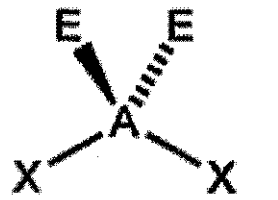

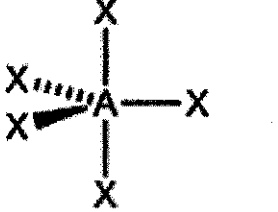
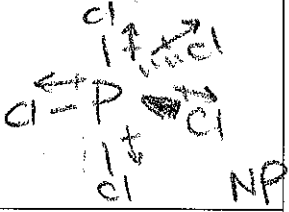
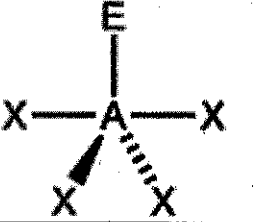

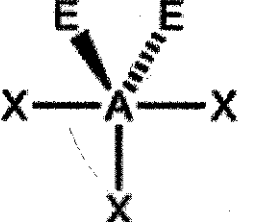
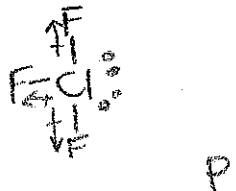
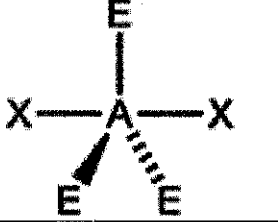
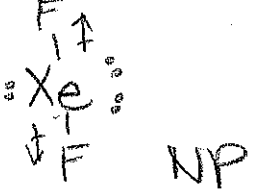
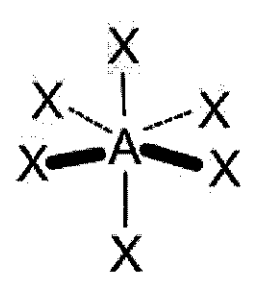
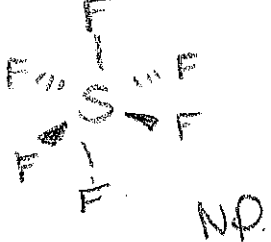
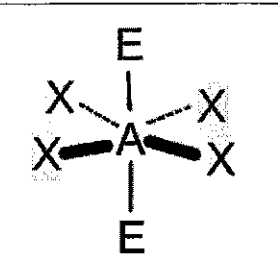
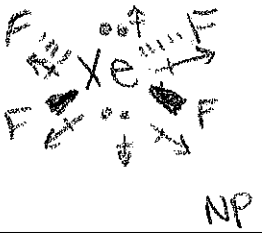
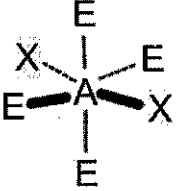


Note, this table utilized the AXE method. "A" is the central atom. "X" are the bonded atoms (peripheral atoms). "E" if the unbound/lone pairs of electrons.

Shape	# Bonding Pair(s)	# Lone electron pairs	Molecular Geometry	Example **Specify whether your example is polar or nonpolar
$A - X$	1	0	linear	$\ddot{N} = \ddot{O}$ NP
$X - A - X$	2	0	linear	$H - C \equiv N:$ P
$\begin{array}{c} X \\ \\ A \\ / \quad \backslash \\ X \quad X \end{array}$	3	0	trigonal planar	$\begin{array}{c} F \\ \\ B \\ / \quad \backslash \\ F \quad F \end{array}$ NP
$\begin{array}{c} E \\ \\ A \\ / \quad \backslash \\ X \quad X \end{array}$	2	1	bent	$\begin{array}{c} \cdot \cdot \cdot \cdot \\ \cdot \cdot \cdot \cdot \end{array}$ P
$\begin{array}{c} X \\ \\ A \\ / \quad \backslash \\ X \quad X \\ \quad \nearrow \\ \quad X \end{array}$	4	0	tetrahedral	$\begin{array}{c} F \\ \\ C \\ / \quad \backslash \\ H \quad H \\ \quad \nearrow \\ \quad H \end{array}$ P
$\begin{array}{c} E \\ \\ A \\ / \quad \backslash \\ X \quad X \\ \quad \nearrow \\ \quad X \end{array}$	3	1	Trigonal Pyramidal	$\begin{array}{c} \oplus \\ \cdot \cdot \cdot \cdot \\ \cdot \cdot \cdot \cdot \end{array}$ H ₂ O Charged

	2	2	Bent	
	5	0	trigonal bipyramidal	
	4	1	seesaw	
	3	2	T-shaped	
	2	3	linear	
	6	0	octahedral	
	4	2	square planar	

	3	3	T shaped	
	5	1	Square pyramidal	
	2	4	Linear	