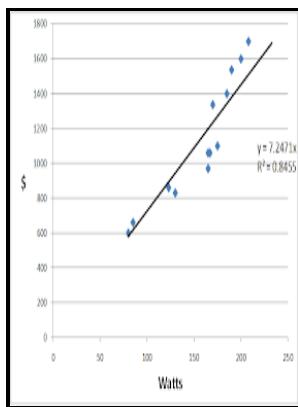


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The is the sum of thermal kinetic energy and thermal potential energy. As a result, there is no change in internal energy; since there is also no change in thermal potential energy, there can be no change in thermal kinetic energy and, therefore, no change in temperature. We will send you an SMS containing a verification code.

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In a Joule—Thomson expansion the enthalpy remains constant.

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The dashed lines demarcate the region where N₂ is a supercritical fluid where properties smoothly transition between liquid-like and gas-like. Remember me on this computer. This equation can only be solved by an iterative trial and error procedure using the tables of for all four components of the products - CO₂, H₂O, O₂, and N₂.

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The blue curves are isenthalps curves of constant specific enthalpy.

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