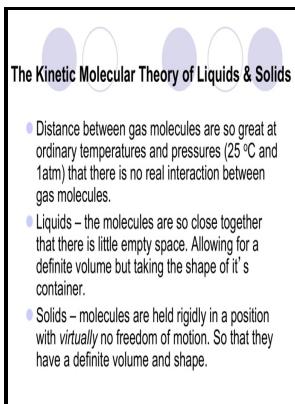


Molecular theory of gases and liquids

Wiley - Molecular theory of gases and liquids. Joseph O. Hirschfelder, Charles F. Curtiss, and R. Byron Bird. John Wiley & Sons, Inc., New York (1954). 1219 pages. \$20, Aiche Journal



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11.2: Solids, Liquids, and Gases

Only by understanding the nature and strength of these forces can we understand how the composition and structure of a substance are related to its physical properties in the liquid or solid state. As the temperature of a gas decreases, the average kinetic energy of its particles decreases, allowing the attractions between the particles to first draw the particles close together, forming a liquid, and then to virtually lock them in place, forming a solid.

The Molecular Theory of Gases and Liquids

Appropriate for graduate students and working scientists alike, this book details advanced concepts without sacrificing depth of coverage or technical detail.

Molecular theory of gases and liquids : Hirschfelder, Joseph Oakland, 1911

What forces cause the lotus leaf to repel water so efficiently? Over the next century, the kinetic theory became more widely adopted among scientists, as part of a growing trend toward scientists adopting the modern view of matter as composed of atoms. Moreover, through this experiment we noticed that the stronger the hydrophobic interactions between two compounds are, it releases more energy, is less ordered, and is more spontaneous; more non-polar atoms had a stronger hydrophobic interaction with the non-polar surface of the gas column Table 7.

Molecular Theory Of Gases And Liquids by Joseph O. Hirschfelder

On the other hand, high melting points are when bonds are very strong such as salts with ionic bonds. Unless otherwise noted, LibreTexts content is licensed by. Atomization and Spray Drying by W.

Kinetic Molecular Theory of Gases

We will study solids in Chapter 12. Clicking on the donut icon will load a page at altmetric.

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