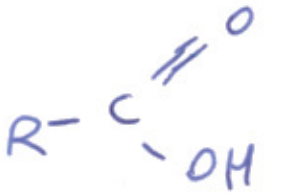
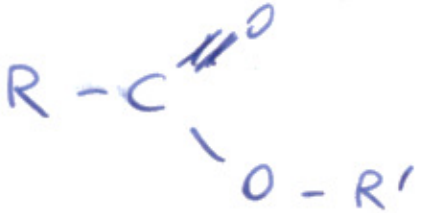
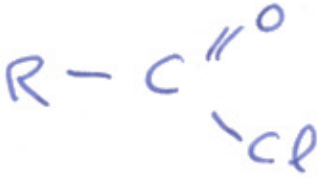


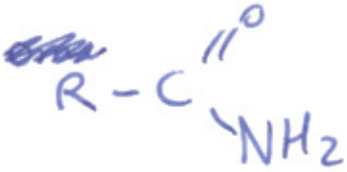
acide carboxylique



ester.



halogénure d'acyle.



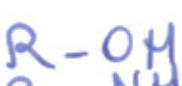
~~oamide~~. -oamide.

nitrile -nitrile

aldehyde -al.

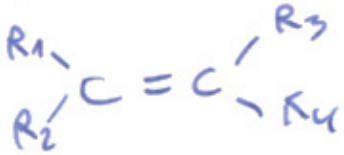
cétone -one.

alcool -ol.



amine -amide.

ether R-oxy-



alcène -ène

alcyne. -yne

- énantiomères : st images par un miroir et non superposables.

- diastéréo :

CIP. S → sens trigo

R → sens antitrigo

C – H aromatique 3000 cm^{-1}

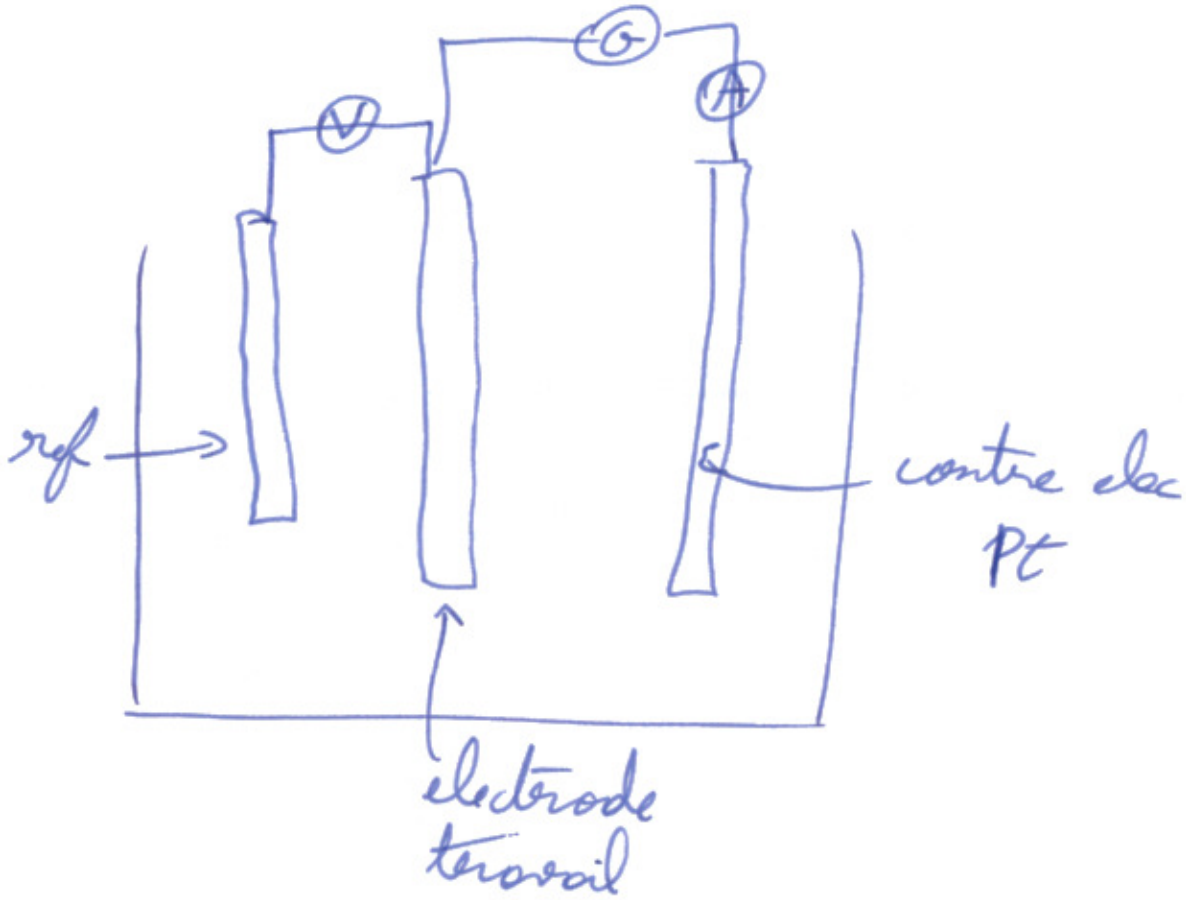
C = C aromatique 1650 cm^{-1}

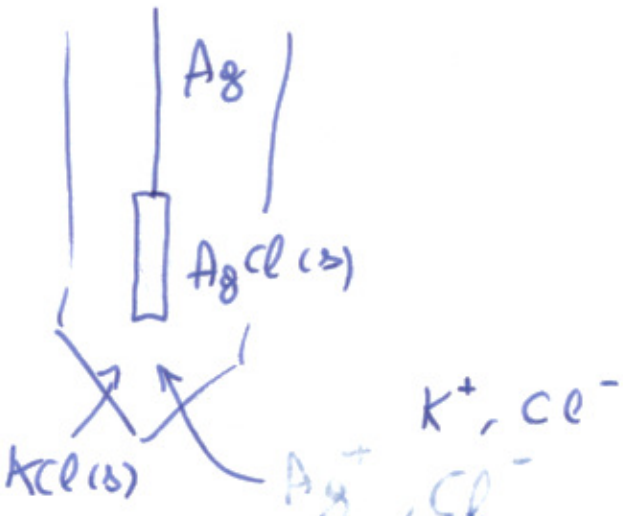
$$Hg_2Cl_{2(s)} + 2e^- = 2Hg_{(l)} + 2Cl^-$$

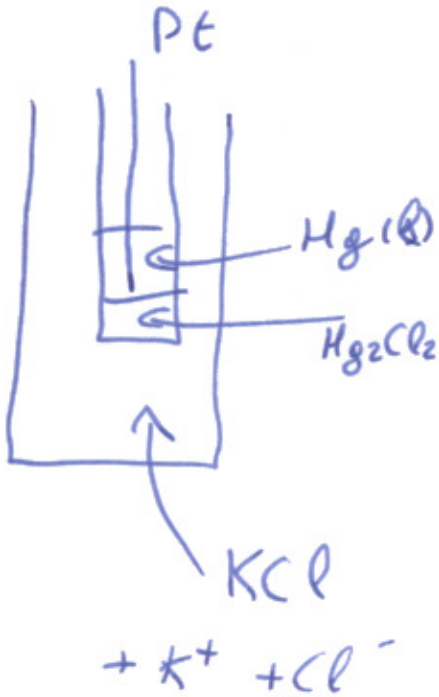
$$E_{Hg_2Cl_2} = E^\circ + \frac{0,06}{2} \log \frac{1}{[Cl^-]^2}.$$

$$K_S = \underbrace{[Cl^-][K^+]}_{\sqrt{K_S}}. = s^2.$$

$$= E^\circ - \frac{0,06}{2} \log K_S.$$

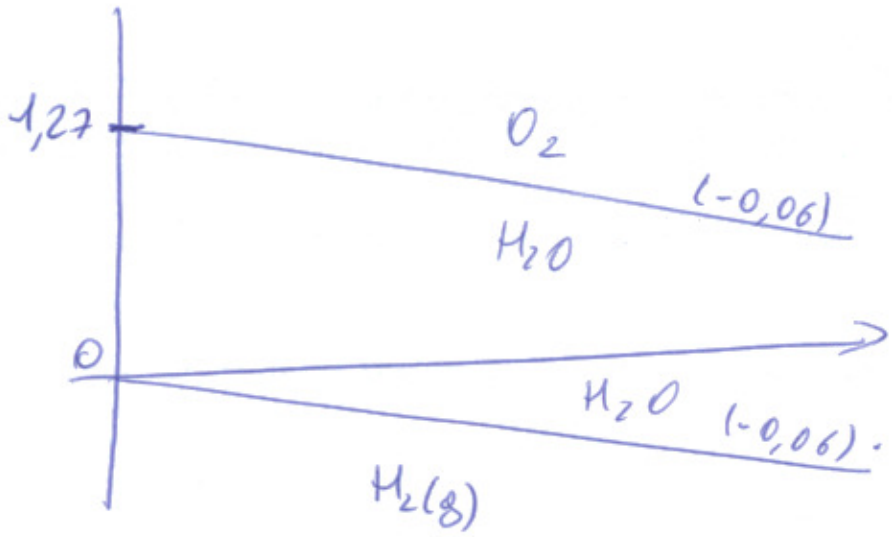

G
V
A
réf
contre élec
Pt
électrode
travail


Ag
AgCl(s)
KCl(s)
K+, Cl-
Ag+, Cl-


Pt
Hg(l)
Hg2Cl2
KCl
+ K+ + Cl-

▷ convention espèce $[I_2] + [I^-] = C_t$

▷ convention atomique $2[I_2] + [I^-] = C_t$



Gibbs Helmholtz :

$$\left.\frac{\partial\left(\frac{\Delta_r G}{T}\right)}{\partial T}\right|_{p,n_i} = -\frac{\Delta_r H}{T^2}$$

Van't Hoff :

$$\frac{d \ln K^\circ(T)}{dT} = \frac{\Delta_r H^\circ}{RT^2}$$

Arrhénius :

$$\frac{d \ln k}{dT} = \frac{E_a}{RT^2}.$$

Si → polaire

↳ composé polaire est plus retenu

↳ ——— apolaire migre ⊕.

Augmenté polarité éluant rot apolaire augmente la competition