CLASS 12 BATCH

FOR CHEMISTRY

LECTURE - 01

ELECTROCHEMISTRY

Today's Goal

Electro Chemical Cell

Electro Chemical Cell

The device which is used to convert chemical energy into electrical energy

Diagram

Salt Bridge

Function of Salt Bridge

Liquid – Liquid Junction Potential

Op Point – KCl can not be used as inert electrolyte when following Metal electrodes are used

Cell Reaction

Cell Representation

Cell Potential Difference

Cell Potential Difference

Standard Hydrogen Electrode

Electrode Potential

Q

Emf of cell Ni | Ni^{2+} (1.0 M) || Au^{3+} (1.0 M) | Au is ..., If E^0 for Ni^{2+} | Ni is −0.25 V, E^0 for Au^{3+} | Au is 1.50 V

A + 1.25 V

B - 1.75 V

C + 1.75 V

D + 4.0 V

Q Saturated solution of KNO_3 is used to make 'salt-bridge' because -

A Velocity of K^+ is greater than of NO_3^-

B Velocity of NO_3^- is greater than that of K^+

C Velocity of both K^+ and NO_3^- are nearly the same

D KNO_3 is highly soluble in water

THANK YOU !!

Homework

REVISE FORMULA OF LAST CHAPTER

DPP Of this Lecture