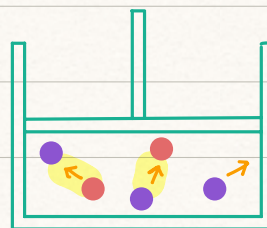
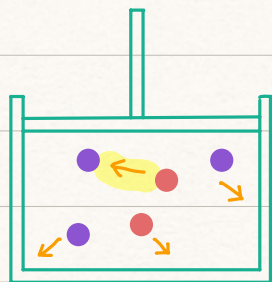


④ Pressure \propto Rate of Reaction [ONLY APPLICABLE FOR GASES]

Pressure \uparrow	Rate of Reaction \uparrow
Pressure \downarrow	Rate of Reaction \downarrow

$P \downarrow$ $V \uparrow$



$P \uparrow$ $V \downarrow$

★ When pressure is increased (for gases only)

- The volume is decreased
- Now there are more particles per unit volume
- More collisions occur per second / frequency of collisions increases
- Greater number of collisions are now successful.
- Hence rate of reaction increases.
- Time taken to complete reaction decreases

★ When pressure is decreased (for gases only)

- The volume is increased
- Now there are less particles per unit volume
- Fewer collisions occur per second / frequency of collisions decreases
- Fewer number of collisions are now successful.
- Hence rate of reaction decreases
- Time taken to complete reaction increases.