Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Period \_\_\_\_\_ Due \_\_\_\_\_\_\_\_\_\_\_\_

AP Chemistry Homework, Unit 1: Chemical Reactions, part 1

**In PENCIL**, write the correct formula of each reactant, then write the correct formula for each product, and then balance the equation.

1. Sodium and iodine react together:
2. Sodium and oxygen react together:
3. Sodium and nitrogen react together:
4. Calcium and iodine react together:
5. Calcium and oxygen react together:
6. Calcium and nitrogen react together:
7. Aluminum and iodine react together:
8. Aluminum and oxygen react together:
9. Aluminum and nitrogen react together:
10. Potassium iodide decomposes:
11. Magnesium iodide decomposes:
12. Aluminum iodide decomposes:
13. Nitrogen triiodide decomposes:
14. Propane is formed from its elements:
15. Ammonia is formed from its elements:
16. Sulfur trioxide is formed from its elements:
17. Iron(II) oxide is formed from its elements:
18. Iron(III) oxide is formed from its elements:
19. Sodium carbonate is heated:
20. Calcium carbonate is heated:
21. Carbonic acid decomposes:
22. Sodium oxide reacts with carbon dioxide:
23. Calcium oxide reacts with carbon dioxide:
24. Carbon dioxide is added to water:
25. Sodium hydroxide is heated:
26. Calcium hydroxide is heated:
27. Sodium oxide is added to water:
28. Calcium oxide is added to water:
29. Sodium sulfite is heated:
30. Calcium sulfite is heated:
31. Sulfurous acid is heated:
32. Sodium oxide reacts with sulfur dioxide:
33. Calcium oxide reacts with sulfur dioxide:
34. Sulfur dioxide is added to water:
35. Ammonium chloride is heated:
36. Ammonium phosphate is heated:
37. Ammonium hydroxide is decomposed:
38. Ammonia reacts with hydrogen chloride:
39. Ammonia is added to water:
40. Ammonium carbonate is heated:
41. Sodium chlorate decomposes:
42. Hydrogen peroxide decomposes:
43. Sodium peroxide decomposes:
44. Methane burns in open air:
45. Ethane burns in open air:
46. Octane burns in open air:
47. Ethane burns in limited air:
48. Methane burns in fluorine:
49. Ethane burns in fluorine: