

1. Read the question and the student's answer carefully.
2. Use the mark scheme to award the student a number of marks and annotate their answer with suggestions to improve.

Stretch: Rewrite the answer to show how it should be done!

Question:

Copper sulfate solution was electrolysed.

Explain why a solid forms at the negative electrode.

(4)

Student answer:

The solid at the anode is sulphur. Things are reduced at the negative electrode because electrons are gained. For example S^+ gains one electron to form $S(s)$.

Marks awarded= _____

Mark scheme:

| Answer | Mark |
|--|------|
| Copper ions (Cu^{2+}) are attracted to the negative electrode/cathode | 1 |
| Copper ions are reduced/gain electrons | 1 |
| To form copper atoms | 1 |
| Copper metal is formed | 1 |