

# Chemistry of cement and concrete.

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## Notation in cement chemistry

Cement and water form a paste that coats each particle of stone and sand—the aggregates. They exhibit good resistance to aggressive agents, including sulfate. If possible, 1 kg of cement, sand, and the aggregate mix should be packed independently for the chemical tests.

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Hydration of C 3S and C 3A also contribute to this peak. The active ingredients are monocalcium aluminate  $\text{CaAl}_2\text{O}_4$   $\text{CaO} \cdot \text{Al}_2\text{O}_3$  or CA in , CCN and  $\text{Ca}_{12}\text{Al}_{14}\text{O}_{33}$   $12 \text{ CaO} \cdot 7 \text{ Al}_2\text{O}_3$ , or C 12A 7 in CCN. This accounts for approximately 5% of anthropogenic  $\text{CO}_2$ .

## Basic chemistry

These qualities explain why one material, concrete, can build skyscrapers, bridges, sidewalks and superhighways, houses and dams.

## How Cement Is Made

A minimum temperature of 5 °C is recommended, and no more than 30 °C. Entrained air in many concrete mixes may also take up another 5 to 8 percent. If you find these basic chemistry short notes useful let us know using the Contact Form and we'll expand them.

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