

**1.3 Assignment: Nomenclature of Compounds**

Complete the table:

Formula	Group	Name
HClO	Covalent acid	Hypochlorous Acid
Formula	Group	Name
C <sub>2</sub> H <sub>6</sub>	Covalent hydrocarbon	Ethane
CuO	Ionic; divalent	Copper(II) Oxide
H <sub>3</sub> PO <sub>4</sub>	Covalent; acid	Phosphoric acid
K <sub>2</sub> CO <sub>3</sub>	Ionic; Monovalent	Potassium Carbonate
C <sub>6</sub> H <sub>10</sub>	Covalent; hydrocarbon	hexyne
C <sub>5</sub> H <sub>10</sub>	Covalent; hydrocarbon	Pentene
Cl <sub>2</sub> O	Covalent; binary	dichlorine monoxide
HF	Covalent; acid	Hydrofluoric Acid
Co(SO <sub>4</sub> )	Ionic; divalent	Cobalt (II) Sulfate
SnS <sub>2</sub>	Ionic; divalent	Tin(IV) Sulfide
XeF <sub>6</sub>	Covalent; binary	Xenon hexafluoride
AsF <sub>3</sub>	Covalent; binary	Arsenic trifluoride
HI	Covalent; acid	Hydroiodic acid
AuCl <sub>3</sub>	Ionic; divalent	Gold(III) Chloride
NaOH	Ionic; mono-valent	Sodium hydroxide