

1. Read the question and the student's answer carefully.
2. Use the mark scheme to award the student a number of marks and annotate their answer with suggestions to improve.

**Stretch:** Rewrite the answer to show how it should be done!

**Question:**

Copper sulfate solution was electrolysed.

Explain why a solid forms at the negative electrode.

(4)

**Student answer:**

The solid at the anode is sulphur. Things are reduced at the negative electrode because electrons are gained. For example  $S^+$  gains one electron to form  $S(s)$ .

Marks awarded= \_\_\_\_\_

**Mark scheme:**

<b>Answer</b>	<b>Mark</b>
Copper ions ( $\text{Cu}^{2+}$ ) are attracted to the negative electrode/cathode	1
Copper ions are reduced/gain electrons	1
To form copper atoms	1
Copper metal is formed	1