

Department of Physics, Chemistry and Biology

Master's Thesis

Optical properties of free-standing cubic silicon carbide

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Linköping, June 4, 2015

LiTH-IFM-A-EX--15/3100--SE



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Avdelning, Institution
Division, Department

Semiconductor physics
Department of Physics, Chemistry and Biology
SE-581 83 Linköping

Datum
Date

2015-06-04

Språk
Language

Svenska/Swedish
 Engelska/English

Rapporttyp
Report category

Licentiatavhandling
 Examensarbete
 C-uppsats
 D-uppsats
 Övrig rapport

ISBN
—
ISRN
LiTH-IFM-A-EX-15/3100--SE

Serietitel och serienummer **ISSN**
Title of series, numbering —

URL för elektronisk version

Titel Optiska egenskaper för fristående kubisk kiselkarbid
Title Optical properties of free-standing cubic silicon carbide

Författare Mattias Jansson
Author

Sammanfattning
Abstract

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I have studied seed growth of nominally undoped cubic material on hexagonal (4H) substrates, and the influence on the grown material from the different faces of the substrate. It is found that it is not possible under the explored conditions to completely cover the growth area with the cubic polytype on the carbon face, but it can be done reproducibly on the silicon face. Reasons for this are discussed. Different doping setups are also explored.

The influence on the material properties from growth conditions is explored. It is shown from absorption measurements that it is possible to grow boron doped cubic silicon carbide using this growth method, whereas optical microscopy studies show that the sample quality degrades with high doping concentrations.

I have explored the luminescence properties of the material. No boron related emission is found with either room temperature or low temperature photoluminescence spectroscopy. Reasons for this are discussed using results from absorption measurements and optical microscopy.

Nyckelord
Keywords semiconductor, silicon carbide, SiC, photovoltaic, water splitting, sublimation

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The properties of free-standing cubic silicon carbide for optoelectronic applications are explored in this work. The main focus of the work is on boron doped cubic silicon carbide, which is proposed as a highly useful material in several optoelectronic applications. The material is grown using sublimation epitaxy and the doped material is grown homoepitaxially on nominally undoped seeds. It is characterized using the experimental setups of photoluminescence spectroscopy, Nomarski interference spectroscopy and absorption spectroscopy.

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Acknowledgments

I wish to thank my supervisor, Jianwu Sun, for teaching me the experimental methods and for invaluable guidance in interpreting the results. I also wish to direct my thanks to Mikael Syväjärvi - thank you for taking me in as a master's student in your lab, and for the help and input in planning my work and writing this thesis. I really appreciate having had the opportunity to be a member of your group during this time. Finally thank you also to Valdas Jokubavicius and Xinyu Liu for your support and guidance in the lab.

*Linköping, June 2015
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1

Introduction

Silicon carbide (SiC) is a semiconductor which has attracted interest in research and industry since the 19:th century, when it was first fabricated and used as an abrasive [1]. Silicon carbide has been found to be a very stable material. It exhibits a high chemical inertness [2], and is currently commonly used in high power, temperature and radiation applications due to its ability to survive in such environments [3].

Silicon carbide is a material which exists in a large number of different polytypes, the most common of which are the hexagonal, cubic and rhombohedral polytypes. The work described in this thesis deals with the cubic polytype, denoted *3C-SiC*. This is one of the structurally most simple polytypes, and the only cubic one. Compared to the hexagonal counterparts 4H- and 6H-SiC, the 3C-SiC polytype has for a long time been difficult to fabricate in good quality and large volume. It is therefore less studied than the hexagonal types. Recently a method of fabricating good quality free standing cubic material using sublimation growth has been reported [4]. This method has been used to grow the samples used in the work described in this thesis. With this method it is possible to grow free-standing 3C-SiC, which means that the material is thick enough to exist on its own after mechanically polishing away the 4H-SiC substrate.

Cubic SiC has many interesting material properties. It has a higher electron mobility compared to the common hexagonal polytypes [5]. It has also attracted attention as a transistor material, since it can achieve a low number of interface defects in such structures [6]. One proposed application for which 3C-SiC is well suited is the use of boron doped 3C-SiC in an intermediate band photovoltaic solar cell. This is suitable for 3C-SiC due to its band gap size of 2.36 eV, together with the binding energy of boron as an acceptor in the material, which is almost ideal for photovoltaic cell material. This would give a significant increase in photovoltaic cell efficiency compared to the conventional single junction alternative [7, 8]. Another proposed application of 3C-SiC is as a photo-electrode in a photoelectrochemical cell used for water splitting [9, 10], where solar energy is used in the decomposition of

water into hydrogen and oxygen gas.

This thesis describes the growth and optical characterization of free-standing cubic silicon carbide, which is vital for the realization of 3C-SiC optoelectronic applications such as the intermediate band solar cell and the water splitting electrode. The work carried out in this project can be divided into two parts. The first part deals with growth of the boron doped material which has been characterized in this thesis. This part includes growth of nominally undoped seeds and growth of the boron doped material. The second part deals with characterization of the grown samples. Characterization has been carried out using photoluminescence spectroscopy, absorption spectroscopy and optical microscopy.

The thesis is divided into several chapters as follows: chapter 2 gives an introduction to silicon carbide, its structure and properties. Chapter 3 describes the process of growing the material. Both growth of undoped and boron doped material is described here. In chapter 4 descriptions of the different characterization methods are given, together with a theoretical description of what the measurements can tell about material properties. Chapter 5 describes how the experiments were done and chapter 6 describes the results obtained from the experiments. The results are discussed in chapter 7. Chapters 8 and 9 discuss what has been learned about the material and how the work should be continued in the future.

2

An introduction to silicon carbide

This chapter describes the properties of SiC which are relevant to this thesis. Section 2.1 describes the atomic arrangement in the material, and some different arrangements are discussed. Section 2.2 discusses the energy band structure of 3C-SiC. Finally section 2.3 describes how doping is achieved in a material and some of its effects in 3C-SiC.

2.1 Crystal structure

Silicon carbide is a crystalline material consisting of silicon and carbon atoms. The crystalline nature of the material means that the atoms are arranged in a periodically repeating structure called a *lattice*. For given chemical elements there may be several different ways to arrange the atoms in a lattice, i.e. different chemical compounds of the same atomic species. This is called *polytypism*, where the different lattice structures are called the *polytypes* of the material. Silicon carbide has a large number of different polytypes - there are more than 250 known polytypes of SiC [11]. The different polytypes can be described as different stacking orders of layers of atoms [12]. Figure 2.1 shows one such layer. The depicted layer is the (111) surface. Each circle symbolizes one carbon and one silicon atom, displaced a small distance from each other. This pair of atoms is called the *base* of the crystal.

The marked hexagon in figure 2.1 marks an area of the crystal plane which can be used to define the different polytypes, where the polytypes are defined by the placement of the base atoms in this area in different layers. Figure 2.2 shows these stacking sequences for three of the most common polytypes. The depicted orders of the layers refer to one period in the periodic structure. The names for the different polytypes are stated at the top of the figure. The digit in the name refers to the number of layers of a period and the letter denotes the crystal symmetry. The letter *H* denotes the hexagonal polytypes, whereas *C* stands for the cubic polytype. Another

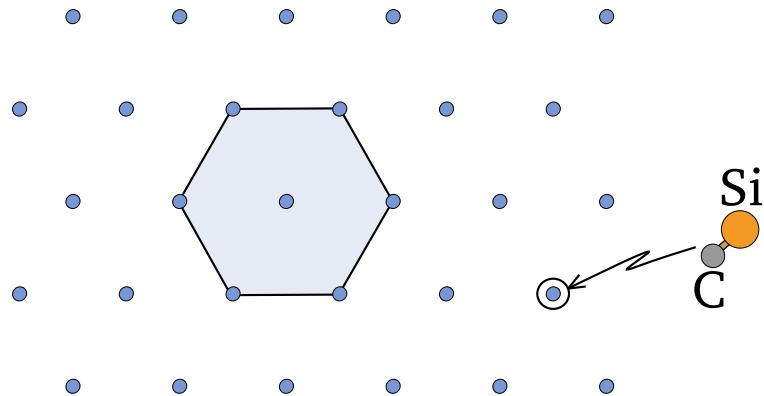


Figure 2.1: Atomic arrangement of the atoms in each (111) layer. Here each sphere corresponds to one carbon and one silicon atom, as shown by the arrow.

common polytype is the 6H-SiC, which thus is a hexagonal structure with a period of six base layers. It should be noted that the 2H-SiC structure is the wurtzite structure and the 3C-SiC is the zincblende structure. The unit cell of the zincblende, or 3C-SiC, crystal has a lattice constant of $a = 4.35 \text{ \AA}$ [13].

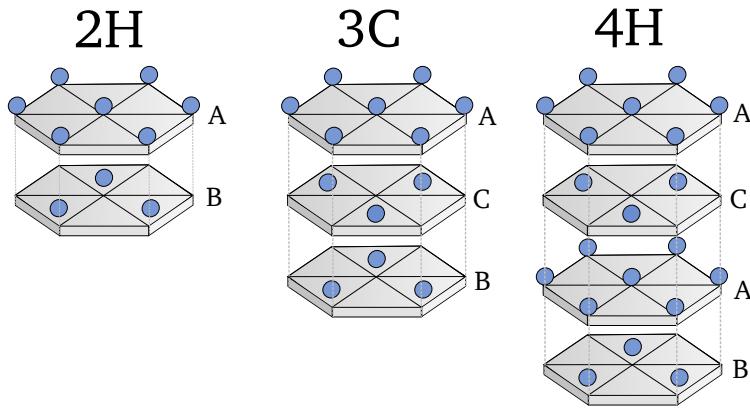


Figure 2.2: Stacking order for the three most simple polytypes.

2.2 Band structure

In SiC the atoms are bonded together with covalent bonds into a crystal. When the atoms are bonded together, the energy levels for the electrons in the material are defined by the material and lattice structure. This means that all materials have characteristic energy levels. When many atoms are bound together in a periodically

repeating manner, as is the case in crystals, the discrete energy levels for the electrons in the atoms merge together to form continuous bands which are called energy bands.

The band structure for 3C-SiC is shown as a band diagram in figure 2.3. In this figure, each curve describes allowed energies for the electrons, and the spaces in between the curves are not allowed. The x-axis shows different points in reciprocal space, or k -space. The marked points are specific points in the first Brillouin zone of k -space, and the intermediate intervals form straight lines between the points. This figure has been obtained using the approximate method of pseudopotentials, as described in [14]. The zero point energy has been chosen to be at the Γ -point in k -space.

The grayed area in the figure is the *band gap* of the material, since there are no allowed electronic states in this area. The band below the band gap is called the valence band, and the band above is called the conduction band. For 3C-SiC, as is characteristic for semiconductors, the Fermi level is in the middle of the band gap. This means that at the temperature of 0 K all electrons will be in the valence band, and the conduction band will be unoccupied. When the conduction band is unoccupied the material will not be able to conduct electricity. If the temperature is raised above 0 K there will be some occupation of the conduction band and some electrical conduction will be possible. With a higher temperature there will be more electrons in the conduction band and thus a higher conductivity of the material. One way to change the position of the Fermi level is to introduce impurities into the crystal, i.e. doping. This will be discussed later.

In the figure 2.3 it can be seen that the smallest energy difference between the valence and conduction bands is between the Γ -point in the valence band and the X-point in the conduction band. This difference is called the *band gap energy*. The difference between the Γ - and X-points is called the indirect band gap, while the difference between the values at the Γ -point is called the direct band gap, as illustrated in the figure. These values are given in table 2.1. The value for the room temperature band gaps are taken from the simulation, whereas the 2 K value is taken from literature.

Table 2.1: Values for the indirect and direct band gap energies. The indirect values are given both for room temperature and low temperature, as this will be of value later in the thesis.

$E_{g,\text{indirect}}$ (300 K)	$E_{g,\text{indirect}}$ (2 K)	$E_{g,\text{direct}}$ (300 K)
2.36 eV	2.42 eV [13]	6.00 eV

When an electron makes the indirect transition from the valence to the conduction band the energy is increased by $E_{g,\text{indirect}}$, and the k -value is changed, as indicated by Δk in figure 2.3. This change of k -value implies a change in momentum of the electron, since k and momentum p is related by

$$p = \hbar k,$$

where \hbar is Dirac's constant. This is of importance when studying interaction between the material and light. Photons are capable of providing the energy needed to transit

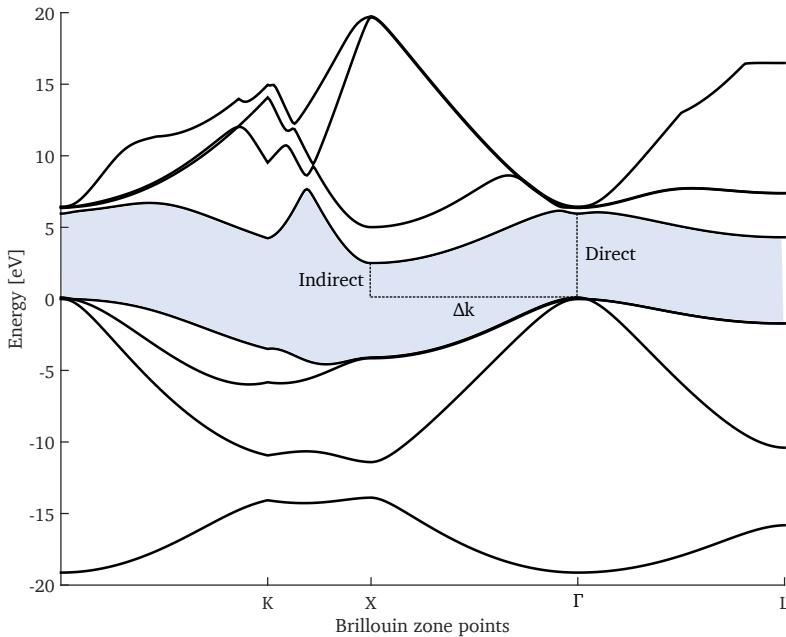


Figure 2.3: The band diagram of SiC at ambient conditions for different positions in the Brillouin zone. The indirect and direct band gaps are marked in the figure. The marked area is the band gap. The energy levels have been obtained using the pseudopotential method, as described in [14].

from the valence to the conduction band, but cannot provide the change of momentum needed. The law of conservation of momentum requires the total momentum of the system to be preserved during the transitions. The extra momentum can come from interaction with *phonons*, which are vibrations in the crystal.

The phonon energy spectrum has four ground modes since there are two different kinds of atoms in the material. There are the acoustic and optical phonon modes. Theoretical calculations of the phonon dispersion curves for 3C-SiC have been done by Karch et al. [15]. They show the four modes: TA, LA, TO, LO, and give the wavelength for each mode. The transition in k-space for the indirect transition is between the Γ - and the X-points, hence it is of interest to know which phonon wavelengths and energies this corresponds to. Table 2.2 shows these computed wavelengths and the corresponding energies, which have been calculated using the fact that

$$E_{\text{meV}} = \frac{hc(\frac{1}{\lambda_{\text{cm}}})}{e} \times 10^5,$$

where h is Planck's constant, c is the speed of light and e is the elementary charge.

Table 2.2: Inverted wavelengths and energies for the Γ -X transition. Values for λ from [15].

Mode	λ^{-1} [cm $^{-1}$]	E [meV]
TA	368	45.6
LA	637	79.0
TO	760	94.3
LO	829	102.9

2.3 Doping in 3C-SiC

Doping is a way to change the electronic structure of a material by substituting some of the native atoms with a foreign element. In the case of SiC this means that either silicon or carbon is replaced by some other element. The band structure of the material changes depending on which impurity is introduced to the crystal. By choosing the dopants it is possible to tailor the band structure to create various new properties of the material.

An important factor for how the band structure is changed is the number of valence electrons in the introduced element, compared to the element it replaces. Silicon and carbon both have four valence electrons, creating four bonds to its neighbouring atoms in the crystal. The electrons are strongly bound to the atomic nuclei, requiring energy corresponding to the band gap to free one electron. If one silicon or one carbon atom is replaced by an element with five valence electrons, the additional electron is not as strongly bound to the nucleus and can easily release an electron to the conduction band. This is called a *donor* atom. Similarly if an atom with only three valence electrons is introduced to the crystal, it can readily bind an electron from the valence band, creating an electron deficiency, or *hole* in the valence band. This is called an *acceptor* atom. Figure 2.4 shows a simplified band diagram with a donor and an acceptor level. The donor level donates an electron to the conduction band when the electron is supplied with the energy E_D . The acceptor accepts an electron from the valence band when the electron is supplied with the energy E_A . The energy required to supply the material with one carrier from a dopant level is called the *binding energy* (or *ionization energy*) of the dopant.

As dopants are introduced in the material, the position of the Fermi level relative to the bands changes. The Fermi level resides in the middle of the band for undoped materials, but will move nearer to the conduction (valence) band as donors (acceptors) are introduced. This change in Fermi energy corresponds to a change in occupancy in the different levels.

2.3.1 Donors

Since silicon and carbon both have four valence electrons, a donor atom in SiC must have five or more electrons. One common donor is nitrogen, which has five valence electrons. The nitrogen atoms take the place of the carbon atoms in the lattice. This

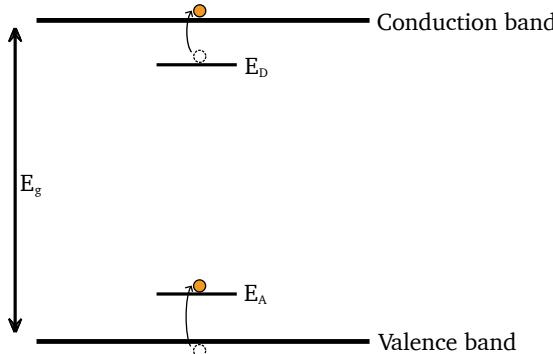


Figure 2.4: Two new energy levels are introduced by doping.

means that each added N-atom can supply the conduction band with one electron. A material with more electron carriers than the hole counterpart is said to be an *n-type* material. Freitas et al. have measured the binding energy of nitrogen in 3C-SiC, and found it to be 54 meV [16]. This means that the lowest N-level is 54 meV below the conduction band.

Nitrogen doping can be achieved intentionally by fabricating the material in a nitrogen atmosphere, but nitrogen is always present in sublimation grown material even at vacuum conditions [17]. This means that material grown by this method is generally n-type, if no other impurity is present. The n-type material can also be created by other elements in group five. Another possible donor in 3C-SiC is phosphorus, which has a binding energy of 48 meV [18]. Reports of doping with As, P and Sb exist, but are far less common than N-doping [19].

2.3.2 Acceptors

Acceptor atoms need to have fewer valence electrons than silicon and carbon, which is why group three elements are the most common acceptors. Examples of acceptor materials are boron and aluminium, which have the binding energies 257 meV [16] and 735 meV [7] respectively. A material with more hole carriers than the electron counterpart is said to be *p-type*. Table 2.3 summarizes the different dopant energy levels.

Table 2.3: Binding energies for some of the most common SiC dopants.

Element	Dopant type	$E_{D/A}$ [meV]	Reference
N	Donor	54	[16]
B	Acceptor	735	[7]
Al	Acceptor	257	[16]

The B-doping energy level is of particular interest for photovoltaic applications, where the binding energy is ideal for an application called *intermediate band photovoltaic solar cell*. In this application, the electrons can make two different energy transitions. One transition can be done between the valence band and the impurity level, and one between the impurity level and the conduction band. It has been shown that the binding energy of the B-acceptor in 3C-SiC is very well matched to the solar spectrum. Much of the emitted solar energy can be captured by the B-doped 3C-SiC [7, 8].

There are several ways to create doped materials. It can be done by ion implantation, where ions are accelerated to high energies and then made to collide with the undoped material [19]. It can also be done as the material is grown, by including the doping element in the ambient. The first method has several advantages: it can be done with good control over doping density and it is possible to select only certain areas of the material to dope. The latter method is done *in situ*, so it requires no additional equipment. It is also not as prone to create defects as the ion implantation method is, where annealing is often required after implantation in order to reduce the number of defects. In the work described in this thesis, the method of doping is to include the doping material during the growth. This method is described in more detail in chapter 3.

During doping some of the native atoms are replaced. This will have effects on the quality of the produced material. Different elements have different atomic radii, so replacing one atom by one of a different element will create some strain in the material. This may lead to defects in the material.

3

Growth technique

Silicon carbide can be fabricated by several different techniques, and the different techniques are advantageous for different polytypes and applications. Some of the most common methods are *sublimation epitaxy*, *liquid phase epitaxy (LPE)*, *chemical vapour deposition (CVD)* and *physical vapour transport (PVT)* [20].

The goal when growing free-standing material is to have high quality material of large enough thickness to allow the substrate and transition layers to be polished away. Up to this point in time, researchers have had more success fabricating high quality 4H-SiC and 6H-SiC material compared to 3C-SiC [3]. The hexagonal polytypes are commonly fabricated using PVT. This method has not been widely adapted to 3C-SiC growth, which is often attributed to the fact that this method uses homoepitaxy, and there are no 3C-SiC seeds widely available. Nishino et al. reported in 1983 a growth rate of approximately 2.5 microns per hour with the CVD technique for 3C-SiC growth [21]. This rate is much too low to create free standing material. The growth rate for CVD growth has been improved since that time, and in 2002 Nagasawa et al. reported a rate of $40 \mu\text{m/h}$ [22], which makes it possible to fabricate free standing 3C-SiC. This growth rate is achieved with silicon as the substrate, which gives rise to a large number of defects. This is due to the lattice mismatch and difference in thermal expansion coefficient between Si and 3C-SiC.

Another growth method is sublimation epitaxy. In this method material is transferred between a source material and a substrate using sublimation. In 1955, Lely demonstrated single crystal growth of SiC using sublimation [23]. This method was later modified by Tairov and Tsvetkov [24]. A more recent method called the sublimation sandwich method was demonstrated by Vodakov et al. [25] and Segal et al. [26]. Recently there have been reports of high quality cubic SiC fabricated by sublimation growth. Growth rates as high as $500 \mu\text{m/h}$ have been reported using this method, while still growing good quality material [4]. This high growth rate is obviously ideal for growth of free-standing (and even bulk) material. This chapter describes in de-

tail the sublimation method in general and more details about the setup used in this work.

3.1 About sublimation and nucleation

Sublimation is the phase transition where a material transitions between solid and gas, without the intermediate liquid phase. At normal conditions SiC does not melt - it sublimes. For SiC to melt a pressure of 100,000 atm and a temperature of at least 3200 °C is required - at atmospheric pressure and below, SiC can only sublime [27]. Sublimation of a SiC source material will give a vapour consisting of different molecules containing silicon and carbon, for example Si, Si₂C, SiC₂ etc. The composition and vapour pressure of the gas can be controlled using parameters such as temperature and pressure of the ambient.

If a substrate is placed in the silicon-carbon vapour, some of the material will be adsorbed on the surface of the substrate, in the form of *adatoms*. When many such atoms are adsorbed at the surface they bind together, initiating a *nucleus* of crystal growth. In a sublimation growth setup, a source of SiC is heated, initiating the sublimation. The vapour then travels to an appropriate substrate by the means of a heat gradient. Material is adsorbed on the substrate which forms the grown crystal. Figure 3.1 shows a schematic of this process. The vapour is transported from the source to the substrate along the temperature gradient. The higher temperature is at the source.

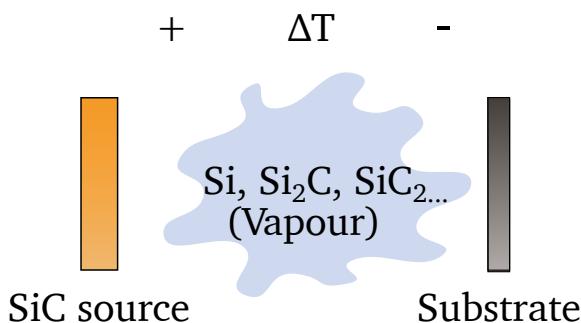


Figure 3.1: A vapour containing silicon and carbon molecules is created by sublimation of a SiC source. The vapour is adsorbed at the substrate.

As the growth begins, the material is adsorbed at the substrate. The crystal growth is governed by the free energy of the whole system. A decrease in free energy works to promote the growth. In crystal growth, a decrease in free energy is an increase in chemical potential μ , meaning that crystal growth is driven by an increase in chemical potential. The system is made up by the vapour and the crystal phases, hence a change in chemical potential is given by

$$\Delta\mu = \mu_v - \mu_c,$$

where μ_v is the chemical potential for the vapour and μ_c is for the crystal. For growth from vapour phase, the chemical potential difference between vapour and crystal is dependent on the *partial vapour pressure* and the *equilibrium vapour pressure* through the formula

$$\Delta\mu = k_B T \log(p/p_e),$$

where p is the partial vapour pressure and p_e is the equilibrium (or saturated) counterpart. Defining the *supersaturation*, σ , as

$$\sigma = \frac{p - p_e}{p_e}$$

thus

$$\Delta\mu = k_B T \log(\sigma + 1).$$

Taking the first order Taylor expansion this becomes

$$\Delta\mu \approx k_B T \sigma.$$

From this it can be concluded that the change in chemical potential is directly proportional to the supersaturation of the system. An increase in supersaturation will hence promote growth and increase growth rate.

Supersaturation is also a measure of how the crystal is formed. Silicon carbide can grow in several ways. One such way is called *spiral growth*. In this growth method there are steps which originate from one point on the surface. As atoms attach to these steps, they move around the center point. This gives a spiral shape on the surface. Another way of growth is *2D nucleation*. This is when atoms are adsorbed to a surface, and the adatoms bind to each other creating a 2D nucleus on the surface. Additional atoms will attach to this nucleus as they are adsorbed to the surface. The value of supersaturation governs which of these two methods is dominating.

Figure 3.2 shows the general relationship between the growth rate (R) and the supersaturation (σ). It should be noted that at certain values of σ the growth rate of spiral growth is higher than that of 2D growth. At other σ values the 2D growth will dominate over spiral growth. At very high values of supersaturation the two growth modes will have similar growth rates.

The growth model for 3C-SiC on off-axis samples is described in detail by Jokubavicius et al. [4]. The model can be divided into three parts. Initially a large terrace (or facet) region is formed at one edge of the substrate. The terrace is grown homoepitaxially. On this terrace small areas of 3C-SiC are formed, and will eventually merge. Lastly the 3C-SiC grows from the terrace onto the rest of the substrate by step flow growth. In step flow growth, the steps inherent in off-axis samples are overgrown with 3C-SiC by enlargement of the steps. This is sometimes also called *lateral enlargement*.

3.2 The sublimation growth setup

The sublimation growth process used in this work uses a setup as shown in figure 3.3. An insulating graphite foam is placed inside a quartz tube which is surrounded

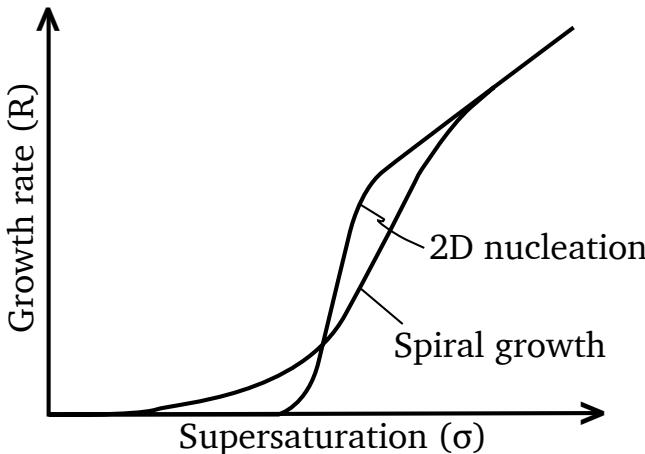


Figure 3.2: The growth rate for different growth modes vary depending on supersaturation. The figure is adapted from [28].

by a copper coil (shown as a cutaway in the figure). The copper coil is connected to an RF-generator capable of creating a current in the coil. The foam cylinder contains a graphite crucible which will couple with the magnetic field created by the current in the copper coil. This will heat the crucible. To measure the temperature of the crucible there is a pyrometer fixed above the quartz tube (not shown in the figure), which can measure the temperature of the crucible through a small hole at the top of the insulating foam.

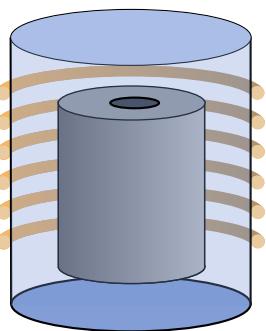


Figure 3.3: An outside view of the reactor setup of the sublimation growth method. A copper coil surrounds a quartz tube containing an insulating graphite foam.



Figure 3.4: The setup inside the crucible. (1): Polycrystalline source material, (2): Graphite spacer, (3): 4H-SiC substrate.

The graphite crucible contains the components of the sublimation process, as shown in figure 3.4. At the bottom of the crucible is a source material which will sub-

lime away during the growth. When growing free-standing 3C-SiC, a polycrystalline SiC wafer piece is used as source material. The source and the substrate are separated by a graphite spacer which is a disc with a hole in it where the vapour can flow freely. The size of the spacer and the hole can be varied depending on the desired size of the grown sample. The substrate can consist of different materials, generally 4H- or 6H-SiC. Wafers of both 4H- and 6H-SiC are commercially available and can thus readily be used as substrates in growth. When the wafers are cut, they can be cut at different angles, i.e. in different crystal planes. This is what is usually referred to as on- and off-axis substrates. The 4H-SiC substrates are typically cut perpendicular to the (0001)-direction for on-axis substrates. For off-axis substrates a small angle from this direction is used.

Growing on on-axis substrates will give many nucleation sites. As these sites expand and merge with each other there will be many defects at the interfaces. These defects are called *double positioning boundaries (DPB)*. The off-axis substrates have a surface which contains many shallow steps. As the 3C-SiC has nucleated at the edge of the growth area, the nucleation site will expand laterally along the step-flow direction. This growth method is called *lateral enlargement mechanism*, and is the main principle of 3C-SiC sublimation growth after nucleation has occurred. The density of DPBs is greatly decreased in this manner. This process is described in detail by Jokubavicius et al. [4].

As described earlier in this chapter, the growth temperature and pressure are vital parameters to determine growth rate and also crystal quality. In sublimation growth these parameters can be controlled and changed during the growth process. The temperature is controlled by modulating the current from the RF-generator. A temperature ramp up scheme is used to facilitate proper step flow growth. It has been shown that the temperature window for 3C-SiC growth is rather narrow [28], hence proper control of the temperature is vital for good quality material. The temperature ramp up scheme is done in two or three parts. First the temperature is increased by increasing the power from the RF-generator. This is done in intervals of several minutes per 100 W. As the desired temperature is reached, it is held for the desired amount of time. For some samples a third step is added, where after growth for some time at the first temperature level, the temperature is again increased to a higher level and held there for some additional time. These schemes are sometimes called one and two step temperature schemes, respectively.

Another important parameter is the pressure in the reactor during growth. To avoid inclusions of foreign material in the crystal, a low pressure during growth is required. Typically a pressure of 10^{-4} to 10^{-5} mbar is used, which is achieved with constant pumping with a vacuum pump during growth. Some nitrogen inclusions will be unavoidable even at these pressures. A residual N-doping in the order of 10^{16} cm⁻³ has been shown in material grown under these conditions [4], which can be compared to 10^{22} cm⁻³ silicon or carbon atoms in the crystal.

For a good quality crystal to form, the proper ratio between silicon and carbon must be present in the vapour. To avoid the problem with excess of carbon, a *carbon getter material* is often used. This is a material which binds carbon from the vapour, thus reducing the C/Si ratio. One such material is tantalum. Tantalum will bind carbon during the growth, thus slowly converting to tantalum carbide. In the growth

setup used in the sublimation process, a piece of thin tantalum foil is often used.

3.2.1 Doping using sublimation growth

It is possible to grow intentionally doped material using the sublimation growth method. To do this, the pure SiC polycrystalline source is exchanged for a doped source with the appropriate dopant concentration. The vapour from the sublimation will then contain the dopant element. Figure 3.5 shows the crucible setup for growth of doped material. To facilitate cubic growth, the substrate contains a 3C-SiC seed before the growth starts. This is called *homoepitaxial growth*.

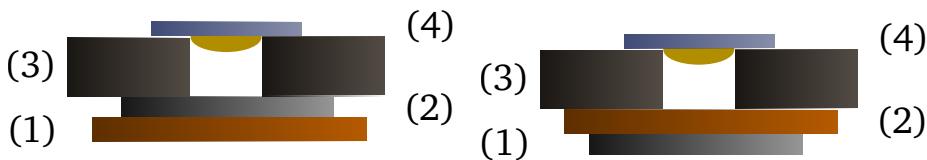


Figure 3.5: The crucible setup when doing doped growth. The source and seed (4) is separated by a graphite spacer (3) from the doped (2) and undoped (1) source material.

Figure 3.6: Indirect growth. Here the doped (1) and undoped (2) source materials have changed places as compared to direct growth.

It is vital to have the proper ratio between all of the elements in the vapour phase, i.e. silicon, carbon and dopant. In this thesis a new method to possibly alter this ratio is introduced. This method is called *indirect doping*, and is shown in figure 3.6. This differs from the direct doping shown in figure 3.5 in that the doped source is now separated from the substrate by a nominally undoped source layer. If the doped source material is not homogeneous, the indirect method is hoped to give a better grown material, as well as reduce the dopant content in the vapour. Direct doping may influence the structural quality of the grown material by a rough morphology.

4

Characterization techniques

To investigate the material properties after growth, the samples must be characterized. There are many established techniques to do this. This chapter gives a description of the characterization techniques used to obtain the results in this thesis. This thesis focuses partly on the morphology of the samples and partly on the optical properties of the grown samples. The specifics of the experiments performed are given in chapter 5.

4.1 Absorption measurements

Absorption measurements are performed to find how much light is absorbed or transmitted when it hits the sample. In this way the band structure can be found even for samples with little or no luminescence. During measurements, light of different wavelengths hits the sample, some of which is absorbed and some transmitted. The transmitted light passes through a spectrometer which measures the light intensity as a function of the wavelength. By measuring the light intensity before (I_0) and after (I) the sample, it is possible to compute the logarithm of the ratio of the two, which constitutes the *absorption* of the sample, i.e.

$$A(\lambda) = \log \left(\frac{I_0(\lambda)}{I(\lambda)} \right). \quad (4.1)$$

Generally some wavelengths are absorbed more than others. In semiconductors, light is absorbed to allow electrons to make a transition from a lower to a higher energy level. The band structure determines which transitions are allowed and which are not. Absorption of a certain wavelength means that there is an allowed transition with the same energy. The energy and wavelength relate through the formula

$$E = \frac{hc}{\lambda}, \quad (4.2)$$

where h is Planck's constant and c is the speed of light.

4.2 Photoluminescence spectroscopy

Photoluminescence (PL) spectroscopy is a very powerful tool in optical characterization of semiconductors. It is a non-invasive and fast technique. Photoluminescence spectroscopy is similar to absorption spectroscopy in that a light source is aimed at the sample to induce electron excitations. As an electron radiatively recombines with a hole, a photon is emitted. These emitted photons are called the *luminescence* of the sample, and is what is measured in PL spectroscopy.

The process of excitation and recombination is shown in figure 4.1, where a photon is captured (a) and generates an *electron hole pair* (*EHP*) (b). The EHP then recombines and emits a photon of energy corresponding to the band gap, E_g (c).

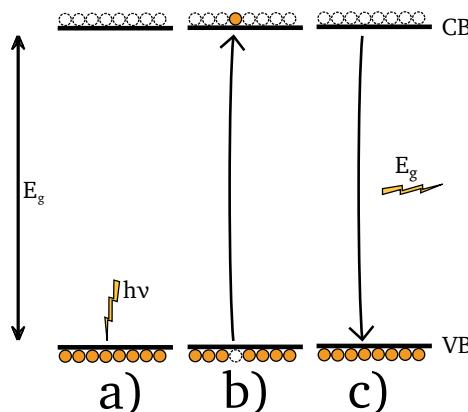


Figure 4.1: When a photon with energy $h\nu$ greater or equal to the band gap energy is captured (a), an EHP is generated (b). The EHP recombines to emit a photon of energy E_g (c).

To generate an EHP by a transition between the valence and conduction bands, the energy of the light must be greater than or equal to the band gap energy. This is in accordance with the principle of conservation of energy. The laser must be chosen accordingly. Not only the energy is conserved in this process, but as is the momentum. Since SiC is an indirect semiconductor (see figure 2.3), an electron moving across the band gap has a change in momentum. The light momentum is far too small to compensate for this change. To allow the EHP to be generated, the change in momentum is contributed by phonons. As the transition occurs, a phonon is either created or annihilated. The phonon energies for the different modes in 3C-SiC are given in section 2.2. The phonons show in PL spectra as lines with energies slightly smaller than the value of the actual electron energy transition, called the *zero phonon line* (*ZPL*). The phonon lines are called *phonon replicas*.

The phonon replica lines can be used to investigate the biaxial stress in the material. The PL intensity of the transverse optical and longitudinal acoustic modes

should be the same if the material is without internal stress. Hence the ratio

$$\frac{I_{LA}}{I_{TO}}$$

should be near unity in stress free samples [17]. In this ratio, I_{LA} and I_{TO} are PL intensities of arbitrary units. It is also possible to use the phonon lines to estimate the doping concentration of a dopant. Camassel et al. have introduced a method where the full width at half maximum (FWHM) value of the TA-peak is used for this purpose. They show that the concentration of nitrogen donors in the sample is proportional to the FWHM, Γ , through the formula

$$\Gamma_{TA} = A[N]^{1/n}. \quad (4.3)$$

Here A and n are constants and $[N]$ denotes the concentration on nitrogen donors. This method has not been validated for acceptor type impurities.

If doping levels are introduced in the band structure, then transitions can occur between such levels and the bands. In this case there will be another set of ZPL and phonon replicas at the new transition energy. The band structure can thus be inferred from the PL spectrum.

As the EHP are generated, the electron and hole will interact with each other through the Coulomb force if they are near enough to each other. Due to their charge difference, they will start to orbit each other. This is called the *exciton* quasiparticle. The energy of recombination of an exciton is slightly shifted from a normal recombination of a EHP, due to the potential energy from the Coulomb interaction. At room temperature the thermal energy from the ambient is generally high enough to split the exciton into a free electron and hole, so the excitons are not seen in the PL spectrum. At very low temperatures there is not enough energy to split the exciton, so the PL spectrum shows this shift in the recombination energy attributed to the Coulomb interaction. Some excitons can freely move in the sample, but some excitons bind to impurities in the sample. There is a small energy difference in the free and bound excitons attributed to the binding energy to the impurity. It is thus possible to distinguish from the free and bound excitons in the PL spectrum.

The PL setup can be described as follows. An optical setup is used to focus a laser beam on a sample. The emitted photoluminescence passes through another optical setup and is focused into a spectrometer. The spectrometer separates the incident light by wavelength, generally by the use of a grating. Photons with a given wavelength can now be collected by a photomultiplier tube or a CCD camera, and the photon intensity for the wavelength spectrum can be computed. Light is captured during a set time (generally ranging between seconds to minutes depending on the sample and setup).

The sample can be placed in a cryostat to measure the luminescence at low temperature, or measurements can be performed in room temperature. Various types of cryostats exist, generally using either liquid nitrogen or helium. In this work results have been obtained using room temperature measurements and by the use of a liquid helium cryostat.

4.3 Nomarski interference contrast microscopy

Nomarski interference contrast microscopy, also known as differential interference contrast microscopy, is a version of optical microscopes designed to be able to investigate the inside of samples. The method utilizes the polarization of light to give slightly different light paths depending on where in the sample the light passes. The setup is shown in figure 4.2.

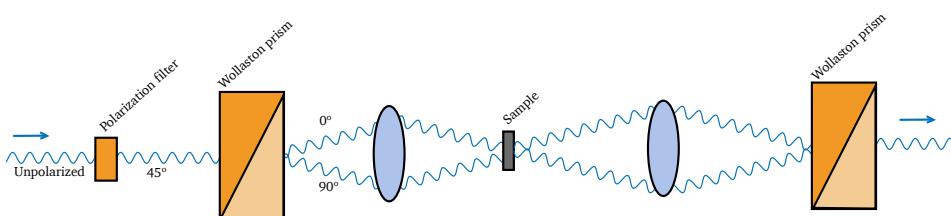


Figure 4.2: A schematic of the Nomarski microscope setup. The light passes through the setup from the left to the right.

The light is initially unpolarized. After passing a 45° polarization filter, the light enters a Wollaston prism. This prism will split the light in different directions depending on the polarization, creating two diverging beams. One beam will have light polarized in the 0° direction and the other in the 90° direction. These two beams are focused on the sample by a lens, but made to hit the sample at slightly different positions. This will give the light beams slightly different paths through the sample. As the beams leave the sample they will enter another Wollaston prism, which will merge the two beams into one, and give the beams the same polarization, 135° . If the beams have travelled different lengths in the sample, or if the sample composition has varied between the two paths, then the beams will be slightly out of phase. This phase difference will through interference between the beams lead to optical contrast.

One possible use for the Nomarski microscope is to detect different thickness in the sample. It can also be used to detect sample composition difference, if the compositions have different refractive indexes. An example would be if a 3C-SiC sample has some hexagonal inclusions, then these will give contrast in the microscope since the hexagonal and cubic polytypes have slightly different refractive indexes.

5

Experimental setup

In this chapter the details of the experimental setups are explained. The chapter is divided into two parts. The first part describes how the samples were grown and prepared. The second part describes details on the experimental setups for the characterization experiments.

5.1 Growth and sample preparation

All samples used in this work are grown using sublimation growth, as described in section 3.2. The nominally undoped samples are grown using two layers of polycrystalline SiC source material. The 4H-SiC substrates with 4° off axis surface were used for undoped growth. The samples were cut along the (0001)-plane, with an off-axis angle towards the <11̄20> direction. The substrates were chemically cleaned before growth. Cleaning was done with acetone and ethanol followed by H₂O:NH₃:H₂O₂ combined in fractions of (5:1:1) and H₂O:HCl:H₂O₂ combined as (6:1:1). Both the carbon and the silicon face of the 4H-SiC substrates were used during growth of undoped samples.

The doped samples were grown using either direct or indirect doping methods, as described in section 3.2.1. The source material for doped growth was doped polycrystalline SiC, doped with boron concentrations in the order of 10¹⁸, 10¹⁹ or 10²⁰ cm⁻³. This was used together with a piece of undoped source material. The doped samples were grown either homoepitaxially on an unintentionally doped seed, or heteroepitaxially on the silicon face of a 4H-SiC 4° off-axis substrate.

A piece of tantalum foil was placed in the bottom of the crucible for use as a carbon getter during growth of all samples. The source and substrate or seed were separated by a 1 mm thick graphite spacer for all samples. The substrate was held in place on the spacer by a graphite plate placed on top. All samples were grown at pressures in the order of 10⁻⁴ to 10⁻⁵ mbar. The pressure is varying slightly during the

growth process due to changes in the temperature. A vacuum pump was connected to the reactor and running continuously during growth. The temperature and growth time were varied between samples, and are given for individual samples in section 6. The temperature was changed by varying the power supplied by the RF generator to the copper coil. The reactor setup is described further in section 3.2. Care was taken to place the insulating foam containing the crucible in the same position in the reactor for each run.

After growth the reactor was cooled in vacuum, by instantly setting the RF-generator to 0 W when the growth time was over. Cooling time was reduced by the use of a fan placed outside of the reactor. Optical microscopy images were done on as-grown samples, but any other characterization was done after chemically cleaning the samples. The cleaning was performed as described for the substrates.

5.2 Characterization experiments

The absorption measurements were done with a PerkinElmer Lambda 950 UV/VIS setup. The software used was PerkinElmer UV WINLAB for molecular spectroscopy. Measurements were done in the range between 2000 to 350 nm. Samples were mounted on a black cardboard piece with a hole for light to pass through. Two hole sizes were used: 3 mm and 4 mm in diameter. Most samples were measured after mechanical polishing of the substrate, i.e. free-standing, but some samples were measured with the substrate remaining. The wavelength step length of the spectrometer was 5 nm.

Low temperature photoluminescence measurements were performed using a liquid helium cryostat. The helium was supercooled to 2 K using a pump. A laser of wavelength 351 nm was used. The laser used was an Ar-ion laser, which had a power output of approximately 2-3 mW. The focused laser spot on the sample surface was approximately 100 μm . The spectrometer used had a Jobin-Yvon HR 460 monochromator, and the luminescence was detected by a CCD camera.

Room temperature luminescence measurements were performed. A laser with wavelength 405 nm was used. The spectrometer was of model Ocean Optics USB2000+, which uses a Sony ILX511B CCD camera to detect the photons.

6

Results

In this chapter, the results of the experiments are shown. The chapter is divided into two parts. The first part shows the results for unintentionally doped seed growth, whereas the second part shows the results of the B-doped samples. The results are further discussed in chapter 7.

6.1 Seed growth

This section contains the results of characterization of unintentionally doped 3C-SiC seeds. To investigate growth of seeds to be used for B-doping, seeds were grown on both carbon and silicon face of the 4H-SiC. Table 6.1 shows the growth conditions and properties of the seeds grown on carbon face. The sample thickness is excluding the substrate thickness. Terrace coverage indicates the percentage of the sample terrace (or facet) covered by 3C-SiC rather than any other polytype. These percentages were estimated from Nomarski and reflection mode micrographs. Not all sample coverages were measured quantitatively, but the percentages in the table are thought to be representative. The silicon face seeds were all grown with a two step temperature ramp up scheme, with temperatures 1850/1925 °C during times of 0:30/1:00 (hours/minutes). Nine out of ten of the silicon face grown seeds had a terrace coverage of 100 %, whereas none of the carbon face samples had complete terrace coverage.

Figures 6.1 and 6.2 show micrographs of the terrace of carbon face grown sample S5. Figure 6.1 shows a reflection mode micrograph. It can clearly be seen that there are a few large areas of different polytypes. The hexagonal polytype regions show spiral growth, which on carbon face 4H-SiC manifests as a star shape. The 3C-SiC regions look smoother, and show some of the characteristic triangle shapes. Figure 6.2 shows a Nomarski mode micrograph which shows a clear contrast between the different polytype regions. These figures are representative of all carbon face

Table 6.1: Growth conditions of seeds of undoped 3C-SiC grown on the carbon face of the substrate. For samples grown using a two stage temperature ramp up, the temperatures and growth times for each step are both shown. The *R*-column indicates the growth rate for the one or two stages.

	Temperature [°C]	Time [h:min]	Thickness [μm]	R [μm/h]	Terrace coverage [%]
S1	1825	4:00	700	175	40
S2	1850/1925	0:30/1:00	800	533	20
S3	1850/1925	0:30/1:00	800	533	
S4	1850/1925	0:30/1:00	900	600	30
S5	1875/1925	0:15/1:00	800	640	50
S6	1925	1:00	700	700	20
S7	1925	1:00	700	700	

grown seeds. In all such samples there are distinctly separate regions of cubic and hexagonal SiC.

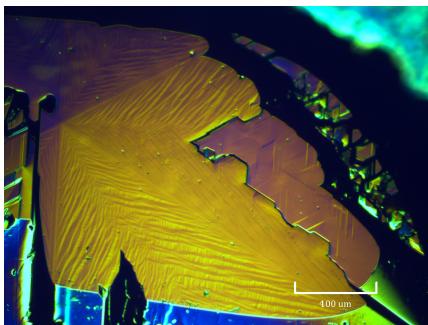


Figure 6.1: Reflection mode micrograph of the facet of C-face seed S5.

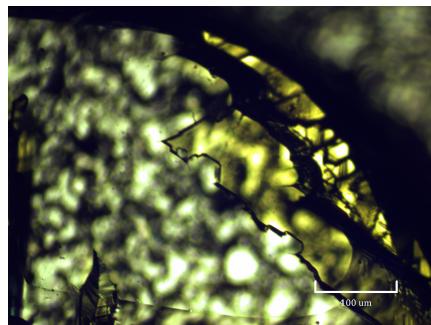


Figure 6.2: Nomarski transmission mode micrograph of the facet of C-face seed S5.

Figure 6.3 shows a reflection mode micrograph of a silicon face grown seed. On this sample the whole terrace is covered with the 3C-SiC polytype, as can be seen from the characteristic triangle shaped features visible on the surface. This micrograph is representative of the Si-face grown seeds, most of which had complete terrace coverage.

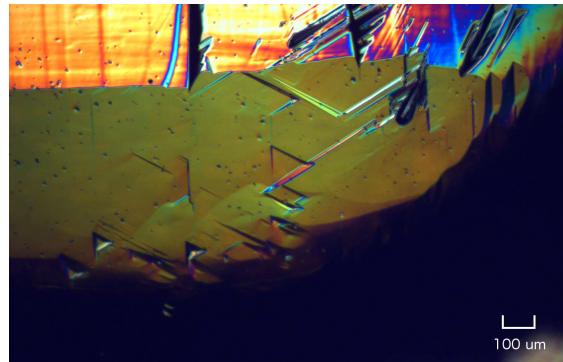


Figure 6.3: Reflection mode micrograph of the facet of a Si-face seed.

6.2 B-doped samples

This section shows the results of characterization of B-doped samples. Table 6.2 shows the growth conditions and sample thicknesses of the grown B-doped samples. The thickness is measured for the doped layer, excluding both the substrate and the undoped seed. The growth mode is the stacking order of the doped and undoped source materials in the crucible, as described in section 3.2.1. All samples except D9 were grown on a 3C-seed. Sample D9 was grown directly on a 4H-SiC 4° off-axis substrate (Si-face). All seeds except for D6 are grown on Si-face substrates.

Table 6.2: Growth conditions of the doped samples are shown in this table. The *R*-column indicates the growth rate of the sample. The growth mode describes which samples are grown using direct or indirect doping methods. These methods are described in section 3.2.1. The doping concentration is given for the source material. The seed column indicates whether the B-doped sample is grown on a 3C-SiC seed or on a 4H-SiC substrate. Sample U1 is an undoped reference sample.

	Temp. [°C]	Time [h:min]	Thickness [μm]	R [μm/h]	Doping [cm ⁻³]	Mode
D1					10^{18}	Direct
D2	1825	2:30	350	140	10^{18}	Direct
D3	1825	3:00	360	120	10^{18}	Direct
D4			440		10^{19}	Direct
D5	1825	2:30	320	128	10^{19}	Direct
D6					10^{20}	Direct
D7	1825	2:30	220	88	10^{20}	Direct
D8	1825	2:30	380	152	10^{18}	Indirect
D9	1825	2:30	380	152	10^{20}	Indirect
U1	1850/1925	0:30/1:00				Undoped

Figure 6.4 shows the surface morphology of an undoped sample (U1) (a), doped sample number D1 (b), doped sample number D4 (c) and doped sample number D6

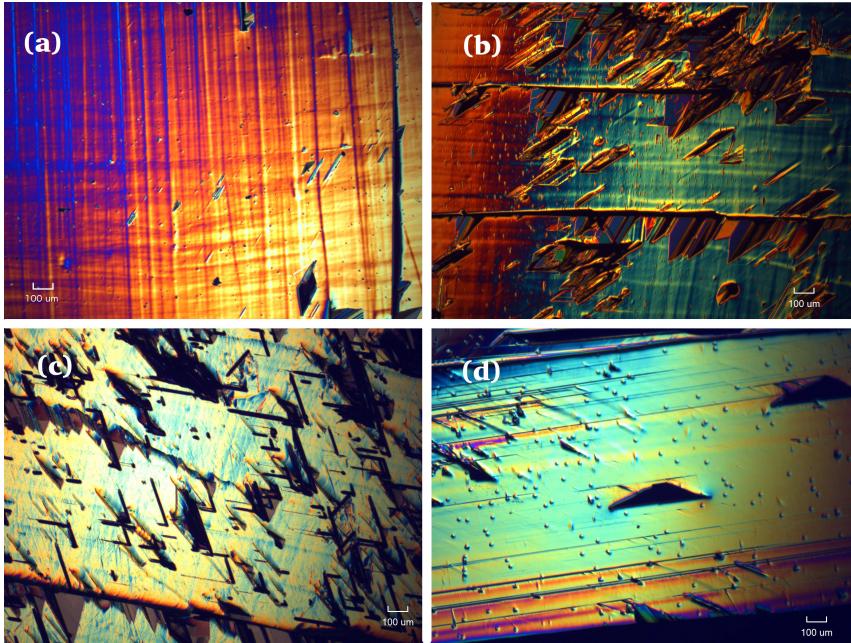


Figure 6.4: Micrographs of an undoped sample (U1) (a), doped sample number D1 (b), doped sample number D4 (c) and doped sample number D6 (d).

(d). It is clearly seen that the undoped sample has a better surface quality compared to (b) and (c), but (d) which is grown with a highly doped source shows good surface quality. Sample D7, grown under similar conditions as D6, shows a much worse surface, comparable with that in (c). The micrographs are representative for the whole surface morphology of the samples. It should be noted that the samples (b)-(d) are grown using a source material doped with dopant concentration of 10^{18} , 10^{19} and 10^{20} cm^{-3} respectively.

Figure 6.5 shows reflection mode micrographs of samples D1 (a) and D8 (b) together with a transmission mode micrograph of sample D8 (c). The two samples are grown with direct and indirect doping methods respectively. It can be seen that the sample surface quality is similar in the two samples. Figure (c) shows no foreign polytype inclusions.

Figure 6.6 shows reflection mode micrographs of samples D6 (a) and D9 (b). These samples are grown with direct and indirect growth methods respectively. Both samples show comparable surface morphology. In (c) a transmission mode micrograph of D9 is shown. It can be seen that there are two different polytypes in the sample. The brighter area in the center is a foreign polytype inclusion.

Absorption measurements were done on directly doped samples with different source doping concentrations, samples D1 and D5. The results of these measurements are shown in figure 6.7 (a), together with a reference spectrum from an undoped sample (dotted line). In both samples we see the band edge absorption at around

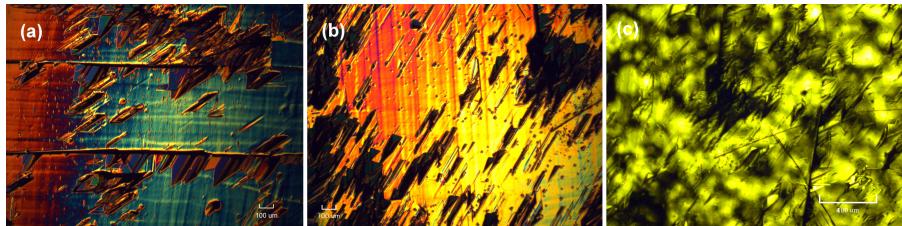


Figure 6.5: Reflection mode micrographs of samples D1 (a) and D8 (b). Transmission mode micrograph of D8 (c).

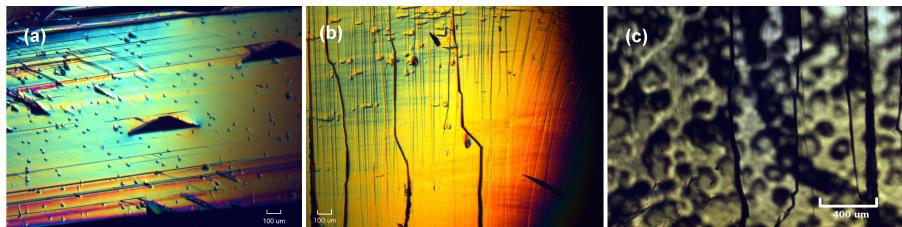


Figure 6.6: Reflection micrographs of samples D6 (a) and D9 (b), and transmission micrograph of D9 (c).

500 nm. A second feature in the doped sample spectra is the peak at around 700 nm, which may be corresponding to the transition between the boron level and CB. This cannot be seen in the undoped sample. The doped samples show no evidence for a VB to boron transition, which should be seen at around 1800 nm, as described in chapter 2.3. Absorption was also measured for doped samples D6 and D7, but neither the band edge nor the boron related peaks were seen in these highly doped samples. The dip and the sharp peak at around 900 nm for sample D5 is attributed to the equipment. The peak is thought to be an artefact due to background subtraction.

Figure 6.7 (b) shows the result of absorption measurements on samples D1 and D8 together with an undoped sample. It should be noted that D8 is an indirectly doped sample. It can be seen that both the doped sample spectra show the same trend. The graphs increase up to a point near the 700 nm B-CB transition energy, where they level out. This is not visible in the undoped sample. It can further be seen that sample D8 has its B-CB maximum at a slightly shorter wavelength compared to sample D1, which is attributed to the fact that measurements were done on D8 before polishing. The peak shift is in fact a substrate artefact. Absorption measurements were also done on sample D9, which did not show the B-CB peak.

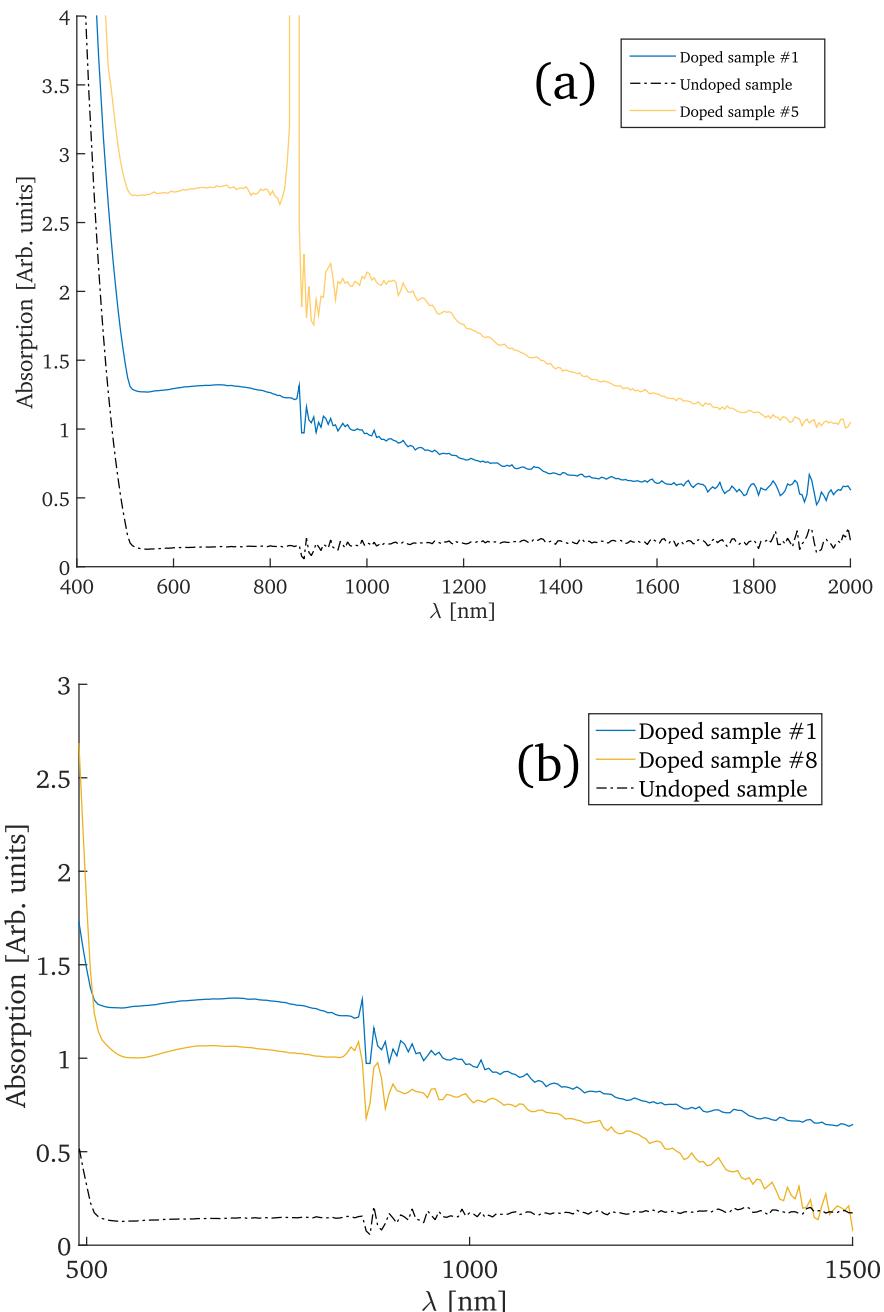


Figure 6.7: (a): Absorption spectrum of undoped sample (dotted line), doped sample D5 (bright yellow line) and doped sample D1 (dark blue line).

(b): Absorption spectrum of doped samples D8 (bright yellow line) and D1 (dark blue line), shown together with an undoped sample (dotted line).

The LTPL-measurements were done on boron doped samples and on an undoped seed. Figure 6.8 shows the resulting spectrum of the undoped seed. Five peaks can clearly be seen in the spectrum, labeled as ZPL and phonon replicas. It should be noted that the ZPL is significantly smaller than the phonon replicas. It can further be seen that several smaller peaks appear on the low energy side of the phonon replicas. These lines originate from exciton complexes.

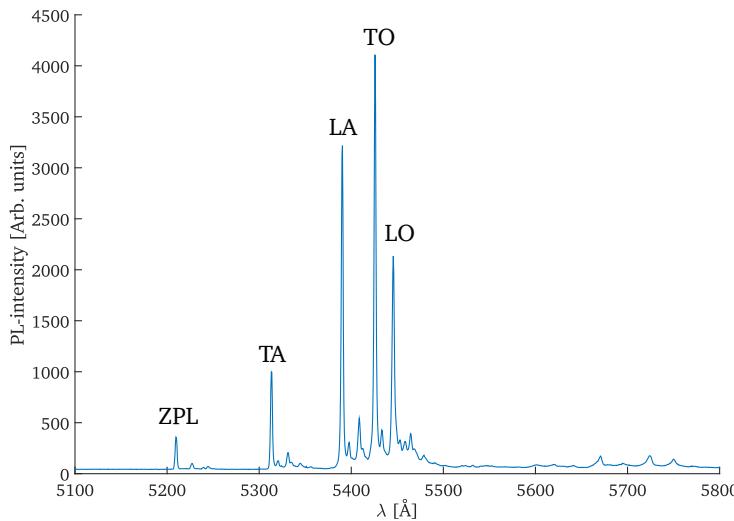


Figure 6.8: LTPL spectrum of an undoped seed. The five main peaks are labeled as ZPL and phonon replicas.

As described in chapter 4.2, the strain in a crystal can be estimated by the fraction between the LA and TO lines. From figure 6.8 this is found to be

$$I_{\text{LA}}/I_{\text{TO}} = 0.783.$$

Again as described in the same chapter, the donor concentration can be found from the full width at half maximum (FWHM) of the TA-peak. The FWHM is found to be $\Gamma_{\text{TA}} = 1.094 \text{ meV}$, giving a nitrogen concentration of

$$[N] \approx 10^{16} \text{ cm}^{-3}.$$

This value was obtained by using the same parameters as proposed by Camassel et al. [29].

Doped samples D1, D4 and D6 were measured using LTPL. Not one of these samples show any luminescence, neither from the boron related energy level, nor from the band to band transition. Sample D9 did show some luminescence. It was the only B-doped sample to do so. The sample showed a peak at 5946 Å. At some points it also showed a peak at 4750 Å. Room temperature PL-spectroscopy was performed on sample D9, but no luminescence was found.

7

Discussion

This chapter discusses the results presented in chapter 6. The chapter is divided into two parts, one part discusses the results of seed growth and the other part discusses the results from characterization of the B-doped samples.

7.1 Seed growth

As shown in section 6.1, the carbon face grown seeds did not have complete coverage of cubic SiC on the terrace, yet the silicon face grown seeds could reproducibly be grown with complete cubic coverage. As seen in figure 3.2, 2D growth is promoted by high supersaturation. A possible explanation of the results is that the supersaturation for the C-face growth was too low. The examined temperatures, as shown in table 6.1 range from 1825 to 1925 °C, which is a large portion of the parameter space in which the 3C-SiC polytype is stable. Changing the temperature much more during growth is unlikely to result in better 3C-coverage, since this will likely be outside the region where 3C-SiC can be formed in a stable manner. Because of this, the supersaturation will not be able to be changed much more by altering the temperature.

From table 6.1 it can be seen that the percentage of the terrace covered by 3C-SiC does not show any trend with regards to temperature. The growth time has been adjusted to get similar sample thickness for all samples. In this way it was possible to vary only the supersaturation between samples. This indicates that for the used growth setup it is not possible to change the supersaturation enough to get complete 3C-SiC coverage of the terrace. However, the pressure conditions for growth have not been explored in this work, and it may be possible to alter the ambient pressure to get more cubic coverage.

A possible explanation of the difference in 3C-SiC coverage on silicon and carbon face substrates is the spiral growth character. On Si-face the spirals take the form of proper spirals, with steps expanding from the spiral arms. The C-face spiral

takes the form of straight lines extending radially from the center, as seen in figure 6.1. The 3C-SiC polytype nucleates on top of the heteroepitaxially grown spirals and grows over the spiral steps. The spiral shape can then be expected to influence how the 3C-SiC is grown. The silicon face spiral may be advantageous for cubic growth, while the carbon face spiral may not be. Another possible explanation is that there may be some difference between how the Si- and C-surfaces look for off-axis cut samples. There may be a difference in the number and density of the steps, which in turn will influence the terrace growth. To get a proper understanding of this, electron microscopy would have to be done on the substrate surface and the terrace after initial nucleation. In this way a proper understanding of the growth mechanisms governing C- and Si-face growth would be possible. This is beyond the scope of this work, but could be included in future work.

The silicon and carbon faces have different surface energies which will influence how material grows on it. Stein et al. have shown that SiC growth on 4H-SiC and 6H-SiC substrates will result in different polytypes grown depending on the face rather than the substrate polytype. This they attributed to the different surface energies of the different faces [30]. This may also be a factor in the problem of growing 3C-SiC on the carbon face. From the micrographs of the carbon face seeds it can be seen that cubic growth always occurs at the edge of the facet, i.e. near the graphite spacer. This is consistent with the hypothesis that the surface energy gives rise to the results since the edge is expected to have a different surface energy compared to somewhere towards the center of the sample.

7.2 B-doped samples

From figure 6.4 it can clearly be seen that B-doping deteriorates the crystal quality. All doped samples have a larger number of defects compared to the nominally undoped sample. For samples D1 and D4 the number of defects increase with increasing source doping. This is to be expected since the boron atoms have a different radius compared to both the silicon and carbon atoms they replace. This size difference will induce strain in the sample and in turn induce defects. From the same figure we can see that sample D6, which was grown with the source material with the highest doping concentration, has a smoother surface compared to the two other doped samples characterized in this thesis. This does not follow the trend. In figure 6.6 (b) can be seen that sample D9, grown from the same source material as D6, shows similarly good surface quality. In contrast sample D8, again grown from the same source material, has a much poorer surface. From the absorption measurements on these samples it can be inferred that these materials do not in fact have very good quality. As described in chapter 6, samples D8 and D9 do not show any B-CB absorption peak, and D6 does not show the band edge absorption. One possible explanation of this is that the source material does not in fact have the high doping concentration it is thought to have. This hypothesis is contradicted by the result of absorption measurement on D6. If the samples would be only lightly doped, then the band edge absorption would still be visible. Another possible explanation is that even though the surface shows a low density of defects, the sample may be of poor crystal quality. This would explain

why sample D6 does not show any band edge absorption. Further characterization of the crystal quality of these samples would give more insight to why the surfaces seem smooth but the absorption measurements yield poor results.

Figure 6.5 shows the difference in surface morphology of doped samples D1 and D8, which have been grown using the direct and indirect growth methods, respectively. The fact that they both show similar numbers of surface defects could indicate that both samples have been doped with comparable concentrations of boron. As shown earlier a higher concentration of boron will induce more defects on the surface. Figure 6.7 (b) gives a further indication that this hypothesis is valid. This figure indicates that the two samples give similar absorption spectra. Most importantly both spectra show the B-CB peak at around 700 nm. From this it can be suggested that the indirect doping method introduced in this work is able to introduce impurities in a grown sample. The density of introduced impurities is from the results thought to be similar, but methods better suited to give impurity concentrations should be used to validate this finding.

The doped samples D1-D5 and D8 all show band edge absorption and the B-CB absorption peak, but none of the samples shows the VB-B peak, which should be present at around 1700 nm. A possible explanation for this would be that the B-level is near or below the Fermi level, so that the occupancy of the level is too high to show much absorption. The Fermi level can be computed if the doping concentrations are known. It can be assumed that the nitrogen doping concentration is the same in doped samples as in unintentionally doped samples, i.e. the donor concentration is $N_D \approx 10^{16} \text{ cm}^{-3}$. Assuming that the B-doped samples are of p-type, i.e. $p >> n$, the neutrality condition is

$$p = N_A^- - N_D^+. \quad (7.1)$$

Since the nitrogen energy level is shallow, it can be assumed that at room temperature all donor atoms are ionized. This gives that $N_D^+ = N_D$. Under the Boltzmann approximation the hole density is

$$p \approx N_V e^{-E_F/kT}, \quad (7.2)$$

where N_V is the valence band density of states, and is given by

$$N_V = 2 \left(\frac{kT m_{d,h}}{2\pi\hbar^2} \right)^{3/2}, \quad (7.3)$$

where $m_{d,h}$ is the density of states hole mass. The number of ionized acceptor atoms is governed by the temperature and the distance between the Fermi level and the acceptor level,

$$N_A^- = \frac{N_A}{1 + 2e^{\frac{E_A - E_F}{kT}}}. \quad (7.4)$$

From equations 7.2-7.4 the neutrality condition 7.1 can be rewritten as

$$f(E_F) \equiv N_V e^{-E_F/kT} + 2N_V e^{(E_A - 2E_F)/kT} + 2N_D e^{(E_A - E_F)/kT} - N_A + N_D = 0, \quad (7.5)$$

where the LHS is denoted $f(E_F)$. By using the values $E_A = 0.735 \text{ meV}$, $m_{d,h} = 0.6m_0$ and $T = 300 \text{ K}$, equation 7.5 gives the plots in figures 7.1 and 7.2. In 7.1, the

graphs show the function $f(E_F)$ for different values of N_A , ranging from 10^{17} (light blue) to 10^{18} (dark blue) cm^{-3} . In figure 7.2 the Fermi level has been computed for concentrations ranging from 10^{16} to 10^{21} cm^{-3} . It can be seen that for concentration around $3 \times 10^{17} \text{ cm}^{-3}$, the Fermi level is at 0.74 eV, which corresponds to the acceptor level. It is likely that the doping concentration in the sample is in the same order of magnitude, or smaller, as in the source material from which it was grown. From this it can be concluded that for doped samples grown with source impurity concentration of 10^{18} the Fermi level likely is near or above the acceptor level at room temperature. If the doping concentration of grown samples is significantly smaller than of the source material, then it is probable that also the samples grown from 10^{19} and 10^{20} cm^{-3} sources have the Fermi level above the acceptor level. This may be the reason why the VB-B peak is not visible in the absorption spectra.

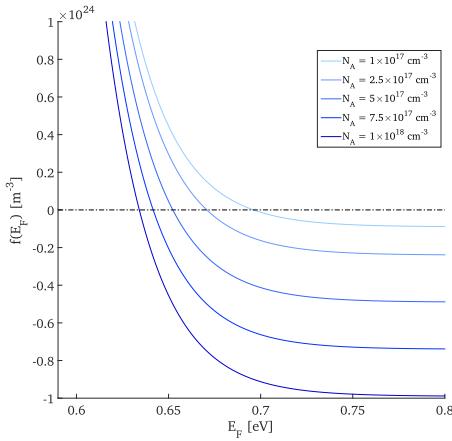


Figure 7.1: Computed values of LHS in equation 7.5, plotted against the Fermi level. The three graphs show acceptor concentrations of 10^{17} , 10^{18} and 10^{19} cm^{-3} .

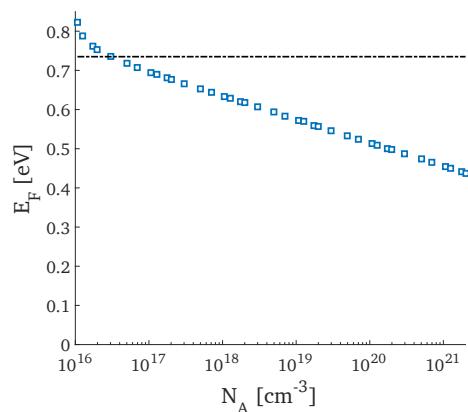


Figure 7.2: Calculations of the Fermi level for different acceptor atom concentrations

Most of the B-doped samples showed no luminescence at all. One possible explanation for this is that the acceptor level is of non-radiative character. It is to be expected that boron, which is a deep level impurity, is of non-radiative character and recombines in a way which does not emit a photon, such as by the Shockley-Reed-Hall process [31, 32]. There have, however, been reports of luminescence from B-doping levels in 3C-SiC, for example by Kubawara et al. [33]. This means that the fact that the B-level may be to some extent non-radiative does not entirely explain why most of the doped samples measured show no luminescence.

Another explanation of the PL-spectra from the doped samples is that the boron doped samples are not of sufficient crystal quality. As shown in figure 6.4 the doped samples contain many defects. It may be that the many defects form non-radiative levels in the band gap and which compete with the boron level. This would explain why the results do not agree to those reported in literature.

One sample (D9) did show luminescence at around 6000 Å which corresponds

to an aluminium impurity level. As aluminium is a shallow acceptor, it will be able to capture carriers more easily compared to the deep levels created by boron. It is likely that more carriers are captured in the Al-level, which decreases the number of exciton recombinations from the B-level. This may be another explanation of why no boron related lines can be seen. There is luminescence from the hexagonal inclusions in this sample (inclusions are shown in figure 6.6 (c)). The luminescence is thought to originate from the B-CB transition. It is likely that both polytypes in the same sample have similar impurity concentrations for both B and Al. This is not consistent with the hypothesis that B-Al competition is responsible for the lack of B-lines in the 3C-SiC spectra.

8

Conclusion

Initially the idea was to investigate how C- and Si-face grown seeds differed in growth of B-doped samples. I was not able to reproducibly grow C-face seeds which were completely cubic. Investigating the seed growth I was able to show that however the temperature conditions were changed, 3C-SiC could not grow to cover the whole growth area on C-face substrates, while on the Si-face it was possible to reproducibly do so. The reason for this is not fully understood, but the reason may be the difference inherent in the surface of C-face and Si-face 4H-SiC, such as the surface energy of the surfaces. It may also be that the steps from the off-axis cut are different on C- and Si-faces, or contributed to the different form of the spirals.

I have shown through optical microscopy that high B-doping leads to deteriorating surface quality of the sample. For most samples the surface quality worsened with increasing B-doping concentration of the source material. From optical microscopy and absorption measurements I have shown that it is possible to grow 3C-SiC material doped with boron using the sublimation growth technique described in this thesis. The absorption measurements indicate the boron to conduction band transition, but not the valence band to boron transition. The lack of the latter is thought to be due to the position of the Fermi level near or above the boron level in the band gap.

I have shown using LTPL-spectroscopy that the boron levels in 3C-SiC grown in this work show no luminescence related to boron. The reason for this may be a combination of the deep level nature of the boron impurity with the poor crystal quality containing non-radiative defects. These PL-measurements further show inclusions of aluminium in the samples, which may form a competing relationship with the boron in carrier capture, further limiting the possibilities of luminescence.

9

Future work

To be able to investigate the growth-face influence on the B-doped material, further study of 3C-SiC growth on the C-face should be done. The pressure is a parameter which has not been explored in this regard. It is possible that by changing the growth pressure the results will improve. Studies of the influence of the surface on growth mechanisms will also be an important tool in understanding cubic growth on the C-face. Experiments such as electron microscopy will be of interest here.

To be able to realize the proposed optoelectronic applications for B-doped 3C-SiC, it is necessary to understand the luminescence properties of the material. Studies of the defects in the grown material should be done, and ways to eliminate them should be explored. This will include both finding better growth conditions and researching ways of improving the quality post-growth. Further studies of the relationship between the Al and B acceptor levels may give some more insight into the optical properties of the material.

More applied studies of the material should also be undertaken in the future. To investigate the material as an electrode material in water splitting applications, photoelectrochemical measurements should be done. Use of boron doped 3C-SiC in photovoltaic applications will need studies of photocurrent generation.

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