

# Geofluids — Part IV

## Introduction to Geochemical and Reactive Transport Modeling

---

Dr. Svetlana Kyas (extension of the lecture notes of Dr. Allan Leal)

May 17, 2021

Postdoc Associate  
svetlana.kyas@erdw.ethz.ch  
NO F 61

## 1. Thermodynamic models of aqueous, gaseous, and mineral phases

### Aqueous phase

Activity coefficient model for aqueous ionic species, Davies model

Activity coefficient model for  $\text{H}_2\text{O}(\text{aq})$ , Davies model

Activity coefficient model for  $\text{CO}_2(\text{aq})$ , Drummond model

### Gaseous phase

### Mineral phase

## Numerical method for chemical equilibrium calculation

---

# Recap of a chemical equilibrium problem

- Calculate the amounts of the species in each phase at equilibrium at given:
  - **temperature** (e.g., 100 °C);
  - **pressure** (e.g, 300 bar); and
  - **amounts of elements H, O, C, Na, Cl**
  - **Note:** the latter is rather given as a recipe that determine the values of H, O, C, Na, Cl by formula matrix  $b = A n$ , e.g.,
    - 1 kg of H<sub>2</sub>O,
    - 1 mol CO<sub>2</sub>,
    - 0.1 mol of NaCl.

Aqueous Phase	Gaseous Phase	Mineral Phases
H <sub>2</sub> O(aq)	CO <sub>2</sub> (g)	NaCl(s, halite)
H <sup>+</sup> (aq)	H <sub>2</sub> O(g)	
OH <sup>-</sup> (aq)		
Na <sup>+</sup> (aq)		
Cl <sup>-</sup> (aq)		
CO <sub>2</sub> (aq)		
HCO <sub>3</sub> <sup>-</sup> (aq)		
CO <sub>3</sub> <sup>2-</sup> (aq)		

# Chemical equilibrium equations, Fundamental condition

**Fundamental condition for chemical equilibrium** reads as follows:

## Gibbs energy minimization (GEM) problem

Given temperature  $T$ , pressure  $P$ , and elements amounts  $b = (b_1, \dots, b_E)$ , find the amounts of the species  $n = (n_1, \dots, n_N)$  that solve the **constrained minimization problem**:

$$\min_n G(n) \quad \text{subject to (s.t.)} \quad An = b \quad \text{and} \quad n \geq 0.$$

How do we transform it into a system of equations from which species amounts  $n = (n_1, \dots, n_N)$  can be calculated?

# Numerical method for chemical equilibrium calculation

---

Unconstrained minimization

# Chemical equilibrium equations, Finding the minimum

- Assume there were **no constraints** of mass conservation and non-negative bounds.
- Then, the **chemical equilibrium equations** would result from the condition

$$\frac{\partial G}{\partial n_i} = 0 \quad (i = 1, \dots, N),$$

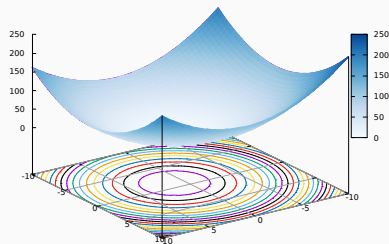
which corresponds to a state in which the **Gibbs energy does not change with infinitesimal changes in the amounts of the species.**

- **But!** some care would still be needed to ensure that a **minimum is found**, and not a **maximum** or **saddle point**.

# Example of finding the minimum

- Consider the function  $f$  defined as:

$$f(x, y) = (x - 1)^2 + (y + 1)^2.$$



- The values of  $(x, y)$  that minimizes  $f$  can be found using:

$$\frac{\partial f}{\partial x} = 0 \implies 2(x - 1) = 0 \implies x = 1$$

$$\frac{\partial f}{\partial y} = 0 \implies 2(y + 1) = 0 \implies y = -1.$$

- Question:** How can we tell, mathematically, that the solution above corresponds to a minimum?

<http://etc.ch/cKvh> or



- Answer:** Check if the derivatives  $\partial^2 f / \partial x^2$  and  $\partial^2 f / \partial y^2$  at  $(1, -1)$  are non-negative.



# Distinguishing between minimum, maximum, saddle point values

- In general, one checks the **eigenvalues of Hessian matrix**

$$H = \begin{bmatrix} \frac{\partial^2 f}{\partial x^2} & \frac{\partial^2 f}{\partial x \partial y} \\ \frac{\partial^2 f}{\partial y \partial x} & \frac{\partial^2 f}{\partial y^2} \end{bmatrix}$$

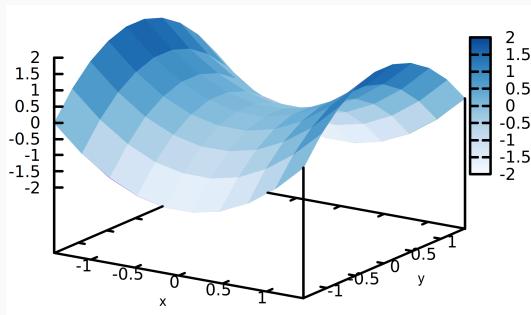
at obtained stationary point  $(x_*, y_*)$ .

- If **the eigenvalues are non-negative** at the  $(x_*, y_*)$  (or Hessian is a positive semi-definite matrix at  $(x_*, y_*)$ ), this point is indeed a **local minimum**.
- If **the eigenvalues are negative** at the  $(x_*, y_*)$  (or Hessian is negative-definite at  $(x_*, y_*)$ ), this point is a **local maximum**.
- If **the eigenvalues are of different signs** at the  $(x_*, y_*)$ , this point is a **saddle point**.

# Example of saddle point problem

- Consider the function  $f$  defined as:

$$f(x, y) = x^2 - y^2.$$



- At  $(x, y) = (0, 0)$ :

$$\frac{\partial f}{\partial x} = \frac{\partial f}{\partial y} = 0,$$

but  $f(0, 0) = 0$  is **neither a minimum nor a maximum**.

- Quiz:** What are the signs of  $\partial^2 f / \partial x^2$  and  $\partial^2 f / \partial y^2$  for this example?



<http://etc.ch/cKvh> or

- Answer:** The values of  $\partial^2 f / \partial x^2 > 0$  and  $\partial^2 f / \partial y^2 < 0$  at  $(0, 0)$ , which implies a **saddle point**.

# Numerical method for chemical equilibrium calculation

---

Constrained minimization

# Lagrangian function of chemical equilibrium problem

- Because of the **mass conservation constraints**,  $An = b$ , and the **non-negative constraints**,  $n \geq 0$ , we need to write the **Lagrange function**  $L$  for the Gibbs energy minimization problem:

$$L(n, y, z) := G(n) - (An - b)^T y - n^T z,$$

where the **unknowns**  $n$ ,  $y$ , and  $z$  correspond to

- the vector of *species amounts*  $n = (n_1, \dots, n_N)$ ,
- the vector of *Lagrange multipliers*  $y = (y_1, \dots, y_E)$ , and
- the vector of *complementary variables*  $z = (z_1, \dots, z_N)$  that satisfy

$$n_i z_i = 0, \quad n_i \geq 0, \quad z_i \geq 0.$$

- **Note:** The above form of the Lagrangian function is called **matrix form** but one can write it explicitly for each unknown  $n_i$ ,  $y_j$ , and  $z_i$ .

## System of equation corresponding to minimization of Lagrangian

- Instead of finding **only**  $n = (n_1, \dots, n_N)$  that solves

$$\frac{\partial G}{\partial n_i} = 0 \quad (i = 1, \dots, N),$$

we find  $n = (n_1, \dots, n_N)$ ,  $y = (y_1, \dots, y_E)$ , and  $z = (z_1, \dots, z_N)$  that solves

$$\frac{\partial L}{\partial n_i} = 0 \quad (i = 1, \dots, N),$$

$$\frac{\partial L}{\partial y_j} = 0 \quad (j = 1, \dots, E),$$

$$n_i z_i = 0 \quad (i = 1, \dots, N),$$

$$n_i, z_i \geq 0 \quad (i = 1, \dots, N).$$

## System of equation corresponding to minimization of Lagrangian ii

We rewrite the system of equations

using the definition of  $\mu_i := \partial G / \partial n_i$ :

$$\frac{\partial L}{\partial n_i} = 0 \quad (i = 1, \dots, N),$$

$$\frac{\partial L}{\partial y_j} = 0 \quad (j = 1, \dots, E),$$

$$n_i z_i = 0 \quad (i = 1, \dots, N),$$

$$n_i, z_i \geq 0 \quad (i = 1, \dots, N),$$

$$\mu_i - \sum_{j=1}^E A_{ji} y_j - z_i = 0 \quad (i = 1, \dots, N),$$

$$\sum_{i=1}^N A_{ji} n_i - b_j = 0 \quad (j = 1, \dots, E),$$

$$n_i z_i = 0 \quad (i = 1, \dots, N),$$

$$n_i, z_i \geq 0 \quad (i = 1, \dots, N).$$

**Note:** Since  $\mu_i$  is a function of unknown  $n_i$ , we obtain **non-linear system of equation**.

## Interior-point method for constrained minimization

---

# Abstract constrained minimization formulation

- Solve the following general **constrained minimization problem**:

$$\min_x f(x) \quad \text{subject to} \quad Ax = b \quad \text{and} \quad x \geq 0.$$

- **Quiz**: What would be  $f(x)$  and  $x$  in a context of chemical equilibrium problem?



<http://etc.ch/cKvh> or

- **Answer**: We will define  $f$  as our Gibbs energy function

$$f(x) = G(x; T, P),$$

where

- $x = (x_1, \dots, x_n)$  represents the **unknown species amounts**,  $n = (n_1, \dots, n_N)$ ,
- $A$  specifies the **formula matrix of the system**, and
- $b$  stands for the **vector of element amounts**.



# Lagrangian of general minimization formulation

- First, we write the **Lagrange function for the general minimization problem**:

$$L(x, y, z) = f(x) - (Ax - b)^T y - x^T z.$$

- Then, we formulate its **first-order necessary conditions for the minimum**:

$$\partial f / \partial x - A^T y - z = 0,$$

$$Ax - b = 0,$$

$$x_i z_i = 0 \quad (i = 1, \dots, n),$$

$$x_i, z_i \geq 0 \quad (i = 1, \dots, n).$$

# Complementarity conditions

- The **complementarity conditions**

$$x_i z_i = 0$$

$$x_i \geq 0$$

$$z_i \geq 0$$

are **challenging to solve numerically**.

- They are sharp conditions, in which

$$x_i \geq 0 \quad \text{and} \quad z_i = 0$$

or

$$x_i = 0 \quad \text{and} \quad z_i \geq 0.$$

# Interior-point method for constrained minimization

---

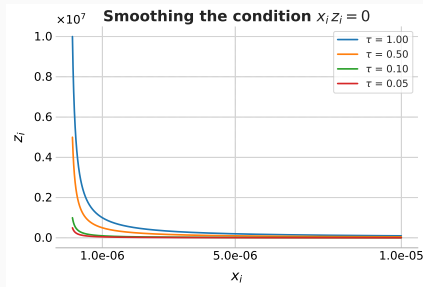
Interior-point perturbation approach

# Interior-point perturbation approach

- Perturb  $x_i z_i = 0$  with a **small and constant number**

$$x_i z_i = \tau.$$

- **Note:** the choice of  $\tau$  affects the smallest amount a species that can exist in equilibrium.
- **Example:** the choice  $\tau = 10^{-20}$  causes species in **unstable phases** to have amounts around this value.
- **Quiz:** What can we say about the amount of mineral dissolved in water?



<http://etc.ch/cKvh> or



- **Answer:** The amount will be of order  $\tau$ .

# Interior-point Newton method

- We'll use **Newton's method** to solve the obtained system of non-linear equations:

$$\frac{\partial f}{\partial x} - A^T y - z = 0,$$

$$Ax - b = 0,$$

$$x_i z_i = \tau \quad (i = 1, \dots, n),$$

$$x_i, z_i \geq 0 \quad (i = 1, \dots, n).$$

- Define the **residual function**  $F$  as follows:

$$F(x, y, z) := \begin{bmatrix} \frac{\partial f}{\partial x} - A^T y - z \\ Ax - b \\ XZe - \tau \end{bmatrix},$$

where

- $e = (1, \dots, 1)^T$ ,
- $X = \text{diag}(x)$ , and
- $Z = \text{diag}(z)$ .

The resulting problem: find  $(x, y, z)$ , with  $x, z \geq 0$ , so that:

$$F(x, y, z) = 0.$$

# Jacobian matrix for the interior-point Newton method

- Newton's method require the Jacobian  $J$  of the residual function:

$$F(x, y, z) := \begin{bmatrix} \frac{\partial f}{\partial x} - A^T y - z \\ Ax - b \\ XZe - \tau \end{bmatrix},$$

which is defined as

$$J := \begin{bmatrix} \frac{\partial F_1}{\partial x} & \frac{\partial F_1}{\partial y} & \frac{\partial F_1}{\partial z} \\ \frac{\partial F_2}{\partial x} & \frac{\partial F_2}{\partial y} & \frac{\partial F_2}{\partial z} \\ \frac{\partial F_3}{\partial x} & \frac{\partial F_3}{\partial y} & \frac{\partial F_3}{\partial z} \end{bmatrix}.$$

- **Exercise:** derive the matrix blocks in Jacobian of  $J$ .

- **Answer:**

$$J = \begin{bmatrix} H & -A^T & -I \\ A & 0 & 0 \\ Z & 0 & X \end{bmatrix}.$$

where

- $I$  be the identity matrix and
- $H$  represent the Hessian matrix of  $f$  (second derivatives of  $f$ )

$$H := \frac{\partial^2 f}{\partial x^2}.$$

- **How can we calculate Hessian matrix  $H$ ?**

# Interior-point method for constrained minimization

---

Hessian matrix of the Gibbs energy  
function

# Hessian matrix of the Gibbs energy function

- The **Hessian matrix of the Gibbs energy** is

$$H := \frac{\partial^2 G}{\partial n^2} \quad \text{or in element-wise notation} \quad H := \{H_{ij}\}_{ij} := \left\{ \frac{\partial^2 G}{\partial n_i \partial n_j} \right\}_{ij}$$

- Recall that

$$\frac{\partial G}{\partial n_i} = \mu_i \quad \text{and} \quad \mu_i = \mu_i^o + RT \ln a_i.$$

- Thus,

$$H_{ij} = \frac{\partial \mu_i}{\partial n_j} = \frac{\partial}{\partial n_j} (\mu_i^o + RT \ln a_i) = RT \frac{\partial \ln a_i}{\partial n_j}.$$

- Expressions for  $\partial \ln a_i / \partial n_j$  can be derived for all previous activity models.



## Partial molar derivatives of activities

- Instead of exact calculation of partial molar derivatives, we **derive them for ideal models**.
- We use activity models to **correct for the non-ideal behavior** of the aqueous and gaseous solutions.
- We use the partial molar derivatives of ideal activity models as it is **simpler to derive**.
- So, in

$$\begin{aligned}\mu_i - \sum_{j=1}^E A_{ji} y_j - z_i &= 0 & (i = 1, \dots, N), \\ \sum_{i=1}^N A_{ji} n_i - b_j &= 0 & (j = 1, \dots, E), \\ n_i z_i &= 0 & (i = 1, \dots, N), \\ n_i, z_i &\geq 0 & (i = 1, \dots, N),\end{aligned}$$

$\mu_i$  will be calculated exactly. But the **Jacobian will be approximated**.

- Ideally, an **automatic differentiation** scheme should be used to compute these derivatives (beyond the scope of this course).

## Partial molar derivatives of activities of aqueous solutes

- Recall that the **activity of an aqueous solute** is

$$a_i = \gamma_i m_i, \text{ where } m_i = 55.508 \frac{n_i}{n_w}.$$

- Its ideal activity model is  $a_i = m_i$ , with  $\gamma_i = 1$ .
- Thus,

$$\begin{aligned} \frac{\partial \ln a_i}{\partial n_j} &\approx \frac{\partial \ln m_i}{\partial n_j} = \text{Exercise} \\ &= \frac{\partial 55.508}{\partial n_j} + \frac{\partial \ln n_i}{\partial n_j} - \frac{\partial \ln n_w}{\partial n_j} = \begin{cases} \frac{1}{n_i} & j = i, \\ -\frac{1}{n_w} & j = w, \\ 0 & \text{otherwise.} \end{cases} \end{aligned}$$

## Partial molar derivatives of activities of aqueous solvent, $\text{H}_2\text{O}(\text{aq})$

- Recall that the **ideal activity of water**,  $\text{H}_2\text{O}(\text{aq})$ , is

$$a_w = -\frac{1 - x_w}{x_w}, \quad \text{where} \quad x_w = \frac{n_w}{\sum n_i}.$$

- Thus,

$$\begin{aligned} \frac{\partial \ln a_w}{\partial n_j} &= \text{Exercise} \\ &= \frac{\partial \ln(x_w - 1)}{\partial n_j} - \frac{\partial \ln x_w}{\partial n_j} = \begin{cases} -\frac{1}{n_w} & j = w, \\ -\frac{1}{n_w} \frac{x_w}{x_w - 1} & j \neq w. \end{cases} \end{aligned}$$

## Partial molar derivatives of activities of $\text{CO}_2(\text{g})$

- Recall that the **activity of  $\text{CO}_2(\text{g})$**  was calculated using

$$a_{\text{CO}_2(\text{g})} = \varphi_{\text{CO}_2(\text{g})} \frac{P}{P^\circ},$$

where the fugacity coefficient  $\varphi_{\text{CO}_2(\text{g})}$  was only a function of  $T$  and  $P$ , i.e.,

$$\varphi_{\text{CO}_2(\text{g})} = \varphi_{\text{CO}_2(\text{g})}(T, P).$$

- Thus, the value of the following derivative

$$\frac{\partial \ln a_{\text{CO}_2(\text{g})}}{\partial n_j} = \text{Exercise} \\ = 0.$$

## Partial molar derivatives of activities – pure minerals

- Recall that the **activities of pure minerals** (e.g., mineral phases with only one mineral end-member),

$$a_i = 1.$$

- Thus, for mineral species,

$$\frac{\partial \ln a_i}{\partial n_j} = 0.$$

# Calculation of the Hessian matrix of the Gibbs energy function in Python i

- Have a look in the Python code of the project to see the implementation of these formulas for when calculating the Hessian matrix  $H$  with entry  $H_{ij}$  given by:

$$H_{ij} = RT \frac{\partial \ln a_i}{\partial n_j} \quad \text{used in} \quad J = \begin{bmatrix} H & -A^T & -I \\ A & 0 & 0 \\ Z & 0 & X \end{bmatrix}.$$

```
1 # Define the function that calculates:
2 #   - G, the Gibbs energy of the system;
3 #   - u, the gradient of the Gibbs energy, or chemical potentials of the
  species
4 #   - H, the Hessian of the Gibbs energy, or partial molar derivatives of
  activities
5 # These quantities are normalized by RT for numerical reasons.
6 def gibbs_energy(T, P, n):
```

# Calculation of the Hessian matrix of the Gibbs energy function in Python ii

```
7   RT = R*T
8   u = chemical_potentials(T, P, n)/RT
9   H = chemical_potentials_ddn(T, P, n)/RT
10  G = n.dot(u)/RT
11  return G, u, H
12
13 # Define the function that calculates the partial molar derivatives of the
14 # chemical potentials of the species.
15 def chemical_potentials_ddn(T, P, n):
16     return R*T*ln_activities_ddn(T, P, n)
17
18 # Define the function that calculates the partial molar derivatives of the
19 # ln activities of all species
20 def ln_activities_ddn(T, P, n):
21     # Create an array with the entries in n corresponding to aqueous species
22     n_aqueous = n[slice_aqueous]
```

## Calculation of the Hessian matrix of the Gibbs energy function in Python iii

```
24 # The matrix with partial molar derivatives of the activities
25 ln_a_ddn = zeros((num_species, num_species))
26
27 ln_a_ddn[slice_aqueous, slice_aqueous] = ln_activities_aqueous_species_ddn(T
    , P, n_aqueous)
28
29 return ln_a_ddn
30
31 # Define the function that calculates the activities of the aqueous species.
32 # Parameters:
33 #   - T is temperature in units of K
34 #   - P is pressure in units of Pa
35 #   - nphase is an array with the mole amounts of the aqueous species
36 # Return:
37 #   - an array with the ln activities of aqueous species
38 def ln_activities_aqueous_species_ddn(T, P, nphase):
39     # The molar amount of H2O(l)
```



# Calculation of the Hessian matrix of the Gibbs energy function in Python iv

```
40  nH2O = nphase[iH2O]
41
42  # The mole fraction of H2O(l)
43  xH2O = nH2O/sum(nphase)
44
45  # Calculate the partial molar derivatives of the solute activities
46  ddn = diag(1.0/nphase)      # 1/ni, j = i
47  ddn[:, iH2O] = -1.0/nH2O    # -1/nw, j = w
48
49  # Calculate the partial molar derivatives of the solvent water activity
50  ddn[iH2O, :] = -1.0/nH2O * xH2O/(xH2O - 1.0) # -1/nw, j = w
51  ddn[iH2O, iH2O] = -1.0/nH2O                  # 1/nw*xw/(xw -1), j<>w
52
53  return ddn
```

**Listing 1:** Calculating Jacobian using Python

# Interior-point method for constrained minimization

---

Numerical solution by the interior-point  
Newton method

# Numerical solution by the interior-point Newton method

**Newton method algorithm** at step 1:

- At the first Newton iteration, start with **initial guess**

$$x^0 = (x_1^0, \dots, x_n^0), y^0 = (y_1^0, \dots, y_m^0) \text{ and } z^0 = (z_1^0, \dots, z_n^0).$$

- Compute  $x^1 = (x_1^1, \dots, x_n^1)$ ,  $y^1 = (y_1^1, \dots, y_m^1)$  and  $z^1 = (z_1^1, \dots, z_n^1)$  using

$$x^1 = x^0 + \alpha_x^0 \Delta x^0$$

$$y^1 = y^0 + \Delta y^0$$

$$z^1 = z^0 + \alpha_z^0 \Delta z^0$$

where  $\alpha_x$  and  $\alpha_z$  are introduced parameters to ensure  $x_i^1 > 0$  and  $z_i^1 > 0$ .

- How do we calculate variations  $\Delta x^0$ ,  $\Delta y^0$ ,  $\Delta z^0$ ?

## Numerical solution by the interior-point Newton method ii

- To calculate the Newton variations  $\Delta x^0 = (\Delta x_1^0, \dots, \Delta x_n^0)$ ,  $\Delta y^0 = (\Delta y_1^0, \dots, \Delta y_m^0)$  and  $\Delta z^0 = (\Delta z_1^0, \dots, \Delta z_n^0)$ , we solve the matrix equation

$$\underbrace{\begin{bmatrix} H^0 & -A^T & -I \\ A & 0 & 0 \\ Z^0 & 0 & X^0 \end{bmatrix}}_{J^0} \begin{bmatrix} \Delta x^0 \\ \Delta y^0 \\ \Delta z^0 \end{bmatrix} = - \underbrace{\begin{bmatrix} g^0 - A^T y^0 - z^0 \\ Ax^0 - b \\ X^0 Z^0 e - \tau \end{bmatrix}}_{F^0}$$

where  $J^0$ ,  $F^0$ ,  $H^0$ ,  $Z^0$ ,  $X^0$  represent to evaluation of respective matrices using the values of  $x$ ,  $y$ ,  $z$  at the 0th iteration, i.e.,

- $J^0 := J(x^0, y^0, z^0)$ ,
- $F^0 := F(x^0, y^0, z^0)$ , etc.

## Numerical solution by the interior-point Newton method iii

**Newton method algorithm** at the step  $k + 1$ :

- The vector steps  $\Delta x^k$ ,  $\Delta y^k$ ,  $\Delta z^k$  at the iteration  $k$  is calculated by

$$J^k \begin{bmatrix} \Delta x^k \\ \Delta y^k \\ \Delta z^k \end{bmatrix} = -F^k$$

- New approximations  $x^{k+1}$ ,  $y^{k+1}$ ,  $z^{k+1}$  is obtained by

$$\begin{aligned} x^{k+1} &= x^k + \alpha_x^k \Delta x^k, \\ y^{k+1} &= y^k + \Delta y^k, \\ z^{k+1} &= z^k + \alpha_z^k \Delta z^k. \end{aligned}$$

- How do we calculate  $\alpha_x^k$  and  $\alpha_z^k$  reassuring positivity of  $x_i^{k+1}$  and  $z_i^{k+1}$ ?

## Numerical solution by the interior-point Newton method iv

- Start with  $\alpha_x^k \leftarrow 1$ .
- Check if

$$x_i^k + \Delta x_i^k \leq 0 \quad \text{for } i = 1, \dots, n$$

- if so,  $\alpha_x^k \leftarrow (\delta - 1) \frac{x_i^k}{\Delta x_i^k}$ , where  $\delta \in (0, 1)$ , e.g.,  $\delta = 10^{-4}$ ;
  - otherwise, go to next component with index  $i$ .
- It results into either a **full Newton step** (for  $\alpha_x^k = 1$ ) or a **partial Newton step** (for  $\alpha_x^k < 1$ ).
- If for certain index  $i$ , parameter  $\alpha_x^k < 1$ , then

$$x_i^{k+1} = \delta x_i^k,$$

and for all other indices

$$x_j^{k+1} = x_j^k + \alpha_x^k \Delta x_j^k, \quad j = 1, \dots, n, j \neq i.$$

## Summary of the Newton method algorithm:

- With initial guess  $x^0$ ,  $y^0$ , and  $z^0$ , repeat

$$\begin{aligned}x^{k+1} &= x^k + \alpha_x^k \Delta x^k \\y^{k+1} &= y^k + \Delta y^k \\z^{k+1} &= z^k + \alpha_z^k \Delta z^k\end{aligned}\quad \text{iteration } k$$

until

$$\|F^k\|_\infty = \|F(x^k, y^k, z^k)\|_\infty < \epsilon_{\text{tol}},$$

where

- $\epsilon_{\text{tol}} = 10^{-6}$  is **selected tolerance** and
- the **norm of a vector**  $v$  is defined by  $\|v\|_\infty = \max(|v_i|)$ .

### References

---

- Peng, D.-Y. & Robinson, D. B. (1976). A New Two-Constant Equation of State. *Industrial & Engineering Chemistry Fundamentals*, 15(1), 59–64.
- Redlich, O. & Kwong, J. N. S. (1949). On the Thermodynamics of Solutions. V. An Equation of State. Fugacities of Gaseous Solutions. *Chemical Reviews*, 44(1), 233–244.
- Smith, J. M., Ness, H. C. V., & Abbott, M. M. (2005). *Introduction to chemical engineering thermodynamics*. McGraw-Hill.
- Soave, G. (1972). Equilibrium constants from a modified Redlich-Kwong equation of state. *Chemical Engineering Science*, 27(6), 1197–1203.
- van der Waals, J. D. (1873). *On the Continuity of the Gaseous and Liquid States*. PhD thesis, Universiteit Leiden.