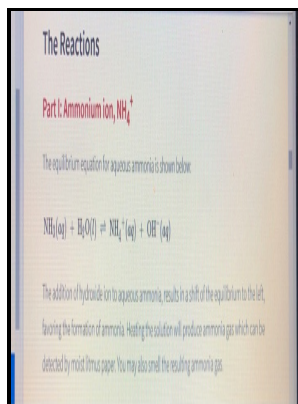


# Qualitative analysis with ionic equilibrium

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Description: -

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Ionic equilibrium

Chemistry, Analytic -- Qualitative. Qualitative analysis with ionic equilibrium

-Qualitative analysis with ionic equilibrium

Notes: Includes bibliographical references and index.

This edition was published in 1991



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Tags: #Qualitative #Analysis #and #Chemical #Equilibrium #(Hogness, #T. #R.; #Johnson, #Warren #C.)

### Wisner, Qualitative Analysis with Ionic Equilibrium

The solution can be filtered, and the filtrate contains  $\text{Ca}^{2+}$  and  $\text{Mg}^{2+}$ . Savvas Learning Company is a trademark of Savvas Learning Company LLC. Consecutive precipitation steps become progressively less selective until almost all the metal ions are precipitated.

### Wisner, Qualitative Analysis with Ionic Equilibrium

Washing When a supernatant is removed from a solid, it is almost certain that some of the liquid has been left behind. The fact that your unknown is insoluble in water may give you an idea as to its identity.

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Complete the entire flow chart to make sure that no other species appear.

### Qualitative Analysis

The various analytical procedures you will do must be performed on ions dissolved in a solution. The pH of a solution is the numerical value of the negative power to which ten must be raised to express the  $\text{H}^+$  ion concentration. Throughout this series of experiments, you will be expected to follow the directions that are presented in the flow charts included in this lab.

### Wisner, Qualitative Analysis with Ionic Equilibrium

For instance, group IV cations, which consist of insoluble carbonates and phosphates such as  $\text{Ba}^{2+}$ ,  $\text{Ca}^{2+}$ , and  $\text{Mg}^{2+}$ , all form white precipitates in the presence of diammonium hydrogen phosphate  $\text{NH}_4\text{H}_2\text{PO}_4$  in a basic solution. Be sure not to add more of a precipitating agent than indicated in the procedure as this may cause undesired side reactions. Each procedure should be performed on approximately 0.

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