

Name:

Class:



## C4.2 Pre-Unit Quiz: Extraction of Metals

1. Choose the correct general word equation for the reaction of metals and acid. [1]

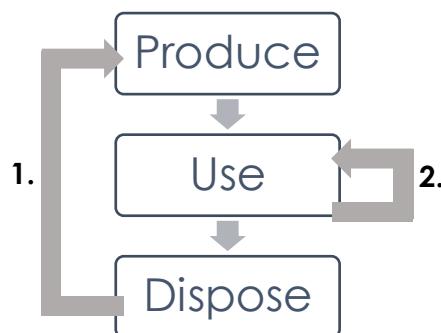
Tick () **one** box.

A. Acid + metal → salt + water

B. Acid + metal → salt + hydrogen

C. Acid + metal → salt + carbon dioxide

2. The diagram below shows ways to save the limited resources that are on Earth.



Complete the gaps in the diagram. [1]

Tick () **one** box.

A. 1 = recycle      2 = reduce

B. 1 = reduce      2 = recycle

C. 1 = recycle      2 = reuse

D. 1 = reuse      2 = recycle

3. The atomic structure of metals relates to their position on the Periodic Table.



Name:

Class:

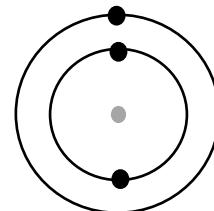


Science  
Master

In which group of the Periodic Table would you find the element represented by this electronic configuration? [1]

Tick () **one** box.

A. Group 2



B. Group 3

C. Group 1

4. A finite resource is a resource that... [1]

Tick () **one** box.

A. is being used up more quickly than it is being made.

B. is being made more quickly than it is being used up.

C. is running out.

5. Choose which describes a property of alkali metals. [1]

Tick () **one** box.

A. Unreactive

B. Very high melting point

C. Soft

6. Choose the product of the following chemical reaction.



Name:

Class:



[1]

Tick (✓) one box.

A. Lithium fluorate

B. Lithium fluoride

C. Lithium fluorine

7. Choose the correct electronic configuration of a sodium atom.

The atomic number of sodium is 11.

[1]

Tick (✓) one box.

A. 2,8,1

B. 2,9

C. 4,7

8. Choose the best explanation for why distillation isn't often used to obtain drinking water from salt-water.

[1]

Tick (✓) one box.

A. The boiling points of water and salt are very close so they are hard to separate accurately by distillation

B. Only small volumes can be distilled

C. It is very expensive due to the energy required

9. This box shows the reactivity series.

Choose a metal that would displace aluminium from

Potassium	Increasing reactivity
Sodium	
Calcium	
Aluminium	
Carbon	
Iron	
Tin	
Lead	
Hg	
U	

Name:

Class:



Science  
**Mastery**

aluminium oxide.

[1]

Tick () **one** box.

A. Iron

B. Platinum

C. Sodium

10. Iron oxide is an iron ore found on Earth.

Choose which method would extract the iron from iron oxide.

[1]

Tick () **one** box.

A. React with a more reactive metal to displace the iron

B. Filter the iron oxide to separate out the iron

C. Add an acid to separate the iron and the oxygen

