

Writing Compound Formulas Worksheet

Write the correct formula for each of the following ionic compound names.

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| 1. Copper (I) phosphate | <u>Cu₃PO₄</u> | 16. Iron (III) sulfate | <u>Fe₂(SO₄)₃</u> |
| 2. Magnesium nitrite | <u>Mg₂(NO₃)₂</u> | 17. Silver (I) dichromate | <u>Ag₂Cr₂O₇</u> |
| 3. Cobalt (II) hydroxide | <u>Co(OH)₂</u> | 18. Lead (II) nitride | <u>Pb₃N₂</u> |
| 4. Iron (III) chromate | <u>Fe₂(CrO₄)₃</u> | 19. Strontium hypochlorite | <u>Sr(ClO)₂</u> |
| 5. Chromium (II) chloride | <u>CrCl₂</u> | 20. Barium chlorate | <u>Ba(ClO₃)₂</u> |
| 6. Zinc (II) fluoride | <u>ZnF₂</u> | 21. Tin (IV) oxide | <u>TiO₂</u> |
| 7. Mercury (II) acetate | <u>Hg₂(C₂H₃O₂)₂</u> | 22. Nickel (II) cyanide | <u>Ni(CN)₂</u> |
| 8. Manganese (II) iodide | <u>MnI₂</u> | 23. Aluminum sulfide | <u>Al₂S₃</u> |
| 9. Copper (II) nitrate | <u>Cu(NO₃)₂</u> | 24. Lead (IV) chromate | <u>Pb(CrO₄)₂</u> |
| 10. Aluminum hydroxide | <u>Al(OH)₃</u> | 25. Potassium permanganate | <u>KMnO₄</u> |
| 11. Iron (II) phosphate | <u>Fe₂(PO₄)₂</u> | 26. Calcium dihydrogen phosphate | <u>Ca(H₂PO₄)₂</u> |
| 12. Magnesium sulfite | <u>MgSO₃</u> | 27. Barium hydrogen sulfate | <u>Ba(HSO₄)₂</u> |
| 13. Tin (II) bromide | <u>TiBr₂</u> | 28. Lithium monohydrogen phosphate | <u>Li₂HPO₄</u> |
| 14. Sodium bicarbonate | <u>NaHCO₃</u> | 29. Silver (I) bisulfite | <u>AgHSO₃</u> |
| 15. Chromium (III) oxalate | <u>Cr₂(C₂O₄)₃</u> | 30. Ammonium carbonate | <u>(NH₄)₂CO₃</u> |

Write the Compound formulas for the covalent molecules below:

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|----------------------------|------------------------------------|--------------------------------|------------------------------------|
| 31. Dinitrogen oxide | <u>N₂O</u> | 35. Tetraphosphorus trisulfide | <u>P₄S₃</u> |
| 32. Carbon tetrachloride | <u>CI₄</u> | 36. Sulfur hexafluoride | <u>SF₆</u> |
| 33. Dioxygen heptachloride | <u>O₂Cl₇</u> | 37. Dinitrogen pentoxide | <u>N₂O₅</u> |
| 34. Phosphorus tribromide | <u>PBr₃</u> | 38. Tetrasulfur decanitride | <u>S₄N₁₀</u> |

Write the compound name for the following covalent molecules.

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|------------------------------------|-------------------------------------|------------------------------------|-------------------------------|
| 39. CS ₂ | <u>Carbon disulfide</u> | 43. N ₂ S ₅ | <u>dinitrogen pentasulfid</u> |
| 40. PF ₅ | <u>phosphorous pentafluoride</u> | 44. S ₂ F ₁₀ | <u>disulfur decafluoride</u> |
| 41. CBr ₄ | <u>carbon tetrabromide</u> | 45. Cl ₂ O ₇ | <u>dichlorine heptoxide</u> |
| 42. P ₄ S ₁₀ | <u>tetraphosphorous decasulfide</u> | 46. I ₂ O ₅ | <u>diiodine pentoaxide</u> |