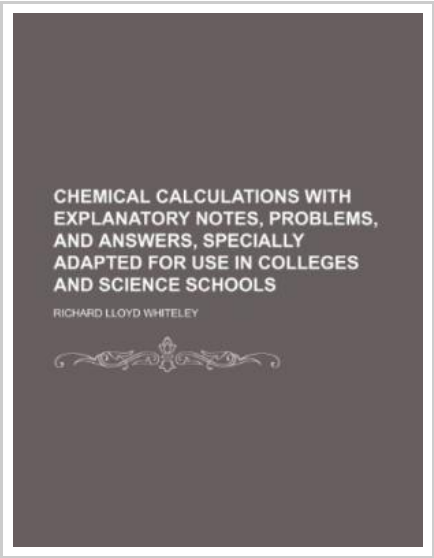




# Chemical Calculations with Explanatory Notes, Problems, and Answers, Specially Adapted for Use in Colleges and Science Schools (Paperback)

By Richard Lloyd Whiteley

Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1896 edition. Excerpt: .are required. Calculate the percentage of chlorine in the specimen. The equations representing the reactions are--KC103 + 3H, = KC1 + 3H20; KC1 + AgN03 = AgCl + KN03. Since the mol. wgt. of AgN03 = 170, a normal solution will contain 170 grams in 1 litre of solution; 1 c.c. will contain 0-170 gram AgN03. If the solution is deci-normal, since only--of 170 J 10 grams is dissolved, 1 c.c. will equal 0-0170 gram AgN03. Now 170 grams AgN03 will precipitate 35-5 grams of chlorine as AgCl;.-. 1 c.c. of N AgN03 (= 0-17 gram) will precipitate (or is equivalent to) 0-0355 gram of chlorine; and. . 1 c.c. of-AgN03 (= 0-017 gram) will precipitate (or is equivalent to) 0 00355 gram of chlorine. N In the example given, 26 c.c. of--AgN03 were used;, 26 x 0-00355 = the weight of 01, to which...



### Reviews

*Comprehensive guide! Its this sort of very good go through. It generally is not going to price too much. Its been designed in an remarkably basic way which is simply following i finished reading this pdf where really changed me, affect the way i really believe.*

-- ***Prof. Jeremie Blanda DDS***

*This publication is indeed gripping and intriguing. It is actually writter in basic terms and not difficult to understand. I am just pleased to explain how here is the greatest publication we have read through during my own lifestyle and could be he best pdf for at any time.*

-- ***Ervin Crona***