

Aromatic nitration

Cambridge University Press - Nitration

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Steiner, Rudolf, 1861-1925.

Bolívar, Simón, 1783-1830.

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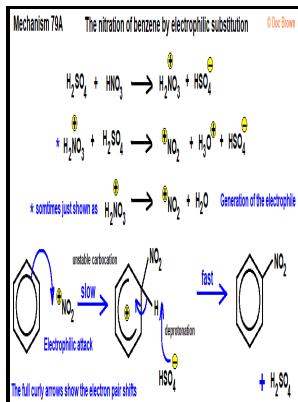
Aromatic compounds.

Nitration.Aromatic nitration

-Aromatic nitration

Notes: Includes bibliographical references and indexes.

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Tags: #Experiment #11: #Electrophilic

#Aromatic #Substitution

18.4 Aromatic Nitration and Sulfonation

The second most shielded peak, peak B, was a triplet, located at 7. So a lone pair of electrons on our water molecule are going to take that proton, which would cause these electrons to move in here to reform your aromatic ring.

Reactions of Aromatic Compounds — Organic Chemistry Tutor

This is going to be the electrophile in our mechanism. From this we may conclude that the nitration reagent is more reactive and less selective than the halogenation reagents. Employing increased reaction temperatures the authors observed an increase in the reactions selectivity toward the synthesis of 5-nitro-salicylic acid 102, obtaining 65% at 50 °C.

Nitration

So this oxygen is still going to have a negative 1 formal charge on it. The aromaticity of the aromatic system is preserved in an electrophilic aromatic substitution reaction. If the dinitrated aromatic or substituted aromatic composition is desired, then stoichiometrically four moles soluble anhydrite are required to adsorb the two moles water generated by the dinitration reaction.

nitration of benzene and methylbenzene

On the other hand, in the resonance structures have negative charges around the ring system see the picture below : The can donate electron density through Attack occurs at the ortho and para positions, which have partial formal negative charges that indicate π electron excess at these positions.

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