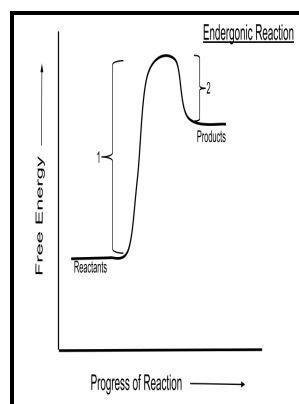


# Energy changes in biochemical reactions

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## Energy Changes in Chemical Reactions

The energy change in a chemical reaction is due to the difference in the amounts of stored chemical energy between the products and the reactants. These NADHs are used instead, their electrons are donated to our friend pyruvate here, these three carbon sugar. We invested two but we harvested six.

## Types of Biochemical Reactions

Learners first draw these curves without knowing about activation energy and then towards the end of the chapter they start adding the activation energy. It is hoped that a reader who has conscientiously worked his way through this volume will acquire not only a cocktail-party knowledge of thermodynamics but will be able to apply it to some simple biochemical or chemical reactions. Now, what do enzymes do? Point X on the diagram is at the lowest energy.

## Chemical reaction

And, therefore, if I were to supply these reactants with energy, for instance, let's say I were to heat up these reactants and therefore give them a higher degree of thermal energy which they might be able to use to move up to this high energy state. To activate a lightstick, you bend the plastic stick, which breaks the glass vial. Cardboard cups will also work fine.

## Biochemical Reactions, Energy, and Enzymes

The difference in the two cases depends on restricting the flow of gasoline.

## Types of Biochemical Reactions

Energy is absorbed to break bonds, and energy is evolved as bonds are made.

## Free energy changes in biochemical systems

The higher the activation enthalpy, the more energy is required for the products to form. And that makes it possible for their equilibrium to be much more favored.

### **Gibbs Free Energy and the Spontaneity of Chemical Reactions**

Evaporation is an endothermic process while condensation is an exothermic process.

### **Chemical reaction**

You may recall this symbol from physics.

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