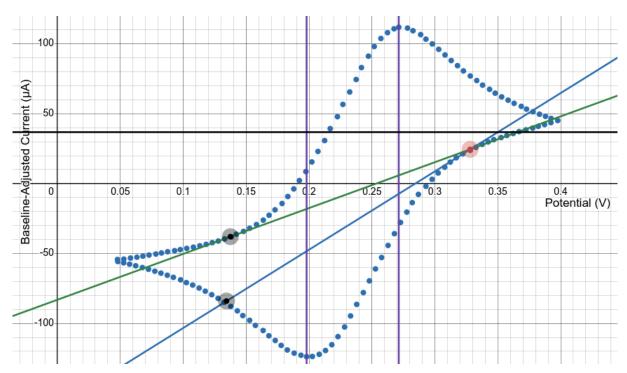
Lab 8: Chronoamperometry and Cyclic Voltammetry

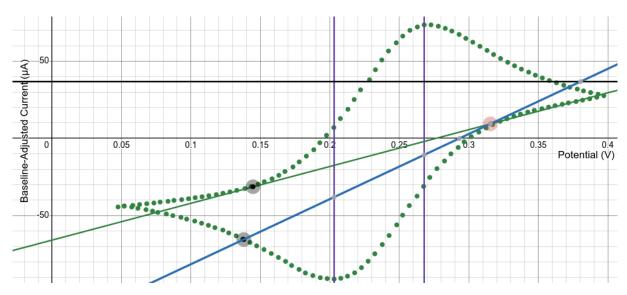
Questions

Part A

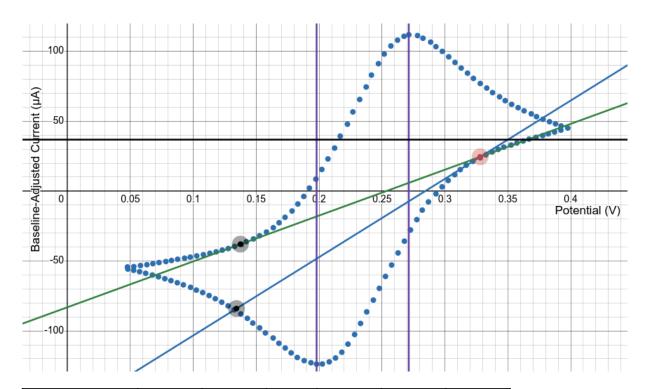
Scan Rate 100 mV/s



Scan Rate 50 mV/s



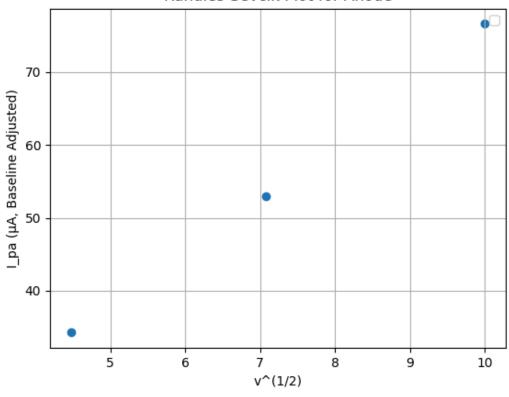
Scan Rate 20 mV/s

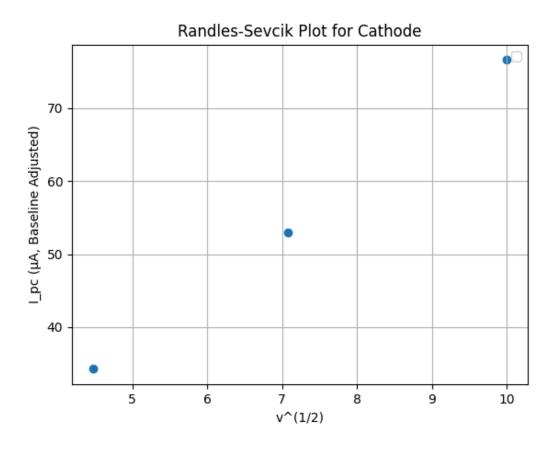


Scan rate, mV/s	$\mathbf{E}_{\mathbf{p_a}}$	E _{pc}	$\mathbf{E}_{\mathbf{p}/2}$	E_{pa} - $E_{p/2}$	I _{pa} (μA)	I _{pc} (μ A)
100	0.271282	0.198041	0.2301	0.041182	118.1728	76.59921
50	0.267902	0.203031	0.2272	0.040702	84.16323	53.01171
20	0.268224	0.20287	0.2270	0.041264	53.21656	34.27879

1. 50:0.06487, 20:0.06535, 100:0.07324 2. 50:0.62987, 20:0.64414, 100:0.6482 3. 50:0.0407, 20:0.04126, 100:0.04118







5. The data does not suggest that the reaction is reversible due to the peak cathode voltage being substantially larger than the peak anode voltage. The data does appear to diffusion controlled as the Randles-Sevcik plot is linear.

Part B

1

Lab Notebook

Exp. No Experiment/Subject (y(1:(Voltummety,	Desk No.	Course & Section No.
Objective: A perform a form of there were no peats Objective: A perform a form of the for ky polish working electrode Muss furricyonive, make 10mc solution with 100 create electro chemical this solution. Remoted approximate different so an varous 100 272 117 268 77 268 46	eratey to calculate technology to calculate points and a rate of with	Betrochemical cell like in part A. Do chromounpean with setting protocul, those modifica	intervals sin nith
Signature A A A A A A A A A A A A A A A A A A A	Date Witness	orta Wirma.	Date