

Electrochemical determination of Gibbs energies of formation of cobalt and nickel sulfides

U.S. Dept. of the Interior, Bureau of Mines - Gibbs' free energy (Concept)

Description: -

- Epistolary fiction, English -- History and criticism.

Richardson, Samuel, 1689-1761 -- Criticism and interpretation.

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Nickel sulphide.

Cobalt sulphide.

Gibbs free energy.

Electrochemistry. Electrochemical determination of Gibbs energies of formation of cobalt and nickel sulfides

- 8588.

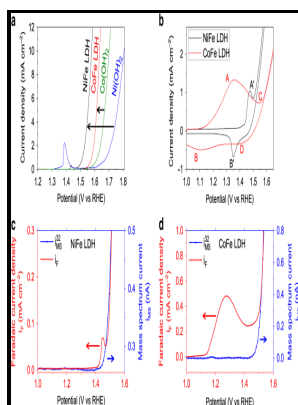
Report of investigations (United States. Bureau of Mines) ;

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Report of investigations / United States Department of the Interior, Bureau of Mines ; Electrochemical determination of Gibbs energies of formation of cobalt and nickel sulfides

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Standard Gibbs free energy of formation for Cu₂O, NiO, CoO, and Fe₃O₄: High resolution electrochemical measurements using zirconia solid electrolytes from 900

Galvanic cells with oxygen-specific solid electrolytes made of calcia-stabilized zirconia have been used to make equilibrium measurements of the standard Gibbs free energy of formation, $\Delta_f G^\circ_m(T)$, for copper I oxide Cu₂O, nickel II oxide NiO, cobalt II oxide CoO, and wüstite Fe₃O₄ over the temperature range from 900—1400 K. Measured equilibrium molal concentrations of H₂ aq at 256, 274, 300, 324 and 355 °C are 0. These values of $\Delta_f S^\circ_m(298)$.

Standard Gibbs free energy of formation for Cu₂O, NiO, CoO, and Fe₃O₄: High resolution electrochemical measurements using zirconia solid electrolytes from 900

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The standard parameters of sphalerite were analyzed by the second law of thermodynamics.

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Division of Solar Energy, Westinghouse Electric Corporation. Borie, Oak Ridge National Laboratory, and U.

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