

# Temperature Structure of the Atmosphere

EES 3310/5310

Global Climate Change

Jonathan Gilligan

Class #6: Friday, January 17 2020

# Review Question

What is the “atmospheric window”?

1. Regions where there are few clouds to block radiation.
2. Desert regions with very little water vapor.
3. Tropical regions with low CO<sub>2</sub> concentrations.
4. A range of wavelengths where no greenhouse gases absorb much.

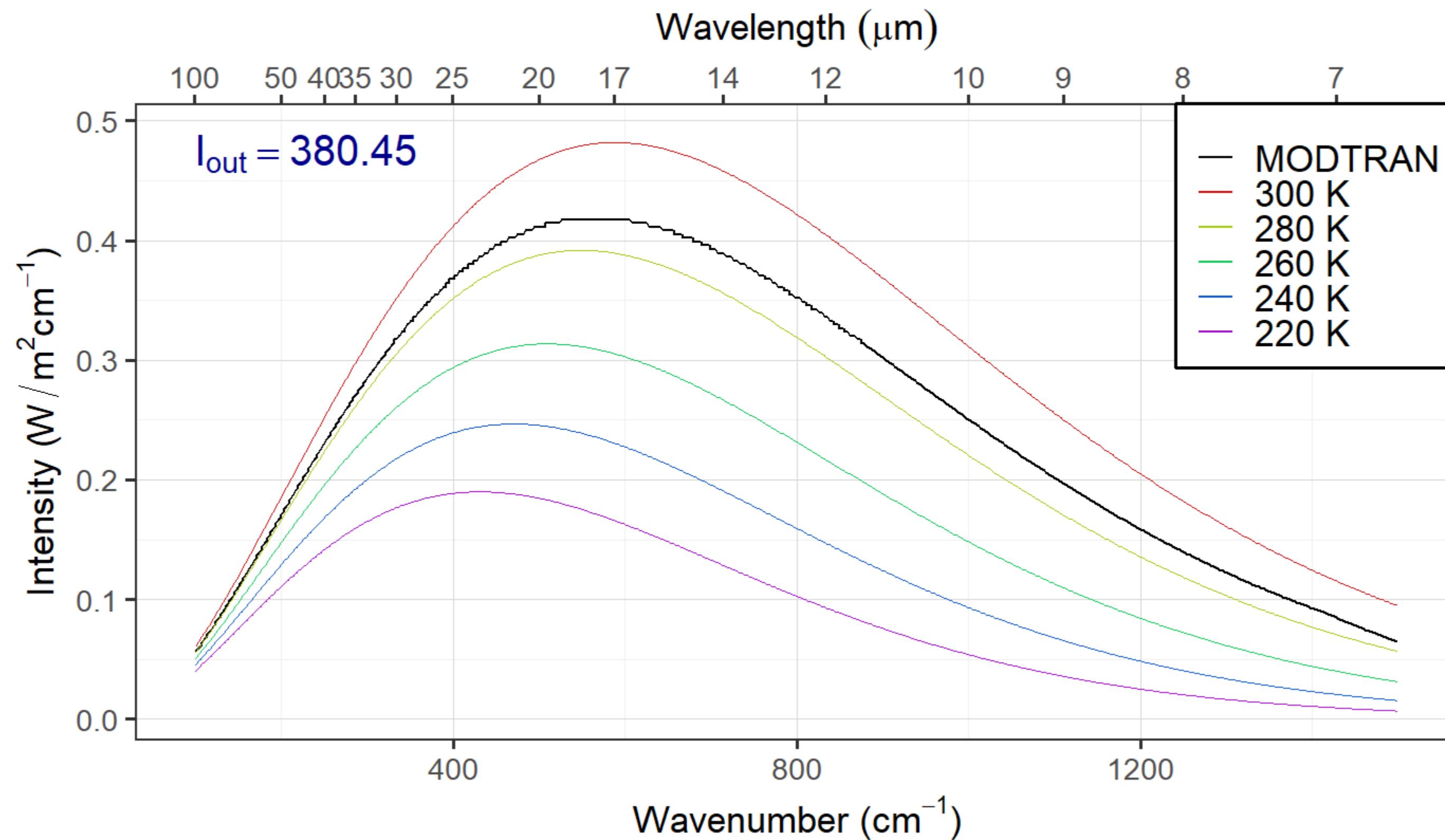
# Measuring Band Saturation

## Set up MODTRAN:

- Go to MODTRAN, set CO<sub>2</sub> to 0.25 ppm, and set all other gases to zero.
- Set altitude to 20 km and location to “1976 US Standard Atmosphere”.
- Press “Save this run to background”
- Note  $I_{\text{out}}$
- Double CO<sub>2</sub> and note the change in  $I_{\text{out}}$
- Keep doubling CO<sub>2</sub> until you get to 1024 ppm.
- Do you notice anything about the changes in  $I_{\text{out}}$ ?

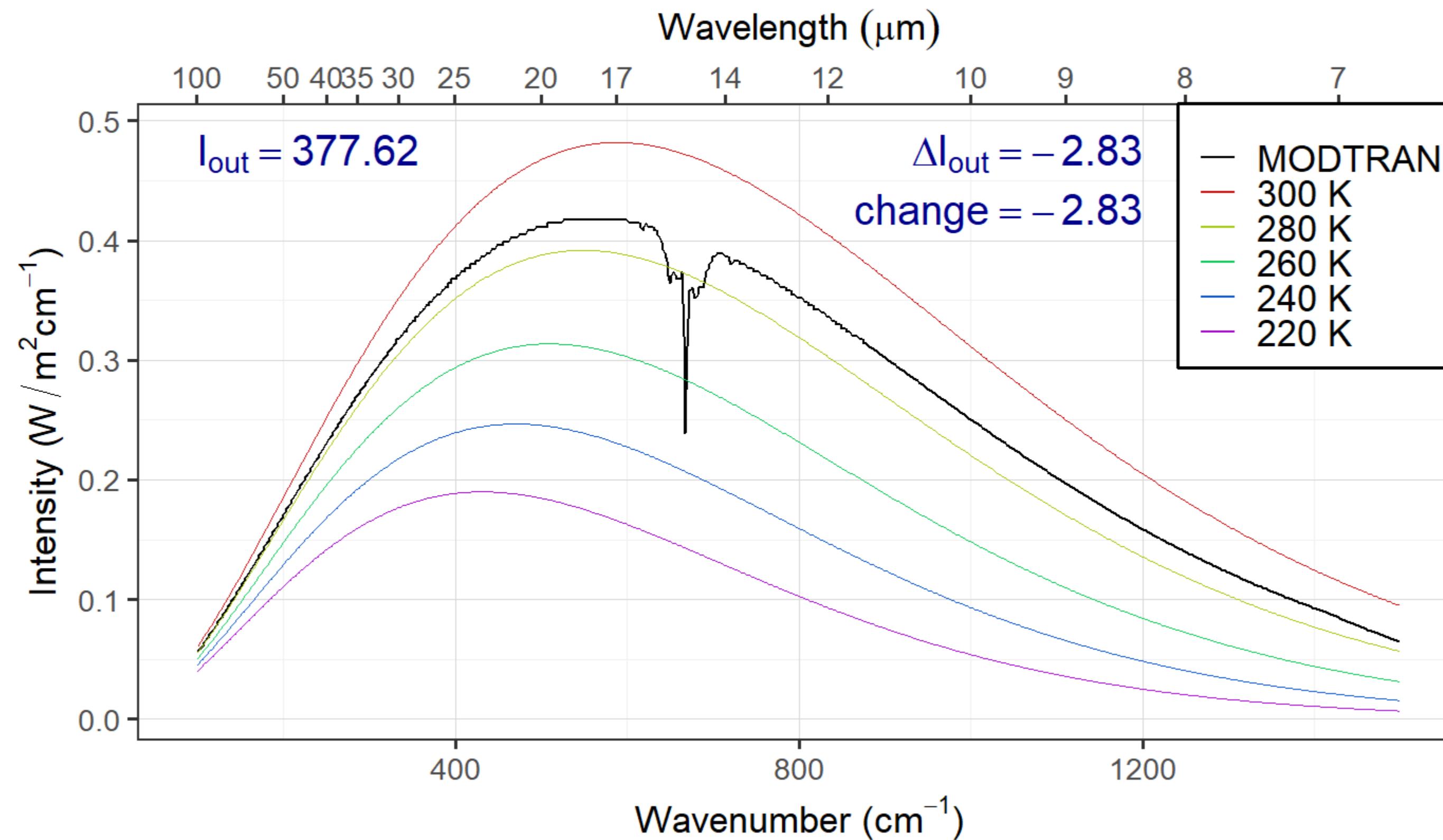
# 0 ppm CO<sub>2</sub>

MODTRAN: 0 ppm CO<sub>2</sub>, 20 km



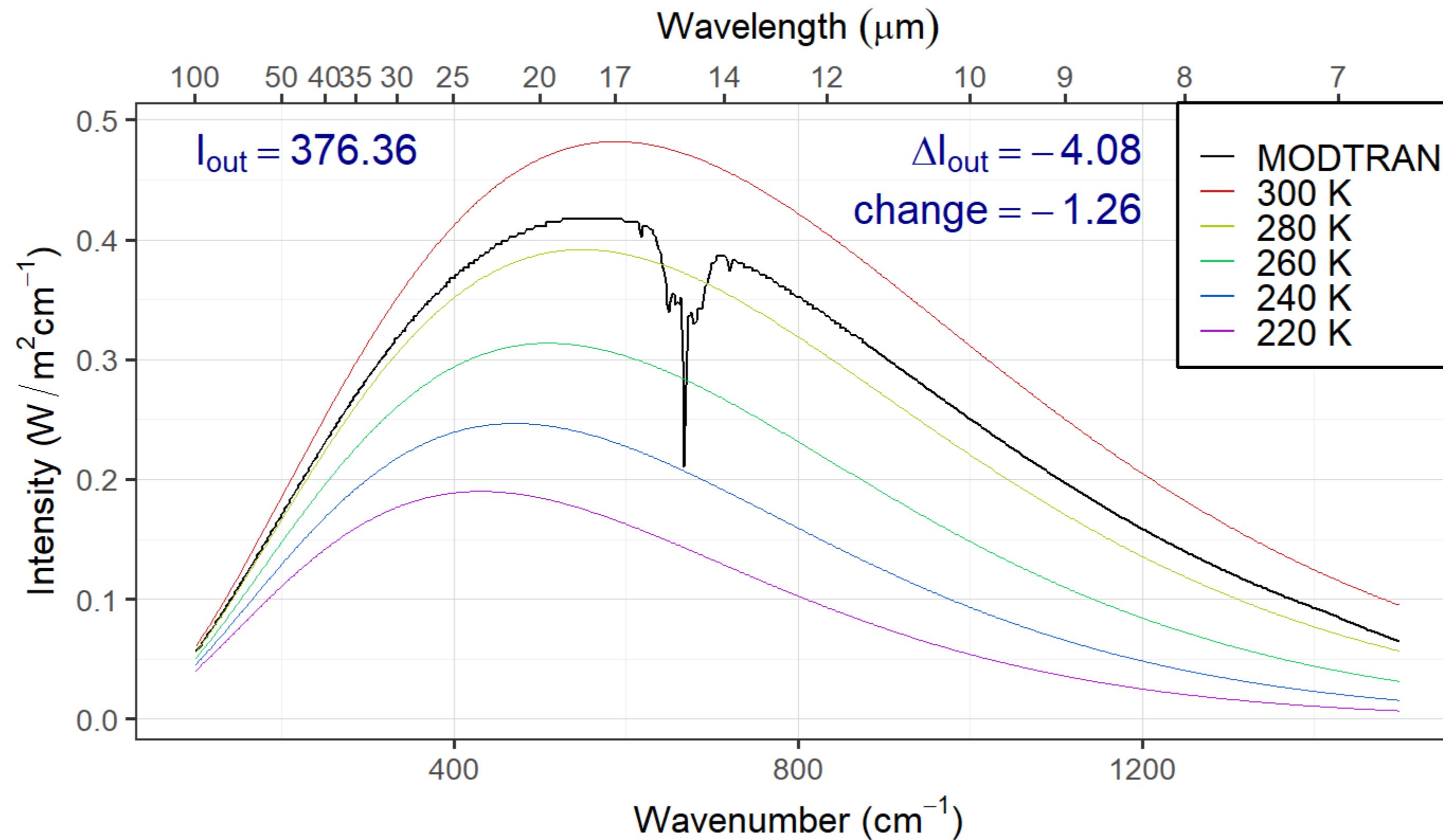
# 0.25 ppm CO<sub>2</sub>

MODTRAN: 0.25 ppm CO<sub>2</sub>, 20 km



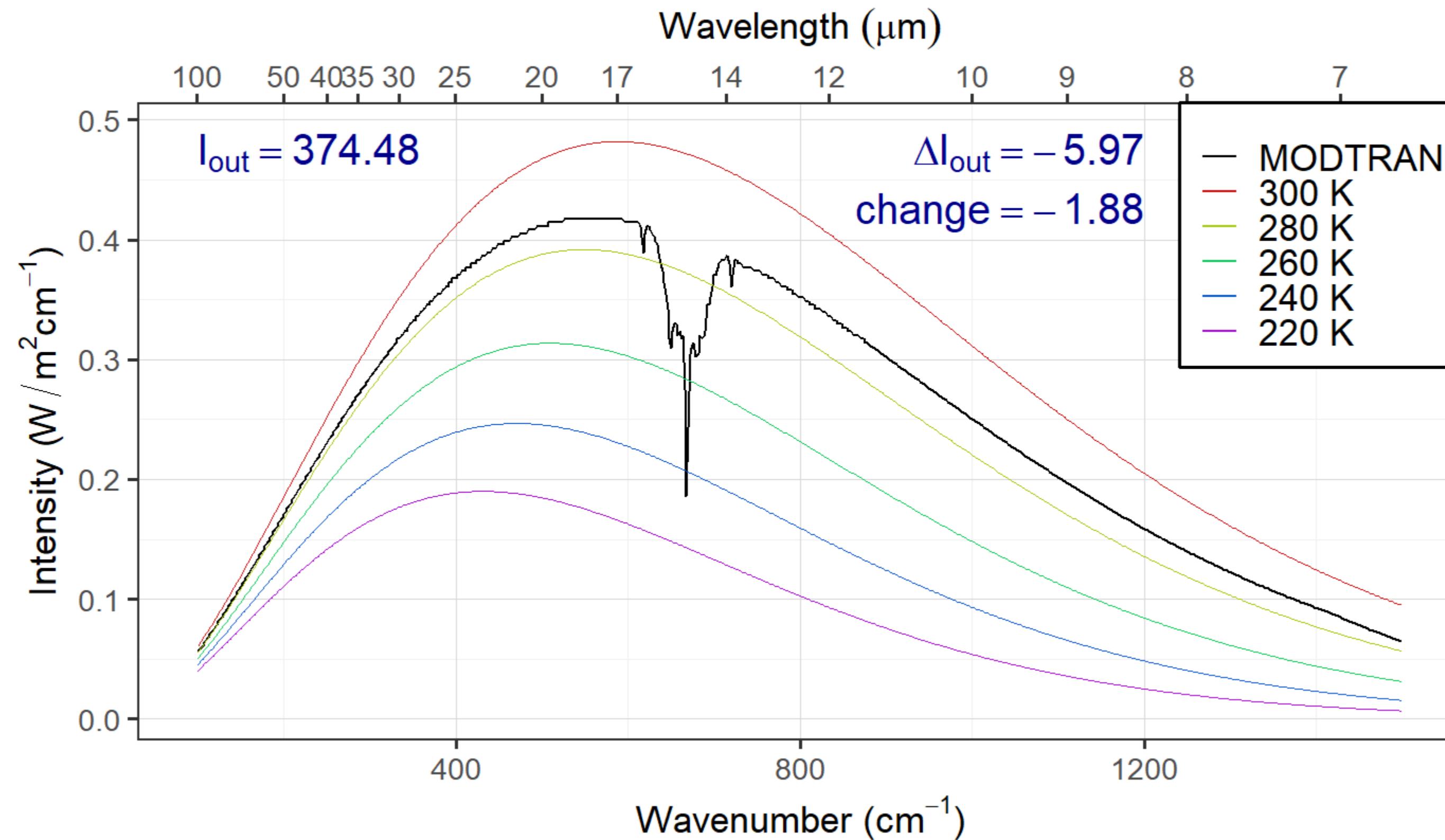
# 0.5 ppm CO<sub>2</sub>

MODTRAN: 0.5 ppm CO<sub>2</sub>, 20 km



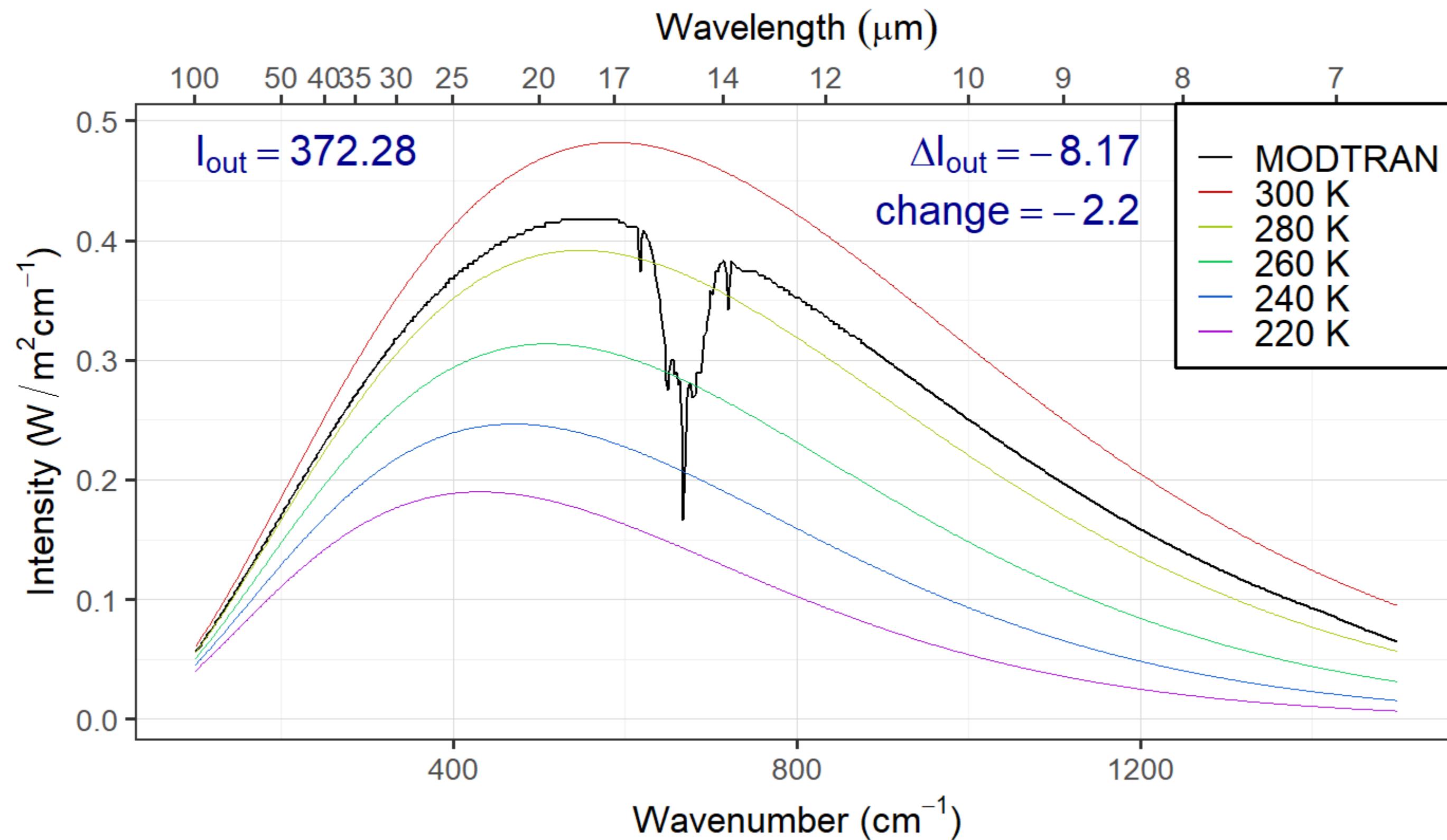
# 1 ppm CO<sub>2</sub>

MODTRAN: 1 ppm CO<sub>2</sub>, 20 km



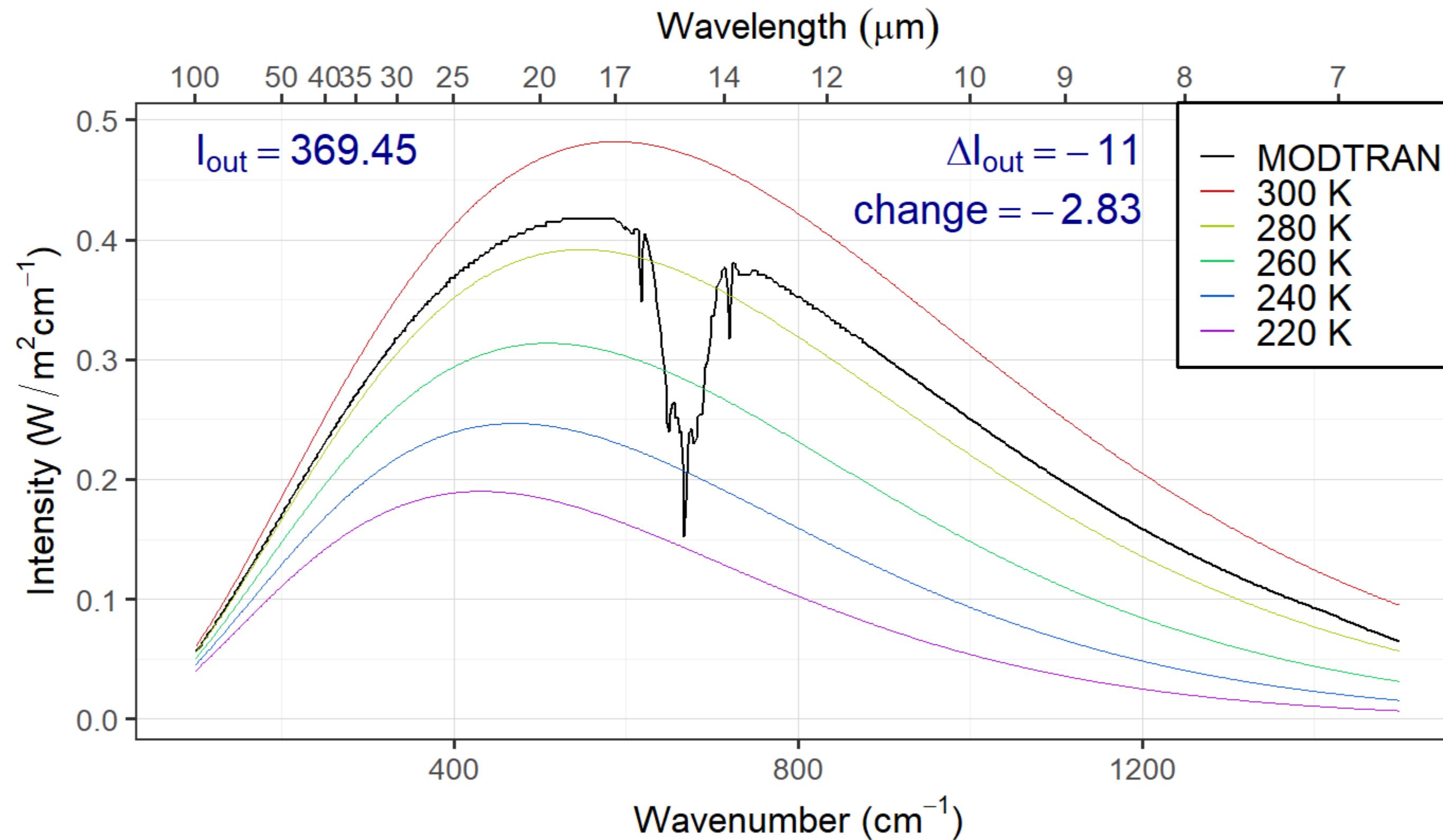
# 2 ppm CO<sub>2</sub>

MODTRAN: 2 ppm CO<sub>2</sub>, 20 km



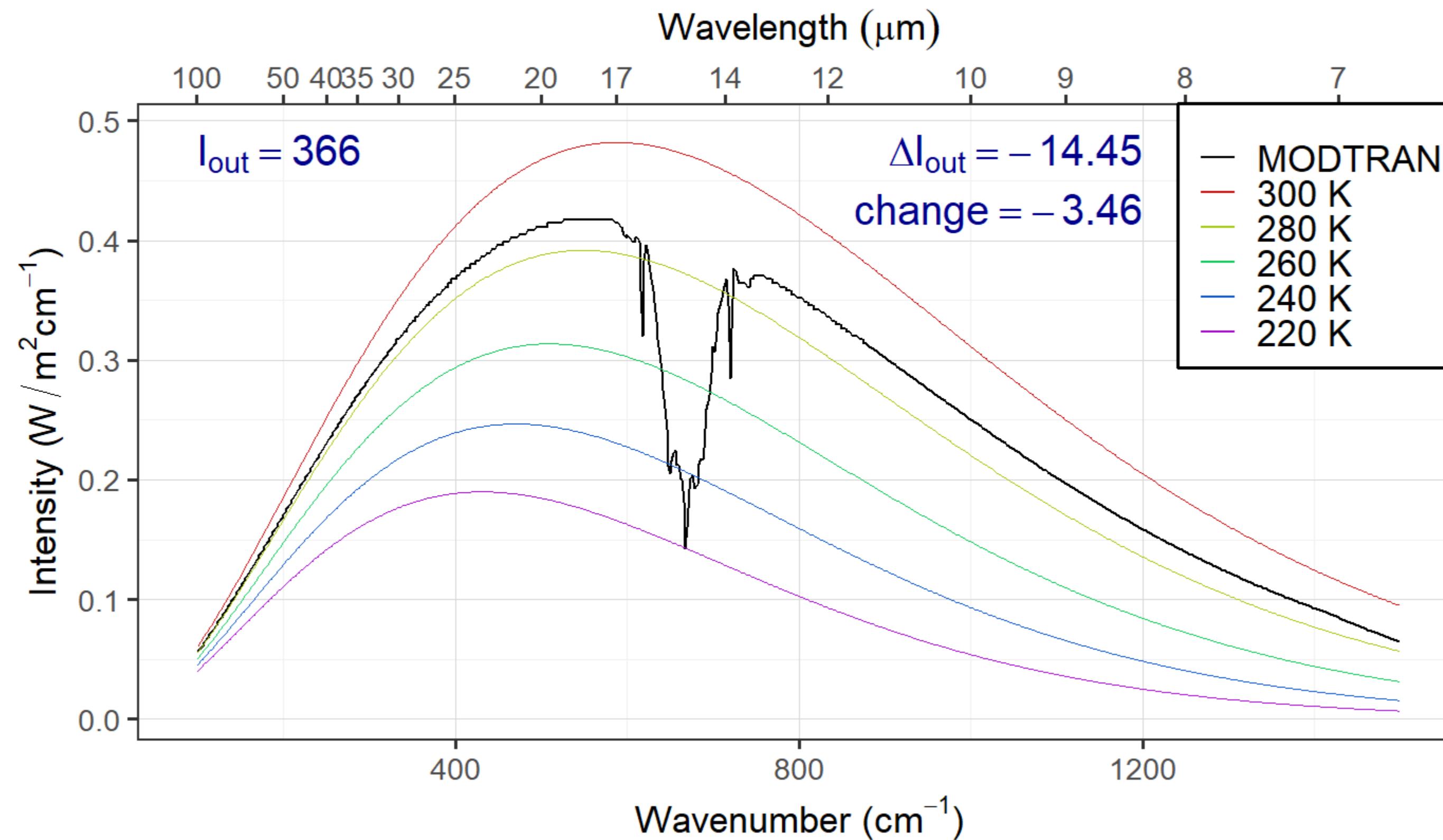
# 4 ppm CO<sub>2</sub>

MODTRAN: 4 ppm CO<sub>2</sub>, 20 km



# 8 ppm CO<sub>2</sub>

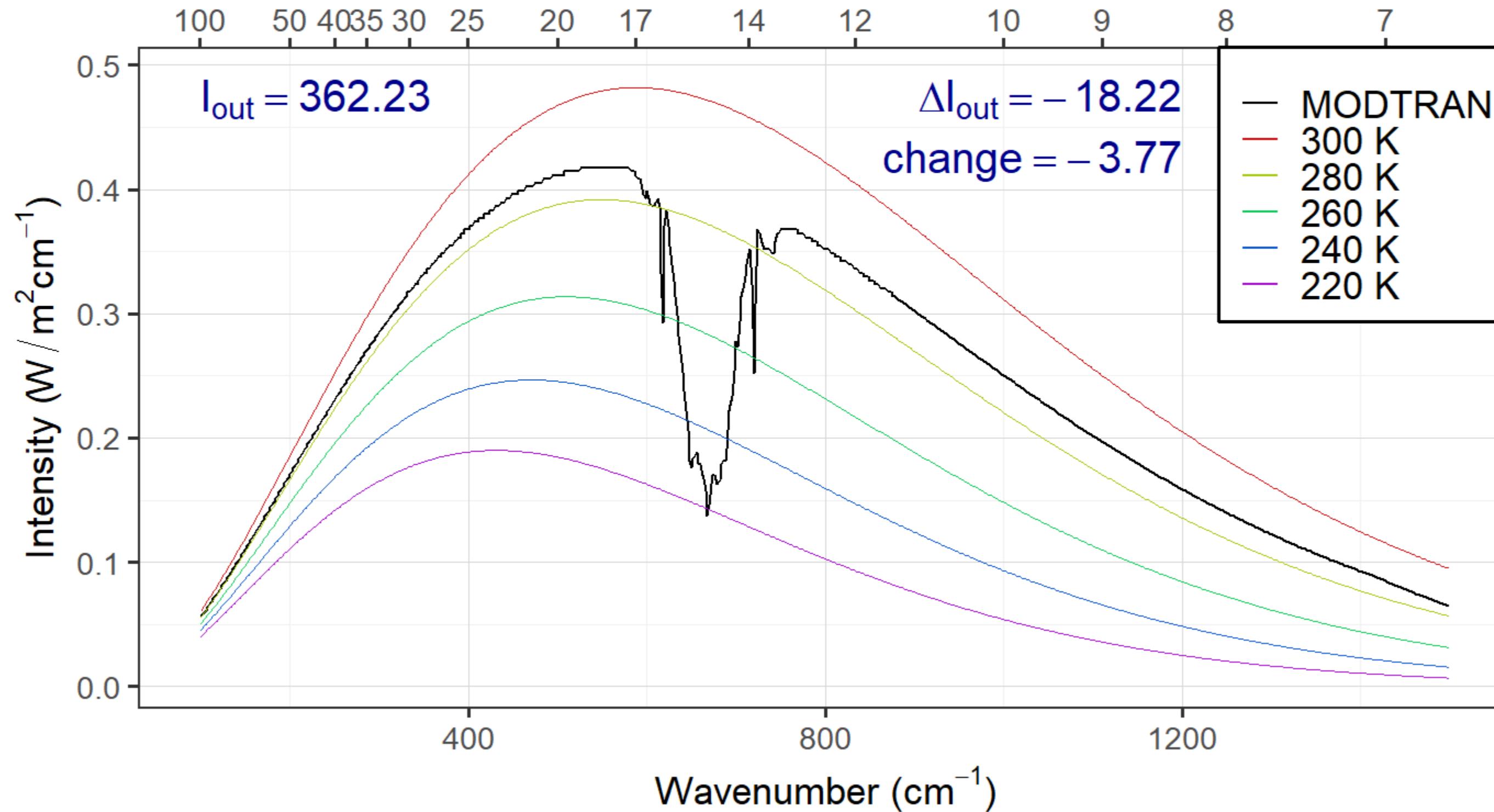
MODTRAN: 8 ppm CO<sub>2</sub>, 20 km



# 16 ppm CO<sub>2</sub>

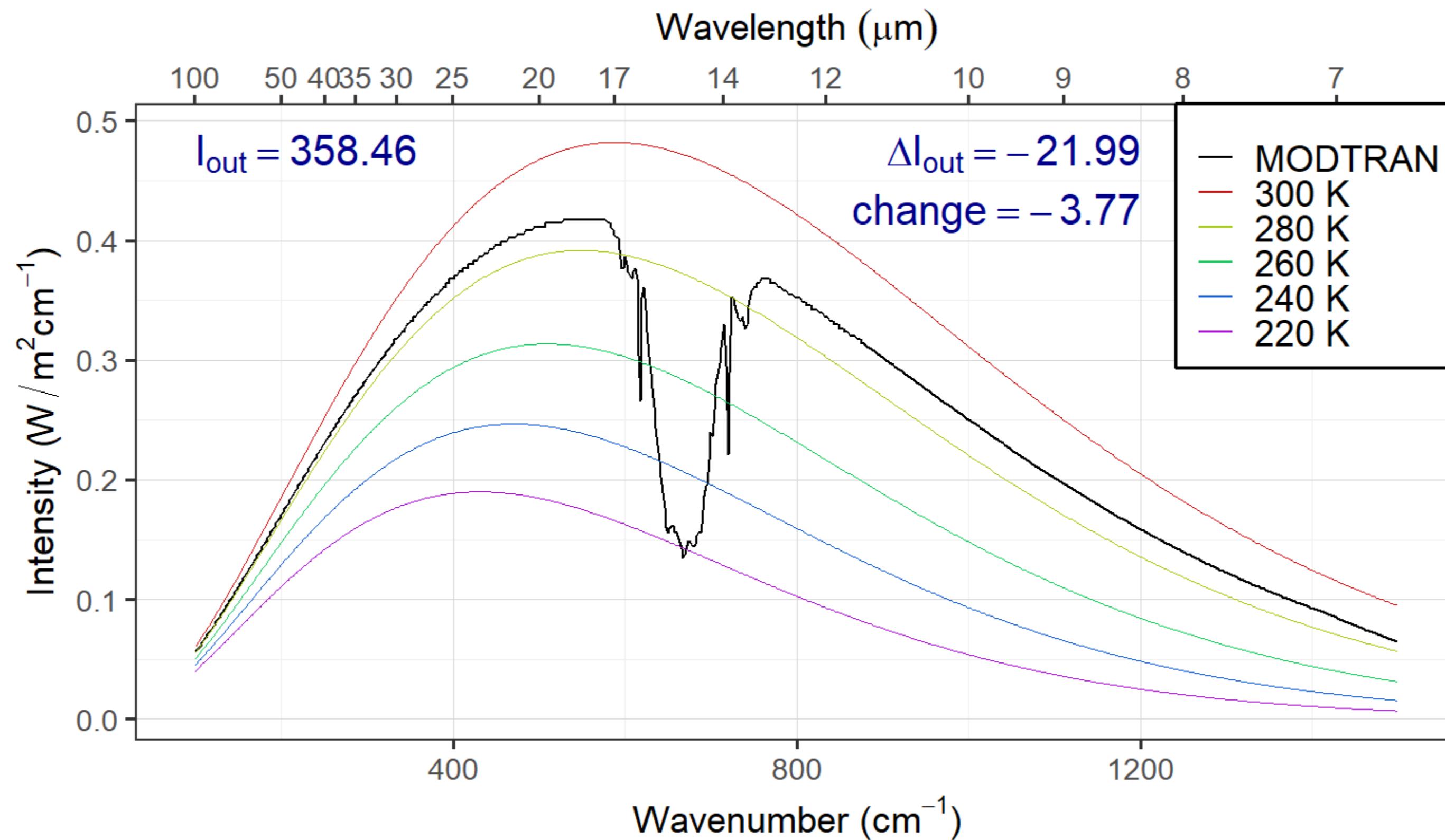
MODTRAN: 16 ppm CO<sub>2</sub>, 20 km

Wavelength ( $\mu\text{m}$ )



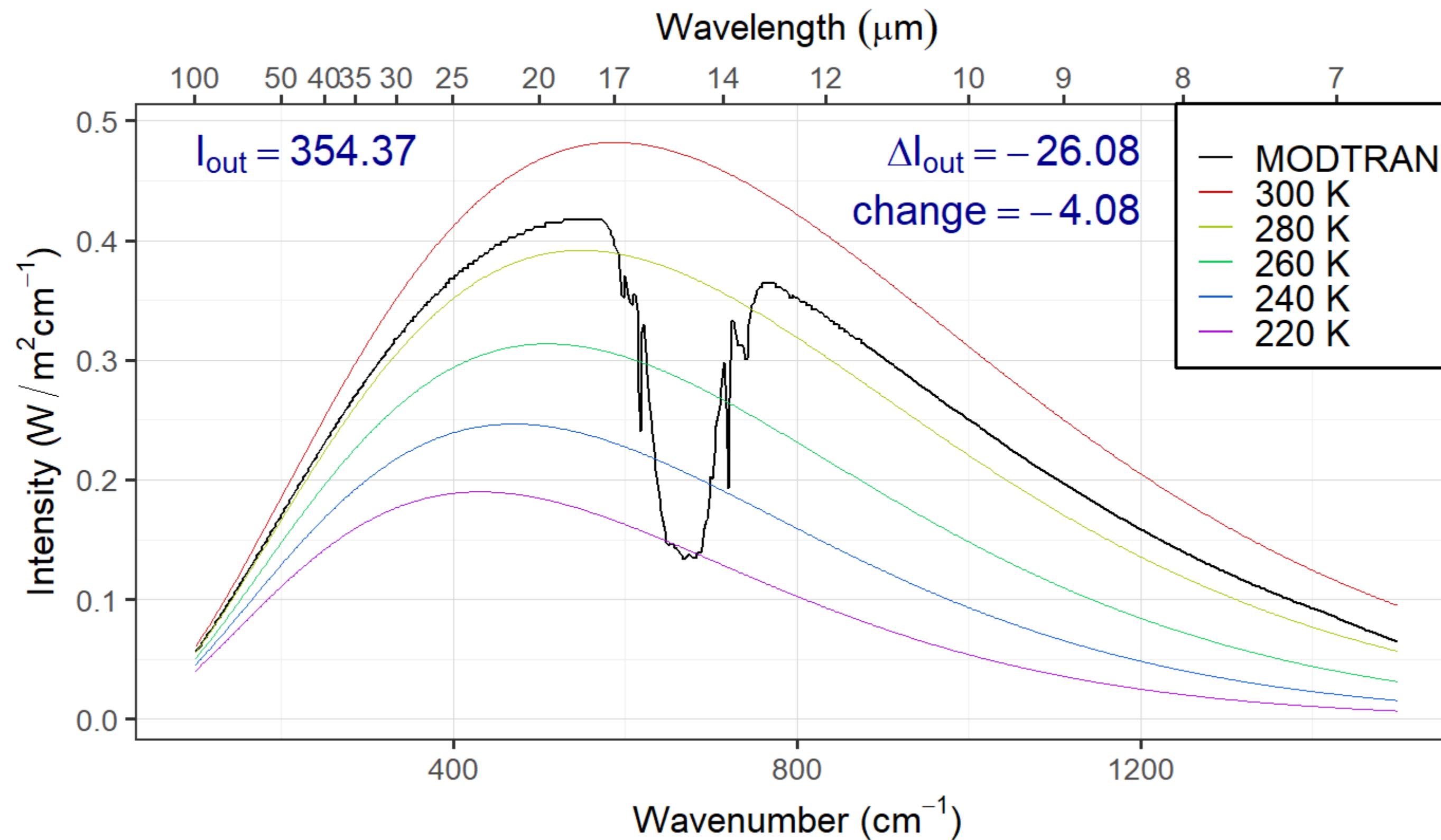
# 32 ppm CO<sub>2</sub>

MODTRAN: 32 ppm CO<sub>2</sub>, 20 km



# 64 ppm CO<sub>2</sub>

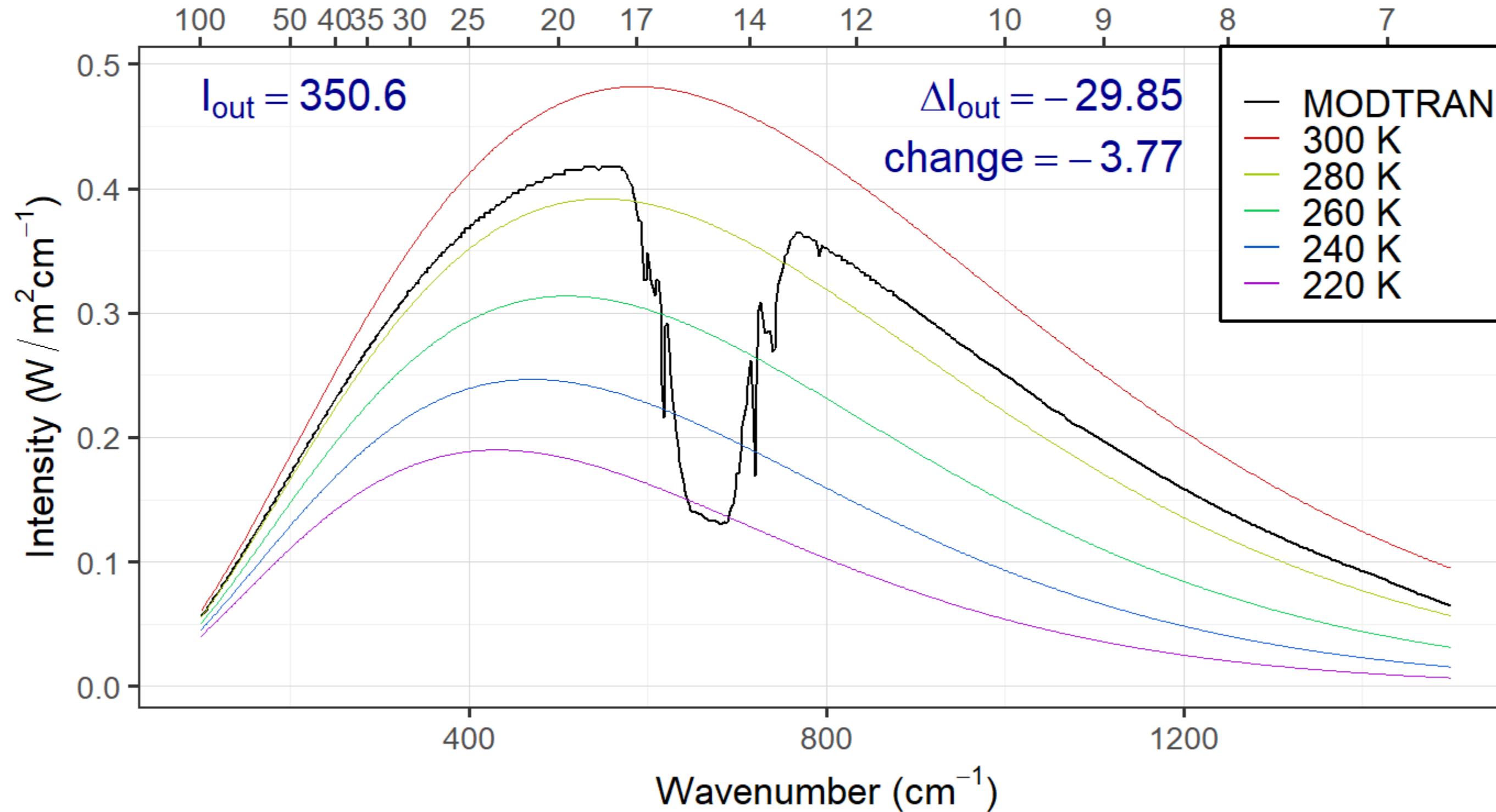
MODTRAN: 64 ppm CO<sub>2</sub>, 20 km



# 128 ppm CO<sub>2</sub>

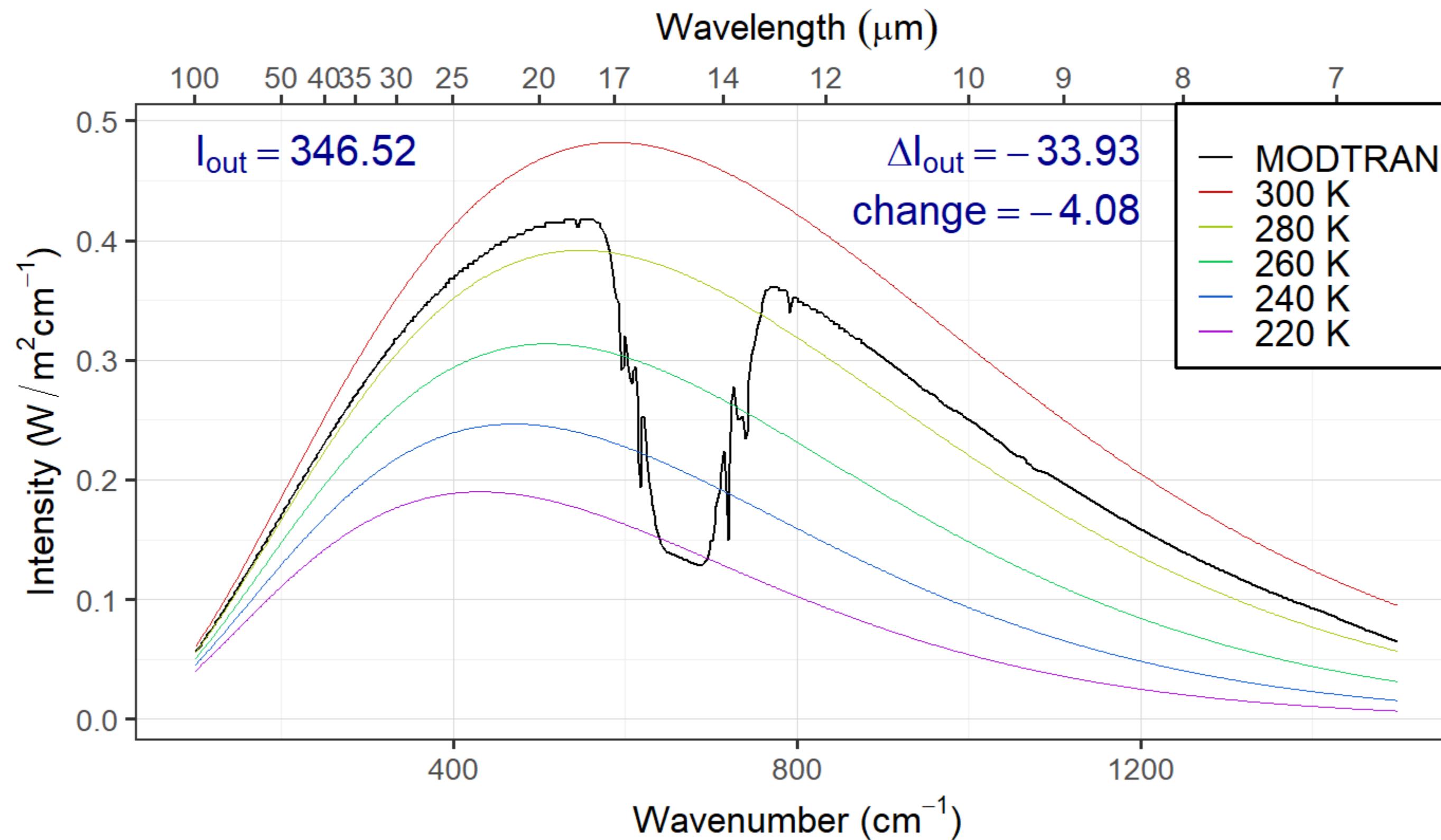
MODTRAN: 128 ppm CO<sub>2</sub>, 20 km

Wavelength ( $\mu\text{m}$ )



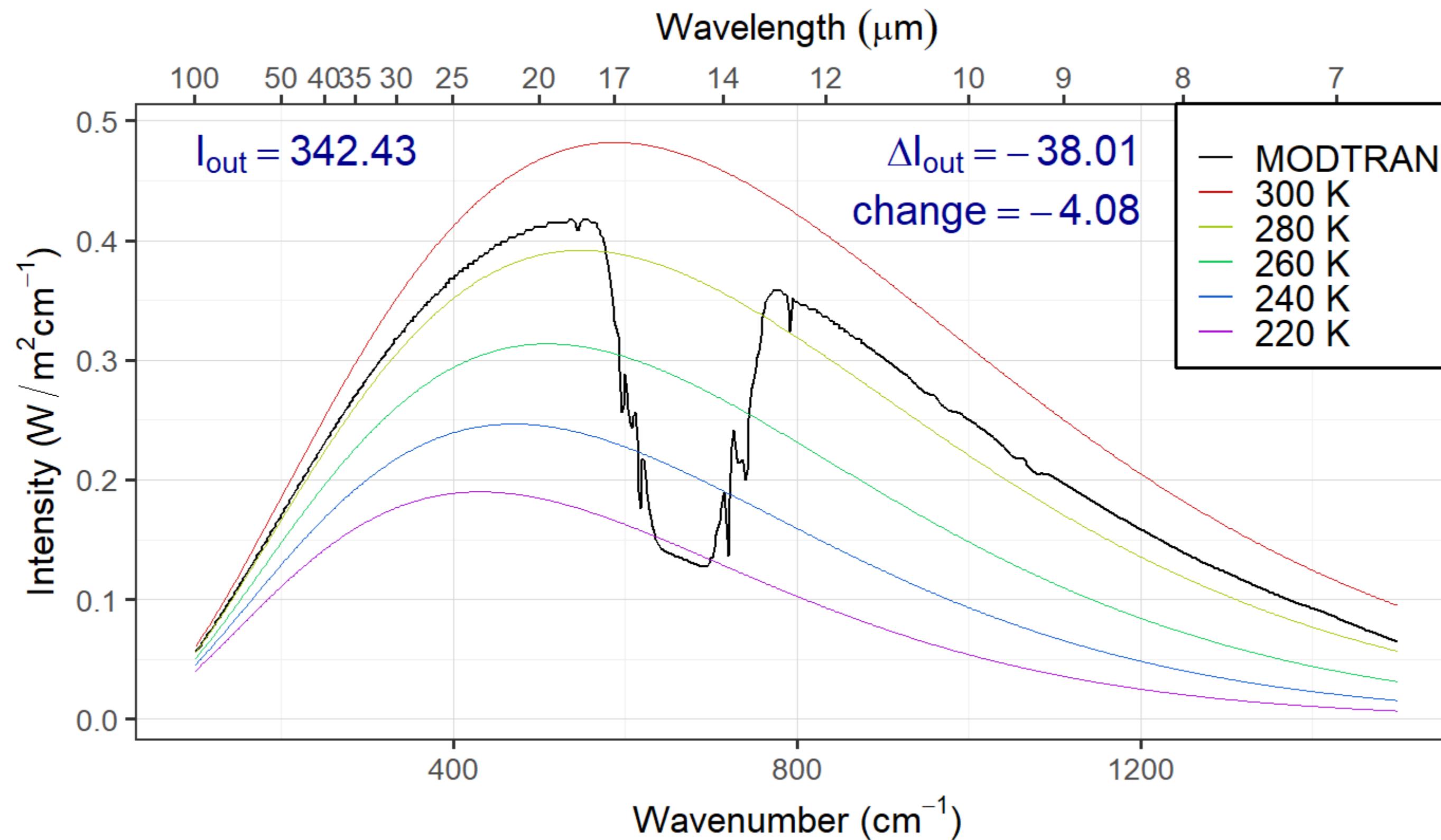
# 256 ppm CO<sub>2</sub>

MODTRAN: 256 ppm CO<sub>2</sub>, 20 km



# 512 ppm CO<sub>2</sub>

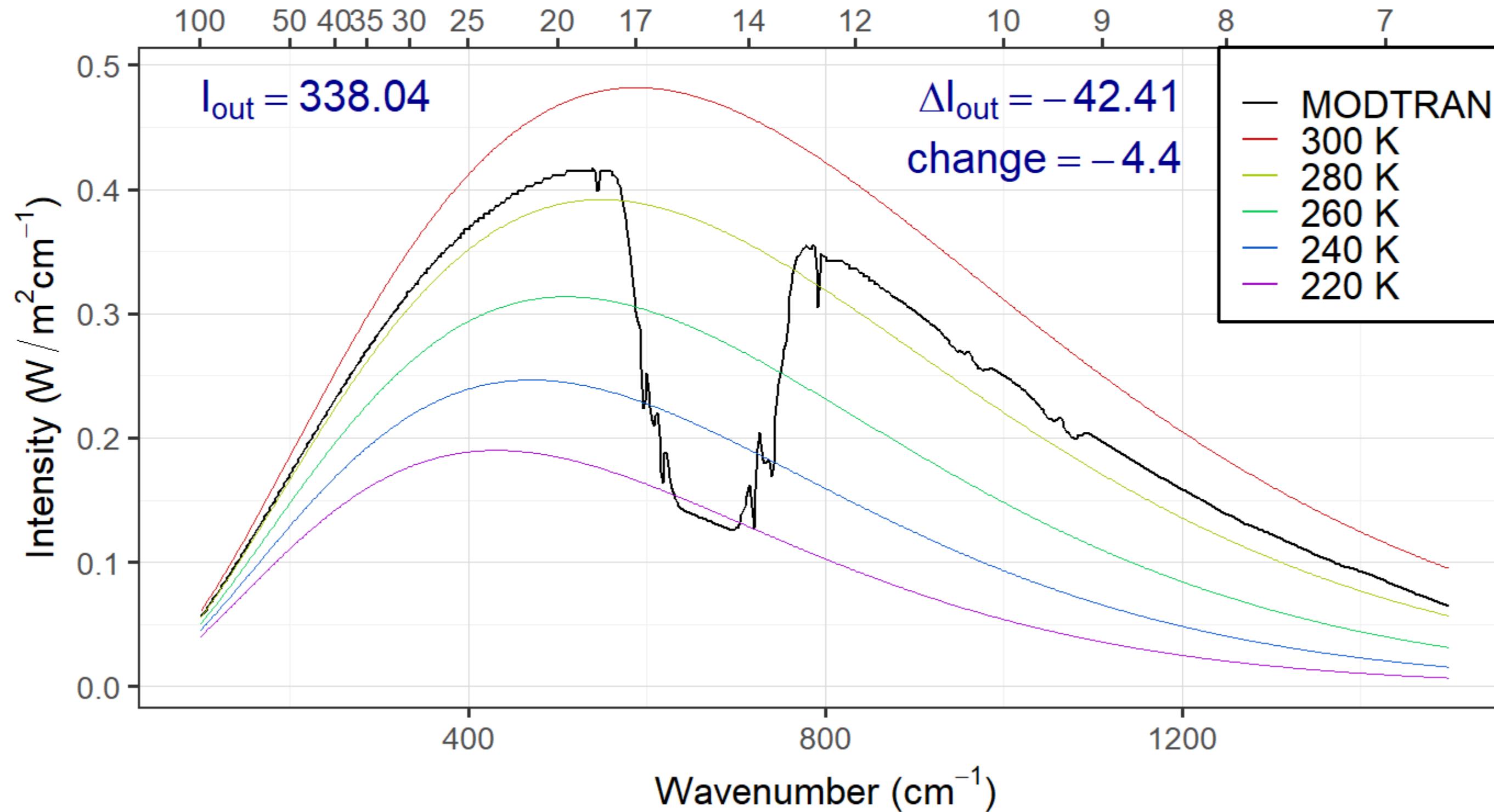
MODTRAN: 512 ppm CO<sub>2</sub>, 20 km



# 1024 ppm CO<sub>2</sub>

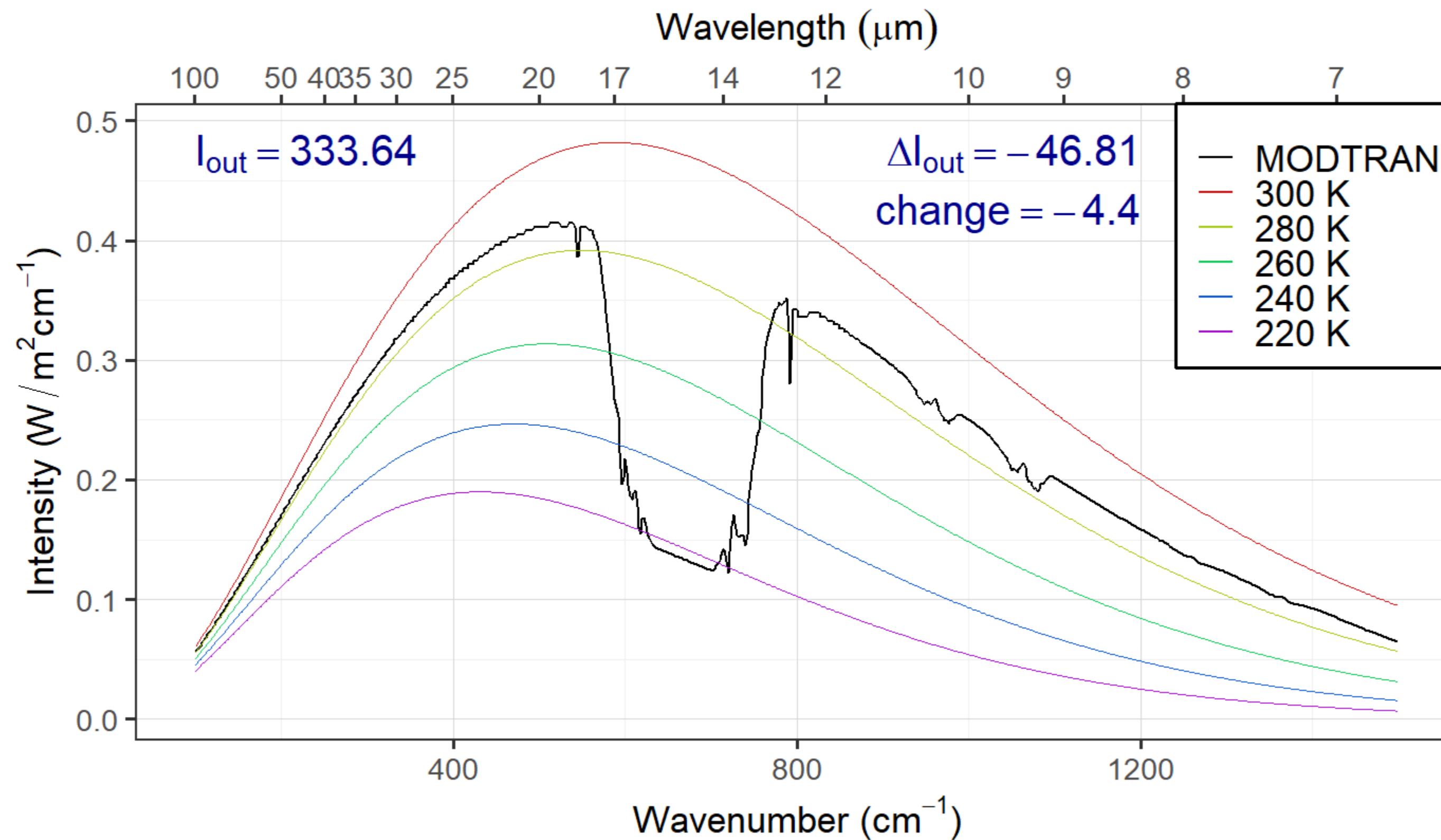
MODTRAN: 1024 ppm CO<sub>2</sub>, 20 km

Wavelength ( $\mu\text{m}$ )

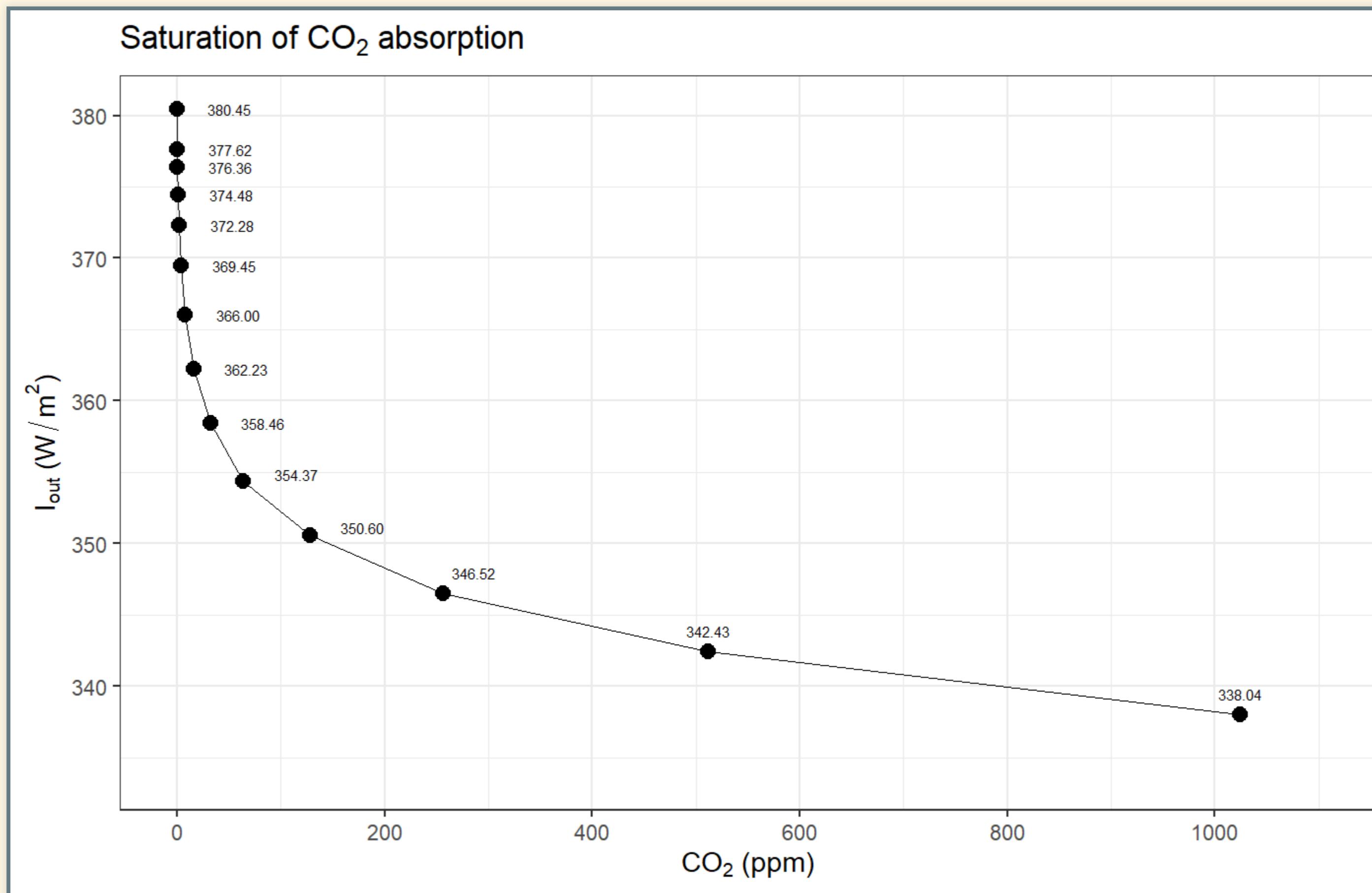


# 2048 ppm CO<sub>2</sub>

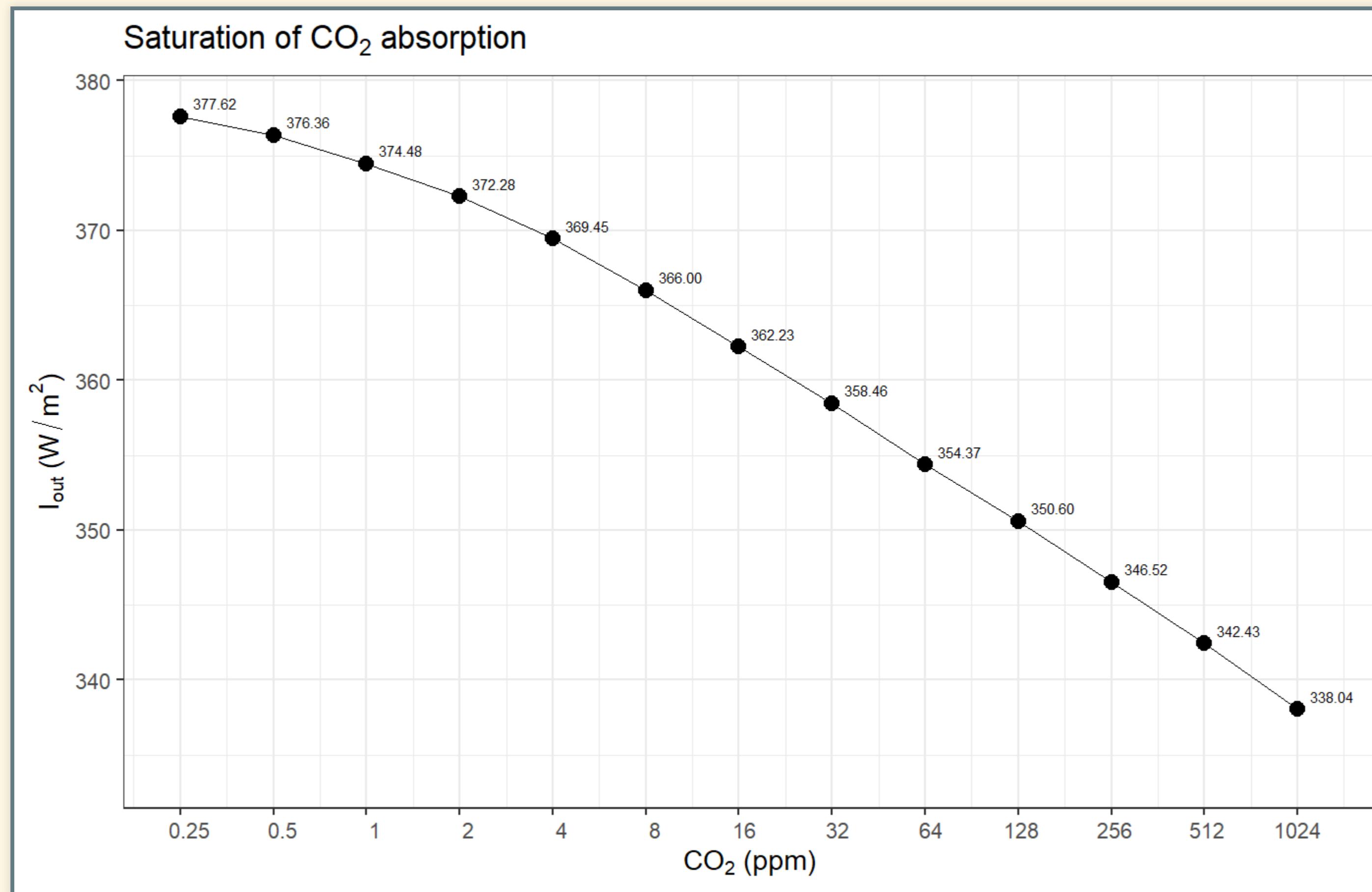
MODTRAN: 2048 ppm CO<sub>2</sub>, 20 km



# Band Saturation ( $I_{out}$ )

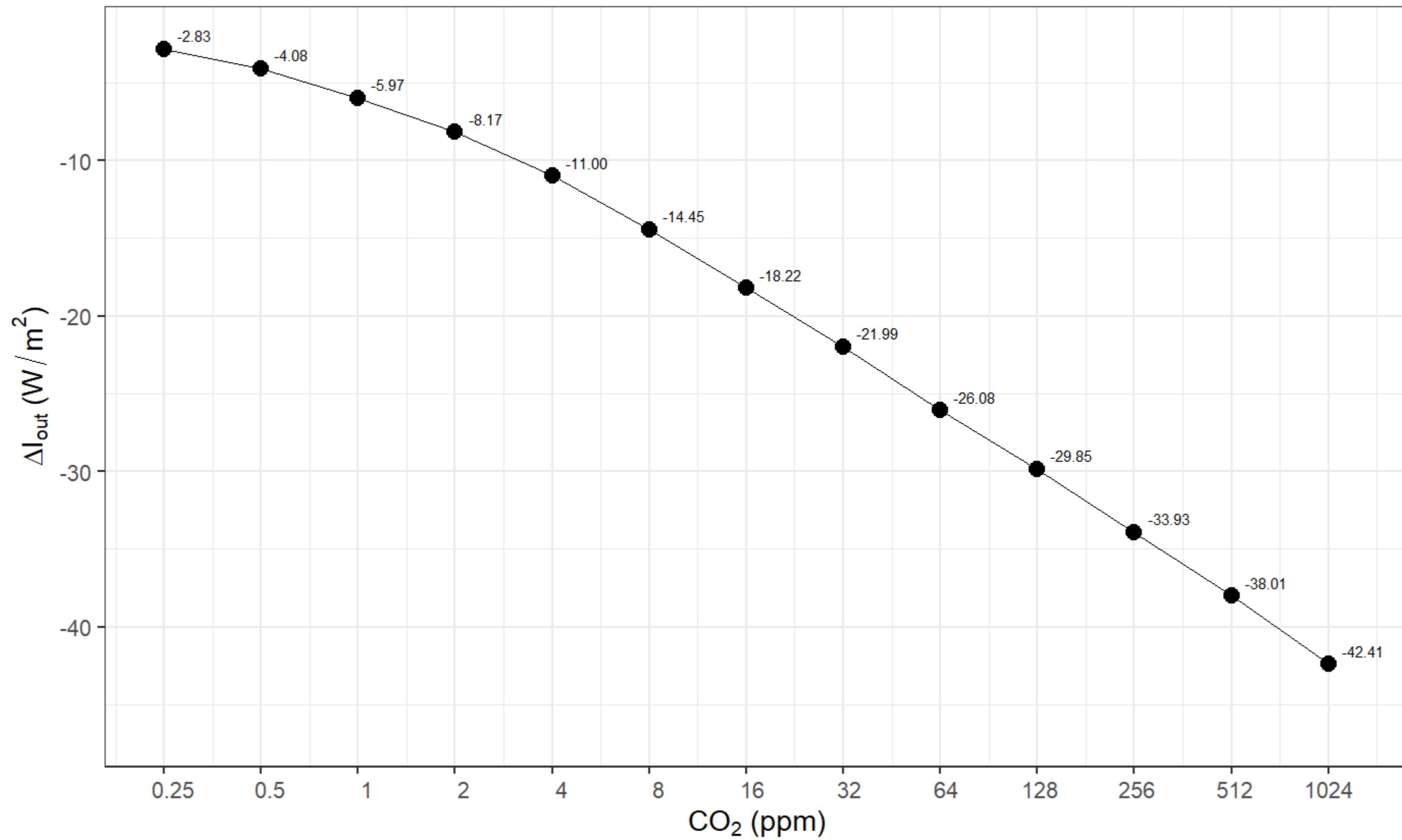


# $I_{\text{out}}$ ( $\text{CO}_2$ on log scale)



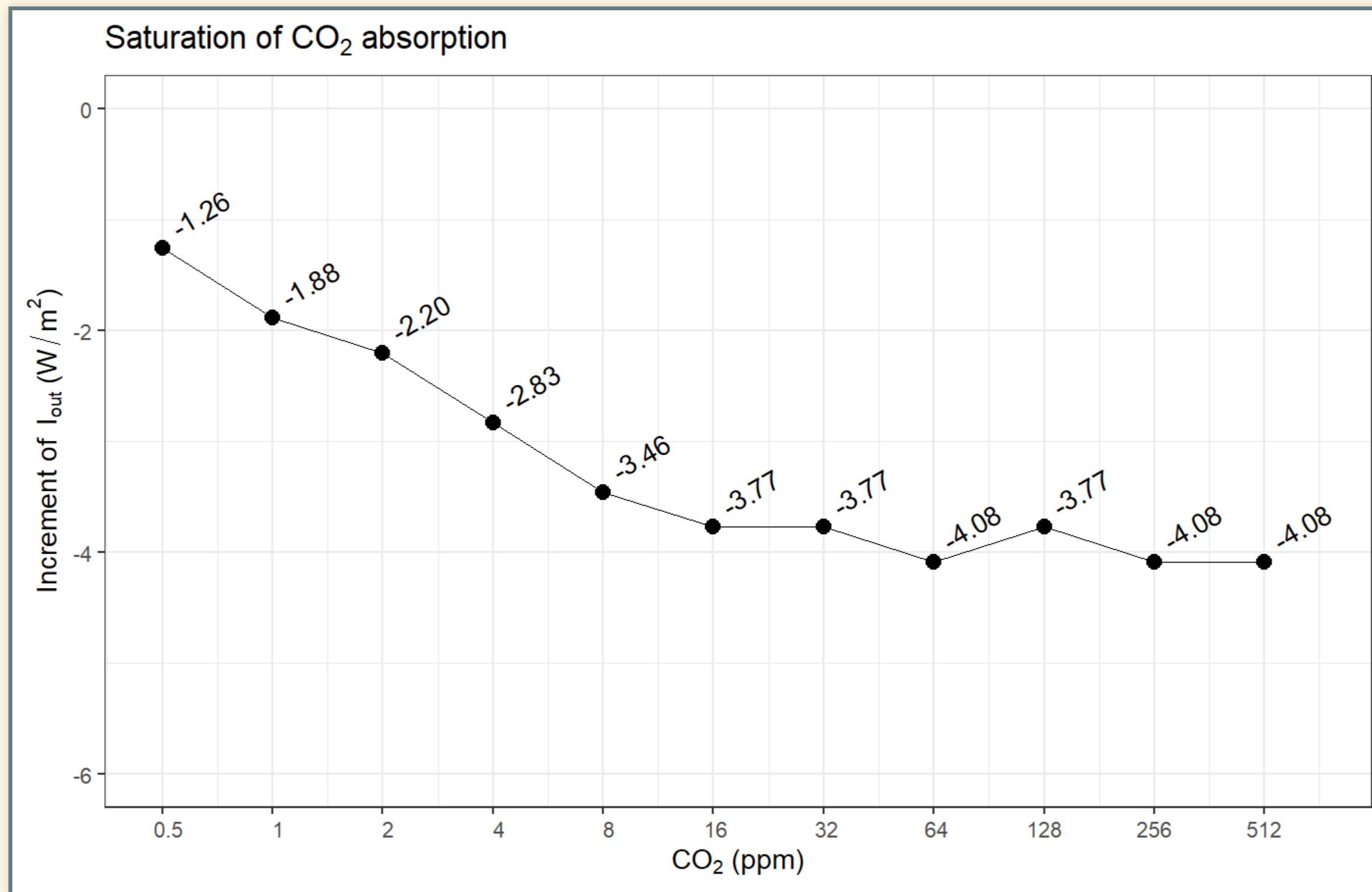
# $\Delta I_{\text{out}}$

## Saturation of CO<sub>2</sub> absorption



# Change in $I_{\text{out}}$ from no CO<sub>2</sub>

# Increments of $I_{\text{out}}$



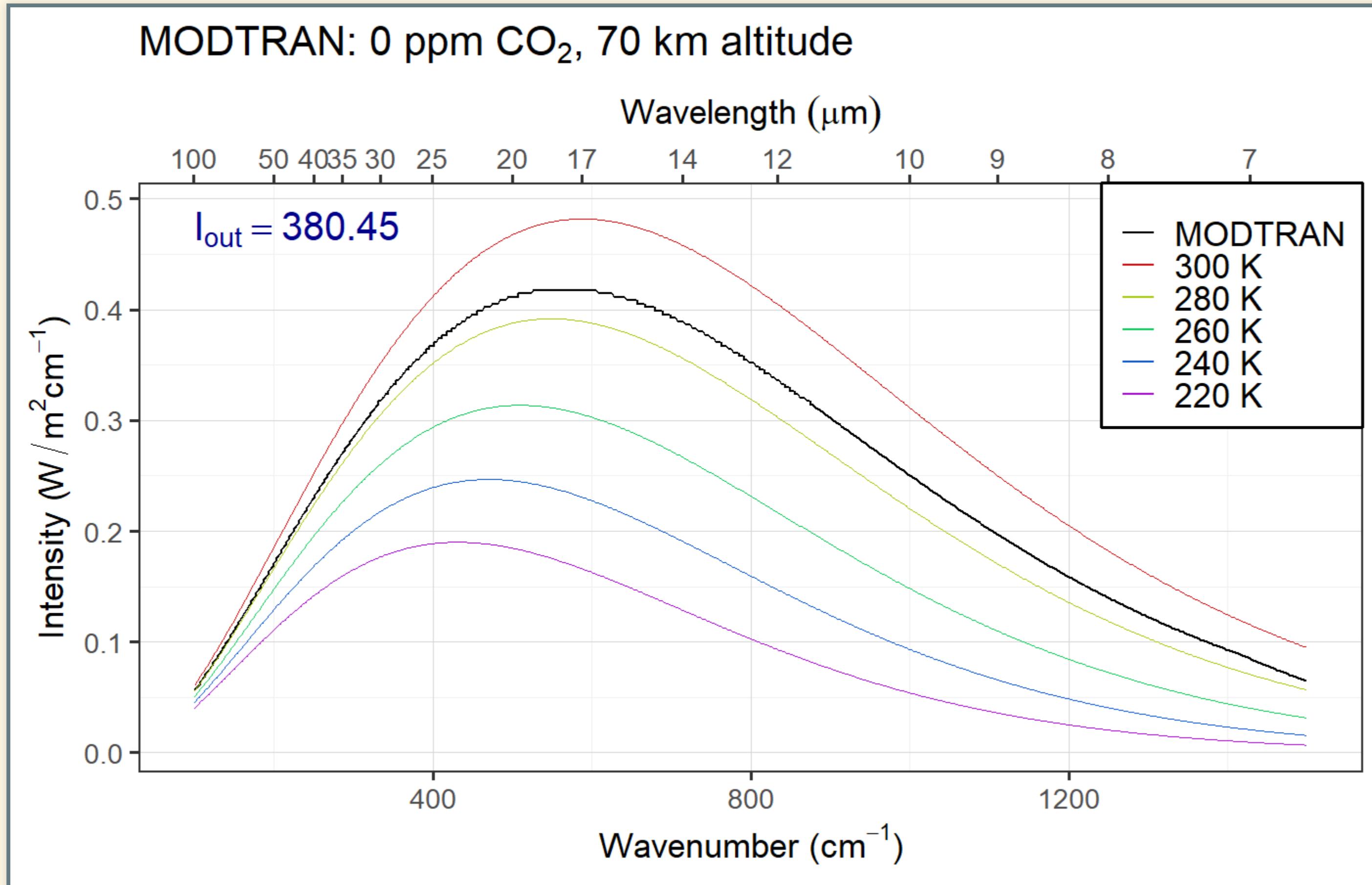
Change in  $I_{\text{out}}$  from previous  $I_{\text{out}}$

# Measuring Greenhouse Effect:

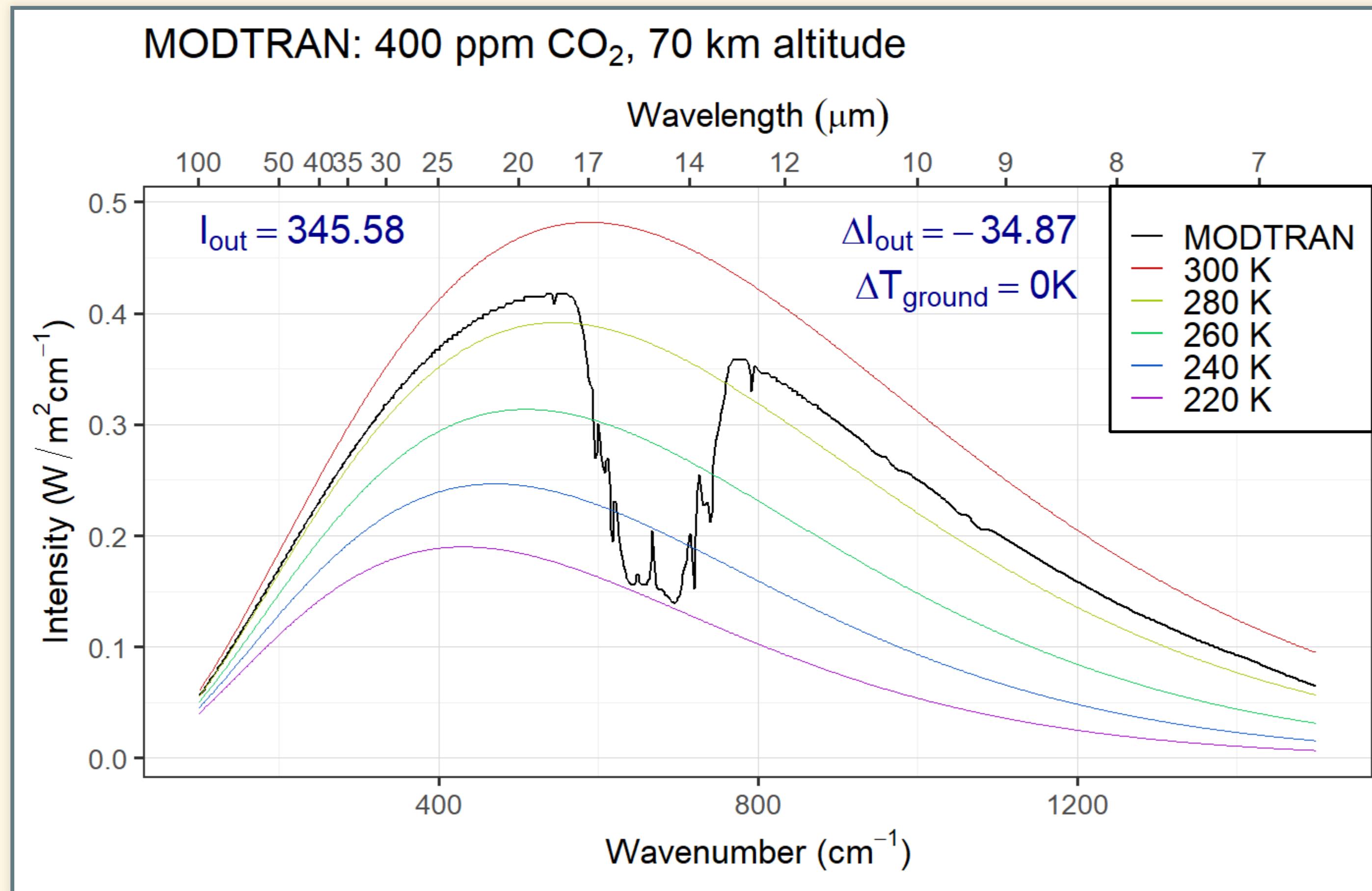
# Measuring Greenhouse Effect:

- Go to MODTRAN, set CO<sub>2</sub> to 0 ppm, and set all other gases to zero.
- Set altitude to 70 km and location to “1976 US Standard Atmosphere”.
- Press “Save this run to background”
- Note  $I_{\text{out}}$
- Set CO<sub>2</sub> to 400 ppm and note the change in  $I_{\text{out}}$
- Adjust the temperature offset to make the difference in  $I_{\text{out}}(\text{New} - \text{BG})$  equal zero.

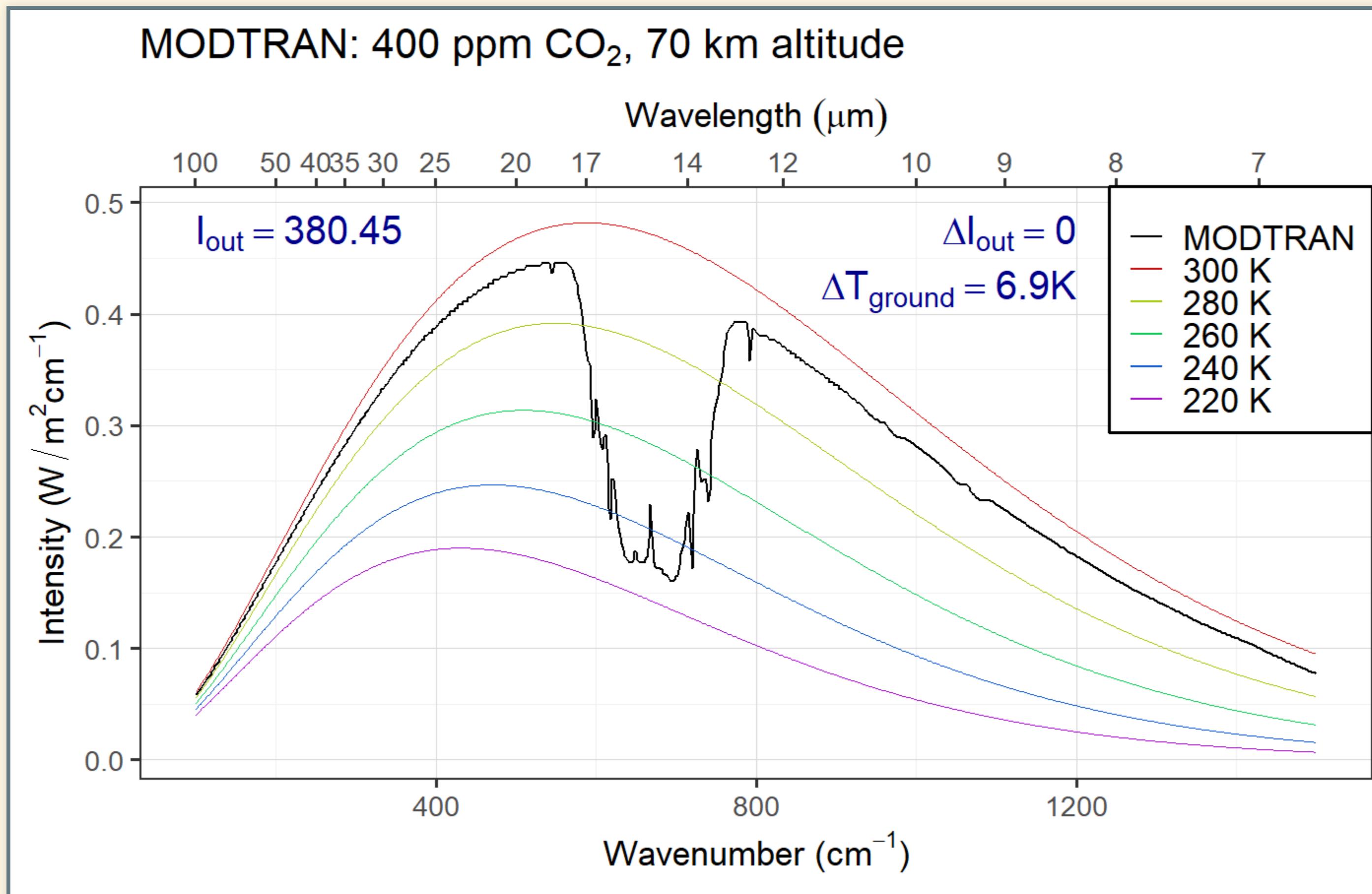
# No Greenhouse Gases



# 400 ppm



# Adjust temperature



# Calculating Global Warming

# Calculating Global Warming

- “Climate sensitivity” =  $\Delta T_{2x}$ 
  - Temperature rise for doubled CO<sub>2</sub>.
  - Uncertain (because of feedbacks)
  - Best estimate:  $\Delta T_{2x} \sim 3.2\text{K}$  (range 2.0–4.5 K)
- Every time you double CO<sub>2</sub>,  $T$  rises by  $\Delta T_{2x}$ .
- For arbitrary change in CO<sub>2</sub>:

$$\Delta T = \Delta T_{2x} \times \frac{\ln\left(\frac{\text{new } p\text{CO}_2}{\text{old } p\text{CO}_2}\right)}{\ln 2}$$

# Global Warming Potential

- Absorption by CO<sub>2</sub> and water vapor are very saturated
- Absorption in the atmospheric window is not saturated
- Therefore, molecule-for-molecule, gases that absorb in the window have a much bigger effect on the climate than adding more CO<sub>2</sub>.
  - One chlorofluorocarbon molecule = thousands of CO<sub>2</sub> molecules
- Global Warming Potential (GWP) of x = how many CO<sub>2</sub> molecules cause the same warming as one molecule of x

# Evolving theory of greenhouse effect

# Greenhouse effect

## 1. Purely radiative (no convection)

- Each layer has uniform temperature
  - a. Single-layer, uniform spectrum ([Mon. 1/13](#))
    - Absorbs 100% longwave light
  - b. Multi-layer, uniform spectrum ([Lab #2](#))
    - More layers  $\Rightarrow$  greater greenhouse effect.
  - c. Realistic spectrum ([Wed. 1/15 & today](#))
    - More realistic
    - Harder to do calculations (need computer)

## 2. Introduce convection ([Today & Wednesday](#))

- Temperature changes with height
- Convection moves heat up and down
- Radiative-convective models are very accurate
  - But require computers

# The Vertical Structure of the Atmosphere

# Greenhouse effect

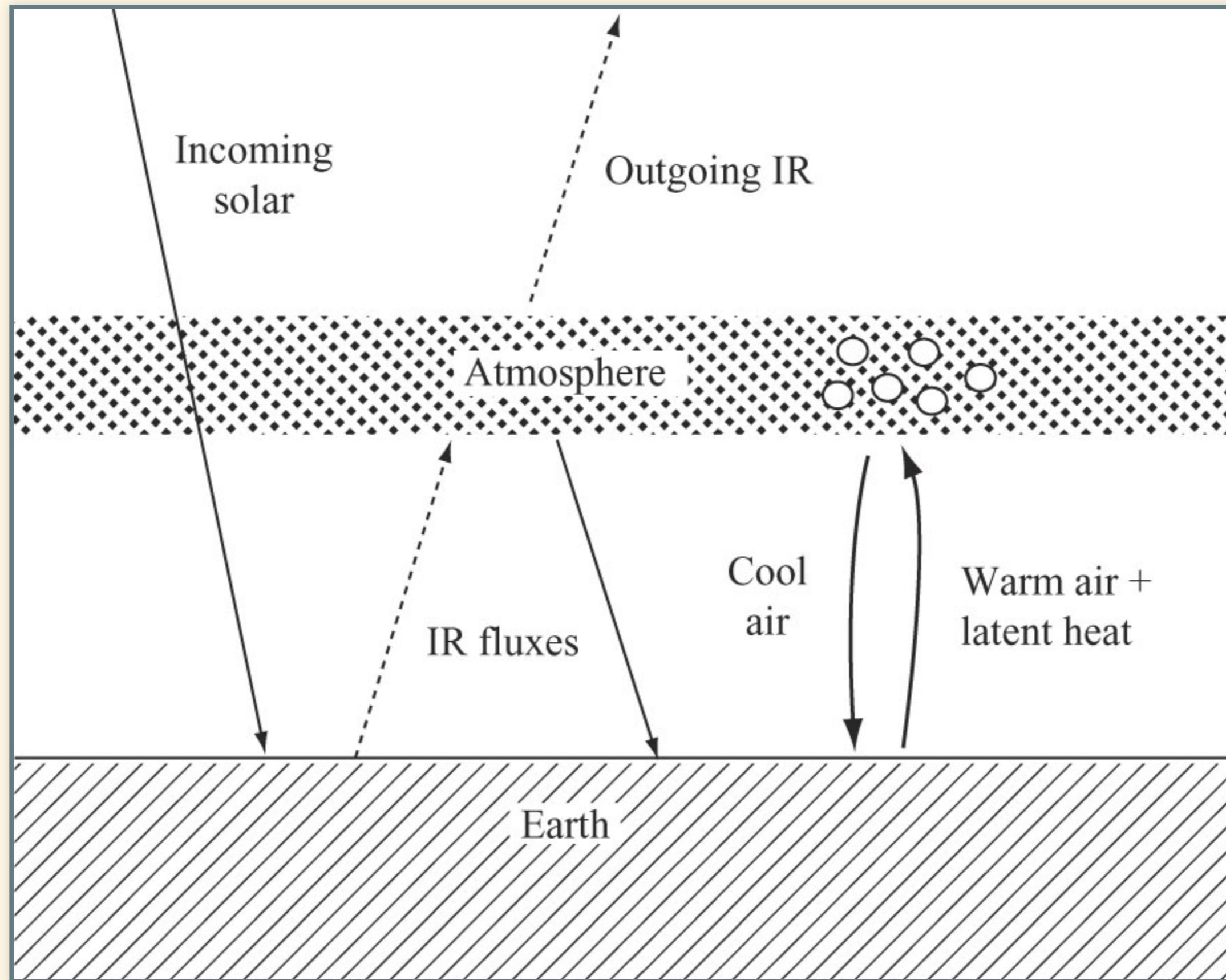
## 1. Purely radiative (no convection)

- Each layer has uniform temperature
  - a. Single-layer, uniform spectrum
    - Absorbs 100% longwave light
  - b. Multi-layer, uniform spectrum
    - More layers  $\Rightarrow$  greater greenhouse effect.
  - c. Realistic spectrum
    - More realistic
    - Harder to do calculations (need computer)

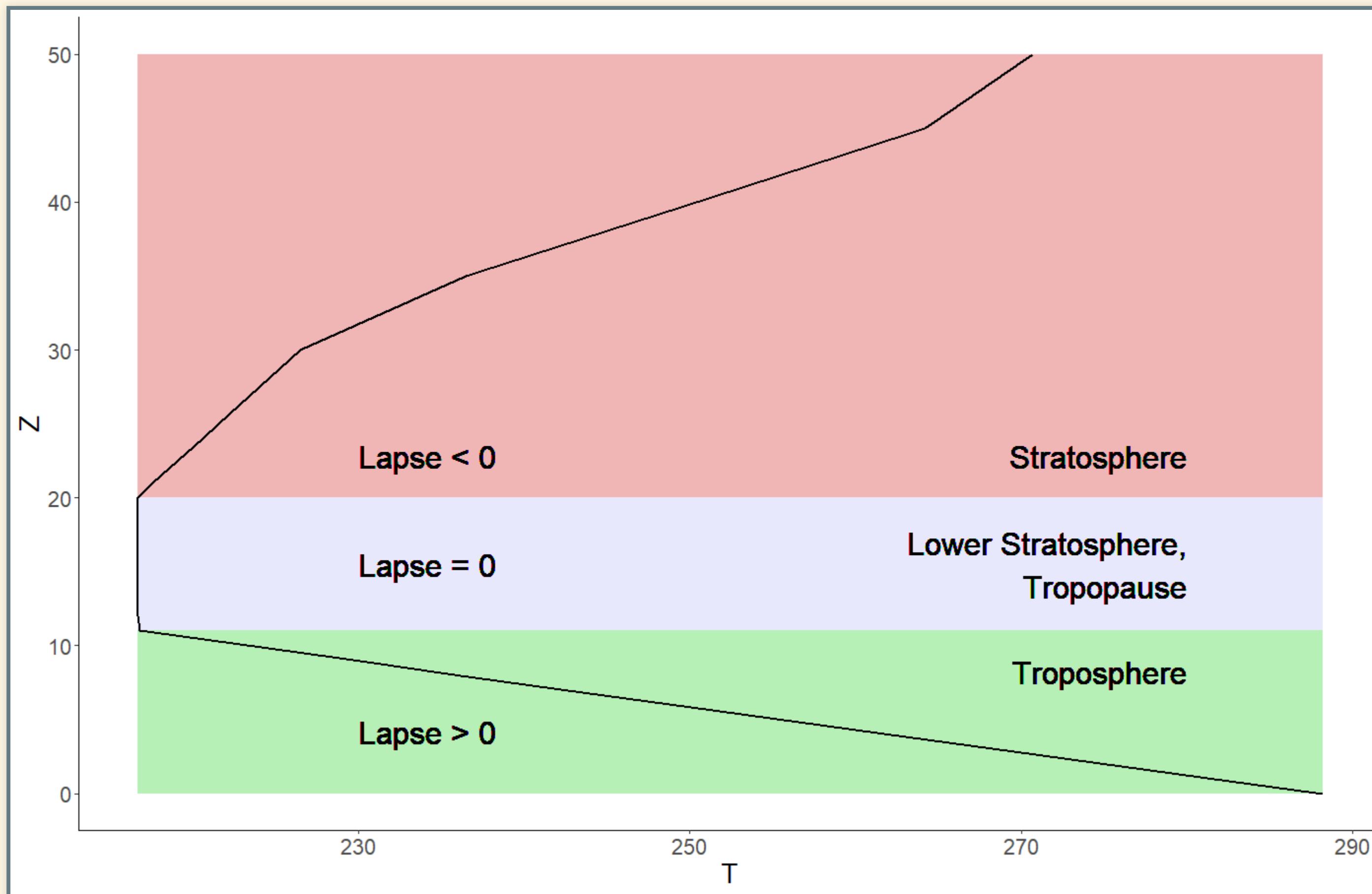
## 2. Convection:

- Temperature changes with height
- Convection moves heat up and down
- Radiative-convective models are very accurate
  - But require computers

# Radiative-Conductive Equilibrium



# Normal Atmosphere:

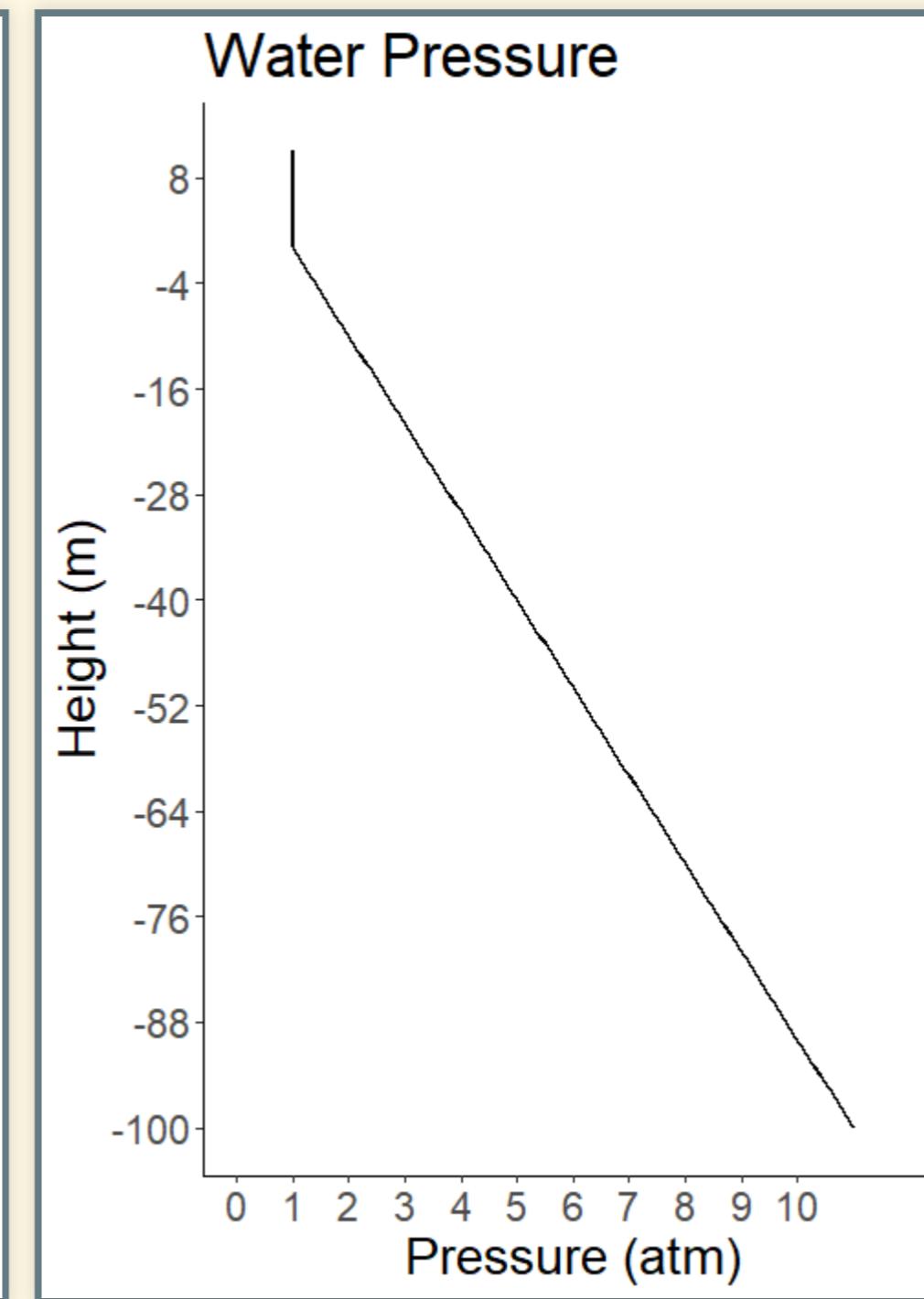
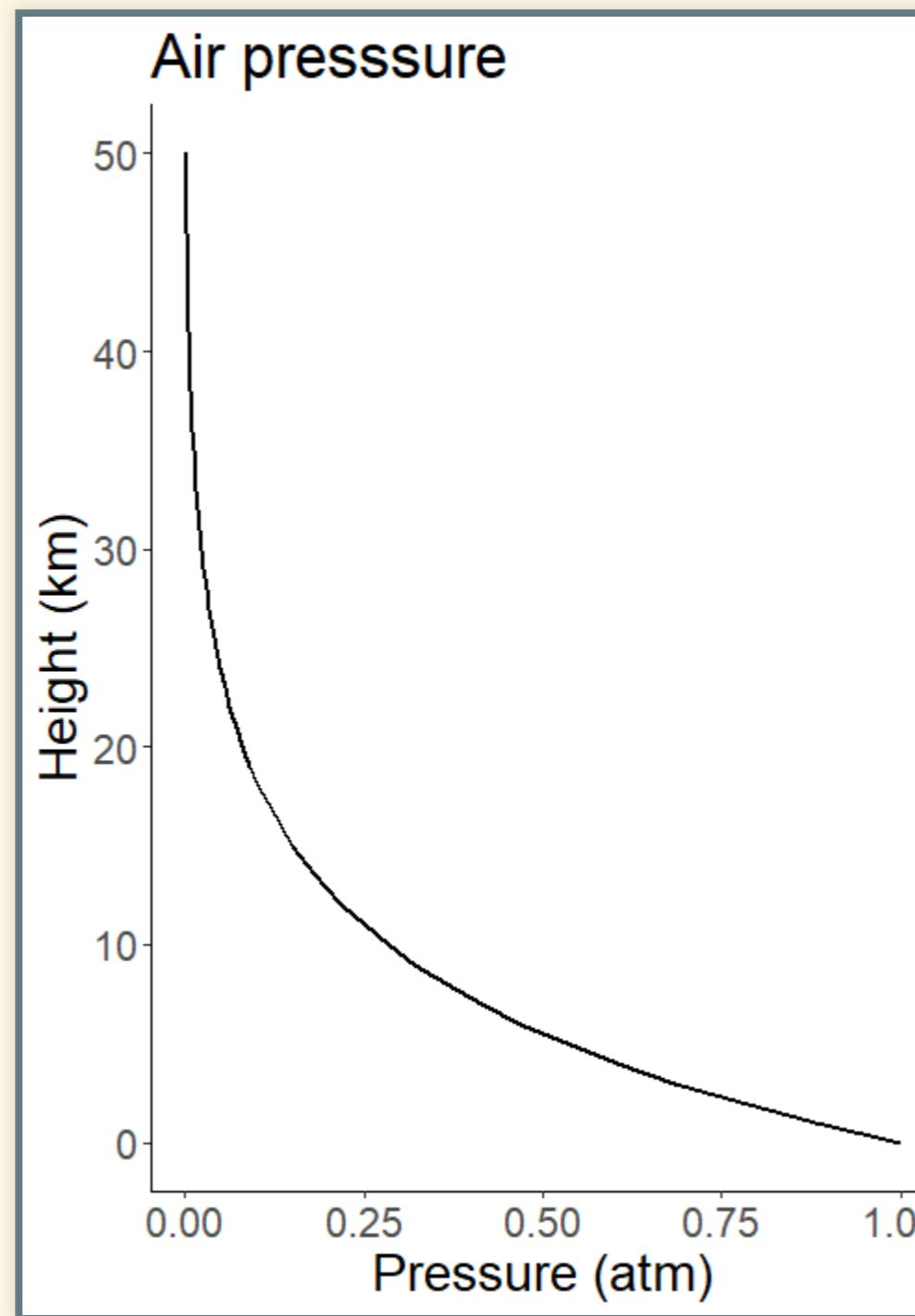


# Vertical Structure

$$\text{Lapse rate} = \frac{-\Delta T}{\Delta \text{height}}$$

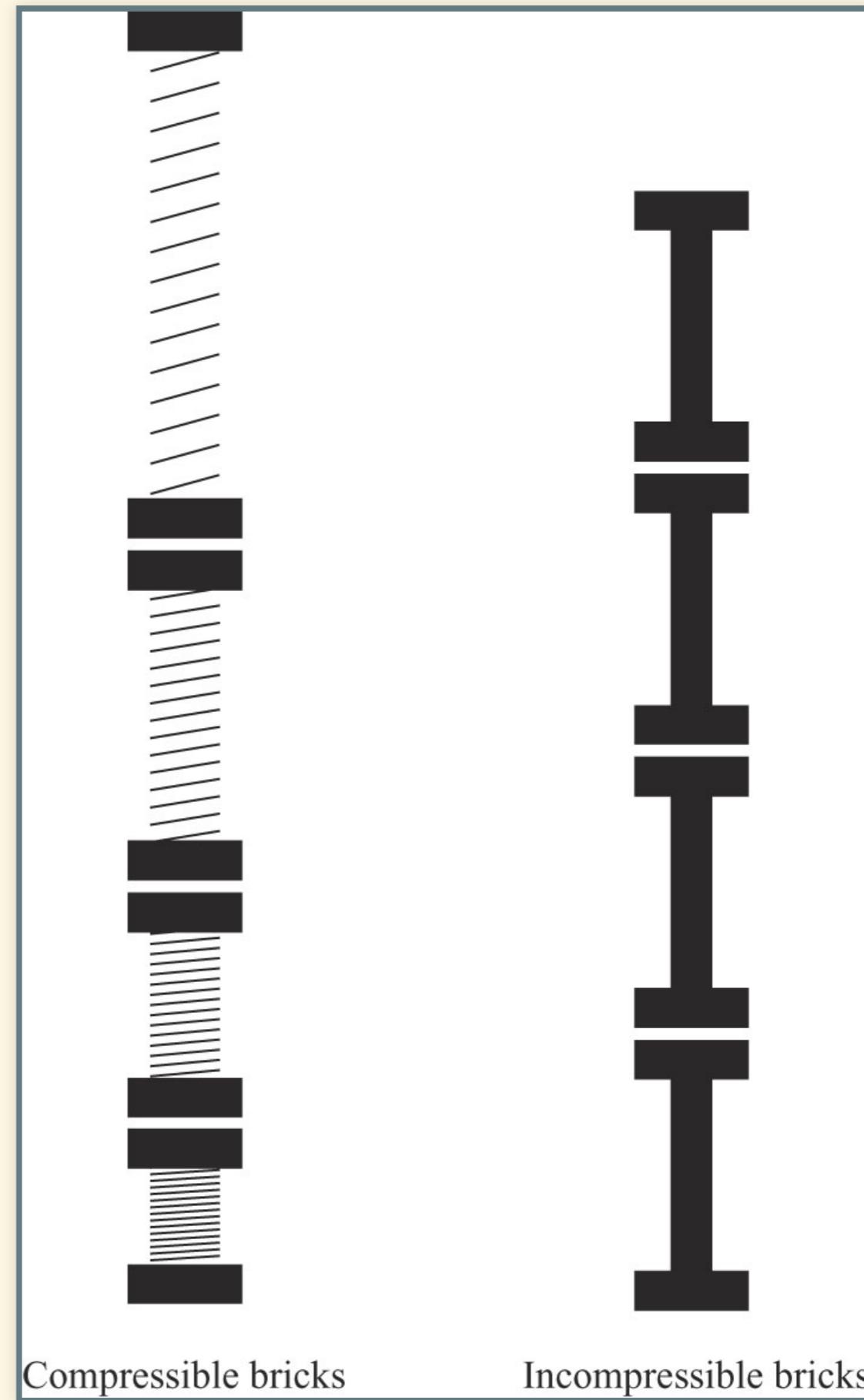
- Positive lapse rate: Air overhead is cooler  
(normal for troposphere)
- Negative lapse rate: Air overhead is warmer  
(abnormal, “inversion”)

# Air vs. Water



# Air vs. Water

- Pressure = weight of everything overhead.
- Air is compressible, water isn't.
- 1 meter height of water weighs  $1000 \text{ kg/m}^2$
- 1 meter height of dry air at sea-level density weighs  $1.3 \text{ kg/m}^2$
- 1 m height of dry air 10 km above sea level weighs  $0.4 \text{ kg/m}^2$

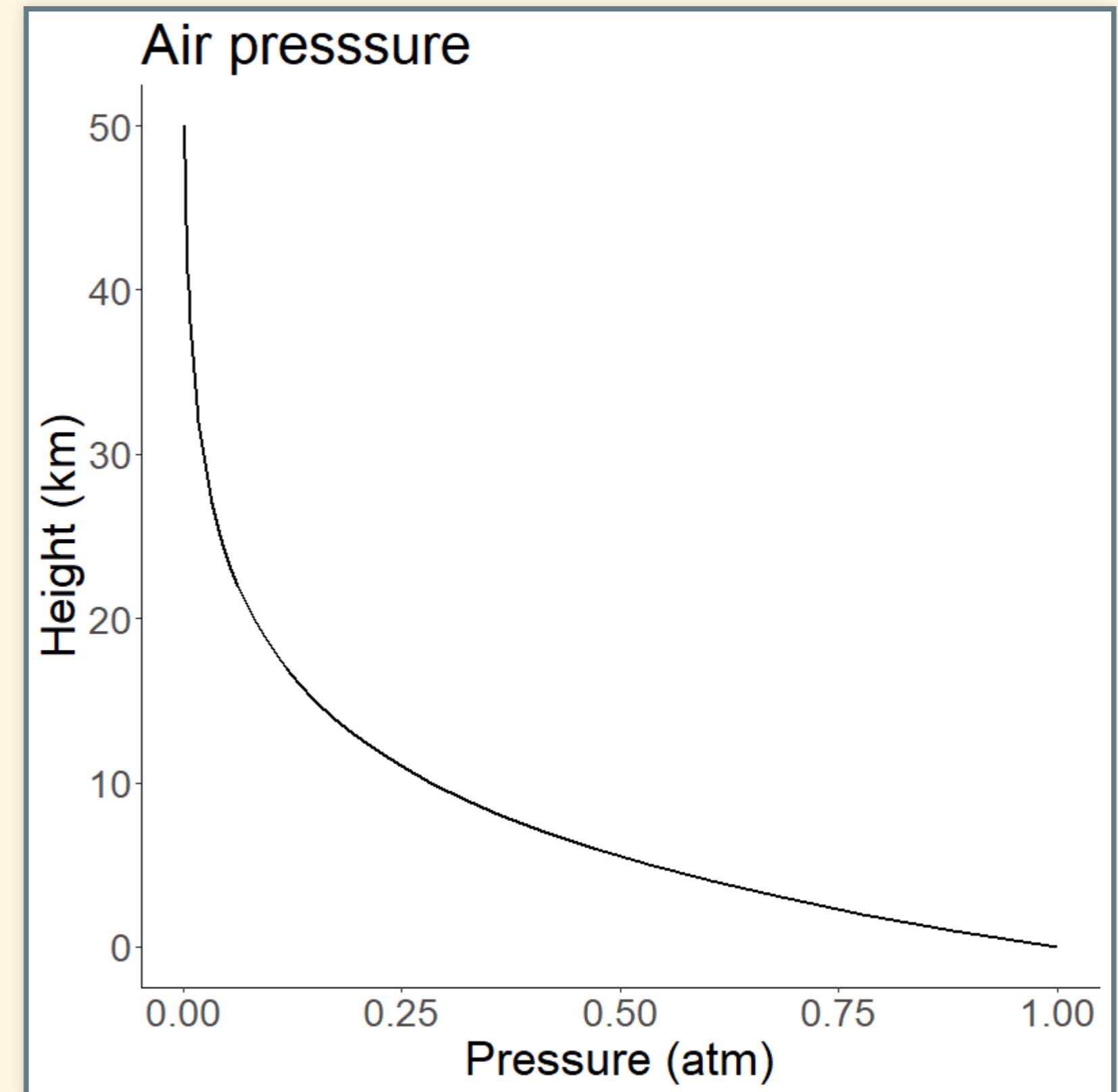


# Air Pressure

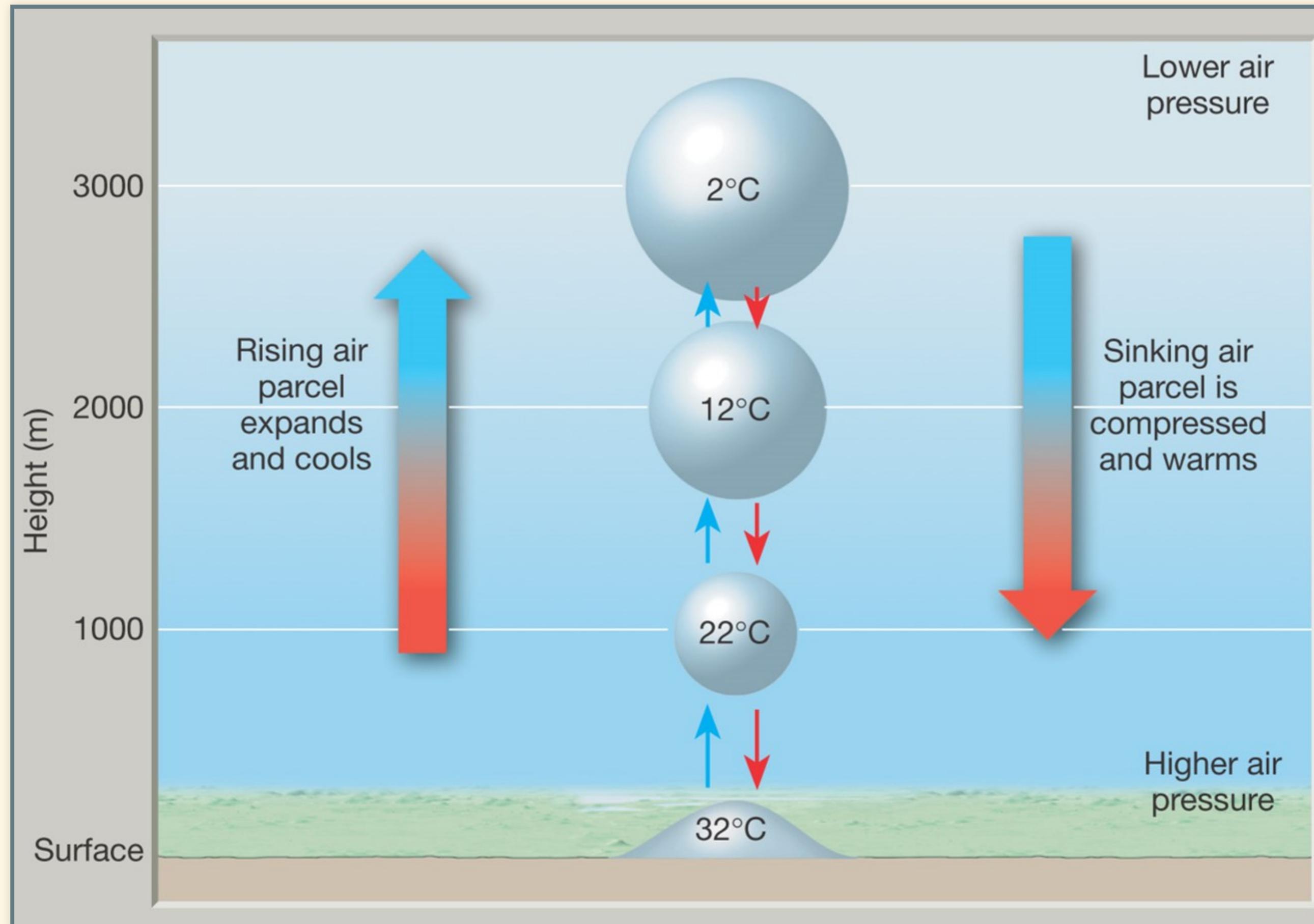
- Pressure at height  $h$ :

$$\begin{aligned}P(h) &= P_0 e^{-h/8.0\text{km}} \\&= P_0 2^{-h/5.5\text{km}} \\&= P_0 \left(\frac{1}{2}\right)^{h/5.5\text{km}}\end{aligned}$$

- Half the air is below 5.5 km.
- $3/4$  is below 11 km
- $7/8$  is below 16.5 km
- **NOTE:** The number 5.5 km is not exact, but it's consistent with the textbook.



# Why is the air cooler higher up?



# Terminology

- **Environmental Lapse**

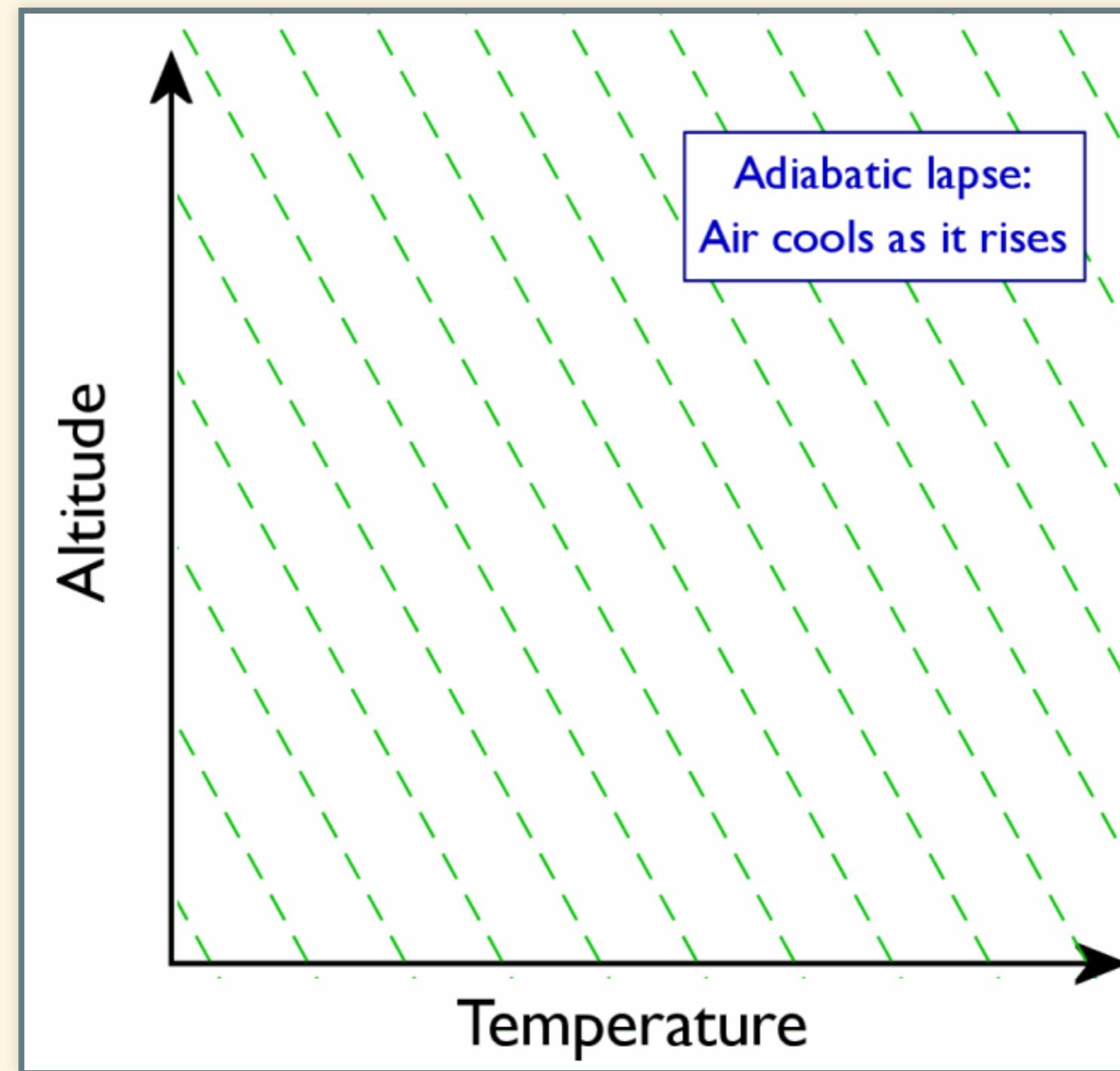
- Measured temperature of actual atmosphere
- Compares one bit of air at one height with another bit at another height.
- Changes from one time and place to another.

- **Adiabatic Lapse**

- Change in a single parcel of air as it moves up or down
- “**Adiabatic**” means no heat flowing in or out
  - **Adiabatic changes are reversible**
  - **Heat flow is irreversible**

# Overview of Convection

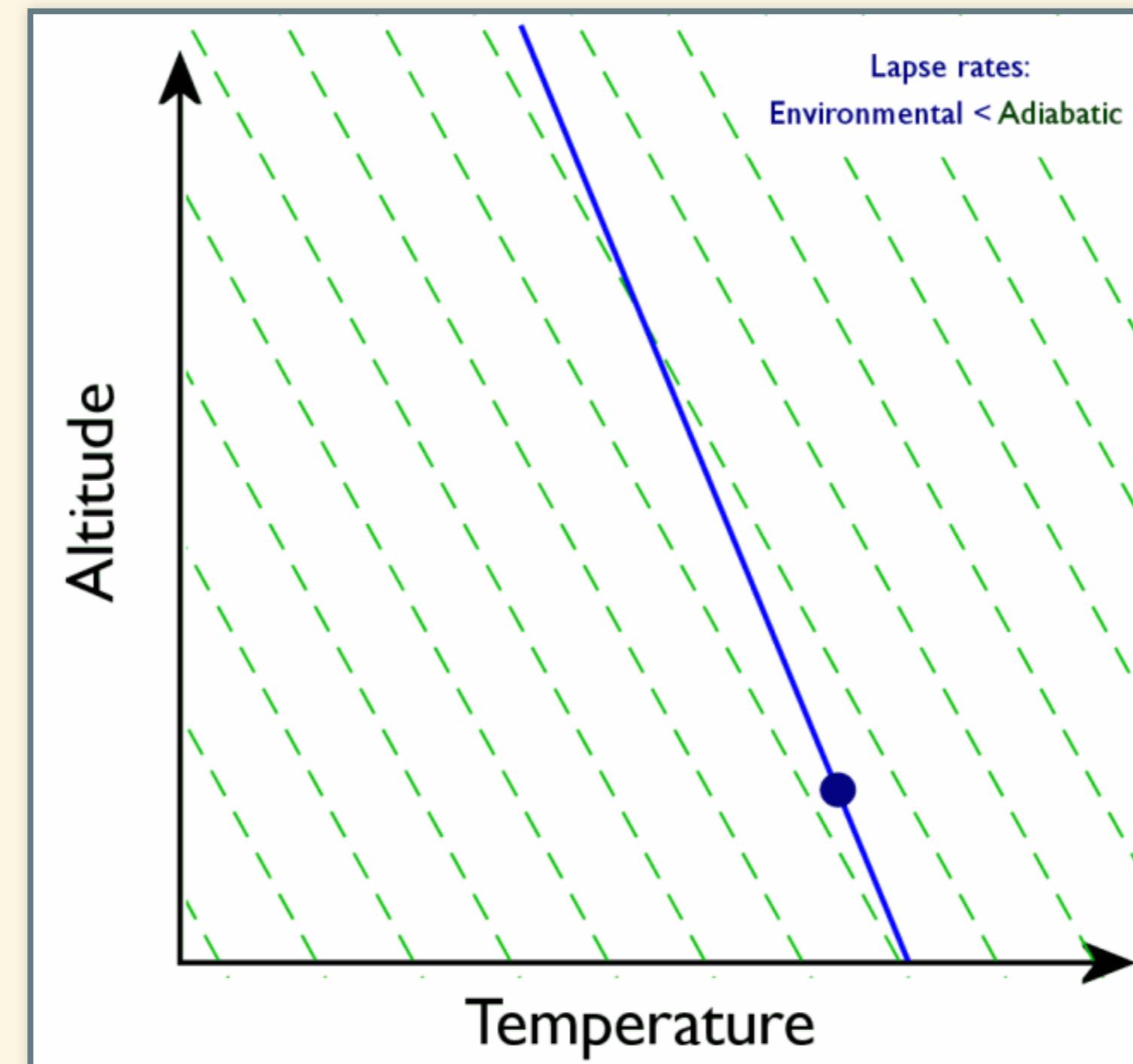
# Overview of convection



- Closer to vertical = smaller lapse rate (vertical = zero)
- Closer to horizontal = larger lapse rate

# Stable Atmosphere

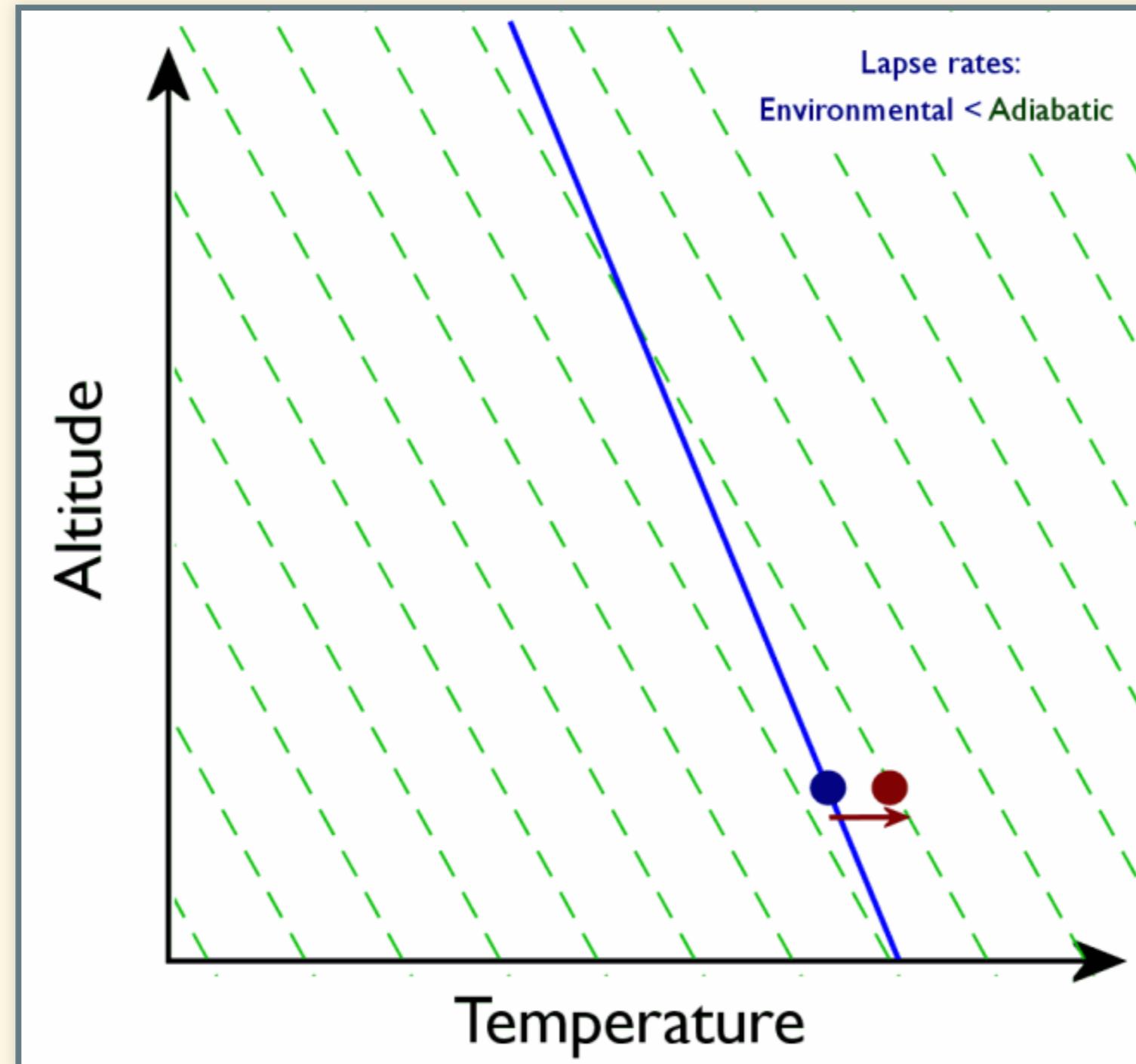
## Initial State



- green = adiabatic lapse
- blue = environmental lapse < adiabatic

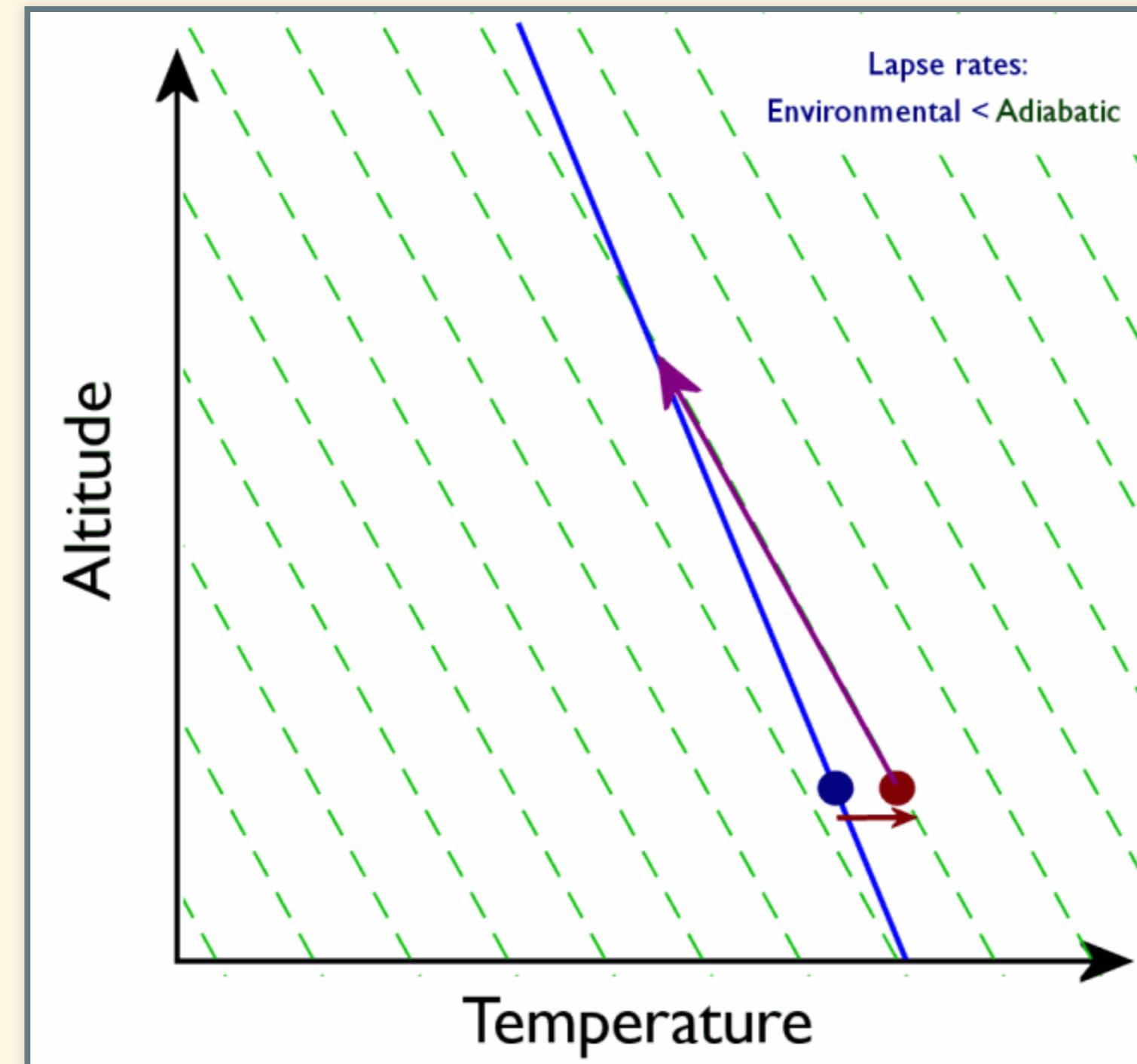
# Stable Atmosphere

## Parcel is heated



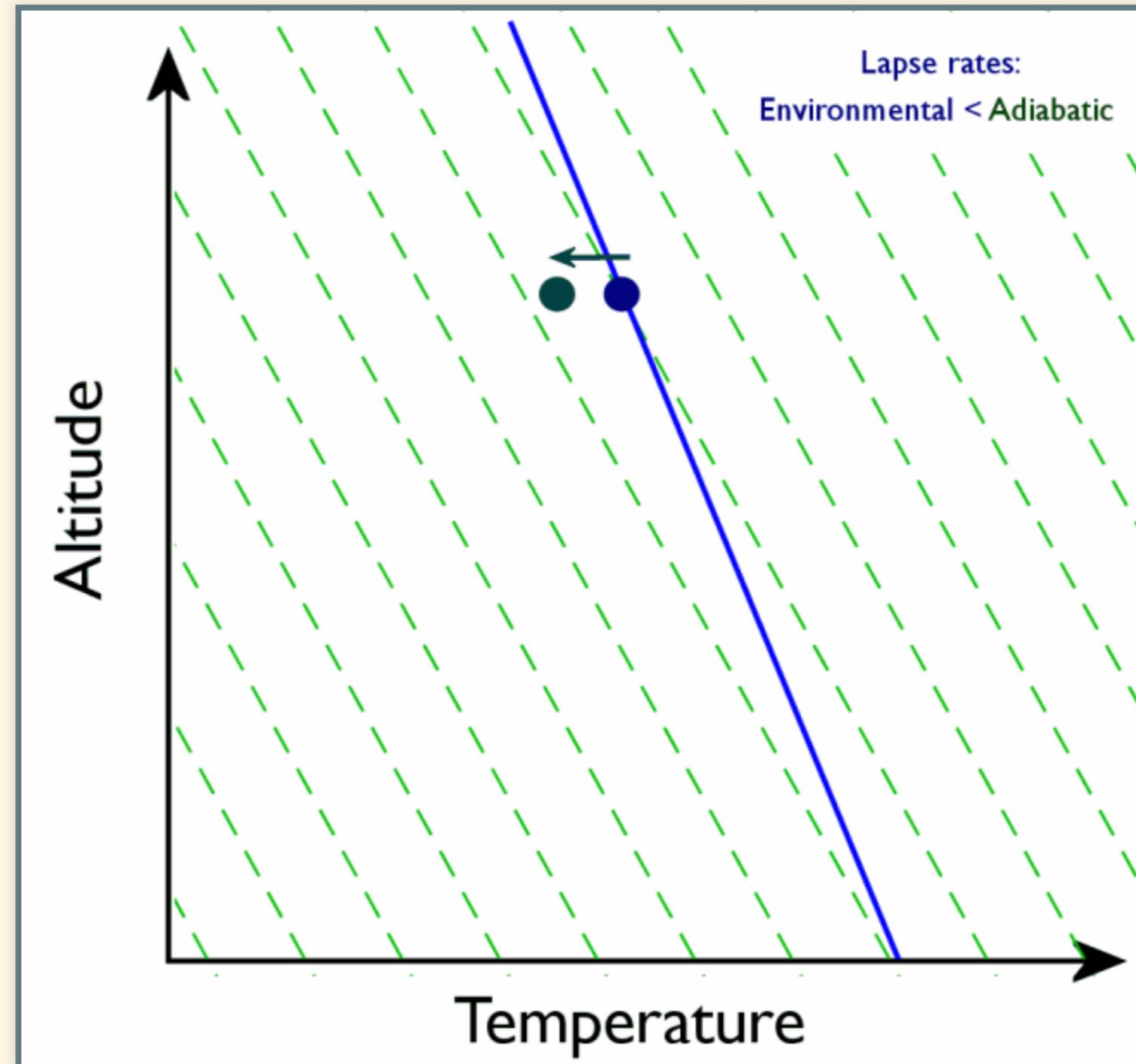
# Stable Atmosphere

## Rises to new equilibrium



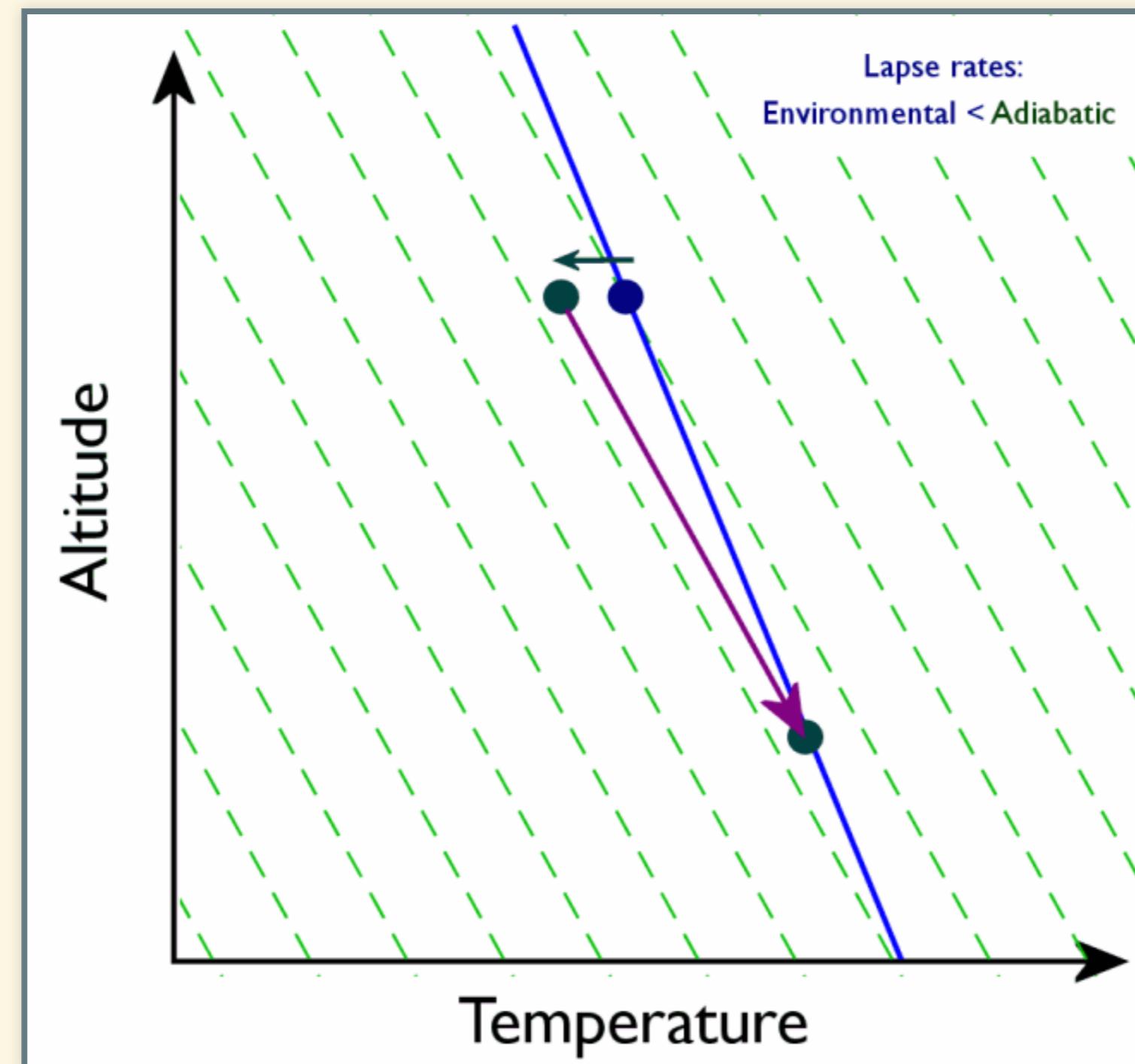
# Stable Atmosphere

## Parcel is cooled



# Stable Atmosphere

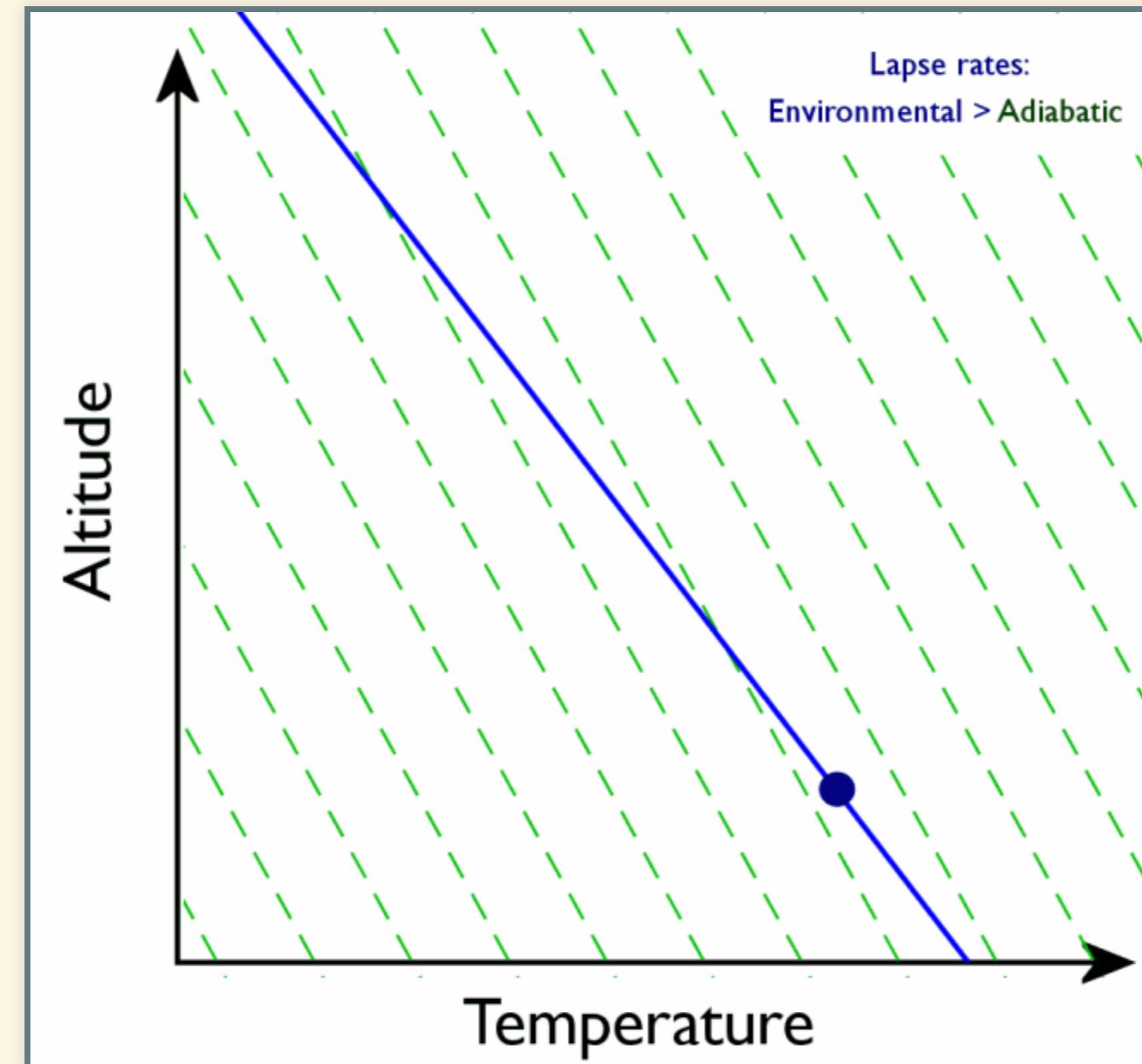
## Sinks to new equilibrium



# Unstable Atmosphere

# Unstable Atmosphere

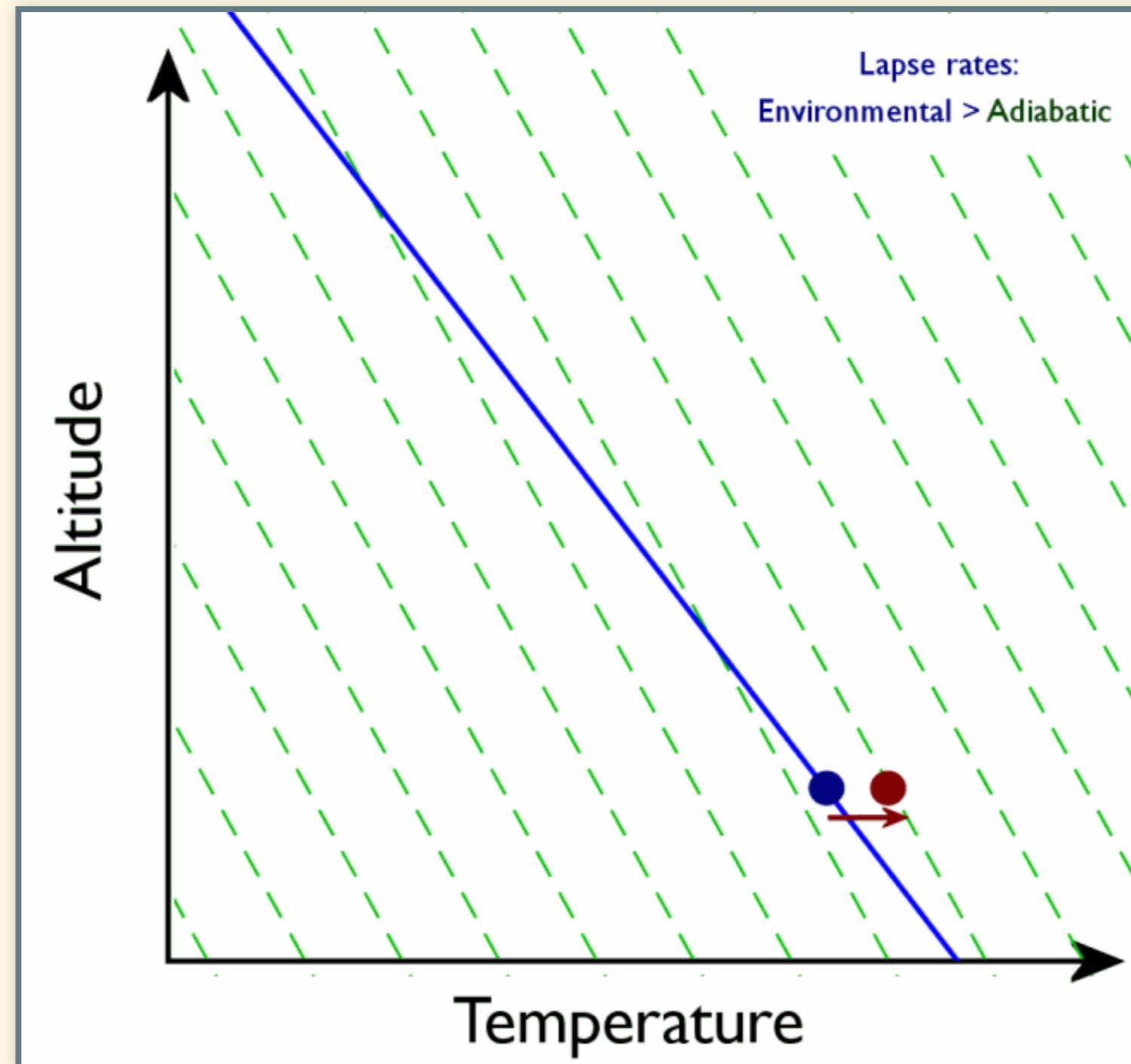
## Initial State



- green = adiabatic lapse
- blue = environmental lapse  $>$  adiabatic

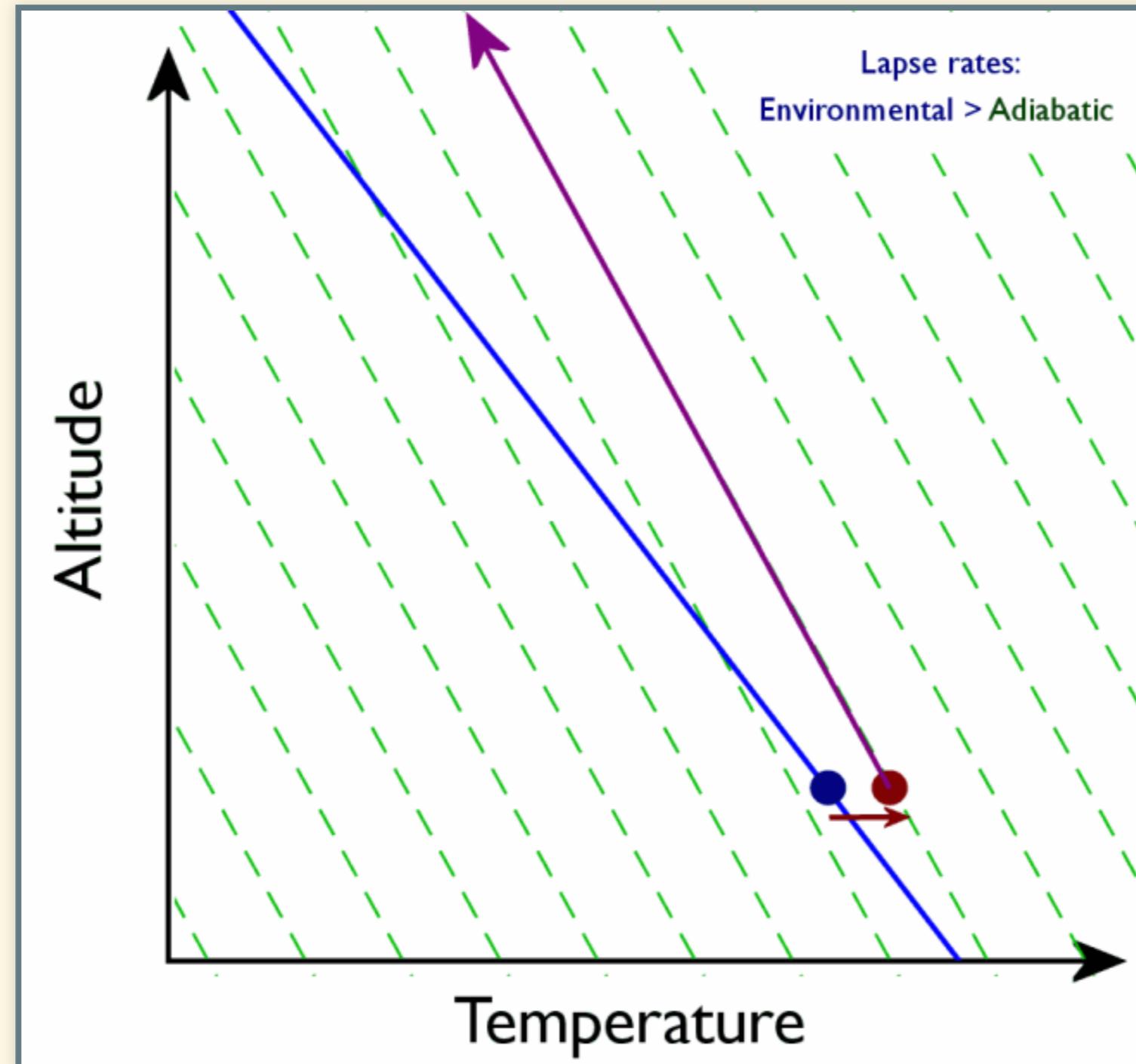
# Unstable Atmosphere

## Parcel is heated



# Unstable Atmosphere

## Rises without stopping



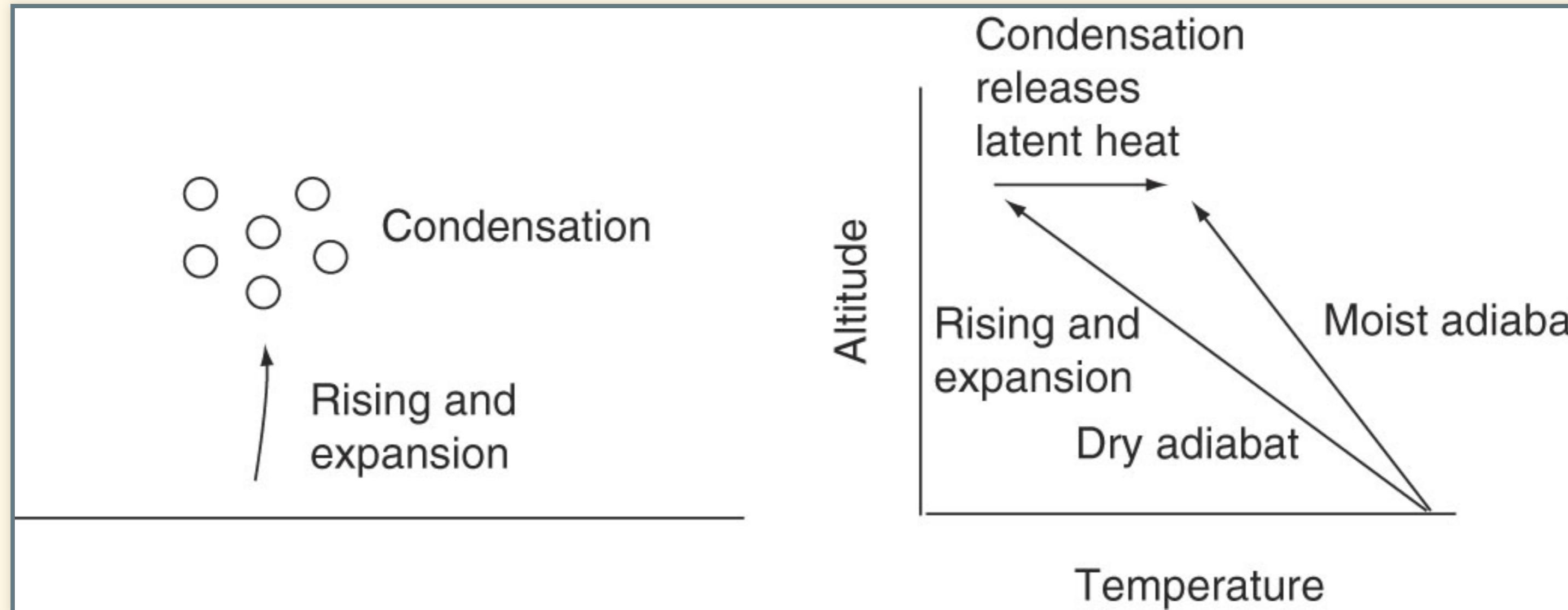
# Summary of Stability

# Summary of stability:

- Stable conditions:
  - Adiabatic Lapse > Environmental Lapse
- Unstable conditions:
  - Adiabatic Lapse < Environmental Lapse
- Why is stability important?
  - A stable atmosphere does not move heat around
  - An unstable atmosphere undergoes **convection**:
    - Hot air rises, cold air sinks
    - Redistributions heat

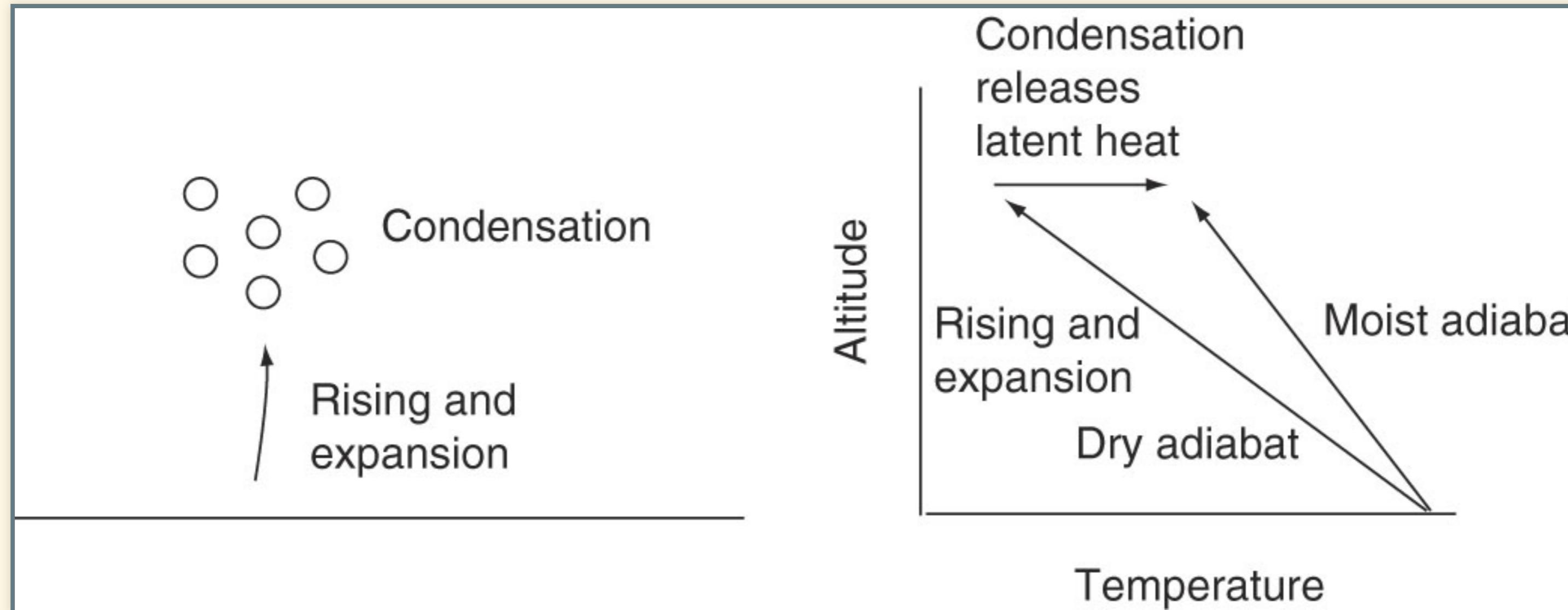
# Moist Convection

# Moist Convection



- Dry air rises and cools
- Cooling  $\Rightarrow$  water vapor condenses to liquid
- Condensation releases latent heat
- Latent heat warms air

# Moist Convection



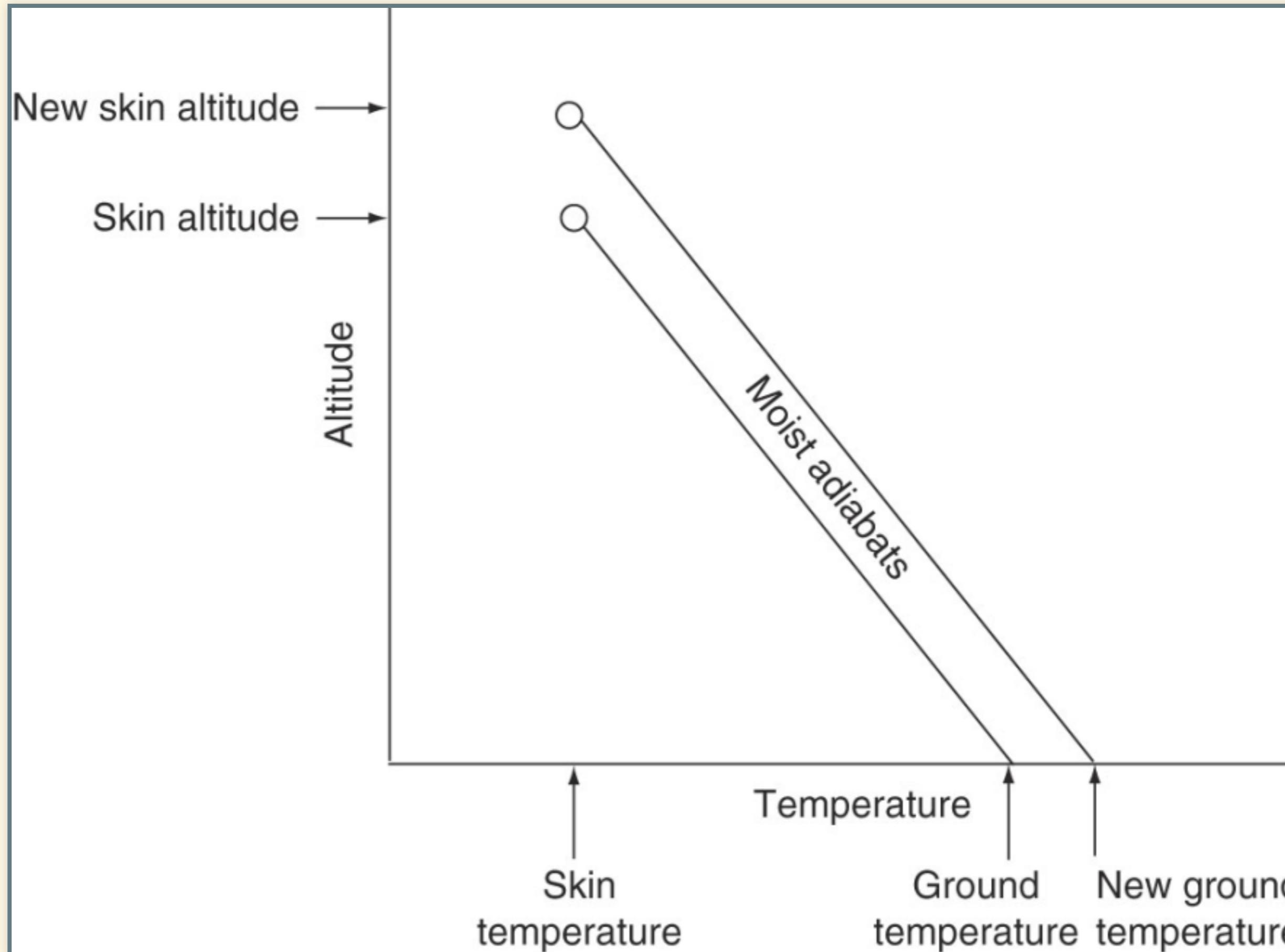
- Latent heat warms air
- Reduces adiabatic cooling
- Moist adiabatic lapse < Dry adiabatic lapse
- Smaller lapse = less stable
- **Humid air is less stable than dry air**

# Perspective

- Stable:
  - Environmental lapse  $\leq$  adiabatic lapse
- Unstable:
  - Environmental lapse  $>$  adiabatic lapse
- Adiabatic lapse:
  - Dry: 10 K/km
  - Moist: 4-8 K/km (depends on humidity)
- Pure radiative equilibrium:
  - Would produce lapse of **16 K/km**: unstable
- Radiative-Convective equilibrium:
  - Convection modifies environmental lapse
  - Normal environmental lapse is roughly **6 K/km**  
*(typical moist adiabatic lapse rate)*

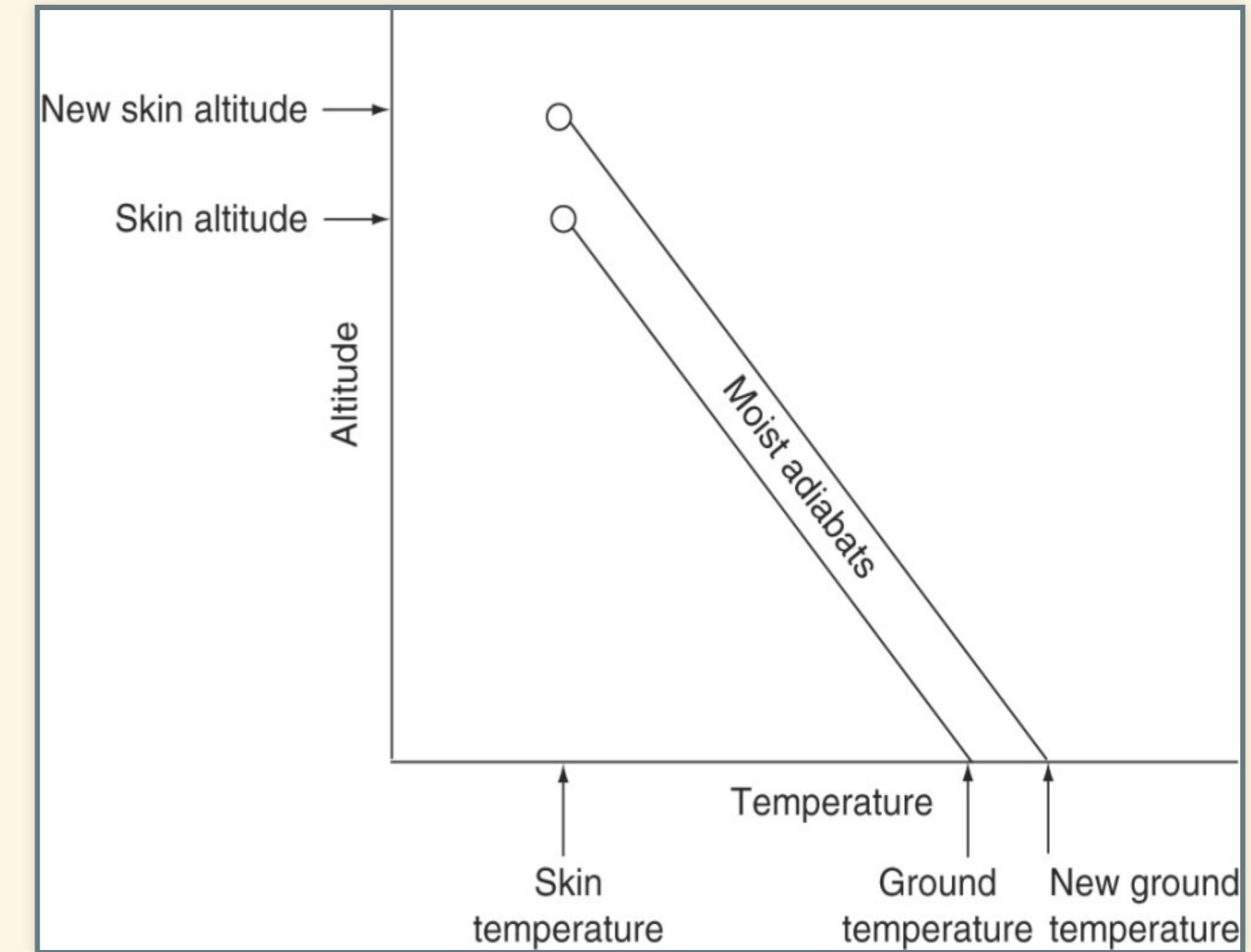
# Greenhouse effect

# Greenhouse effect



# Greenhouse effect

1.  $T_{\text{skin}} = 254 \text{ K}$
2.  $T_{\text{ground}} = T_{\text{skin}} + \text{lapse rate} \times h_{\text{skin}}$
3. Increase greenhouse gases
4. Skin height rises by  $\Delta h_{\text{skin}}$
5.  $T_{\text{ground}}$  rises by lapse rate  $\times \Delta h_{\text{skin}}$



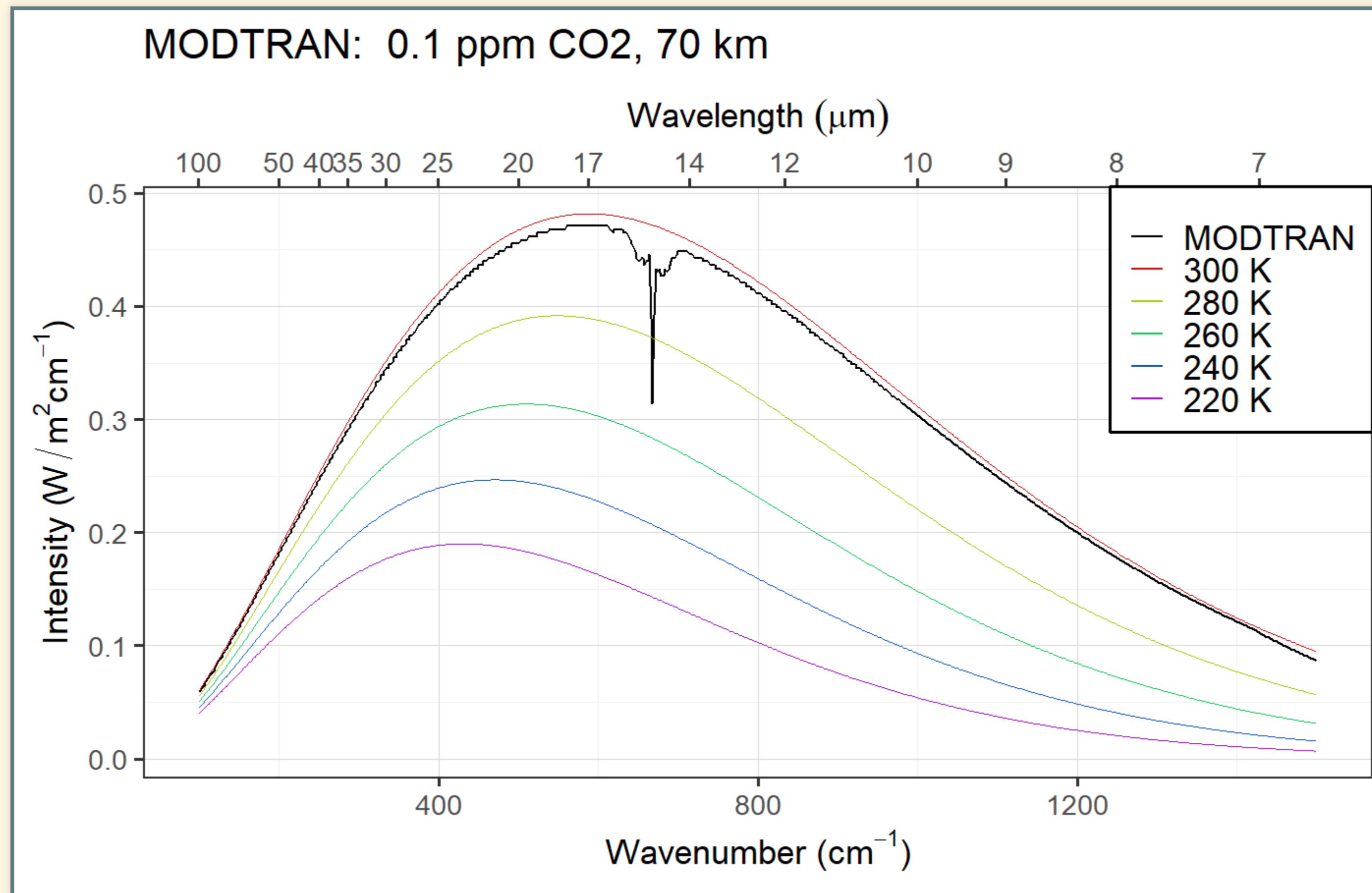
# Vertical Structure and Saturation

## Set up MODTRAN:

Go to MODTRAN (<http://climatedmodels.uchicago.edu/modtran/>)

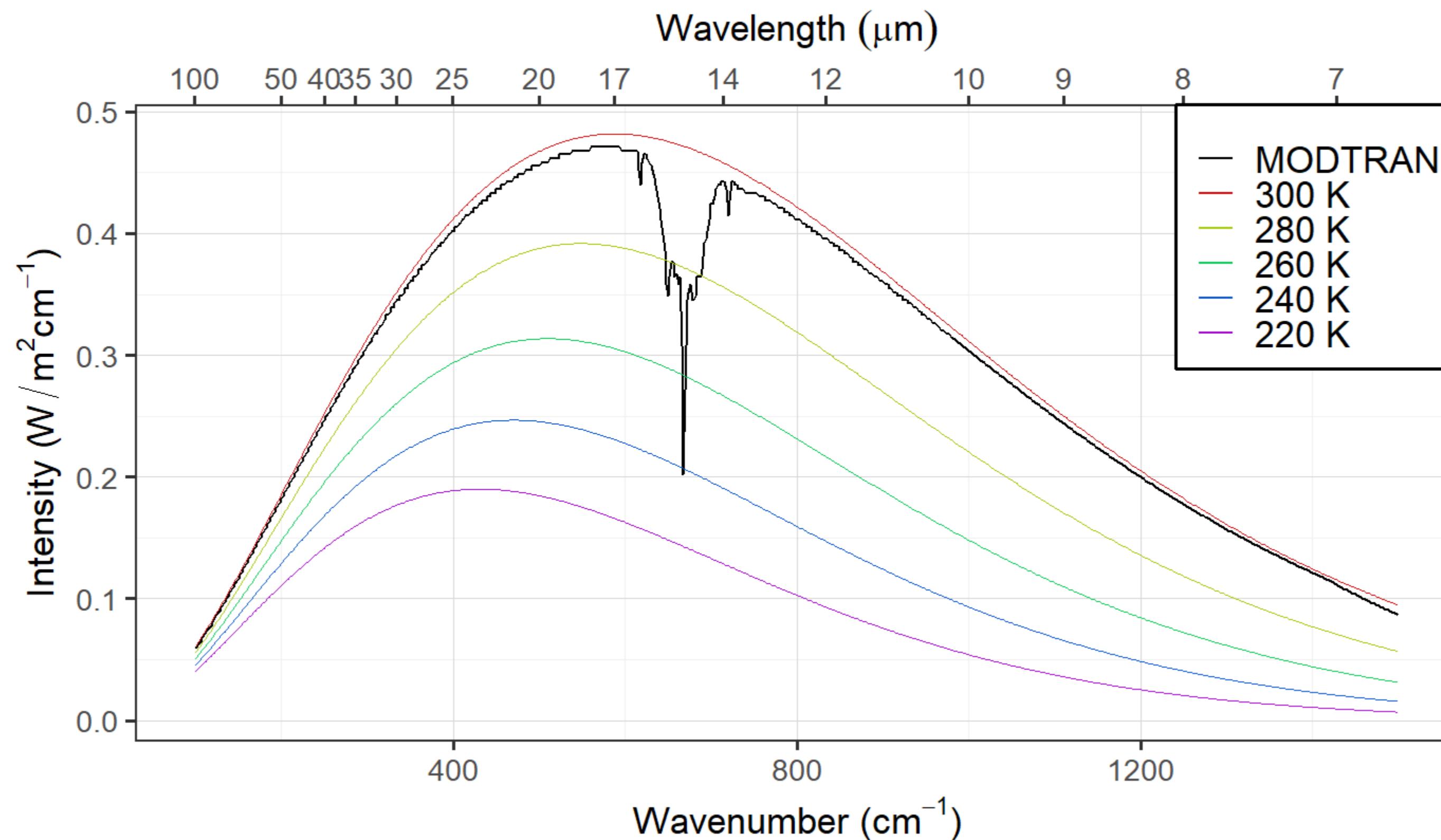
- Set altitude to **70 km** and location to “1976 U.S. Standard Atmosphere”.
- Set CO<sub>2</sub> to 1 ppm, all other gases to zero.
- Now increase by factors of 10 (10, 100, 1000, ...)

# 0.1 ppm CO<sub>2</sub>



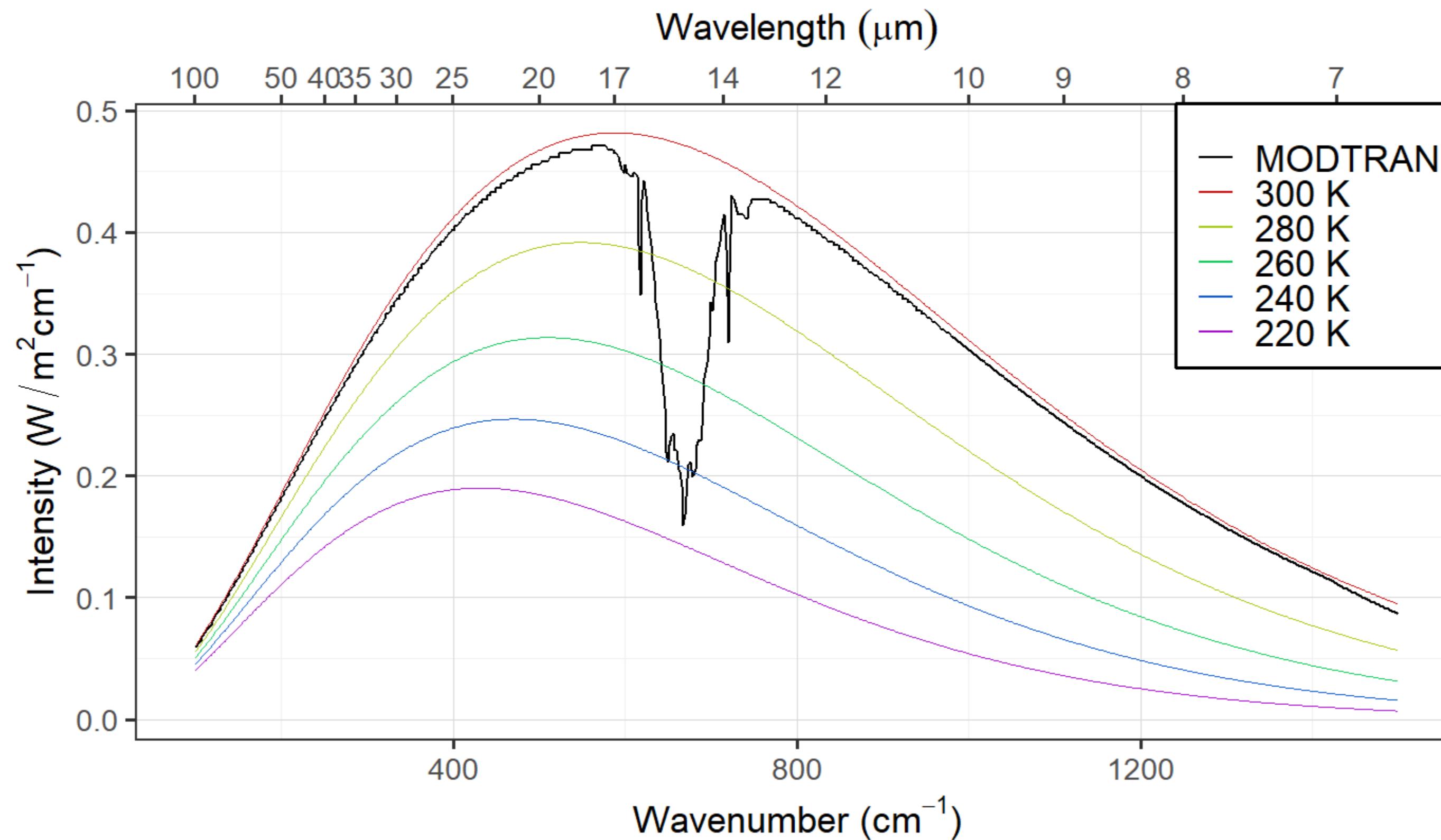
# 1 ppm CO<sub>2</sub>

MODTRAN: 1 ppm CO<sub>2</sub>, 70 km



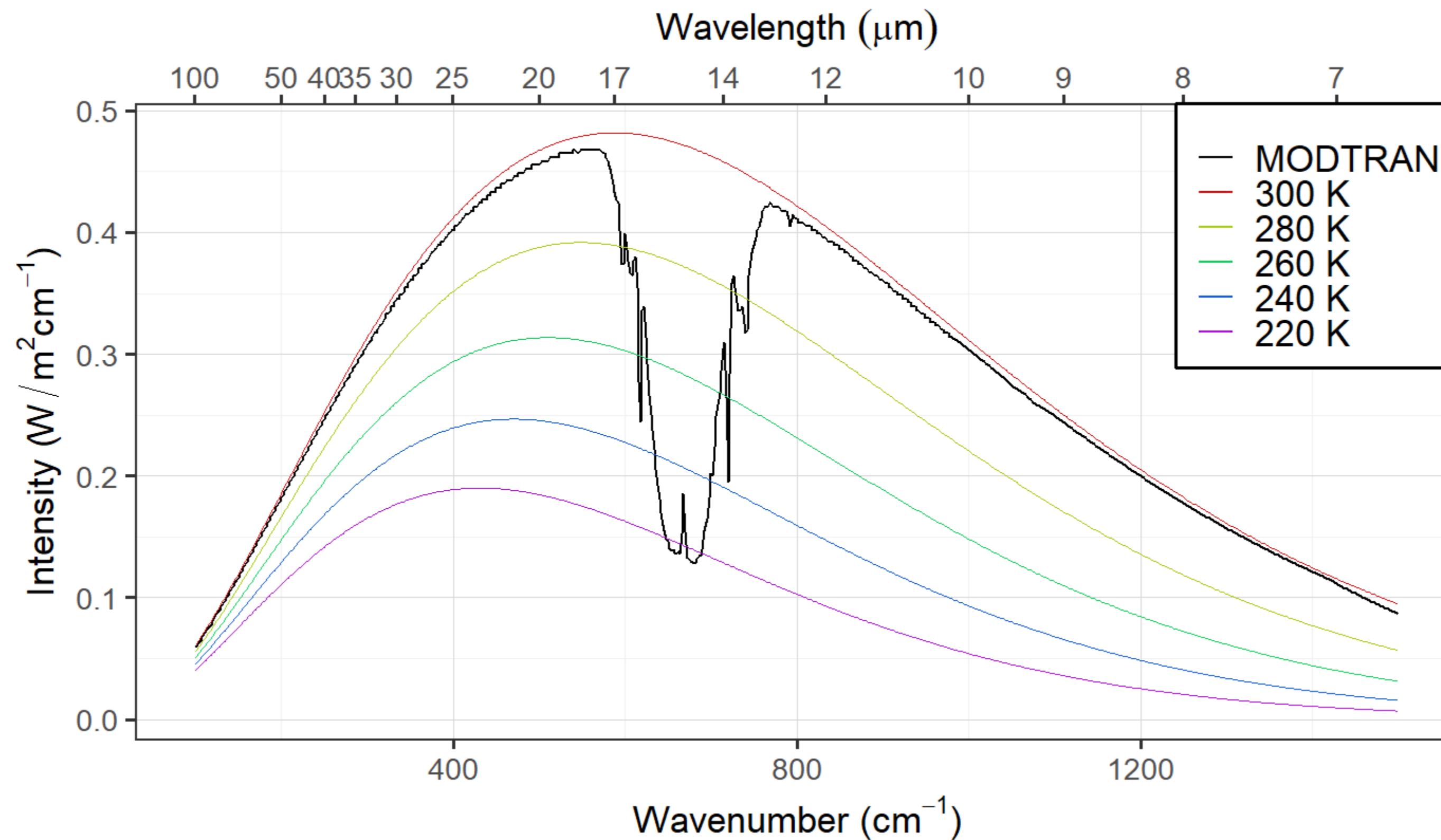
# 10 ppm CO<sub>2</sub>

MODTRAN: 10 ppm CO<sub>2</sub>, 70 km



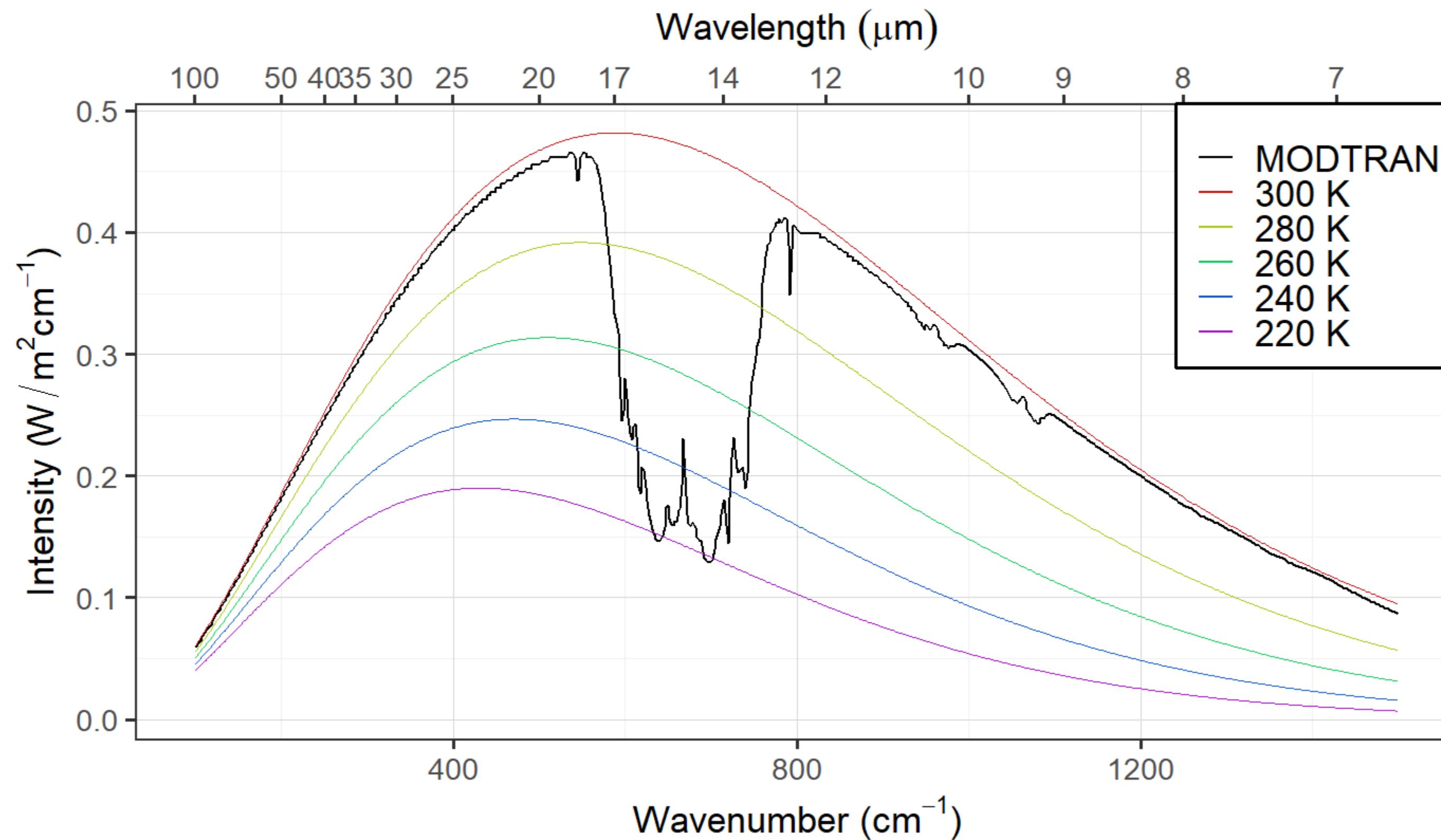
# 100 ppm CO<sub>2</sub>

MODTRAN: 100 ppm CO<sub>2</sub>, 70 km

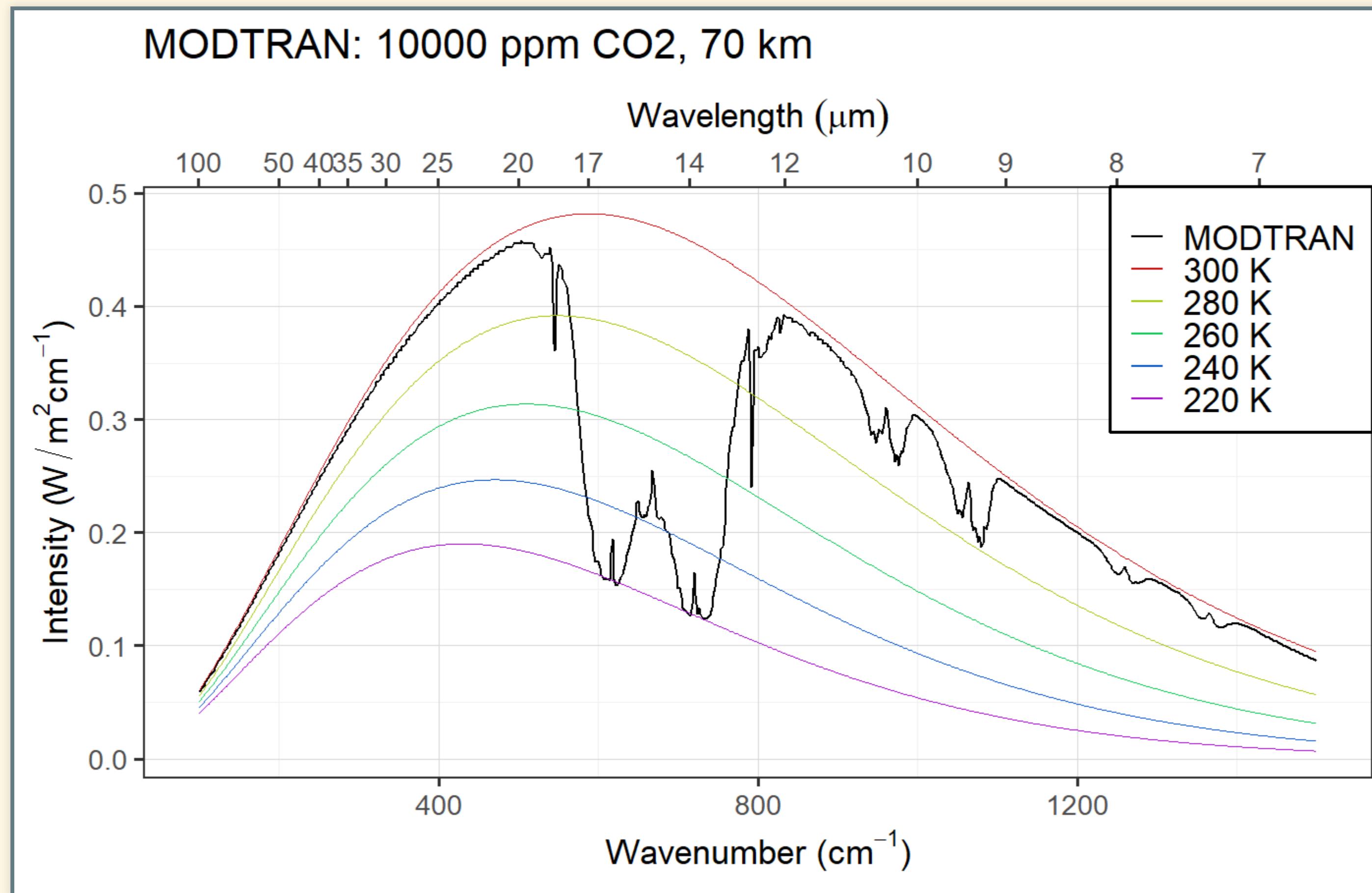


# 1000 ppm CO<sub>2</sub>

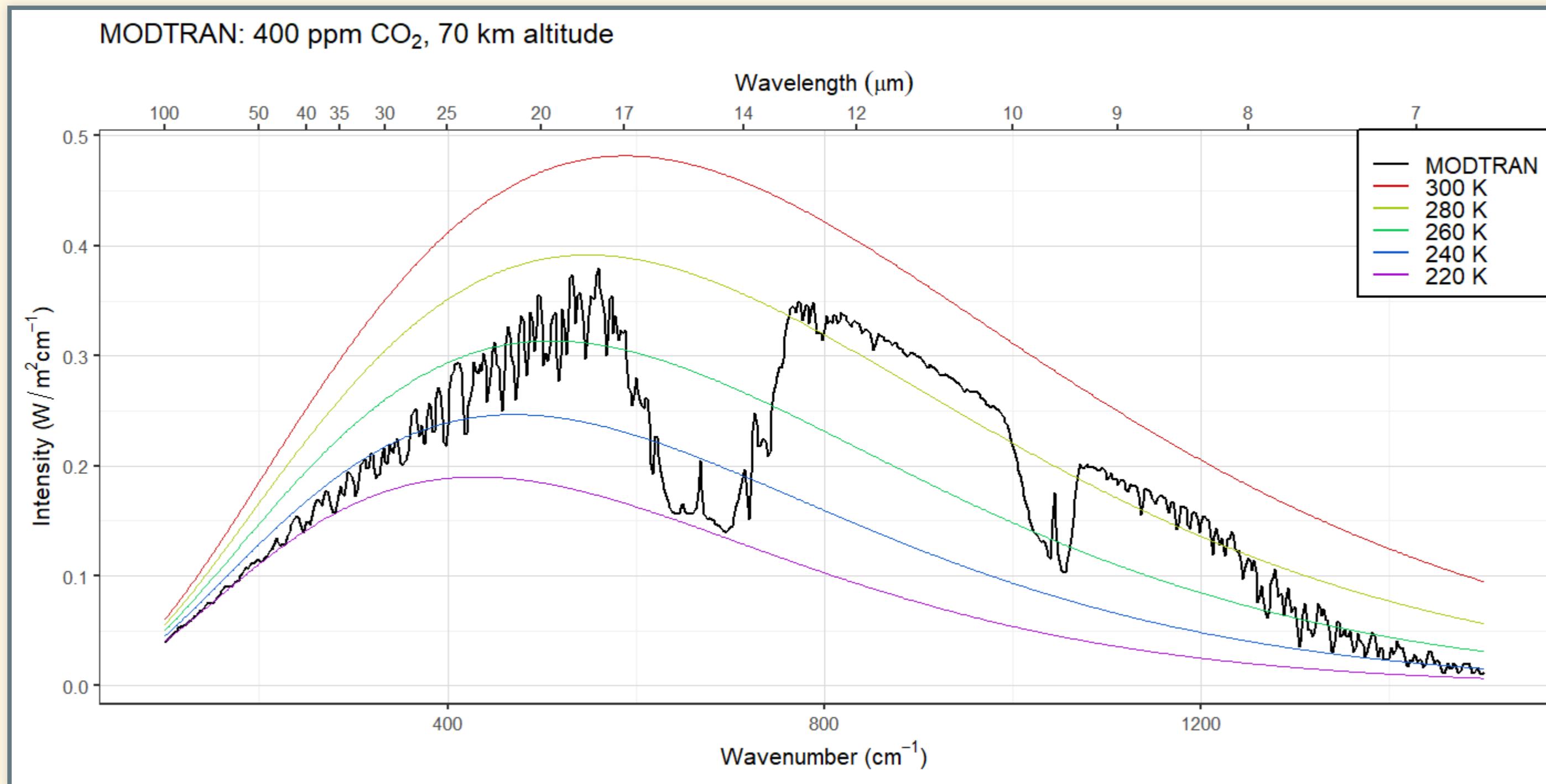
MODTRAN: 1000 ppm CO<sub>2</sub>, 70 km



# 10,000 ppm CO<sub>2</sub>

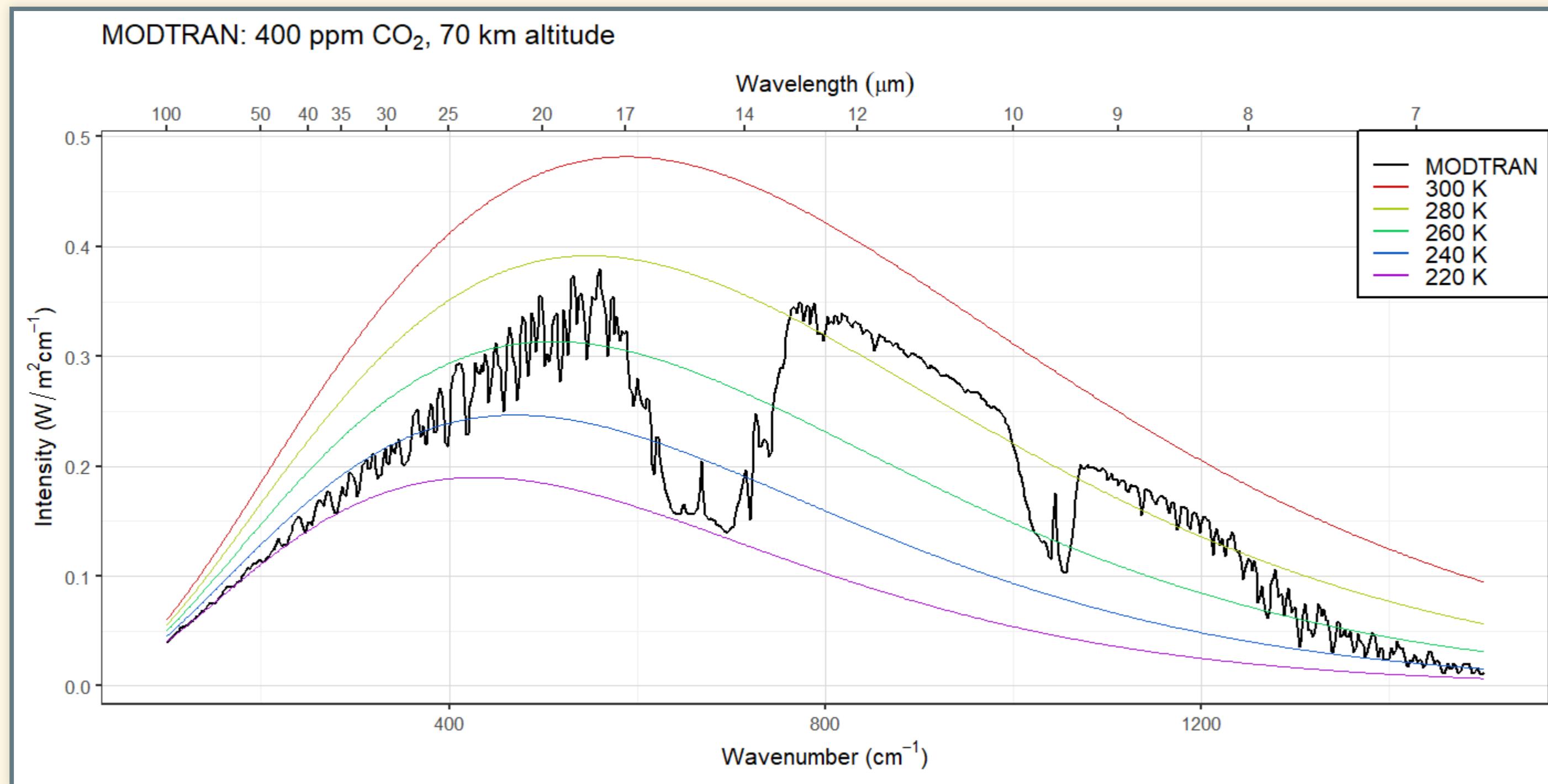


# Question



- Why do we see the spike in the middle of the CO<sub>2</sub> absorption feature?

# Question



- Water vapor absorption is completely saturated.
  - Why does water vapor emit at warmer temperatures than CO<sub>2</sub>?