

Qualitative inorganic analysis - a new physico-chemical approach.

Methuen - [PDF] CHEMISTRY WITH INORGANIC QUALITATIVE ANALYSIS SECOND EDITION PDF

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Qualitative inorganic analysis

Most chromatography in inorganic chemistry is on solutions e. The precipitates of these cations are almost indistinguishable, except for , which is yellow.

Characterization of Inorganic Compounds

This should not be tried with azide, perchlorate or fulminate salts because these tend to be explosive. The chlorides of these elements cannot be distinguished from each other by their colour - they are all white solid compounds. Therefore, HCl in concentrations which completely precipitate Hg^{2+} and Ag^{+} may not be sufficient to do the same to Pb^{2+} .

Difference Between Qualitative Analysis and Quantitative Analysis

| Experiment | Observations | Inference |
|------------------|---|--|
| 1. Colour | Blue or bluish green Greenish Light green Dark brown Pink Light pink, flesh colour or earthy colour White | Cu^{2+} or Ni^{2+} Ni^{2+} Fe^{3+} Fe^{3+} Cr^{3+} Mn^{2+} Shows the absence of Co^{2+} , Ni^{2+} , Fe^{3+} , Fe^{2+} , Mn^{2+} , Cr^{3+} |
| 2. Smell | Ammoniacal smell Vinylic like smell Smell like that of rotten eggs | NH_4^+ CH_3COO^- S^{2-} |
| 3. Density | (i) Heavy (ii) Light fluffy powder | Salt of Pb^{2+} or Ba^{2+} Carbonate |
| 4. Deliquescence | Salt absorbs moisture and becomes paste like | (i) If coloured, may be $\text{Cu}(\text{NO}_3)_2$, FeCl_3 (ii) If colourless, may be $\text{Zn}(\text{NO}_3)_2$, chlorides of Zn^{2+} , Mg^{2+} etc. |

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The reagent usually used is at around 0.

Qualitative Inorganic Analysis : A brief look into the Why's and How's !

For complex compounds involving large organic moieties, the IR becomes more difficult to interpret. CuS dissolves in ammonia forming an intense blue solution, while CdS dissolves forming a colourless solution. Once again, the precipitates are collected by filtration or centrifugation.

17.7: Qualitative Analysis for Metallic Elements

In this experiment a sample is irradiated with microwave radiation and the field is swept until resonance occurs. The test with the sulfide ion must be conducted in the presence of dilute HCl.

Qualitative Inorganic Analysis : A brief look into the Why's and How's !

Chlorides are confirmed by the test.



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