1. Potassium Ferricyanide has chemical formula $K_3Fe(CN)_6$[1]. Thus, the following computation can be performed:
   1.55 ppm $K_3Fe(CN)_6$ = $\frac{1.55\text{ g }K_3Fe(CN)_6}{10^6 g\text{ solution}} \times \frac{1\text{ g solution}}{1\text{ mL solution}} \times \frac{1000\text{ mL}}{1\text{ L}} \times$

   $\frac{3\times 39.098\text{ g K} + 55.845\text{ g Fe} + 6\times(12.011\text{ g C} + 14.007\text{ g N})}{1\text{ mol }K_3Fe(CN)_6}$

Bibliography

(1) *Ferricyanide*. https://en.wikipedia.org/wiki/Ferricyanide (accessed 2025-01-15)