

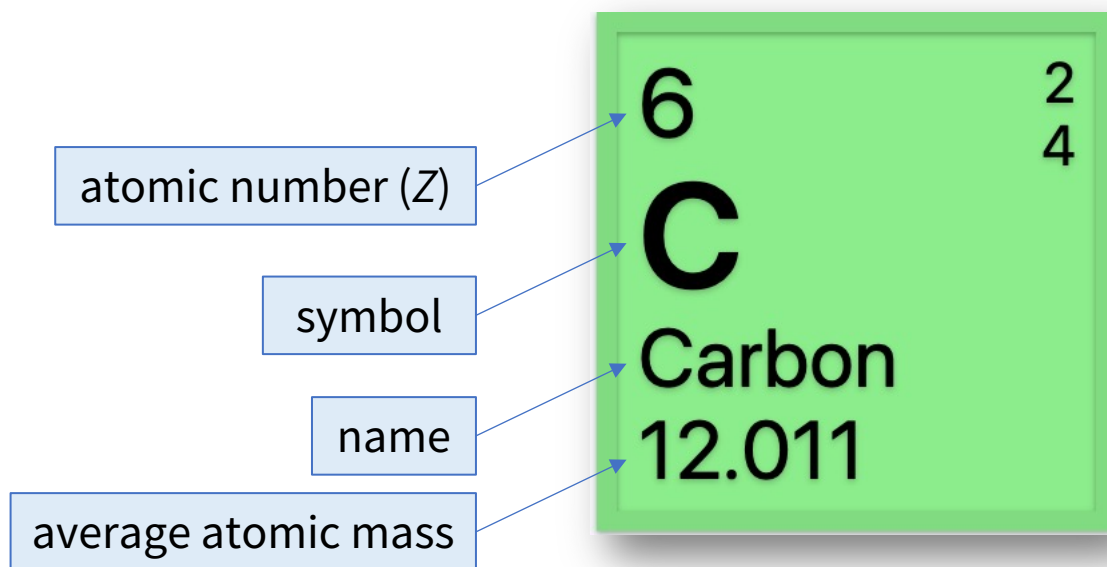
# **Module 2: What is a “Chemical Species”?**

## **The Periodic Table**

Fundamentals of Chemistry Open Course

1. State and apply the laws of chemical combination.
2. State and apply the tenets of the modern atomic theory.
3. Visualize the subatomic particles that constitute the atom using a simple planetary model; count subatomic particles using atomic number ( $Z$ ) and mass number ( $M$ ).
4. Represent an atom or ion using an atomic symbol.
5. Represent the number ratios of atoms in a compound using a chemical formula.
6. Visualize and distinguish between submicroscopic models of molecular and ionic compounds.
7. Use the periodic table to efficiently find information about a chemical element.
8. Recognize key collections of elements on the periodic table.
9. Determine the name of a binary ionic compound from the chemical formula and *vice versa*.
10. Determine the name of a simple molecular compound from the chemical formula and *vice versa*.

- Trends in the properties and chemistry of the elements observed throughout the 19<sup>th</sup> century led Mendeleev to organize them into a table, which has evolved to become the modern **periodic table**.
- Each entry in the table includes information about a single chemical element:
  - Name
  - Symbol
  - Atomic number
  - **Average atomic mass:** a weighted average of the masses of naturally occurring isotopes
  - ...and sometimes more!



# Organization of the Chemical Elements

- Each column (**group**) includes elements with analogous properties and reactivity.
- Across each row (**period**), consistent trends in properties such as the size of the atom are generally observed.

Group Period →	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
1	1 H																	2 He
2	3 Li	4 Be											5 B	6 C	7 N	8 O	9 F	10 Ne
3	11 Na	12 Mg											13 Al	14 Si	15 P	16 S	17 Cl	18 Ar
4	19 K	20 Ca	21 Sc	22 Ti	23 V	24 Cr	25 Mn	26 Fe	27 Co	28 Ni	29 Cu	30 Zn	31 Ga	32 Ge	33 As	34 Se	35 Br	36 Kr
5	37 Rb	38 Sr	39 Y	40 Zr	41 Nb	42 Mo	43 Tc	44 Ru	45 Rh	46 Pd	47 Ag	48 Cd	49 In	50 Sn	51 Sb	52 Te	53 I	54 Xe
6	55 Cs	56 Ba	* 71 Lu	72 Hf	73 Ta	74 W	75 Re	76 Os	77 Ir	78 Pt	79 Au	80 Hg	81 Tl	82 Pb	83 Bi	84 Po	85 At	86 Rn
7	87 Fr	88 Ra	* 103 Lr	104 Rf	105 Db	106 Sg	107 Bh	108 Hs	109 Mt	110 Ds	111 Rg	112 Cn	113 Nh	114 Fl	115 Mc	116 Lv	117 Ts	118 Og
			* 57 La	58 Ce	59 Pr	60 Nd	61 Pm	62 Sm	63 Eu	64 Gd	65 Tb	66 Dy	67 Ho	68 Er	69 Tm	70 Yb		
			* 89 Ac	90 Th	91 Pa	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No		



- Most of the elements are **metals**:
  - Silver-colored or dark-colored solids
  - Electrically and thermally conductive
  - Malleable
  - Oxidizable (cation forming)
  - Found mostly in ionic compounds



- A much smaller number are **nonmetals**:
  - Solids, liquids, or (most commonly) gases with various colors
  - Electrical insulators
  - Reducible (anion forming)
  - Found in molecular or covalent compounds



- A few elements are **metalloids**, with properties of both metals and nonmetals:
  - Electrically semiconductive
  - Found in both ionic and molecular or covalent compounds

# Metals, Nonmetals, and Metalloids

- Most of the elements are metals; metals are generally found to the left and at the bottom of the table.
- Nonmetals are found in the top right of the table. **Hydrogen is a nonmetal.**
- The elements boxed in red along the “diagonal” are metalloids.

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5	37 Rb	38 Sr	39 Y	40 Zr	41 Nb	42 Mo	43 Tc	44 Ru	45 Rh	46 Pd	47 Ag	48 Cd	49 In	50 Sn	51 Sb	52 Te	53 I	54 Xe
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			* 89 Ac	90 Th	91 Pa	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No		

# Blocks of Elements

- The table is divided into four rectangular **blocks**: the *s* block, *p* block, *d* block, and *f* block.

Group Period	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
1	1 H		<b>s block</b>										<b>p block</b>					2 He
2	3 Li	4 Be											5 B	6 C	7 N	8 O	9 F	10 Ne
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			* 57 La	58 Ce	59 Pr	60 Nd	61 Pm	62 Sm	63 Eu	64 Gd	65 Tb	66 Dy	67 Ho	68 Er	69 Tm	70 Yb	<b>f block</b>	
			* 89 Ac	90 Th	91 Pa	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No		

# Diatomic Elements

- Seven nonmetals are diatomic in their elemental forms; these are worth committing to memory.
- $\text{Br}_2$ ,  $\text{I}_2$ ,  $\text{N}_2$ ,  $\text{Cl}_2$ ,  $\text{H}_2$ ,  $\text{O}_2$ ,  $\text{F}_2$ : “Brinklehoff”

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- Why is the periodic table arranged the way it is? For example, why are there two columns in the  $s$  block, six in the  $p$  block, ten in the  $d$  block, and fourteen in the  $f$  block?
- One reason the periodic table is useful is that the elements display **periodic trends** in their properties. What are these properties and how do they vary across the periodic table?