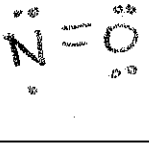
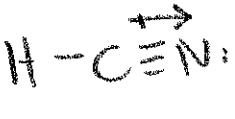
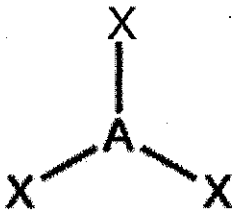
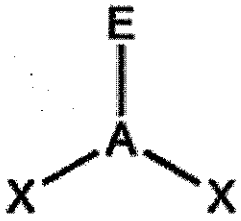
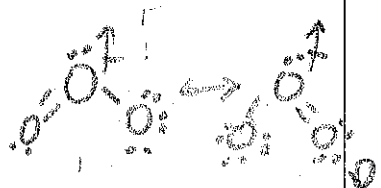
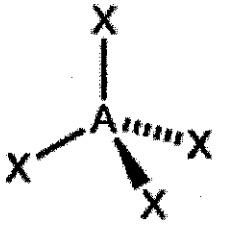
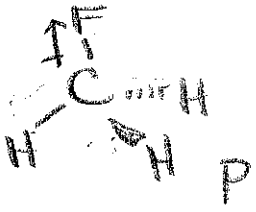
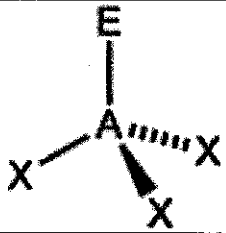
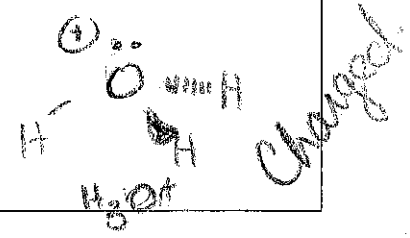
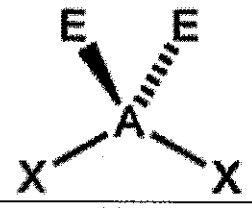
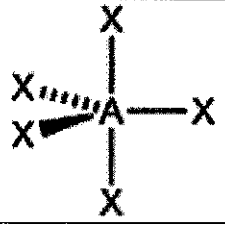
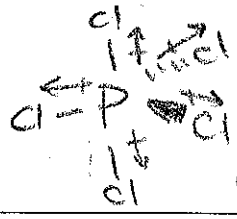
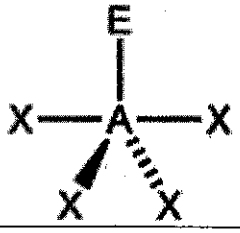
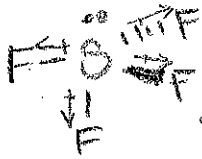
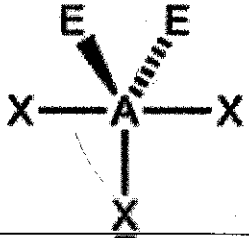
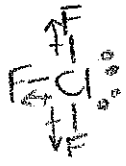
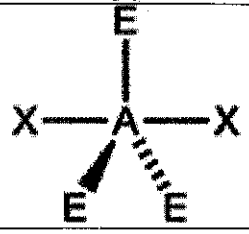
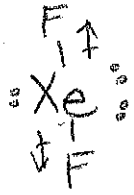
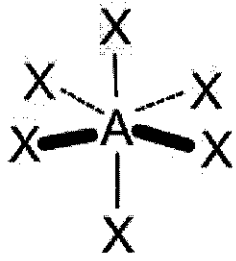
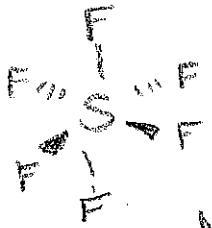
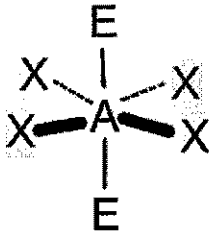
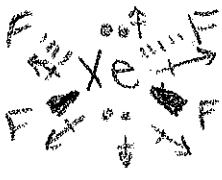
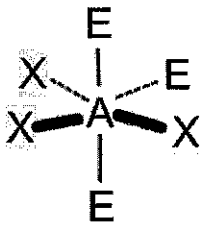
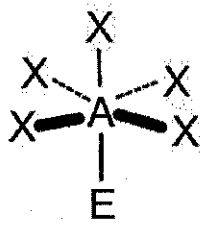
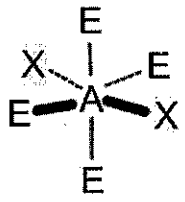
Note, this table utilized the AXE method. "A" is the central atom. "X" are the bonded atoms (peripheral atoms). "E" if the unbound/lone pairs of electrons.

Shape	# Bonding Pair(s)	#Lone electron pairs	Molecular Geometry	Example **Specify whether your example is polar of nonpolar
A — X	1	0	linear	 NP
X—A—X	2	0	linear	 P
	3	0	trigonal planar	 NP
	2	1	bent	 P
	4	0	tetrahedral	 P
	3	1	Trigonal pyramidal	 Charged!

	2	2	Bent	 P.
	5	0	trigonal bipyramidal	 NP
	4	1	seesaw	 P.
	3	2	T-shaped	 P
	2	3	linear	 NP
	6	0	octahedral	 NP
	4	2	square planar	 NP

	3	3	T Shaped	
	5	1	Square Pyramidal	
	2	4	Linear	