

Air N_2, O_2, Ar

$$n_{N_2} + n_{O_2} + n_{Ar} = n_{total}$$

mole percent

$$\frac{P_{N_2}}{P_{total}} \times 100 = \frac{n_{N_2}}{n_{total}} \times 100$$

Mole Fraction

$$P_{N_2} + P_{O_2} + P_{Ar} = P_{total}$$

Dalton's Law

$$\frac{n_{N_2}}{n_{total}} = \frac{P_{N_2}}{P_{total}} \Rightarrow P_{N_2} = \left(\frac{n_{N_2}}{n_{total}} \right) \cdot P_{total}$$

$$P_{N_2} V = n_{N_2} R T$$

$$P_{total} V = n_{total} R T$$

<u>% comp</u>	<u>78% N_2</u>	<u>21% O_2</u>	<u>1% Ar</u>
<u>fractions</u>	0.78	0.21	0.01

$$\text{today Air pressure} = 1.006 \text{ atm}$$

$$P_{N_2} = (0.78) 1.006 \text{ atm} = 0.785 \text{ atm}$$

$$P_{O_2} = (0.21) 1.006 \text{ atm} = 0.211 \text{ atm}$$

$$P_{Ar} = (0.01) 1.006 \text{ atm} = \frac{0.010}{\text{atm}}$$