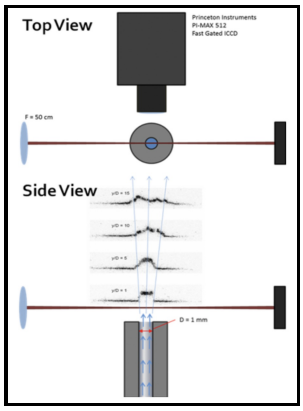
Electronic excitation of nitrogen.

- - Electron



Description: -
-electronic excitation of nitrogen.
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Notes: Thesis (Ph. D.)--The QueensUniversity of Belfast, 1954.
This edition was published in 1954

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Phys. Rev. A 75, 022705 (2007)

This causes a lowering of energy of the ground state and not the excited state. Once we have the molecular orbital energy diagram for benzene, we can assign symmetries to each orbital arrangement of the ground state. They tend to have molar absorptivities on the order of 10,000 and undergo a red shift with solvent interactions a shift to lower energy and longer wavelengths.

Electronic Spectroscopy: Interpretation

The electronic excitation temperature diagnosed using the two-line intensity ratio method of NII 391. So I have a question about the the affect of the electron energy levels of Nitrogen when it becomes diatomic. To do this we look up the Tanabe and Sugano diagrams for Octahedral fields.

Phys. Rev. A 75, 022705 (2007)

The non-polar solvents can interact though polarizability via London interactions also causing a blurring of the vibronic manifold. We express this by modifying the transition moment integral from an integral of eigenstates to an orthogonally expressed direct product of the symmetries of the states.

Phys. Rev. 125, 229 (1962)

When the excited state emerges, the solvent molecules do not have time to rearrange in order to stabilize the excited state. The A 1g to E 1u transition is fully allowed and therefore the most intense peak. Another advantage is the simple experimental implementation.

Electronic Spectroscopy: Interpretation

These transitions arise because of the low-lying energy of the ligand orbitals.

Phys. Rev. A 75, 022705 (2007)

Provide details and share your research! If the transition is allowed, then it should be visible with a large extinction coefficient.

Electron

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Excitation of nitrogen molecular ions in a strong laser field by electron recollisions

Abstract Experimental data from several sources concerning the formation of metastable excited nitrogen molecules on electron impact have been examined critically in order to develop an interpretation that is consistent with all available electron impact information. Three types of transitions are important to consider are Metal to Ligand Charge Transfer MLCT , Ligand to Metal Charge Transfer LMCT , and d-d transitions.

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