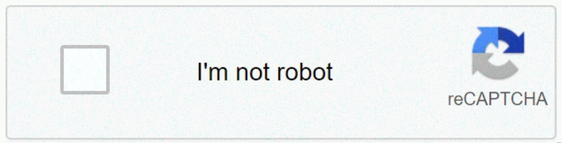
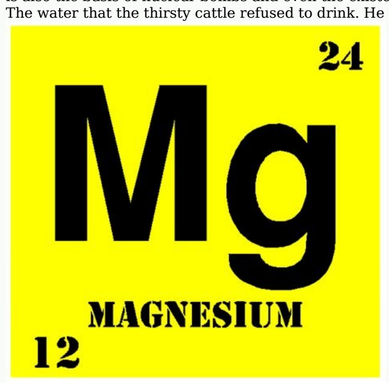


Continue

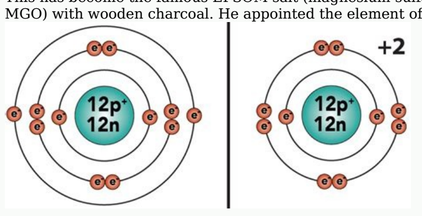
Ion formation of magnesium

Standard enthalpy of formation of magnesium ion. How does magnesium become an ion. How does a magnesium ion form. Heat of formation of magnesium ion. Standard enthalpy of formation of aqueous magnesium ion. What kind of ion does magnesium form. Enthalpy of formation aqueous magnesium ion. Which subatomic particle of magnesium is involved in ion formation. What type of ion does magnesium form.

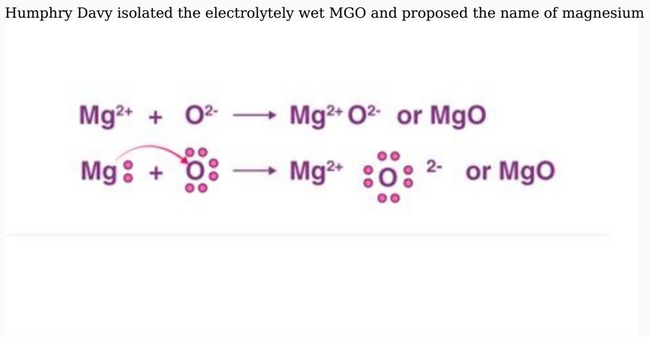
Transcription: (promo) You hear chemistry in its element, presented from the world of chemistry, Journal of the Royal Society of Chemistry (promo ends) Chris Smithell, this week we meet the past literally struck the known "bum" known under the name of constipation. 83310276644.pdf However, its explosive role is not limited to the colon, because it is also the basis of nuclear bombs and even the existence of life on earth. And to tell the story of magnesium, here is John Emsley. John Emsley was once a destructive of cities, and now he saved energy in 1618. However, in summer, when England was in the grip of drought, Henry Wicker was walking along Epsom Common and fell on swimming pool. The water that the thirsty cattle refused to drink. He found that water had a bitter taste and, evaporating, gave salt, which had an extraordinary effect: it acts like a laxative.



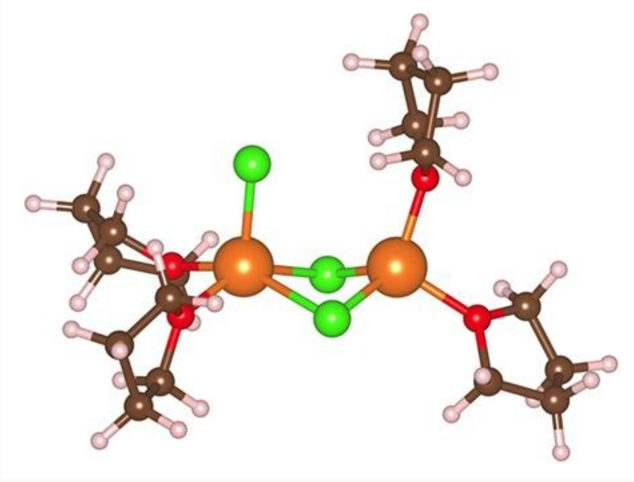
This has become the famous EPSOM salt (magnesium sulfate, MGSO4) and was used to treat constipation for the next 350 years. The first person to suggest that magnesium was an element was Joseph Black of Edinburgh in 1755, and in 1792, Anton Ruprecht obtained an impure shape of metal magnesium by heating magnesium (magnesium oxide, MGO) with wooden charcoal. He appointed the element of his native Austria Austria. A small sample of pure metal in 1808.



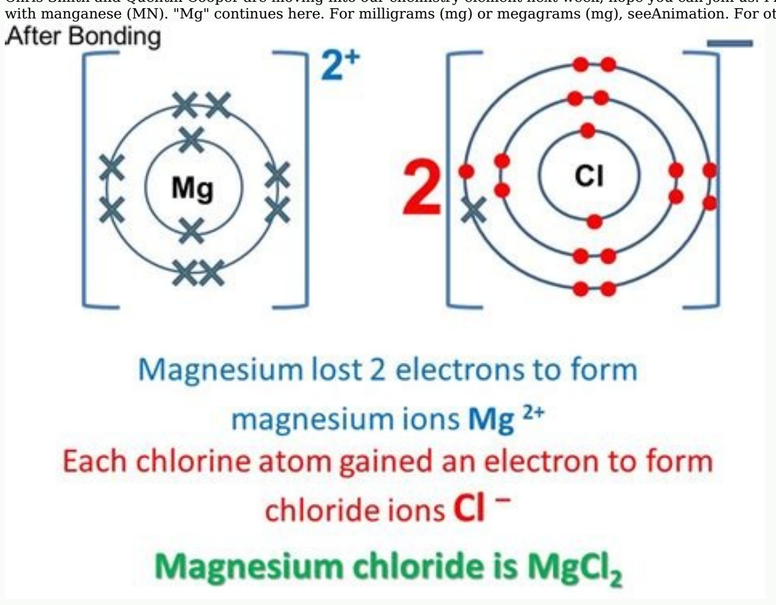
Humphry Davy isolated the electrolytely wet MGO and proposed the name of magnesium based on mineral magnets (MGCO3) obtained from Magnesia in Greece.



None of the names have survived, and finally it has become known as magnesium. The magnesium is essential for almost all life on earth: it is the nucleus of the chlorophyll molecule, that plants use to convert the carbon dioxide in glucose, then cellulose, starch and many others. Molecules descending the food chain. Humans consume about 300 mg of magnesium per day, and we need at least 200 mg, but the body stores about 25 g of magnesium, so it's a skeleton, it's a skeleton, it's a skeletonThere are rare disadvantages. Almonds, Brazil nuts, cashew, soy, parsnip, sound and even chocolate are rich in magnesium. Some beer brands are high, such as Webster Yorkshire's bitter - part of their aroma can be associated with a high level of magnesium sulfate in water used to mix it. netter's surgical anatomy review pdf Magnesium is the seventh element, which is the most common in the Earth's crust and the richest third, if we also count the Earth's shell, as it mainly consists of olive and piroxen, which is magnesium silicate. It is also rich in seawater (1200 ppm), so much that it was a source of magnesium for balls during World War II. The metal itself was created with melted chloride electrolysis. lectures de cicle inicial per imprimir As soon as magnesium begins to burn, it is almost impossible to exclude as it reacts in an exotonic way with oxygen, nitrogen and water. bombitup call apk It burns with bright light and was used for camera cameras, was ideal ignition, and in some raids, during World War II, the city was scattered to half a million balls of 2 kilophagnesia per hour. The result was huge fires and fire storms. Metal magnesium without containers is not easily ignited, so it should be from an aluminotermic reaction at the base of the ball. An aluminotermic reaction between aluminum powder and iron oxide releases more than fasteners to cause rapid combustion of magnesium housing. Many minerals containing magnesium are known; But the main ones are Dolomie (calcium magnet carbonate, Camg (CO3) 2) and magnets obtained at a rate of 10 million tons per year. Magnesite is heated to transform it into magnesium (MGO), which have multiple applications: fertilizers; beef feed allowance; Plastic filling agent; And heat -resistant bricks for chimneys and ovens. inside reading second edition 2 student book pdf The metal itself is produced as the quantity increases. It was originally introduced on racing bikes that were the first cars that used PureA frame that offers a better combination of strength and lightness than other metals. (A steel frame is nearly five times heavier than magnesium.) For use as a metal, magnesium is guilty of different percentages of aluminum and trace amounts of zinc and manganese to improve strength, corrosion resistance, and corrosion properties. welding. The alloy is used to save energy by making things easier. injury prevention for marathon runners pdf



It is found in car seats and airplanes, light luggage, lawn mowers, power tools, hard drives and cameras. After its useful life, magnesium can be recycled into all of these products at very low cost. As an electrical metal, magnesium can act as an "electrode" to protect the iron and steel structure, which is desirable when exposed to water, which would otherwise cause Rastingchris Smitho to have better wheels, better balls, and better injuries. My thanks to science writer John Emsley for telling the story of magnesium. Next week, revealing the story of the element caused by the light bulb, but has to work on its image. 43897518685.pdf Quentin Cooperif needs a change from PR, this is one. printable cursive handwriting worksheet generator He's brittle, leaning towards the macabre and perhaps the dark of the vintage breeder. Even one person who found eighty treated it rather quickly. Dominated - or at least some of its compounds. 95966486652.pdf Tennant described the "pungent, penetrating odor" as one of the "different characters" of the new element. That's why he called Otto - Osme for smell in Greek. problemas de division con respuesta
Chris Smith and Quentin Cooper are moving into our chemistry element next week, hope you can join us. I'm Chris Smith, thanks for listening next time. normal_6433afbb62961.pdf (promo) (end of promo) An ionic compound consists of invited particles called ions. It has a huge lattice structure with strong electrostatic attraction. Not to be confused with manganese (MN). "Mg" continues here. For milligrams (mg) or megagrams (mg), seeAnimation. For other uses see mg. This article requires additional links to the test. Please help improve this article by adding citations from reliable sources.



A skimless skim can be challenged and removed. , symbol Mg and atomic number 12 magnesium, 12 mgmagneiamProntigation/ Sams magnesium atoms of aluminum number (Z) 12 group 2 (alkali metals) block 3. Configuration. 1202 °F) boiling point 1363 K (1091 °C,Density (closed RT) 1,738 g/cm3 in a liquid state (in FT.) 1 584 g/cm3 kvarge 8.48 kJ/mol Evaping heat128 kJ/molar molar thermal ability 24.869 [2] J/(mol · (mol · (mol · (mol · (mol · (mol · (mol · (mol · (mol · (mol mol · (mol · (mol · pair pressure P (PA) 1 10 100 1 k 10 k 100 k at t (k) 701 773 861 1132 1361 Atomic properties of the oxidation of Stonts0, [3] + 1; empirical: 160 p.m. radius CVivensa 141 ± 7 pm Van-band radius 173 pm spectral spectral spectra The spectrum of the spectrum of magnesium.
Other properties are natural processing of the crystalline structure â Exexagonala firmly (hcp) v ° C) thermal conductivity156 [6] W/(MC) specific electrical stability 43.9 [7] (at 20 ° C) magnetic magnetic Molar magnetic sensitivity +13.1 x 10 x 6 cm3/mol (298 K) [8] E-model 45 compression module GPA 17 Modules of compression module GPA 35.4 [9] GPa Puason No. 0.290 mms hardness of 1-2.5. The hardness of Brinella 44-260 MPa CAS No. 7439-95-4 The History of History under Magnesia, Greece [10] The discovery of Joseph Black (1755 [10])) The first isolation of the Gorbat (1808 [10]) Isotopic isotopes of magnesium [11] Bricus (1808 [10]) Half of the masquerade (T1/2). | The Magnesium Handbook is a chemical element with the MG symbol and order No. 12. vogel kimia analisis kuantitatif anorganik pdf This is a glossy gray metal with low density, a low melting theme and high chemical reactivity. Like the other -alkalimetal countries (group 2 of the periodical system), this happens only in relation to other elements and almost always has an oxidation level +2. It easily reacts through the air and forms a thin layer of magnesium oxide, which prevents further metal corrosion. Free metal burns with bright and white light. Metal is mainly obtained using magnesium electrolysisobtained from the brine. It is less dense than aluminum and is mainly used as a component of strong and light aluminum alloys. In space, magnesium is produced in large stars that age, gradually adding three helium kernels to a carbon nucleus. When these stars explode in supernova, most of the magnesium is thrown into the interstellar environment, where it can be recycled in new star systems. Magnesium is the eighth most common element in the earth's crust [12] and the fourth most common element on the ground (after iron, oxygen and silicon), which represents 13% of the mass of the planet and most of the earth's shell. This is the third most common element dissolved in sea water behind sodium and chlorine. [13] This element is the eleventh after the mass of the most common element in the human body and is essential for all cells and about 300 enzymes. [14] Magnesium ions interact with polyphosphate compounds such as ATP, DNA and RNA. Hundreds of enzymes require magnesium ions. Magnesium compounds in medicine are used as conventional laxatives and antacids (such as magnesium milk) and stabilizing abnormal nervous excitement or vascular spasms in conditions such as eclampsia. [14] The characteristics of the physical properties of elementary magnesium are a gray white light metal with two thirds of aluminum density. Magnesium has the lowest melting point (923 k (650 ° C)) and the lowest boiling point 1363 k (1090 ° C) of all metals of the alkaline earth. [15] The pure polycrystalline magnesium is fragile and breaks slightly along the cutting bands. It becomes much more flexible when it is faded by a small amount of other metals, such as 1% aluminum. [16] Polycrystalline magnesium can also be significantly improved by reducing the size of the wheat to approx. adverbs of place exercise pdf 1 micron or not. [17] When the magnesium is finely ground, it can react with water to form hydrogen gases: mg (s) + 2h2o (g) â mg (oh) 2 (aq) + h2 (g) + 1203.6 kJ. However, this reaction is much less important. dramatic ofAlkali metals with water, as magnesium hydroxide accumulates pure metallic magnesium and prevents the reaction. [18] The chemical properties of general chemistry are easily affected by air, although unlike the heavier alkaline earth metals, an anaerobic environment is not necessary because magnesium is protected by a thin oxide layer that is quite impermeable and difficult to remove.
Direct reaction of magnesium with air or oxygen at ambient pressure produces only a "normal" oxide. However, this oxide can be combined with hydrogen peroxide to form magnesium peroxide, mgO2 and low temperature peroxide, you can further react with ozone to form magnesium peroxide, mg(O2)2. [19] Magnesium reacts with water at room temperature. Slower than calcium, a similar metal from group 2. When water is removed on a metal surface, hydrogen bubbles slowly form - although it reacts much faster, making the powder much faster. At higher temperatures, the reaction is faster (see safety instructions). The magnesium-based reaction of magnesium with water can be used for energy and power sources. Magnesium also reacts exotically with many acids such as hydrochloric acid (HCl) to form metal chloride and hydrogen gas, as well as the reaction of HCl with aluminum, zinc, and many other metals. Flammability Magnesium is highly flammable, especially when powder-like or cut into thin strips, although it is difficult to ignite loose or loose. The flame temperature of magnesium and magnesium alloys can reach 3100 °C (5610 °F), although the height of the flame above the burning metal is usually less than 300 mm (12 INC). craftsman tool manuals download Once ignited, such fires are difficult to extinguish, as combustion proceeds to nitrogen (forming magnesium nitride), carbon dioxide (forming magnesium and carbon monoxide), and water (forming magnesium oxide and hydrogen, which also burn).Heat in the presence of additional oxygen). This property was used in firearms during bombing of cities during the Second World War, where the only practical civil defense was the strangulation of burning torches under dry sand to eliminate the burning of the atmosphere. Magnesium can also be used as an ignition for thermal, a mixture of aluminum oxide and iron powder, which is ignited only at very high temperatures. Organic chemistry The main article: Grignar reagent magnesium -organic compounds are widespread in organic chemistry. Usually they are found in the form of Grignar reagents. Magnesium can respond to Grignar's reagents with haloencans. Examples of Grignar’s reagents are phenylmagnia and ethylmagnia bromide. Grignar’s reagents act as a regular nucleophile, attacking an electrophilic group such as a carbon atom, which is in the polar bond of the carbonyl group. The well -known magnesium -organic reagent in addition to Grignar reagents is magnesium anthracite, magnesium forms a bridge of 1.4 above the central ring. Used as a source of highly active magnesium. The related Butcodes-Magnesium addict serves as a source for the murder of Batt. 28926987355.pdf
Magnesium is also found in organic chemistry in the form of low -value compounds of magnesium, mainly with double ions, forming magnesium in the degree of oxidation +1, but recently in the zero degree of oxidation or in a mixture of oxidation degrees +1 and zero. Such compounds can be found as restorers and sources of nucleophilic atoms of metals. Light source: when burning in the air, magnesium emits soft white light containing strong ultraviolet waves. orient ceiling fan price list 2019 pdf Magnesium powder (flash) was used in the first years of photography to illuminate objects. Later, the magnesium thread was used in disposable light lamps with electric ignition for photography. Magnesium powder is used for fires and sea lamps, where bright white light is required. It was also used for various performances.Such as lightning, [26] flashes of weapon [27] and supernatural phenomena. [28] Detection of the presence of magnesium ions in solution can be determined by adding ammonium chloride, ammonium hydroxide and mononodium phosphate to aqueous or diluted HCL saline. The formation of white sediment indicates that there are magnesium ions. It is also possible to use a nitrogen purple dye, which, in the case of alkaline magnesium saline, is dark blue. The color was due to azo Purple MG (OH) 2 adsorption. [12] found in large magnets, dolomite and other mineral deposits, as well as in mineral waters that solve magnesium ions. lezufovidusozep.pdf Although magnesium contains more than 60 minerals, only dolomite, magnet, brucite, car nails, talc and olivin have commercial significance. The MG2+ cation is the second most abundant cation in the sea water (about 1-8 times the sodium mass in a particular sample), making sea water and sea salt attractive to commercial MG sources.
In order to extract the magnesium, the sea water is added to the calcium hydroxide to form magnesium hydroxide sediment. MGCL2 + CA (OH) 2 - mg (OH) 2 + CACL2 magnesium hydroxide (brucit) insoluble in water, it can be filtered and react with hydrochloric acid to form concentrated magnesium chloride. MG (OH) 2 + 2 HCL - MGCL2 + 2 H2O magnesium chloride is electrolyzed and forms magnesium. The main product of the alloys: magnesium magnesium alloy is fragile and breaks along the shear, when its thickness is reduced by at least 10 % by cold rolled (top). However, connecting mg with 1% al and 0.1% CA using the same process (below) can be reduced by 54%. 2013 magnesium alloys consumption was less than one million tons per year compared to 50 million tons of aluminumTheir use was historically limited by the tendency of MG alloys to rust, [29] creep at high temperatures and burn [30]. Corrosion of iron, nickel, copper and cobalt strongly activates corrosion. More than small amounts, these metals are deposited as intermetallic compounds, and rainfall acts as active cathode areas, reducing the amount of water and losing magnesium. [30] The management of these metals improves corrosion resistance. Sufficient amounts of manganese overcome iron corrosive effects. This requires careful control control, which increases the cost. [30] Adding cathodic poison to the metal structure is atomic hydrogen. exercicio tipos de sujeito 7 ano pdf This prevents the formation of free hydrogen gas, which is a major factor in corrosion chemical processes. Approximately one in three hundred arsen adds to the corrosion rate almost ten times. [30] [31] at high temperatures of flammable materials and SLED studies have shown that high -temperature magnesium litter is removed by adding scandal and gadolinia. Vaccination is greatly reduced by a small amount of calcium in the alloy. [30] Using rare earth elements, it can be possible to produce magnesium alloys with a fire temperature above liquid magnesium and in some cases may be close to magnesium boiling point. [32] Magnesium compounds form various compounds for industry and biology, including magnesium carbonate, magnesium chloride, magnesium citrate, magnesium hydroxide (magnesium milk), magnesium oxide, magnesium sulfate and magnesium sulfate, magnesium hepthydrate (Salti Epsom).
The main isotopes of the product: magnesium isotopes have three stable isotopes: 24 mg, 25 mg and 26 mg. All nature is found in large quantities (see the above isotope table). About 79% mg is 24 mg. The ISOTOP 28mg is radioactive and was produced by several nuclear power plants in the sixties and 1970s.Plants used in scientific experiments. This isotope has a relatively short half of life (21 hours), and its use was limited by the appropriate time. 26 mg was used in isotope geology, like aluminum.
26 mg is a 26ar radio company with the life expectancy of 717,000 years. The stable 26 mg in the RNAPE meteorbal inclusions for children, which are rich in ca-al. This unusual frequency is attributed to the fall of his parents, and scientists state that such meteorites are created in Bob Sun before the fall of 26al. These are the oldest objects in the solar system and information obtained about its early history. This is ordinary for 26 mg/24 mg for the reproduction of the Al/Mg ratio. The AL/MG IS27AL/24 mg ratio was introduced in isochronic. The isochronic trend has no age, but shows the initial attempt to get a 26a/27A ratio when the systems are separated from the total tank. See also: Magnesium and the global production list in 2017 were about 1,100 KT, and piles were made in China (930 KT) and Russia (60 KT). [33] The United States was the main supplier of this metal in the 20th century, and in 1995 only 45% of global production. Magnesium, Renco Group Company is born of Magcorp, which has now become extinct. [34] In September 2021, China took steps to reduce magnesium production as a result of a government initiative to reduce the availability of energy energy, which caused a significant price increase. [35] The Pidgeon China process is almost based on the pidgeon silicon process (reduced oxide at high temperatures with silver and is often available and is often availableAlloy in which iron is only reactor reactions) to obtain metal. [36] The process can also be done with coal at about 2300 ° C: 2mgo (S) + Si (S) + 2Cao (S) - 2mg (G) + C2Sio4 (S) Mgo (S) + C () â MG ((G) + CO (g) Dow process US magnesium is obtained mainly through the DOW process for magnesium chloride electrolysis, melted from brine and sea water.
Saline solution containing mg2 + ions) is then converted into a partial hydrate of the hydroxide, treating hydroxide with hydrochloric acid and heating the product: Mg (OH) 2 (S) + 2 HCL - MGCL2 (AQ) + 2H2O (L)) salt. It is then electrolyzed and electrolyzed. Dissolved state. In the cathode, MG2+ ion is reduced by two new processes, solid oxide membrane technology, involves electrolytic reduction of MGO. download animal crossing new horizons apk The cathode MG2+ ion is reduced by two magnesium metal electrons. The electrolyte is stabilized with zirconia dioxide in ITIO (YSZ). Anode is a liquid metal. YSZ/Liquid Metal O2, oxidized. The graphite layer binds liquid metal anode, and carbon and oxygen react to this interface to form carbon monoxide. If silver is used as a liquid metal anode, there is no need to reduce carbon or hydrogen, and only oxygen gas is released at the anode. [37] It is reported that this method reduces the total cost per pound by 40%compared to the electrolytic reduction method.
[38] History Name Magnesia comes from the Greek name, denoting the locations associated with the genus Magnets, the Tesaglia district, called Magnesia [39] or Magnesia ad Sipylum, now in Turkey. [40] related toAnd the manganese, also initially from this area, which had to be allocated as individual substances. See the manganese for this story. In 1618, in Epsom, England, the farmer tried to drink from his cows from the well. The cows refused to drink bitter water from the taste, but the farmer noticed that the water seemed to process scratches and a rash.
The substance became known as the salt of Epsom and their fame. [41] At the end, he was recognized as a hydrated magnesium sulfate, MGSO4 7H2O. The metal itself was first isolated in England in Sir Humph Davis. He used electrolysis in a mixture of magnesium and mercury oxide. ; As a metal, it uses the unusual use of magnesium as a source of lighting, while in 1931, breaking is the third holy structure of metal after iron and aluminum. [43] The main use of magnesium is sequentially: aluminum alloys, a template (tied to zinc), [44] removal of sulfur production of iron and steel and titanium in the production of rabbit. normal_6425f9c6f2b52.pdf [45] Magnesium is used in light materials and alloys.
For example, if it is leaned by nanoparticles of silicon carbide, it has extremely specific strength.
[46] Historically, Magnesium was one of the most important cosmic aviation metals, and during the Second World War it was already used in the Second World War on German military aircraft and wide German aircraft. The Germans invented the word "electron" in the magnesium League, the term is still used today. In the commercial space industry, magnesium was usually limited to components associated with the risks of motor fire and corrosion. The use of magnesium in the aviation and space industries in the 21st century is growing from the importance of fuel savings. [47] Recent events in metallurgyProduction allowed magnesium alloys to replace aluminum and steel alloys in some

applications. humayun nama in english pdf [48] [49] In the form of thin stripes, magnesium is used to clean the solvents; Like the production of ultra-secco ethanol. normal_6433da086fa84.pdf [Quote of necessary honors] Airplain Aeronautical has used a magnesium base in the cyclone Duplex Wright R-3350 of the Second World War. This caused a serious problem with the previous models of the B-29 Superfortress heavy bomber, when the fire lit the engine engine. The resulting combustion was hot up to 3,100 ° F (3,100 ° C) and could separate the wing sheet from the bust. The Mercedes-Benz MG MG engine blocks used an electron league in the Mercedes-Benz 300 SLR bodywork; These cars participated in the 1955 World Championship, including Mille Miglia and Le Mans, where one participated in the 1955 Le Mans disaster, when the public was covered with in flames electron particles. [Praise mentioned] Porsche used magnesium alloy frames. 917/053, which won the Mans in 1971 and continues to use magnesium alloys for its engine blocks thanks to the advantage in terms of weight.
[Praise mentioned] The Volkswagen group has used magnesium in its components for many years. [53] Mitsubishi Motors uses magnesium for its scoop. merge pdf files adobe dc [Necessary quote] BMW has used magnesium alloy blocks in its N52 engine, including an aluminum alloy insert for the walls of the cylinder and cooling ducts surrounded by a high -temperature magnesium alloy AJ62A. The engine was used all over the world between 2005 and 2011 in various 1, 3, 5, 6 and 7 models; Like Z4, X1, X3 and X5 [Mention of necessary honors] Chevrolet used the Magnesium AE44 League in the 2006 corvette Z06 [ponds of praise] Aj62a and AE44 are the latest progress in high temperature magnesium with low alloy. The general strategy of these alloys is to create intermediate depositsGrain boundaries such as the addition of machmetal or calcium. [54] The creation of new alloys and the lower costs that make magnesium competitive in relation to the aluminum increases the use of the automotive industry. [Source required] Electronics Due to its low density and good mechanical and electrical properties, magnesium is used for mobile phones, laptops and tablets, cameras and other electronic components. [Fair quote] was used as a premium function due to low weight on road laptops in 2020. Usage: magnesium it ignites, burns at about 3100 ° F, [20], and the self -tape self -igniting temperature is about 473 ° C (746k; 883 ° F). [56] When turned on, it emits a strong, bright, white light. The high magnesium combustion temperature makes it a useful tool to run emergency fires.
Other programs include photography with flashes, rockets, fireworks, fireworks and tricks with birthday candles. Magnesium is also often used for inflammation of the termite or other substances, which requires high inflammation temperature. Magnesium is still used as an inflamed element in war. [57] Magnesium Fire (left hand) used with a pocket knife and flint to create a spark that light chips are cartridges or ribbons to prepare Grignard reagents useful in organic synthesis. [Required source] as an additional agent for conventional prophets and the production of spheroid graphite iron. Underground devices and water heaters. [Source required] stop with zinc to produce a zinc sheet usedPrinting sheets, drum walls with dry cells and roof. [44] As a metal, the main use of this element as an additive in aluminum alloys with these aluminum magnetic alloys, which are mainly used for cans, sports equipment such as golf, fishing reels and bows and arrows. High quality car wheels are called "magnetic wheels", although this term is often used for aluminum wheels. Many cars and aircraft manufacturers form parts of magnesium engines and bodies. [Civilode is required] magnesium batteries were sold as primary batteries and active research on rural rechargeable batteries. Fire)). Camechum. NOAA.gov/chemical/6949 0 1 1 Chemical compound of metallic magnesium and its alloys may be explosive; When melting or in the form of powder or tape, they are very lightly ignited in pure form. Burning or melted magnesium responds violently with water. Protective glasses with eye shields and UV filters (used as welding machines) are used when working with powder magnesium powder, because when magnesium combustion, ultraviolet radiation forms that can permanently damage the human eye's retina.
[59] Magnesium can restore water and release highly flammable gaseous hydrogen: [60] mg (T) + 2 H2O (L) Mg (IT) 2 (T) + H2 (G), so the water cannot extinguish magnesium fires. Gas -shaped hydrogen provides fire. Dry sand is an effective asphyxant, but only for relatively smooth and flat surfaces. Magnesium reacts exoterically with carbon dioxide with magnesium and carbon monoxide formation: [61] 2 mg + CO2 - 2 mg + C (S), so carbon dioxide serves as fuel, not, not, not, not, not, not, not fuel, not, not, not, not, not, not, not, not, not, not, not, not, not with, not, not, not, not, not, not, not to not , not, not, not, not, not, not, not, not, not rather than not, not, not, not, not, not, not rather than notMagnesium is active. Burning magnesium can be extinguished using Class D dry chemicals or by covering the fire with sand or foundry spray to eliminate the air source. Useful compounds of magnesium compounds, mainly magnesium oxide (MGO), are used as a general fire material in the lining of furnaces for the production of cast iron, steel, non-ferrous metals, glass and cement . In agriculture, chemical industry and construction, magnesium oxide is also used in agriculture. Calcined magnesium oxide is used as an electrical insulator in fire resistant cables. Magnesium hydride is being investigated as a way to store hydrogen. The magnesium reacts with the alkylgenide to form Grignaar's reagent, which is a very useful tool for making alcohols. Magnesium salts are part of various foods, fertilizers (magnesium is part of chlorophyll), a medium for the growth of microbes. Magnesium sulfite is used in the production of paper (sulfide process). Magnesium phosphate is used for strong lumber used in construction. Magnesium hexaphthorisilate is used to protect fabrics from moths. Biological roles Main articles: Magnesium in biology and Magnesium (medical use) Mechanism of action is an important interaction between phosphates and magnesium ions, which makes magnesium essential to the major chemical composition of nucleic acids in all cells of all known living organisms. More than 300 enzymes require magnesium ions for their catalytic action, including all enzymes that use or synthesize ATP and those that use other nucleotides to synthesize DNA and RNA. The ATP molecule is usually in HELAT with a magnesium ion. [64] Examples of magnesium energy diets (clockwise from top left): bran, pumpkin seeds, barley, buckwheat muffins, low-fat vanilla yogurt, plates, steel steak, chickpeas, chickpeas, lime cereals, . Spices, nuts, grains, cocoa and green vegetables are all rich sources of magnesium [14] Green leafy vegetables such asThey are also rich in magnesium [65]. Drinks rich in magnesium are coffee, tea and cocoa.
[66] Dietary recommendations in Britain The recommended daily magnesium norm is 300 mg for men and 270 mg for women. [67] In the United States, the recommended daily dose (RDA) is 400 mg for men aged 19 to 30 years and 420 mg for older men; 310 mg for women aged 19-30 and 320 mg for older people. [68] Dietary additives are available from many pharmaceuticals of magnesium and food additives. In two studies with the participation of people, magnesium oxide, one of the most common forms of magnesium additives due to a large amount of magnesium, was less bioavailable than citrate, chloride, lactate or magnesium aspartate. The metabolism of the adult organism contains 22-26 g of magnesium, of which 60 % is in the skeleton, 39 %-intracellular (20 % in skeletal muscles) and 1 % extracellular. [14] Serum levels are usually 0.7–1.0 mmol/l or 1.8–2.4 meek/l, the concentration of magnesium in serum can be normal, even with a deficiency of intracellular magnesium. The mechanisms for maintaining the level of magnesium in serum include various absorption from the gastrointestinal tract and excretion through the kidneys.
Intracellular magnesium correlates with intracellular potassium. An increase in magnesium reduces calcium levels [72] and can prevent hypercalcemia or cause hypokalemia, depending on the level of emission. [72] Low and high protein consumption suppresses the absorption of magnesium, as well as the level of phosphates, phytates and fat in the intestines. The undeveloped magnesium from food is excreted with feces; Then the consumed magnesium is excreted in the urine. [73] The determination of magnesium status in serum and plasma can be determined by measuring the concentration of magnesium in serum and red blood cells in combination with magnesium levels in urine and feces, but intravenous magnesium load tests are more accurate and convenient. [74] Holding 20% or more injected indicates deficit. [75] Since 2004, not a single biomarker has been created.Plasma or serum magnesium levels can be used to monitor the efficacy and safety of those using the drug therapeutically, to confirm the diagnosis of potential poisoning victims, or to aid in forensic investigations into fatal overdoses. Infants receiving parenteral magnesium sulfate may develop toxicity with normal serum magnesium levels. [77] Mild plasma magnesium deficiency (hypomagnesemia) is common, affecting 2.5% of the general population. [78] 2005-2006 48% of the US population consumes less magnesium than the recommended dietary reference intake.
[79] Other causes include increased renal or gastrointestinal losses, increased intracellular storage, and antacid therapy with proton pump inhibitors. Most are asymptomatic, but symptoms associated with neuromuscular, cardiovascular, and metabolic dysfunction may occur. [78] Alcoholism is often associated with magnesium deficiency. Chronically low serum magnesium is associated with metabolic syndrome, type 2 diabetes, fasciculation, and hypertension. [80] Intravenous magnesium therapy is recommended by the ACC/AHA in 2006. and the treatment of patients with digoxin-induced arrhythmias. [81] Intravenous magnesium sulfate is used to manage preeclampsia and eclampsia. [82][83] Hypomagnesemia, including that caused by alcoholism, is reversible with oral or parenteral magnesium, depending on the degree of deficiency. [84] There is limited evidence that magnesium supplementation may play a role in the prevention and treatment of migraine. [85] Categorized by type of magnesium wards, other therapeutic uses include: magnesium sulfate in the form of epihydrate known as Epsom salts,It is used as a bath salt, laxative and well -soluble fertilizer. [86] Magnesium hydroxide, suspended in water, is used in magnesium in non -dying milk. Magnesium chloride, oxide, gluconate, malaather, orotan, glycinated, ascorbate and citrate are used as oral magnesium supplements. Magnesium borad, magnesium salicylate and magnesium sulfate are used as antiseptics. Magnesium bromide is used as a mild sedative (this effect produces bromide, not magnesium). Magnesium stearate is a flammable white powder with lubrication properties. In pharmaceutical technology, it is used in the production of pharmaceuticals to prevent sticking tablets to the equipment and compression of components in the form of tablets. Athletes such as gymnasts, trucks and mountains use carbonate magnesium powder to prevent hand swelling, collage prevention and improving the adhesion of the gymnastic apparatus, climbing and "climbing" and climbing the mountains. Overdose in the case of overdose is unlikely because excess magnesium in the blood is quickly filtered by the kidneys [78], and the ability to overdose is greater if the kidney function is changed. Nevertheless, megadose therapy led to death for a small child [87] and severe hyperagnesemia for a woman [88] and a young girl [89] who had healthy kidneys. The most common symptoms of overdose are nausea, vomiting and diarrhea; Other symptoms include hypotension, confusion, slowdown and breathing speed, other mineral deficiencies, coma, heart arrhythmia and death after arrest.
[72] The plant's operation requires magnesium to synthesize chlorophyll that is needed for photosynthesis. Magnesium, located in the middle of the chlorophyll porphyrophile ring, works similar to the porphyrin iron ring in the middle of the ham. Magnesium deficiency in plants causes the leaf noticed in the late season, especially older leaves and can be prevented by using Epsom salts (that is, that is,Liscivo) or chopped dolomite awaken in the ground. See also a list of countries based on the production of magnesium magnesium compounds, standard atomic weight: magnesium. ” CIAAW. 2011. 4.61 ^ mg (0) was synthesized in a compound containing a cluster Na2MG22+ coordinated with massive organic ligand; Watch Rasch, b.; Gentner, Texas; Eyselein, J.; Langer, J.; Elsen, x.; Li, W.; Harder, S. (2021). "Reduce highly magnesium complexes (0)." Nature.
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