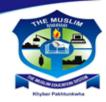




CHEMISTRY 9th (PRE-BOARD EXAM. 2024-25)

Nam Roll 1			والا دائره بھریں۔	ے دیئے گئے ہیں، صرف صحیح جواب	ا۔ ہر سوال کے سامنے چار دائر۔
Class Time	s/Section: e Allowed: 03 Hours ss: 65		.کریں۔	لئے نیلے یاکا لے رنگ کا پن استعال	۲۔ دائروں کوشیر (بھرنے)کے
Time	Allowed: 15 minutes	Section	on "A"	Mar	ks: 12
1.	Anion is formed by gaining one or more	O Neutron	O Proton	O Electron	O Positron
2.	Why is potassium iodide (Kl) (considered as powerful reducing agent?	O Iodide ion is oxidized to iodine	Olodide ion is reduced to iodine	O Iodine is reduced to iodine ion	Olodine is oxidized to iodine ion
3.	The f-subshell can accommodate a maximum of	14 electrons	O10 electrons	06 electrons	O 02 electrons
4.	Which group of elements contains incomplete s-subshell?	O Group I-A	O Group II-A	O Group III-A	O Group IV-A
5.	In Sodium Chloride, Chlorine after gaining one electron attains the electronic configuration of:	O Neon	O Argon	Krypton	O Xenon
6.	Helium obeys duplet rule. How many electrons are in its valance shell:	O 1	O 2	O 3	O 4
7.	Which of the given two elements will form ionic bond:	O Na and K	O K and Ca	O Ca and Cl	O Cl and C
8.	Which of the given solution is more dilute?	O 10 ⁻² M	10^{-3} M	10^3 M	10^{-4} M
9.	Halogens react with metals to form	O Halides	O Oxide	O Halogen sulphides	O Hydrogenated compound
10.	Which one of the given acts as oxidizing agent in the given reaction? $Br_2 + H_2S \rightarrow 2HBR + S$	O Br ₂	О Н	\circ s	\bigcirc H ₂ S
11.	Reduction of alkali metals ion take place at:	O Anode	○Walls of cell	Battery	O Cathode
12.	The fifth element of the first transition series is:	OChromium	O Titanium	O Maganees	O Vanaidum





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Note: Time allowed for Section – B and Section – C is 2 Hours and 45 minutes.

Section – B Marks: 32

Q.II Answer any EIGHT parts. Each part carries FOUR marks.

- Q. 1 Differentiate between physical chemistry and analytical chemistry.
- Q. 2 Compare alkalli and alkaline earth matals with refrence to their loaction in periodic table.
- Q. 3 Write factors which affect the ionization energies of atoms.
- Q. 4 Which rule is followed by Oxygen and Helium to complete their valance shell and why they do so?
- Q. 5 Draw a diagram to show the ionic bond in MgO?
- Q. 6 With the help of dot cross structure show the formation of CH₄ and NH₃.
- Q. 7 A gas occupies 4.31 liters at a pressure of 0.755 atm. Determine the volume if the pressure is increased to 1.25 atm.
- Q. 8 Calculate the molarity of a solution containing 0.5 moles of HCl in 1.5 dm³.
- Q. 9 Explain the chemical reaction of electroplating.
- Q. 10 Explain cell reaction in voltaic cell. Write half-cell reactions and overall reaction.
- Q. 11 Interpret the chemical reaction between
 - i. Potassium and iodine
- ii. Hydrogen and iodine

Section-C

Marks: 21

NOTE: Attempt any THREE questions. All questions carry equal marks.

- Q. 12. (a) Calculate gram molecular mass of H₂SO₄.
 - (b) What do you know about Auf-Bau's principle? How many electrons can accommodate in

M shell of an element?

Q. 13. (a) Which group is known as Boron family in periodic table? Write names of elements

present in this group.

- (b) Write the role of electronegativity in the formation of bonds.
- Q. 14. (a) Explain solid state with reference to attractive forces, motion and gravity.
 - (b) Complete the following table

Class of Colloid	Phase	Example	
Gell	?	?	
?	?	Butter	

- Q. 15 (a) Explain how Alloying and cathodic protection helps to prevent corrosion.
 - (b) Explain the formation of cations from the given elements Na, Ca and Al