3 parameter Arrhenius form

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 k_f (T) is the temperature dependent reaction rate constant

$$k_f(T) = A * \left(\frac{T}{T_0}\right)^B \exp\left(\frac{E_A * 41840000}{R_u T}\right)$$

Where

 E_A is units of cal/mol (1 calorie = 41840000 erg)

T is Kelvin

The reference temperature: $T_0 = 1$ Kelvin

B is dimensionless

The universal gas constant: $R_U = 83145100 \text{ erg/(mol*K)}$

Units of A are cgs (cm, sec, K, mol), the exact units depend on the reaction order

Sample Reaction

 $H+O_2=O+OH$

 $A = 3.547*10^{15}$ units?

B = -0.406

 $E_A = 1.6599*10^4 \text{ cal/mol}$

$$r_f = k_f * [H]^1 * [O_2]^1$$

the reaction order with respect to reactant H is 1, the overall reaction order is 2 r_f needs units mol/(cm³*sec)

[] is mole concentration of species in units mol/cm^3 $\,$

Therefore for this second order reaction A has units of

$$\frac{mol}{cm^3*sec}*\left(\frac{cm^3}{mol}\right)^2 = \frac{cm^3}{sec*mol}$$

(For a first order reaction it has units of s⁻¹, for that reason it is often called frequency factor.)

Equilibrium constant gives us k_b

$$K_c(T) = \left(\frac{P^0}{R_u T}\right)^{\sum_{i=1}^{N} (v_i^{\prime\prime} - v_i^{\prime})} \exp\left[\frac{-\Delta \bar{g}_r^0(T)}{R_u T}\right]$$

$$k_b = k_f / K_c$$

$$\Delta \bar{g}_r^0(T) = \sum_{i=1}^{KK} (v_i'' - v_i') \bar{g}_i^0(T)$$

The following relations express the net production rate of each species in a multistep mechanism:

 \dot{w}_i is the production rate of the j-th species, with units of mol/(cm³*sec)

$$\dot{w}_j = \sum_{i=1}^{II} v_{ji} q_i \qquad for \quad j = 1, 2, \dots, KK$$

i is reaction #, II is total number of reactions i is species #, KK is total number of species

 v_{ii} is the net stoichiometric coefficient matrix, v_{ii} is reactants, v_{ii} is products, (unit dimensionless)

$$v_{ii} = \left(v_{ii}^{\prime\prime} - v_{ii}^{\prime}\right)$$

$$v_{ji} = \left(v_{ji}^{"} - v_{ji}^{"}\right)$$

$$q_i \text{ is the rate-of-progress variable, with units of mol/(cm^3*sec)}$$

$$q_i = k_{f,i} \prod_{j=1}^{KK} c_j^{v_{ji}} - k_{b,i} \prod_{j=1}^{KK} c_j^{v_{ji}}$$

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