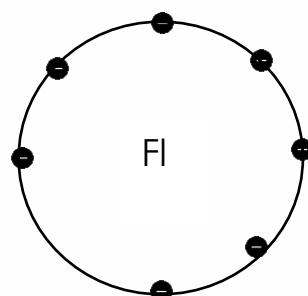
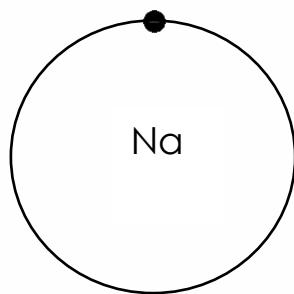


Read the exam style question carefully, then fill in each section below.

Question:

The diagram below shows the outer electrons in an atom of sodium and an atom of fluorine.



Sodium forms an ionic compound with fluorine.

Describe what happens when an atom of sodium reacts with an atom of fluorine.

You should describe the transfer of electrons in your answer.

Give the formulae of the ions formed.

(6)

Section 1: At first glance

1. What **command words** are used in this question? Circle them clearly.
2. **Underline the key information** in the question above.
2. **How many marks** is this question worth?

Section 2: Thinking ahead

Read the question again.



What do you need to know in order to answer this question really well?

Can you split the question into two or more parts?

Are there any labelled diagrams that might help you to show your answer?

What are the key words that you should include in your answer?

Section 3: Space to plan

Use this space to plan your answer.

Section 4: Answer the question



Section 5: Check your answer

Answer	Mark
Sodium atom <u>loses</u> electron	1
(to form) a 1+ sodium ion/ Na^+ ion	1
Fluorine atom <u>gains</u> electron(s)	1
(to form) 1- <u>fluoride ion</u> / F^- ion	1
<u>1</u> electron transferred	1
(positive and negative) ions are held together by electrostatic forces of attraction	1