# Derive how to find amount of oxygen

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Contact info for Max Plomer email: maxplomer@gmail.com cell: 203-945-8606

How to find amount of oxygen  $(O_2)$ ,  $\alpha$ , at Phi=1?

$$C_x H_Y O_Z + \alpha * O_2 = A * CO_2 + B * H_2 O$$

balance C: X=A balance H: Y=2\*B

balance 0:  $Z+2*\alpha=2*A+B$ 

$$\alpha = A + \frac{B}{2} - \frac{Z}{2} = X + \frac{Y}{4} - \frac{Z}{2}$$

where: X is the total number of carbon atoms in the fuel species Y is the total number of hydrogen atoms in the fuel species Z is the total number of oxygen atoms in the fuel species

My code finds the amount of oxygen  $(O_2)$  at phi=1, this is the variable 'amount\_oxidant'. It can handle multiple fuel species, only one fuel ratio at a time.

## Pseudo code:

- -Find what number C,H,O are on the list of elements, put this into variables c\_order, h\_order, o\_order
- -go through the fuel species and add the amount of C/H/O to the to  $c\_tot$ ,  $h\_tot$ ,  $o\_tot$  variables
- -then use the equation derived above to calculate amount\_oxidant, which is the amount of  $O_2$  for phi =1 (fuel equivalence ratio)
- -then the actual amount of  $O_2$  used in the reactor is: amount oxidant/phi

# Equivalence ratio:

$$\phi = \frac{fuel - to - oxidizer \ ratio}{(fuel - to - oxidizer \ ratio)_{st}} = \frac{m_{fuel}/m_{ox}}{\left(m_{fuel}/m_{ox}\right)_{st}} = \frac{n_{fuel}/n_{ox}}{\left(n_{fuel}/n_{ox}\right)_{st}}$$

st means stoichiometric which is the same as saying phi=1 lets say we want to find out how much oxygen to put in to methane and air at phi=2? Well. at stoichiometric:

$$CH_4 + 2 * O_2 = CO_2 + 2 * H_2O$$

we find that

$$\left(\frac{n_{fuel}}{n_{ox}}\right)_{ct} = \frac{1}{2}$$

therefore at phi=2

$$\frac{n_{fuel}}{n_{ox}} = \phi \left(\frac{n_{fuel}}{n_{ox}}\right)_{st} = 2 * \frac{1}{2} = 1$$

if you just assume that the amount of fuel will be constant, then

$$n_{ox} = \frac{n_{ox,st}}{\phi}$$

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