

# Student Booklet

P4.1 Matter

Combined Science

Science  
**Mastery**



Ark**Curriculum+**

# The Big Idea

## Matter

*Why do some substances exist as solids and others as gases?  
What is the difference between solids and liquids? Why do some  
objects float on water and others sink? How can the density of  
objects be compared?*

Matter makes up everything. All objects in the universe are made of particles and it is the arrangement of these particles that determines their properties. Different materials can exist as solids, liquids or gases at room temperature, which means their particles are arranged in different ways.



This is the **third** unit we are studying as part of the big idea: **Energy is Conserved**

In this unit we will learn about energy in particles and how they are arranged in the different states of matter and their properties. We will look at density and how to measure it, as well as how particles exert pressure.

We will develop our mathematical skills in this unit by practising substitutions into equations.

We will develop our practical enquiry skills in this unit by doing an investigation into how the density of regular and irregular shaped objects can be measured.

### TASKS:

What subject will this unit focus on?  
(circle the correct subject)

BIOLOGY

CHEMISTRY

PHYSICS

There are lots of keywords underlined above. List these into the two columns:

Words I know	Words I haven't seen before

**To answer before the unit:**

1. What are you most excited to learn about in this topic?

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2. What do you already know about this topic?

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3. Why do you think it's important to learn about how energy is conserved?

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4. What knowledge from previous science lessons might help us?

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5. What questions do you have about this topic?

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**To answer at the end of the unit:**

1. Tick off any words in the 'words I haven't seen before' column that you are now confident with. Circle any you still need more practice to use.

2. What have you most enjoyed about this unit?

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3. What more would you like to learn about forces as part of the big idea: 'energy is conserved'?

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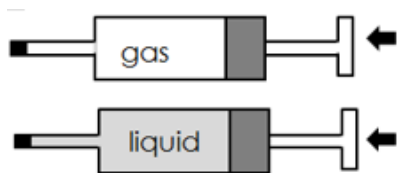
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# Pre-Test

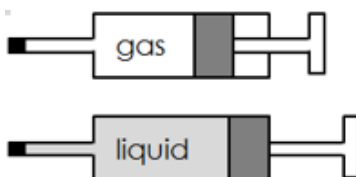
1. In which state(s) of matter are particles free to move? [1]  
Tick (✓) **one** box.
- (a) Solids and liquids ☐
- (b) Liquids and gases ☐
- (c) Gases only ☐
2. In which state of matter do particles have the most energy? [1]  
Tick (✓) **one** box.
- (a) Solid ☐
- (b) Liquid ☐
- (c) Gas ☐
3. Which state(s) of matter can be compressed? [1]  
Tick (✓) **one** box.
- (a) Gases only ☐
- (b) Liquids and gases ☐
- (c) Solids, liquids and gases ☐
4. Water is boiled in a kettle. Which statement correctly describes what has happened to the particles? [1]  
Tick (✓) **one** box.
- (a) Water particles (liquid) have turned into steam particles (gas) ☐
- (b) Particles have more energy so they are able to move more quickly in random directions ☐
- (c) Water particles have chemically reacted with oxygen to form steam ☐

5. A student filled and sealed two syringes, one with a gas and the other with a liquid, as shown in **figure 1**.



**Figure 1**

**Figure 2** shows the syringes after the plungers were pushed.



**Figure 2**

Choose the statement that explains why the gas was compressed more than the liquid. [1]

Tick (✓) **one** box.

- (a) The particles in gases are more spread out than the particles in liquids
- (b) The particles in gases have less mass than the particles in liquids.
- (c) The particles in gases are softer than the particles in liquids

☐☐☐

6. Which best explains why oil floats on water?

[1]

Tick (✓) **one** box.

- (a) Oil is lighter than water
- (b) Oil is less dense than water
- (c) Oil is more dense than water

☐☐☐

7. Which state of matter has the greatest density?

[1]

Tick (✓) **one** box.

- (a) Solid
- (b) Liquid
- (c) Gas

☐☐☐

8. Choose the correct definition of internal energy.

[1]

Tick (✓) **one** box.

(a) The total kinetic energy of all the particles in a system

☐

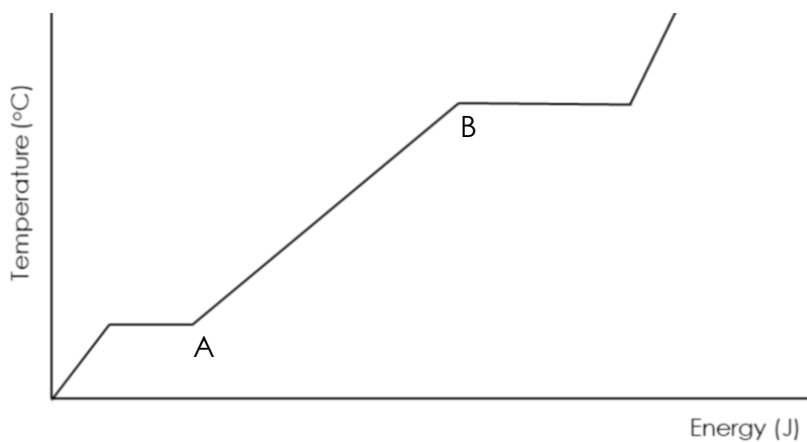
(b) The total kinetic and potential energy of all the particles in a system

☐

(c) The total energy transferred when a substance changes state

☐

9. A student has plotted a heating curve of water.



Choose the statement that is correct for line AB.

Tick (✓) **one** box.

(a) Kinetic energy and potential energy are increasing

☐

(b) Kinetic energy and internal energy are increasing

☐

(c) Potential energy and internal energy are increasing

☐

10. An object has a mass of 10 g and a volume of 2 cm<sup>3</sup>.

Choose the density of this object. [1]

Tick (✓) **one** box.

A. 0.2 g/cm<sup>3</sup>

☐

B. 5 g/cm<sup>3</sup>

☐

C. 20 g/cm<sup>3</sup>

☐

**Total = \_\_\_\_ /10**

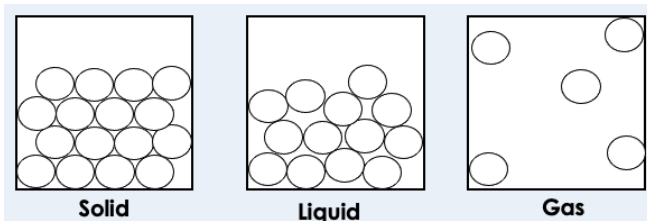
10	B	1
9	B	1
8	B	1
7	A	1
6	B	1
5	A	1
4	B	1
3	A	1
2	C	1
1	B	1
<b>Qu</b>	<b>Answer</b>	<b>Marks</b>

End of Unit Pre-Test. Turn over to see the answers. Give yourself a mark out of 10.

# Knowledge Organiser

## Particles and Density

1. **Particle diagrams** can be used to represent the arrangement and movement of particles in solids, liquids and gases.



2. Solids are the most dense state of matter as the particles are held most closely together due to the **forces of attraction**.

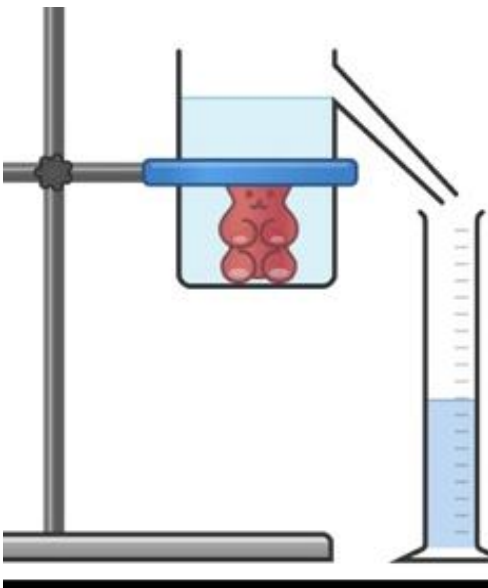
3. **Density** is the **mass per unit volume**.

4. Density can be calculated using the equation:

$$\text{Density} = \text{mass/volume}$$
$$\rho = m/V$$

with density,  $\rho$ , in kilograms per metre cubed,  $\text{kg/m}^3$ ; mass,  $m$ , in kilograms,  $\text{kg}$ ; volume,  $V$ , in metres cubed,  $\text{m}^3$

5. The density of a **regular** shaped solid can be calculated by measuring its **mass** and **volume**, then using the equation.
6. The density of an **irregular** solid or liquid can be determined using its **mass** and **displacement** of liquid.



## Gas Pressure

7. A **fluid** is a substance with **no fixed shape** - a liquid or a gas

8. Gas pressure is caused by **collisions** of particles with the walls of a container

9. Pressure is measured in Pascals (Pa)

10. Changing the temperature of a gas at constant volume changes the pressure exerted by the gas

11. Particles at **higher temperatures**, have higher thermal energy and move more quickly, so they have a **higher pressure**

# Glossary

<b>Degrees Celsius</b>	The unit used for temperature. <i>The melting point of water is 0 <b>degrees Celsius</b> (°C).</i>
<b>Density</b>	The mass per unit volume. <i>Warm fluids have a lower <b>density</b> than cold fluids, causing them to rise.</i>
<b>Displacement</b>	The movement of something from its original position. <i>Irregularly shaped solids can be placed into a <b>displacement</b> can to determine the volume of water that is displaced.</i>
<b>Fluid</b>	A substance with no fixed shape: a liquid or gas. <i>Convection is thermal transfer when particles in a heated <b>fluid</b> rise.</i>
<b>Incompressible</b>	Cannot be compressed (has a fixed volume). <i>Liquids and solids are <b>incompressible</b> but gases can be compressed.</i>
<b>Internal energy</b>	The total kinetic energy and potential energy of all the particles in a system. <i>When a substance is heated, its <b>internal energy</b> increases.</i>
<b>Irregular shape</b>	An object that has sides and angles of any length and size, so is not a cube, cuboid, cylinder etc. <i>A jelly baby has an <b>irregular shape</b>.</i>
<b>Kinetic energy</b>	A store of energy that any object or particle has when moving. <i>Particles in a gas have the greatest store of <b>kinetic energy</b>.</i>
<b>Mass</b>	The amount of matter in an object. <i><b>Mass</b> is measured in kilograms (kg).</i>
<b>Potential energy</b>	A store of energy related to the position of objects or particles. <i>Particles in a gas have the greatest store of <b>potential energy</b>.</i>
<b>Pressure</b>	The amount of force exerted per unit area. <i>Particles in a fluid exert <b>pressure</b> on any surface.</i>
<b>Regular shape</b>	An object that has sides and angles of equal sizes and lengths.

A **cube** has a regular shape.

**State**

The physical form in which a substance is in: solid, liquid or gas.  
*Melting and boiling are examples of changes of **state**.*

**System**

A body, object or group of bodies.  
*When looking at the internal energy of a **system**, you must consider the kinetic and potential energy of all of the particles in it.*

**Temperature**

Related to the average kinetic energy of particles in a system.  
**Temperature** is measured in °C.

**Volume**

The amount of space that a substance or object takes up.  
*Liquids and solids have a fixed **volume**.*

# Density

## Do Now

1. Put the states of matter in order of increasing internal energy.
2. In which state of matter are particles held most tightly together?
3. Compare the forces of attraction between particles in liquids and gases.
4. Define internal energy.
5. Name a piece of apparatus that can be used to measure the volume of a liquid.

## Drill:

1. State the formula to calculate the volume of a cube.
2. Calculate the volume of a cube with sides of length 4 cm.
3. Calculate the surface area of the cube in Q2.

## Read Now:

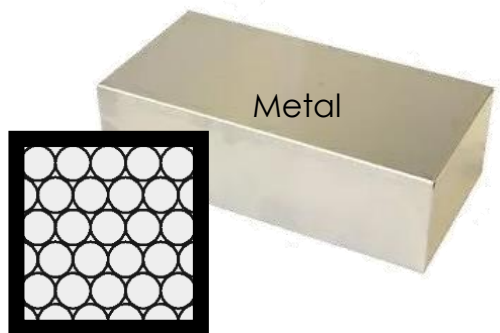
Matter makes up everything, and matter is made of atoms. Atoms can exist individually or as part of compounds or molecules. When we talk about matter, we can describe all these different arrangements of atoms as 'particles'. Density is a fundamental property of matter. It describes how close together the particles are, specifically the mass of particles within a given volume. The SI unit for mass is the kilogram, but often in chemistry the masses we work with are so small that we can use the gram.

1. Explain what matter is made of.
2. Define density.
3. State the SI unit of mass.
4. Identify which state of matter is the most dense.
5. How many grams are in 1 kg?

## Density

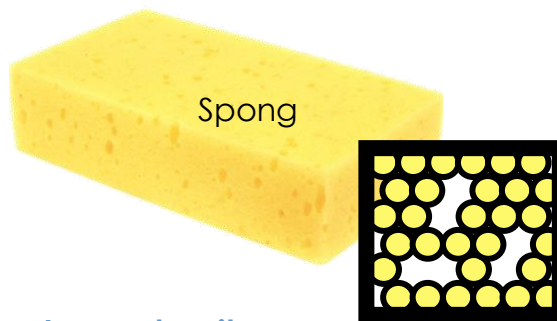
**The density of an object is the mass per unit volume.**

This means how much matter (particles) is packed into a certain volume.



**Higher density:**

More particles in a given volume



**Lower density:**

Fewer particles in a given volume

In order to calculate the density of an object, we need to know its **mass** and **volume**.

**Mass (m)** is measured in **kilograms (g)**.

**Volume (V)** is measured in **metres cubed (m³)**.

$$\rho = \frac{m}{V}$$

Diagram illustrating the density equation  $\rho = \frac{m}{V}$ :

- $\rho$  is labeled **Density** measured in  $(\text{kg}/\text{m}^3)$ .
- $m$  is labeled **Mass** measured in  $(\text{kg})$ .
- $V$  is labeled **Volume** measured in  $(\text{m}^3)$ .

This is an equation you will need to remember.

### Think About:

These three cubes are made of different materials.

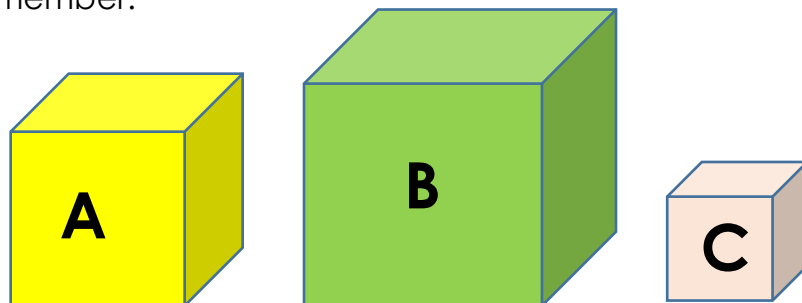
They all have the same mass.

What can be said about the density of these cubes?

Describe a method you could use to calculate the density of each.

Stretch:

If all the cubes had the same density, what would you expect to note of their masses?



## Drill

1. Define density.
2. Explain what density means in terms of particles.
3. Identify the state of matter which has the greatest density.
4. State the equation used to calculate density.
5. State the SI unit for mass.
6. State the SI unit for volume.
7. State the equation used to calculate the volume of a regularly-shaped solid.

**I Do:** Calculating mass, volume and density

This block of steel has a mass of 23 550 g and a volume of 3000 cm<sup>3</sup>.

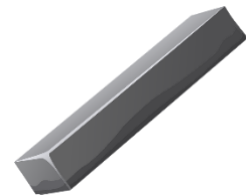
Calculate the density of the steel.

*HINT: What is the equation that links density, mass and volume?*

$$\text{Density} = \frac{\text{Mass}}{\text{Volume}}$$

$$\text{Density} = \frac{23\,550}{3\,000}$$

$$\text{Density} = 7.85 \text{ g/cm}^3$$



Another block of the same type of steel has a mass of 14 130 g.

Calculate its volume.

$$\text{Density} = \frac{\text{Mass}}{\text{Volume}}$$

$$7.85 = \frac{14\,130}{\text{Volume}}$$

$$\text{Volume} = \frac{14\,130}{7.85}$$

$$\text{Volume} = 1\,800 \text{ cm}^3$$

**You Do:** Calculating mass, volume and density

This plastic building block has a volume of  $2.2 \text{ cm}^3$ . Its mass is  $1.25 \text{ g}$ .

Calculate the density of the block.

*HINT: What is the equation that links density, mass and volume?*

The same material is used to make a brick that is double the volume.

Calculate the mass of the bigger brick.

## Exit Ticket

1. Which statement is correct?

- ☐ A. Density is a measure of how heavy the particles in a substance are
- ☐ B. Density is a measure of how many particles are in a given volume
- ☐ C. Density is a measure of how many particles there are in a substance

2. Two cubes have the same volume but different masses. Which statement is correct?

- ☐ A. The cube with the least mass is the most dense
- ☐ B. Both the cubes have the same density
- ☐ C. The cube with the greatest mass is the most dense

3. What is the mass of a block of iron that has a volume of  $1 \text{ m}^3$ ? Iron has a density of  $7800 \text{ kg/m}^3$ .

- ☐ A.  $7800 \text{ kg}$
- ☐ B.  $7.8 \text{ kg}$
- ☐ C.  $0.000128 \text{ kg}$

# Measuring Density

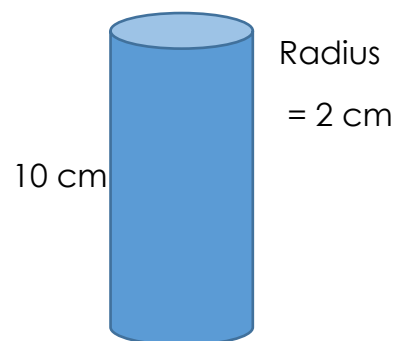
## Do Now:

1. Define density.
2. State the formula to calculate the volume of a cube.
3. Which state of matter has the lowest density?
4. State the volume shown in this image.
5. Convert this volume into litres.



## Drill:

1. State the formula to calculate area of a circle.
2. Calculate the area of one end of this cylinder.
3. Calculate the volume of this cylinder.



## Read Now:

Density is the mass per unit volume, calculated using the mass of a substance divided by its volume. For regularly shaped solids, such as cubes, cuboids and cylinders, we can use mathematical formula to calculate their volumes. For irregularly shaped solids we must use a different method, as these formulae do not apply. We can use a piece of equipment called a Eureka can to measure how much liquid is displaced when the object is placed into it. You can think about this like when you get into the bath and the water rises. If the bath was full to the top and you got into it, the volume of water that spilled out of the bath would be equal to your volume.

1. Define density.
2. Describe how to calculate the volume of a regularly shaped solid.
3. Give an example of a regularly shaped solid.
4. Name the piece of equipment that can be used to measure the volume of liquid displaced by an irregularly shaped solid.
5. This paragraph describes how to measure the volume of different solids. How could you measure the volume of a liquid?

## Measuring Mass

To measure the mass of an object we use a **balance**.

**Method for measuring mass:**

1. Make sure the balance is wiped clean. **Why?**
2. Check that the balance is on a level surface.
3. Turn on the balance.
4. Ensure that the unit on the display is grams (g).
5. Tare the balance by pressing the 'zero' or 'tare' button. **Why?**
6. Place the item you want to measure onto the balance and note the reading on the screen.

**Reading liquid volumes accurately**

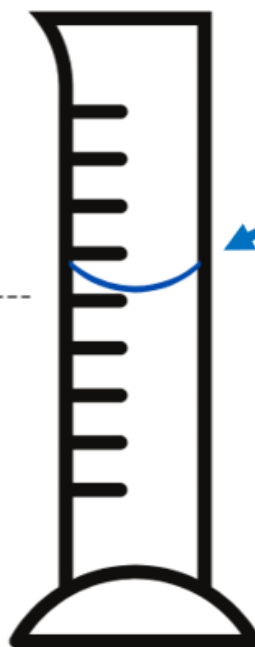
**1. Read from eye level**

*Bend down so your eyes are level with the water level*



**2. Read the scale from the bottom of the meniscus**

*Look at the line on the scale which is closest to the bottom of the 'U' shape.*



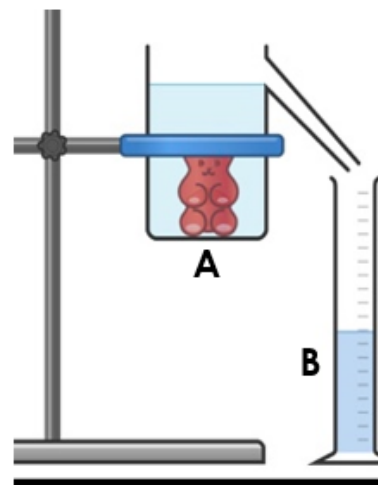
**The meniscus**

*Liquids in long thin containers form a 'U' shaped meniscus.*

## Using a displacement/eureka can

If an object has an **irregular shape**, the volume can be measured using a **displacement can**, or **Eureka can**.

1. Fill the displacement can (A) with water and wait until the water stops running out of the spout.
2. Place a graduated cylinder (B) under the spout.
3. Carefully lower the irregular object into the displacement can, collecting the displaced water in the graduated cylinder.



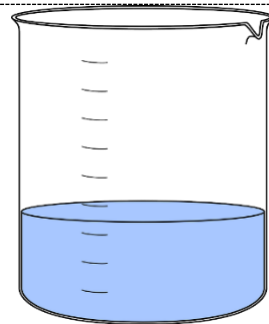
The **displaced** water in the cylinder occupies the same amount of space as the irregular object. **The volume of water in the graduated cylinder is equal to the volume of the object!**

## Drill

1. State the equation that links density, mass and volume.
2. Name the piece of apparatus used to measure mass.
3. Describe how to calculate the volume of a regularly shaped solid.
4. Describe how a displacement/eureka can is used to measure the volume of an irregularly shaped solid.
5. List any units that could be used to measure mass.
6. List any units that could be used to measure volume.
7. Describe how to read the measurement of a liquid.

**I: Describe: to recall facts, events or processes in an accurate way**

Describe how to determine the density of a liquid.



**Model answer:**

- Measure the volume using a measuring cylinder
- Measure the mass of the empty measuring cylinder using a mass balance
- Measure the mass of the measuring cylinder when the liquid is in it and subtract the mass of the empty cylinder
- Use the equation  $\text{density} = \text{mass}/\text{volume}$

**You: Describe: to recall facts, events or processes in an accurate way**

Describe how to determine the density of an irregularly shaped object.



**Answer:**

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**Exit Ticket**

1. Which measurement is taken using a displacement or eureka can?

- ☐ A. Density
- ☐ B. Mass
- ☐ C. Volume

2. Which statement is correct?

- ☐ A. The volume of water displaced by an irregularly shaped object is the same as the volume of the object
- ☐ B. The volume of water left in the eureka can is the same as the volume of the object
- ☐ C. Using the volume of water displaced by an object only works for irregularly shaped objects

3. What is the density of an object with mass 5 g that displaced 25 cm<sup>3</sup> of water?

- ☐ A. 125 g/cm<sup>3</sup>
- ☐ B. 0.2 g/cm<sup>3</sup>
- ☐ C. 5 g/cm<sup>3</sup>

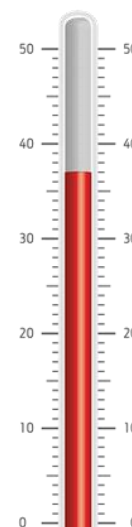
# Gas Pressure

## Do Now

1. Define density.
2. State the equation used to calculate density.
3. Describe the movement of particles in a gas.
4. Define internal energy.
5. Describe what happens to the movement of particles in a gas as they are heated.

## Skills Drill:

1. State the unit usually used to measure temperature.
2. State the temperature shown on the thermometer.
3. State the resolution of this thermometer.



## Read Now:

Particles in a gas can take the shape of their container because they are constantly moving randomly throughout the container. The gas particles frequently collide with other gas particles, as well as with the walls of the container. Gas pressure is a measure of how often and how hard the particles of a gas collide with the walls of the container. As gas particles are moving quickly, they exert a force on the wall of the container when they hit it, and the faster the particles are moving, the greater this force will be. On aerosol canisters, such as spray deodorants, there will be a safety warning that instructs the user not to heat the canister. This is because if the particles are heated, they will move more quickly and hit the walls of the container with greater force, causing the gas pressure to build up so much that the canister may explode! State the definition of 'micro-' as a prefix.

1. State two properties of gases.
2. Describe the movement of particles in a gas.
3. Explain what is meant by gas pressure.
4. Explain what happens to the force exerted by particles when they are heated.
5. Explain why it is not safe to heat an aerosol canister.

## Gas Pressure

Particles in a gas move at **random speeds** in **random directions**.

They **collide** with each other and with the walls of their container.

When they **collide** with the walls of the container, they exert a force on it. This is **gas pressure**.

Pressure is a measure of the **force** exerted per unit **area**.

When a substance is heated, its **internal energy** increases.

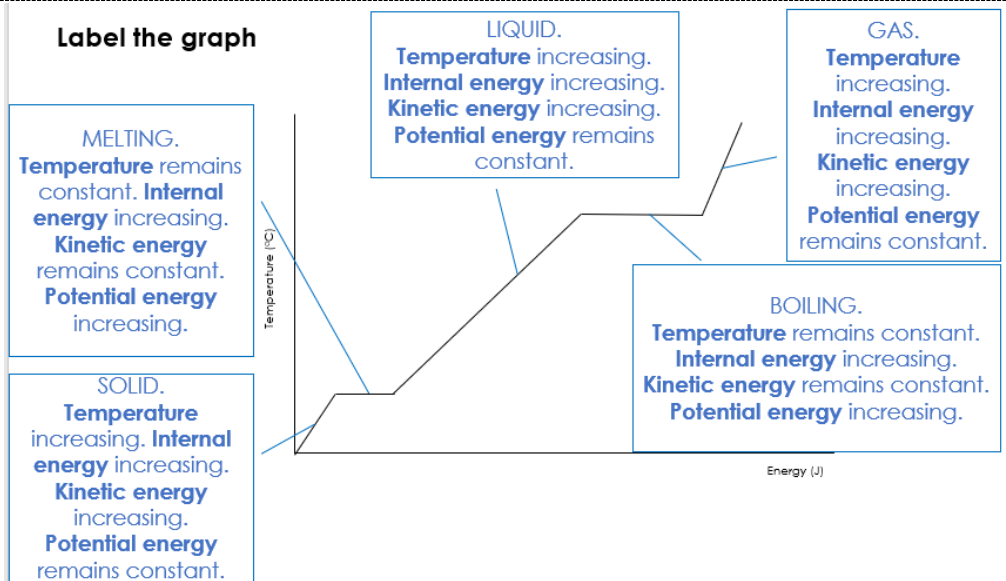
**Temperature** is a measure of the average **kinetic energy** of all the particles in a system.

When temperature is increased, gas particles **gain kinetic energy** and move more quickly.

The gas particles are moving **more quickly** so **collide more often** with the walls of the container and with **more force**.

This means they exert **more pressure**.

### Label the graph



## Drill

1. Describe the movement of particles in a gas.
2. Describe the properties of gases.
3. Describe what happens to the internal energy of a gas when it is heated.
4. Describe what happens to the movement of particles when a substance is cooled.
5. Describe the relationship between temperature and pressure.
6. Explain the relationship between temperature and pressure.

**Higher Tier only**

**Work is the transfer of energy by a force.**

When a force is applied to a gas, **work** is done on the gas.

This increases the **internal energy** of the gas and can cause an **increase** in its **temperature**.

**I: Explain: to use scientific understanding to make something clear or state the reason for something happening**

Explain what happens to the pressure of a gas as temperature increases.

**Model answer:**

- Increasing the temperature increases the kinetic energy of the particles in a gas
- This means the particles start to move more quickly and hit the surface of the container more often and with more force
- So, this increases the pressure exerted by the gas

**We: Explain: to use scientific understanding to make something clear or state the reason for something happening**

*Explain what happens to the pressure of a gas as temperature decreases.*

**You: Explain: to use scientific understanding to make something clear or state the reason for something happening**

*Explain the relationship between temperature and pressure of a gas.*

## Exit Ticket

1. Which describes the movement of particles in a gas?

- ☐ A. All particles move randomly at the same speed
- ☐ B. All particles move at different speeds in a pattern
- ☐ C. All particles move at random speeds in random directions

2. What would happen to the pressure in a sealed pot if it was heated?

- ☐ A. The pressure would increase
- ☐ B. The pressure would decrease
- ☐ C. The pressure would stay the same

3. Which statement is correct?

- ☐ A. When particles are heated they expand
- ☐ B. When particles are heated their internal energy increases
- ☐ C. When particles are heated they exert less pressure on the walls of a container



# Independent Practice

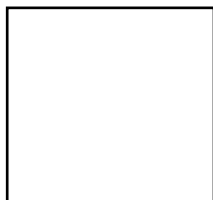
Density:.....	24
Measuring Density .....	27
Gas Pressure: .....	29

## Density:

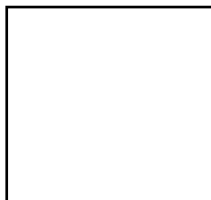
### Section A

1. There are three states of matter.

a. In the boxes below, sketch particle diagrams for the three states of matter.



Solid



Liquid



Gas

b. Which state of matter has the highest density?

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c. Use your particle diagrams to explain why.

---

---

d. Which is the correct equation to calculate density? Tick (✓) **one** box.

Density = mass x volume

☐

Density = mass

volume

☐

Density = volume

mass

☐

e. Density can be measured in  $\text{kg/m}^3$ . Which of these is another unit that could be used to measure density? Tick (✓) **one** box.

$\text{g/cm}$

☐

$\text{cm/g}^3$

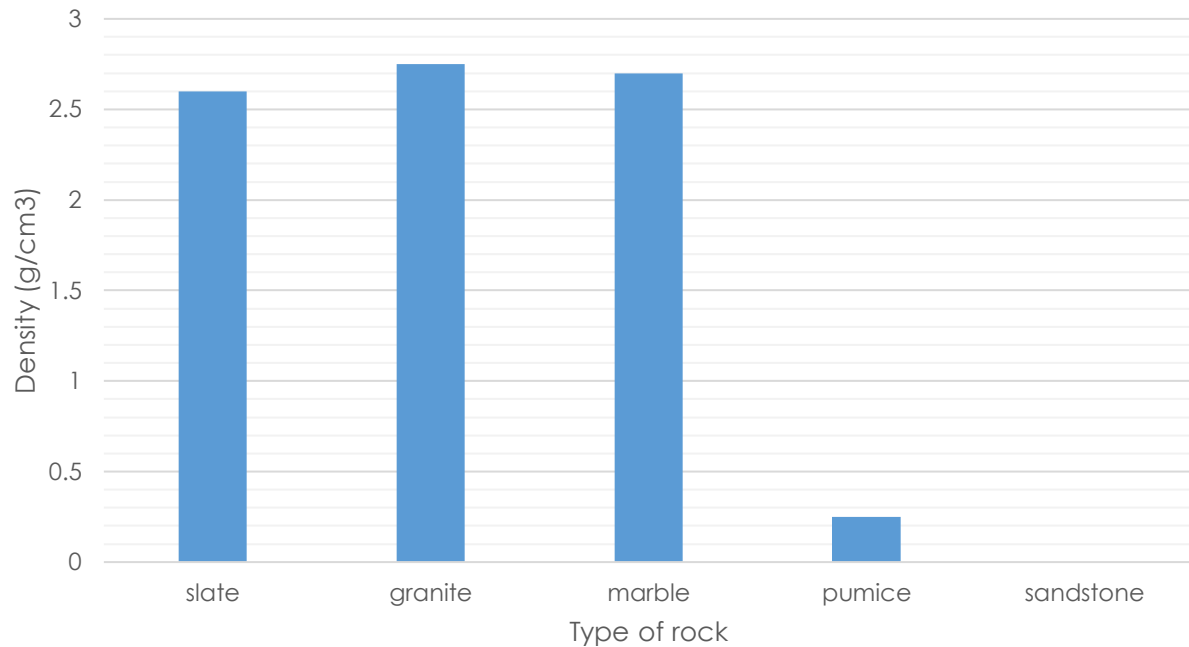
☐

$\text{g/cm}^3$

☐

## Section B

2. The graph below gives the density of several types of rock.



a. Which type of rock would contain the most particles in a 1 cm<sup>3</sup> sample?

---

b. What type of graph is this?

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c. Why is this type of graph suitable?

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d. A scientist has a sample of sandstone with a mass of 0.011 kg. What is this mass in grams?

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e. The sample of sandstone has a volume of 5 cm<sup>3</sup>. Calculate the density of the sandstone.

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f. Add this value to the graph above.

g. Calculate the volume of a 1.375 g piece of granite, using information from the graph.

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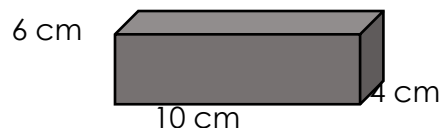
h. Calculate the mass of this block of slate, using information from the diagram and the graph.

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3. Helium-filled balloons float in air. Explain what this shows about the relative densities of air and helium and their particle arrangements.

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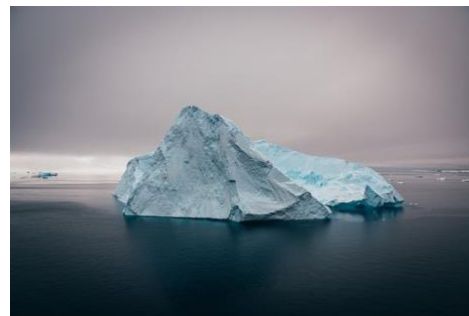
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### Section C

4. Water is a useful substance to demonstrate changes of state because its melting and boiling points can be achieved in the lab.

- State what water is called when it exists as each different state of matter.
- State the melting and boiling points of water.
- Describe what happens to the particles when water changes from liquid to solid.
- Define internal energy.
- In which state of matter do particles have the most internal energy.
- Describe what happens to internal energy when a substance melts from a solid to liquid.
- An iceberg has a volume of  $35 \text{ m}^3$  and a mass of 32 000 kg. Calculate the density of the iceberg.
- Explain why the iceberg floats on water.
- The iceberg will eventually melt. Explain why.



## Measuring Density

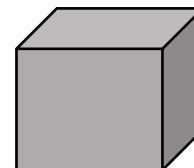
1. Read the question and the student's answer carefully.
2. Use the mark scheme to award the student a number of marks and annotate their answer with suggestions to improve.

**Stretch:** Rewrite the answer to show how it should be done!

### Question:

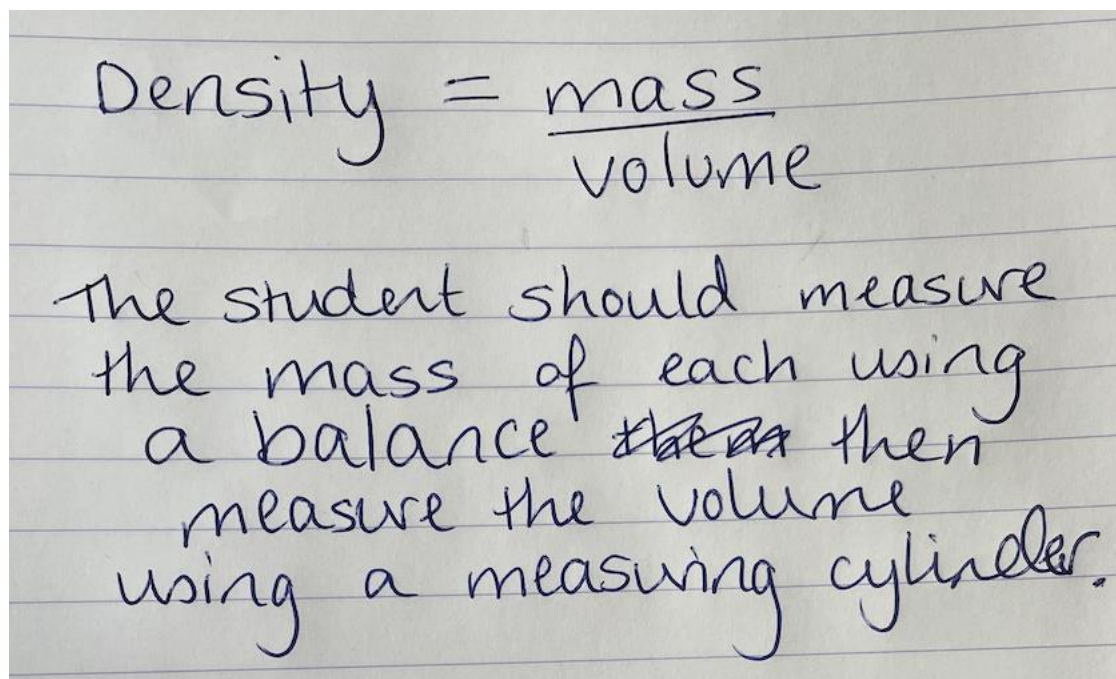
A student wants to compare the density of two objects.

Describe the method(s) that the student could use to measure the density of each object.



(6)

### Student answer:



**Marks awarded=** \_\_\_\_\_

## Mark scheme:

Level 0 (0 marks)	No relevant content
Level 1 (1-2 marks)	<b>Basic</b> descriptions of the measurements that should be taken for <b>one method</b> but no mention of how to use them, <b>or</b> descriptions of the quantities needed to calculate density but no information on how to measure them.
Level 2 (3-4 marks)	<b>Clear</b> description of <b>one method</b> to measure density <b>or</b> basic descriptions of both methods
Level 3 (5-6 marks)	<b>Clear</b> descriptions of <b>both methods</b> that would allow the density of each to be measured

For both

- Measure mass using a balance
- Calculate density using  $\rho = \frac{m}{V}$

Cube (regularly shaped solid)

- Measure length of sides using a ruler
- Calculate volume using  $lbh$

Screw (irregularly shaped solid)

- Use a displacement can/Eureka can
- Measure the volume of water displaced
- Volume of water displaced = volume of screw

## Gas Pressure:

### Section A

1. This question is about gas pressure.

a. Which is the correct description of how particles move in a gas?

Tick (✓) **one** box.

At random speeds in the same direction

☐

At random speeds in random directions

☐

At the same speed in the same direction

☐

At the same speed in random directions

☐

b. Gas pressure is exerted when particles collide with a surface of a container. Explain what this means.

---

c. What is the relationship between gas pressure and temperature?

Tick (✓) **one** box.

Increasing temperature increases pressure

☐

Increasing temperature decreases pressure

☐

Increasing temperature does not affect pressure

☐

d. Which property of gas particles is the temperature related to?

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## Section B

2. The image shows a syringe that contains air. The syringe is placed into an ice bath, which decreases the temperature of the air in the syringe.

a. What would happen to the average kinetic energy of the particles when the temperature is decreased?

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b. Explain why the average kinetic energy is used.

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c. What would happen to the average speed of the gas particles if the temperature decreased?

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d. What would happen to the gas pressure inside the syringe if the temperature decreased?

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e. The plunger of the syringe is pulled outwards. Why does air move into the syringe as the plunger is pulled outwards?

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f. The syringe is then placed into a hot water bath. Complete the table by ticking the correct box to identify what would happen to the following quantities.

	Increases	Decreases	Stays the same
Mass of the gas			
Density of the gas			
Average speed of particles of gas			
Pressure exerted by the gas			



## Section C

3. A group of scientists were investigating the behaviour of air particles at different temperatures. They used the same container throughout, which had a volume of  $1 \text{ m}^3$ . They cooled the air from  $22^\circ\text{C}$  to  $0^\circ\text{C}$ .
- j. Describe and explain what would happen to the particles of air in the container.
  - k. Describe and explain what would happen to the pressure within the container.
  - l. Is air an element, a compound or mixture? Explain your answer.
  - m. Water within the air froze at  $0^\circ\text{C}$ . The specific latent heat of fusion of water is  $330 \text{ kJ/kg}$  and the change in internal energy of the water was  $0.85 \text{ kJ}$ .  
Calculate the mass of ice that formed.  
Use the equation:  $\Delta E = m L$
  - n. The air also contains oxygen and nitrogen. The melting and boiling points of each substance is shown in the table.

	Melting point ( $^\circ\text{C}$ )	Boiling point ( $^\circ\text{C}$ )
Oxygen	-218	-183
Nitrogen	-210	-195

The apparatus the scientists were using allowed them to cool the air to  $-190^\circ\text{C}$ . What state would the water, the oxygen and the nitrogen be in at  $-190^\circ\text{C}$ ?