

# **Design Rules for Solid State Fluorescence Exploiting Excited State Intramolecular Proton Transfer**

Michael Dommett

School of Biological and Chemical Sciences  
Queen Mary University of London

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## **Abstract**

This is the abstract.

# Contents

<b>List of Figures</b>	<b>4</b>
<b>List of Tables</b>	<b>5</b>
<b>List of Abbreviations</b>	<b>6</b>
<b>1 Connecting Chromophore Design with Crystal Morphology</b>	<b>7</b>
1.1 Introduction . . . . .	7
<b>2 Bibliography</b>	<b>9</b>
<b>Appendices</b>	<b>10</b>

# List of Figures

A1	CASSCF space used in Chapter ?? . . . . .	11
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# List of Tables

1.1	Molecular structures and their QEs ( $\Phi$ ) in the solid state <sup>???</sup> . . . . .	8
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# List of Abbreviations

<b>AIE</b>	aggregation induced emission
<b>HC</b>	2'-hydroxychalcone
<b>HP</b>	2-hydroxyphenylpropenone
<b>PES</b>	potential energy surface

# 1 Connecting Chromophore Design with Crystal Morphology

## 1.1 Introduction

In Chapters ?? and ?? the potential energy surfaces (PESs) of a range of 2'-hydroxychalcone (**HC**) derivatives were mapped, in both vacuum and the crystalline form. Through the topology of the PESs and the associated energy differences between states, we determined the radiative and nonradiative relaxation channels and rationalise the observed aggregation induced emission (AIE) of **HC1** and the nonemission of **HC5**. By elucidating the AIE mechanism of **1** and the nonemission of **5**, we isolated three design principles to increase the quantum yield of fluorescence for ESIPT chromophores in the solid state.

In this Chapter, the scope of the study is extended as the design rules established in Chapter ?? are applied to a new set of ESIPT systems. Into the test set we add two fluorene-substituted **HC** derivatives, **HC6** and **7**.<sup>?</sup> Furthermore, and most pertinently, four completely new compounds with lasing properties are considered. Closely related to **HCs** are the family of 2-hydroxyphenylpropenone (**HP**) derivatives.<sup>?</sup> In contrast to **HCs**, and other organic fluorophores, **HP** compounds contain only a single aryl group and have remarkable QEs, ranging from 0.72-0.84. This has been qualitatively attributed in experimental studies to the herringbone packing mode and molecular rigidity reducing nonradiative decay. The increased quantum yield of the **HPs**, with respect to the **HCs**, make them prime candidates to test the efficacy of our design rules. The eleven compounds studied in this Chapter are summarised in .

Table 1.1: Molecular structures and their QEs ( $\Phi$ ) in the solid state<sup>???</sup>

	HC.pdf				HP.pdf		
	R <sub>1</sub>	R <sub>2</sub>	$\Phi$		R <sub>3</sub>	R <sub>4</sub>	$\Phi$
<b>HC1</b>	H	H	0.32	<b>HP1</b>	H	H	0.74
<b>HC2</b>	CH <sub>3</sub>	H	0.25	<b>HP2</b>	F	H	0.84
<b>HC3</b>	OCH <sub>3</sub>	CH <sub>3</sub>	0.26	<b>HP3</b>	H	OCH <sub>3</sub>	0.77
<b>HC4</b>	H	CH <sub>3</sub>	<0.01	<b>HP4</b>	H	F	0.72
<b>HC5</b>	H	OCH <sub>3</sub>	<0.01				
<b>HC6</b>	F	H	0.41				
<b>HC7</b>	H	F	0.10				



## 2 Bibliography

- X. Cheng, Y. Zhang, S. Han, F. Li, H. Zhang and Y. Wang, *Chem. Eur. J.*, 2016, **22**, 4899–4903.
- B. Tang, H. Liu, F. Li, Y. Wang and H. Zhang, *Chem. Commun.*, 2016, **52**, 6577–6580.
- X. Cheng, K. Wang, S. Huang, H. Zhang, H. Zhang and Y. Wang, *Angew. Chem., Int. Ed.*, 2015, **54**, 8369–8373.

# **Appendices**

## Appendix A

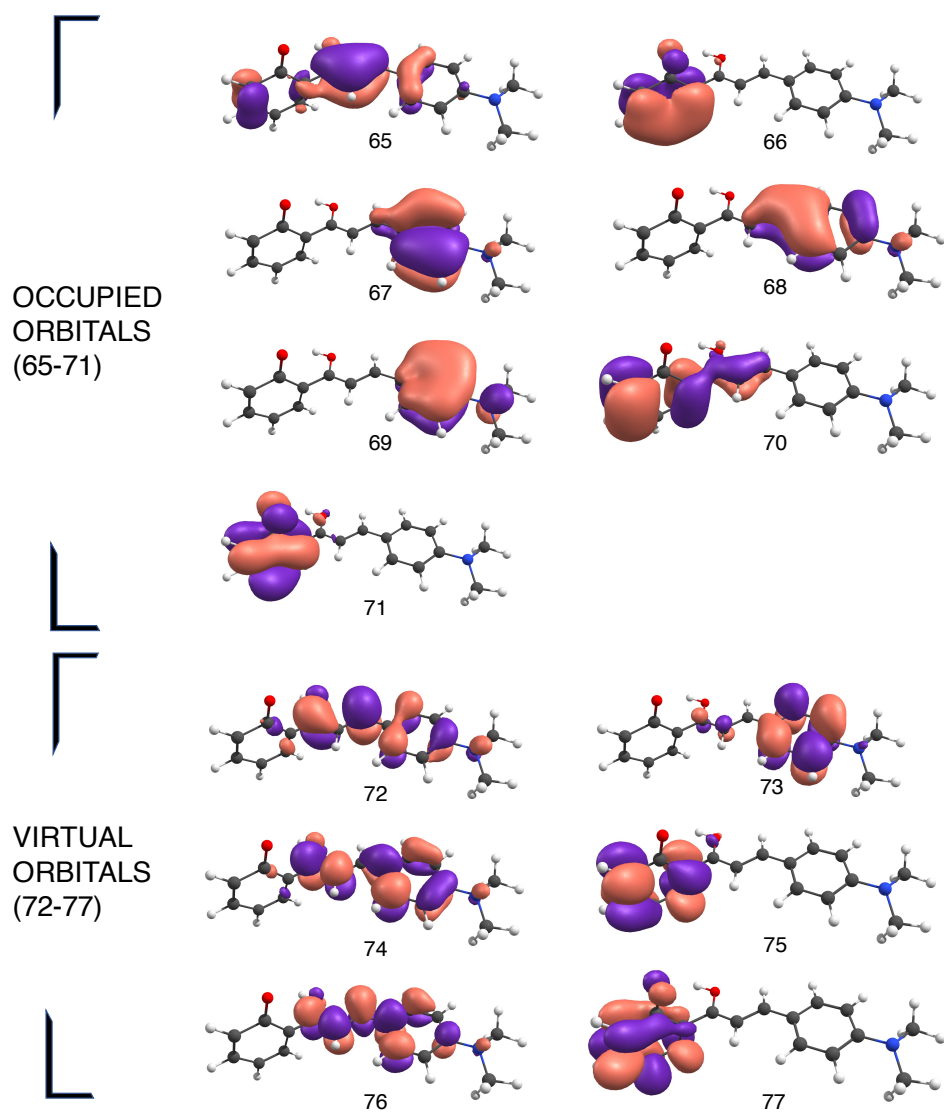


Figure A1: CASSCF space used in Chapter ??.