CAS number  RTECS number  - TR1580150  - TR1580150  - TR1580150  - Hermann Schnell  Details  - Bayer chemist synthesized polycarbonate in the first attempt; the experiment was based on analysis of previous research and selection of right building blocks; in order to succeed with implementation, he had to overcome scepticism of his pears that such reaction and outcome are possible; polymer was patented immediately but production begun in 1958  SYNTHESIS  Monomer(s) CAS number(s)  Monomer(s) Structure  - CH <sub>3</sub> Monomer(s) molecular weight(s)  Monomer(s) expected purity(les)  %  98; 99  Method of synthesis  - Bisphenol A is treated with NaOH, which is then reacted with phospene. It can also be manufactured by transesterification of bisphenol A with diphenyl carbonate.  **C  30-40 (polycondensation); 160-240 (transesterification)  Couper, J.R. Penney, W.R. Fair, J. F. Walas, S.M. Chemical Equipment, 2262-26, 2010.  **C  Catalyst  4 Interest and the control of the control				
Common name   -	PARAMETER	UNIT	VALUE	REFERENCES
December   Polytoxycarbonyloxy-1.4-phenylene(dimethylmethylmethylene)-1.4-phenylene	GENERAL			
ACS name  - carbonic acid, polymer with 4.4" (1-methylethylidene) bis[phenol]; polycarbonates  - CAS number - 25037-45-0; 25766-59-0  RTECS number - TR1580150  History  Person to discover - Hermann Schnell Datal  - 1953; 1958  Bayer chemist synthesized polycarbonate in the first attempt; the experiment was based on analysis of previous research and selection of right building blocks; in order to succeed with implementation, he had to overcome scepticism of his peers that such reaction and outcome are possible; polymer was patented immediately but production begun in 1968  SYNTHESIS  Monomer(s) CAS number(s)  Monomer(s) cas packed purity(ties)  Monomer(s) work of the succeed purity(ties)  Monomer(s) work of the succeed purity(ties)  Monomer(s) work of the succeed purity(ties)  Method of synthesis  - Sisphenol A is treated with NaOH, which is then reacted with phospene. It can also be manufactured by transesterification of bisphenol A with diphenyl carbonate.  Temperature of polymerization  Temperature of polymerization  Pa Sisphenol A is treated with NaOH, which is then reacted with phospene. It can also be manufactured by transesterification of bisphenol A with diphenyl carbonate.  Temperature of polymerization  Pa Sisphenol A is treated with NaOH, which is then reacted with phospene. It can also be manufactured by transesterification of bisphenol A with diphenyl carbonate.  Temperature of polymerization  Pa Sisphenol A is treated with NaOH, which is then reacted with phospene. It can also be manufactured by transesterification of the phospene of the phosp	Common name	-	polycarbonate	
Acronym  - PC  CAS number  - 25037-45-0; 25766-59-0  RTECS number  - TR1580150  - Hermann Schnell  Date  - 1953; 1958  Bayer chamist synthesized polycarbonate in the first attempt; the experiment was based on analysis of previous research and selection of right building blocks; in order to succeed with implementation, he had to overcome scepticism of his peers hat such reaction and outcome are possible; polymer was patented immediately but production begun in 1958  SYNTHESIS  Monomer(s) CAS number(s)  Monomer(s) cAS number(s)  Monomer(s) expected purity(les)  Monomer(s) expected purity(l	IUPAC name	-		
CAS number  - 25037-45-0; 25766-59-0  RTECS number  - TR1580150  - TR1580150  - Hermann Schnell  - 1953; 1958  Bayer chemist synthesized polycarbonate in the first attempt; the experiment was based on analysis of previous research and selection of right building blocks; in order to succeed with implementation, he had to overcome scepticism of his peers that such reaction and outcome are possible; polymer was patented immediately but production begun in 1958  SYNTHESIS  Monomer(s) CAS number(s)  Monomer(s) atructure  - CH <sub>3</sub> Monomer(s) molecular weight(s)  Monomer(s) expected purity(ies)  % 98; 99  98; 99  98; 99  98; 99  Method of synthesis  - Bisphenol A is treated with NaOH, which is then reacted with phospene. It can also be manufactured by transesterification of bisphenol A with diphenyl carbonate.  **C  30-40 (polycondensation); 160-240 (transesterification)  Couper, J.R. Penney, W.R. Fair, J. F. Walas, S.M. Chemical Equipment, J. Mim, Y. J. Kim, J. J. Kim, Y. J. Kim, J. J. Kim, Y. J. Kim, J. J. Kim, J	ACS name	-		
Formula  - TR1580150  - Hermann Schnell  1953; 1958  Bayer chemist synthesized polycarbonate in the first attempt; the experiment was based on analysis of previous research and selection of right building blocks; in order to succeed with implementation, he had to overcome scepticism of his peers that such reaction and outcome are possible; polymer was patented immediately but production begun in 1958  SYNTHESIS  Monomer(s) CAS number(s)  Monomer(s) structure  - B0-05-7; 75-44-5  daiton, g/ 28 29, 99.92  mol, amu  Monomer(s) expected purity(ies)  Method of synthesis  - Bisphenol A is treated with NaOH, which is then reacted with phosgene. It can also be manufactured by transesterification of bisphenol A with diphenyl carbonate.  Temperature of polymerization  **C  30-40 (polycondensation); 180-240 (transesterification)  **C Cuper, J R; Penney, W R; Fair, J R; Wallas, S M, Chemical Equipment, 2nd Ed., Elsevice; 2010, pp. 501-564, Km, J Kmn, V J, B; Roberts, G W; Oh, S-G, Polymer, 250-26, 2010.  **Time of polymerization  Peasure of polymerization  Pa atmospheric  Catalyst  daiton, g/  4 17,500-41,300	Acronym	-	PC	
HISTORY  Person to discover  Date  Details  - Hermann Schnell  Details  - Bayer chemist synthesized polycarbonate in the first attempt; the experiment was based on analysis of previous research and selection of right building blocks; in order to succeed with implementation, he had to overcome scepticism of his peers that such reaction and outcome are possible; polymer was patented immediately but production begun in 1958  SYNTHESIS  Monomer(s) structure  - GH3 OH CCl20  Monomer(s) cAS number(s)  Monomer(s) expected purity(ies)  Method of synthesis  - Bisphenol A is treated with NaOH, which is then reacted with phosgene. It can also be manufactured by transesterification of bisphenol A with diphenyl carbonate.  Temperature of polymerization  - C 30-40 (polycondensation); 160-240 (transesterification)  Time of polymerization  Pa 30-40 (polycondensation)  atmospheric  - Couper, J.R. Penney, W.R. Fair, J. R. Walss, S.M. Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-840, Km., J.Km. J.P. Ahmed, T.S. Dong, L.B. Roberts, G.W. Oh, SG. Polymer, 2520-26, 2010.  Time of polymerization  Pa atmospheric  Catalyst  - Bersylriethylammonium chloride (polycondensation); LiOH R. Walss, S.M. Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-840, Km., J.Km. J.P. Ahmed, T.S. Dong, L.B. Roberts, G.W. Oh, SG. Polymer, 2520-26, 2010.  Number average molecular  dalton, g/ 17,500-41,300	CAS number	-	25037-45-0; 25766-59-0	
HISTORY  Person to discover - Hermann Schnell  Date - 1953, 1958  Bayer chemist synthesized polycarbonate in the first attempt; the experiment was based on analysis of previous research and selection of right building blocks; in order to succeed with implementation, he had to overcome scepticism of his peers that such reaction and outcome are possible; polymer was patented immediately but production begun in 1968  SYNTHESIS  Monomer(s) CAS number(s)  Monomer(s) expected purity(ies)  Monomer(s) expected purity(ies)  Method of synthesis  - Bisphenol A is treated with NaOH, which is then reacted with phosgene. It can also be manufactured by transesterification of bisphenol A with diphenyl carbonate.  Temperature of polymerization  Catalyst  - Cuper, J.R. Penney, W.R. Fair, J. R. Wallas, S.M. Chemical Equipment, J. P. M. Wallas, S.M. Chemical Equipment, J. P.	RTECS number	-	TR1580150	
Person to discover Date Date Date Details Details Details  - 1953, 1958 Bayer chemist synthesized polycarbonate in the first attempt; the experiment was based on analysis of previous research and selection of right building blocks; in order to succeed with implementation, he had to overcome scepticism of his peers that such reaction and outcome are possible; polymer was patented immediately but production begun in 1958  SYNTHESIS  Monomer(s) CAS number(s)  Monomer(s) CAS number(s)  Monomer(s) molecular weight(s)  Monomer(s) molecular weight(s)  Method of synthesis  - 80-05-7; 75-44-5 228.29; 98.92  Method of synthesis  - Bisphenol A is treated with NaOH, which is then reacted with phosyene. It can also be manufactured by transesterification of bisphenol A with diphenyl carbonate.  Temperature of polymerization  Temperature of polymerization  Temperature of polymerization  A 0.25-4 (polycondensation): 160-240 (transesterification)  Time of polymerization  Pa atmospheric  benzyltriethylammonium chloride (polycondensation); LiOH (transesterification)  Pressure of polymerization  A through Table T	Formula		$\begin{array}{c c} & CH_3 \\ \hline \\ CH_3 \\ \hline \end{array} \begin{array}{c} O \\ O \\ -C \\ \end{array} \begin{array}{c} O \\ -C \\ \end{array}$	
Person to discover Date Date Date Details Details Details  - 1953, 1958 Bayer chemist synthesized polycarbonate in the first attempt; the experiment was based on analysis of previous research and selection of right building blocks; in order to succeed with implementation, he had to overcome scepticism of his peers that such reaction and outcome are possible; polymer was patented immediately but production begun in 1958  SYNTHESIS  Monomer(s) CAS number(s)  Monomer(s) CAS number(s)  Monomer(s) molecular weight(s)  Monomer(s) molecular weight(s)  Method of synthesis  - 80-05-7; 75-44-5 228.29; 98.92  Method of synthesis  - Bisphenol A is treated with NaOH, which is then reacted with phosyene. It can also be manufactured by transesterification of bisphenol A with diphenyl carbonate.  Temperature of polymerization  Temperature of polymerization  Temperature of polymerization  A 0.25-4 (polycondensation): 160-240 (transesterification)  Time of polymerization  Pa atmospheric  benzyltriethylammonium chloride (polycondensation); LiOH (transesterification)  Pressure of polymerization  A through Table T				
Date  Details  - 1953; 1958 - 1958 sayer chemist synthesized polycarbonate in the first attempt; the experiment was based on analysis of previous research and selection of right building blocks; in order to succeed with implementation, he had to overcome scepticism of his peers that such reaction and outcome are possible; polymer was patented immediately but production begun in 1958  SYNTHESIS  Monomer(s) structure  - 2H <sub>3</sub> OH CCl <sub>2</sub> O  Monomer(s) molecular weight(s) dallon, g/ mol, amu  Monomer(s) expected purity(ies)  Method of synthesis  - Bisphenol A is treated with NaOH, which is then reacted with phosgene. It can also be manufactured by transesterification of bisphenol A with diphenyl carbonate.  Temperature of polymerization  **C**  30-40 (polycondensation); 160-240 (transesterification)  **C**  **Outper, J.R. Penney, W.R. Fair, J.R. Walas, S.M. Chemical Equipment 2nd Ed. Elsewier, 2010, pp. 581-840; Kim, J. Kim, Y. Jr. Kim, J. Jr. Jh. Ahmed, T.S. long, L.B. Roberts, G.W. (oh, SG. Polymer, 2520-26, 2010.  **Note: Set of polymerization**  **Dessure of polymerization**  **Pressure of polymerization**  **Dessure of polymeriz	HISTORY			
Bayer chemist synthesized polycarbonate in the first attempt; the experiment was based on analysis of previous research and selection of right building blocks; in order to succeed with implementation, he had to overcome scepticism of his peers that such reaction and outcome are possible; polymer was patented immediately but production begun in 1958  SYNTHESIS  Monomer(s) structure  - B0-05-7; 75-44-5  dalton, g/mol, amu  Monomer(s) expected purity(ies)  Method of synthesis  - Bisphenol A is treated with NaOH, which is then reacted with phospene. It can also be manufactured by transesterification of bisphenol A with diphenyl carbonate.  Temperature of polymerization  Temperature of polymerization  Pa atmospheric  Denzylriethylammonium chloride (polycondensation); LiOH (transesterification)  Pressure of polymerization  Pa atmospheric  benzyltriethylammonium chloride (polycondensation); LiOH (transesterification)  Pa (T, Wolas, S, M, Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-840; Km, J, Km, J, J, Km, J, J, J, Km, J, J, J, J, Km, J,	Person to discover	-	Hermann Schnell	
the experiment was based on analysis of previous research and selection of right building blocks; in order to succeed with implementation, he had to overcome scepticism of his peers that such reaction and outcome are possible; polymer was patented immediately but production begun in 1958  SYNTHESIS  Monomer(s) structure  -	Date	-	1953; 1958	
Monomer(s) structure  - HO ← CH <sub>3</sub> ← CH <sub>2</sub> Monomer(s) CAS number(s)  - 80-05-7; 75-44-5  Monomer(s) molecular weight(s)  dalton, g/ mol, amu  Monomer(s) expected purity(ies)  Method of synthesis  - Bisphenol A is treated with NaOH, which is then reacted with phosgene. It can also be manufactured by transesterification of bisphenol A with diphenyl carbonate.  Temperature of polymerization  CC  30-40 (polycondensation); 160-240 (transesterification)  R; Walas, S M, Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-640; Kim, J; Kim, Y-J; Kim, J-D; Ahmed, TS; Dong, LB; Roberts, G W; Oh, S-G, Polymer, 2520-26, 2010.  Time of polymerization  Pa atmospheric  Catalyst  - benzyltriethylammonium chloride (polycondensation); LiOH (transesterification)  R; Walas, S M, Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-640; Kim, J; Kim, Y-J; Kim, J-D; Ahmed, TS; Dong, LB; Roberts, G W; Oh, S-G, Polymer, 2510-266, 2010.  Catalyst  dalton, g/  17,500-41,300	Details	-	the experiment was based on analysis of previous research and selection of right building blocks; in order to succeed with implementation, he had to overcome scepticism of his peers that such reaction and outcome are possible; polymer was	
Monomer(s) structure  - HO ← CH <sub>3</sub> ← CH <sub>2</sub> Monomer(s) CAS number(s)  - 80-05-7; 75-44-5  Monomer(s) molecular weight(s)  dalton, g/ mol, amu  Monomer(s) expected purity(ies)  Method of synthesis  - Bisphenol A is treated with NaOH, which is then reacted with phosgene. It can also be manufactured by transesterification of bisphenol A with diphenyl carbonate.  Temperature of polymerization  CC  30-40 (polycondensation); 160-240 (transesterification)  R; Walas, S M, Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-640; Kim, J; Kim, Y-J; Kim, J-D; Ahmed, TS; Dong, LB; Roberts, G W; Oh, S-G, Polymer, 2520-26, 2010.  Time of polymerization  Pa atmospheric  Catalyst  - benzyltriethylammonium chloride (polycondensation); LiOH (transesterification)  R; Walas, S M, Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-640; Kim, J; Kim, Y-J; Kim, J-D; Ahmed, TS; Dong, LB; Roberts, G W; Oh, S-G, Polymer, 2510-266, 2010.  Catalyst  dalton, g/  17,500-41,300	OVALTUEOLO			
Monomer(s) CAS number(s)  Monomer(s) molecular weight(s)  datton, g/ mol, amu  Method of synthesis  - Bisphenol A is treated with NaOH, which is then reacted with phosgene. It can also be manufactured by transesterification of bisphenol A with diphenyl carbonate.  Temperature of polymerization  Catalyst  - Catalyst  HO				
Monomer(s) molecular weight(s)       dalton, g/ mol, amu       228.29; 98.92         Monomer(s) expected purity(ies)       %       98; 99         Method of synthesis       -       Bisphenol A is treated with NaOH, which is then reacted with phospene. It can also be manufactured by transesterification of bisphenol A with diphenyl carbonate.       Couper, J R; Penney, W R; Fair, J R; Walas, S M, Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-640; Kim, J; Kim, Y J; Kim, J-D; Ahmed, T S; Dong, L B; Roberts, G W; Oh, S-G, Polymer, 2520-26, 2010.         Time of polymerization       h       0.25-4 (polycondensation)         Pressure of polymerization       Pa       atmospheric         Catalyst       -       benzyltriethylammonium chloride (polycondensation); LiOH (transesterification)       Couper, J R; Penney, W R; Fair, J R; Walas, S M, Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-640; Kim, J; Kim, Y J; Kim, J-D; Ahmed, T S; Dong, L B; Roberts, G W; Oh, S-G, Polymer, 2520-26, 2010.         Number average molecular       dalton, g/       17,500-41,300	Monomer(s) structure	-	HO— OH CCI <sub>2</sub> O	
Monomer(s) expected purity(ies)       98; 99         Method of synthesis       -       Bisphenol A is treated with NaOH, which is then reacted with phosgene. It can also be manufactured by transesterification of bisphenol A with diphenyl carbonate.         Temperature of polymerization       °C       30-40 (polycondensation); 160-240 (transesterification)       Couper, J R; Penney, W R; Fair, J R; Walas, S M, Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-640; Kim, J-D; Ahmed, T S; Dong, L B; Roberts, G W; Oh, S-G, Polymer, 2520-26, 2010.         Time of polymerization       Pa       atmospheric         Catalyst       -       benzyltriethylammonium chloride (polycondensation); LiOH (transesterification)       Couper, J R; Penney, W R; Fair, J R; Walas, S M, Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-640; Kim, J; Kim, Y J; Kim, J-D; Ahmed, T S; Dong, L B; Roberts, G W; Oh, S-G, Polymer, 2520-26, 2010.         Number average molecular       dalton, g/       17,500-41,300	Monomer(s) CAS number(s)	-	80-05-7; 75-44-5	
Bisphenol A is treated with NaOH, which is then reacted with phosgene. It can also be manufactured by transesterification of bisphenol A with diphenyl carbonate.    Couper, J R; Penney, W R; Fair, J R; Walas, S M, Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-640; Kim, J; Kim, J-D; Ahmed, T S; Dong, L B; Roberts, G W; Oh, S-G, Polymer, 2520-26, 2010.   Couper, J R; Penney, W R; Fair, J R; Walas, S M, Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-640; Kim, J; Kim, Y J; Kim, J-D; Ahmed, T S; Dong, L B; Roberts, G W; Oh, S-G, Polymer, 2520-26, 2010.   Couper, J R; Penney, W R; Fair, J R; Walas, S M, Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-640; Kim, J; Kim, J-D; Ahmed, T S; Dong, L B; Roberts, G W; Oh, S-G, Polymer, 2520-26, 2010.   Number average molecular   dalton, g/ 17,500-41,300	Monomer(s) molecular weight(s)		228.29; 98.92	
phosgene. It can also be manufactured by transesterification of bisphenol A with diphenyl carbonate.  Temperature of polymerization  °C  30-40 (polycondensation); 160-240 (transesterification)  Couper, J R; Penney, W R; Fair, J R; Walas, S M, Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-640; Kim, J; Kim, J-D; Ahmed, T S; Dong, L B; Roberts, G W; Oh, S-G, Polymer, 2520-26, 2010.  Time of polymerization  Pa atmospheric  - benzyltriethylammonium chloride (polycondensation); LiOH (transesterification)  Catalyst  Couper, J R; Penney, W R; Fair, J R; Walas, S M, Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-640; Kim, J; Kim, J-J; Kim, J-J; Ahmed, T S; Dong, L B; Roberts, G W; Oh, S-G, Polymer, 2520-26, 2010.  Number average molecular  dalton, g/ 17,500-41,300	Monomer(s) expected purity(ies)	%	98; 99	
R; Walas, S M, Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-640; Kim, J; Kim, Y J; Kim, J-D; Ahmed, T S; Dong, L B; Roberts, G W; Oh, S-G, Polymer, 2520-26, 2010.  Time of polymerization  Pressure of polymerization  Pa atmospheric  Catalyst  - benzyltriethylammonium chloride (polycondensation); LiOH (transesterification)  Couper, J R; Penney, W R; Fair, J R; Walas, S M, Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-640; Kim, J; Kim, J-D; Ahmed, T S; Dong, L B; Roberts, G W; Oh, S-G, Polymer, 2520-26, 2010.  Number average molecular  dalton, g/ 17,500-41,300	Method of synthesis	-	phosgene. It can also be manufactured by transesterification	
Pressure of polymerization  Catalyst  - benzyltriethylammonium chloride (polycondensation); LiOH (transesterification)  - Couper, J R; Penney, W R; Fair, J R; Walas, S M, Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-640; Kim, J Kim, Y J; Kim, J-D; Ahmed, T S; Dong, L B; Roberts, G W; Oh, S-G, Polymer, 2520-26, 2010.  Number average molecular  dalton, g/ 17,500-41,300	Temperature of polymerization	°C	30-40 (polycondensation); 160-240 (transesterification)	R; Walas, S M, Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-640; Kim, J; Kim, Y J; Kim, J-D; Ahmed, T S; Dong, L B; Roberts, G W; Oh, S-G, Polymer,
Catalyst  - benzyltriethylammonium chloride (polycondensation); LiOH (transesterification)  - Couper, J R; Penney, W R; Fair, J R; Walas, S M, Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-640; Kim, J; Kim, Y J; Kim, J-D; Ahmed, T S; Dong, L B; Roberts, G W; Oh, S-G, Polymer, 2520-26, 2010.  Number average molecular  dalton, g/ 17,500-41,300	Time of polymerization	h	0.25-4 (polycondensation)	
(transesterification)  R; Walas, S M, Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-640; Kim, J; Kim, Y J; Kim, J-D; Ahmed, T S; Dong, L B; Roberts, G W; Oh, S-G, Polymer, 2520-26, 2010.  Number average molecular  dalton, g/ 17,500-41,300	Pressure of polymerization	Pa	atmospheric	
	Catalyst	-		R; Walas, S M, Chemical Equipment, 2nd Ed., Elsevier, 2010, pp. 581-640; Kim, J; Kim, Y J; Kim, J-D; Ahmed, T S; Dong, L B; Roberts, G W; Oh, S-G, Polymer,
	Number average molecular weight, $\mathbf{M}_{\mathrm{n}}$		17,500-41,300	

PARAMETER	UNIT	VALUE	REFERENCES
	dalton, g/ mol, amu	19,000-56,000	
Polydispersity, M <sub>w</sub> /M <sub>n</sub>	-	1.3-3.2	
Molar volume at 298K	cm <sup>3</sup> mol <sup>-1</sup>	calc.=213.8; exp.=211.9	
Van der Waals volume	cm <sup>3</sup> mol <sup>-1</sup>	calc.=136.21; exp.=138.36	
Molecular cross-sectional area, calculated	cm <sup>2</sup> x 10 <sup>-16</sup>	19.8	
Radius of gyration	nm	13	Fornes, T D; Baur, J W; Sabba, Y; Thomas, E L, Polymer, 47, 1704-14, 2006.
STRUCTURE			
Crystallinity	%	typically amorphous because of its rigid backbone; 20-42	Farmer, R, A Study of Crystallization in Bisphenol A Polycarbonate. Dissertation, Virginia Polytgechnic Institute, 2001; Kim, J; Kim, Y J; Kim, J-D; Ahmed, T S; Dong, L B; Roberts, G W; Oh, S-G, Polymer, 2520-26, 2010.
Entanglement molecular weight	dalton, g/ mol, amu	calc.=1,734-2,495; exp.=2,495-4,800	
Lamellae thickness	nm	3.18-5.39	Farmer, R, A Study of Crystallization in Bisphenol A Polycarbonate. Dissertation, Virginia Polytgechnic Institute, 2001.
COMMERCIAL POLYMERS			
Some manufacturers	-	Bayer; Sabic	
Trade names	-	Makrolon; Lexan	
PHYSICAL PROPERTIES			
Density at 20°C	g cm <sup>-3</sup>	1.19-1.22; 1.04 (melt); 1.25-1.52 (10-40% glass fiber)	
Bulk density at 20°C	g cm <sup>-3</sup>	0.66 (pellets)	
Refractive index, 20°C	-	calc.=1.5773-1.587; exp.=1.586-1.587	
Birefringence	-	-0.001 (surface) to 0.001 (bulk)	Lin, T H; Isayev, A I, Antec, 1664-68, 2006.
Transmittance	%	82-91	
Haze	%	<0.8 to 3	
Melting temperature, DSC	°C	255-267	
Storage temperature (max)	°C	93	
Cecomposition onset temperature temperature	°C	420	
Thermal expansion coefficient, -40 to 95°C	°C <sup>-1</sup>	6.0-7.5E-5; 1.6-6.5E-5 (10-40% glass fiber)	
Thermal conductivity, melt	W m <sup>-1</sup> K <sup>-1</sup>	0.19-0.24; 0.2-0.22 (10-40% glass fiber)	
Glass transition temperature	°C	calc.=134-158; exp.=137-154	
Specific heat capacity	J K <sup>-1</sup> kg <sup>-1</sup>	1,200; 1,040-1,210 (10-40% glass fiber)	
	kJ kg <sup>-1</sup>	100-115 (crystal); 22.95-23.11 (heat to melt)	DeLassus, P T; Landes, B G; Har-

- C poryour normato			
PARAMETER	UNIT	VALUE	REFERENCES
Maximum service temperature	°C	-40 to 130	
Temperature index (50% tensile strength loss after 20,000 h/5000 h)	°C	120	
Heat deflection temperature at 0.45 MPa	°C	133-166; 141-154 (10-40% glass fiber)	
Heat deflection temperature at 1.8 MPa	°C	119-156; 135-146 (10-40% glass fiber)	
Vicat temperature VST/A/50	°C	138-150; 144-149 (10-35% glass fiber)	
Vicat temperature VST/B/50	°C	144-170; 154 (10-40% glass fiber)	
Hansen solubility parameters, $\boldsymbol{\delta}_{\mathrm{D}}$ , $\boldsymbol{\delta}_{\mathrm{P}}$ , $\boldsymbol{\delta}_{\mathrm{H}}$	MPa <sup>0.5</sup>	19.6, 8.8, 5.7	
Interaction radius		10.2	
Hildebrand solubility parameter	MPa <sup>0.5</sup>	22.2	
Surface tension	mN m <sup>-1</sup>	calc.=33.0-37.9; exp.=28.4-42.9	Liao, C-C; Wang, C-C; Shih, K-C; Chen, C-Y, Eur. Polym. J., 47, 911-24, 2011.
Dielectric constant at 100 Hz/1 MHz	-	3.2/2.9	
Relative permittivity at 100 Hz	-	2.9-3.2; 3.1-3.6 (10-40% glass fiber)	
Relative permittivity at 1 MHz	-	2.8-3.1; 3.0-3.6 (10-40% glass fiber)	
Dissipation factor at 100 Hz	E-4	5-310; 8-13 (10-40% glass fiber)	
Dissipation factor at 1 MHz	E-4	90-120; 67-90 (10-40% glass fiber)	
Volume resistivity	ohm-m	1E12 to 1E15	
Surface resistivity	ohm	1E15 to 1E17	
Electric strength K20/P50, d=0.60.8 mm	kV mm <sup>-1</sup>	15-67; 18-36 (10-40% glass fiber)	
Comparative tracking index, CTI, test liquid A	-	200-250; 175 (10-35% glass fiber)	
Comparative tracking index, CTIM, test liquid B	-	125M	
Coefficient of friction	ASTM D1894	0.21 (chrome steel); 0.41-0.54 (aluminum)	Maldonado, J E, Antec, 3431-35, 1998.
Permeability to nitrogen, 25°C	cm <sup>3</sup> m <sup>-2</sup> 24 h <sup>-1</sup> bar <sup>-1</sup>	130 (100 μm film), 510 (25.4 μm film)	
Permeability to oxygen, 25°C	cm <sup>3</sup> m <sup>-2</sup> 24 h <sup>-1</sup> bar <sup>-1</sup>	700 (100 μm film), 2,760 (25.4 μm film)	
Permeability to water vapor, 25°C	g m <sup>-2</sup> 24 h <sup>-1</sup> bar <sup>-1</sup>	15	
Contact angle of water, 20°C	degree	81.3-84.0	
Surface free energy	mJ m <sup>-2</sup>	42.3	
Speed of sound	m s <sup>-1</sup>	38	
Acoustic impedance		2.69-2.77	
Attenuation	dB cm <sup>-1</sup> ,	22.1-24.9	

PARAMETER	UNIT	VALUE	REFERENCES
MECHANICAL & RHEOLOGICAL P	ROPERTIE	S	
Tensile strength	MPa	55-88; 45-158 (10-40% glass fiber)	Chang, M C O; Garrett, P D, Antec, 2588-93, 1996.
Tensile modulus	MPa	2,200-3,100; 3,800-9,400 (10-35% glass fiber)	
Tensile stress at yield	MPa	57-74	Chang, M C O; Garrett, P D, Antec, 2588-93, 1996.
Tensile creep modulus, 1000 h, elongation 0.5 max	MPa	1,700-1,900; 2,900-8,500 (10-35% glass fiber)	
Elongation	%	66-140; 1.8-15 (10-40% glass fiber)	Chang, M C O; Garrett, P D, Antec, 2588-93, 1996.
Tensile yield strain	%	5.7-7; 8 (10-40% glass fiber)	
Flexural strength	MPa	94-120; 103-186 (10-40% glass fiber)	
Flexural modulus	MPa	2,220-2,600; 3,440-9,600 (10-40% glass fiber)	
Elastic modulus	MPa	1,600	
Compressive strength	MPa	70	
Young's modulus	MPa	2,390-2,600	Chang, M C O; Garrett, P D, Antec, 2588-93, 1996.
Charpy impact strength, unnotched, 23°C	kJ m <sup>-2</sup>	no break; 40-150 (10-35% glass fiber)	
Charpy impact strength, unnotched, -30°C	kJ m <sup>-2</sup>	no break	
Charpy impact strength, notched, 23°C	kJ m <sup>-2</sup>	11-80; 8-12 (10-35% glass fiber)	
Charpy impact strength, notched, -30°C	kJ m <sup>-2</sup>	9-14	
Izod impact strength, unnotched, 23°C	J m <sup>-1</sup>	12-90 to NB; 1280-2140 (10-40% glass fiber)	
Izod impact strength, notched, 23°C	J m <sup>-1</sup>	12-736; 8-133 (10-40% glass fiber)	
Izod impact strength, notched, -30°C	J m <sup>-1</sup>	55-618	
Shear modulus	MPa	805	Ozcelik, B; Sonat, I, Mater. Design, 30, 367-75, 2009.
Abrasion resistance (ASTM D1044)	mg/1000 cycles	11 (10-40% glass fiber)	
Poisson's ratio	-	calc.=0.424; exp.=0.401-0.420	Lin, T H; Isayev, A I, Antec, 1664- 68, 2006.
Rockwell hardness	-	L89; M85, R124 (10-40% glass fiber)	
Ball indention hardness at 358 N/30 S (ISO 2039-1)	MPa	116	
Shrinkage	%	0.4-0.9; 0.2-0.55 (10-35% glass fiber)	
Melt volume flow rate (ISO 1133, procedure B), 300°C/1.2 kg	cm <sup>3</sup> /10 min	1.25-36	
Pressure coefficient of melt viscosity, b	G Pa <sup>-1</sup>	26.6	Aho, J; Syrjala, S, J. Appl. Polym. Sci., 117, 1076–84, 2010.
Melt index, 300°C/1.2 kg	g/10 min	6-80	
Water absorption, equilibrium in water at 23°C	%	0.12-0.40; 0.23 (10-40% glass fiber)	
Moisture absorption, equilibrium 23°C/50% RH	%	0.09-0.3	
23 G/30 /0 KH			

PARAMETER	UNIT	VALUE	REFERENCES
CHEMICAL RESISTANCE			
Acid dilute/concentrated	-	good/poor	
Alcohols	-	good	
Alkalis	-	good-poor	
Aliphatic hydrocarbons	-	good	
Aromatic hydrocarbons	-	poor	
Esters	-	poor	
Greases & oils	-	good-poor	
Halogenated hydrocarbons	-	poor	
Ketones	-	poor	
⊙ solvent, ⊙-temp.=170°C	-	n-butyl benzyl ether	
Good solvent	-	acetophenone (hot), aniline (hot), benzene (hot), chloroform, cresol, 1,2-dichloroethane, methylene chloride	
Non-solvent	-	amyl alcohol, ethylene glycol, heptane, isopropyl alcohol	
Chemicals causing environmental stress cracking	-	acetone, benzyl alcohol, carbon tetrachloride, cyclohexanone, nitrobenzene	Wang, H T; Pan, Q G; Du, Q C; Li, Y Q, Polym. Test., 22, 125-28, 2003.
Effect of EtOH sterilization (tensile strength retention)	%	98-100	Navarrete, L; Hermanson, N, Antec, 2807-18, 1996.
EL AMMA DIL ITV			
FLAMMABILITY	00	400	
Ignition temperature	°C	480	
Autoignition temperature	°C	550	
Limiting oxygen index	% O <sub>2</sub>	25-30; 30-43 (flame retardant grades)	
Minimum ignition energy	J	0.025	
Heat release	kW m <sup>-2</sup>	479-548; 124-385 (with fire retardant)	Yu, B; Liu, M; Lu, L; Dong, X; Gao, W; Tang, K, Fire Mater., 34, 251-61, 2010.
NBS smoke chamber	Ds	190	Padey, D; Walling, J; Wood A, Polymers in Defence and Aero- space 2007, Rapra, 2007, paper 15.
Burning rate (Flame spread rate)	mm min <sup>-1</sup>	passed	
Char at 500°C	%	21.7	Lyon, R E; Walters, R N, J. Anal. Appl. Pyrolysis, 71, 27-46, 2004.
Heat of combustion	J g <sup>-1</sup>	31,060-31,530	Walters, R N; Hacket, S M; Lyon, R E, Fire Mater., 24, 5, 245-52, 2000.
Volatile products of combustion	-	carbon monoxide, carbon dioxide, bisphenol A, diphenyl carbonate, phenol and phenol derivatives. Traces of aliphatic and aromatic hydrocarbons, aldehydes and acids.	
UL rating	-	HB to V-2; V-0 (some flame retardant grades); HB to V-0 (10-35% glass fiber)	
WEATHER STABILITY			
Spectral sensitivity	nm	<275 (completely absorbed); 260-300 (photo-Fries rearrangement); 280-305; 330-360	
Activation wavelengths	nm	290-320; 310 (chain scission), 330-360	
Important initiators and accelerators	-	4-hydroxystilbene; products of thermal degradation; bisphenol A, stilbene-like structures, water, bis(3-hydroxyphenyl)ether structures in main chain, some inorganic pigments	

PARAMETER	UNIT	VALUE	REFERENCES
Products of degradation	-	photo-Fries rearrangement; chain scissions, crosslinks, free radicals, hydroxyl groups, ethers, unsaturations (photolysis); chain scissions, hydroperoxides, free radicals, hydroxyl groups, carbonyl groups (photooxidation)	
Stabilizers	-	UVA: 2-hydroxy-4-octyloxybenzophenone; 2,2',4,4'-tetrahydroxybenzophenone; 2-(2H-benzotriazol-2-yl)-p-cresol; 2-benzotriazol-2-yl-4,6-di-tert-butylphenol; 2-(2H-benzotriazole-2-yl)-4,6-di-tert-pentylphenol; 2-(2H-benzotriazole-2-yl)-4,6-di-tert-pentylphenol; 2-(2H-benzotriazol-2-yl)-4,6-bis(1-methyl-1-phenylethyl) phenol; 2,2'-methylenebis(6-(2H-benzotriazol-2-yl)-4-1,1,3,3-tetramethylbutyl)phenol; 2-(2H-benzotriazol-2-yl)-4-1,1,3,3-tetramethylbutyl)phenol; 2-(2H-benzotriazol-2-yl)-6-dodecyl-4-methylphenol, branched & linear; 2-(2'-hydroxy-5'-methacryloxyethylphenyl)-2H-benzotriazole; 2-(2H-benzotriazol-2-yl)-4,6-bis(1-methyl-1-phenylethyl) phenol; 2-(3-sec-butyl-5-tert-butyl-2-hydroxyphenyl) benzotriazole; reaction product of methyl 3(3-(2H-benzotriazole; reaction product of methyl 3(3-(2H-benzotriazole-2-yl)-5-t-butyl-4-hydroxyphenyl propionate/PEG 300; 2-(4,6-diphenyl-1,3,5-triazin-2-yl)-5-hexyloxy-phenol; 2-[4,6-bis(2,4-dimethylphenyl)-1,3,5-triazin-2-yl]-5-(octyloxy) phenol; 1,3-bis-[(2'-cyano-3',3'-diphenylacryloyl)oxy]methyl-propane; propanedioic acid, [(4-methoxyphenyl)-methylene]-dimethyl ester; 2,2'-(1,4-phenylene)bis[4H-3,1-benzoxazin-4-one]; Screener: zinc oxide; Phenolic antioxidant: 2-(1,1-dimethylethyl)-6-[[3-(1,1-dimethylethyl)-2-hydroxy-5-methylphenyl] methyl-4-methylphenyl acrylate; Phosphite: tris (2,4-di-tert-butylphenyl) phosphite; isodecyl diphenyl phosphite; di(p-butoxyphenyl) cyclohexylphosphine oxide; Optical brightener: 2,2'-(2,5-thiophenediyl)bis(5-tert-butylbenzoxazole; Plasticizer: dicyclohexyl phthalate (reduces yellowing on exposure to gamma radiation	
Results of exposure		yellowness index: 45 after 60 month exposure in Florida without stabilization and 22 after 60 month exposure in Florida with stabilization; 13 and 26 after 120 month exposure in Engerfeld, Germany with and without stabilization, respectively	
Low earth orbit erosion yield	cm <sup>3</sup> atom <sup>-1</sup> x 10 <sup>-24</sup>	4.29	Waters, D L; Banks, B A; De Groh, K K; Miller, S K R; Thorson, S D, High Performance Polym., 20, 512-22, 2008.
BIODEGRADATION			
Colonized products		compact disks, medical devices	
Typical biodegradants	-	microorganisms producing esterase are capable of damage; marine environment	Artham, T; Doble, M, J. Polym. Environ., 17, 170-80, 2009.
Stabilizers	-	silver-based biocides	
TOXICITY			
NFPA: Health, Flammability, Reactivity rating	-	0/1/0	
Carcinogenic effect	-	not listed by ACGIH, NIOSH, NTP	
OSHA	mg m <sup>-3</sup>	5 (respirable), 15 (total)	
Acceptable daily intake	μg kg <sup>-1</sup> body weight day <sup>-1</sup>		
Oral rat, LD <sub>50</sub>	mg kg <sup>-1</sup>	>10,000	
Skin rabbit, LD <sub>50</sub>	mg kg <sup>-1</sup>	>2,000	
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PARAMETER	UNIT	VALUE	REFERENCES
ENVIRONMENTAL IMPACT			
Aquatic toxicity, Daphnia magna, $LC_{50}$ , 48 h	mg l <sup>-1</sup>	5,000	Lither, D; Damberg, J; Dave, G; Larsson, A, Chemosphere, 74, 1198 <sup>-1</sup> 200, 2009.
Cradle to grave non-renewable energy use	MJ/kg	111	
Cradle to pellet greenhouse gasses	kg CO <sub>2</sub> kg <sup>-1</sup> resin	7.8	
PROCESSING			
Typical processing methods	-	blow molding, calendering, electrospinning, extrusion, gas- assisted injection molding, injection molding, solution casting, thermoforming	
Preprocess drying: temperature/ time/residual moisture	°C/h/%	120-135/2-12/0.02	
Processing temperature	°C	280-355; 310-340 (10-40% glass fiber)	
Processing pressure	MPa	0.3-0.7 (back pressure)	
Additives used in final products	-	Fillers: boric oxide, glass beads, carbon black and graphite fibers for EMI shielding, glass fiber, graphite, molybdenum sulfide, nanosilica, nickel-coated graphite fibers, PTFE, single-walled carbon nanotubes, steel fibers, titanium dioxide, wollastonite; Plasticizers: dibutyl phthalate, dioctyl phthalate, dicyclohexyl phthalate, mineral oil, pentaerythritol tetraborate, trimellitic acid tridecyloctyl ester, tritolyl phosphate, tetraethylene glycol dimethyl ether, tri-(2-ethylhexyl) phosphate; Antistatics: carbon black, carbon nanotubes, copper oxide, glycerol mono-iso-stearate, indium tin oxide, nickel-coated carbon fiber, polyetheresteramide, polyoxyethylene fatty acid ester; Antiblocking: amorphous silica, calcium carbonate, dimethylsiloxane grafting, siloxane particles; Release: glycerol monostearate, pentaerythritol tetrastearate, siloxane, zinc stearate; impact modifier; UV stabilizers; release agents	
Applications	-	bearings, blood collector containers, camera components, computer printers, copying machines, corrective eyeglasses, dental applications, data storage (CD, DVD, etc.), dinnerware, drinking cups, disposable syringes, head lamp covers and housings, gears, glazing, goggles, golf tees, guide pins, helmets, instrument panels, laminated walls, lenses, medical tubing, microfibers, needles syringes, optical lenses, pacemaker components, projection screens, rollers, roofing, safety glasses, skylights, speedometer needles, solar modules, tool boxes, toys, water bottles, windows, windscreens	
BLENDS			
Suitable polymers	-	ABS, PBT, PET, PHEMA, PLA, PVDF	
ANALYSIS			
FTIR (wavenumber-assignment)	cm <sup>-1</sup> /-	C-H – 3047, 2877; C-O-C – 989; C=O – 1778	Abdel-Salam, M H; Nouh, S A; Radwan, Y E; Fouad, S S, Mater. Chem. Phys., 127, 305-9, 2011.
Raman (wavenumber-assignment)	cm <sup>-1</sup> /-	C=O – 1769, 1597, 712; CH <sub>3</sub> – 1453	Hoeller, T L, Antec, 3124-30, 2007.