

## LECTURE - 03 ELECTROCHEMISTRY

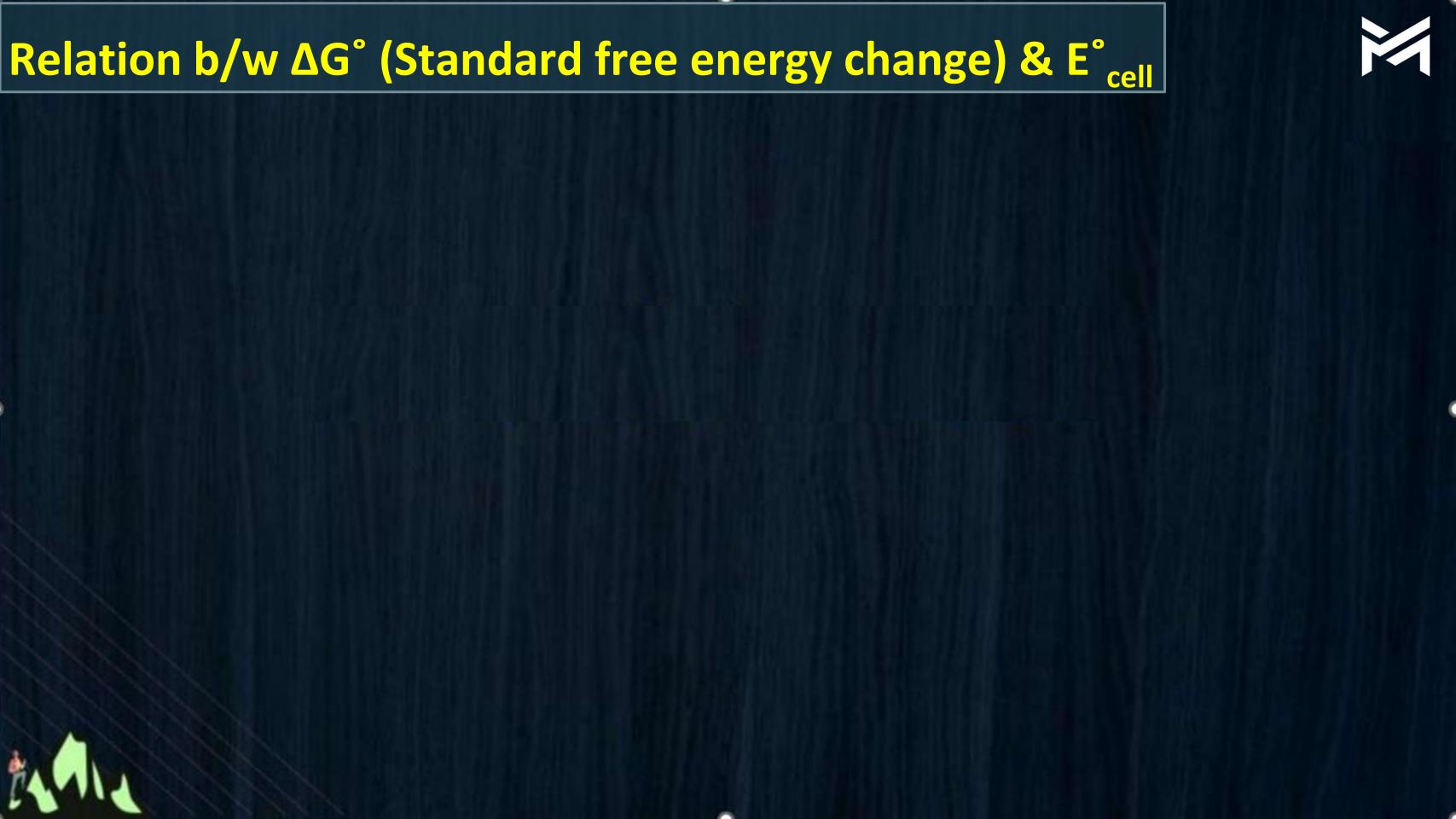


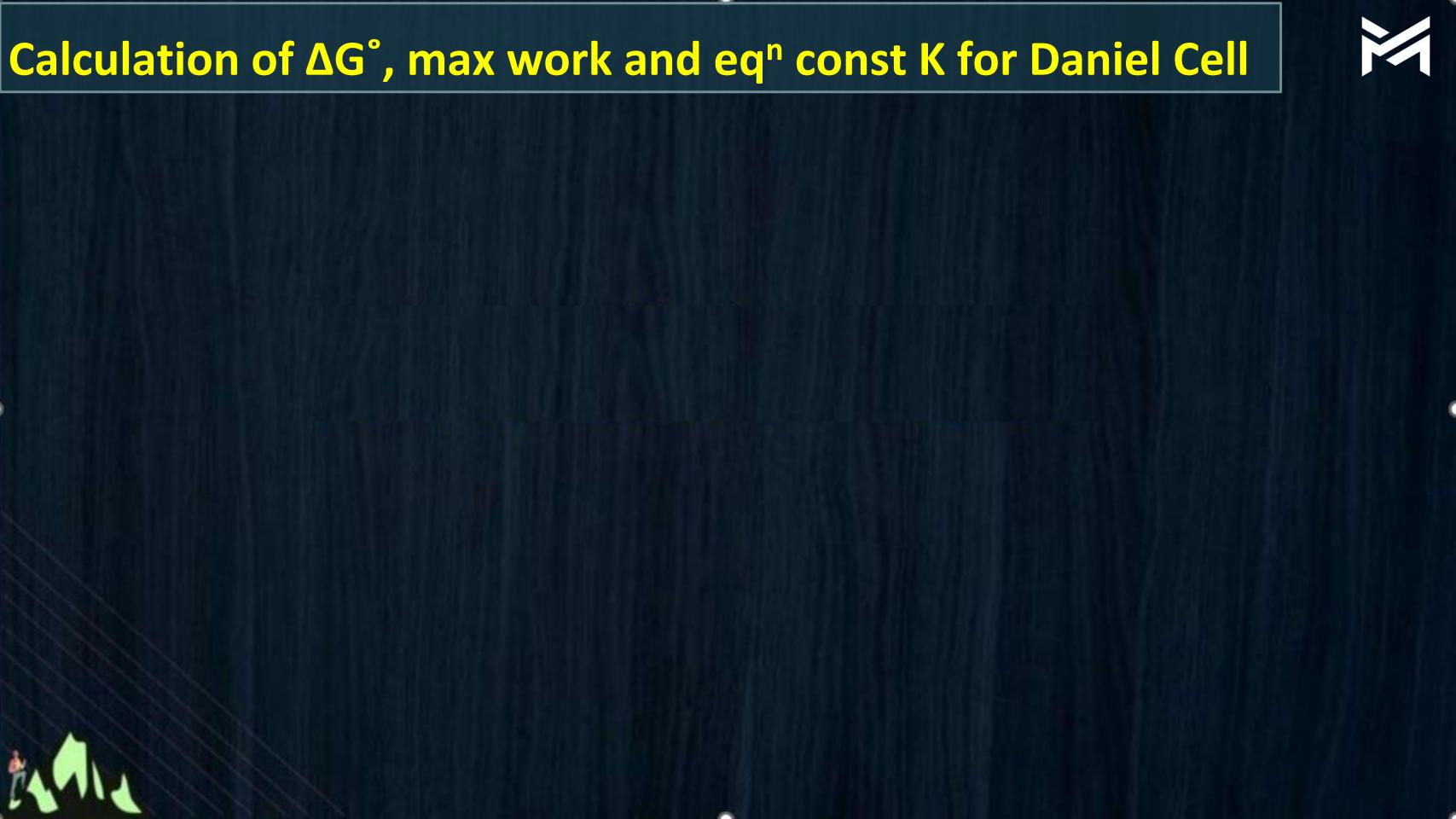
### Today's Goal



# Free Energy relation Equilibrium Constant Concentration cell







Q

If the E° for a given reaction has a negative value, which of the following gives the correct relationships for the values of  $\Delta G^{\circ}$  and  $k_{eq}$ ? (NEET - II 2016)



$$\Delta G^{\circ} < 0$$
; Keq > 1



$$\Delta G^{\circ} < 0$$
;  $K_{eq} < 1$ 



$$\Delta G^{\circ} > 0$$
;  $K_{eq} < 1$ 

$$\Delta G^{\circ} > 0$$
; Keq > 1





### THANK YOU!!

#### Homework

REVISE FORMULA OF LAST CHAPTER
DPP Of this Lecture

