

Chem1

Your Name

Academic Year 2025-2026

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Lecture 1: GasLaw

Fri 05 Dec

Molar Volume at STP is $22.4 \frac{L}{mol}$ for any gas

Gas mixture is composed of He and N_2 and has pressure of 0.50 atm

Wht is the partial pressure of He if mole fraction is N_2 is 0.32

$$P_{total} = P_{He} + P_{N_2}$$

$$\begin{cases} \text{KE} &= \frac{1}{2} mv^2 \\ \text{KE} &= \frac{3}{2} \frac{RT}{N_2} \end{cases}$$

Lecture 2: ChemFigPractice

Fri 05 Dec

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