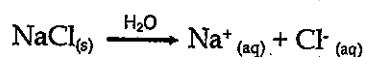


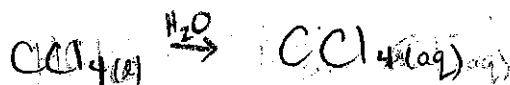
- Characteristics of Solutions Teacher.notebook

2. Determine whether each of the following substances will undergo dissociation, ionization or dissolving. Then write solvation equations for each of the substances when mixed with water. Show physical states of the products and balance the equations.

1. sodium chloride - Ionic; therefore will undergo dissociation in water

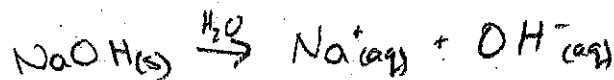


2. carbon tetrachloride - molecular; dissolving

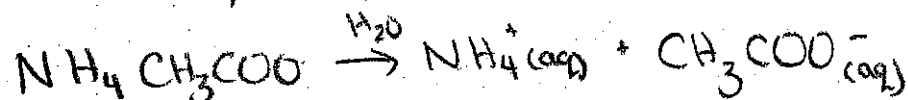


* diatomic elements don't have diatomic ions.

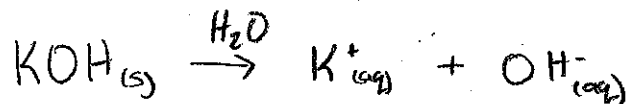
3. sodium hydroxide - ionic; dissociation



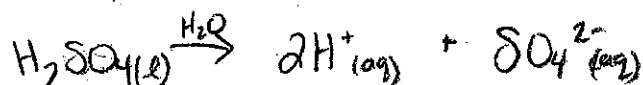
4. ammonium acetate ionic; dissociation



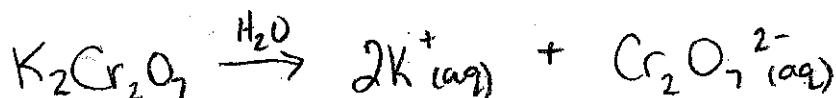
5. potassium hydroxide ionic; dissociation



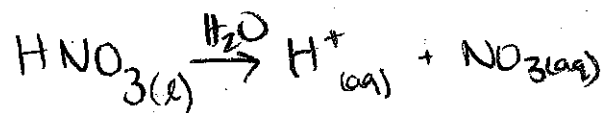
6. sulfuric acid molecular (acid); ionization



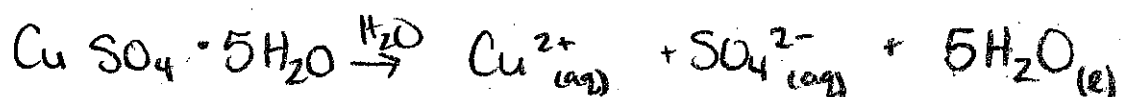
7. potassium dichromate ionic; dissociation



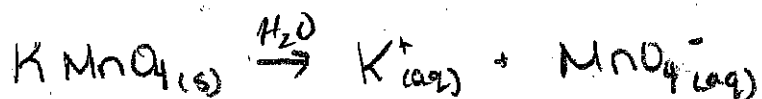
8. nitric acid molecular (acid); ionization



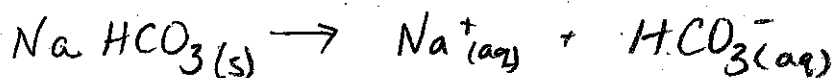
9. copper (II) sulfate pentahydrate ionic; dissociation



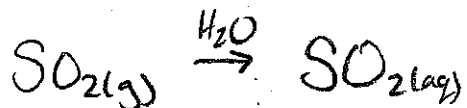
10. potassium permanganate ionic; dissociation



11. sodium bicarbonate ionic; dissociation



12. sulfur dioxide molecular; dissolving



13. hydrochloric acid molecular (acid); ionization

