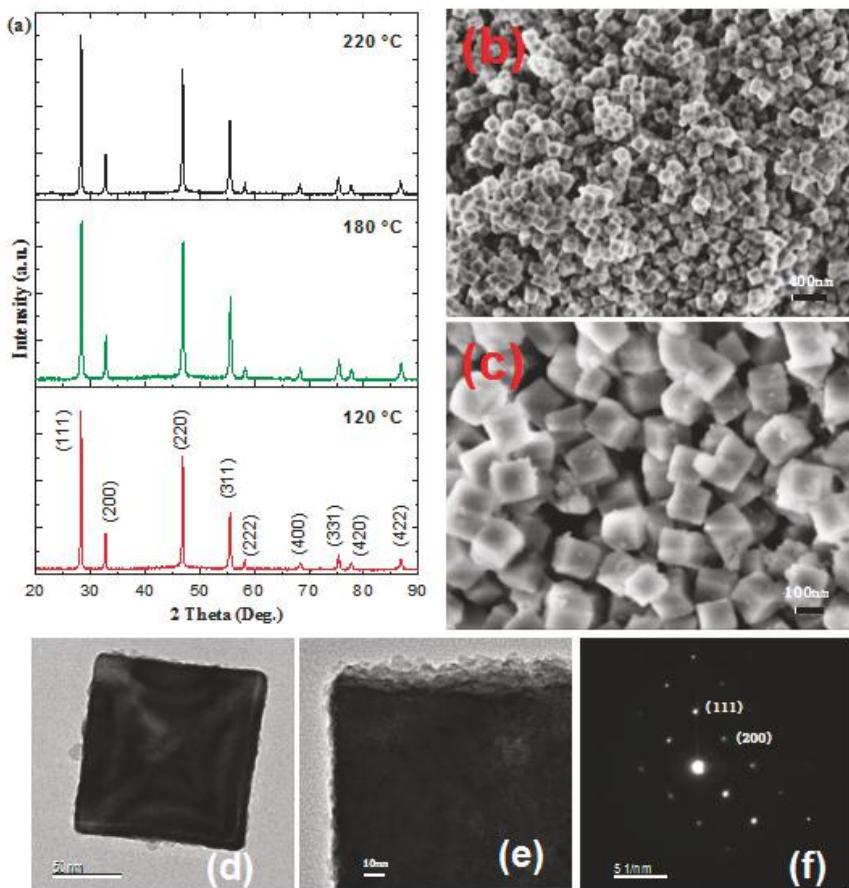


Figure 1. (a) X-ray diffraction (XRD) patterns of sodium yttrium fluoride (NYF):Yb³⁺,Er³⁺ (20/2 mol%) sample at various solvothermal temperatures using diethylene glycol (DEG) as the reaction medium (all of the diffraction peaks are attributed to cubic-phase NaYF₄), field-emission scanning electron microscope (FE-SEM) images ((b) low-magnification; (c) high-magnification), (d,e) TEM images, and (f) selected area electron diffraction (SAED) pattern of as-obtained NYF:Yb³⁺,Er³⁺ (20/2 mol%) submicrocrystals at 120 °C. (Note the nanoparticles aggregated to form submicrocubes).

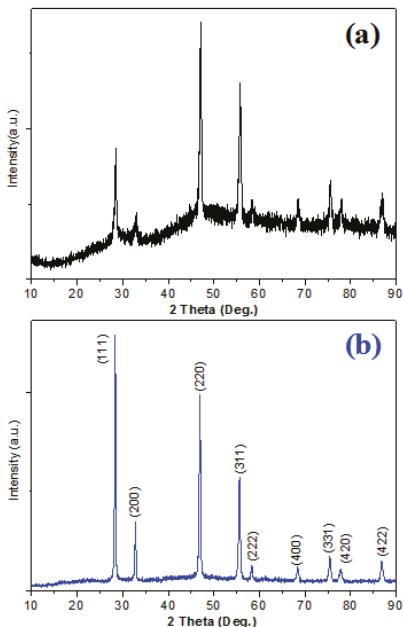


Figure 2. XRD patterns of (a) NaYF₄ host, and (b) NYF:Yb³⁺,Er³⁺ (20/2 mol%) samples using 1,2-ethanediol as solvent.

For comparison, the preparation of NYF:Yb³⁺,Er³⁺ was also conducted through an LSS procedure using NH₄F as the F[−] source and the mixture of ethanol–H₂O–oleic acid as the medium at 120 °C and 220 °C [11,21]. Figure S2 (see Supplementary Materials) revealed the XRD patterns of as-obtained NYF:Yb³⁺,Er³⁺ (20/2 mol%) at 120 °C and 220 °C. Obviously, the product obtained at the lower temperature can be ascribed to a pure α -phase NaYF₄, as expected for NaYF₄ synthesized under mild conditions [21,23,32]. However, in the case of higher temperature (220 °C), only β -phase NaYF₄ was fabricated. These results demonstrated that promoting the reaction temperature can induce the $\alpha \rightarrow \beta$ phase change of NaYF₄, which is consistent with previous reports [9–12]. However, as shown in Figure 1a, even if the reaction temperature reached 220 °C, the as-prepared nanoparticles via the fluoride ions-slow-release procedure unambiguously remained in a pure cubic phase. In a word, regardless of the treatment temperature, an α -phase NaYF₄ can be obtained by this slow-release strategy.

As mentioned above, without the tri-doping of Gd³⁺, the XRD pattern of the NYF:Yb³⁺,Er³⁺ (20/2 mol%) sample matched a cubic phase of NaYF₄ (PDF No.77-2042). As for the NYF:Yb³⁺,Er³⁺ (20/2 mol%) sample, previous works revealed that introducing lanthanide ions (such as Gd³⁺) with a larger size than the Y³⁺ ion in the NaYF₄ lattice not only induced an alteration from the α phase to the β phase, it also dominated the forming of pure β -phase NaYF₄ NCs [1,20,50]. However, in this work, as revealed in Figure 3a, the pure cubic phase of NaYF₄ remained when Gd³⁺ of 15 mol% was incorporated into host lattices. With the further increasing of the Gd³⁺ ion content (Figure 3b), no impurity diffraction peaks were found, showing the forming of a homogeneous solid solution, which is due to the small structural difference between the cubic-phase NaGdF₄ and NaYF₄. Obviously, phase transformation did not occur upon a higher-level doping of the dopant.

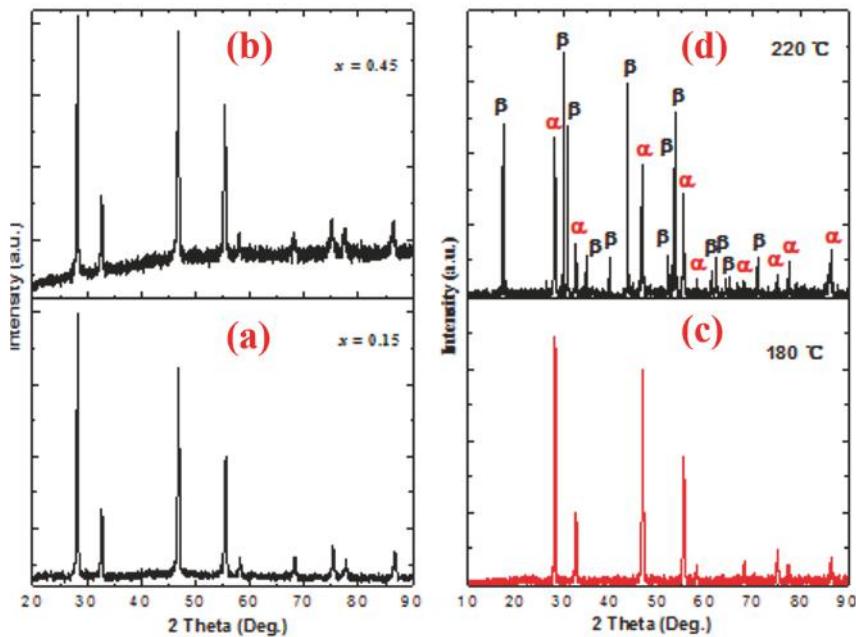


Figure 3. XRD patterns of $\text{Na}(\text{Y}_{0.78-\text{x}}\text{Gd}_{\text{x}})\text{F}_4:\text{Yb}^{3+}$, and Er^{3+} (20/2 mol%) samples with different tri-doping levels of Gd^{3+} ((a) $\text{x} = 0.15$, (b) $\text{x} = 0.45$), and $\text{Na}(\text{Y}_{0.48}\text{Gd}_{0.30})\text{F}_4:\text{Yb}^{3+},\text{Er}^{3+}$ (20/2 mol%) samples obtained at higher solvothermal temperatures ((c) $180\text{ }^\circ\text{C}$, (d) $220\text{ }^\circ\text{C}$; the symbol α and β represent cubic and hexagonal phases, respectively).

High-level doping usually leads to an $\alpha \rightarrow \beta$ phase transition of NaYF_4 in the LSS reaction system [1,20]. However, in present work, by using BmimPF_6 and polyol as the F^- source and reaction medium, respectively, as shown in Figure 3c, the as-synthesized submicrocubes remained in the cubic phase of NaYF_4 in spite of higher total doping concentrations (52 mol%) as well as a higher solvothermal treatment temperature ($180\text{ }^\circ\text{C}$). Even if the total doping contents were set as high as 52 mol%, and the treatment temperature simultaneously approached $220\text{ }^\circ\text{C}$ (near to the work-limited temperature of the PTFE vial), the α -phase NaYF_4 still existed in the products (see Figure 3d).

According to He et al. [51], the $\alpha \rightarrow \beta$ phase change of NaYF_4 can be attributed to the elevated content of F^- and the alteration to the reaction environment of Y^{3+} ions. In an LSS system involving oleic acid and a high active F^- source such as NH_4F and NaF , it is found that the oleate anions are more likely to be combined with Y^{3+} in comparison with Na^+ ions [4]. The interaction between oleate anions and Y^{3+} could effectively lower the energy barrier of the $\alpha \rightarrow \beta$ phase transition [4]. Moreover, effective concentration of F^- ions was elevated, resulting from the rapid supply of F^- . All of these could effectively promote the $\alpha \rightarrow \beta$ phase transformation of NaYF_4 [4,52,53]. Ultimately, β -phase NaYF_4 was formed in an LSS system [9–12,21].

However, in the case of BmimPF_6 as the F^- source in the presence of polyol, the interaction between PF_6^- ions and Y^{3+} was quite limited compared with the case of the above-mentioned LSS system [17,18]. In addition, noting the solubility product of NaYF_4 , $a_{\text{Na}}a_{\text{Y}}a_{\text{F}}^{-4}$, the supersaturation degree ($a_{\text{Na}}a_{\text{Y}}a_{\text{F}}^{-4}/K_{\text{SP}}$) drastically varies with the content of fluoride ions, with an exponential relationship. Thus, the content of fluoride ions in a reactive system is of importance to the phase control of NaYF_4 [15,19,33]. Herein, BmimPF_6 slowly decomposes and hydrolyzes to create the required F^- during the elevation of the reaction temperature [23,45,47]. Therefore, BmimPF_6 is a low-active F^- source relative to NH_4F and NaF . Since the equilibrium constant of the hydrolyzed reaction (Equation (1)) is extremely small, the effective concentration of F^- ions was relatively low in

the reaction system, which consequently results in a very slow precipitation with Na^+ and Y^{3+} [45] (Equation (2)).



In such circumstances, the supersaturation degree of the reaction system is not adequate to form the nuclei of the hexagonal phase NaYF_4 [23]. Therefore, the formation of cubic phase NaYF_4 was favored [51].

As mentioned above, polyol can complex with Na^+ and Y^{3+} . In addition, according to Chaumont et al. [54], PF_6^- (IL) can coordinate with Y^{3+} . Consequently, when BmimPF_6 was uniformly dispersed in 1,2-ethanediol (or DEG) solution containing Na^+ and Y^{3+} ions, these metal ions were believed to be simultaneously bonded by PF_6^- anions as well as polyol [45]. In this case, Na^+ and Y^{3+} ions were in the same shell surrounded by the imidazolium cation of BmimPF_6 [23]. Upon thermal treatment, PF_6^- slowly hydrolyzed and released F^- , which was accompanied by forming NaYF_4 nanosized grains; this can be evidenced by the nanoparticles that were attached on the surface of the as-obtained mesocrystals (Figure 1e). Subsequently, the polyol and IL co-stabilized nanoparticles aggregated to form NaYF_4 mesocrystals, which possibly occurred through oriented attachment or mesoscale assembly processes due to the coexistence of a Coulombic force, van der Waals interaction, and hydrogen bonds in the system of polyol and BmimPF_6 [55,56]. Finally, it should be pointed out that the above-proposed forming course is only one of several possible mechanisms. Further studies about this issue are underway, and will be reported in future work.

Under the excitation of a 980-nm laser, α -phase $\text{NYF}: \text{Yb}^{3+}, \text{Er}^{3+}$ (20/2 mol%) mesocrystalline emitted bright yellow fluorescence, which demonstrated its photo functionality performance. The related luminescence spectrum is shown in Figure 4a. The green-emitting bands at about 521 nm and 540/552 nm are due to the ${}^2\text{H}_{11/2} \rightarrow {}^4\text{I}_{15/2}$ and ${}^4\text{S}_{3/2} \rightarrow {}^4\text{I}_{15/2}$ energy-level transitions of Er^{3+} , respectively, while the red band at around 651/669 nm is assigned to the ${}^4\text{F}_{9/2} \rightarrow {}^4\text{I}_{15/2}$ transition of Er^{3+} . The related Commission Internationale de l'Eclairage (CIE) coordinates are calculated as ($x = 0.3984$, $y = 0.5854$), which are situated in the region of yellowish light (point “ \times ” in Figure 4b), revealing that it emitted yellowish light.

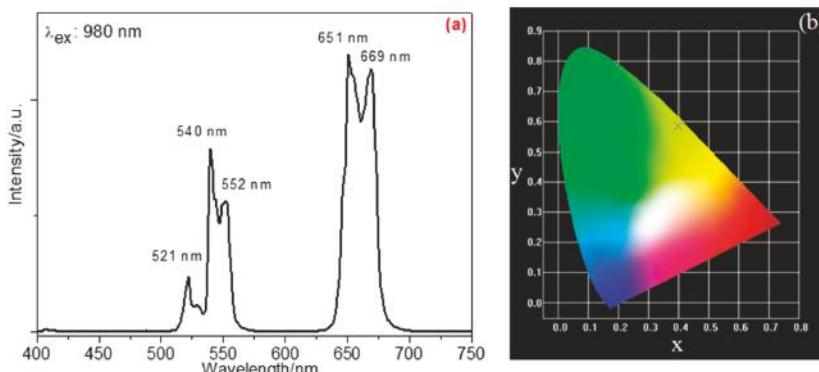


Figure 4. (a) Up-conversion luminescence spectrum at room temperature and (b) Commission Internationale de l'Eclairage (CIE) chromaticity diagram of $\text{NYF}: \text{Yb}^{3+}, \text{Er}^{3+}$ (20/2 mol%) sample ($\lambda_{\text{ex}}: 980 \text{ nm}$).

4. Conclusions

In summary, cubic-phase well-defined NaYF_4 based photofunctional mesocrystallines were successfully prepared at relatively low temperature by using IL BmimPF_6 and viscous polyol as the fluorine source and reaction medium, respectively. Combining slow-releasing fluoride via the decomposition and hydrolysis of fluoride-containing IL and the assistance of polyol, the formation of

cubic-phase NaYF₄ was favored, despite the higher treatment temperature or/and higher content of dopant. We believed that the key to the formation of uniform α -NaYF₄-based mesocrystals is the use of fluorine-containing IL as a fluorine source as well as the existence of a polyalcohol. Our contribution offers a new alternative in constructing mesocrystal and other hierarchical nanostructured materials with an object phase under mild conditions.

Supplementary Materials: The following are available online at <http://www.mdpi.com/2079-4991/9/1/28/s1>, Figure S1: XRD pattern of the sample using BmimBF₄ as fluorine source (The bar represents the standard cards PDF#70-1935 for YF₃), Figure S2: XRD patterns of NYF:Yb³⁺,Er³⁺ (20/2 mol%) NCs via LSS method at (a) 120 °C and (b) 220 °C (The bars in (a) and (b) represent the standard cards PDF#77-2042 and #PDF16-0334, respectively).

Author Contributions: Supervision, X.H.; data curation, X.H.; writing—original draft preparation, X.H.; writing—review and editing, Y.Z.; investigation, Y.F.; formal analysis, N.L.; methodology, Z.L.

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Conflicts of Interest: The authors declare no conflict of interest.

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Article

Ionic Nanocomplexes of Hyaluronic Acid and Polyarginine to Form Solid Materials: A Green Methodology to Obtain Sponges with Biomedical Potential

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Abstract: We report on the design, development, characterization, and a preliminary cellular evaluation of a novel solid material. This material is composed of low-molecular-weight hyaluronic acid (LMWHA) and polyarginine (PArg), which generate aqueous ionic nanocomplexes (INC) that are then freeze-dried to create the final product. Different ratios of LMWHA/PArg were selected to elaborate INC, the size and zeta potential of which ranged from 100 to 200 nm and +25 to −43 mV, respectively. Turbidimetry and nanoparticle concentration analyses demonstrated the high capacity of the INC to interact with increasing concentrations of LMWHA, improving the yield of production of the nanostructures. Interestingly, once the selected formulations of INC were freeze-dried, only those comprising a larger excess of LMWHA could form reproducible sponge formulations, as seen with the naked eye. This optical behavior was consistent with the scanning transmission electron microscopy (STEM) images, which showed a tendency of the particles to agglomerate when an excess of LMWHA was present. Mechanical characterization evidenced low stiffness in the materials, attributed to the low density and high porosity. A preliminary cellular evaluation in a fibroblast cell line (RMF-EG) evidenced the concentration range where swollen formulations did not affect cell proliferation (93–464 µM) at 24, 48, or 72 h. Considering that the reproducible sponge formulations were elaborated following inexpensive and non-contaminant methods and comprised bioactive components, we postulate them with potential for biomedical purposes. Additionally, this systematic study provides important information to design reproducible porous solid materials using ionic nanocomplexes.

Keywords: self-assembly; sponges; ionic nanocomplexes; polyarginine; hyaluronic acid; cell proliferation

1. Introduction

Aqueous ionic nanocomplexes (INC) are structures formed in water by the association of high-molecular-weight molecules, such as polymers, with complementary charged low- or high-molecular-weight polyions, such as ionic cross-linkers (tripolyphosphate), dyes, oligomers, or polyelectrolytes [1–10]. INC can be formed following the very simple procedure of mixing aqueous solutions of oppositely charged components at room temperature. This procedure avoids the use of organic/toxic solvents and the application of high mechanical energies, thus being ideal for biological uses. In addition, the procedure is low cost, both economically and environmentally, which facilitates the adoption of INC at the industrial level. The main nanostructures that are obtained as INC are nanogels, massive nanoprecipitates, or swollen aggregates [8,11–14]. Opposite to gelation (where only the larger component is able to be allocated in the surface, determining the surfacial charge), ionic complexation between two polymeric species allows the net charge of the INC to be selected by simply varying the ratio between oppositely charged species, thus allowing the production of formulations with tuneable electrostatic characteristics [13]. Depending on the binding forces between the ionic reactants and the ratio between their corresponding apparent charge concentrations, colloidal suspensions of the INC can be achieved just by tuning the absolute amount of matter in the mixture. At high electroneutralization regimes, the system tends to produce macroprecipitates, so the colloidal suspensions need to be highly diluted to keep colloidal stability [8]. However, an excess of one of the components allows stability of the colloidal suspension at a more concentrated regime, producing colloidal particles charged enough to ensure stability through electrostatic repulsions. Although the molecule to add in excess is normally of high molecular weight [8], examples are also found in which a low-molecular-weight component is added in excess, which, instead of diffusing out of the colloidal particles, keeps associated to them, determining the net charge of the particle and thus being responsible for the mixture stability [4].

Interestingly, upon removal of water from colloidal suspensions of INC, a solid material may arise [7,8]. In this sense, freeze-drying appears as a suitable technique to obtain micro- and nanoporous materials from suspensions of INC. Sponges are solid porous structures that can easily be manipulated and applied to selected biological tissues [15–17]. Polymeric materials prepared in a spongy form can be very useful in tissue engineering as scaffolds, which can reinforce, replace, and support some organs of the body, and also as non-scaffold materials that are able to promote cell growth [18]. If they are enriched with drugs, they are also usable as active drug delivery systems. In any case, such formulations must possess several essential properties, such as biocompatibility, biodegradability (if necessary), and absence of cytotoxicity, which primarily depend on the composition and the elaboration method of the material. Nowadays, different techniques have been developed to prepare sponge-like structures from polymers, such as phase separation, electrospinning, freeze-drying, etc.

Glycosaminoglycans, such as hyaluronic acid (HA), keratan sulfate, and chondroitin sulfate, are negatively charged biopolymers and good candidates to provide biological and functional properties as they are part of the extracellular matrix in a variety of tissues. HA is composed of repeating units of disaccharides, which include D-glucuronic acid and N-acetylglucosamine molecules linked by -(1–4) and -(1–3) glycoside bonds. This compound is involved in numerous processes occurring in the body, such as wound healing, ovulation, fertilization, signal transduction, and tumor physiology [19]. The biocompatibility and related negligible side effects make HA one of the more readily available compounds used in many fields of medicine as a biologically active molecule and as excipient in drug delivery systems [19,20].

Polyaminoacids are promising macromolecules for the development of biological active compounds and drug delivery systems. This is based on the fact that these molecules are structurally similar to polypeptides and are thus degraded by human enzymes; their accumulation within the

organism is minimal. Interestingly, the cationic polyaminoacid polyarginine (PArg) shows interesting biological properties, such as being able to translocate through cell membranes, thereby promoting the uptake of molecules associated with it [21,22]. This interesting feature of PArg has been exploited to develop drug delivery systems used for gene therapy [23], protein/vaccine delivery [24], and cancer treatment [25]. Furthermore, PArg enhances the absorption of drugs across epithelia [26], a property that may be utilized for mucosal drug delivery. Interestingly, the toxicity of the bare polycationic PArg may be minimized by its complexation with polyanionic species, such as HA [27].

The aim of this work was to study, for the first time, the formation of reproducible sponges using INC comprising LMWHA and PArg prepared at different LMWHA/PArg as the input. The methodology to prepare the INC, which was further used to test sponge formation, was similar to the one previously developed by us [13] but explored new combinations. The colloidal suspensions were studied in terms of turbidity, nanoparticle concentration, apparent hydrodynamic diameter, zeta potential, and shape. The solid materials were prepared by freeze-drying (using standard conditions: 0.02 mbar, -54°C , and 24 h) the obtained INC suspensions, a strategy that is significantly different from others focused at obtaining solid materials from solubilized components [28–31]. Optical and scanning electronic microscopy (SEM) images of the solid material (sponge) were used to support the analysis of the final product. Mechanical characterization evidenced low stiffness in the materials, attributed to the low density and high porosity. Finally, we conducted a preliminary cellular evaluation in fibroblast (RMF-EG cell line). This evidenced the concentration range where swollen formulations did not affect cell proliferation at 24, 48, or 72 h, thus projecting potential doses to be administered in further *in vitro* or preclinical/clinical studies.

2. Materials and Methods

2.1. Chemicals

Low-molecular-weight hyaluronic acid (LMWHA, Mw ~29 kDa) was purchased from Inquieroma (Barcelona, Spain). The equivalent weight of the LMWHA considering the number of possible charges was 403 g/mol of ionizable groups. Polyarginine (PArg, Mw ~5–15 kDa) was purchased from Sigma Aldrich (St. Louis, MO, USA). The equivalent weight of the PArg considering the number of possible charges was 192.5 g/mol of ionizable groups. For cell culture studies (RMF-EG cells), we used Dulbecco's Modified Eagle Medium (DMEM-HG, Gibco, Paisley, UK), fetal bovine serum (FBS, Gibco, Paisley, UK), and 1% penicillin–streptomycin (Gibco, Paisley, UK). 3-(4,5-dimethylthiazol-2-yl)-5-(3-carboxymethoxyphenyl)-2-(4-sulfophenyl)-2H-tetrazolium inner salt (MTS) and phenazine methosulfate (PMS) were purchased from Promega (Madison, WI, USA). All other reagents were of the highest analytical grade. Milli-Q water was used for experimentation.

2.2. Solid Material Preparation and Characterization

The method comprised three steps: (1) prepare suspensions of polymeric INC containing LMWHA and PArg; (2) freeze-dry the obtained INC suspensions to form the solid material; and (3) sterilize the solid material resulting from freeze-drying the INC suspensions.

INC suspensions were prepared following a procedure similar to the one described by Oyarzun-Ampuero et al. [13]. Different mass of LMWHA (5, 10, 12, 13, 15, 20, 25, 30, 35, and 50 μmol) were dissolved in 4.5 mL of Milli-Q water and added to a solution prepared by dissolving 12 μmol of PArg in 4.5 mL of Milli-Q water. The mixing was done in 50-mL cylindrical plastic containers with 38 mm diameter while stirring at room temperature. Magnetic stirring was maintained for 10 min to enable complete stabilization of the systems. The formed INC suspensions were then characterized. Size and zeta potential were determined by photon correlation spectroscopy and laser Doppler anemometry using a Zetasizer Nano-ZS (Malvern Instruments, Malvern, UK). Each batch was analyzed in triplicate. Turbidimetry studies were done using a UV–Vis spectrophotometer (Perkin Elmer, Lambda 25, Waltham, MA, USA), choosing a wavelength where the individual components (LMWHA and PArg) did not

present absorption bands. In this research, the wavelength of 540 nm was selected. The nanoparticle concentration was determined by nanoparticle tracking analysis (NTA) using a NanoSight NS300 (Malvern Instruments, Malvern, UK). Each batch was diluted from 10 to 1000 times with Milli-Q water to achieve an optimum concentration range of 10^7 – 10^9 particles/mL. A minimum of five videos (one minute each one) of the particles moving under Brownian motion were captured by the NanoSight. The videos were then analyzed for size distribution and particle concentration using the built-in NTA v 3.0 software (Malvern, UK). The morphology of the INC suspensions was determined by scanning transmission electron microscopy (STEM), model Inspect F-50 (FEI, Hillsboro, OR, USA). STEM images were obtained by sticking a droplet (20 μL) of the formulation to a copper grid (200 mesh, covered with Formvar) for 2 min, then removing the droplet with filter paper (avoiding the paper from touching the grid), then washing the grid twice with a droplet of Milli-Q water for 1 min, and then removing the droplet with filter paper. Subsequently, the sample was stained with a solution of 1% (w/v) phosphotungstic acid by adding a droplet of this solution to the grid for 2 min and then removing with filter paper. Finally, the grid was dried at room temperature for at least 1 h before being analyzed.

In order to prepare the solid materials, 9 mL of the prepared INC suspensions were frozen at -20°C for 24 h in the cylindrical plastic container they were produced in and then transferred to a freeze-dryer (Christ, Alpha Plus 1-2 LD, Osterode am Harz, Germany). The sublimation proceeded at 0.02 mbar for 24 h (condenser temperature of -54°C). The morphology and porosity of the solid materials were examined with naked eye and by SEM (LEO 420, Cambridge, England). For SEM analysis, the samples were cut with a razor blade and coated with a gold layer. Porosity threshold of the sponges was analyzed theoretically after their non-floatability in cyclohexane, a low-density organic solvent, was corroborated so that the maximum volume limit value for the solid part of the sponges could be easily calculated from their mass. The porosity threshold was then calculated as follows:

$$\text{Porosity} = \frac{\text{volume of the sponges} - \text{maximum volume limit of the solid part of the sponges}}{\text{volume of the sponges}}$$

For mechanical characterizations of the sponges, a series of 200- μL samples were transferred in 96-well plates, frozen (-20°C , 24 h), and then freeze-dried (using standard conditions: 0.02 mbar, -54°C , and 24 h). The analyses were performed in the compression mode, and hardness and apparent Young's modulus (E_{app}) were obtained. The materials were placed on a fixture base table (TA-BTKIT, Brookfield) and compressed, carefully centered, with a cylindrical TA-39 probe of 2 mm diameter. The resolution of the texture analysis system was 0.1 g and 0.1 mm. The test speed was set at 0.7 mm s^{-1} , the load trigger value ranged from 0.7 to 1.0 g, and the maximum load was set at 2 g. For the final E_{app} analysis, data for 25 sponges were considered. Finally, the stability of the sponges in water and PBS was evaluated by optical microscopy (Olympus CKX41, Arquimed, Tokyo, Japan) using a digital camera (Digital Sight DS-Fi2, Nikon, Tokyo, Japan) with the Micrometrics SE Premium® software. The samples were placed at 37°C (room temperature) on a slide and under the 4 \times objective. Subsequently, 10 μL of Milli-Q water or PBS was added, and the behavior of each sponge was recorded using the Open Broadcaster software (v.23.0.2, OBS Studios Contributors). To sterilize the solid material resulting from freeze-drying the INC suspensions, the sponges were sterilized under ultraviolet 25 W light for 4 min (UV lamp, Biolight, Santiago, Chile), sealed in plastic bags in laminar flow hood, and stored in desiccators containing dried silica gel in order to avoid moisture before their use. This method was previously used by our group, and the resulting sponges were demonstrated to maintain their biological potential in *in vitro* [32,33] and *in vivo* [34] studies.

2.3. Cellular Studies

2.3.1. Material Preparation and Administration to Cells

A formulation of INC comprising 12 mg of LMWHA and 2.4 mg and PArg (charge ratio LMWHA/PArg = 2.4), was prepared following the procedure described in Section 2.2. Aliquots of

the above preparation were diluted in Milli-Q water in order to obtain 93, 186, 464, and 1856 μM in 100 μL ; transferred in 96-well plates; and freeze-dried as described in Section 2.2. For the transfer of the material to the cells, 100 μL of culture medium was added to the sponges in the plates, and swollen formulations were then pipetted to the 96-well plates containing the cells.

2.3.2. Proliferation Assay

Five thousand RMF-EG cells were seeded in 96-well plates. After the cells adhered to the plates (2 h), the culture medium (100 μL) was extracted, and the selected formulation (charge ratio LMWHA/PArg = 2.4, prepared and transferred as described in Section 2.3.1) was added in order to achieve different concentrations (93, 185, 464, and 1856 μM). After the evaluation in terms of cellular proliferation at different time intervals (24, 48, and 72 h), the medium was replaced with 80 μL of serum-free medium plus 20 μL of MTS:PMS (20:1), mixed, and then incubated for 2 h at 37 °C. We used proliferation kit MTS (Promega, Madison, WI, USA). The reduction of MTS to formazan was determined by measuring the absorbance of this solution at 490 nm by spectrophotometry. This methodology is in accordance with the ISO 10993-5 guidelines.

2.4. Statistical Analysis

The results are shown as the mean \pm standard error of the mean for $n = 3$. The results were analyzed using one-way ANOVA tests and Dunn posttests. Statistical significance was set at $p < 0.05$.

3. Results and Discussion

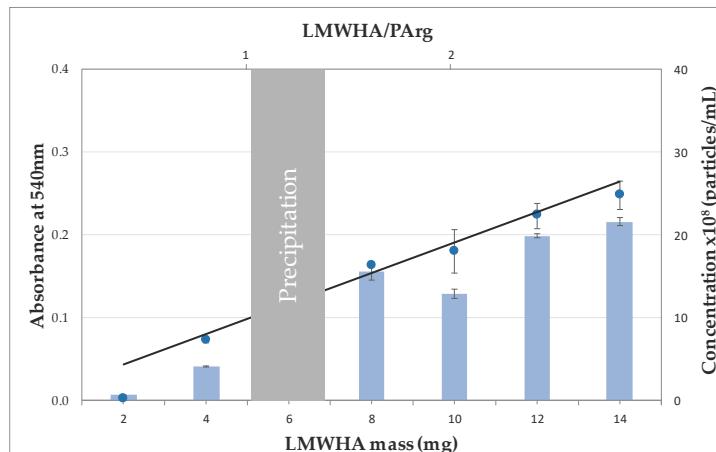
3.1. LMWHA/PArg INC Suspensions

Table 1 shows the results of apparent size and zeta potential of the prepared LMWHA/PArg INC. The methodology followed involved fixing the amount of PArg and varying the amount of LMWHA in order to achieve different charge ratios. It can be seen that most of the mixtures showed apparent size in the range between 100 and 200 nm, with low polydispersity indexes. Interestingly, macroprecipitates were formed at a LMWHA/PArg ratio of 1.1, expressed in relative number of equivalents, evidencing the highest electroneutralization between negatively and positively charged polymers. In this respect, it can also be seen that, at lower ratios, stable INC with positive zeta potential were obtained due to the excess of the positively charged PArg, while at higher ratios, the INC showed negative zeta potential due to the excess of LMWHA in the particles. In addition, there were appreciable differences in size, zeta potential, and polydispersity. These characteristics were influenced by the specific properties of each component (i.e., rigidity, linear charge, and molecular weight) and by the total mass of the polymeric formulations. In fact, formulations developed under the same strategy but using high-molecular-weight polymers showed higher size and higher polydispersity [13,35,36], indicating the role of the polymeric molecular weight on the homogeneity of the mixtures. The high correlation between the charge ratio of the components and the physicochemical properties of the formulations could be attributed to the low molecular weight of the polymers and also to their low molecular mass polydispersion. These selected parameters (concentration, charge ratio, and molecular weight of polymers) are ideal for designing specific nanoformulations in terms of size, low size polydispersity, and net charge.

Table 1. Physicochemical properties of formulations prepared with different ratios of low-molecular-weight hyaluronic acid/polyarginine (LMWHA/PArg) and evaluated in Milli-Q water (mean \pm SD, $n = 3$).

Mass Ratio LMWHA/PArg	Charge Ratio [LMWHA]/[PArg]	Size (nm)	Polydispersity Index	Zeta Potential (mV)
2.0/2.4	0.4	126 \pm 29	0.3–0.4	21.6 \pm 5
4.0/2.4	0.8	128 \pm 2	0.1–0.2	24.35 \pm 3
4.7/2.4	0.9	138 \pm 23	0.1–0.2	23.3 \pm 3
5.4/2.4	1.1	Precipitation	—	—
6.0/2.4	1.2	141 \pm 10	0.1–0.2	-17.9 \pm 3
8.0/2.4	1.6	145 \pm 9	0.2–0.3	-33.1 \pm 3
10/2.4	2.0	149 \pm 32	0.2–0.3	-35.0 \pm 4
12/2.4	2.4	146 \pm 18	0.2–0.3	-36.2 \pm 3
14/2.4	2.8	186 \pm 71	0.1–0.2	-39.7 \pm 0.4
20/2.4	4.0	166 \pm 13	0.1–0.2	-42.6 \pm 0.1

With the aim of more in-depth characterization of the colloidal suspensions, turbidity and nanoparticle tracking analyses of the formulations were studied. Due to the fact that stable colloidal suspensions maintain INC homogeneously dispersed in the aqueous phase, turbidity may give information about the stability of the suspensions as well as a qualitative idea regarding the interplay of size and amount of INC formed [37]. Figure 1 shows the values of the apparent absorbance of the colloidal suspensions at 540 nm, where functional groups of LMWHA and PArg did not show absorption bands. It can be seen that there was almost a linear increase in turbidity as the total mass of LMWHA increased (not considering the precipitation zone) due to an increase in the mass of the suspensions. The increase in the size of the INC as more LMWHA was added was moderate, as can be seen in Table 1. Therefore, the increase in turbidity was presumably mainly caused by the increase in the number of formed nanoparticles.

**Figure 1.** Turbidity (dots) and nanoparticle concentration (bars) as a function of the LMWHA mass and ratio of the components (LMWHA/PArg) (mean \pm SD, $n = 3$).

The above presumption was corroborated by nanoparticle tracking analysis, which showed that the concentration of nanoparticles in the colloidal suspensions increased similarly to the turbidity. In fact, a linear tendency related to the increases in LMWHA mass and LMWHA/PArg ratio was also observed. Because the amount of PArg was fixed in these experiments, the increase in the number and size of particles indicate that the formation mechanism of the INC allowed the incorporation of more LMWHA to preexisting particles, making them generally bigger and showing higher zeta

potential (in absolute value), and that the low-molecular-weight polymers were subjected to interaction equilibrium, allowing a higher number of contacts as the absolute concentration of the reactants increased. Importantly, as evidenced above, turbidity (analyzed by UV–Vis spectrophotometer) represents a simple and inexpensive methodology to preliminarily study (in terms of colloidal behavior) nanoparticle concentration and stability to design new nanoparticle formulations. It could also be useful to analyze batch-to-batch reproducibility for routinary analyses.

STEM microscopy experiments were developed in order to visualize selected formulations (LMWHA/PArg: 0.4, 0.8, 1.2, 1.6, 2.0, 2.4). As evidenced in Figure 2, the nanoparticles showed a spheroidal shape with an apparent size between 100 and 200 nm. In addition, an aggregation/agglomeration pattern could be seen as the LMWHA/PArg ratio increased. The samples at charge ratio of ≤ 1.6 showed more homogenous distribution with lower aggregation between the particles, while an agglomeration pattern between the nanoparticles was observed at higher ratios. It can be expected that the observed agglomeration pattern at higher mass of LMWHA could influence the formation of the solid material from the LMWHA/PArg nanocomplexes.

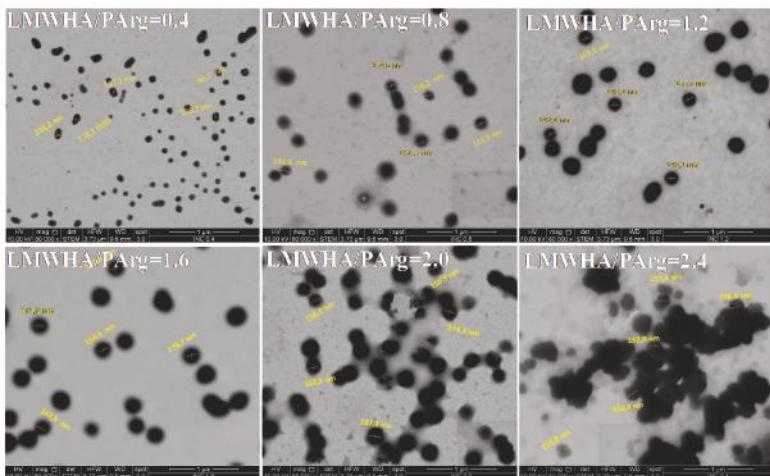


Figure 2. Scanning transmission electron microscopy (STEM) images of ionic nanocomplexes (INC) containing LMWHA and PArg at a ratio of 0.4, 0.8, 1.2, 1.6, 2.0, and 2.4 (scale bar of 1 μ m).

3.2. Solid Materials

Figure 3 shows the obtained materials after freeze-drying 9 mL of the colloidal suspensions, whose characteristics are shown in Table 1 and Figures 1 and 2, in 50-mL cylindrical containers with 38 mm diameter. The selected LMWHA/PArg ratios were 0.4, 0.8, 1.2, 1.6, 2.0, and 2.4. It can be noticed that the formation of well-structured sponges failed in the case of the two compositions bearing excess of PArg as well as when freeze-drying the highly-neutralized mixture showing a LMWHA/PArg ratio of 1.2. In contrast, the suspensions showing an excess of LMWHA presented better characteristics as sponges, showing quasi cylindrical shape with diameter of around 30 mm, which was slightly lower than that of the container they were produced in, and height of 3–6 mm, as a result of shrinking on their z axis. The formed sponges were low-density, highly porous materials with density lower than 10^{-3} g/cm³ and porosity higher than 99%. Interestingly, we observed by compression analyses that the sponges had low stiffness. Their stress-strain curves showed a nonlinear behavior with viscoelastic characteristics (Figure S1, Supplementary Material). Hardness of around 8 g was obtained between 7–15% deformation, and the E_{app} values were around 253 ± 77 kPa ($n = 25$), a fact related to the low density and high porosity of the materials. In addition, water and PBS were added to the sponges at 37 °C, and the behavior was observed by microscopy. As shown in the videos (video S1,

video S2), the microfibers were rapidly hydrated, and the material were dispersed in small pieces into the medium. Materials with similar behavior when exposed to biorelevant media have been proposed for in vitro/in vivo testing for wound healing purposes [32,34,38] and/or to be enriched with active molecules for therapeutic purposes [39–41]. Furthermore, other interesting characterizations regarding the interplay between cells and solid materials and obtaining the stiffness have been published. In this sense, the proposal from Liverani et. al. (2017) could represent an approach to be considered in the future to characterize this and other solid materials [42].

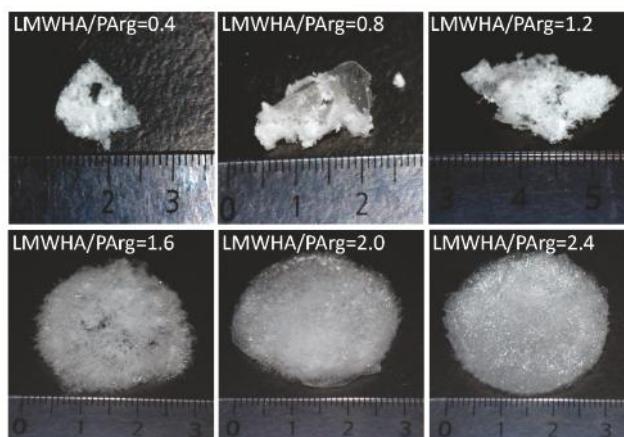


Figure 3. Optical images of materials obtained after freeze-drying 9 mL of INC aqueous suspensions containing LMWHA and PArg at a ratio of 0.4, 0.8, 1.2, 1.6, 2.0, and 2.4 in cylindrical plastic containers with 38 mm diameter.

The production of well-formed sponges from a colloidal suspension depends on several factors, among which we can name the shape and dimensions of the container compared to the volume of the colloidal suspension, the amount of components of the colloidal suspensions, and the nature of the material components. Upon freezing colloidal suspensions, the colloidal particles are concentrated during ice formation at the boundary of ice crystals, submitted to the out-of-equilibrium process called “ice-segregation-induced self-assembly” (ISISA) [8,43,44]. The migration of the solutes during freezing is determined by their hydrophilicity and mutual interactions at increasing local concentrations and lower temperatures. In this sense, as freezing at temperatures around $-20\text{ }^{\circ}\text{C}$ is produced from outside to inside the limits of the suspension volume, small ions may migrate with liquid water and concentrate at the inner part of the frozen cylinder. Molecules and particles with lower diffusion coefficients, interacting with complementary charged polyions, and amphiphilic or hydrophobic components may migrate more slowly, producing structured deposits at the boundary of ice crystals. Molecular rigidity may enhance the cohesive forces between particles and molecules subjected to electrostatic interactions. Sublimation of ice crystals after freezing during freeze-drying once the solutes are rearranged at the boundary of ice crystals then allows well-structured porous materials to be obtained. In this sense, several facts may explain the results found in this investigation. PArg is a more flexible, more amphiphilic polymer compared with the more rigid, more hydrophilic HA. In addition, the compositions with a lower content of LMWHA present less total mass of solutes. All this favors migration of the components, a weak structure at the boundary of ice crystals during freezing, and a tendency of the system to collapse during sublimation due to attractive interactions and gravity. In contrast, the higher total mass of the LMWHA-rich compositions and molecular rigidity of the polysaccharide favor the reinforcement of the structures around ice crystals and achieve the necessary tensile strength to keep the porous structure during and after sublimation.

The obtained materials presented a porous structure made of a combination of micrometric morphologies, such as microfibers and microsheets, showing a high surface area, as can be seen from the SEM images in Figure 4. As evidenced, materials with LMWHA/PArg ratios of ≤ 1.2 showed a more entangled and holey network of microstructures and lower pore size. As the LMWHA content increased in the materials, higher pores were observed, which was related to the formation of more extended, non-collapsed microstructures, in accordance with the facts revealed by optical observation. In addition, it can be seen that, at ratios of ≥ 1.6 , larger fibers and microsheets were increasingly formed, which is in agreement with the tendency described for the nanoparticles evidenced by STEM (Figure 2) and for the sponges evidenced by optical images showing more structured materials (Figure 3). Interestingly, the images did not show texturing on the surface of the laminar structures. This fact suggests that there were no remaining molecules (such as uncomplexed polymers). The above can be explained by the tendency of the solutes to migrate in the freezing process, i.e., once ice is produced upon freezing, uncomplexed chains (if any) will migrate and bind the aggregates as the local concentration increases, making them part of the final product.

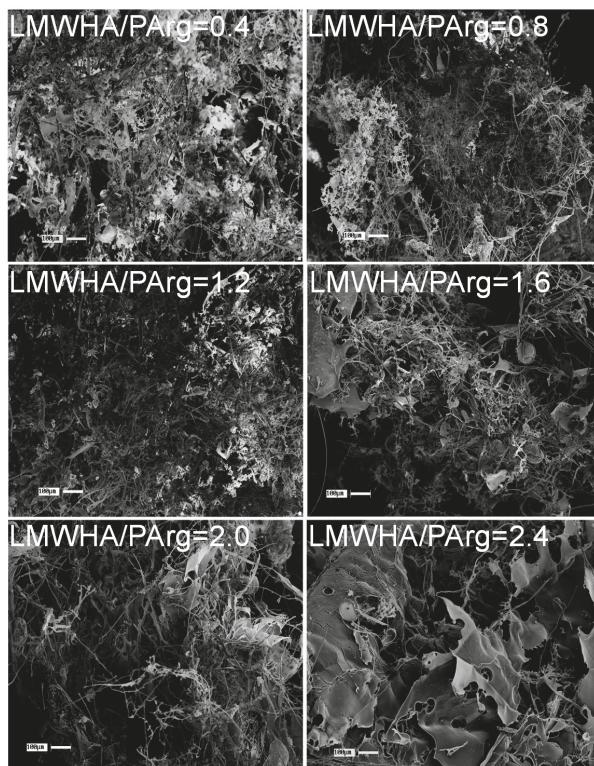


Figure 4. SEM images of materials obtained after freeze-drying 9 mL of INC containing LMWHA and PArg at different ratios of 0.4, 0.8, 1.2, 1.6, 2.0, and 2.4 in cylindrical plastic containers with 38 mm diameter. Scale bar of 100 nm.

3.3. Cellular Studies

Fibroblasts represent an adequate model cell to test the safety of new materials proposed to be applied in the skin for different therapeutic purposes [45–47]. Sponges based on a LMWHA/PArg ratio of 2.4 were selected due to their more structured characteristics (Figures 3 and 4) and also because they show very fast swellability when exposed to aqueous media. The abovementioned

characteristics are important because they favor a better manipulation of the solid formulations to be applied or when swollen formulations are administered (as in the present study). Different doses of the sponges treated with culture medium were transferred to cells in order to achieve concentrations of 93, 185, 464, and 1856 μM (see Section 2.3). As evidenced in Figure 5, the concentrations of 93, 185, and 464 μM did not significantly affect the fibroblast cell proliferation at 24, 48, or 72 h. The absence of toxicity for those doses could be used as a reference for further in vitro or preclinical/clinical studies. In contrast, the highest concentration (1856 μM) decreased the cell proliferation from 48 h. Additionally, cell morphology could be clearly seen in phase-contrast microphotographs (Figure S2). The results were similar to those observed by MTS, demonstrating that a considerable amount of dead fibroblast was observed at the highest concentration (1586 μM). In fact, at 932 μM , cells started to accumulate, showing a less healthy appearance. This behavior could be reasonably explained due to the larger content of polymeric species in the culture, possibly affecting the homeostatic equilibrium between the cells, the culture medium, and the transfer of O_2/CO_2 to the environment.

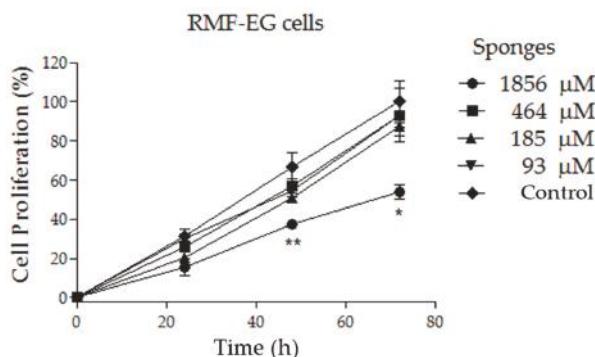


Figure 5. Proliferation curve of fibroblast (RMF-EG cell line) exposed to different doses of sponges of LMWHA/PArg = 2.4. Significant differences were obtained when comparing the values from each time with respect to the respective control value (mean \pm SD, $n = 3$) and are indicated (** $p \leq 0.01$, * $p \leq 0.05$).

3.4. Final Remarks

Low-density sponge-like materials, as those presented in this paper, are solid pharmaceutical forms to be administered in tissues and can easily be obtained from INC. The formation of INC between biocompatible ionic polymers therefore emerges as a useful tool to achieve solid, highly porous materials with biomedical potential. Thus, the strategy for this type of formulations is (i) to formulate INC that provides good balance of charges and intimate mixing of oppositely charged polymeric species and (ii) to fabricate low-density sponges by freeze-drying (avoiding cryoprotectants). This has several advantages for better manipulation, such as high stability, easy storage and transportation [38], together with high therapeutic potential [7,8,32,34,38,39,41] despite their low stiffness and high lability toward hydration (see video S1, video S2). Importantly, the hydrated material stands as a gel-like mucus (if intending to dry INC by freeze-drying and further reconstitute them through hydration, cryoprotectants are required [36,48,49]), and cell cultures studies (done following the ISO 10993-5 guidelines) correspond to the cellular response to this hydrated hydrogel. The physical and mechanical characteristics of the sponges facilitate their administration to patients as they are able to be directly applied into selected tissues [34,50–54]. Due to the need to keep stability of the colloidal suspensions of INC to be freeze-dried, the total maximum concentration of reactants is normally low, furnishing low density to the final solid materials and avoiding excessive metabolic stress when applied, which is adequate for therapeutic purposes. Another advantage of these materials regarding possible commercial purposes is that they are elaborated under mild conditions in aqueous medium, avoiding toxic excipients or covalent cross-linkers. In addition, these materials can also serve as drug

carriers due to the countless number of active molecules that can be incorporated in the colloidal suspensions, giving rise to the solid materials. The incorporation of the extra active molecules must be studied case-by-case for their possible interactions with the polymeric reactants, their influence on INC formation and stability [7,41], and their behavior in the freeze-drying process.

Here, we have shown the formation of low-density sponge-like materials made from INC of low-molecular-weight polymers, among which one was an anionic polysaccharide (LMWHA) and the other was a cationic polyaminoacid (PArg). Both higher stability as a solid material and good swellability in culture media were furnished by an excess of LMWHA over PArg (LMWHA/PArg ratio of 2.4). In vitro experiments showed a limit of cytotoxic concentration of this selected sponge for RMF-EG fibroblasts in the range of 464–1856 μM . This affords potential safe doses for further studies on the application of this material for medical purposes. The critical role of PArg in these materials as a biodegradable and biocompatible polycationic substrate able to interact with LMWHA, allowing the formation of the corresponding INC, could also be further extended to specific uses where its cell-penetrating capacities could promote the intracellular access of selected drugs and genes attached to the solid material [55–57].

4. Conclusions

In the present work, we demonstrated, for the first time, that aqueous INC formed by LMWHA and PArg were able to generate sponges after freeze-drying of the nanosuspensions. Interestingly, only those INC comprising an excess of HA were able to form sponges. NTA showed the formation of an increasing number of INC as the excess of LMWHA increased up to 2.4 over PArg. STEM experiments showed an increasing tendency of the particles to agglomerate. This phenomenon may be attributable to the higher total mass of these formulations, together with the higher rigidity afforded by HA. Mechanical characterization evidenced materials with low stiffness, attributed to the low density and high porosity. Finally, we provided a preliminary cellular evaluation in fibroblasts, evidencing the concentration range where a selected formulation did not affect cell proliferation up to 72 h, thus projecting potential doses to be administered in further in vitro or preclinical/clinical studies. Considering that the generated materials are composed of biodegradable and biocompatible compounds, we postulate them as candidates with potential for biomedical purposes. Additionally, this systematic study provides important information for researchers to design reproducible porous solid materials using INC of selected compositions as input.

Supplementary Materials: The following are available online at <http://www.mdpi.com/2079-4991/9/7/944/s1>, Figure S1: Typical strain–stress behavior of the sponges analyzed; Figure S2: Phase contrast microphotographs of fibroblast (RMF-EG cell line) exposed to different doses of sponges of LMWHA/PArg = 2.4; Video S1: Behavior of the sponges in 10 μL of Milli-Q water and at 37 °C; Video S2: Behavior of the sponges in 10 μL of PBS at 37 °C.

Author Contributions: M.G.V.-S., A.G., and V.M. developed LMWHA/PArg INC suspensions and sponges; they also contributed to the characterization using several techniques. S.M. and F.P. performed cellular assays. J.G.L. contributed to the mechanical characterization. I.M.-V., S.L.O., and F.A.O.-A. guided the INC and sponge formation. A.V., M.C., and L.L. guided the biological evaluation. All authors contributed with the writing of the manuscript.

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Review

Green Micro- and Nanoemulsions for Managing Parasites, Vectors and Pests

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Abstract: The management of parasites, insect pests and vectors requests development of novel, effective and eco-friendly tools. The development of resistance towards many drugs and pesticides pushed scientists to look for novel bioactive compounds endowed with multiple modes of action, and with no risk to human health and environment. Several natural products are used as alternative/complementary approaches to manage parasites, insect pests and vectors due to their high efficacy and often limited non-target toxicity. Their encapsulation into nanosystems helps overcome some hurdles related to their physicochemical properties, for instance limited stability and handling, enhancing the overall efficacy. Among different nanosystems, micro- and nanoemulsions are easy-to-use systems in terms of preparation and industrial scale-up. Different reports support their efficacy against parasites of medical importance, including *Leishmania*, *Plasmodium* and *Trypanosoma* as well as agricultural and stored product insect pests and vectors of human diseases, such as *Aedes* and *Culex* mosquitoes. Overall, micro- and nanoemulsions are valid options for developing promising eco-friendly tools in pest and vector management, pending proper field validation. Future research on the improvement of technical aspects as well as chronic toxicity experiments on non-target species is needed.

Keywords: agricultural pests; dengue; filariasis; insecticides; larvicides; mosquito control; stored product insects

1. Introduction

1.1. Micro- and Nanoemulsions

Over the past decades, pharmaceutical, food and agricultural research has focused the attention on the development of delivery systems able to encapsulate, protect and deliver lots of different compounds. One of the most versatile tools is represented by colloidal dispersions, which are heterogeneous systems in which the inner phase is dispersed into a continuous medium. Micro- and nanoemulsions (MEs and NEs respectively) are self-emulsifying colloidal systems, having the internal phase usually smaller than 100 nm, dispersed in a liquid medium [1]. This characteristic enhances some physicochemical properties, i.e., stability and bioavailability. In fact, the small size of the internal phase allows the system to bypass the problems related to the gravity force, avoiding phenomena as creaming or sedimentation. Moreover, the low surface and interfacial tensions promote suitable spreading and penetration of the active compounds [2].

MEs and NEs are generally composed of an aqueous phase, an oily phase, a surfactant agent and a possible cosurfactant. For this reason, they are able to incorporate both hydrophilic and lipophilic

compounds [3]. The choice of MEs and NEs components are strictly related to their application. For example, it is possible to select several oily phases between synthetic oils, (ethyl oleate, squalene and triglycerides), mineral oils and vegetable oil (e.g., olive, sunflower and soybean oil). Generally, the oily phase is used to solubilise and carry lipophilic molecules, but sometimes the oily fraction, as in the case of plant essential oils (EOs), can also be the active ingredient. EOs have been widely used in traditional medicine around the world since the Middle Ages, mainly for their antimicrobial and antioxidant properties.

A fundamental aspect about the formulation of EOs is the selection of suitable surfactant agents. The amphiphilic properties of a surfactant are represented by the hydrophilic-lipophilic balance (HLB) value. The choice of the suitable HLB value depends on the nature of the continuous phase. However, it should be desirable to select a surfactant with an intermediate value because it will partition between the aqueous and the oily phase, lowering the interfacial tension and conferring the optimal curvature of the layer, to guarantee the formation and stabilisation of the droplets. Depending on the chemical properties, surfactants can be divided into different classes: anionic, cationic, non-ionic and zwitterionic. The most diffused are polisorbates (anionic), such as Tween 80 (HLB 16.7) and Span 80 (HLB 8.6). In recent years there has been a growing interest in exploiting the surfactant properties of natural products such as polysaccharides, proteins (lectin) and sugar esters, which are desirables for the development of eco-friendly formulations. MEs and NEs have been deeply investigated, since they possess some practical advantages: easiness of formulation, industrial scale-up and high potential for use in several applications.

Apart from the terminology, these two systems present some substantial differences that it is necessary to highlight to better understand the mechanisms of their formation: (i) physicochemical behaviour, (ii) properties and (iii) applications. A summary of the main features of MEs and NEs is reported in Figure 1.

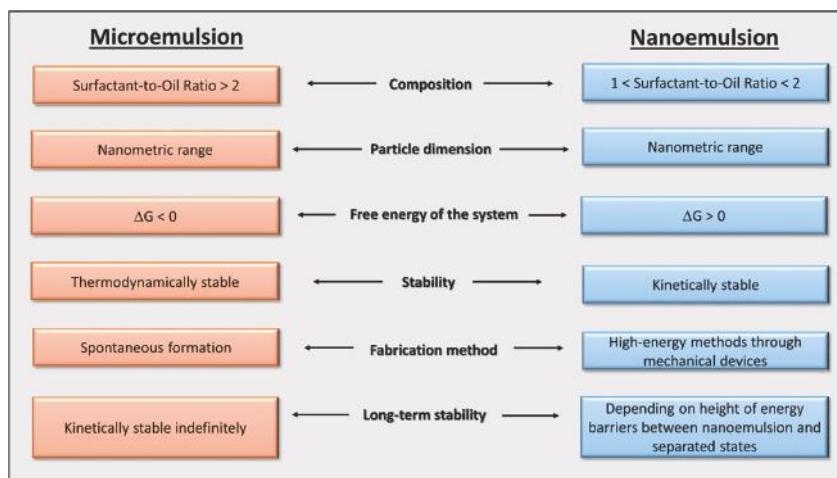


Figure 1. Comparison of the main physicochemical properties between micro- and nanoemulsion.

First, it is important to highlight that, despite the prefixes ‘micro’ and ‘nano’ define two different orders of magnitude, i.e., 10^{-6} and 10^{-9} , respectively, the size of the dispersed phase (generally oily droplets) for both of these two systems fall in the nanometric range. According to the literature it is not possible to exactly define a range of particle size distribution, since different authors report different results within the nanometric order of magnitude [3,4]. In any case, it has been reported that MEs are characterised by a smaller size of the dispersed phase respect to NEs [5,6].

ME has been defined as “a system of water, oil and amphiphile, which is a single optically isotropic and thermodynamically stable liquid solution” [7]. Introduced for the first time in 1944 by Hoar and Schulman, MEs were initially investigated for oil recovery from underground reservoirs [8,9]. Furthermore, the interest around them spread into several application fields. MEs were studied in detail in the pharmaceutical field as promising drug delivery systems for lipophilic compounds. As previously mentioned, they show several advantages such as solubilization of lipophilic compounds, enhancement of physicochemical stability respect to the related macro-systems (emulsions), improvement of the active ingredients bioavailability, achievement of a controlled drug delivery system, easiness of preparation and scale-up [10]. However, their real use is limited by the high amount of surfactant requested for the formation of such system, being these agents irritant against mucous membranes and potentially hazardous for the environment [11,12].

On the contrary, one of the most important advantages of NEs is the presence of low amounts of surfactant, generally less than 10%, compared to almost 15% in MEs, and a low surfactant-to-oil ratio (SOR) necessary for their formation, that is, >2 in MEs and comprised between 1 and 2 in NEs [2,5,6]. Briefly, NE is defined as “a thermodynamically unstable colloidal dispersion consisting of two immiscible liquids, with one of the liquids being dispersed as small spherical droplets ($r < 100$ nm) in the other liquid” [5]. It can be considered as a conventional emulsion, with the only difference of a smaller size of the dispersed phase. However, the most influential parameter varying in these two nanostructured colloidal dispersions is their free energy, conferring them different features in terms of preparation, formulation and stability.

As reported in the previous definitions, MEs are thermodynamically stable while NEs are kinetically stable. This is due to the free energy possessed by the separate state (oil + water) respect to the colloidal systems. MEs are energetically favoured, with ΔG values lower than the respective separate phases. On the contrary, NEs (oily droplet in water) possess higher free energy than those of the separate phases, water and oil.

The preparation methods of MEs and NEs are a direct consequence of this aspect. In fact, being the formation of MEs favourable, they can be obtained spontaneously by mixing oil, water and surfactant, without any external energy input. However, the application of magnetic stirring or heating could be convenient to expedite the process in order to overcome the kinetic barriers.

The energetic process that drives the MEs formation is based on the following formula [13].

$$\Delta G = \gamma \Delta A - T \Delta S \quad (1)$$

where ΔG is the free energy of the final system (ME), γ is the interfacial tension oil–water, ΔA is the variation of the interfacial area, T is the temperature and ΔS is the variation of the system entropy.

Briefly, ΔG must be negative so that a process occurs spontaneously. Since ΔA is very high in a ME (because of the formation of lots of small oily droplets that increases the interfacial area), this process is promoted by a very slow interfacial tension (γ) and by the entropy of the system that rises for the transition of the separate phases into only one containing a large number of particles; this allows obtaining a negative ΔG value.

The formation of MEs is strictly dependent on the sensitive SOR and, to determine the optimal one, is used to build a pseudoternary phase diagram. This kind of system, in fact, needs a very low interfacial tension and a favourable packaging of surfactant molecules, given by the relative interaction between their hydrophobic tails and the oil phase. This allows the formation of a fluid film at the oil–water interface [14]. Usually, the addition of cosurfactant agents is required, generally alcohols, to facilitate this phenomenon, useful to reduce the amount of surfactant as well [15]. Being MEs dynamic systems, we have to take into account that the interface is continuously subjected to a rearrangement of its structure and to the Brownian motion of the internal phase, with a possible variation of its radius [16].

Since NEs are thermodynamically unstable, the free energy of the systems, ΔG (Formula (1)), will be always positive. Thus, to exceed this value, an external energy input results to be necessary.

Depending on the physicochemical mechanisms, the methods used for NE preparation can be divided into high-energy and low-energy methods. The first ones use mechanical devices able to provide the force needed for the disruption of the dispersed phase into very small droplets, in the range of nanometres ($r < 200$ nm). Generally, NE formation follows a two steps procedure. In the first phase there is the formation of a macroemulsion through a mechanic stirrer. In the second one the macroemulsion is converted into a NE.

The most common devices used for this process are microfluidizer, sonicator, and high-pressure homogenizer. This last device uses high pressure value to pump the macroemulsion in a very narrow orifice that promotes the breaking of big droplets into smallest ones. The same result is achieved through ultrasound waves that lead to the dispersion process by means of cavitation phenomenon.

Although these approaches seem to be robust, they show some limitations concerning costs, process implementation and industrial scale-up [17].

On the contrary, low-energy methods are simpler, cheaper and more effective in producing smaller droplets. However, they require an accurate knowledge of the process parameters, showing some limitations in the ingredients and conditions [6,18].

Generally, low-energy methods are based on the phase inversion, transforming a W/O macroemulsion into an O/W NE through the variation in composition (emulsion inversion point (EIP)) or temperature (phase inversion temperature (PIT)). At the inversion point, the interfacial tension is so low that very fine droplets can be obtained, only with the support of low energy input.

Briefly, the phase inversion due to the PIT method is linked to the presence of surfactants that, based on a temperature change, modify their affinity for the hydrophilic or lipophilic phase. With EIP method there is a modification in the composition (water, surfactants, electrolytes) of the final system, which leads to a variation of the lipophilic-hydrophilic balance, with a consequent change in the curvature of surfactant layer. The free energy of the system influences the long-term stability behaviour as well. MEs should remain stable indefinitely, if the initial conditions about the chemical composition and storage will keep unchanged.

NEs, instead, will remain in a metastable state that will guarantee the stability of the systems if the energy barrier between the two different energy states remains high enough to avoid the reversion of the system and the phase separation. This occurs because of such instability phenomena such as coalescence, flocculation and Ostwald ripening, which, bringing growth of droplets, lead to creaming. It represents the migration of the dispersed phase influenced by buoyancy.

Coalescence is due to the merger of small droplets into bigger ones, while, in flocculation, droplets become very closer to move as a unique phase. These phenomena are related to the surfactant layer on the droplets surface that guarantees the steric stabilisation as much as the thickness of the layer is comparable with the droplets size. For this reason, NEs are not particularly affected by coalescence and flocculation, as compared to a traditional emulsion.

On the contrary, NEs are more prone to Ostwald ripening. This phenomenon can be defined as: “the process of disappearance of small particles or droplets by dissolution and deposition on the larger particles or droplets. The driving force for Ostwald ripening is the difference in solubility between the small and the large particles. The smaller particles (with higher radius of curvature) are more soluble than the larger ones (with lower radius of curvature). With time, the smaller particles or droplets dissolve, and their molecules diffuse in the bulk and deposit on the larger ones. This results in a shift of the particle or droplet size distribution to larger values” [19]. It is a thermodynamic process, being larger particles energetically favoured over the smaller ones.

Since the aqueous solubility of the oily droplets strongly influences the occurrence of this phenomenon, a suitable solution could be the addition of non-polar compounds that condition positively the distribution of the droplets in the oily phase. Some of the most used “ripening inhibitors” are medium-chain triglycerides (MCT), corn oil and sunflower oil [20–22].

Concluding, some of the most influential parameters on the NE stability are:

- (i) The SOR and relative concentrations; they influence the interfacial tension. It is not possible to stabilize a fixed relationship between these parameters because they are strictly related to the nature of the compounds that confer unique properties to the systems, which, in turn, differ from each other.
- (ii) The ionic strength of the dispersion medium; it affects the repulsive forces between the droplets of the dispersed phase. As the ionic strength increases, the repulsive forces decrease and the systems will be prone to instability.
- (iii) The solubility of the dispersed phase; it allows droplets to move towards the continuous phase with the appearance of Ostwald ripening.
- (iv) The temperature; it affects the solubility with the above-mentioned consequences. Moreover, it influences the energy balance of the system as well.

1.2. Applications

Thanks to the previously mentioned advantages, such nanosystems have been widely exploited in different fields as a tool for oil recovery, fuel and reaction medium in chemical applications [23–25]. However, in this section we are going to focus the attention on their applications in food, agrochemical, cosmetic and pharmaceutical fields.

About the food area, they have been developed to improve and extend the use of low water-soluble compounds or food-derived bioactive compounds with poor bioavailability. Such delivery nanosystems seem to be a suitable tool to solve this kind of problems. A significant example has been reported by Yu and Huang [26]. They demonstrated that curcumin showed a 9-fold increase in oral bioavailability when encapsulated into NEs. Moreover, it was faster digested as well, through lipolysis, respect to the unformulated compound. In the last years, NEs have been considered as a fundamental tool for the delivery of functional substances in functional foods or fortified beverages such as fatty acids, polyphenols, vitamins, micronutrients, antioxidants and others [27]. For example, O/W NE was exploited in order to encapsulate and deliver Omega-3 fatty acids in yoghurts [28].

Being extremely stable in a wide range of pH, MEs and NEs are very useful for encapsulating nutrients and protecting them from environmental conditions such as temperature or light-mediated oxidation and from possible transformation by means of enzymatic reactions and hydrolysis [29]. They formulated a valid solution to maintain suitable organoleptic properties of foods and beverages. In fact, MEs and NEs can encapsulate volatile molecules and control the release of flavours. Moreover, they can be used to prevent contamination of products and to prolong their shelf-life, both directly, for example by adding a preservative NE inside food, or indirectly, by functionalizing the packaging system in the same way [30,31]. Besides these advantages, MEs and NEs in food chain show some limitations, due to the nature of the components. For examples, in a food product the oily phase should be a triglyceride. Since the solubilization of a long chain triglyceride (LCT) is hard to obtain, it should be preferable to choose between a medium and short chain triglyceride [32].

Actually, the real limiting step in food grade nanosystem formulation is related to surfactant, because many of them are not allowed for human consumption or just at very low concentrations. Some of the admitted ones are sugar esters, monoglycerides, lecithins, glycolipids, fatty alcohols and fatty acids [33]. This issue is, nowadays, a great object of study. A large number of authors in fact, through the building of pseudoternary phase diagram of food grade components, tried to find suitable and stable formulations based only on food-grade compounds [32].

Regarding the pharmaceutical field, modern technology is progressing toward developing efficient drug delivery tools, with particular attention to an increase of bioavailability, a controlled release of the drug, a targeted biological effect and good storage stability over time. All these goals could be pursued by the exploitation of MEs and NEs. Being composed of hydrophilic and lipophilic domains, they are versatile systems able to incorporate and solubilise drugs of both natures. Araya et al. proved that MEs enhanced the oral bioavailability of poor water soluble drugs, as Ibuprofen and Ketoprofen, increasing their solubility and their plasma concentration from 60 to 20,000 times [34].

Since MEs/NEs can raise the bioavailability, the administered dose of drugs could be reduced minimizing possible side effects. These formulations behave as controlled release tools, both in O/W and W/O systems. In fact, in the first case, the oily phase acts as reservoir of active compounds, while when the oil is the external phase limits the diffusion [35]. However, the rate of drug release is influenced by the composition of the environment, such as pH and ionic strength, and features of the nanosystems, i.e., droplet dimension, type of MEs or NEs, nature of the drug and route of administration. Moreover, a limiting step is represented by the ability of the drug to cross the biological barrier, such as mucosa cells or skin [36].

Oral delivery of such nanosystems should be very useful to carry on poor water-soluble drugs, since they allow to overcome the dissolution issue on gastric fluids, which generally is strictly related to bioavailability. Moreover, they reduce the hepatic first-pass metabolism favouring the passage of the drugs in the bloodstream [35]. The small size of the internal phase and the presence of surfactants improve the drug absorption in the gastrointestinal (GI) tract, in the first case, enhancing the permeability of biological barriers, and, in the second case, promoting a wide and deep distribution [3,37].

Yin et al. showed how a ME, composed by Capryol 90 (oil), Cremophor EL (surfactant) and Transcutol (cosurfactant), increased the bioavailability of docetaxel, as compared to the related commercial product, after oral administration in rats. This result was obtained through the cumulative effect of enhanced drug solubility, improved permeability and inhibition of P-glycoprotein (P-gp) efflux [38]. Thanks to their low viscosity and possibility to be sterilised by filtration, MEs and NEs are very favourable in parenteral administration as well [4]. Moreover, they showed an appreciable physical stability in plasma [39].

Both O/W and W/O systems are suitable for parenteral formulations. Generally, O/W systems are used to deliver lipophilic compounds in order to obtain a controlled release of the drugs. Thus, they are administered by the intravenous, intramuscular or subcutaneous routes. On the contrary, W/O systems, applied as subcutaneous or intramuscular administration, are suitable to encapsulate hydrophilic drugs in order to obtain prolonged release delivery systems [4].

Dordevic et al. optimised a risperidone-based NE and monitored the pharmacokinetic parameters of the active ingredient. After intraperitoneal administration in rats, they obtained a 1.2–1.5-fold increase of bioavailability, 1.1–1.8-fold decrease in liver distribution, and 1.3-fold increase of brain uptake of risperidone as compared to the drug solution [40]. MEs and NEs are widely studied and used for topical, ocular and nasal administration as well [13,41]. The topical route has been investigated mainly in the cosmetic field, exploiting these systems in order to obtain a better penetration of the active molecules through the skin barrier [42].

Intranasal route should be exploited to deliver active molecules directly on the brain. Vyas et al. developed a mucoadhesive clonazepam-based ME for the epilepsy treatment. The concentration of this molecule in the brain was found to be 2-fold higher when compared with intravenous administration, indicating an enhanced distribution and bioavailability of the active ingredient in the site of action [43].

As in other nanosystems, functionalisation of MEs and NEs allows to build up targeted drug delivery systems, which are able to address the activity mainly in a desired target site.

Shiokawa et al. reported the formulation of aclacinomycin A, a lipophilic antitumour-antibiotic drug, through a ME linked to folate molecules. They showed that, the use of folate, helpfully modified with PEG molecules, can be considered as an effective strategy to target MEs on tumour cells [44].

Another interesting field of application of nanosystems is the agricultural one. In particular, nanotechnology is starting to revolutionise the pest management, providing innovative tools, i.e., nanoemulsions, nanoparticles and nanocapsules for the delivery of pesticide compounds (Figure 2).

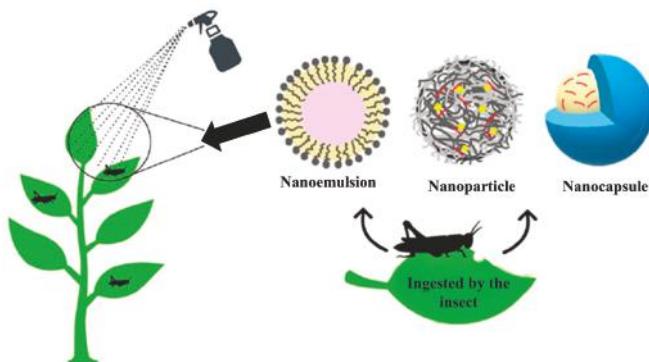


Figure 2. The most used nanosystems in insect pest control (adapted from Medina-Pérez et al. [45], with permission of Elsevier, 2019).

Among several nanodelivery systems, MEs and NEs are the easiest ones to handle and formulate. In particular, they are necessary in the presence of compounds with low water solubility that require a delivery system for their application in the field [46]. Du et al. carried out a systematic study about the formation of O/W NE based on methyl laurate as oil phase and alkyl polyglycoside and polyoxyethylene 3-lauryl ether as surfactants [47]. Moreover, they evaluated the effect of β -cypermethrin on the stability and physicochemical properties of the system.

The encapsulation process improves the physicochemical stability of pesticides and prevents the degradation of active agents [48]. Song et al. (2009) proved that the encapsulation of triazophos—an organophosphorus insecticide—is able to prevent the hydrolysis of the active compound [49]. In terms of bioactivity, these compounds result to be more effective. Nanosystems are able to ensuring their release to the target site, also providing a controlled release of the molecules at the site of action and thus reducing the required concentration of applied pesticides [2,49]. Moreover, thanks to the small size of the dispersed phase, the active compounds could improve their spreading, deposition and permeation on the target site.

2. Green Micro- and Nanoemulsions

In the last years, the growing interest of the global community on the planet fate is leading towards a more responsible and sustainable exploitation of natural resources. In particular, the worth of plants, as primary sources of ingredients for the realisation of a great variety of products, has been revaluated. In fact, some plant-based materials offer superior performance characteristics as compared to the synthetic ones. Nowadays, they have started to be applied in several fields such as pharmaceuticals, nutraceuticals, cosmetics and agrochemicals. Relying also on longstanding uses in the traditional medicine systems, they are generally employed as essential oils (EOs) and extracts, acting as flavouring agents, dyes, fortifying agents in functional foods or actual active ingredients [50].

EOs are mixtures of volatile and lipophilic molecules (mainly terpenoids and phenylpropanoids), produced in secretory structures of aromatic plants, in particular those belonging to angiosperms, such as Apiaceae, Asteraceae, Geraniaceae, Lamiaceae, Lauraceae, Myrtaceae and Verbenaceae, as products of their secondary metabolism [51].

EOs have been widely employed in the flavour and fragrance industry. They also find industrial application in foodstuffs (e.g., soft drinks, food and packaging) and cosmetics (e.g., perfumes, skin and hair care products). Regarding their medical properties, EOs are mainly used as antimicrobial agents.

Recent studies have attested pesticide properties of several EOs, natural pure compounds and extracts. The use of plant sources in crop protection dates back to 2000 years ago [52]. However, in the 20th century a wide spread of synthetic pesticides started to take hold. They were favourable thanks to a high and long-lasting efficacy. If, on the one hand, they increased crop yield, on the other hand,

their overuse led to toxic effects on humans and the environment with occurrence of resistance in pests [53–55].

The current limitations of their use are pushing discovery and development of less harmful products. One of the most promising solutions is the exploitation of plant-based pesticides. In fact, if the synthetic pesticide market is expected to decline by 1.5% per year, biopesticides have been estimated to reach the 20% of the pesticide market by 2025 [56,57].

The oldest and most widely used biopesticide is pyrethrum, a pure compound derived from the dried flowers of *Tanacetum cinerariifolium* (Trevir.) Sch.Bip. (Asteraceae) [58]. Actually, it has taken around 80% of the biopesticide market [59]. By virtue of its low toxicity against both mammals and environment, it presents a high safety profile [60]. However, its synthetic derivatives, also known as pyrethroids, have been designed to emulate the activity of the natural molecule. Despite their efficacy, they showed to be hazardous for the environment because of their long-lasting effects and high toxicity against non-target organisms [61].

Nicotine and the other alkaloids of tobacco represent another class of botanical pesticides. They act on the nervous system of pest, mimicking the neurotransmitter acetylcholine. Their use is now declined for their proved toxicity on human beings. The same problem has been observed for rotenone, isolated from *Derris elliptica* (Wall.) Benth. roots. Even though it is one of the most effective biopesticides, its high toxicity towards aquatic organisms and mammals deeply limited its use [62].

Neem (*Azadirachta indica* A. Juss.) is source of a very interesting compound, azadirachtin, a limonoid with considerable pesticide activity. It has shown bactericidal, fungicidal, and insecticidal properties, acting as a feeding and oviposition deterrent and as a growth inhibitor [63]. A fundamental aspect is its safety profile: no persistence in soil, no adverse effects on water or groundwater organisms, no toxicity to mammals [64,65].

Eco-friendly alternatives in biopesticides include the wide group of EOs. One of the most promising aspects in the exploitation of EOs is their lack of toxicity on mammals; they are generally harmless for the environment when compared with synthetic pesticides [66]. Their safety profile is guaranteed by the fact that most of EOs have been recognised as Generally Recognised As Safe (GRAS) substances by the Food and Drug Administration (FDA) and by the Environmental Protection Agency (EPA) of the United States [67]. For these reasons, a possible residue of EO-based pesticides on crop does not constitute a risk for human health.

It has been reported that EOs, such as thymol-containing EOs or EOs compounds, such as eugenol or α -terpineol, showed LC₉₀ values two or three order of magnitude higher as compared to synthetic commercial products, such as endosulfan, against Juvenile Rainbow Trout [68]. Pavela et al. reported that Apiaceae EOs have no toxicity against non-target organisms, as adult microcrustaceans *D. magna* and adult earthworms *E. fetida*, unlike α -cypermethrin that, even in much lower concentrations, caused almost 100% mortality [69].

Beyond the proofs about their safety, in the last years several studies have been carried out on the pesticide efficacy of EOs. Results showed that such substances exert a marked activity against pests, both in direct and indirect way. They act as chemosterilant, fumigant, ovicidal and repellent agents, altering growth, development and feeding behaviour [70–73]. In a recent review, Pavela collected the results published about the pesticide activity of EOs deriving from around 122 different species. Their efficacy could be expressed by an exciting data: 77 EOs showed LC₅₀ < 50 ppm [74]. Their bioactivity is strictly linked to the presence of different compounds present in the mixture of each EO, monoterpenes and sesquiterpene hydrocarbons, phenolic monoterpenes, oxygen containing mono- and sesquiterpenes and phenylpropanoids [75].

The main mechanism of action is linked to the ability of EOs to interfere with the cell membrane. Their accumulation leads to the disruption of the cell wall, leakage of the cellular contents and perturbation of homeostasis [76,77]. All these alterations lead to cell death. It has been reported that several EO constituents act in this manner [78,79]. Nevertheless, EOs, as well as plant extracts, are able to interfere with the nervous system of pests and vectors, inducing even death [80].

For example plant extracts, in particular alkaloids, can act at different levels of the pest nervous system [81]. They can function as competitive inhibitors of the acetylcholinesterase (AChE) enzyme, with consequent accumulation of the neurotransmitter in the synapses, followed by a state of permanent stimulation of the postsynaptic membrane [82]. Moreover they could be antagonist of GABA receptors as well, causing hyperexcitation, convulsion and death of the pest due to reduction of neuronal inhibition [83]. However, the most important target site of EOs is the octopaminergic system [80,84]. Octopamine is a neuromodulator and the absence of octopamine receptors in mammals is the factor that determines the distinction between target and non-target organisms. Acting on the octopaminergic system, the active compounds will be harmless for non-target organisms [72,85].

In addition to the above-mentioned advantages on the exploitation of EOs as biopesticides, a fundamental aspect is their synergistic effect. Synergism occurs in EOs since they are a mixture of 20–60 compounds, where all the components cooperate to enhance the bioactivity [86,87]. This results in a high efficacy since they act with different and complementary mechanisms of action and the combined effect is usually higher than those of the single components, allowing the reduction of the effective dose. Moreover, the mutual synergism represents a suitable tool to fight the development of resistance phenomenon, which is common with synthetic pesticides, that normally have only a target site [75].

Since EOs showed to be among the best candidates as botanical pesticides, we can ask why its commercial spread is still limited. The reason is strictly linked to their physicochemical properties, such as lipophilic nature and thus poor water solubility, scarce stability, high volatility, thermal decomposition and oxidative degradation [88]. These aspects translate into reduced efficacy and handling difficulties [72,85]. Moreover, being volatile compounds, EOs show low persistence in the environment and a scarce accumulation in soil and water [89].

All these reasons are encouraging researchers to find out suitable solutions to protect and deliver EOs. Currently, the selected strategy is the encapsulation method. Encapsulation is a process through which an active compound is coated or entrapped into a matrix. In this way, the bioactive molecule is isolated and protected by the matrix from the surrounding environment and its release depends on the external conditions and the matrix nature as well [88].

In this respect, in the last years nanotechnology revealed to be the best approach for the exploitation of EOs, allowing to overcome the limitations related to their use [48,90–95]. Although nanotechnology represents an innovative tool able to revolutionise pest management science, it remains a big, but exciting, challenge. An example of EOs stabilisation has been reported by Cespi et al. [96]. They found a suitable solution allowing the use of *Smyrnium olusatrum* L. EO, an oil difficult to handle for stability problems related to the high concentrations of its main constituent, isofuranodiene, which easily undergoes crystallisation. After a systematic study based on an experimental design, they found the best ME capable of encapsulating and protecting EO thanks to the presence of ethyl oleate that avoids the crystallisation issue. Moreover, this formulation proved to be stable over one year and maintained unchanged the bioactivity of EO. Pavela et al. used the same strategy to vehiculate isofuranodiene, the main active compound of *S. olusatrum* EOs [97]. Isofuranodiene-based ME (0.75%) has been tested against *Culex quinquefasciatus* Say showing potent larvicidal effects, with LC₅₀ value of 17.7 mL·L⁻¹.

The advantage of MEs and NEs to deliver EOs is not only related to the enhancement of the physicochemical stability but also to the improvement of bioavailability [2,49]. For this reason, the bioactivity of EO-based nanosystems is often higher than those of free EOs. Osman Mohamed Ali et al. carried out a study on the encapsulation of neem and citronella EOs in O/W NEs, to exploit their pest control properties. Stunning *in vivo* results were obtained towards phytopathogenic fungi *Rhizoctonia solani* (Cooke) Wint. and *Athelia rolfsii* (Curzi) C.C. Tu & Kimbr.; EO-based NEs showed exceptional effectiveness, which was higher than those of free EOs [98]. The higher activity of EO-based MEs compared to free EOs (*Trachyspermum ammi* (L.) Sprague ex Turrill, *Pimpinella anisum* L. and *Crithmum maritimum* L.) has been also demonstrated by Pavoni et al. on different species of bacteria and fungi [99]. Moreover, Liang et al. tested the antibacterial activity of peppermint EO NE and the relative free EO on *Listeria monocytogenes* and *Staphylococcus aureus* [100]. Although they showed comparable MIC

values, the surprising difference was related to the long-term inhibition growth given by NE. Such formulation, by increasing the stability and solubility of EO, was capable of establishing a sustained release. The dispersed phase in fact acts as a nanotank releasing active ingredient over time [21].

Furthermore, the small size of the internal phase improves mobility and penetration with an increase of the activity, and the high surface area of the oily drops enhances the efficacy [101]. Salvia-Trujillo et al. demonstrated the advantageous bioactivity of EO-based nanosystem as compared to the related coarse emulsion [102]. In this case, the difference has been made by the size of the oily droplets, highlighting once again the great advantages generated by such nanosystems.

More explicative examples of EO-based MEs/NEs as biopesticides will be reported in-depth in the following sections.

3. Green Micro- and Nanoemulsions as Insecticides

3.1. Hemiptera

Hemiptera is an order of insects comprising ~68,000 species. Some of them, including many aphids, are important agricultural pests, damaging crops by the direct action of sucking sap, but also harming them indirectly by being the vectors of bacteria, phytoplasmas, spiroplasmas and viruses. They often produce copious amounts of honeydew which encourages the growth of sooty mould. Significant pests include the cottony cushion scale, a pest of citrus fruit trees, the green peach aphid and other aphids which attack crops worldwide and transmit plant diseases. Although several studies have been reported on the activity of EOs against Hemiptera species, only few authors investigated their effectiveness on the same target when encapsulated into MEs or NEs [103–106].

Among the few examples available, Fernandes et al. developed an insecticidal NE based on *Manilkara subsericea* (Mart.) Dubard extract [107]. The efficacy of hexane-soluble fraction from ethanolic extract of *M. subsericea* on *Dysdercus peruvianus* has been previously reported by the same authors [108]. *D. peruvianus* is an Hemiptera species (Pyrrhocoridae) that acts on cotton crops causing huge harvest losses [109]. Since the apolar fraction of the extract is water insoluble, the exploitation of NE technology seemed to be a favourable strategy. After a wide screening on the suitable HLB value of surfactants and the mean droplet size, the following NE composition has been chosen: 5% of *M. subsericea* extract solubilised in 5% octyldecyl myristate (oil phase) and 5% of surfactants (sorbitan monooleate/polysorbate 80). This NE, characterised by mean droplet size of 155 nm and PDI value of 0.15, proved to be a good insecticide. In fact, it showed its activity since the first day of treatment (12% of mortality), that was sustained over time, with a mortality index of 66% of the insect population after 30 days. Moreover, the safety of this NE was confirmed noting the lack of effects against acetylcholinesterase as well as no acute toxicity on mice.

As said before, aphids represent ones of the world's major insect pests, causing serious economic damage to a range of temperate and tropical crops. This ranges from grain crops and brassicas to potato, cotton, vegetable and fruit crops. For this reason, the investigation on botanical remedies to manage these pests gained great importance and generated several studies on a wide number of EOs and aphids species [110]. Santana et al. (2012) tested the activity of *Thymus vulgaris* L. and *Lavandula latifolia* Medik. on different aphid species, namely *Rhopalosiphum padi* (L.) and *Myzus persicae* Sulzer [111]. Isman (2000) evaluated the fumigant toxicity of four EOs on *Aphis gossypii* Glover, the pest that affects mainly cotton crops, as well as a variety of plants such as citrus, coffee, cocoa, pepper, potato and many ornamental plants [85,112]. On the same target Kalaitzaki et al. tested a formulation of natural pyrethrins, a combination of six esters extracted from the flowers of *T. cinerariifolium* [113]. They solubilised pyrethrins in lemon oil obtaining, initially a W/O ME that was suddenly diluted in water, leading to the formation of an O/W NE. Results about insecticidal activity showed lower LC₅₀ and LC₉₀ values of pyrethrin-based NE as compared to those of pyrethrum commercial products (761.8 vs. 965.5 mg/mL and 4011.2 vs. 5224.0 mg/mL, respectively).

Pascual-Villalobos et al. performed a wide screening of the repellence activity of 10 EOs and 18 pure compounds against *R. padi*, the major pest of cereal crops on a world scale [114,115]. To face the volatility issue related to the nature of EOs, authors encapsulated the most active ones in NEs, in particular aniseed and peppermint EOs, as well as geraniol, *cis*-jasmone and farnesol. The effectiveness of NEs were evaluated in terms of repellence (RD_{50} and RD_{90}) and mortality after 24 h. Interestingly, some results showed that the smaller were the oil droplets the higher was the repellence activity. In particular, citral-based NE at 2%, having a particle size of 99 nm, showed a repellence index of 66, while the same formulation with larger particles (816 nm) exerted low activity.

3.2. Mosquitoes

Mosquitoes are the vectors of pathogens and parasites of medical and veterinary importance leading to the spread of diseases such as malaria, filariasis, dengue, yellow fever, Japanese encephalitis and Zika virus, just to cite the most important, some of them are lethal, especially in developing countries [116]. Thus, the effective management of these vector populations is a worthy challenge. At the moment, the main approaches to control their spread are: (1) killing adult species through the use of insecticides, (2) reduction of adults population interfering with their fecundity and oviposition or (3) killing mosquito young instars [74].

Although several pesticide products are available on the market, their dangerous effects on the environment along with the development of resistance bring to the need of new sustainable and eco-friendly tools. In the last years, research focused the attention on those EOs suitable as active ingredients in botanical larvicides. Pavela reported, from the literature, the activity of 122 EOs as mosquito larvicides [74]. Interestingly, 77 of them showed LC_{50} value < 50 ppm. Moreover, Pavela assessed the acute toxicity of 30 aromatic compounds of EOs against *C. quinquefasciatus* [87], which is the main vector of the lymphatic filariasis and has been investigated as a vector of Zika virus as well [117,118]. For this reason, several authors investigated the effect of different EOs encapsulated into MEs/NEs against this target.

Oliveira et al. improved the water solubility of *Pterodon emarginatus* Vogel oleoresin through its dispersion in a polisorbate 80/sorbitan monooleate NE, at 1:1 oil–surfactant ratio [119]. This formulation caused the death of around 100% of *C. quinquefasciatus* larvae after an exposure time of 48 h at the concentrations of 100 and 200 mg/L, probably due to morphological alterations on the final abdomen segment of the larvae. Since the *P. emarginatus*-based NE did not exert any toxicity on the green algae *Chlorella vulgaris* Beijerinck, it can be considered an eco-friendly botanical product. The effect of EOs formulations on non-target organisms have been investigated in depth by Pavela et al. on the microcrustacean *Daphnia magna* Straus, the aquatic worm *Tubifex tubifex* (Müller) as well as the earthworm *Eisenia fetida* (Savigny) [69,97]. Moreover, they proved the larvicidal activity of MEs based on Apiaceae EOs, as those of *T. ammi*, *C. maritimum* and *P. anisum*, and on isofuranodiene, the major volatile compound of *S. olusatrum* EO, evaluating the chronic and acute toxicity on *C. quinquefasciatus*. These formulations showed remarkable efficacy, with LC_{50} values of 1.57, 2.23, 4.01 and 17.7 mL/L, respectively.

Several studies have been conducted on the effectiveness of OEs-based MEs and NEs against *Aedes aegypti* L. larvae, the major vector of dengue and yellow fever. In particular, *Rosmarinus officinalis* L. and *Ocimum basilicum* L.-based NEs showed evident efficacy on larval mortality, in a time and dose-dependent manner [120,121]. Interestingly, several authors reported how the exploitation of nanotechnology in pest management could be useful to enhance, not only the stability of EOs, but also their efficacy as pesticide agents.

Balasubramani et al. [122] reported a study based on the larvicidal activity of *Vitex negundo* L. EO on *A. aegypti*. The encapsulated EO showed higher toxicity as compared to the free one, with lower LC_{50} and LC_{90} values. MEs and NEs, in fact, providing a higher dispersion of the lipophilic phase into an aqueous one, could increase the concentration of active ingredients dispersed at the interface leading to direct improvement of the interaction with the target [123].

An important parameter related to the EOs activity is the size of the oily droplets. In fact, Anjali et al. [124] observed that the smaller was the droplets size, the higher was the formulation efficacy. In particular, neem oil NE with a medium diameter of 31 nm caused the mortality of 86% of *C. quinquefasciatus* larvae after 24 h, while NEs of 93 and 251 nm showed a percentage of mortality of 73% and 48%, respectively.

Sugumar et al. [125] compared the activity of *Eucalyptus globulus* Labill. EO encapsulated both in NE and bulk emulsion against *C. quinquefasciatus*. It was observed that, at the concentration of 250 ppm, NE caused 100% of mortality after only 4 h, while the bulk emulsion obtained the same result after 24 h. It is possible to suppose that the size reduction of oil droplets, and thus the increment of the surface area, lead to a better interaction and penetration of the active ingredients into the target organisms [126].

3.3. Stored Product Beetles

Cereal crops can be still considered a main food source for mankind [127]. However, their yield could be compromised by pest infestations during storage. This leads to an extensive loss of crops in term of quality and quantities. In fact pests, not only reduce the amount of grains, but also create suitable environmental conditions for the growth of moulds [128]. The most widespread insect of stored products is *Tribolium castaneum* Herbst, also known as the red flour beetle, which is able to release carcinogenic substances [129].

Botanical research found out several EOs able to fight stored product pests, in particular *T. castaneum*, acting through contact, fumigant, growth inhibitory, antifeedant and repellent actions [130]. Starting from this knowledge, several authors worked on the development of suitable formulation of EOs for their real application. Hashem et al. encapsulated *P. anisum* EO, known to be effective against *T. castaneum*, into a NE, in order to enhance its physicochemical properties [131]. 10% EO-based NE showed a mortality index of 81.33% after 12 days of exposure. Moreover, such system was able to significantly affect the development of progeny and reduce the grain weight loss (%). Morphological and histological evaluations showed that the EO-based NE adhered to several body parts and penetrated through the cuticle, causing cellular necrosis. On the same target, other authors tested EOs obtained from three species of *Achillea*, *A. biebersteinii* Afan., *A. santolina* Falk and *A. millefolium* E.Mey. [130]. They showed how the EO bioactivity depends on the kind of exposure and thus, the mechanism of action. In fact, fumigant toxicity proved to be more effective respect to the topical and contact ones. In particular, the EO-based NE showed significant higher fumigant toxicity as compared to the free EOs, with almost one order of magnitude lower LD₉₅ values. Moreover, authors proved that these nanosystems were more effective, in terms of mortality, on adults as compared to larvae, although they strongly affected their growth and development.

Interestingly, Pant et al. added a new ingredient to EO-based NEs that was proven to enhance the effectiveness of the system [132]. They formulated 10% eucalyptus EO NE to test against *T. castaneum*, using karanja and jatropha aqueous filtrates (at increasing concentration from 20% to 60%) in place of water. Such filtrates, obtained from the de-oiled seed cakes, showed to possess insecticidal properties [133,134]. This study reported how the presence of aqueous filtrates improves the physicochemical properties of the formulations, reducing the medium size of the dispersed phase and the PDI value. Moreover, they enhanced the shelf-life of EO for long periods of time reducing its volatility. In fact, after two months, in presence of filtrates, the concentration of EO active ingredients remained unchanged, while in presence of water it decreased to 5%.

Eucalyptus globulus-based NE has been investigated against the species *Sitophilus granarius* L., as well [135]. This formulation showed higher efficacy on this pest when compared with free EO. In addition, such NE showed to be safe, since it did not show mortality and did not cause biochemical alterations in rats.

Choupanian et al. investigated the activity of neem oil NEs against *T. castaneum* and *Sitophilus oryzae* L., also known as the rice weevil [136]. Authors underlined as the effectiveness of a system could

depend, not only on the presence and amount of active ingredients, but rather on the formulation parameters. In this case, the choice of the surfactant was carefully evaluated. In fact, polysorbate and alkylpolyglucoside have been compared. NEs obtained with polysorbate showed smaller droplets size and enhanced stability as compared to those containing the other surfactant. Moreover, by their reduced size, they showed higher activity since the active ingredient could penetrate the insect cuticle and come in contact with the target. Moreover, the study reported higher pest mortality of NEs as compared to commercial products and the crude oil extract. These results could be ascribed again to the reduced droplets size of the NEs that caused 100% of mortality in both species after 48 h. Although the previous mentioned species are the most common pests that affect stored products, researchers investigated EO-based NEs against other species as well, obtaining encouraging results about the effectiveness of such nanosystems on the preservation of cereal crops from the infestation of several different pests species [130].

4. Green Micro- and Nanoemulsions as Insect and Tick Repellents

As detailed in the paragraph above, hematophagous insects act as main vectors of several diseases, such as Zika virus, dengue, malaria and yellow fever, causing more than one million deaths per year [137,138]. There is need of new specific drugs or vaccines to treat or prevent such diseases; however, one possible approach to control them is represented by reliable vector control tools, with proven epidemiological impact. One of the simpler ways to deal with this is the employ of repellent products. Repellents are chemical molecules able to prevent the arthropod landing on the skin and the consequent bite [139]. They act through a topical action forming a vapour layer having an intolerable odour for a given arthropod species, preventing its contact with human skin. It is desirable that such molecules do not penetrate in the bloodstream but, rather remain in the stratum corneum [140].

The ideal arthropod repellent should possess some key features: (i) broad spectrum of activity, (ii) long-lasting effect (>8 h), (iii) no toxicity for human being and environment, no skin irritation and low penetration, (iv) odourless to humans and unbearable to arthropods [139]. Generally, repellents are lipophilic volatile molecules, thus they need a suitable vehicle or formulation to be administered.

Now only five/six compounds have been recognised and approved by the Environment Protection Agency (EPA) and the Center for Disease Control and Prevention (CDC) as active repellent ingredients. They have been admitted for skin products thanks to their low toxicity [138]. Three of them are synthetic compounds. The most known and used, since 1957, is *N,N*-diethyl-3-methylbenzamide (DEET). Despite its high efficacy and long-lasting effect, several studies proved its toxicity due to high skin absorption [141,142]. Its overuse may cause encephalopathy, dermal toxicity, cardiovascular diseases and psychosis and hence, its use has been now restricted and forbidden for pregnant women and children [143]. Other recognised synthetic compounds are ethyl butylacetylaminopropionate (IR3535) and picaridin. The first one is not harmful if ingested, inhaled, or used onto the skin and thus, it can be accepted for human use. Picaridin can be compared to DEET in terms of efficacy and long-lasting effect but it showed only slow toxicity [138,139].

Given the toxicity and resistance issues related to synthetic repellents, one of the biggest challenges for the scientific community is the identification of new efficient and safe compounds [142]. Since ancient times human being has used plants as means to protect himself from insects and pests, by burning or bruising them or by applying their extracts directly on the skin [144,145]. In fact, plants can produce some by-products properly to defend themselves against bloodsucking arthropods. Generally they act binding the odorant-binding proteins in the arthropod's antennae for cuing, preventing their approach [145].

At present, research is focused on the exploitation of EOs to find out new effective natural repellents [146]. Their activity seems to be related to the presence of isoprenoid molecules. In particular, the combination of monoterpenes and sesquiterpenes in the mixture of EOs is considered to be responsible for their repellent activity [147]. Several studies reported that monoterpenes as citronellol, limonene, camphor and thymol showed effective repellent activity [148–150]. Citronellal

and eucalyptus EOs have been recognised as skin treatments by EPA while PMD (*p*-menthane-3,8-diol), a compound of *Corymbia citriodora* (Hook.) K.D.Hill & L.A.S.Johnson. EOs, is the only natural repellent recommended by CDC, showing no adverse effects on human health [146]. Although EOs efficacy and safety have been widely proved, their use is still restricted due to some drawbacks related to their physicochemical properties. In fact, they showed rapid evaporation and a short action. Moreover, the application of pure EOs on the human skin could cause irritation [139].

To overcome these limitations the best strategy could be the encapsulation of such active ingredients to develop suitable formulations able to protect and control the release of EOs. The main systems developed for the formulation of repellent EOs are micro-/nanocapsules, MEs/NEs, liposomes, solid lipid nanoparticles and polymeric micelles [139]. Containing oily and water insoluble substances, MEs and NEs could be considered among the best choices as EOs vehicle.

Nowadays the classical repellent formulations on the market are spray solutions and lotions. The first ones require a high amount of alcohol to solubilise the active ingredients while the second ones are emulsions with low stability. On the contrary, NEs and MEs are able to overcome these issues. In fact, they are highly stable, low viscous to be easily spread on the skin and physiologically acceptable in terms of composition [139].

Nuchuchua et al. carried out a study on NEs based on citronella (*Cymbopogon citratus* (DC.) Stapf), hairy basil (*Ocimum americanum* L.) and vetiver (*Vetiveria zizanioides* (L.) Nash) EOs [151]. They evaluated their physicochemical properties, the *in vitro* release, the *in vivo* efficacy on *Ae. aegypti* and the toxicity against normal human foreskin fibroblast (NHF) cells. They compared the different formulations before and after high-pressure homogenisation. After this high-energy process, smaller oily droplets, in the range of 150 to 160 nm, were obtained. They resulted to have a better stability, expressed as zeta potential values, after 2 months. Moreover, the small size of the oily droplets showed to play an important role in the formulation efficacy. In fact, NEs showed a higher release rate, based on a diffusion mechanism, and longer repellent activity. Authors supposed that formulations having smaller size should be able to form a whole film on the skin to prolong the activity. The best formulation was the NE composed of 10% citronella, 5% hairy basil and 5% vetiver EOs, in terms of size, stability and efficacy (4.7 h of protection). Also, Sakuluku et al. investigated the effects of high pressure homogenisation, concentration of surfactant and presence of glycerol on the physicochemical properties and mosquito repellent activity of 20% citronella EO NEs [152]. The best conditions to obtain effective NEs were as follows: concentration of surfactant at 2.5% and water:glycerol at 0:100 ratio. In fact, they demonstrated to influence the kinetic release and the activity against *Ae. aegypti*, as well as the droplet size and the long-term stability. The high amount of glycerol, and thus the high viscosity of the system, delayed the release of EOs, resulting in a prolonged repellent activity on time.

Drapeau et al. formulated PMD based-MEs to evaluate against *Ae. aegypti* [138]. They compared a “surfactantless” ME, composed of water, propanol and PMD and a classical ME, obtained through the construction of a ternary phase diagram. The presence of surfactants led to a prolonged activity, that increased from 315 min of the “surfactantless” ME to 385 min of the classical ME, as well as the reduction of the amount of propanol. The selected formulation was composed of: 46% of H₂O, 20% (w/w) of PMD, 25% of PrOH, 2% of Cremophor RH40 (surfactant), 3% of Texapon N70 (surfactant), 1% of 2-ethylhexane-1,3-diol (cosurfactant) and 3% of ethyl (−)-(S)-lactate (cosolvent). The addition of these two additives seemed to increase the activity of PMD. The cosurfactant has been selected for its repellent properties, while ethyl (−)-(S)-lactate could act as lactic acid competitor on human skin, a good attractant for mosquitos [153–155].

Lastly, Navayan et al. showed how MEs could be a suitable tool to prolong the repellent activity of EOs [156]. In fact, 5%, 10% and 15% eucalyptus EO-based MEs showed a protection time against Culicidae of 82, 135 and 170 min, respectively, while free EO at the same concentrations showed lower time of activity, i.e., 34, 47 and 59 min, respectively. The results obtained through the encapsulation of EO were similar to those of DEET at the same concentrations. Notably, this work outlined how

nanosystems could be a desirable tool to increase EOs protection, reduce their volatility, promote their release and prolong the activity on time.

5. Green Micro- and Nanoemulsions as Acaricides

Mite control is economically important for assuring the survival of several vegetables and ornamental plants in greenhouses. For this purpose, conventional pesticides have been widely applied. They include organotin compounds, mitochondrial electron transport inhibitor-acaricides (fenazaquin, fenpyroximate, pyridaben and tebufenpyrad) and pyrethroids. Although they resulted to be very effective, their use has been limited due to the development of pest resistance and the non-target, environmental and human toxicity. These issues have highlighted the need to find out new alternatives for pest management. Botanical pesticides seem to be a valid alternative to the synthetic ones, and are in the field of acaricides products as well. In particular, EOs showed to be the most important natural sources of compounds with acaricidal activity [157–160].

Choi et al. tested the activity of fifty-three EOs against eggs and adults of *Tetranychus urticae* Koch as well as adults of the biocontrol agent *Phytoseiulus persimilis* Athias-Henriot [161]. This study revealed that the most active EOs were: caraway (*Carum carvi* L.) seed, citronella java (*Cymbopogon winterianus* Jowitt), lemon eucalyptus (*C. citriodora*), pennyroyal (*Mentha pulegium* L.), and peppermint (*M. x piperita* L.) EOs showed >90% of toxicity against adults of both mite species. From the obtained results, authors supposed that EOs were delivered and acted on the vapour phase, affecting the respiratory system of mites.

Although their safety and effectiveness, EOs showed a short lasting effect related to their rapid volatilisation and/or degradation [125]. Thus, their encapsulation in liquid sprayable MEs and NEs could be a suitable solution.

Concerning mite species of public health importance, Xu et al. investigated the acaricidal activity of neem oil against *Sarcoptes scabiei* expressed as the speed of kill (min) [162]. Authors compared the effectiveness of pure EO, the EO-based emulsion and the EO-based ME. Neem EO-ME demonstrated the highest acaricidal activity with a lethal time of 192 min followed by 212 min of EO-emulsion and 337 min of pure EO. As expected, the encapsulation process and the small size of the dispersed phase enhanced the activity of EOs and the interaction with target organisms. Moreover, the study reported that ME without active ingredient showed the ability to kill mites. It has been supposed that it could be due to the presence of sodium dodecyl benzene sulfonate (SDBS) in the mixture of surfactants. In fact, given its activity, it has been used to enhance the efficacy of the active ingredients [162].

Research aimed to the effective management of tick species has also been carried out. Chaisri et al. tested the activity of citronella EO on *Rhipicephalus microplus* (Canestrini) [163]. In this study, results have been expressed as larval and adult mortality. ME showed higher acaricidal efficacy compared with the pure citronella EO. In particular, larval mortality after 24 h occurred at the concentration of 0.78% EO-based ME in respect to the concentration of 3.125% of free EO. Also in this case, it could be supposed that the small size of oily droplets, <50 nm, and the presence of surfactants, Tween 20/propylene glycol 3:1, gave a synergistic effect. In particular, surfactants could interfere with the lipids of mites epicuticle, favouring the penetration of active ingredients [153,164].

dos Santos et al. proved the use of cinnamon (*Cinnamomum verum* J. Presl) EO as efficient tool to control ticks on cattle [165]. Indeed, this EO was evaluated against *R. microplus* through both *in vitro* and *in vivo* tests, the latter performed on infested dairy cows. Authors also formulated nanocapsules and NEs. They resulted to be very useful for the exploitation of cinnamon EO acaricidal activity. In fact, nanoencapsulated EO showed to be effective at low concentration (0.5%), ten times lower than that of pure EO (5%). Thus, such nanosystems at 0.5% were able to reduce infestation, oviposition and fertility of *R. microplus*. In fact, the encapsulation of EOs produced an improvement of the active ingredient stability and of its protection and guaranteed sustained release over time.

Nevertheless, the advantages of nanotechnology cannot be ever observed. Galli et al. investigated the activity of *E. globulus* EO [166]. For the purpose, they used the same formulation, concentrations,

target and procedures of those previously reported. In this case, pure EO showed to be effective decreasing the reproduction of ticks. On the contrary EO-based nanocapsules and NEs exerted low efficacy. However, it is possible to find an explanation of this result on the short exposure time (30 s) of the pests to nanosystems. This time should be not sufficient for the release of EO [167].

Mossa et al. recently investigated the acaricidal activity of emulsion and NE based on garlic (*Allium sativum* L.) EO on two eriophyid olive mites: *Aceria oleae* (Nalepa) and *Tegolophus hassani* (Keifer) [168]. After several stability studies, they found out a suitable and stable formulation, respect to the classical emulsion giving phase separation after two days. It was composed of 5% of garlic EOs, oil/Tween20 at 1:1.2 ratio and it was obtained through a sonication process for 35 min. Beyond the stability issue, garlic EO-based NE was demonstrated to be more effective than the respective emulsion. In fact, NE showed LC₅₀ values of 298.22 and 309.634 µg/mL on *A. oleae* and *T. hassani*, respectively, over to 584.878 and 677.830 µg/mL of the emulsion. Moreover, they proved to be safe for mammal administration as they did not produce toxicity in rats.

Badawy et al. formulated four different NEs based on two EOs—*Callistemon viminalis* (Sol. ex Gaertn.) G.Don and *Origanum vulgare* L.—and two monoterpenes—R-limonene and pulegone [169]. They investigated the activity of 10% concentrated NEs on *T. urticae* in terms of contact toxicity, fumigant toxicity and on bean plants under greenhouse conditions. Although all the formulations showed high efficacy, the monoterpene-based NEs proved to be more toxic against the target organism and with a more rapid outbreak of the activity. Moreover, the fumigant toxicity was more pronounced than contact toxicity. As mentioned above, this could be explained by the fact that such compounds are delivered on vapour phase and act mainly on the respiratory system [161].

6. Green Micro- and Nanoemulsions for Developing Antiparasitic Drugs

Micro- and nanoemulsions can also be useful tools to boost the bioactivity and increase the stability of antiparasitic drugs [170]. In the following paragraphs, we will outline the major achievements in the development of green micro- and nanoemulsions targeting both protozoan and helminth parasites.

6.1. Parasitic Protozoa

6.1.1. *Toxoplasma gondii*

The apicomplexan *Toxoplasma gondii* (Nicolle & Manceaux) infects approximately two billion people worldwide [171]; however, seroprevalence is declining in Western Countries [172].

New drugs are needed for the treatment of toxoplasmosis, particularly in immunocompromised patients or in congenitally infected subjects [173]. Among the new possible drugs, atovaquone is under evaluation for its ability to suppress protozoan parasites with a broad-spectrum activity. However, the use of this drug is limited by its extremely low water solubility and bioavailability. NEs prepared with atovaquone, based on grape seed oil using spontaneous emulsification method, showed increased bioavailability and efficacy for treatment of toxoplasmosis. In fact, *in vitro* this NE resulted active against *T. gondii*, using both RH and another strain (namely, the so-called Tehran strains), cultured on HeLa cells. Such results were confirmed in *in vivo* studies in mice treated orally; these resulted with a lower number of tissue cysts compared to animals treated with the standard preparation, by virtue of better bioavailability [174].

6.1.2. *Leishmania* spp.

They are vector-borne parasites belonging to *Leishmania* genus, order Trypanosomatida. They cause diseases with different clinical pictures: cutaneous (CL), mucocutaneous (MCL) and visceral (VL) [175].

Studies have been carried out on the effects of aromatic/heterocyclic sulphonamides, in the low nanomolar range, on the β-carbonic anhydrase (CA, EC 4.2.1.1) of *Leishmania* spp., which resulted effectively inhibited, without, however, any effect on parasite viability. The same drugs, formulated as

NEs in clove oil, inhibited the growth of either *Leishmania infantum* Nicolle or *Leishmania amazonensis* Lainson & Shaw, being less cytotoxic than the widely used antifungal amphotericin B, as revealed by haemolytic assay [176].

NEs as a delivery system for copaiba (*Copaifera* sp. Linnaeu) and andiroba (*Carapa guianensis* Aublet) oils (nanocopa and nanoandiroba with an average particle size of 76.1 and 88.1 nm, respectively) were tested on *L. infantum* (VL) and *L. amazonensis* (CL). Nanocopa and nanoandiroba resulted toxic to promastigotes of both *Leishmania* species. In particular, ultrastructural analyses by scanning electron microscopy showed a shift of the parasite to oval shape and the retraction of flagella, as early as 1 h after treatment, with concentrations near the IC₅₀ values. Furthermore, the treatment with such NEs reduced infectivity of the two species in macrophage cultures. Beneficial results were obtained also in mice experimentally infected with *L. amazonensis* or *L. infantum* (i.e., reduction in lesion size, parasite burden and inflammation). Animals affected by CL treated for eight weeks with NEs showed delay in lesion development. In VL model, around 50% reduction in parasite burden in liver and spleen of mice treated with nanocopa and nanoandiroba was found as compared with control untreated animals [177].

Nanotechnology has allowed the advancement of photodynamic therapy (PDT). In fact, many photosensitisers (PS), insoluble in water, need a nanocarrier as a physiologically acceptable carrier. NEs are efficient in solubilising liposoluble drugs, like the PS, in water. A zinc phthalocyanine (PS) oil-in-water NE, essential clove oil and polymeric surfactant (Pluronic® F127) for the formulation of a topical delivery system for use in PDT was used against *L. amazonensis* and *L. infantum*. The toxicity in the dark and the photobiological activity of the formulations were evaluated *in vitro* on *Leishmania* and macrophages. The zinc phthalocyanine NE was effective in PDT against *Leishmania* spp. with several advantages compared to other topical treatments like paromomycin and amphotericin B. These drugs have many disadvantages like local side effects and a very high cost, often limiting their use [178].

The antiparasitic activity of nanoemulsionated EO of a *Lavanudula* species was tested against *Leishmania major*, a species responsible for CL. In particular, NE with EO of *L. angustifolia* Mill. (where 1,8-cineol and linalool were the major components), as well as of *Rosmarinus officinalis* L., induced significant mortality of the parasite [179]. The NE of *L. angustifolia* and *R. officinalis* EOs showed antiparasitic effects that were much more significant than those obtained with the nonemulsionated EO of *R. officinalis* [180].

A taxonomically related parasite to *Leishmania* is *Trypanosoma evansi* Steel, the etiological agent of the disease known as "Surra" and "Mal das Cadeiras" which affects horses in Brazil, and sometimes also humans. The *in vitro* trypanocidal activity of the nanoemulsified *Schinus molle* L. EO was tested; this NE reduced the number of living parasites even totally, when the highest concentration was used (1%) contrary to the non-emulsified EO, which gave only 68% of mortality as a maximum [181].

6.1.3. *Plasmodium* spp.

Plasmodium parasites cause malaria, a disease which represents one of the major public health problem at global level with 219 million cases of malaria and 435,000 deaths estimated in 2017, particularly concentrated in Africa [182].

NEs loaded with arteether (ART), a semisynthetic derivative of artemisinin, by virtue of their solubility and consequently bioavailability, enhanced efficacy against *Plasmodium yoelii nigeriensis*, in a mouse model of experimental malaria. The *in vitro* release profile of the ART-NEs showed 62% drug release within 12 h; no significant effect on cell viability was observed. The authors focused the attention on a particular NE, loaded with ART (ART-NE), ART-NE-V, which showed a significantly enhanced bioavailability. This NE was well tolerated in the experimentally infected mice with no abnormality in behaviour, food/water consumption and general activity of the animals throughout the treatment and post treatment period. ART-NE-V, administered orally, had an 80% curative rate in comparison to the 100% cure rate achieved by intramuscular route at the same dose and to the 30% curative rate obtained in mice treated with ART in ground nut oil [183].

6.2. Helminths

Echinococcus granulosus

This parasite is the aetiological agent of cystic echinococcosis (CE), a zoonotic infection with economic and public health importance worldwide distributed. CE can result in a substantial human disease burden and have a relevant economic impact on animal productivity [184,185].

EOs from *Zataria multiflora* Boiss. were tested on the cestode *Echinococcus granulosus sensu lato* [186]. The effect was tested on the protoscoleces, isolated in liver hydatid cysts collected from naturally infected sheep. NEs at different concentrations (1–2 mg/mL) induced mortality levels up to 100% after 20 and 10 min, respectively, a scolicidal activity significantly higher than that obtained with nonemulsified oil [187]. *In vivo* studies in infected mice showed that the largest cysts were significantly reduced in size, as well as their total number, in animals treated with NE, compared to those treated with nonemulsified oil [186].

The *in vitro* and *ex vivo* activity of *Melaleuca alternifolia* (Maiden & Betche) Cheel oil (tea tree oil (TTO)), its NE formulation (NE-TTO) and its major component (terpinen-4-ol) were evaluated for their effects against *Echinococcus ortleppi* (another *Echinococcus* species, also known as G5 and clearly closely related to the genotypes of *E. canadensis*). This *Echinococcus* species infects cattle, which represents the principal intermediate host, mainly distributed in Europe, Africa, some areas of Asia and South America [188]. In *ex-vivo* studies the TTO, NE-TTO and the terpinen-4-ol were directly injected in the cysts isolated from cattle. The protoscolicidal action of the TTO major compound, terpinen-4-ol, resulted very promising. In fact, just after 5 min of exposure, non-viable *E. ortleppi* protoscoleces were obtained, at the concentration of 2 mg/mL. The results obtained in this study showed protoscolicidal effect at all tested formulations and concentrations. However, the effects of TTO were higher than those of NE-TTO but this latter had the ability to reduce the volatilisation of the compound and consequently to increase the protoscolicidal effect at the action site [189].

7. Green Formulations against Nematodes Attacking Plants

Meloidogyne spp.

The root-knot nematodes (*Meloidogyne* spp.) are key pests threatening several crops of economic importance. Their control is mainly based on the use of chemical nematicides. However, following the withdrawal of several synthetic nematicides because of their detrimental effects on soil biodiversity, natural products of botanical origin have been investigated for their possible use against these agricultural pests. Indeed, besides effectiveness for nematode control, botanicals assure beneficial effects on structure and residual life (e.g., microorganisms) of the soil. Among the most promising natural substances with nematicidal activity, glucosinolates, isothiocyanates, aliphatic acids (e.g., acetic, butyric, hexanoic and decanoic acids), alkaloids, piperamides, flavonoids (e.g., quercetin-7-glucoside), limonoids (azadirachtin, meliacins), quassinoids (e.g., chaparrinone, glaucarubolone, klaineanone, samaderines B and E), saponins and triterpene acids (e.g., 11-oxo triterpenic, pomolic, lantanolic, lantoic, camarin, lantacin, camarinin and ursolic acids), cyanogenic glycosides, polyacetylenes, phenolic acids (e.g., salicylic, gallic, *p*-hydroxybenzoic, vanillic, caffeic, and ferulic acids), fatty acids (e.g., linoleic and oleic acids) and volatile compounds (e.g., ascaridole, 2-undecanone, furfural, benzaldehyde, thymol, geraniol, eugenol, linalool, decenal and decadienal) are the most important ones [190–193]. Among them, isothiocyanates and neem azadirachtin have been encapsulated in marketed formulations effective against the growth and development of *Meloidogyne* spp. with limited effects on soil biodiversity [193,194]. Also, the EOs from *Foeniculum vulgare* Mill., *Pimpinella anisum* L., *Eucalyptus melliodora* A Cunn ex Schauer, *Origanum vulgare* L., *O. dictamnus* L., *Mentha pulegium* L. and *Melissa officinalis* L. were effective against *M. incognita* (Kof. & White) Chitwood showing EC₅₀ values of 0.2, 0.3, 0.8, 1.6, 1.7, 3.2 and 6.2 $\mu\text{L}\cdot\text{mL}^{-1}$, respectively [194,195]. Among their main constituents, benzaldehyde, γ -eudesmol, methyl chavicol, carvone, pulegone and (*E*)-anethole were ideal candidate ingredients for

nematicidal formulations [194,195]. On the other hand, efforts about formulating these botanical active ingredients in micro- and nanoemulsions remain limited, outlining the urgent need of future research.

8. Green Micro- and Nanoemulsions in the Real World

As reported above, researchers in entomology and parasitology are making great efforts for the improvement of pest control in terms of efficacy and safety for environment and human being. The potential of EOs and plant extracts as biopesticides and their exploitation through nanoencapsulation opened new challenging strategies for Integrated Pest/Vector Management (IPM/IVM). From the literature analysis (Scopus database, 27 June 2019), it can be observed that in the last 20 years approximately 100 documents were published concerning the employment of MEs and NEs for the vehiculation of pesticides (Figure 3). Interestingly, MEs were firstly studied and the maximum interest was reached around 2010. On the contrary, the use of NEs as pesticide formulations was more recent, reaching the highest attention in the last 2–3 years. Another aspect to be highlighted is represented by the nature of active ingredients employed as pesticide. Regarding MEs, the use of botanical and synthetic pesticides is almost the same along the years, while for NEs there is always a stronger prevalence (~70%) of studies on natural pesticides. These results seem to highlight a temporal correlation between the diffusion of biopesticides and the development of NEs for their application.

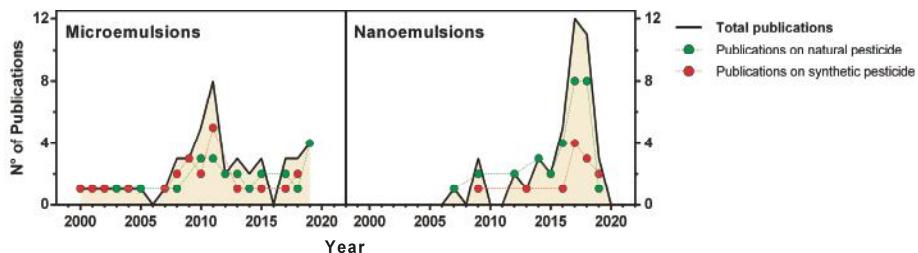


Figure 3. Number of publications on micro-(left) and nanoemulsion (right) vehicles for natural and synthetic pesticides per year.

Even though literature reported several studies with effective results, in the real world the exploitation of EO-based MEs and NEs is still limited. Currently, the pronounced effectiveness of chemical pesticides is still predominant respect to the eco-friendly advantages of the botanical ones. However, the common awareness about the need of a more sustainable world will likely lead towards a radical change in favour of the exploitation of green solutions in the near future.

Although some patents reported the nanoformulation of chemical pesticides or the nanoencapsulation of EOs [196,197], only few of them describe EO-based MEs or NEs as biopesticides. Enan et al. patented MEs as tool for the encapsulation and delivery of two or more EOs for pest control [198]. In particular, they used unsaturated C12-C26 fatty acids and/or salts and saturated C6-C14 fatty acids and/or salts as surfactants to enhance the activity of the ingredients, resulting in an improvement of the pesticide efficacy. According to the authors, this approach brings to a reduction of the active ingredient amount required to obtain an effective pest control.

Since scientific studies showed promising results, in the last years some botanical pesticides started to be available on the market. For example, Prev-Am® Plus is a fungicide and insecticide, based on orange (*Citrus x aurantium* L.) EO, that acts for direct contact. Since Prev-Am® Plus biodegrades rapidly and does not have a high environmental persistence, it is an excellent product for the Integrated Pest Management (IPM) programs, helping in the management of resistance and ensuring a minimal impact on beneficial insects. It can be used on a wide range of crops such as olive trees, vines and citrus fruits and it is allowed in organic agriculture.

Given the well-known repellent activity of EOs, a personal repellent based on EOs has been commercialised. Repel® is a spray containing 30% of lemon eucalyptus EO. It was proven to be able to

repel mosquitoes, in particular the vectors of Zika, West Nile, Dengue and Chikungunya viruses, for up to six hours.

Also, Bayer® launched on the market Requiem® EC, an emulsifiable concentrated formulation based on terpenes originally discovered in an insecticidal plant—*Dysphania ambrosioides* (L.) Mosyakin & Clemons. It is a contact insecticide/acaricide for use in the control or suppression of many foliar-feeding species, including aphids, thrips, plant-feeding mites, whiteflies, mealy bugs, leafminers, Lygus bugs, leafhoppers and moths attacking crops such as citrus, grapes, potatoes and others. Its low toxicity on mammalian and non-target organisms makes it a reduced-risk insecticide.

9. Regulatory Remarks

The EU regulates the botanical products used for the control of parasites, arthropod pests and vectors through two different regulations, the EC No. 1107/2009 and the EU No. 528/2012. The first one regards the plant protection products, addressing their risk evaluation and regulating the authorisation of commercialisation in the crop protection field. The second one, named Biocidal Products Regulation (BPR), takes into account “any substance or mixture exerting a controlling effect on any harmful organism by any means other than mere physical or mechanical action”.

Interestingly, while EC No. 1107/2009 does not mention nanomaterials at all, the BPR poses specific issue, stating that “where nanomaterials are used in that product, the risk to human health, animal health and the environment has been assessed separately”. For this reason, BPR excludes the possibility of “simplified authorisation procedure” followed for “low-risk” products, in the case of biocide containing nanomaterials. Moreover, the BPR highlights the necessity of a proper methodology for the risk evaluation for nanomaterials.

Although nanotechnology showed to be a great opportunity to achieve a more rational Integrated Pest Management (IPM), the lack of knowledge on the fate and effects on humans and environment of nanomaterials represents, nowadays, an important limitation on their widespread exploitation. It is needed an increased regulatory oversight to ensure their appropriate identification and risk assessment evaluation. In this direction, the European Community is addressing innovative methodologies able to evaluate the risk of nanopesticides and nanomaterials in general. In particular, the European Chemicals Agency (ECHA) is starting to define the guidelines for the monitoring and the evaluation of nanomaterials in the environment, and for the support about their registration procedure (four appendices for nanomaterials applicable to Chapters R.6, R.7a, R.7b and R.7c of the IR&CSA guidance) [199].

Among the different risk assessment procedures, the Quantitative Structure-Activity Relationship/Quantitative Structure-Property Relationship (QSAR/QSPR) appears one of the most promising tools for chemicals. In this regard, the scientific community is moving towards an innovative tool, nano-QSAR/QSPR, introducing the computational approach in the risk assessment of nanomaterials. Several studies focused on how nano-QSAR/QSPR should be supported by the development of new interpretative descriptors for the nanosystems. Moreover, they highlighted the need to model different classes of nanomaterials, given their wide variability in the molecular structure and mechanism of toxicity [200,201].

Currently, the most studied nanomaterials through nano-QSAR/QSPR for risk evaluation are metal oxide and carbon nanoparticles [202–204].

Although nano-QSAR/QSPR is showing to be a useful approach on the risk assessment on nanomaterials, it should be improved by increasing the experimental data on the toxicity of all the different nanomaterials classes, that are still restricted, allowing nano-QSAR/QSPR to be a real tool for the prediction of nanomaterials fate.

Even though much progress has been made, the efforts that are underway to improve the risk assessment procedures of nanomaterials should continue. A pragmatic and internationally accepted nanomaterial decision framework is necessary in order to clarify all the potential toxicological issues, opening to a large-scale diffusion of all the nano-based products.

10. Conclusions and Key Challenges for Future Research

Control of pests and vectors is a highly current issue since they are known to affect the health of the planet. Acting as vectors of devastating pathogens, many pests constitute a threat for the health and survival of living beings, as plants, animals and, above all, human beings. Although in the last decades chemical pesticides have been considered the solution to this problem, nowadays we are becoming aware that they are nothing more than a palliative. In fact, their efficacy has been overshadowed by two main drawbacks, the environmental hazards and the resistance development, linked to their overuse.

Nowadays, a possible solution has been found on the exploitation of botanical compounds, in particular EOs, which showed to possess antiparasitic, insecticidal, larvical, acaricidal, ovicidal, fumigant, repellent and chemosterilant effects among other biological properties. They could ensure a sustainable and eco-friendly way to control parasite and pest spreading. In this direction, several efforts have been done in the scientific research fields. For example, several botanical species have been deeply investigated to find out a high number of new active compounds. Anyway, suitable and innovative solutions could be reached only through a multidisciplinary approach. In fact, the physicochemical limits of biological compounds could be overcome only thanks to the development of suitable formulations. For this reason, technological research could offer the real solution to exploit the great advantages and the effectiveness of botanical compounds. Besides insecticides and acaricides, this is also true also for the development of new nematicides, as well as to develop drugs against parasites of public health importance.

In this scenario, nanotechnologies represent the tool of choice. Since they can encapsulate the active compound in a suitable way to protect them and, at the same time, to exalt their efficacy, botanical compound-based nanosystems could represent the turning point in the pest management. Among the different nanosystems available, the MEs and NEs proved to be the most suitable as vehicles for botanicals when those are characterised by high lipophilicity.

Although promising results have been reported in the literature, a strong gap between the theoretical research and the practical application still persists. In this direction, in the near future it is necessary to improve and examine in-depth different aspects of green nanotechnologies; in particular, (i) industrialisation of botanical species plantation in order to increase the amount and the yield of active ingredients, (ii) standardisation of products in terms of quali-quantitative composition, (iii) optimisation of the formulation process to enhance the stability and efficacy of nanosystems, (iv) reduction of the costs of production, (v) evaluation of the real long-term effects of the new products on the environment and non-target organisms and (vi) definition of a clear normative framework able to facilitate the commercial authorisation of botanical compound-based nanosystems.

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