



CLASS 12 BATCH

FOR CHEMISTRY

LECTURE - 03

ELECTROCHEMISTRY



Today's Goal



**Free Energy relation
Equilibrium Constant
Concentration cell**



Relation b/w ΔG° (Standard free energy change) & E°_{cell}



Calculation of ΔG° , max work and eqⁿ const K for Daniel Cell





If the E° for a given reaction has a negative value, which of the following gives the correct relationships for the values of ΔG° and K_{eq} ? (NEET - II 2016)



$\Delta G^\circ < 0; K_{eq} > 1$



$\Delta G^\circ < 0; K_{eq} < 1$



$\Delta G^\circ > 0; K_{eq} < 1$



$\Delta G^\circ > 0; K_{eq} > 1$





THANK YOU !!

Homework

REVISE FORMULA OF LAST CHAPTER
DPP Of this Lecture

