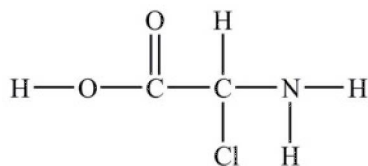


MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question.

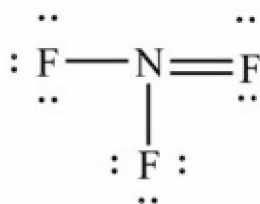
- 1) How many lone pairs of electrons need to be added to complete this Lewis structure? 1) _____



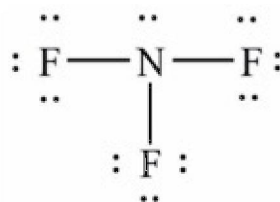
- A) 6 B) 1 C) 8 D) 5 E) 16

- 2) Select the correct Lewis structure for nitrogen trifluoride, NF_3 . 2) _____

A)



B)



- 3) The number of resonance structures for the sulfur dioxide molecule that satisfy the octet rule is 3) _____

- A) 1.
C) 3.

- B) 2.
D) None of the answers is correct.

- 4) Using the VSEPR model, what is the predicted molecular geometry of the PCl_3 molecule? 4) _____

- A) trigonal planar
C) trigonal pyramidal

- B) tetrahedral
D) linear

- 5) Using the VSEPR model, what is the electron geometry of the PCl_3 molecule? 5) _____

- A) bent
B) trigonal pyramidal
C) tetrahedral
D) trigonal planar
E) linear

- 6) Predict the molecular geometry and polarity of the CS_2 molecule. 6) _____

- A) linear, nonpolar
C) tetrahedral, nonpolar

- B) bent, polar
D) linear, polar

- 7) Which one of the following molecules is polar? 7) _____

A) CH_4

B) F_2

C) CO_2

D) CHBr_3

E) CBr_4

Answer Key

Testname: MINI-PRAC CH6

- 1) C
- 2) B
- 3) B
- 4) C
- 5) C
- 6) A
- 7) D