

LECTURE - 01 ELECTROCHEMISTRY



Today's Goal



Electro Chemical Cell



Electro Chemical Cell

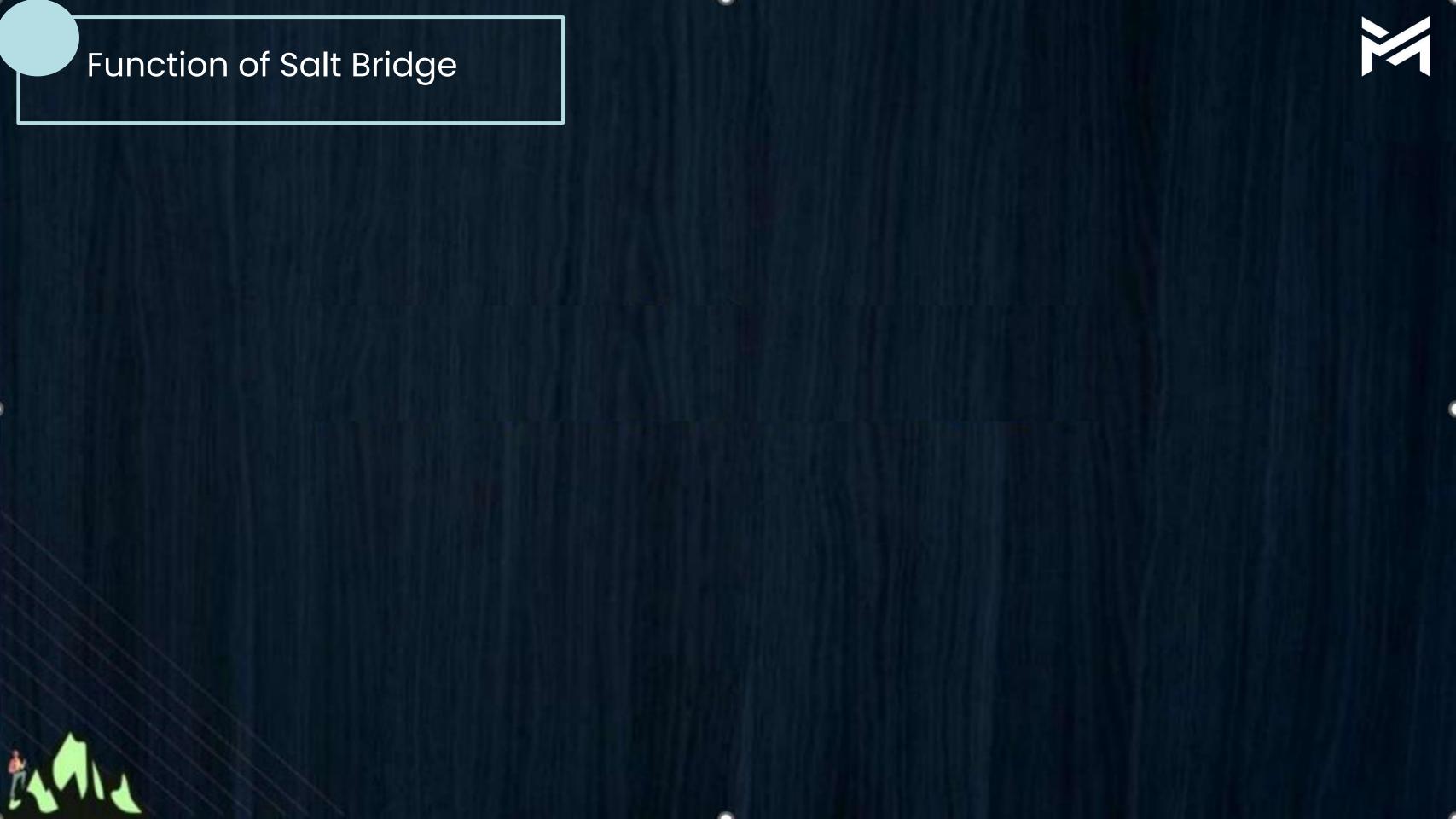


The device which is used to convert chemical energy into electrical energy









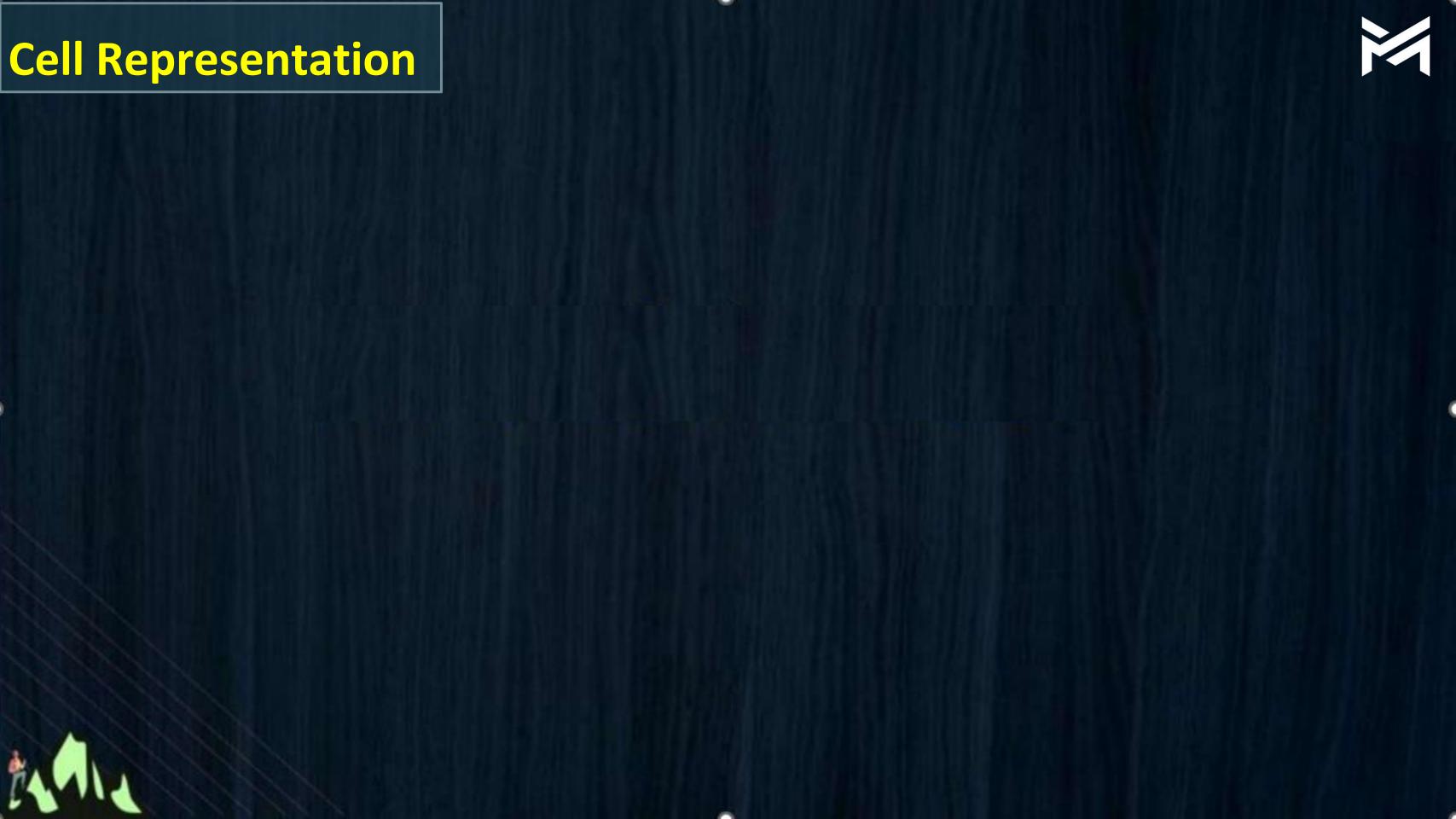
Liquid – Liquid Junction Potential

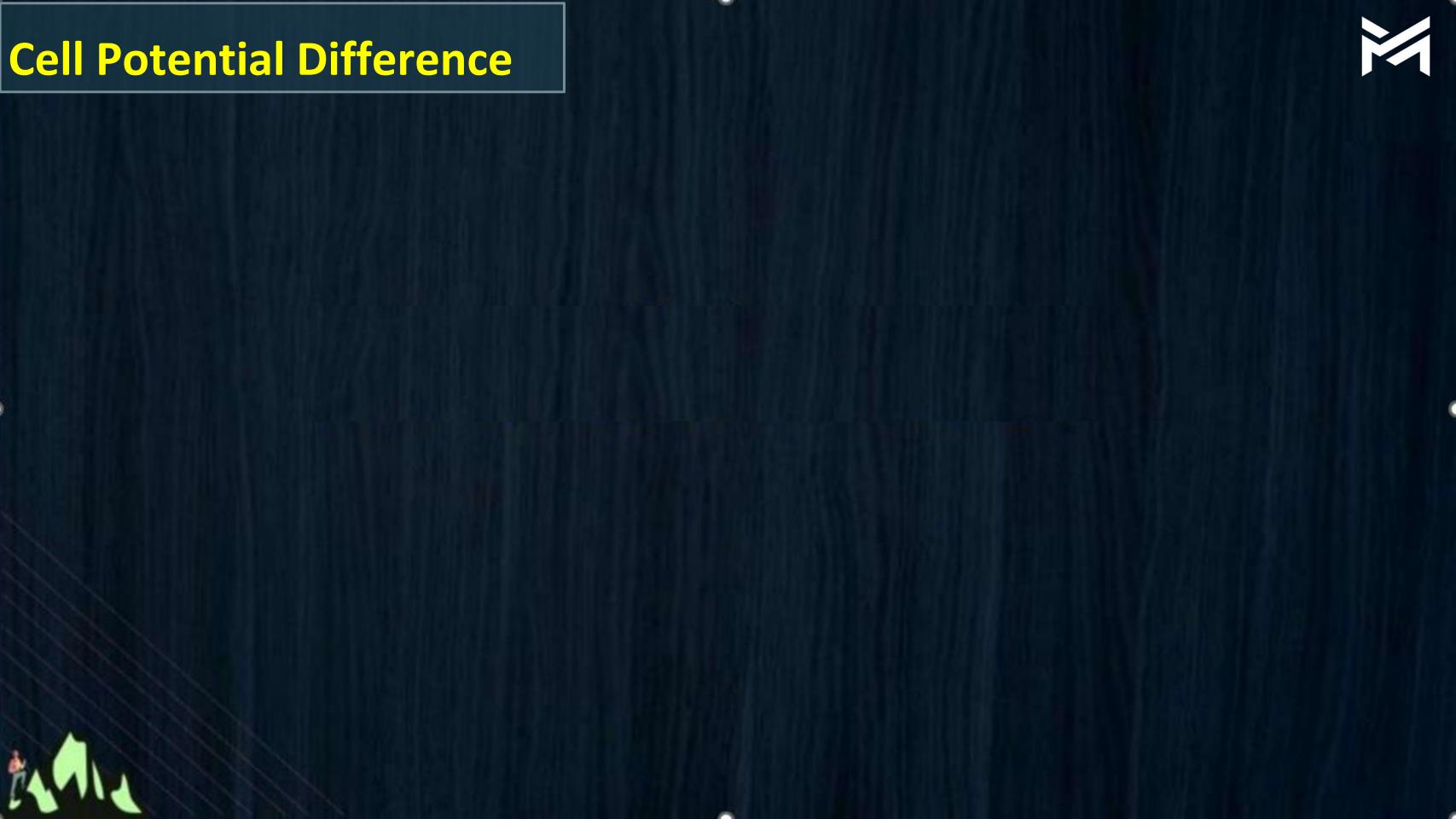


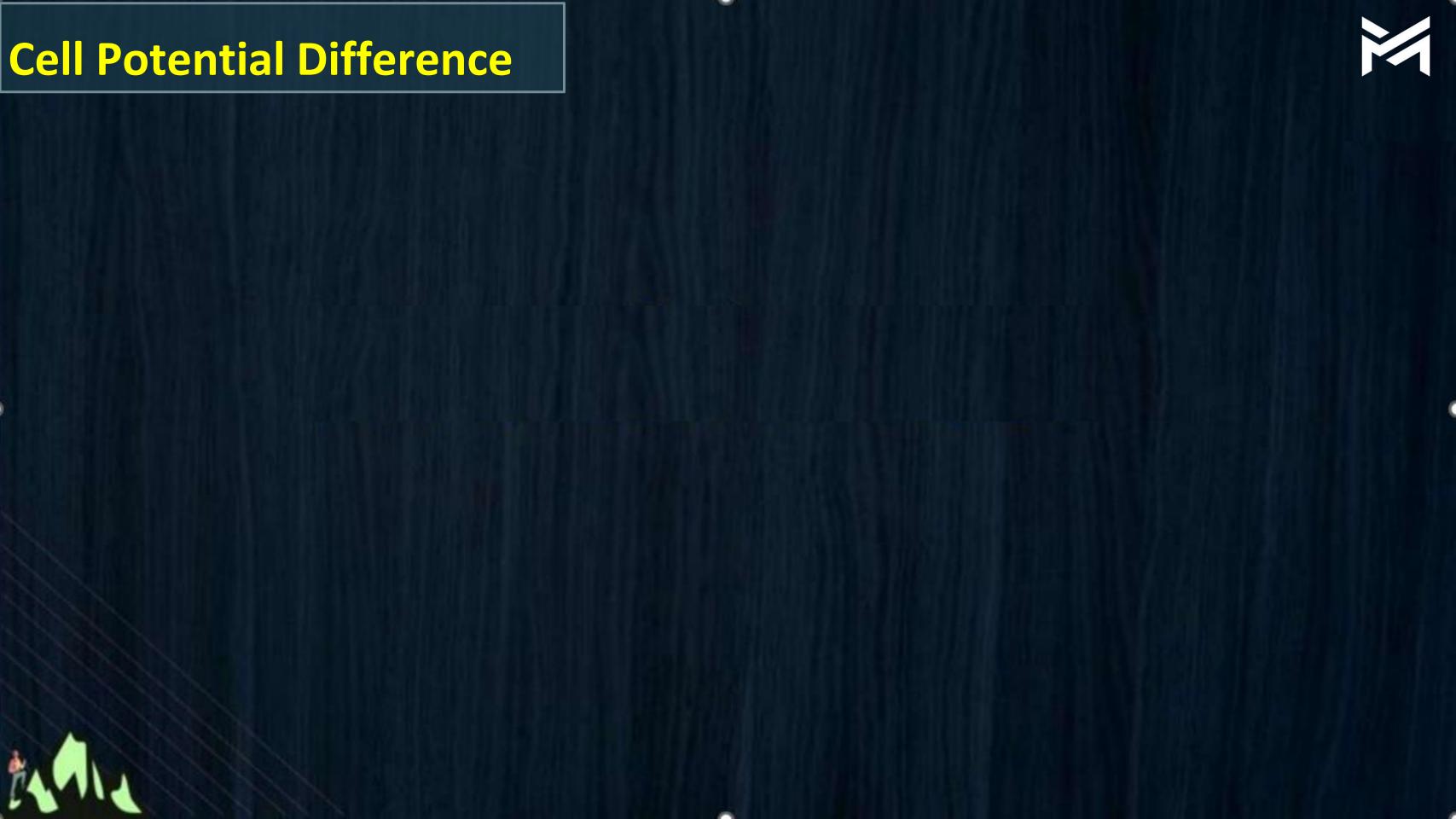
Op Point – KCl can not be used as inert electrolyte when following Metal electrodes are used



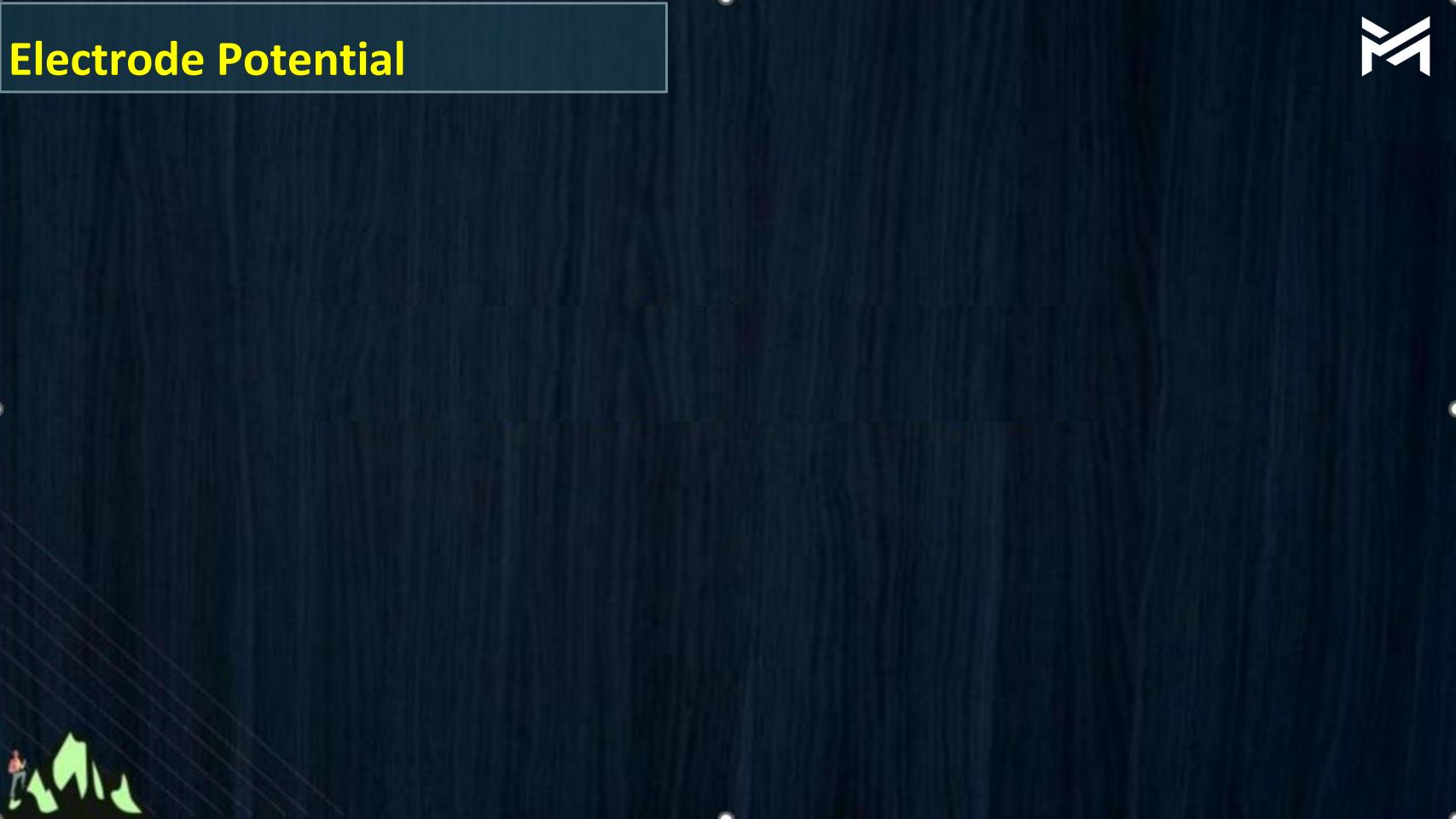














Emf of cell Ni | Ni²⁺ (1.0 M) || Au³⁺ (1.0 M) | Au is ..., If E^0 for Ni²⁺ | Ni is -0.25 V, E^0 for Au^{3+} | Au is 1.50 V





+ 1.25 V



+ 1.75 V



- 1.75 V



+ 4.0 V





Saturated solution of KNO₃ is used to make 'salt-bridge' because -





Velocity of K⁺ is greater than of NO₃⁻



Velocity of NO₃ is greater than that of K⁺



Velocity of both K⁺ and NO₃ are nearly the same



KNO₃ is highly soluble in water





THANK YOU!!

Homework

REVISE FORMULA OF LAST CHAPTER
DPP Of this Lecture

