

06/07/2018

2-B1 $VOLUME = 50 \text{ m}^3 = 50,000 \text{ cm}^3$ Helium density = $0.1785 \times 10^{-3} \text{ g/cm}^3$

DENSITY = $\frac{MASS}{VOLUME} \rightarrow MASS = \text{density} \times \text{volume}$

$m = 5 \times 10^7 \text{ cm}^3 (0.1785 \times 10^{-3} \text{ g/cm}^3) = 8,925 \times 10^3 \text{ g}$

$He(\text{molar_mass}) = 4.0026 \rightarrow \frac{4.0026}{N_A} = 6,64607721 \times 10^{-24}$

$\frac{8,925 \times 10^3 \text{ g}}{6,64607721 \times 10^{-24}} = 1,34289243 \times 10^{27} \text{ atoms}$

2-B2 WATER IN OCEANS OF EARTH (ESTIMATE) = $10^{21} \text{ kg} = 1 \times 10^{24} \text{ g}$

$O = 15.9994 / N_A$ $H = 1.00797 / N_A$ $H_2O = 2(H) + (O) = 3,9913914 \times 10^{-23} \text{ g}$

$= 2,656644 \times 10^{-23} \text{ g} = 1,67367372 \times 10^{-24} \text{ g} = 2,99133914 \times 10^{-23} \text{ g}$

$DYE = 2 \text{ g} \rightarrow \frac{2 \text{ g}}{N_A} = 3,32082003 \times 10^{-22} \text{ g} \text{ (or } 10^{-24})$

$MASS_OCEANS / H_2O = 3,74298437 \times 10^{27} \text{ molecules of } H_2O \text{ in the OCEANS}$

$MASS_DYE / DYE = 6,0225000 \times 10^{21} \text{ molecules of DYE (or } 10^{23})$

$\frac{\text{molecules_dye}}{\text{molecules_oceans}} = 1,8057400 \times 10^{-28} \text{ dye/H}_2\text{O (or } 10^{-26})$

$300 \text{ g/H}_2\text{O} = 1,20347437 \times 10^{25} \text{ atoms} \times \text{dye/H}_2\text{O} = 2,1681 \times 10^{-3} = \frac{1}{500} \text{ (or } 2,1681 \times 10^{-1} = \frac{1}{50})$

2-B3 $\frac{1}{2} H = 1.008 \text{ g/mole} = 1,67372354 \times 10^{-24} \text{ g}$

$1,008 \text{ g} / 6,0225 \times 10^{23} = \frac{1}{500}$

2-B4 $^{239}_{94}\text{Pu}$ $239 / N_A = 3,96845164 \times 10^{-22} \text{ g/atom}$

1 Mol = 1 Million Atoms = $3,9684574 \times 10^{-16}$

2-B5 10L H_2O (water) SPECIFIC GRAVITY = 1.00 g/cm^3 $\frac{1.00}{1.26} = 0,7937$

5L Glycerin SPECIFIC GRAVITY = 1.26 g/cm^3 $\frac{1.26}{1.00} = 1,26$

$\frac{39,6825397}{100} = 0,396825397$ $\frac{1,26 \text{ g/cm}^3}{1,00 \text{ g/cm}^3} = 1,26$ $\frac{1,26}{0,396825397} = 3,175$

2-B6 $^{10}_5\text{B}$ 19% of t AVG WEIGHT of Boron = 10,81 V

$^{11}_5\text{B}$ 81% of t $(10 \times 0,19) + (11 \times 0,81) = 10,81$

2-B7 1000 cm^3 Liquid CARBON (CCl_4) $\Delta MASS = 1375 \text{ g}$

$1,595 \text{ g/cm}^3$ $(\text{C}/N_A)(\text{Cl}/N_A) \times 4 = 6,2047662 \times 10^{-24}$

2-B8 $A = 26,9815$ $4,48011623 \times 10^{-23} \text{ g}$

2-C1 $F = F_1 + F_2 = \frac{10^{-10}}{x^2} + \frac{10^{-80}}{x^2}$ $x = \text{distance in meters}$

a) calculate F_1 , F_2 and the total force or F for the following values:

$x = 10^{-8} \text{ m}, 10^{-9} \text{ m}, 10^{-10} \text{ m}, 10^{-11} \text{ m}, 10^{-12} \text{ m}$

$F = \frac{10^{-10}}{(10^{-8})^2} + \frac{10^{-80}}{(10^{-8})^2}$ $F = 1 \times 10^{-6}$ $F = \frac{10^{-10} + 10^{-80}}{(10^{-8})^2}$ $F = 1 \times 10^{-1}$

$F = 1 \times 10^{-6} + 1 \times 10^{-8}$ $F = 1 \times 10^{-1} + 1,1 \times 10^{-80}$

06/07/2018 (SOLFITO) - 2

$9,05 \div 6,146 = 1,472$ $1,36 \times 10^{27}$

06/08/2018 + Canada

Lynne be reviewed

06/08/2018

10/09/2018

2AS

2AS

1,267

24/09/2018 2AS

01/10/2018


20/10/18