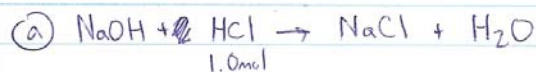


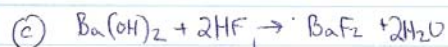
24. ~~24~~



1:1 ratio

1.0 mol HCl $\boxed{1 \text{ mol NaOH}}$

(b) You Try (answer .75 mol HNO_3)



.20 mol

.20 mol HF

$$\frac{1 \text{ mol Ba(OH)}_2}{2 \text{ mol HF}} = \boxed{.10 \text{ mol Ba(OH)}_2}$$

(d) You Try (answer 0.90 mol H_2SO_4)

24.

25



15.0 mL
 $2.5 \times 10^{-2} \text{ M}$

10.0 mL
X M

25

(1) Find moles of H_2SO_4

$$M = \frac{\text{mol}}{\text{L}}$$

$$2.5 \times 10^{-2} = \frac{x}{.015 \text{ L}}$$

25

24.

$$.00075 \text{ mol} = x \text{ mol H}_2\text{SO}_4 \times \frac{2 \text{ mol KOH}}{1 \text{ mol H}_2\text{SO}_4} = .00075 \text{ mol KOH}$$

(2) moles of KOH

$$(3) \text{ Molarity KOH} = \frac{\text{moles KOH}}{\text{L}} = \frac{.00075 \text{ mol}}{.010 \text{ L}} = \boxed{.075 \text{ M KOH}}$$

26

Try yourself answer $3.02 \times 10^{-2} \text{ M HNO}_3$