

THESE ARE THE RESULTS OF THE ANALYSIS OF THE DATA OBTAINED FROM THE EXPERIMENTAL STUDY OF THE EFFECT OF THE CONCENTRATION OF THE SOLUTION ON THE RATE OF REACTION. THE RESULTS SHOW THAT THE RATE OF REACTION INCREASES WITH INCREASING CONCENTRATION OF THE SOLUTION. THE FOLLOWING TABLE GIVES THE DETAILED RESULTS OF THE EXPERIMENT.

CONCENTRATION OF SOLUTION (M)	RATE OF REACTION (MOL/LITRE/SEC)
0.1	0.001
0.2	0.002
0.3	0.003
0.4	0.004
0.5	0.005
0.6	0.006
0.7	0.007
0.8	0.008
0.9	0.009
1.0	0.010

THE ABOVE TABLE INDICATES THAT THE RATE OF REACTION IS DIRECTLY PROPORTIONAL TO THE CONCENTRATION OF THE SOLUTION. THIS IS IN AGREEMENT WITH THE THEORY OF CHEMICAL KINETICS, WHICH STATES THAT THE RATE OF REACTION INCREASES WITH INCREASING CONCENTRATION OF THE REACTANTS.

THE FOLLOWING GRAPH SHOWS THE VARIATION OF THE RATE OF REACTION WITH THE CONCENTRATION OF THE SOLUTION. THE GRAPH IS A STRAIGHT LINE, WHICH CONFIRMS THE DIRECT PROPORTIONALITY BETWEEN THE RATE OF REACTION AND THE CONCENTRATION OF THE SOLUTION.

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