



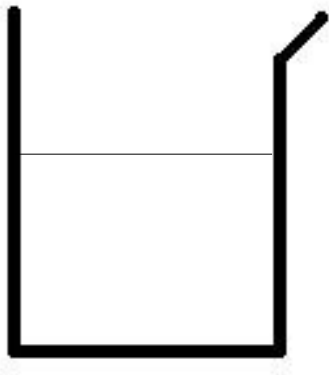
LC.27 Solubilité

Maria Ubero Gonzalez

Du sel dans l'eau

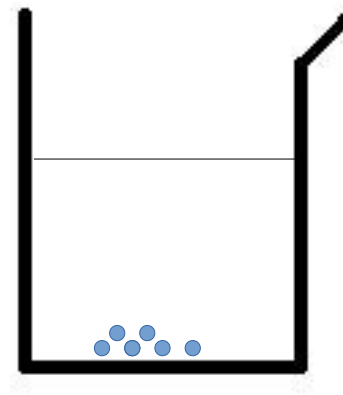
10 mL H₂O

1 g NaCl_(s)

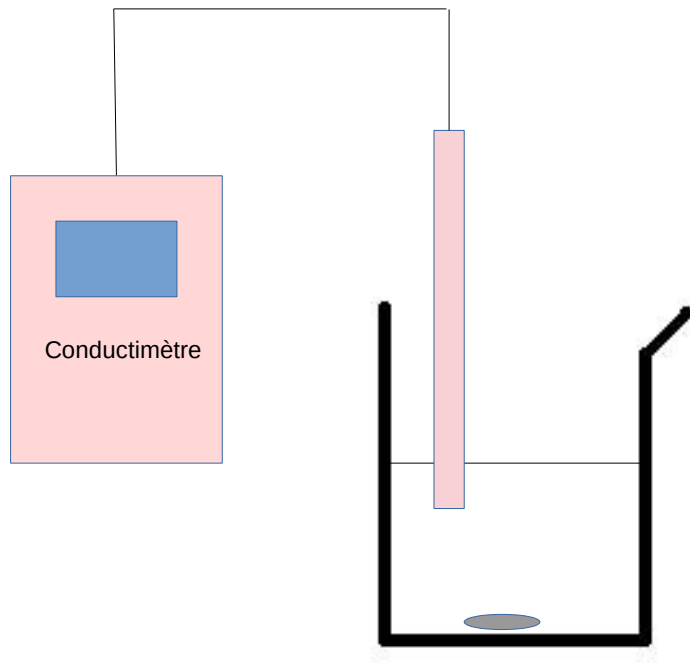


10 mL H₂O

5 g NaCl_(s)

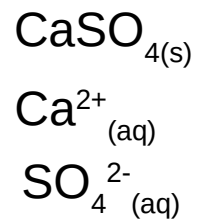


Dilution du sulfate de calcium dans l'eau



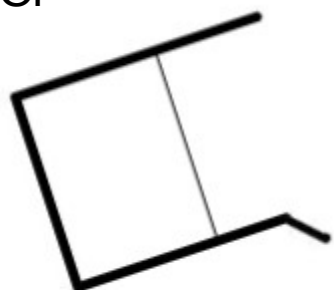
Bécher thermostaté à 25°C
Application de la loi de Kohlrausch

$$\sigma = \lambda^o(\text{Ca}^{2+})[\text{Ca}^{2+}]_{\text{eq}} + \lambda^o(\text{SO}_4^{2-})[\text{SO}_4^{2-}]_{\text{eq}}$$



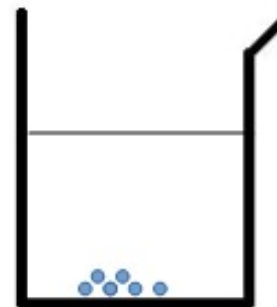
Effet d'ion commun

1 mL de
solution saturée de
NaCl

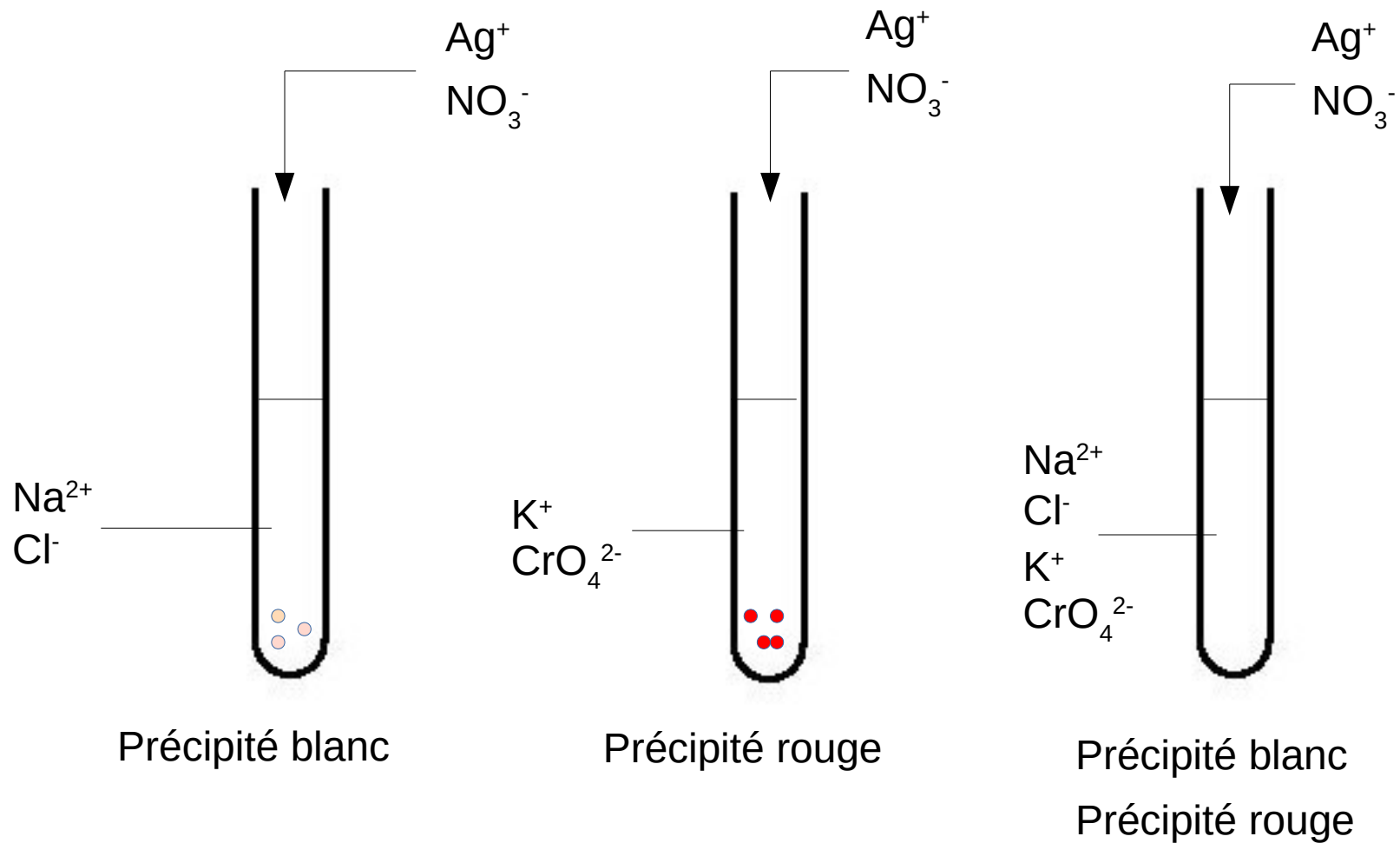


HCl
 $V = 10 \text{ mL}$
 $C = 5 \text{ mol/L}$

On observe un précipité



Méthode de Mohr



Titration des ions chlorure par la méthode de Mohr

