

# Atomic and free radical reactions.

## Reinhold - 5.5. The Free

A second option for Cl-Cl bond breakage: "homolytic" cleavage

Note the "single-barbed" curved arrows, each depicting the movement of 1 electron

"Free radicals"

Note that each Cl ends up with one electron from the Cl-Cl bond

The resulting species are called "free radicals"

Description: -

- Atomic and free radical reactions.
- Atomic and free radical reactions.

Notes: In 2 volumes.

This edition was published in 1954



Filesize: 51.95 MB

Tags: #Free

## Atomic and Free

These enzymes, like catalase, degrade hydrogen peroxide. However, some physical properties of BHT and BHA such as their high volatility and instability at elevated temperature, strict legislation on the use of synthetic food additives, carcinogenic nature of some synthetic antioxidants, and consumer preferences have shifted the attention of manufacturers from synthetic to natural antioxidants.

## Free Radicals and Reactive Oxygen

Although the precise mechanism and site of radical formation in vivo are not well elucidated yet, the metal-induced decompositions of hydroperoxides and hydrogen peroxide must be one of the important sources. Advances in Heterocyclic Chemistry 62. Synthetic antioxidants such as butylated hydroxytoluene and butylated hydroxyanisole have recently been reported to be dangerous for human health.

## 5.5. The Free

Induction of mutagenesis, the best known of the biological effect of radiation, occurs mainly through damage of DNA by the HO.

## Unit 5: Radicals and Radical Reactions

The presence of an unpaired electron results in certain common properties that are shared by most radicals.

## Related Books

- [Parliament.](#)
- [Teaching research methods in the social sciences](#)
- [Joy diet - 10 steps to a happier life](#)
- [Architecture - an introductory reader](#)
- [Future of natural fibres - papers presented at a Shirley Institute Conference on 29-30 November 1977](#)