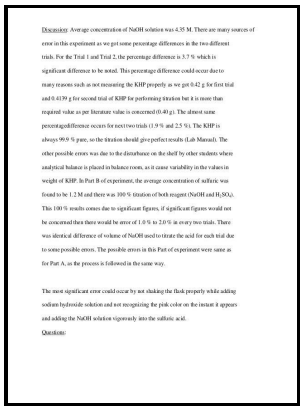
Chemical equilibrium and analysis

Addison-Wesley Pub. Co. - 11: Chemical Equilibrium



Discussion: Average concentration of NaOH solution was 4.35 M. There are many sources of error in this experiment as we got some percentage differences in the two different trials. For the Trial 1 and Trial 2, the percentage difference is 3.7 % which is significant difference to be noted. This percentage difference could occur due to many reasons such as not measuring the KHP properly as we got 0.42 g for first trial and 0.4139 g for second trial of KHP for performing titration but it is more than required value as per literature value is concerned (0.40 g). The almost same percentagedifference occurs for next two trials (1.9 % and 2.5 %). The KHP is always 99.9 % pure, so the titration should give perfect results (Lab Manual). The other possible errors was due to the disturbance on the shelf by other students where analytical balance is placed in balance room, as it cause variability in the values in weight of KHP. In Part B of experiment, the average concentration of sulfuric was found to be 1.2 M and there was 100 % titration of both reagent (NaOH and H_2SO_4). This 100 % results comes due to significant figures, if significant figures would not be concerned then there would be error of 1.0 % to 2.0 % in every two trials. There was identical difference of volume of NaOH used to titrate the acid for each trial due to some possible errors. The possible errors in this Part of experiment were same as for Part A, as the process is followed in the same way.

The most significant error could occur by not shaking the flask properly while adding sodium hydroxide solution and not recognizing the pink color on the instant it appears and adding the NaOH solution vigorously into the sulfuric acid.

Questions:

Description: -
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Chemical equilibrium.
Chemistry, Analytic.Chemical equilibrium and analysis
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Addison-Wesley series in chemistryChemical equilibrium and analysis
Notes: Includes index.
This edition was published in 1981

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Tags: #Sections #in #this #Course

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Chemical equilibrium

The LibreTexts libraries are and are supported by the Department of Education Open Textbook Pilot Project, the UC Davis Office of the Provost, the UC Davis Library, the California State University Affordable Learning Solutions Program, and Merlot. Any fully detailed theory of reaction rates should start at the molecular level, and derive an expression for k T by performing correctly weighted summations or integrations over the variables that characterize the molecular events.

Chemical equilibrium analysis of energetic materials using Particle Swarm Optimization

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Therefore the anion part of the salt will hydrolyze as follows.

Chemical reaction equilibrium analysis (1982 edition)

Sometimes the analyte or titrant will serve this function auto indicating. In this reaction, Cl - is neither added to the flask nor is it consumed in the reaction in the course of the titration process. For any equilibrium reaction, the ratio of concentrations of the substances on the right to the concentrations of those on the left equals a constant appropriate for that specific reaction.

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