Thermodynamics.

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Energy, Enthalpy, and the First Law of Thermodynamics

One of the simplest is the Clausius statement, that heat does not spontaneously pass from a colder to a hotter body. Three numbers describe 10 26 or 10 30 nominal independent variables.

The Four Laws of Thermodynamics

But if the temperatures were equal, no heat would have flowed in the first place.

Flameless Heater

So you see, this simple gas of noninteracting particles actually has an amazing ability: it can convert thermal energy into mechanical work, and vice versa! It also springs up in chemistry too. What would happen if we created a set of conditions under which no work is done by the system on its surroundings, or vice versa, during a chemical reaction? Basic Concepts of Heat Transfer Broadly speaking, the heat of a material is understood as a representation of the energy contained within the particles of that material. The states that in a natural , the sum of the of the interacting never decreases.

The Four Laws of Thermodynamics

Entropy in any closed system always increases; it never decreases. If heat flows into a system or the surroundings do work on it, the internal energy increases and the sign of q and w are positive. Boltzmann himself spent years trying to prove the second law, and several times thought he had succeeded.

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