

Tuning Threshold Voltage in Organic Electrochemical Transistors by Varying Doping of the Conjugated Polymer p(g3T2-T)

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Preface

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Abstract

Organic Electrochemical transistors (OECTs) exhibit advantageous properties, such as high transconductance and steep-slope switching, while operating at very low voltages. Although, their switching speed is comparatively slower than solid-state devices, it remains sufficient for applications in bioelectronics [?]. The gold standard semiconductor for p-type OECTs is PEDOT:PSS. However, its main drawback lies in its depletion-mode operation, which requires power to turn off the device. To minimize power consumption and improve stability, efforts have been made to the design conjugated polymers that allow accumulation-mode devices. One such polymer, 3-(2-(2-methoxyethoxy)ethoxy)thiophene (p(g3T2-T)) has demonstrated negative threshold voltages close to zero and high transconductance [1]. Furthermore, by doping p(g3T2-T) at various levels and drop-casting it as a gate, it has been possible to fine-tune the threshold voltage [2]. This study aims to adapt a microstructuring method for fabricating side-gated OECT devices that comprises different doping levels of F₄TCNQ and F₆TCNNQ in p(g3T2-T) and a solid-state electrolyte [3], the latter is deposited by inkjet printing. Additionally, the study aims to adjust the threshold voltage by utilizing these varying doping levels, while analyzing the stability and performance of the doped devices.

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Chapter 1

Introduction

The field of organic electronics has witnessed significant advancements in recent years due to its potential biocompatibility, mechanical compliant, and other application-specific characteristics. Among the numerous types of organic devices, Organic Electrochemical Transistors (OECTs) have attracted considerable attention due to their unique capabilities such as high transconductance and steep-slope switching at low operation voltages, which give them potential for use in energy storage, bioelectronics and neuromorphic devices.

Accumulation-mode transistors, devices that are normally in the OFF state at zero-gate-biased condition, rely on the use of undoped conjugated polymers. In contrast, the ability to precisely control and tune threshold voltage of an OECT can be achieved by manipulating the doping level of the mentioned conjugated polymer as shown by Tan et al. [2]. However, they fabricated devices that did not follow a complete microstructuring technique, limiting their integration into circuits.

The primary objective of this research is to address this missing information by developing a microstructuring method to fabricate accumulation-mode OECTs with controlled doping levels and enable their seamless integration into circuits. However, during the pursuit of this objective, it was identified that stability of dopants in an electrochemical environment may pose challenges that need to be addressed as well. Therefore, this research project specifically aims to:

1. Characterize 3-(2-(2-(2-methoxyethoxy)ethoxy)ethoxy)thiophene (p(g3T2-T)) with varying doping levels of F₄TCNQ and F₆TCCNNQ. This involves chemically modifying the conjugated polymer with different concentrations of dopant and analyzing their electronic structure, morphology and electrical properties using techniques such as UV-Vis spectroscopy, Ultraviolet Photoelectron Spectroscopy (UPS), Van Der Pauw method, Electrical Impedance Spectroscopy (EIS) and Cyclic Voltammetry (CV),
2. fabricate OECT devices, which involves utilizing the conjugated polymer at different doping levels, and adapting an existing method that combines electrode patterning techniques, spin-coating, photolithography and inkjet printing. The devices will be carefully optimized to ensure reproducibility and stability,

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3. assess doped polymer stability in OEET, which involves performing conductivity measurements over time and selecting an electrolyte composition that allows an stable performance, and
4. investigate the shift in threshold voltage through electrical characterization of the fabricated OEET devices with varying doping levels of F₄TCNQ and F₆TCNNQ.

The thesis is structured as follows: Chapter 1 provides an overview of organic electronics and the importance of the OEET threshold voltage, and outlines the motivation, goals, and structure of the thesis. Chapter 2 presents a comprehensive review of the relevant background information on Organic Semiconductors (OSCs), Organic Mixed Ionic Electronic Conductors (OMIECs), and Organic Electrochemical Transistors, and relevant research on tuning the threshold voltage of OEETs. Chapter 3 illustrates the Experimental Methods used in this research, describes the materials, equipment, software, and procedures to chemically doped the films and characterization method employed in the study. Finally, it outlines the fabrication and characterization process of OEET devices. In Chapter 4, the experimental results obtained are presented, from the characterization of the conjugated polymer at different doping levels and fabricated OEET devices. Analysis of the relationship between the doping level and the shift in threshold voltage. Finally, it discusses the implications of the findings and their relevance to the field. Lastly, Chapter 5 provides a summary of the research objectives and the extent to which they were achieved, suggests future research direction, and potential applications.

By addressing the aforementioned research goals and following the proposed thesis structure, this study aims to contribute in adapting an existing protocol to the conjugated material p(g3T2-T), in understanding the doping-dependent tuning of OEET threshold voltage and establish a foundation for the development of high-performance organic electronic devices.

Chapter 2

Background

2.1 Organic Semiconductors

Unlike inorganic semiconductors, organic semiconductors are lightweight, easy to chemically tune, mechanical flexible, and possess low-cost and low-temperature processing. All of these characteristics are responsible for the increased attention to this type of materials in the field of organic electronics.

2.1.1 Electronic Structure and Transport

Organic semiconductors are π -conjugated molecules that comprise mostly carbon and hydrogen atoms, with alternating multiple (sp^2 hybridization) and single (sp hybridization) bonds. The wavefunction of the sp^2 orbitals overlap so that electrons can be delocalized. Based on the size of the conjugated system, organic semiconductors can be divided into conjugated polymers and small molecules, the latter has the advantage of being ordered and its synthesis allow to obtain high purity material [4].

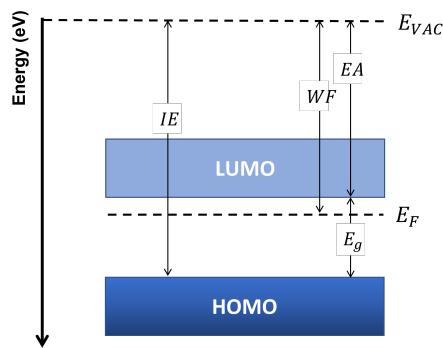


FIGURE 2.1: Energy level diagram of an n-type organic semiconductor.

The energy structure of organic semiconductors comprises a highest occupied molecular orbital (HOMO) and a lowest unoccupied molecular orbital (LUMO), that are analogous to the valence and conduction bands, respectively, from inorganic

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semiconductors. The difference between both energy levels corresponds to the band-gap (E_g) of the material, as represented in Figure 2.1. We could also defined i) the Fermi level (E_F) using the material work function (WF), ii) the ionization energy (IE, also refer as ionization potential, IP), using the HOMO energy, and iii) the electron affinity (EA), using the LUMO energy.

An organic semiconductor material could also be classified depending on whether the ground state is degenerate or non-degenerate. The former describes monomers that are energetically equivalent in the ground state, and in the latter, the opposite happens, commonly seen due to the energy difference of aromatic (benzoid) and quinoid structures [5], as exhibited in Figure 2.2.

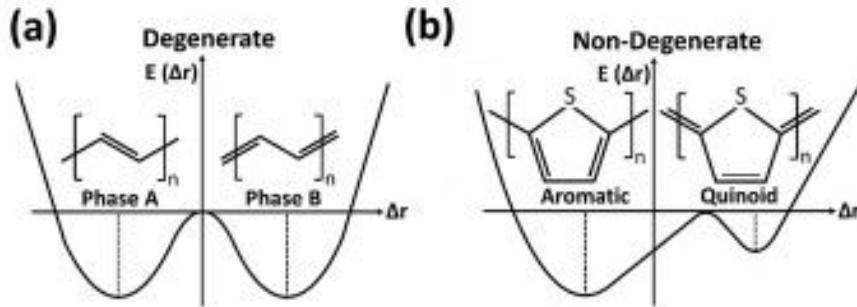


FIGURE 2.2: Potential energy change (electronic plus lattice distortion energy) in a) degenerate and b) non-degenerate ground-state-conjugated polymers. Extracted from reference [6].

Transport in an organic semiconductor creates changes in the bond-length alternation pattern. Since charge transport come together with a chain distortion, new quasiparticles are defined: solitons in degenerated molecules and polarons in non-degenerated molecules [7].

2.1.2 Molecular Doping

The basic principles are similar than in inorganic materials. A donor or acceptor entity is added to generate electrons or holes, as shown in Figure 2.3. While n-type dopants donate electrons to the lowest unoccupied molecular orbital (LUMO) states, the p-type dopants extract electrons from the highest occupied molecular orbital (HOMO) states, hence creates holes [8]. In other words, the Fermi level E_F of the polymer will shift towards the LUMO (or HOMO) level when there is n-type (or p-type) doping. This shift can be measured by spectroscopy techniques such as Ultraviolet Photoelectron Spectroscopy (UPS) at room temperature (RT) [9], although a bit limited by the penetration depth of the incoming electrons.

When doping occurs, the formation of a new quasiparticle named bipolaron occurs. For instance, if p-type dopant is introduced to a degenerated molecule (Figure 2.4), electrons are removed from double bonds, if we focus on just one event, while one carbon is left positively charged, the adjacent has an unpaired electron. This unpaired electron shifts down the polymer chain until reaching another unpaired

2.1. Organic Semiconductors

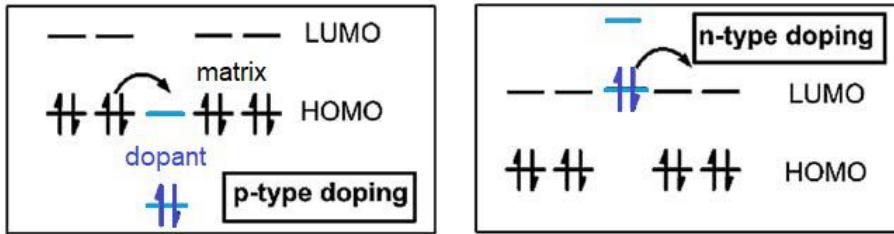


FIGURE 2.3: Schemes for p-type (left) and n-type (right) doping processes. Extracted from reference [8].

electron from a separated doping event, together they form a double bond, leaving aside two charged polarons: a bipolaron, which unlike a single polaron, it is stable [10].

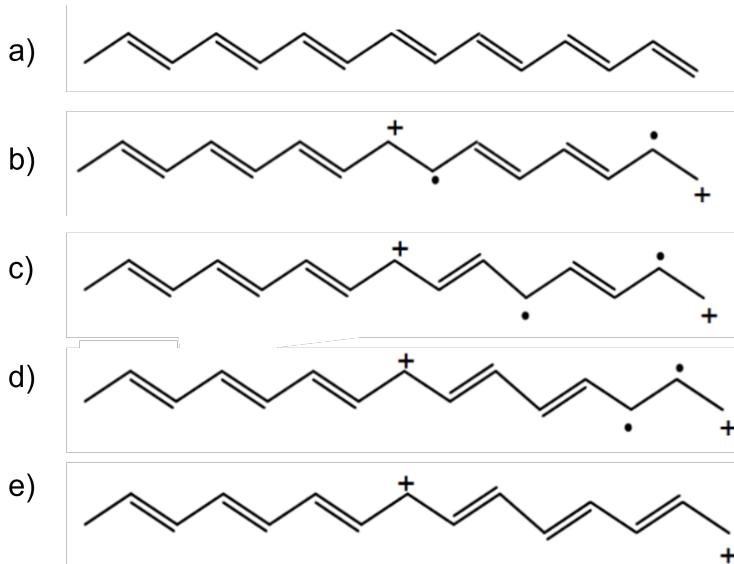


FIGURE 2.4: a) Structure of degenerated molecule. b) Electron extraction via doping. c) Unpaired electron shifting down. d) Unpaired electrons meet. e) Formation of double bond and two charged polarons: bipolaron.

The energy levels generated upon the formation of polarons and bipolarons are represented in Figure 2.5. As the doping increases, the amount of bipolaron states increases as well, the overlapping of them led to the formation of bipolaron bands, and the energy difference between the two in-gap states i and i^* decreases. In addition to the new electronic states, the doped polymer will also exhibit distinct optical transitions that could be disclosed by UV-Vis-NIR spectroscopy, although not clearly defined quantitatively [6].

The use of small molecules is commonly reported as dopants for organic materials. Some strong acceptor (or electron-deficient) molecules that are widely used are 2,3,5,6-tetrafluoro-7,7,8,8-tetracyanoquinodimethane (F_4TCNQ) or 1,3,4,5,7,8 hexafluoro-

2. BACKGROUND

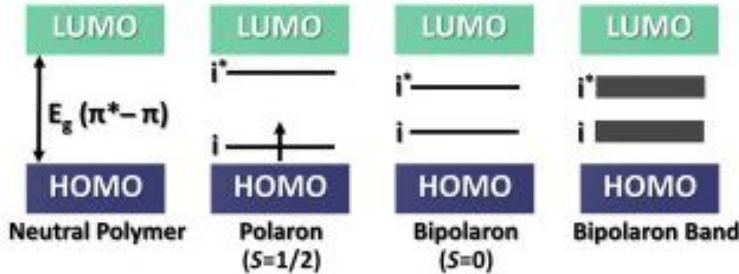


FIGURE 2.5: Potential energy change of aromatic and quinoid in the non-degenerate ground-state neutral conjugated polymer, and formation of polaron, bipolaron, and bipolaron bands upon doping. Extracted from reference [6].

7,7,8,8-tetracyanonaaphthoquinodimethane (F_6TCNQ), which also take part in this work. The latter exhibits a higher electronic affinity (-5.3 eV) or deep HOMO than F_4TCNQ (-5.2eV) meaning that it can abstract electrons more efficiently, specially from polymers with low ionization potential (less than 5eV) or shallow HOMO [11].

Among the different methods to molecular doped a conjugated polymer, Jacobs et al. compared a solution-mixed and solution sequential doping of P3HT (a thiophene-based polymer) doped with F_4TCNQ , both straightforward and easy methods for doping. Their work demonstrated that solution-mixed films are considerably rougher than solution-sequential films, affecting negatively to its conductivity [12]. The fact that solution-sequential doped films allows better homogeneity, make it also more compatible with microstructuring processes such as photolithography. At the expense of having less control over the doping levels compared to solution-mixed films [13].

2.2 Organic Mixed Ionic/Electronic Conductors (OMIECs)

Organic Mixed Ionic/Electronic conductors are organic semiconductors that allow the conduction of electrons (or holes) and ions, the latter set them apart from other organic semiconductors. Commonly structured with polar side chains, they have been identified as a promising class of materials for the field of bioelectronics [14]. Initially investigated for batteries and super capacitors [15] [16], where the induction of charges in a semiconducting polymer was the main objective. OMIECs have rapidly grown to include other applications, among them, our focus: OECTs.

Paulsen et al. classified OMIECs into six different categories according to whether they *intrinsically contain ionic charge*” (I, III, V) or not (II, IV, VI), the latter “*contain polar moieties that can solvate ions*”. Another distinction among the categories is whether the conjugated system comprises a single material (homogeneous, type V and VI) or two-component, more complex systems or block co-polymers materials (heterogeneous, type I, II, III, IV)[17], an schematic representation is shown in Figure 2.6.

2.2. Organic Mixed Ionic/Electronic Conductors (OMIECs)

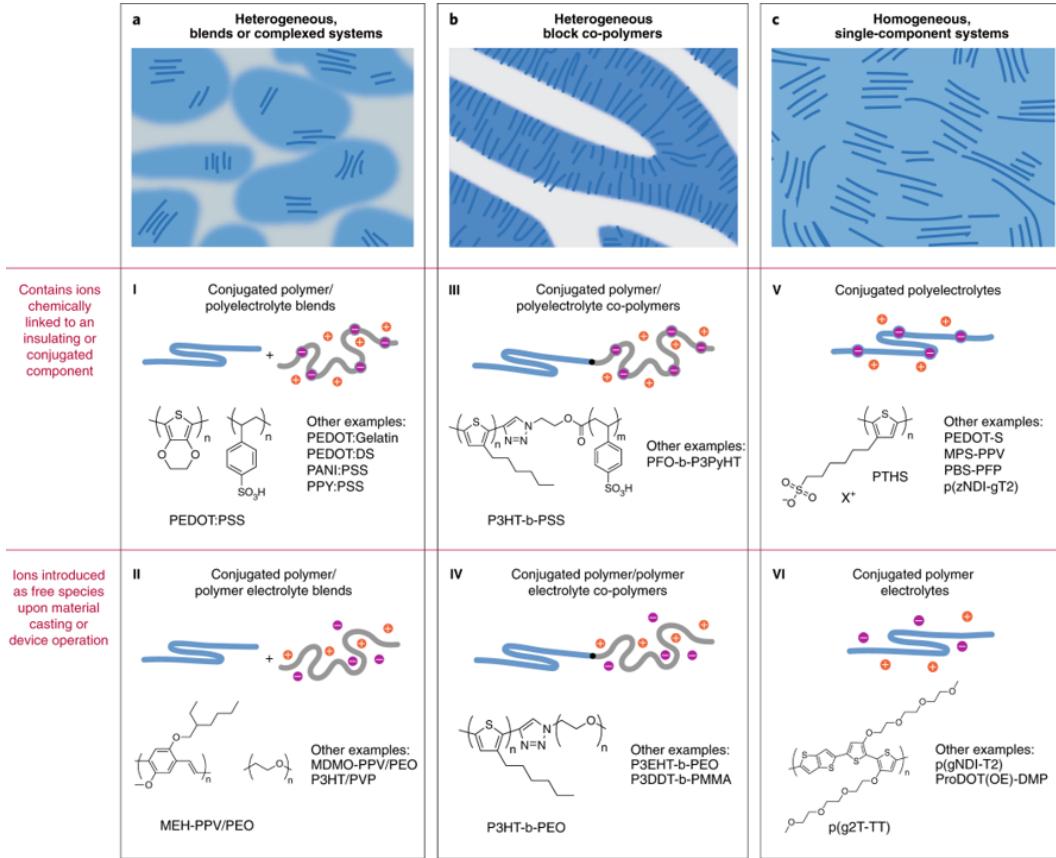


FIGURE 2.6: OMIECs classes. a) Heterogeneous blends of an conducting conjugated polymer with (I) a polyelectrolyte or (II) an ion-solvating polymer electrolyte. b) Heterogeneous block copolymers of a conducting conjugated polymer with (III) a polyelectrolyte or (IV) an ion-solvating polymer electrolyte. c) Fully conjugated (V) polyelectrolytes and (VI) ion-solvating polymer electrolytes. Extracted from reference [17].

2.2.1 Processes in OMIECs

2.2.1.1 Ionic-electronic interactions

The presence of electronic charge in OMIECs requires also the presence of excess ionic charge, so charge in the system remain balance. In the case of types II, IV and VI OMIECs, the so-called stabilizing electrochemical doping, is achieved by the presence of mobile ions, that act as free charges, the remaining type of OMIECs, on the other hand, have this stabilizing charges fixed (chemically bounded), they are inherently doped.

The amount of coupling between electronic charge and excess ionic charge in OMIECs can be modulated with an applied bias when coupled through an electrolyte [17]. This is the basic principle of OECTs, and we will be further discuss in section 2.3.

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2.2.1.2 Electronic transport

Electronic charge transport mechanisms present in OMIECs are thermally-activated hopping and band-like transport, as represented in Figure 2.7, which are not different from other conjugated polymers. The former as its name suggests, it is driven by thermal energy and it is limited by the degree of structural disorder, where the weakly-bond electrons in delocalized π -orbital move along an adjacent π -orbital within the length of the conjugated polymer, or even between molecules where there is sufficient π - π overlapping, if this mechanism is predominant is due to a low electronic charge carrier density and the density of accessible hopping states is low, hence, low mobility and electrical conductivity [17].

The band-like transport, on the other hand, occurs normally in OMIECs with increasing doping levels, where the activation energy of charge hopping decreases and carrier mobility increases, leading to diffuse band-like charge transport and it is shown within the polymer-stacking [18].

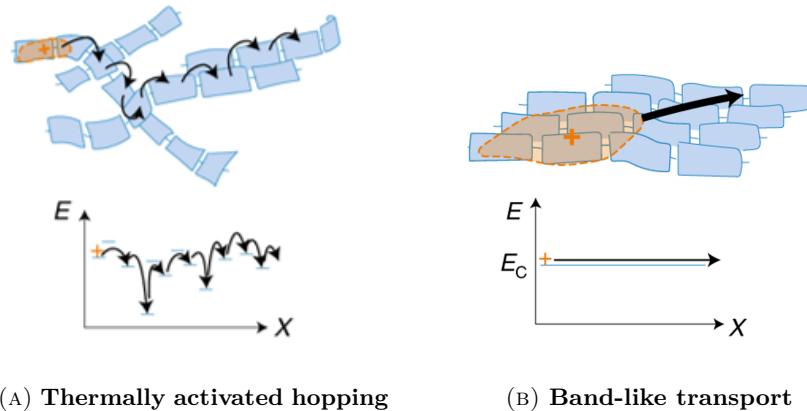


FIGURE 2.7: Schematic representation of electronic charge transport mechanisms: A) Thermally-activated hopping transport and B) band-like transport, where the electronic charge carrier is relatively localized and delocalized, respectively. Extracted from reference [17].

2.2.1.3 Ionic transport

Although transport of charged anions and cations can be seen analogous to electrons and holes, they remain more complex, since they can be “*multi-valent, and form pairs and larger clusters; moreover, they are sensitive to solvent and solvation*” [17].

Ion transport is unipolar for dry OMIECs of type I, III and V, since they are fixed on a polyelectrolyte, which is not the case for type II, IV and VI, both anions and cations are mobile. OMIECs in contact with an electrolyte swells, which allows the penetration of excess ions from the electrolyte, therefore both mobile anions and cations would contribute to ion transport to ensure electroneutrality.

Then two types of ionic charge transport can be differentiated: ion hopping and vehicular solvated-ion transport. In dry or minimally hydrated OMIECs only

ion hopping assisted by segmental motion of the OMIEC side chains or backbone, occurs, as shown in Figure 2.8A; whereas, when OMIEC is in contact with solvent or liquid electrolyte, both mechanisms are present but predominantly solvated ion vehicle transport. For instance, in water-based electrolytes, proton diffusion occurs via the Grotthuss mechanism, where protons in the water molecules diffuses within the neighboring molecules, this implies a transfer of an ionic effect through the hydrogen-bonded network, as shown in Figure 2.8B.

Finally, due to the existing ionic-electronic coupling, both ionic and electronic transport in OECTs and other OMIEC-electrolyte based applications, are not independent but are rather complex and needs to consider side effects, such as hydration and electrolyte swelling [17], which will be further discuss in section 2.3.

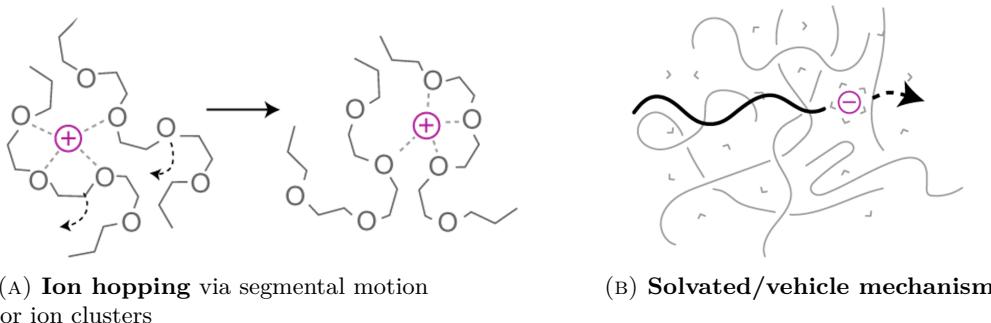


FIGURE 2.8: Schematic representation of ionic charge transport mechanisms: A) Ion hopping via segmented motion and B) vehicular solvated-ion transport. Extracted from reference [17].

2.3 Organic Electrochemical Transistors (OECTs)

Organic Electrochemical Transistors (OECTs) consists of metallic source, drain and gate electrodes, an organic conjugated polymer channel (specifically an OMIEC as described in previous sections) and an electrolyte that couples channel and gate, as represented in Figure 2.10A). Devices that have received increasingly attention due to their mechanically compliance, biocompatibility, and are sensitive to biochemical modules [19].

2.3.1 Device Physics

Although the structure of OECTs are different from conventional metal-oxide-semiconductor field-effect transistors (MOSFETs). The basic understanding of the latter can give us a clear idea of how OECTs operate. Unlike MOSFETs, OECTs are coupled with an electrolyte rather than an insulator, as seen in Figure 2.9. So, when applying a gate voltage, instead of polarizing the dipoles in the insulator, creating a field that causes accumulation of carriers at the interface of the semiconductor/insulator as it happens in a MOSFET, in an OECT, the gate voltage

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drives ions to penetrate the bulk of the channel, therefore accumulation of carriers will happen throughout the whole volume of the OMIEC film. This explains the large gate-channel capacitance in these devices compared to MOSFETs, and why drain-source current takes into account a volumetric capacitance [20].

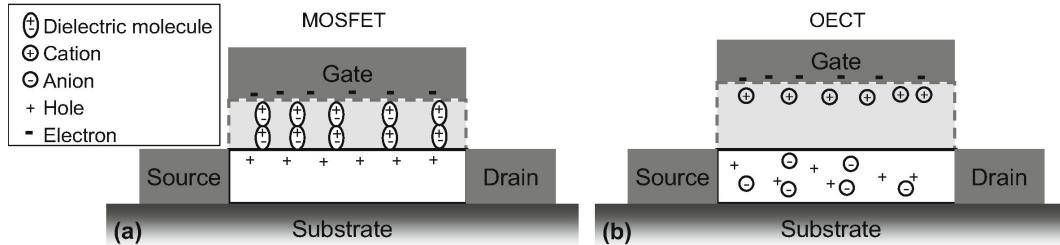


FIGURE 2.9: Comparison of p-type A) MOSFET and B) OECT. Where the light-gray region represents an insulator and a electrolyte, respectively. Extracted from reference [20]

Bernards and Malliaras implemented a model based on a p-type depletion-mode OECT (based on PEDOT:PSS since it is widely fabricated and investigated, further discussion in the following section). The model divides the behavior of the OECT into an electronic (source-channel-gate structure) and an ionic circuit (gate-electrolyte-channel structure). The electronic circuit is treated as a *variable* resistor therefore it is modeled using Ohm's law, its variability lies on the fact that upon the application of a positive gate voltage, de-doping of the semiconductor occurs, analogous to the compensation doping of Silicon, cations from the electrolyte penetrate the polymer, compensating one acceptor. Meanwhile, the ionic circuit consists of a resistor that represents the flow of ions in the electrolyte, in series with a capacitor, representing the storage of ions in the channel as shown in Figure 2.10B) [21][22].

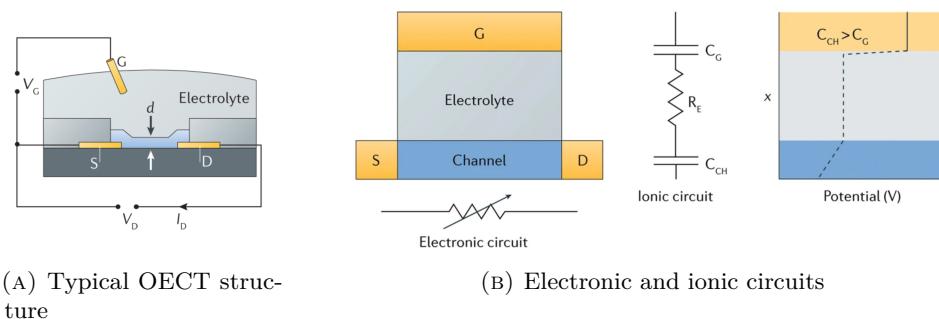


FIGURE 2.10: A) Typical structure of an organic electrochemical transistor (OECT). B) (Left) Electronic circuit modelled as a resistor with a variable resistance. (Right) Ionic circuit consisting of channel (C_{CH}) and gate (C_G) capacitors, coupled with a resistor corresponding to the electrolyte (R_E). Extracted from reference [21].

2.3.2 Operation Modes

Analogous to conventional MOSFETs, depending on whether the device needs a gate potential to turn it ON, it will describe two operation modes: depletion and enhancement (the latter commonly named as accumulation for OECTs). These operation modes have a strict relationship to the channel material.

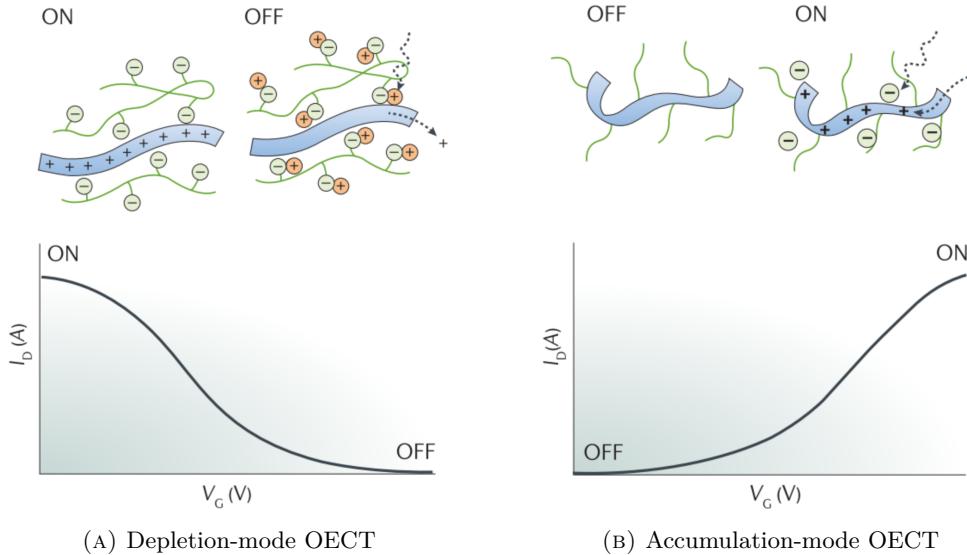


FIGURE 2.11: (A) Transfer curve showing depletion-mode operation of a p-type OECT with a conducting polymer channel. (B) Transfer curve showing accumulation-mode operation of a p-type OECT with a semiconducting polymer channel. Images extracted from reference [21].

As seen in Figure 2.11A) and B), the polymer can be conductive or semiconducting. In the first scenario, the OMIEC already possesses anions which have had induced charges within its backbone, therefore it needs the injection of cations to counteract this effect, hence turn off the device. The opposite scenario is presented for a semiconducting polymer channel, a zero-gate biased OECT would have no charges in its backbone hence the device will be in off state, it will need the application of a gate voltage to drive anions into the polymer and induce charges.

2.3.2.1 Standard material for depletion-mode OECTs

Poly(3,4-ethylenedioxythiophene) poly(styrene-sulfonate) (PEDOT:PSS) is a “*degenerately doped*” [22] or conductive polymer that is widely used in multiple applications in organic electronics. Classified as type I OMIEC, as seen in Figure 2.6, it is a blend between a conjugated polymer (PEDOT) and a polyelectrolyte (PSS), the latter possesses chemically linked ions and serves as a polymeric acid template to allow dispersable suspensions [17].

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Due to its commercial availability, operational stability, and relatively high performance, PEDOT:PSS became a standard material for p-type OECTs. Its main drawback lays in its depletion-mode operation, since, as explained in the previous section, requires power to turn off.

With the aim of minimizing power consumption, there is a special interest to fabricate accumulation-mode devices with high performance [1][13][23][24]. This type of devices have the advantage of dissipating less static power when the device is not operated, due to low OFF current [14].

2.3.2.2 Prospective materials for accumulation-mode OECTs

PEDOT:PSS was not descarted as a possible accumulation-mode OECT, Keene et al. used a series of amines to de-dope PEDOT:PSS and obtain OECTs with negative turn-on voltages [24]. However, synthetically modifying PEDOT:PSS in a control manner remains complicated. Parallel to these efforts, the design of new semiconducting polymers with the aim of not only having acummulation-mode OECTs but enhancing performance is also studied. Nielsen et al. reported a series of semiconducting polymers with Triethylene glycol (TEG) side chains with good performance, among the five thiophene- and benzodithiophene-based polymers, they found out that the backbone consisted of 2,2'-bithiophene polymerized with other thiophene molecule (g2T-T), as seen in Figure 2.12, shown the highest performance [1].

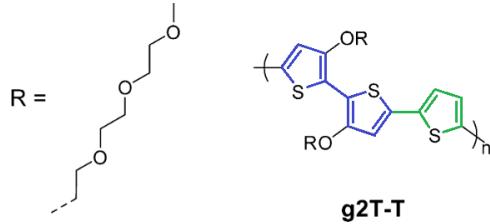


FIGURE 2.12: Chemical structure of polymer with backbone g2T-T, R representing the side chain. Extracted from reference [1]

Moser et al. took the same backbone and studied the impact of the length of the ethylene glycol (EG) side chain on the performance of OECTs. They reported that reducing chain length maximized both the capacitance and mobility; nonetheless, it was unfavorable for ion-polymer interaction. Finally, they suggested an optimum-side-chain length of 3 monomers (over 2, 4 and 6 monomers), demonstrating that an OECT with 3-(2-(2-(2-methoxyethoxy)ethoxy)ethoxy)thiophene (p(g3T2-T)) has a turn-on voltage close to zero, and better performance compared to other thiophene-based species (as seen in Figure 2.13), even better than PEDOT:PSS OECTs [25].

The structural tuning of this polymer was though to have a backbone that warrants reversibility during electrochemical redox reactions and good electronic transport, meanwhile having side chains that enable its stability in aqueous electrolytes and efficient transport of ionic and electronic charge carrier [26].

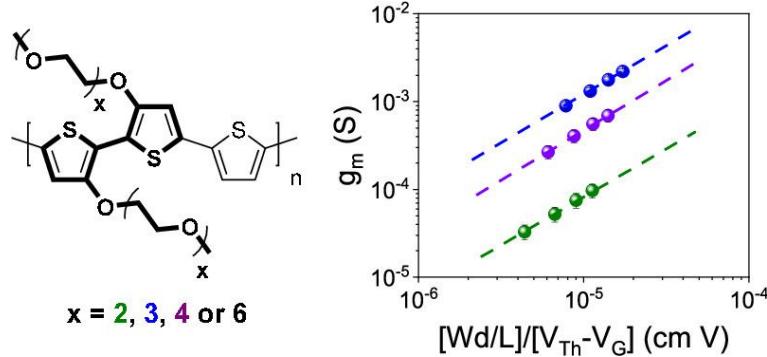


FIGURE 2.13: (Left) Chemical structure of the repeat units for p(gxT2-T). (Right) Transconductance vs channel geometry and operating parameters of p(gxT2-T) for $x = 2, 3$ and 4 . Extracted from reference [25].

Under the classification shown in Figure 2.6, p(g3T2-T) can be identified as a type VI OMIEC that comprises a conjugated polymer with ions introduced as **free species** whereas PEDOT:PSS' ions are **chemically linked** to the polyelectrode (PSS). This structural characteristic make p(g3T2-T) display larger magnitudes of ionic-electronic coupling and volumetric capacitance than biphasic OMIECs (such as PEDOT:PSS) [17] and at the same time will be important for understanding the challenges on having a stable OEET, further discussion in Section 2.3.4.

2.3.3 Important Figures of Merit

2.3.3.1 Transconductance

Considered as the most important parameter to measure any transistor's amplification capability, calculated as the first-order derivative of the output current (drain-source current) with respect to the input voltage (gate-source voltage): $g_m = \partial I_{DS} / \partial V_{GS}$. Bernards and Malliaras calculated this parameter from their implementation of a mathematical model for depletion-mode p-type OEETs [22], as previously seen in Section 2.3.1. Including also n-type, the following equation expresses the transconductance:

$$g_m = \frac{Wd}{L} \mu C * |(V_{th} - V_G)|, \quad (2.1)$$

where W , L and d are the channel width, length and thickness, respectively, $\mu C*$ product, the gate voltage (V_G) and threshold voltage (V_{th}), which will be discussed in the next subsections.

Commonly, the maximum value of transconductance is reported ($g_{m,max}$), which falls into the saturation regime. Inal et al. reported the maximum transconductance of various channel materials at different device geometry parameters using a Ag/AgCl pellet as reference/gate electrode and 0.1M NaCl as electrolyte, showing OEETs with polymerized-g2T-backbones, the best performances, within the range of 1 to 30 mS, depending on the geometry (Figure 2.14A) [23].

2. BACKGROUND

2.3.3.2 μC^* product

The most important parameter for benchmarking OECT channel materials, that represents both ionic and electronic transport properties [23]. It is the product of two important parameters, μ , the electronic mobility, and C^* , the volumetric capacitance, the latter encloses the ion penetration, transport, and storage ability of the OMIEC film.

Along with the transconductance, Inal et al. extracted the values of μC^* from the calculating the linear slope of the maximum transconductance and channel geometry (Figure 2.14A), following Equation 2.1. And correlated with the independent calculation of both parameters, showing again polymerized-g2T-backbones with the highest value of $261 \pm 29 F cm^{-1} V^{-1} s^{-1}$, and among the materials that have closest 1:1 relation for both methods of calculation μC^* product (Figure 2.14B) [23].

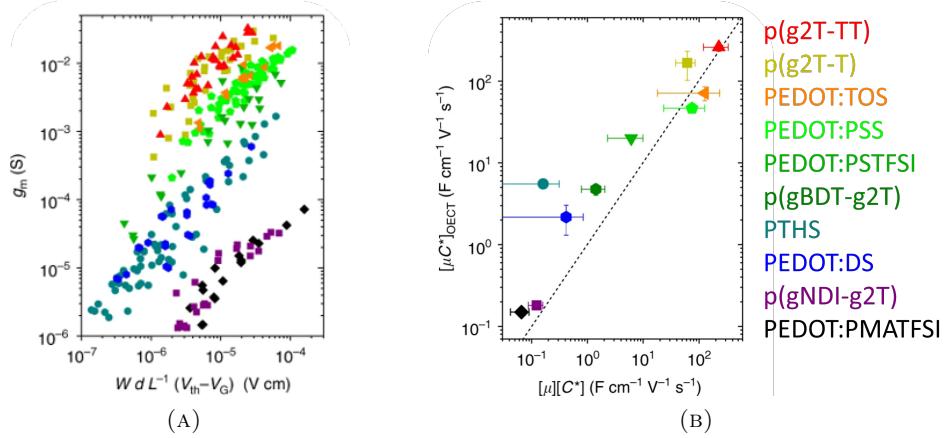


FIGURE 2.14: A) Plot of μC^* product calculated by the linear slope between transconductance and channel geometry. B) Calculated slope from A) in function of the product of independent determination of μ and C^* , dotted line represents the 1:1 relation between both methods of calculation. Extracted from reference [23].

The independent calculation of μ is yet a bit tricky due to the presence of ionic species, normally calculated in transient regimes, taking advantage of ions slower mobility, which is out of the scope of this work. Among the method to calculate C^* , performing Electrochemical Impedance Spectroscopy (EIS) is a straightforward one, and by using Equation 2.2 to calculate the capacitance, at low frequency ranges where the capacitance should describe a plateau, since the modulation of AC is slow enough to fully populate the OMIEC with ions. Finally, divide the calculated capacitance by the film volume to obtain C^* [27].

$$C = \frac{1}{2\pi \cdot f \cdot |Z^{img}|}, \quad (2.2)$$

where $Z^{img}(\Omega)$ is the imaginary part of the impedance and f is the frequency (Hz).

The calculation of this capacitance is under a fixed-gate biased condition, it is important to remember that the modulation of the degree of electrochemical doping in OECTs with an applied biased, will manifest a potential-dependent capacitance (C)[23].

2.3.3.3 Threshold voltage

From the steady-state characteristics of an OECT, specifically the transfer characteristics (I_{DS} vs V_{GS}), one can calculate not only the transconductance but also the ON/OFF ratio, and the threshold voltage (V_{th}). The latter can be determined by plotting the square root of the I_{DS} as a function of V_{GS} and extrapolate the linear portion of the slope, where the intersection with the x-axis will give this parameter [27].

In MOSFETs, it represents the voltage that fills trap states in the semiconductors. In OECTs, it represents the “film’s readiness for ion penetration” [27]. Controlling and/or shifting this parameter is wished, specially to integrate transistors and meet any circuit requirement and control operation, noise margins and power consumption.

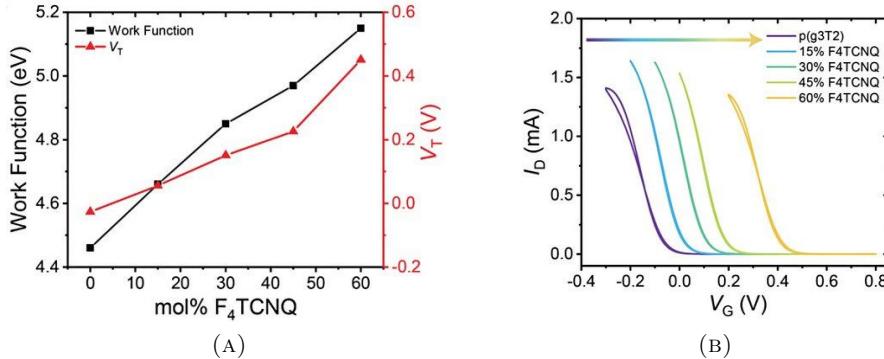


FIGURE 2.15: Controlling OECT threshold voltage by chemical doping of p(g3T2-T) gate electrode with F_4TCNQ . A) Plot of threshold voltage and gate work function for doped gates of different dopant concentrations. B) Transfer curves of p(g3T2-T) channel OECT with p(g3T2) gates of various F_4TCNQ dopant concentrations. . Extracted from reference [2].

The chemically de-doping of PEDOT:PSS by Keene et al. described in previous sections, is also an approach to shift the threshold voltage, until reaching negative values, characteristic for accumulation-mode OECTs [24]. Tan et al., on the other hand, explored a different approach, rather than modifying the doping level of the channel, they tuned the doping level of the gate to shift the threshold voltage. They used p(g3T2-T) and obtained a 400mV change with 60% mol ratio of 2,3,5,6-Tetrafluoro-7,7,8,8-tetracyanoquinodimethane (F4TCNQ) dopant, as seen in Figure 2.15. The advantage over this approach is i) protecting the material from oxidation with air, since the Fermi level is brought towards the highest occupied molecular orbital (HOMO), and ii) no interference with the channel which helps to leave the transconductance unaffected [2].

2. BACKGROUND

2.3.4 Side Reactions

2.3.4.1 Water uptake and swelling

Water-based electrolytes are widely used, and even in the pursue of solid-state OECTs, precursors have a certain degree of water content [3][28], hence, side reactions of OMIECs upon water contact is inevitable and therefore, studies to understand the impact are necessary. Water uptake in OMIECs cause their increase of mass (swelling) and change in morphology. Water uptake happens since OMIECs will need to compensate their intrinsic doping. Therefore, “*the effect of doping-induced hydration on the OMIEC morphology must be taken into account when designing OECTs*” [29].

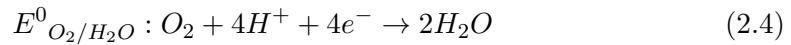
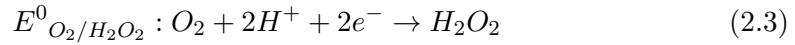
Savva et al. study the influence of water on the performance of PEDOT:PSS OECT, the water uptake led to 10-13% mass increase under non biased conditions. As the concentration of water decrease (NaCl_{aq} 10 mM, 100mM, 1M, and 6M) ionic charging got faster; however, the fastest response time is not achieved by the highest salt concentration but rather NaCl_{aq} 1 M. This due to attractive forces between counter ions which hinders the drift of anions, hence delays the ion injection from the electrolyte [30].

In another study, Savva et al. shown that certain level of hydration is necessary for facile ionic transport, but can negatively impact electronic charge transport, in glycol-based side chains [29], commonly used in enhancement-mode OECTs and where ionic transport is already enhancement by side-chain engineering [26].

2.3.4.2 Oxygen Reduction Reaction (ORR)

Oxygen Reduction Reaction is a common undesirable side reaction at environmental conditions, specially among devices fabricated with polymers with low ionization potential (IPs). The utilization of OMIECs with low IP are frequent in accumulation-mode OECTs.

The ORR is a non-capacitive faradaic reaction between the OMIEC and molecular oxygen, where electron-transfer occur. The reaction yields either hydrogen peroxide (H_2O_2) or water (H_2O), as described in 2.3 and 2.4, or oxidation (p-doping) of the OMIEC that acts as the catalyst [14].



As seen in Figure 2.16, η_1 and η_2 represent free energy difference between reactants and the reaction products H_2O_2 and H_2O , respectively. If this free energy is negative, the reaction is endergonic (the opposite is exergonic), and needs energy to be driven. PEDOT:PSS and p(g3T2-T) exhibit exergonic reactions. For instance, a polymer with higher IP ($> 4.9\text{eV}$) would be needed to prevent oxygen peroxide to form.

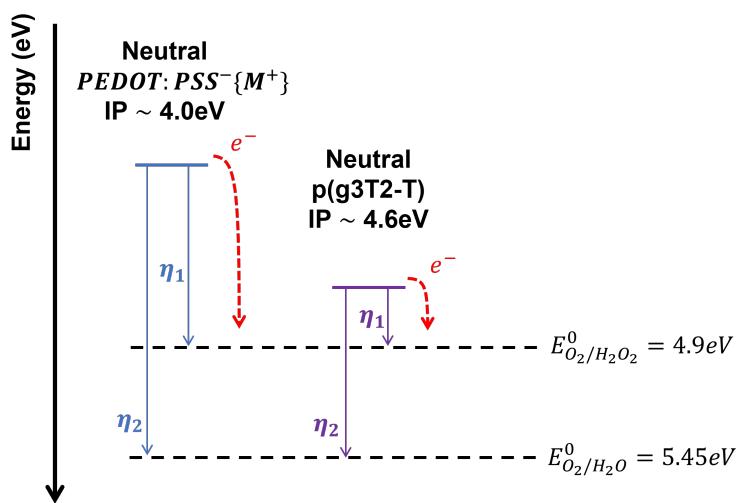


FIGURE 2.16: Simplified mechanism of two- and four-electron Oxygen Reduction Reactions with neutral states of OMIECs: p(g3T2-T) and PEDOT:PSS. The free energy difference between reactant and the reaction products is represented by η . Image adapted from reference [14], where values are defined from cyclic voltammetry measurements using Ag/AgCl as reference electrode and NaCl as electrolyte.

Chapter 3

Experimental Methods

3.1 Materials

All reactives were purchased from commercial suppliers and no further chemical modification or purification was done unless stated before.

- Chromium etchant: Standard, Sigma Aldrich
- Gold etchant: Standard, Sigma Aldrich HHPAA (2-Hydroxy-4'-(2-hydroxyethoxy)-2-methylpropiophenone), 98%, Sigma Aldrich
- Developer: AZ 726 MIF Developer, Merck performance Materials GmbH
- EG: Ethylene glycol, ≥ 95%, Sigma Aldrich
- [EMIM][EtSO₄]: (1-Ethyl-3-methylimidazolium ethyl sulfate), ≥95%, Sigma Aldrich
- MBBAm (N,N'-Methylenebisacrylamide), 99%, Sigma Aldrich
- NIPAm: (N-Isopropylacrylamide), 97%, Alfa Aesar
- Sacrificial Layer 1: Sacrificial Layer 1 (SL1), Orthogonal Inc.
- Orthogonal Photoresist for undoped species: NLOF 2020, commercial negative-tone photoresist, Microchemical.
- Orthogonal Photoresist for doped species: OSCoR 4020 Photoresist, Orthogonal Inc.
- Developer for SL1: Developer HF 7300, Orthogonal Inc.
- Orthogonal Developer for OSCor 4020: Orthogonal Developer 103a, Orthogonal Inc.
- Photoresist: AZ 1518 Photoresist, Merck Performance Materials GmbH & Microchemical GmbH

3. EXPERIMENTAL METHODS

- Orthogonal Stripper: Orthogonal Stripper 900, Orthogonal Inc.
- p(g3T2-T): 3-(2-(2-methoxyethoxy)ethoxy)ethoxythiophene
- Dopants: 1,3,4,5,7,8-hexafluorotetracyanonaphthoquinodimethane ($F_6\text{TCNNQ}$) and 1,3,4,5-tetrafluorotetracyanonaphthoquinodimethane ($F_4\text{TCNQ}$)
- Silane A174 (3-(Trimethoxysilyl)propyl methacrylate), TCI

3.2 Equipment

- Baking: All baking steps were carried out on a Stuart SD160 digital hotplate (Stuart Equipment, UK).
- Electrical characterisation (ambient): Device characterisation under ambient conditions was performed on a Everbeing C-6 Probe Station (Everbeing Int'l Corp., Taiwan), connected to a Keithley 4200-SCS Semiconductor Characterisation System (Keithley Instruments, USA).
- Electrical characterisation (glovebox): Device characterisation was performed in a nitrogen-filled glovebox. Probing needles were connected to two Keithley 236 Source Measure Units (Keithley Instruments, USA).
- Cyclic voltammetry and Impedance measurements: Cyclic voltammetry and Impedance measurements were carried out by using a Metrohm Autolab PGSTAT302N potentiostat/galvanostat (Metrohm AG, Switzerland) in the range of 10^{-1} to 10^5 Hz.
- Micrographs: Micrographs were taken on a Nikon Eclipse LV100ND microscope, equipped with a DS-Fi2 camera (Nikon, Japan).
- Photolithography: Photolithography was carried out on a SÜSS Microtec MJB4 maskaligner system (SÜSS Microtec AG, Germany).
- Photomasks: Photomasks were custom made by Compugraphics Jena in a 4-inch format (soda-lime glass covered with chromium) and held several mask designs (Compugraphics Jena GmbH, Germany).
- Plasma cleaning: O₂-plasma cleaning was performed by using a Harrick PDC-002 plasma cleaner (Harrick Plasma, USA), connected to a Leybold Heraeus Combitron CM 330 Vacuum Gauge Controller (Leybold GmbH, Germany).
- Plasma etching: O₂-plasma etching was performed by using a Diener electronic ATTO plasma cleaner (Diener electronic GmbH & Co. KG, Germany).
- Profilometry: Profilometry was performed on a Veeco Dektak 150 surface profiler (Veeco Instruments Inc., USA).

- Absorption measurements: Absorption measurements were performed with UV-Visible-NearInfraRed Spectroscopy on a SolidSpec-3700 UV-Vis-NIR spectrometer from Shimadzu.
- Workfunction measurements: Workfunction was measured using Ultraviolet Photoelectron Spectroscopy (UPS) by using a PHOIBOS 100 from Specs, a Helium plasma discharge lamp (UVS10/35, Specs) for UV excitation line of 21.22 eV.
- Spincoating: Samples were coated with a SAWATEC SM-180-BT spincoater (SAWATEC AG, Switzerland).

3.3 Software

- Data processing: All data was processed by customised scripts written in the Python programming language. Mathematical computations (e.g. fits, integration) were carried out by employing the Pandas, NumPy, and SciPy libraries. Visualisations were performed using the Matplotlib library. All is seen in GitHub Repository.
- Electrical characterisation: Electrical characterisations were performed by controlling SMUs through the in-house developed SweepMe! software.
- Profilometry: Profilometry was performed by using the Dektak software (Veeco Instruments Inc., USA).
- Cyclic voltammetry and Impedance measurements: Impedance measurements were performed by using the NOVA software (Metrohm AG, Switzerland).

3.4 Experimental Procedures

All fabrication steps were performed under standard cleanroom conditions.

3.4.1 Preparation of films

Dynamic spin-coating of p(g3T2-T). Unlike reference [2], in which drop cast is used to deposit undoped and doped p(g3T2-T), in order to perform photolithography and being able to do a miniaturization process, homogeneous films are needed. Dynamic spin-coating was previously established by BioSens group members at IAPP, due to the high volatility of p(g3T2-T) solvent: chloroform. Substrates were cleaned using subsequent steps of ultrasonic bath with acetone for 15 minutes, IPA rinsing, N₂ drying and O₂-plasma cleaning. Then 70 μ L of 10 mg/mL of p(g3T2-T) mixed at 60°C for 20 minutes, was dynamic-spin-coated at 3000 RPM for 60s, to yield approximately 70nm-thick films.

3. EXPERIMENTAL METHODS

Dynamic spin-coating of F₄TCNQ dopant. Different doping levels of p(g3T2-T) was achieved by dynamic spin-coating dopants at different concentrations 5, 10 and 20 mg/mL, previously mixed in acetonitrile at 60°C for 20 minutes.

3.4.2 Doping characterization of films

UV-Vis-NIR Spectroscopy. After the preparation of the films, Transmittance (T) and reflectance (R) was measured using the UV-Vis-NIR Spectrometer in the range of 285 to 1600 nm, using a slit width of 5,0 nm. Then absorption (A) was calculated via

$$A = 1 - T - R, \quad (3.1)$$

normalized with respect to the incoming light.

Additional measurements were taken, after a couple of days of storage under ambient conditions.

Ultraviolet Photoelectron Spectroscopy. After the preparation of the films (air exposure) the energy of the highest occupied molecular orbital cutoff (E_{HOMO}) and the high binding energy cutoff E_{HBEC} was measuring using UPS. The pressure in the chamber during measurements was about $5 \cdot 10^{-9}$ mbar, while the base pressure is in the range of $10 \cdot 10^{-10}$ mbar. The workfunction (WF) and ionization energy (IE) are given by the following equations:

$$WF = h\nu - E_{HBEC}, \quad (3.2)$$

$$IE = h\nu - (E_{HBEC} - E_{HOMO}), \quad (3.3)$$

where $h\nu = 21.22\text{eV}$, the main He I excitation line.

Profilometer. The films were scratched 4 times to remove part of material, and measure the cavity deep to obtain the film thickness.

Van Der Paw

3.4.3 Fabrication of Organic Electrochemical Transistors

Doping OECT channel

Undoped p(g3T2-T) OECT

Doped p(g3T2-T) OECT

3.4.3.1 Stability on air of p(g3T2-T)

3.4.3.2 Reverse oxidation of undoped p(g3T2-T)

Applying a gate biased

Applying a baking step

3.4.3.3 Stability of solid OECTs using undoped p(g3T2-T)

OECTs with dropcast Solid-State Electrolyte

OECTs with photopatternable Solid-State Electrolyte

OECTs with inkjet-printed Solid-State Electrolyte

3.4.3.4 Solid OECTs using doped p(g3T2-T)

3.4.3.5 Threshold voltage shift

Chapter 4

Results and Discussion

4.1 Doping characterization

Upon increase of the dopant concentration that is deposited and bake on top of the p(g3T2-T) film, the reflection hue, as observed in Figure 4.1 shift towards more yellowish colors.

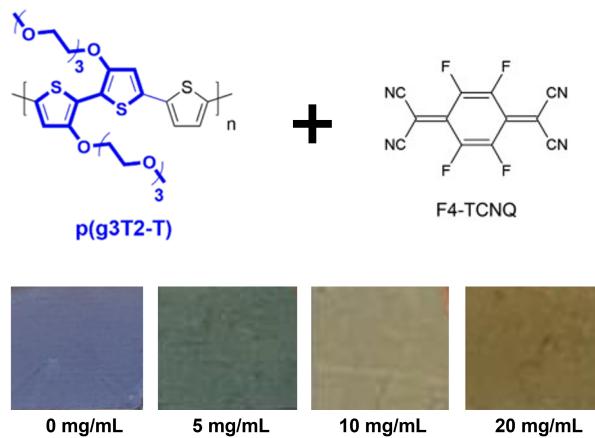


FIGURE 4.1: (Top) Chemical structure of p(g3T2-T) and F₄TCNQ dopant. (Bottom) Color change upon increasing dopant concentration from 0 to 20 mg/mL.

4.1.1 Absorbance and dopants diffusion

The visible color hue shift can be described quasi-quantitatively by the absorbance spectra of the samples, as seen in Figure 4.2, where the undoped-p(g3T2-T) shows a prominent absorption peak at 588 nm (red), which diminishes with doping, characteristic from oxidation, and it is more intense with higher doping levels. New absorption peaks are seen at around 860 nm, this is due to generation of polarons that lead to new optical transitions, as explained in section... from the Background chapter. Tan et al. described the apparition of new absorption peaks within 300

4. RESULTS AND DISCUSSION

to 600 nm, the higher energetic (lower wavelength) one is generated by unreacted neutral dopant species (TCNQ^0), while the second is due to the new dopant anions (TCNQ^-), the latter induce charges in our polymer [2]. It is shown also, that after some days of storage, the initially dominant peak of unreacted neutral dopants is now lower in intensity than the anions peak, suggesting that the dopants were still diffusing through the polymer along the days.

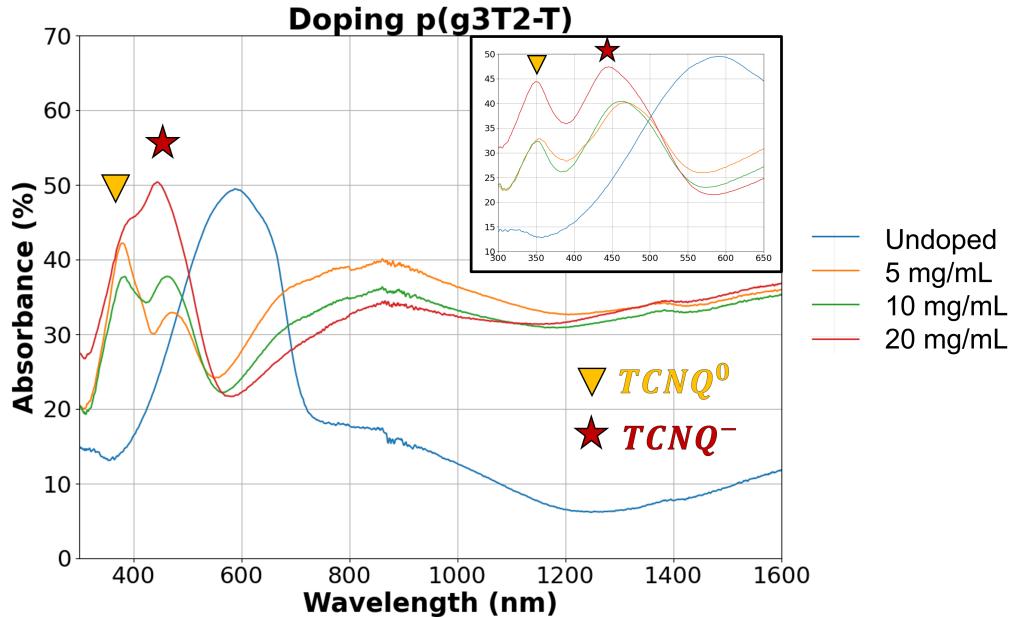


FIGURE 4.2: Spectra of undoped and doped-p(g3T2-T) at different doping levels, corresponding to samples on Figure 4.1. Inlet represents absorbance after two weeks of storage in ambient conditions.

Absorbance values are directly linked to the density of states of these new optical transitions [10]. From our spectra, the absorption value of the lowest doped-p(g3T2-T) (5 mg/mL) is higher (around 40%), which may be counter-intuitive. However, as doping concentration increases, hence more electrons are taken out of the polymer chain, a bipolaron is the energetically more favorable quasiparticle to generate. According to references [2][31], bipolaron formation is seen as a shift to lower energies in the broad absorbance the mid-IR region (wavenumber 1000-1600 cm^{-1}) so further analysis of hole bipolaron formation can be conducted with Fourier Transform InfraRed (FTIR) spectroscopy.

4.1.2 Workfunction

Although, some air exposure and contamination was not possible to be avoid during the preparation of the films, it is possible to distinguish an increase in the workfunction (shift of Fermi level) of the polymer. The Fermi level shift towards the HOMO level

4.2. Fabrication of Organic Electrochemical Transistors

is higher at higher dopant levels, as represented in Figure 4.3, characteristic of p-type doping. Some possible contamination did not allow to measure sample with 20 mg/mL of dopant.

TABLE 4.1: Workfunction calculation from UPS measurements.

	Undoped	5 mg/mL	10 mg/mL
E_{HBEC} [eV]	17.35	16.47	16.36
E_{HOMO} (vs E_F)	4.28	3.27	3.24
WF [eV]	3.87	4.28	4.86

It is important to consider that UPS is a surface sensitive measurement, the penetration depth of the ultraviolet range electrons are around 2 nm, which is significantly lower than our polymer thickness (around 70 nm). So this technique limits our understanding of the dopants diffusion through the whole volume of the polymer. Although, we had some insights of this topic from the previous section, further analysis could be done with X-Ray Photoelectron Spectroscopy that has a higher penetration depth (around 5 nm), also limited.

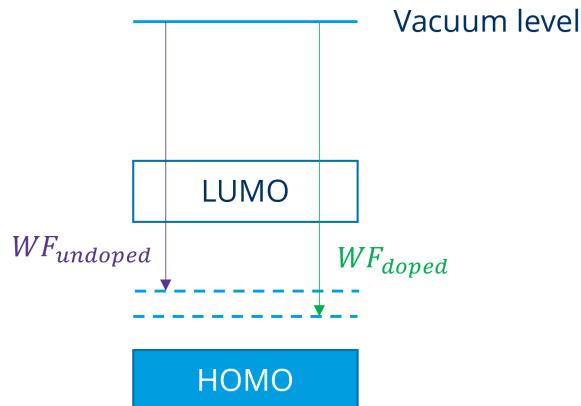


FIGURE 4.3: Representation of the Fermi level shift for doped-p(g3T2-t)

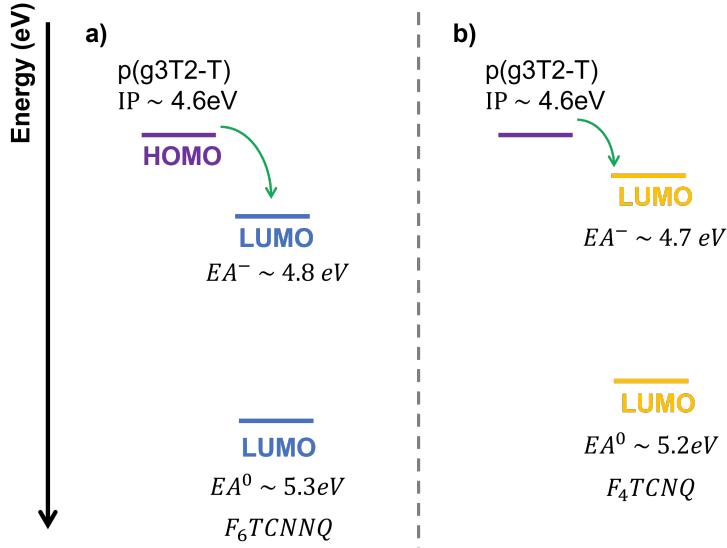


FIGURE 4.4: Energy diagram of p(g3T2-T) with a) F_6TCNNQ and b) F_4TCNNQ . Ionization potential of p(g3T2-T) extracted from reference [2] and electron affinities of the neutral species (EA^0) and of the anion (EA^-) of the dopants are extracted from cyclic voltammetry measurements in reference [11].

4.1.3 Thickness, sheet resistance and resistivity

4.2 Fabrication of Organic Electrochemical Transistors

4.2.1 Influence of doping OECT channel

4.2.2 Stability on air of p(g3T2-T)

4.2.3 Reverse oxidation of undoped-p(g3T2-T)

4.2.3.1 By Electrochemical Dedoping

4.2.3.2 By Heating

4.2.4 Solid-OECTs using undoped p(g3T2-T)

4.2.4.1 Dropcasted Solid-State Electrolyte

4.2.4.2 Photopatternable Solid-State Electrolyte

4.2.4.3 Inkjet-Printed Solid-State Electrolyte

4.2.5 Solid-OECTs using doped-p(g3T2-T)

Chapter 5

Conclusions and Outlook

Appendices

Appendix A

Absorption of p(g3T2-T) doped with F6TCNNQ

Appendix B

PEDOT:PSS vs doped p(g3T2-T) as channel

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