

UNIVERSITY OF PRETORIA

DOCTORAL THESIS

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5           **Fast temperature programmed gas  
chromatography coupled to supercritical  
fluid chromatography (SFCxGC)**

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by

**Daniel MALAN**

10         Submitted in partial fulfilment of the requirements for the  
degree

**Doctor of Philosophy**

In the Faculty of Natural & Agricultural Sciences  
University of Pretoria  
Pretoria

15

30 August 2019



## **Declaration of Authorship**

I, Daniel MALAN, declare that the thesis, which I hereby submit for the degree PhD (Chemistry) at the University of Pretoria, is my own work and has not previously  
20 been submitted by for a degree at this or any other tertiary institution.

SIGNATURE: .....

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## **25      Ethics statement**

The author, whose name appears on the title page of this thesis, has obtained, for the research described in this work, the applicable research ethics approval.

The author declares that he/she has observed the ethical standards required in terms of the University of Pretoria's Code of ethics for researchers and the Policy  
30 guidelines for responsible research.



*"Learning is a peculiar compound of memory, imagination, scientific habit, accurate observation, all concentrated, through a prolonged period, on the analysis of the remains of literature. The result of this sustained mental endeavour is not a book, but a man."*

Mark Pattison



UNIVERSITY OF PRETORIA

## *Abstract*

Natural and Agricultural Sciences  
Department of Chemistry

Doctor of Philosophy

<sup>40</sup> **Fast temperature programmed gas chromatography coupled to supercritical fluid chromatography (SFCxGC)**

by Daniel MALAN

The Thesis Abstract is written here (and usually kept to just this page). The page is kept centered vertically so can expand into the blank space above the title too...



45

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## *Preface*

This thesis is therefore with the following audience in mind:

110 A young South African, who has passed through the South African school system, and has just completed an Honours course at the University of Pretoria, and is now embarking on a postgraduate degree. This student is highly trained in chemistry, in a fragmentary fashion, and has not had exposure to a variety of scientists from other institutions or other disciplines.

115 This thesis therefore starts from a very general base, and attempts to put the technical work done strongly in context of technology, history and society.

I beg my examiners to forgive me this indulgence.



# Contents

<b>Declaration of Authorship</b>	<b>iii</b>
<b>Ethics statement</b>	<b>v</b>
<b>Abstract</b>	<b>ix</b>
<b>Acknowledgements</b>	<b>xi</b>
<b>Preface</b>	<b>xiii</b>
<b>1 Introduction: Climate, fuel, and biodiesel</b>	<b>1</b>
1.1 Energy, fuel and the atmosphere . . . . .	1
1.2 Carbon footprints and carbon neutrality . . . . .	5
1.3 Internal combustion engines. . . . .	7
1.3.1 Scaling and efficiency . . . . .	7
1.3.2 Engine design and thermodynamic cycles. . . . .	8
Otto engine and Otto cycle . . . . .	9
Diesel Cycle . . . . .	10
Gas turbines and the Brayton Cycle . . . . .	11
1.3.3 Noxious pollution from internal combustion engines. . . . .	12
Unburnt fuel . . . . .	12
Contaminants and additives. . . . .	13
Side-reactions . . . . .	13
1.3.4 Mitigating pollution from internal-combustion engines. . . . .	13
Fuel formulation . . . . .	14
Exhaust gas cleanup . . . . .	14
Engine management . . . . .	14
Efficiency implications . . . . .	14
1.3.5 Avoiding pollution from internal-combustion engines . . . . .	15
Electrification . . . . .	15
Carbon-neutral fuels . . . . .	16
1.4 Biofuels . . . . .	16
1.4.1 Bio-gas . . . . .	17
1.4.2 Bio-ethanol . . . . .	17
Brazilian ethanol . . . . .	18
US maize . . . . .	19
1.4.3 Fischer-Tropsch fuel from biomass . . . . .	19
1.4.4 Hydrotreated vegetable oil: “Green diesel” . . . . .	19
1.4.5 Biodiesel . . . . .	19
1.5 The significance of biodiesel and its quality control. . . . .	20
1.6 Conclusion: the role of chromatography. . . . .	21

155	<b>2 Introduction: Carbon dioxide and chromatography</b>	23
	2.1 The chemical industry . . . . .	23
	2.1.1 "Green chemistry" . . . . .	25
	2.1.2 Carbon dioxide as a green chemical . . . . .	26
160	2.2 Extractions using carbon dioxide . . . . .	27
	2.2.1 Commercial extractions . . . . .	27
	Plant oils . . . . .	27
	Hops . . . . .	27
	Coffee . . . . .	27
	2.2.2 Analytical Extractions . . . . .	28
165	2.2.3 Why carbon dioxide? . . . . .	28
	Modifiers . . . . .	30
	Practical extractions . . . . .	30
170	2.3 Supercritical Fluid Chromatography . . . . .	31
	2.3.1 SFC and FID . . . . .	32
	2.3.2 SFC×GC . . . . .	34
175	<b>3 Biodiesel standards</b>	37
	3.1 Introduction to standards . . . . .	37
	3.2 SANS 1935: An overview . . . . .	39
	3.3 SANS 1935: Properties, requirements and methods. . . . .	40
180	3.4 SANS 1935: Physical properties . . . . .	40
	3.4.1 Density . . . . .	40
	3.4.2 Kinematic viscosity . . . . .	41
	3.5 SANS 1935: Aggregate properties: Specialized instrumentation . . . . .	42
	3.5.1 Cetane number . . . . .	42
	3.5.2 Flash point . . . . .	43
	3.5.3 Oxidation stability . . . . .	43
	3.5.4 Cold filter plugging point . . . . .	44
	3.5.5 Copper strip corrosion . . . . .	45
185	3.6 SANS 1935: Chemical properties: Classical determination . . . . .	45
	3.6.1 Sulfated ash . . . . .	45
	3.6.2 Carbon residue . . . . .	45
	3.6.3 Total contamination . . . . .	45
	3.6.4 Acid value . . . . .	46
	3.6.5 Iodine value . . . . .	46
190	3.7 SANS 1935: Chemical properties: Electrochemical determination . . . . .	47
	3.7.1 Water content . . . . .	47
	3.8 SANS 1935: Chemical properties: Spectroscopic determination . . . . .	47
	3.8.1 Group I metals . . . . .	47
	3.8.2 Group II metals . . . . .	47
195	3.8.3 Phosphorus . . . . .	48
	3.8.4 Sulfur . . . . .	48
	3.9 SANS 1935: Chemical properties: Chromatographic determination . . . . .	48
	3.9.1 Methanol content . . . . .	48
	3.9.2 Ester content . . . . .	48
200	3.9.3 Glyceride content . . . . .	49
	3.9.4 Free Glycerol . . . . .	50
	3.9.5 Polyunsaturated methyl esters . . . . .	50
	Linolenic acid methyl ester . . . . .	50
	Highly unsaturated fatty acid methyl esters . . . . .	50

205	3.10 Comparison of SANS 1935 and SANS 342 . . . . .	51
	3.11 Conclusion: room for innovation. . . . .	52
210	<b>4 Instrumentation: Supercritical Fluid Chromatography</b>	53
	4.1 SFC . . . . .	53
	4.2 Mobile phase . . . . .	53
	4.3 Pump . . . . .	53
	4.3.1 Pressure control system . . . . .	55
	4.3.2 Modifier Control . . . . .	56
	4.4 Sample injection . . . . .	57
	4.4.1 Column . . . . .	57
	4.4.2 Stopped-flow . . . . .	59
215	4.5 Pressure Relief . . . . .	59
	4.6 Detector . . . . .	63
220	<b>5 Instrumentation: Fast GC</b>	65
	5.1 Speed of analysis . . . . .	65
	5.2 Fast gas chromatography theory . . . . .	66
	5.3 The General Elution Problem. . . . .	70
	5.4 Temperature ramp rates . . . . .	71
	5.4.1 Resistive heating . . . . .	72
	5.5 Calibration . . . . .	75
225	5.5.1 Temperature uniformity . . . . .	75
	Resistivity's dependence on temperature . . . . .	75
	Thermal conduction . . . . .	75
	Radiation . . . . .	75
	Convection . . . . .	75
230	Examining thermal uniformity by imaging . . . . .	76
	5.5.2 Temperature calibration . . . . .	76
	5.5.3 Thermocouple Welding . . . . .	79
	5.5.4 Thermocouple probe construction . . . . .	81
	5.5.5 Thermocouple interfacing . . . . .	81
235	5.5.6 Calibration procedure . . . . .	81
	5.5.7 Cold spots . . . . .	84
	5.5.8 Cold column . . . . .	85
	Cryogen supply . . . . .	86
	5.5.9 Column mounting . . . . .	86
240	5.5.10 Heating control . . . . .	87
	Temperature monitoring . . . . .	87
	PID tuning . . . . .	87
	Heating rates . . . . .	88
	Cooling rates . . . . .	88
245	5.6 Detector . . . . .	89
	5.7 Data acquisition and control software . . . . .	89
	5.8 Data structure . . . . .	89
	5.9 Data visualization . . . . .	90

<b>6</b>	<b>Investigating biodiesel feedstock using SFC<math>\times</math>GC</b>	<b>93</b>
250	6.1 Introduction . . . . .	93
	6.2 SFC of FAMEs . . . . .	93
	6.3 Coaxial heater performance. . . . .	94
	6.3.1 Column lifetime . . . . .	94
	Mechanical degradation . . . . .	94
255	Stationary phase degradation . . . . .	96
	Column bleed . . . . .	96
	6.3.2 Thermal shock . . . . .	96
	6.3.3 Thermal fatigue . . . . .	97
	6.3.4 Corrosion . . . . .	97
260	6.3.5 GC retention time precision . . . . .	97
	6.4 Study of biodiesel by SFC $\times$ GC . . . . .	98
	6.4.1 Samples . . . . .	98
	6.4.2 Sample preparation . . . . .	98
	Method . . . . .	98
265	6.4.3 SFC . . . . .	99
	6.4.4 Modulation . . . . .	99
	6.4.5 GC . . . . .	99
	6.4.6 Results and discussion . . . . .	99
	6.4.7 Sunflower oil . . . . .	101
270	6.4.8 Coconut oil . . . . .	102
	6.4.9 Flax seed oil . . . . .	103
	6.5 The study of biodiesel blends by SFC $\times$ GC . . . . .	105
	6.5.1 Introduction . . . . .	105
	6.5.2 Sample . . . . .	105
275	6.5.3 SFC . . . . .	105
	6.5.4 Modulation . . . . .	105
	6.5.5 GC . . . . .	105
	6.5.6 Results and discussion . . . . .	106
	6.6 Conclusion . . . . .	108
280	<b>7 Application: Modified SFC</b>	<b>111</b>
	7.1 Introduction . . . . .	111
	7.2 SFC $\times$ GC with modifier . . . . .	112
	7.2.1 Sample . . . . .	112
	7.2.2 SFC . . . . .	113
285	7.2.3 Modulation . . . . .	113
	7.2.4 GC . . . . .	113
	7.2.5 Results and discussion . . . . .	113
	7.2.6 Conclusion . . . . .	115
	<b>8 Conclusion</b>	<b>117</b>
290	8.1 Introduction . . . . .	117
	8.2 Synopsis . . . . .	117
	8.2.1 Chapter 1: Introduction: Climate, fuel, and biodiesel . . . . .	117
	8.2.2 Chapter 2: Introduction: Carbon dioxide and chromatography	117
	8.2.3 Chapter 3: Biodiesel standards . . . . .	118
295	8.2.4 Chapter 4: Instrumentation: Supercritical Fluid Chromatography . . . . .	118
	8.2.5 Chapter 5: Instrumentation: Fast GC . . . . .	118

	8.2.6 Chapter 6: Investigating biodiesel feedstock using SFC $\times$ GC . . . . .	118
	8.2.7 Chapter 7: Application: Modified SFC . . . . .	118
300	8.3 Contributions of this study . . . . .	118
	8.3.1 Coaxial heater cooling . . . . .	119
	8.3.2 Coaxial heater reliability . . . . .	119
	8.3.3 Length of SFC column . . . . .	119
	8.4 Voids in this study . . . . .	119
305	8.4.1 Lack of optimization . . . . .	119
	8.4.2 Luck with voltage ranges . . . . .	119
	8.5 Design . . . . .	119
	8.6 Disasters . . . . .	119
	8.6.1 About the isolation amplifier . . . . .	119
310	8.6.2 Thinking that we'd elute alcohols . . . . .	119
	8.6.3 Variable modulation period . . . . .	119
	8.7 Limitations of this SFC $\times$ GC design . . . . .	119
	8.7.1 Cooling control . . . . .	119
	8.8 Suggested design improvements . . . . .	120
315	8.8.1 Four-wire resistance measurement . . . . .	120
	8.8.2 Legs heating and cooling integrated in detector and inlet . . . . .	120
	8.8.3 Siltek-treated coaxial heater . . . . .	120
	<b>A Frequently Asked Questions</b> . . . . .	<b>123</b>
	A.1 How do I change the colors of links? . . . . .	123
320	<b>B Glossary</b> . . . . .	<b>125</b>
	<b>C Thermocouple Production</b> . . . . .	<b>127</b>
	<b>D Engineering Drawings</b> . . . . .	<b>129</b>
	<b>E Fatty acid profiles</b> . . . . .	<b>137</b>
	<b>Bibliography</b> . . . . .	<b>141</b>



# List of Figures

1.1	The Otto cycle . . . . .	9
1.2	The Diesel cycle . . . . .	10
1.3	The Brayton cycle . . . . .	12
1.4	Transesterification . . . . .	20
330	2.1 Cyanocobalamin . . . . .	24
	2.2 The carbon dioxide phase diagram . . . . .	29
	2.3 SFE system diagram . . . . .	31
	2.4 SFC system diagram . . . . .	32
	2.5 FID diagram . . . . .	33
335	3.1 Discussion of SANS 1935. . . . .	41
	3.2 Cetane and isocetane . . . . .	42
	3.3 The peroxidation of lipids. . . . .	44
	3.4 Derivitization . . . . .	49
	3.5 Overlap of requirements . . . . .	51
340	4.1 Filling a CO <sub>2</sub> pump. . . . .	54
	4.2 A process control system . . . . .	56
	4.3 A cutaway diagram of the mixing chamber . . . . .	57
	4.4 Schematic diagram of the injection valve. . . . .	58
	4.5 Schematic diagram of a continuous-flow valve. . . . .	59
345	4.6 A diagram of a integral restrictor. . . . .	60
	4.7 A electron microscope photo of a restrictor orifice . . . . .	61
	4.8 A electron microscope photo of a blocked restrictor orifice . . . . .	62
	5.1 Schematic diagram of a the chromatographer's trilemma. . . . .	67
350	5.2 A plot of H vs. F . . . . .	69
	5.3 A temperature-rate table from the Agilent7890B data sheet . . . . .	72
	5.4 Coaxial heater resistance heater . . . . .	73
	5.5 A photograph of the setup used to record the thermal video . . . . .	77
	5.6 A thermograph of the coaxial heater at temperature . . . . .	77
	5.7 Welding process schematic . . . . .	80
355	5.8 Fine-wire welder . . . . .	80
	5.9 A microphoto of a thermocouple twist ready to be welded. . . . .	81
	5.10 A cartoon explaining how to construct a long, thin thermocouple probe. . . . .	82
	5.11 Calibration curve of the coaxial heater . . . . .	83
	5.12 Measurement curve of the coaxial heater . . . . .	84
360	5.13 Effect of controller tuning . . . . .	88
	5.14 Heating rate illustration . . . . .	89
	5.15 The main LabVIEW VI . . . . .	90
	5.16 A single fast GC chromatogram . . . . .	91
	5.17 A 2D SFCxGC chromatogram . . . . .	91

365	5.18 A 2D SFC $\times$ GC chromatogram . . . . .	92
	6.1 Separation of FAME by GC . . . . .	95
	6.2 SFC $\times$ GC of canola oil . . . . .	100
	6.3 SFC $\times$ GC of sunflower oil . . . . .	101
	6.4 3D Bar chart of fatty acid profile . . . . .	102
370	6.5 SFC $\times$ GC of coconut oil . . . . .	103
	6.6 SFC $\times$ GC of flax seed oil . . . . .	104
	6.7 Biodiesel separated from petrodiesel. . . . .	106
	6.8 An example of a petrochemical fuel chromatogram. . . . .	107
	7.1 Modifiers in SFC . . . . .	114
375	D.1 Manifold dimensions. . . . .	130
	D.2 Manifold assembly. . . . .	131
	D.3 Coolant reservoir . . . . .	132
	D.4 Technical drawing of coaxial heater mounting rails . . . . .	133
	D.5 Technical drawing of coaxial heater mounting. . . . .	134

## List of Tables

3.1	A few well-known standards organizations . . . . .	38
6.1	A summary of retention time repeatability of alkanes separated on the fast temperature programmed chromatograph. . . . .	97
6.2	Oils used for FAME analysis . . . . .	98
385	E.1 The fatty acid profile of sunflower oil sample . . . . .	138
	E.2 Fatty acid profile of coconut oil according to the FAO (Joint FAO/WHO Codex Alimentarius Commission 2019) . . . . .	139



## List of Abbreviations

<b>EROEI</b>	Energy returned on energy invested.
<b>LCA</b>	Life Cycle Analysis
<b>LAH</b>	List Abbreviations Here
<b>WSF</b>	What (it) Stands For
<b>UNFCCC</b>	United Nations Framework Convention on Climate Change
<b>IPCC</b>	Intergovernmental Panel on Climate Change
<b>CCGT</b>	Combined cycle gas turbine
<b>HPLC</b>	High performance liquid chromatography
<b>GC</b>	Gas chromatography
<b>SFC</b>	Supercritical fluid chromatography
<b>CCGT</b>	Combined cycle gas turbine
<b>CCGT</b>	Combined cycle gas turbine



## Physical Constants

Constant Name    $Symbol = ConstantValue$  with units



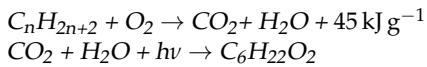
## List of Symbols

$a$	distance	m
$P$	power	$\text{W} (\text{J s}^{-1})$
$\omega$	angular frequency	rad



*Dedicated to the memory of Dr. Barbara Lotze*





Hydrocarbons give heat and sunshine makes sugar

# 1

## Introduction: Climate, fuel, and biodiesel

### *<Chapter1>* 1.1 Energy, fuel and the atmosphere

395 Industrialized societies depend on reliable sources of energy. For the purpose of this discussion, the main types of energy can be counted as electricity and fuels, although they are interconvertible. Fuels for industrialized societies are predominantly found as underground mineral deposits, from where they are extracted by mining or drilling. They are found as solids, liquids and gases, which the energy  
 400 industry refers to as *coal*, *crude oil*, and *natural gas*. Because these fuels are of biological origin, deposited during previous geological eras, and then metamorphosed and preserved by geological processes, these fuels are commonly referred to as 'fossil fuels'.

The large-scale exploitation of fossil fuels started in the middle of the 18th century, when the plentiful coal from the coalfields of Great Britain drove a development that is known to history as the Industrial Revolution. This development is closely associated with steam engines (Rosen 2012). The use of crude oil started in the middle of the 19th century, and is associated with the development of the automobile (Watts 2005, p. 42). The use of natural gas started in the middle of the 410 20th century and is associated with the introduction of gas-fired central heating in for homes in cold climates (Hanmer and Abram 2017) Fossil fuels also serve as feedstock to the chemical industry, but that topic lies outside the scope of this discussion.

There is no doubt that the use of fossil fuels as an energy source greatly improved the human lot. The mechanization of agriculture and the easy distribution of food 415 by motorized transport have eliminated famine as a natural disaster (Angelis et al. 2007). The distribution of medical supplies by motorized transport and the rapid deployment of medical personnel have limited the impact and spread of epidemics (Ministere de la santé 2018). Heating and cooling of buildings have increased the habitable zone on earth. Artificial light has increased the hours available for mental activity, in particular extending the reach of entertainment, art and education.  
 420

However, in the context of chemistry, the uncontrolled use of fuels has at least two major problems.

The first problem with fossil fuel is that it is finite. There is only a certain amount of fossil fuel on earth. If all of it is extracted it will no longer be a reliable source 425 of energy, and the existence of industrialized societies and the complex civilizations

that depend on them will be in jeopardy. It is tempting to think that civilizations would have the foresight to prepare for such an eventuality, but the historical record shows that societies can collapse when at the height of their powers (Diamond 2006).

The second problem with fuels is that they produce pollution wherever they are produced, processed, transported and used. Pollution is injurious to the health and well-being of individuals, societies and nature. If pollution is serious enough, it not only degrades society and the environment, but also negates the benefits brought by the application of energy: modern hospitals are energy-intensive, but if they are filled with victims of pollution there is no nett benefit.

The first approach to pollution from fuels has been to ignore it. Pictures from the early industrial revolution shows English towns coated with soot and choked with smoke (Flick 1980), and rivers became toxic sewers (Halliday 2001).

The development of public health, social responsibility (Sreter 2003) and an embryonic environmental movement (Williams 1965) lead to political pressure for the implementation of pollution controls, which governments gradually introduced and increased in strictness.

The first generation of pollution control offered essentially two options: concentrate or dilute.

I will illustrate these two options using a typical South African coal-fired power station. In such a power station coal is burned to produce heat, which converts liquid water into high-pressure steam. The steam is allowed to expand through a turbine, which converts the energy in the steam into rotary motion, which is used to turn an alternator that produces an electric current by rotating a set of electrical conductors in a magnetic field. The furnaces of such a power station produce a flow of waste. This waste is an *aerosol*: finely divided solids suspended in a mixture of gases.

The solid part is typically separated from the gas-phase part by filter bags and electrostatic precipitators: the collected material is known as *fly ash*. The gas-phase part might be sent through scrubbers to remove some of the gas-phase pollutants, capturing it in a solid form. The solid part of the power station's furnace waste has now been concentrated. It is transported to a storage site, where it is stored indefinitely. It goes without saying that concentrated pollutants should be encapsulated during storage in some way, otherwise they just become more sources of pollution.

The gas-phase part of the pollutant is handled by diluting it. The outlet for the gas-phase stream of waste is through a tall stack, which ends high above ground level<sup>1</sup>. At this altitude the wind is strong and steady, which rapidly carries the gases and remaining aerosols away and disperses them.

The dilution of pollutants might seem like an abdication of responsibility, but it is an reasonable response to pollution. At low enough concentrations pollutants that enter the biosphere are broken down by sunlight and microbes which render it harmless. This makes dilution a reasonable first attempt at controlling pollution.

The devil is, of course, in the detail. For example, mercury that find its way into the environment is eventually converted by microbes to methyl mercury, which concentrates in aquatic animals. All rivers in the continental USA are now polluted by airborne mercury that originate from coal-fired power stations (Wentz et al. 2014). Some persistent organic pollutants, which also concentrate in the food chain, originate in fuel combustion. So while dilution was a reasonable first attempt at controlling pollution, it it certainly not the final answer.

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<sup>1</sup>The chimneys of those sooty Victorian towns were not there to disperse the smoke, but to create a 'draft', a flow of air created by the buoyancy of hot air. The better the draft, the more efficient the fire.

The majority of fuels provide their energy as heat, which can be converted into useful work. This heat is obtained by combining the chemical compounds found in the fuel with atmospheric oxygen to form compounds with lower internal energy. The maximum amount of work that can be extracted from a given fuel can be estimated by examining the Gibbs free energy equation:

$$\Delta G = \Delta H - T\Delta S \quad (1.1) \text{ ?eq: Gibbs?}$$

For a given compound fuel compound,  $\Delta H$  is determined by the difference in enthalpy of formation of the product waste compounds and the enthalpy of formation of the reactant fuel compounds. Since the reactant fuel compounds are given,  $\Delta H$  is maximized by having product compounds with very low enthalpies of formations.

To maximize  $\Delta S$ , the products should be as disordered as possible. This implies that *gas-phase* products composed of *small molecules* will yield more work.

The temperature  $T$  should also be as high as possible.

Because the industrial machinery in economically competitive, capital-intensive industries are highly efficient, the maximum amount of energy is extracted from their fuels for the lowest cost. Following Gibbs, the major compounds left over after extracting the energy from a fuel should have very low enthalpies of formation and be in the gas phase at the temperature of the process.

Because all fuels contain carbon as a major component, the extraction of energy yields compounds containing carbon. Most fuels also contain hydrogen. Reacting these fuels with the oxygen in air to extract maximal work will therefore yield water ( $H_2O$ ) and carbon dioxide ( $CO_2$ ).

(Of course the argument is not that Victorian engineers designed steam engines using the Gibbs energy equation, but the variation-selection process (Vincenti 1990, Chapter 8) by which engineering improvements accrue would inevitably drive the development of heat engines fuelled by fossil fuels to emit large quantities of carbon dioxide gas.)

Water is of course not a pollutant at all, and carbon dioxide is only toxic at very high concentrations, and therefore, for most of the industrial era, carbon dioxide was easily dealt with by diluting it in the atmosphere, where it is also naturally present at low levels. To the extent that carbon dioxide was considered a pollutant it was assumed that it would be absorbed by the biosphere.

Photosynthesis, the process by which green plants capture the energy from sunlight to live and grow, does indeed remove carbon dioxide from the atmosphere, and an elementary model of the carbon cycle that assumes stability would seem to indicate that excess carbon dioxide in the atmosphere would be captured by photosynthesis and sequestered, leaving the carbon dioxide concentration in the atmosphere stable. Of course this model contains testable assumptions, which scientists could, and did, test.

The most famous of these tests is probably the “Keeling Curve” (Harris 2010). This is a continuous record of measurements of concentration of atmospheric carbon dioxide in the pristine air of the Pacific Ocean. This record starts in 1958 and shows that the carbon dioxide concentration of the atmosphere is increasing.

Paleoclimatologists have studied the hypothesis that the carbon dioxide concentration is stable over time. Not only have they found that the carbon dioxide concentration is *not* stable, they have also determined that the pre-industrial concentration of atmospheric carbon dioxide was lower than it was in 1958 (Petit et al. 1999).

520 So it is clear that the carbon dioxide concentration in the atmosphere is rising because of the biosphere is not absorbing all the carbon dioxide produced by the combustion of fossil fuels.

525 The projected concentration of carbon dioxide in the atmosphere is, however, still not at or even near toxic levels, which might make the continued dilution of carbon dioxide in the atmosphere seem a viable disposal method.

530 Emitted gases, however, do not only have chemical properties and biological impacts, they also have physical properties. Of concern for the current discussion is carbon dioxide's absorbance of electromagnetic radiation, in particular the radiation arriving from the sun, and the radiation from the earth's surface out to space, both which must pass through the atmosphere. A molecule can absorb parts of this radiation by having its electrons excited, or by changing its vibration. Carbon dioxide is electronically very stable, and therefore absorbs only extreme ultraviolet radiation, just like the major gases in the atmosphere, molecular nitrogen and molecular oxygen. However it poorly absorbs near ultraviolet and visible light, in which it is also similar to the major gases. Because there is a dipole moment between the carbon atom and the oxygen atoms of carbon dioxide, it has vibrational modes which can absorb radiation in the infrared region of the electromagnetic spectrum. Oxygen and nitrogen do not have dipole moments, and therefore do not absorb infrared radiation strongly. This means that carbon dioxide will absorb infrared radiation in the atmosphere much more strongly than the major gases. The energy from the absorbed infrared radiation is of course turned into vibration, a form of kinetic energy, and this kinetic energy is randomly further distributed among the gases of in the atmosphere, appearing as an increase in temperature as the energy distribution moves towards equilibrium.

540 545 The most cursory understanding of the absorption of infrared radiation in the atmosphere by carbon dioxide therefore seems to say that an increase in the concentration of carbon dioxide in the atmosphere would lead to a higher average temperature in the atmosphere. This increase in temperature was first estimated by Svante Arrhenius in 1896 (Arrhenius 1897).

550 (The other major product of extracting energy from fuel is water. Its molecule also has a dipole moment, and it also absorbs infrared radiation strongly. However, its intermolecular properties sets an upper limit to its concentration: at high enough concentrations it will either condense into water or crystallize into ice and precipitate from the atmosphere to end up as surface water.)

555 560 Further research has only confirmed that industrial processes extracting energy from fuel is increasing the amount of carbon dioxide in the atmosphere by gigatons every year, and that this increase in concentration is leading to higher temperatures in the atmosphere. It had also become increasingly certain that this increase in temperature is inevitable and significant: it will change earth's climate<sup>2</sup>. This projected increase in temperature and the accompanying change in global climate is bound to have impact on societies within the lifetime of people alive today, and a larger impact on future generations. Some of these projected changes are incompatible with the maintenance of the complex civilizations that are supported by industrialized societies (IPCC 2014).

565 Scientists do not inhabit ivory towers, and their research is funded by public money in the expectation that the resulting science will benefit their societies. As the

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<sup>2</sup>The atmosphere is in contact with the hydrosphere, and the two exchange carbon dioxide. The carbon dioxide diluted in the atmosphere is therefore also diluted in the oceans. This leads to *ocean acidification*, a gigaton problem in its own right, but one that falls outside the scope of this discussion.

understanding of carbon pollution grew, it became obvious to scientists that industrialized societies could not just continue diluting gigatons of carbon dioxide in the atmosphere, and they alerted policymakers. The iconic moment of this development  
570 was the testimony of Dr James E. Hansen, at the time director of NASA's Goddard Institute of Space Studies, before the Senate of the US Congress in 1988 (Shabecoff  
[1988](#)). In the same year the scientists of the world came together and established the Intergovernmental Panel on Climate Change (IPCC) to create a coherent body of knowledge to inform decision-making. The IPCC has so far produced a series of five  
575 Assessment Reports, which assesses the science of climate change, its impacts, and ways to mitigate it (Allen et al. [2014](#)).

(Carbon dioxide is not the only greenhouse gas and the IPCC reports consider each in detail, but because the other gases are not produced in significant quantities by fuels and the energy industry they fall outside the scope of this discussion.)

580 Industrialized societies currently depend on energy from fossil fuels. Huge investments have been made in infrastructure, large numbers of people find employment in this industry, and available alternative sources of energy are not nearly sufficient. The health and well-being of the greater part of the world's population depend on it. But the future health and well-being of those same societies also depend on  
585 ending that dependence. It would be in the best interest of industrialized societies to create plans to reduce the dependence on energy from fossil fuels and exploit alternative energy sources.

But industrialized societies are also in economical competition with each other.  
590 This means that any society that spends resources on risky, expensive alternative sources of energy risks falling into a competitive disadvantage, and there is a surfeit of short-sighted politicians who will take advantage of this risk to create fear and so prevent planning for and investment in changes in energy production and use<sup>3</sup>.

Fortunately there are enough leaders who have vision, and based on the recommendations of the IPCC the governments of the world have come together to create  
595 the UNFCCC. The United Nations Framework Convention On Climate Change is an international agreement that structures the response of nations to limit their emissions of greenhouse gases. Two international treaties have been agreed to so far: the Kyoto Protocol in 1997 ([Kyoto Protocol to the United Nations Framework Convention on Climate Change. 1997](#)) and the Paris Agreement in 2015 ([Paris Agreement 2015](#)).

600 These agreements attempt to reduce the amount of carbon dioxide released into the atmosphere by dividing the cost of reducing emissions fairly. While the Kyoto Protocol invoked a complex carbon credit trading scheme, in the Paris Agreement nations pledge to reduce emissions of carbon dioxide into the atmosphere, and each is free to do so in a way that suits them best.

605 The Republic of South Africa is a signatory to the Paris Agreement, and therefore the country is legally bound to limit and reduce its emissions of carbon dioxide and other greenhouse gases. The country has so far promised to stop increasing emissions between 2020 and 2025 ([South Africa's Intended Nationally Determined Contribution \(INDC\) 2016](#)).

## 610 1.2 Carbon footprints and carbon neutrality

Ending the emission of carbon dioxide by industrialized societies is an exercise in balancing effort with consequences. Every reduction in carbon dioxide emissions

<sup>3</sup>The impact of political corruption is not negligible, but I will abstain from commenting on it.

must necessarily have an effect on the economy, and conversely, every change in economic activity will have an effect on carbon dioxide emissions.

615 A rational, discerning society in searching for ways to carbon dioxide emissions will therefore attempt to change their economy in such a way that the emission of carbon dioxide is limited or reduced. But in most societies there are many ways to reduce carbon dioxide emissions, and decisions have to be made on which ones to implement. Such decisions must be based on sound information, and one way to generate that information is a discipline called *life cycle analysis* (LCA).

620 Life cycle analysis can yield rigorously-calculated data and comparisons, but it is a very general method. For analyses that are very similar and differ only in context, it is possible to develop simplified, standardized life-cycle analysis tools.

One such tool is the *carbon footprint*. Every economic activity in an industrialized 625 society generates emissions of greenhouse gases. These emissions might be far removed from the activity in space and time. A most obvious example of this is the use of electricity: At the moment I'm using a computer to compose this paragraph and electric light illuminates the desk I'm working on here in my office in Pretoria. The electricity that powers the computer and the lamps are generated hundreds of 630 kilometres from here, most of it on the Mpumalanga Highveld, in coal-fired power stations which emits carbon dioxide in the process. So my activity is far removed from the associated emissions **in space**. But even before the coal was burned, the machines that mined it and transported it to the power station emitted carbon dioxide, so my activity is also removed from the emissions **in time**. The sum of all these 635 emissions connected with my activity constitute the carbon footprint of my activity.

Although carbon footprinting is conceptually straightforward, it is analytically rigorous and computationally complex. But there are standard protocols (World Resource Institute 2004) that can be followed, so that carbon footprints from different activities and different organizations are transparent and comparable.

640 Once an organization's activities have been footprinted, the organization can examine its operations and look at ways to change activities that will reduce that footprint.

If my university, for example, decides to install photovoltaic panels on the roof 645 of my building, and use the electricity to illuminate my office, the activity will now most likely have a reduced carbon footprint, because photovoltaic electricity usually has a lower carbon footprint than coal-derived electricity.

To prevent a climate disaster, at some point industrialized societies will have to change all their activities until they no longer emit any greenhouse gases. Such activities will have carbon footprints of zero, and will be known as 'carbon neutral' 650 activities'. There are also activities that are nett removers of carbon from the atmosphere, which would be 'carbon negative'.

In truth, in industrialized societies there are very few carbon neutral activities. Even if I'm just sitting quietly in a pristine nature reserve, the food I am digesting and metabolizing have a positive carbon footprint: carbon was emitted to fix nitrogen 655 from the air, which was used as fertilizer to help grow the food I ate, and is now in the proteins of my body. The very calcium and phosphorus in my bones were sourced from mines that were powered by fossil fuels.

## 1.3 Internal combustion engines.

Carbon footprinting exercises quickly show that a major source for emissions is transport. Most transport in industrialized societies is powered by internal-combustion engines. They are called ‘internal-combustion’ engines because the chemical transformation that extracts the energy from the fuel is internal to the engine. This is in contrast to the power station steam turbines mentioned earlier, where the extraction of the chemical energy from the fuel happens outside of the engine. The engine is the device that produces the mechanical energy. These engines are usually powered by liquid fuels derived from crude oil, although some designs run on gas, and of course in South Africa Sasol and PetroSA supply us with liquid fuels derived from coal and natural gas.

This discussion will use the term *noxious pollution* to describe pollution that is directly harmful to humans and *carbon pollution* to describe pollution that leads to industrial climate disruption.

The simplest way to reduce an activity’s carbon footprint is to use the fuel that drives that activity more efficiently. Fortunately market forces are aligned with this, because more efficient use of a fuel also reduces running costs. Discussing engine designs therefore almost always invokes *efficiency*. Efficiency of an engine is calculated by taking the ratio output power:input power. The greek letter  $\eta$  is used as a symbol for efficiency:

$$\eta = \frac{P_{out}}{P_{in}} \quad (1.2) \text{?eqn:efficiency?}$$

Conservation of energy dictates that the output power cannot be more than the input power, and therefore  $\eta \leq 1$ . The power is ‘lost’ to two factors: firstly entropy, as dictated by the Second Law of Thermodynamics, and secondly losses, such as friction.

### 1.3.1 Scaling and efficiency

(par:scaling) The efficiency of engines scales disproportionately with its linear dimensions, by which is meant that the larger an engine is, the more efficient it is. There are three factors that improve efficiency in larger engines (Brown, Menon, and Hagen 2015):

1. Longer residence time: In larger engines, the reacting fuel spends more time in the engine. This gives reactions more time to reach equilibrium, and hence more energy can be extracted.
2. Smaller surface area to volume ratio: The volume of an engine increases with the cube of its linear dimension, but the surface area increases with the square of its linear dimensions. This means that larger engines have a smaller surface-to-volume ratio than smaller engines. Since losses take place at the surface, the loss per unit volume is therefore larger for smaller engines than it would be for larger engines of the same type. If the losses in larger engines are comparatively smaller, we can expect larger engines to be more efficient than smaller ones.
3. Larger Reynolds number: The Reynolds number is a number that predicts the onset of turbulent flow, and is calculated from the viscosity, velocity and a characteristic linear dimension. The higher the Reynolds number, the higher the likelihood of turbulent flow. Turbulent flow promotes mixing, which leads

to improved combustion and hence higher efficiency. The larger the engine, the larger the characteristic linear dimension, and therefore the higher the likelihood of turbulent flow.

From the viewpoint of reducing carbon pollution, one should therefore aim to  
705 use a few large engines rather than a multitude of small engines.

### 1.3.2 Engine design and thermodynamic cycles.

*(par:efficiency)* The final design of a successful internal combustion engine is determined by the variation-selection process described by Vincenti (Vincenti 1990). Such a design sufficiently satisfies a wide range of requirements. These requirements might be explicitly expressed in documentation, or they may be implied, or they may be practical.  
710 These requirements would include, but are not limited to, capital cost, running cost, maintenance cost, noise, sound, ease of maintenance, power output, emissions, surface finish, weight, mounting method, size, supply chain capability, shape, colour, fuel availability, operating temperature, altitude tolerance and torque.

From designers' attempts to fulfil these requirements arise the myriad of different engine designs, delivering anything from milliwatts to megawatts of power to anything from model aircraft to oil tankers. It is the good fortune of researchers who  
715 study engine efficiency that there are only a few conceptual systems that explain how the internal combustion engines convert the chemical energy of the fuel into mechanical work.

The principles on which the different engines operate are named 'cycles', because the way they convert heat into work can be described in terms of a series of events that are repeated endlessly. The concept of a cycle is further entrenched because it is customary to explain the thermodynamic processes involved using at  $pV$  diagram,  
720 showing the changes of state of the working fluid during engine operation.

Theoretical thermodynamic cycles are useful because they can predict the performance of heat engines, allowing for comparison. In particular, it allows for comparison with theoretical maxima. The maximum efficiency of a heat engine is delivered by an engine running on the Carnot cycle. This describes a hypothetical engine that  
730 uses only reversible processes to extract useful work from the temperature difference between two reservoirs of heat. It has been shown that the Carnot cycle offers the maximum possible heat extraction, and that the maximum efficiency ( $\eta_C$ ) is determined solely by the temperature difference between the hot ( $T_h$ ) and the cold ( $T_c$ ) reservoirs.

$$\eta_C = 1 - \frac{T_c}{T_h} \quad (1.3) \text{ ?eqn:Carnot?}$$

The Carnot cycle, however, can only deliver an infinitesimal amount of work, because the heat transfer must be reversible, and therefore infinitesimal. A more realistic maximum efficiency is given by the Chambadal-Novikov efficiency ( $\eta_{CN}$ ) (Hoffmann 2008):  
735

$$\eta_{CN} = 1 - \sqrt{\frac{T_c}{T_h}} \quad (1.4) \text{ ?eqn:Chambadal-Novikov?}$$

This theoretical efficiency takes into account the irreversible processes that are  
740 inevitable in engines delivering finite amounts of power.



FIGURE 1.1: The pV diagram of the thermodynamic Otto cycle. The isochoric heat addition step (BC) corresponds to the burning of the homogeneous air-fuel mixture. Work is extracted from the engine during the adiabatic expansion CD (Wolfram | Alpha 2019a).

?<fig:otto-cycle>?

Barring revolutionary new discoveries and inventions, there will be only three thermodynamic cycles and their corresponding internal combustion engines on the market, as we enter the low-carbon era.

### Otto engine and Otto cycle

<sup>745</sup> The oldest of the internal combustion engines is the Otto engine. It is named after Nikolaus Otto, who developed the first working engine of this kind in 1876 (Cummins 1989, Chapter 9). In South Africa these engines are usually called ‘petrol engines’.

<sup>750</sup> In the Otto engine, a homogeneous mixture of atmospheric air and vaporized fuel is compressed. This mixture is then ignited by an electric spark. The energy released by the chemical reaction between the oxygen in the air and the fuel vapour causes the temperature of the compressed air to rise, and consequently the pressure. If this, now hot, gas is allowed to expand in an expandable vessel, the motion of the vessel can be captured to perform useful work. Once the useful work is extracted, <sup>755</sup> the vessel can be collapsed again to expel the exhausted air, and re-filled with a compressed air/fuel vapour mixture. This completes the cycle. The expandable vessel is usually in the form of a cylinder and piston, with the piston connected to a crank that drives a shaft that transfers the work from the engine to the machine being powered.

<sup>760</sup> (These legs of the cycle are only loosely related to the ‘strokes’ of a four-stroke engine, and should not be confused for them.)

The theoretical efficiency of the Otto Cycle is given by



FIGURE 1.2: The pV diagram of the thermodynamic Diesel cycle. The isobaric process BC corresponds to the combustion of the finely divided fuel particles in air. Work is extracted from the engine both during this process and during the adiabatic process CD. (Wolfram | Alpha 2019b)

`(fig:diesel-cycle)`

$$\eta = 1 - \left( \frac{1}{r^{\gamma-1}} \right) \quad (1.5) \quad \text{eqn:otto-efficiency}$$

where  $r$  is the *compression ratio*,  $\frac{V_1}{V_2}$ , and  $\gamma$  is the heat capacity ratio, which can be considered a constant for the purposes of this discussion (Wolfram | Alpha 2019a).

This means that the efficiency of an Otto engine can be improved by increasing the degree of compression of the intake air before combustion.

### Diesel Cycle

In the Diesel engine, named after Rudolf Diesel who demonstrated the first engine of this type in 1897 (Cummins 1989, Chapter 14), air is compressed, and then a finely divided solid or liquid fuel is injected into the system. The fuel then reacts with the oxygen in the air. The energy released in the reaction appears as a higher temperature in the gas, and — following Gay-Lussac's Law — the pressure of the gas rises. If the gases expand in the confines of a collapsible vessel, useful work can be extracted from expansion of the vessel. Once the work has been extracted, the vessel can be collapsed again to remove the now inert ('exhausted') gas and prepare for receiving the next charge of air, completing the cycle.

The theoretical cycle used to analyse the performance of the Diesel Engine is called the Diesel Cycle. (See figure 1.2)

The Diesel cycle differs from the Otto cycle in the heat addition step. In the Otto engine, the heat addition takes place when the homogeneous air/fuel mixture combusts. This combustion takes place in a short space of time during which the

engine parts move only a negligible distance. Hence the pressure rises rapidly. In the Diesel engine, the fuel is injected into a volume of compressed air, where it combusts. (There is no separate ignition source: the temperature of the adiabatically compressed air is higher than the fuel's ignition temperature, so that the fuel ignites upon injection. While this is another difference between the two engines, it is of second-order importance when discussing thermodynamic cycles.) Because the combustion takes place at the surface of fuel particles, the combustion rate is lower than the combustion rate of the homogeneous mixture in the Otto cycle. Hence the rising temperature is balanced by the motion of the engine, and the heat addition is essentially isobaric.

The efficiency of the Diesel cycle is given by Equation 1.6

$$\eta = 1 - \frac{1}{r^{(\gamma-1)}} \left( \frac{\alpha^\gamma - 1}{\gamma(\alpha - 1)} \right) \quad (1.6) \quad \text{eqn:diesel-efficiency}$$

where  $\alpha$  is the *cut-off ratio* and  $\gamma$  is the compression ratio.

The term  $\frac{\alpha^\gamma - 1}{\gamma(\alpha - 1)}$  is always larger than 1, and therefore, when we compare equation 1.6 with equation 1.5 it is clear that for a given compression ratio the Diesel cycle is always less thermally efficient than the Otto cycle.

However, at high compression ratios Otto engines start suffering from 'knocking', when the homogenous air/fuel mixture detonates, instead of burning smoothly. The shock waves from this detonation will damage the engine and lead to faster wear. (The *octane number* of a fuel indicates the compression ratio it can accommodate.) Because Diesel engines do not suffer from this problem, they can, and usually are, designed to operate at higher compression ratios than Otto engines. Otto engines normally operate with a compression ratio of up to 9:1, whereas Diesel engines have compression ratios of up to 25:1. This makes Diesel engines significantly more efficient than Otto engines.

### Gas turbines and the Brayton Cycle

The third internal-combustion engine important to industrialized society is the *gas turbine*. Gas turbines are used rarely in road transport, more often in marine and stationary applications, but thousands take to the sky every day, propelling aircraft (H. Morris 2017) carrying billions of airline passengers every year.

The first gas turbines were developed during wartime urgency to deliver pure jet thrust for military aircraft, but this proved to be an inefficient use of the available energy. Most modern turbine engines drive a shaft to extract rotational work. This shaft might drive a bypass fan (as used in airliner engines), a propeller (as used in smaller, low-speed aircraft), or a shaft which might drive a helicopter rotor or an electrical generator.

In operation, a gas turbine compresses air with one or more compressor stages. The compressed air passes through a combustion chamber, where fuel is added and combusted. As the now hot gases expand they pass through one or more turbine stages, which is connected to drive the compressor and the output shaft, which extracts work from the system. After the gas has passed through the turbine it returns to atmospheric pressure, optionally doing work on the engine in the form of thrust.

The theoretical thermodynamic cycle that is customarily used to analyse the gas turbine is called the Brayton cycle, named after George Brayton, who successfully manufactured a reciprocating engine based on this cycle. (See figure 1.3) Such engines are no longer manufactured (Cummins 1989, Chapter 10).

The efficiency of the Brayton cycle is given by equation 1.7



FIGURE 1.3: The pV diagram of the thermodynamic Brayton cycle. In the gas turbine the continuous combustion of the injected fuel and free expansion of the air through a turbine means that the heat addition (BC) is isobaric. Work is extracted during the adiabatic expansion process CD (Wolfram | Alpha 2019c).

`(fig:brayton-cycle)`

$$\eta = 1 - \frac{T_1}{T_2} = 1 - \left(\frac{P_1}{P_2}\right)^{\frac{1}{\gamma}} \quad (1.7) \quad \text{eqn:brayton-effic}$$

where  $\gamma$  is the compression ratio (Wolfram | Alpha 2019c).

- Because the fuel is continuously added to the compressed air and the combustion takes place in a heterogeneous mixture, gas turbines do not suffer from 'knocking' and therefore there is no upper limit to the compression ratio. The main figure determining the efficiency of the engine is the turbine inlet temperature, (*i.e.* the outlet temperature of the combustion chamber) which reaches 1600 °C in modern engines.

### 1.3.3 Noxious pollution from internal combustion engines.

- Carbon pollution is not the only pollution emitted by internal combustion engines. There is also noxious pollution, which affects the societies in which they are used. There are ways to reduce or eliminate such pollution, but the implementation of such measures affect the efficiency of the engine.

Noxious pollution from internal-combustion engines come from different sources and have different effects.

#### Unburnt fuel

- The combustion reactions in internal-combustion engines are very fast, but they are never at equilibrium. This means that at the end of a cycle, the exhaust gases expelled from the engine contains, besides the carbon dioxide, also chemical intermediates and unreacted fuel molecules.

Unreacted fuel is a noxious pollutant, not only in its own right, but also, during the decomposition process in the environment, chemical reactions induced by sunlight produces ozone, a reactive oxygen species that leads to respiratory problems in victims of pollution (Davidson 1998).

Some of the fuel-like pollutants from internal-combustion engines are not present in the fuel. These compounds are formed in a process that might be called *pyro-synthesis*. They are themselves more stable than the compounds from which they originate, but not as stable as the combustion end products, carbon dioxide and water. Because they are themselves quite stable, they require high temperatures to combust, which might not be achieved before they leave the engine as waste. Notable pollutants from this source are the polycyclic aromatic hydrocarbons (PAH) and, if trace amounts of chlorinated substances are present in the fuel, dioxins

One of the products of pyro-synthesis is particles of soot. These particles are agglomerations of nano-sized particles of pure, amorphous carbon. The agglomerations might include adsorbed PAHs and acids. Soot particles are very stable and only react at very high temperatures. Diesel engines in particular have a reputation for producing soot (Mohankumar and Senthilkumar 2017).

The final pollutant that can be classed as originating from partially combusted fuel is carbon monoxide. It forms readily when fuels are burned in oxygen-poor environments. Modern, well-maintained engines is a relatively rare cause of acute carbon monoxide poisoning (Reumuth et al. 2018), but chronic exposure to low levels of environmental carbon monoxide has harmful effects (Wright 2002).

### Contaminants and additives.

Crude oil and coal contains not only carbon and hydrogen, but also other elements, most notably sulfur and nitrogen. Depending on the refining process, these elements might find their way into fuels. Organic nitrogen and sulfur will easily oxidize and form stable oxides, yielding energy. Once outside the engine, however, the volatile oxides will dissolve in any water in the air, forming acids. (Duncan et al. 2016)

Fuel manufacturers add additives to fuels for various reasons, such boosting octane rating or preventing corrosion. The prime example of additives as a source of noxious pollution was tetraethyllead, which was added to petrol as an octane booster. The emitted lead compounds was shown to be a pollutant that lead to neurological impairment in its victims, and its use was phased out (Needleman 2000).

### Side-reactions

Most internal combustion engines use atmospheric air as oxidant. This air is 21% oxygen, and 78% nitrogen. Although nitrogen is very stable and inert, at high temperatures it can react with oxygen. These reactions produce the oxides of nitrogen: NO, NO<sub>2</sub>, NO<sub>3</sub> and N<sub>2</sub>O, and N<sub>2</sub>O<sub>5</sub>. They are collectively called NO<sub>x</sub>.

These oxides can participate in the cycle that causes the photochemical smog. They will also dissolve in atmospheric water and form acids: this might happen far from the point of emission, resulting in *acid rain*.

#### 1.3.4 Mitigating pollution from internal-combustion engines.

Pollution from internal combustion engines is a social ill, and most governments have regulations in place to limit and reduce this pollution. Engine and fuel manufacturers are working hard to reduce this pollution.

The noxious pollution from internal combustion engines can be reduced by various improvements, but there are three complementary approaches: fuel formulation (Gertler et al. 1999), cleanup (Braun et al. 2018) and engine management (Reif 2015).

### Fuel formulation

- 895 Adding oxygenates improves the octane rating of fuel and reduces NO<sub>x</sub> formation.

### Exhaust gas cleanup

(par:cleanup) Exhaust gases from internal combustion engines can be “cleaned”, and there are basically two approaches. The first is catalytic conversion, in which the exhaust gases that contain the pollutants are passed over a catalyst bed. The catalyst (a proprietary formulation of platinum, palladium and/or rhodium), adsorbs the uncombusted volatile organic carbons, and oxidizes them. It simultaneously catalytically decomposes NO<sub>x</sub> to molecular nitrogen and oxygen.

900 Secondly, filter systems can be used to remove particulate matter.

### Engine management

(par:engine-management) As described above, the main source of noxious pollution from internal-combustion engines is uncompleted or undesired chemical reactions, and is not fundamental to the operation of the engine. By carefully managing the engine system, noxious pollution can be reduced.

910 The oldest engine management technology is the oxygen or ‘lambda’ sensor, which measures the oxygen in the exhaust gases. Such a sensor, coupled to an engine-management computer, allows the metering of the exact amount of fuel needed for optimum combustion (Frauhammer, Schweinsberg, and Winkler 2014).

915 A newer technology is known as exhaust gas recirculation. This mixes the intake air of the engine with exhaust gases, effectively diluting the oxygen. This reduces peak temperatures, and thereby NO<sub>x</sub> formation.

In *stratified charge* engines the distribution of fuel in the volume of intake air is controlled by selective fuel injection. Carefully injecting the fuel at the right place at the right time can allow for higher compression ratios without inducing pinging. *Lean-burn* engines are Otto engines that use extremely high air:fuel ratios.

920 Electronic engine management systems result in much more efficient and less-polluting engines than non-managed ‘mechanical’ engines, but because engines for automotive applications endure such a wide range of operating conditions, they can at best achieve a compromise between power, efficiency and emissions. It was this unsatisfactory compromise that lead to the Volkswagen emissions scandal: manufacturers chose to cheat on emissions test rather than admit to the relatively poor performance of a managed engine optimized for low emissions (Mansouri 2016).

### Efficiency implications

Attempts to mitigate noxious pollution from internal-combustion engines mostly lead to losses in efficiency. For example:

- 930 • Exhaust gas flow through catalytic converters and filters consume energy.
- The engines cannot approach their theoretical maximum efficiencies, because reducing NO<sub>x</sub> emissions is handled by limiting maximum combustion temperatures.

- 935
- NO<sub>x</sub> reduction catalysts require the presence of hydrocarbons, *i.e.* incomplete reactions, which implies that not all energy is extracted from the fuel.
  - Every treatment system added to the engine adds weight to the vehicle, which reduces payload and hence the total efficiency.

940 Before carbon dioxide pollution was a concern, it made sense to accept lower efficiencies as a necessary cost of reducing noxious pollution, but in a low-carbon future we cannot just keep on trading less noxious pollution for more carbon dioxide production.

### 1.3.5 Avoiding pollution from internal-combustion engines

945 ar: carbon-neutral) The efficiency and cleanliness of internal-combustion engines have dramatically increased over the last century, and more improvements are being implemented. But these improvements have not been fundamental to the engines in any way, and have been driven mostly by government regulation, at the cost of increased complexity and a higher purchase price.

It would seem obvious that it would be a good idea to introduce alternative technologies.

950 **Electrification**

In principle, there is nothing special about internal-combustion engines: they are not an end in themselves. They merely deliver a source of torque, which can be coupled to machinery to do useful work. Before the industrial era such torque was available from windmills and waterwheels, and today an alternative is the electric motor.

955 Electric motors are engines that use the interaction between electric current and magnetic fields to deliver useful torque to drive machines. Because they use electricity as a source of energy, they have no noxious emissions where they operate. Because they are not heat engines, their efficiencies are not subject to the Carnot limit, and efficiencies exceeding 95% are standard (Li and Curiac 2012).

960 Electricity, of course, is not necessarily carbon-neutral. Most electricity is generated in power plants that use fossil fuels as a primary source of energy. But because these plants are huge, and efficiency scales disproportionately with size, the energy output by the electric motor has a similar or lower carbon footprint than an equivalent internal-combustion engine (Doucette and McCulloch 2011).

965 Electricity grids are also increasingly being fed by solar and wind power, which are carbon-neutral at source. These renewable plants are also smaller and more flexible than their behemoth fossil-fuel counterparts, with lower capital costs and extremely low running costs.

Hence, in a low-carbon future, there is every reason to support or mandate the use of electrical motors for stationary applications wherever possible.

970 In automotive applications, *i.e.* in cases where the engine is used to move itself in addition to some form of payload, the application of electric motors is more demanding. In this case it is not easy to bring the electricity to the motor, although electric trains and buses fed by overhead conductors are splendid examples of the electrification of transport. So for electric vehicles to use the existing road network, they need to carry a source of electricity with them.

975 This source of electricity can be either a chemical battery, or a fuel cell. In a chemical battery power from the grid is stored in the form of reversible electrochemical reaction, and in a fuel cell the chemical energy from a fuel is directly converted

980 into electricity. There are fuel cells that can use hydrogen as a fuel, and fuel cells that can use methanol as a fuel. (It goes without saying that the hydrogen and the methanol need to be sourced from low-carbon sources for fuel cells to count as low-carbon energy sources.)

985 The storage and transport of hydrogen remain hurdles to the large-scale adoption of hydrogen-fuelled automobiles, although a market seems to be developing for hydrogen-fuelled electric trains (Agence France-Presse 2018). Hydrogen fuel cells emit no carbon or noxious pollution at point of use.

990 Direct methanol fuel cells can react methanol with atmospheric oxygen in an electrochemical cell to yield electricity, with carbon dioxide as a waste product. Little is known about possible noxious pollution.

995 At this time it seems that the electrification of road transport will be by chemical batteries. Lithium-ion batteries can now store enough energy and deliver enough power to make electric motor vehicles practical and attractive (Hayes et al. 2011), and some governments are considering plans to no longer allow the production of passenger vehicles propelled by internal combustion engines (Burke-Kennedy 2018) (Reuters 2018) (Gabbatiss 2018).

### Carbon-neutral fuels

1000 Another way to avoid the carbon pollution associated with internal-combustion engines is to change the fuel. Not all fuels are fossil fuels, and it is possible to use fuels that are carbon neutral, and in some cases carbon negative.

Apart from using hydrogen in a fuel cell, as described above, *hydrogen* can also be used as a fuel in Otto engines, because it will combust in air to yield heat. The emissions are water and NO<sub>x</sub>. Presently there are no hydrogen-fuelled Otto engines on the market.

1005 Methanol is a common product of the fossil fuel industry, but work is underway to produce methanol by reducing carbon dioxide using solar energy. Such *solar methanol* might be used in direct conversion fuel cells, or in internal-combustion engines.

1010 It is possible to harness the energy in contained in the reduced carbon in biological materials and use it as fuel. Such fuels are known as *biofuels*.

## 1.4 Biofuels

Biological processes are an integral part of the carbon cycle, because photosynthesis in plants reduces carbon dioxide in the atmosphere to sugars, which are converted by plant physiology into structural cellulose and other metabolites.

1015 The prototypical biofuel is of course wood, used in all societies for cooking and heating. This familiarity makes biofuels seem an obvious and viable source of energy, but details matter, and switching from fossil fuels to biofuels to reduce carbon footprints of human activities is not a simple choice.

1020 Firstly, the efficiency of photosynthesis is notoriously low: not above a few per cent, whereas the efficiency of a mass-produced silicon-based solar panel is in excess of 20%. In general it is much more efficient to capture solar energy in a PV panel and use an electric motor to provide the necessary mechanical power than to create a biofuel and use it to drive an internal-combustion engine.

1025 Second, increased biofuel production has numerous impacts on the environment and society, which cannot be ignored.

Discussing all the factors that need to be studied to make such a decision is outside the scope of this work, but as an example a report prepared for stakeholders in the Netherlands (Smeets et al. 2006) uses the following criteria:

- GHG [greenhouse gas] emissions – the use of biofuels should cause reductions of GHG emissions. The comparison should be done regarding the average use of fossil fuels, considering the life cycle of fossil and biofuels (i.e., well-to-wheel basis) and in case of biofuels reduction should be at least 30%
- Impacts over food supply – the production of biomass for energy must not endanger the food supply and other local biomass applications. The analysis should be developed considering possible changes of land use in the region of biomass production.
- Biodiversity – Biomass production must not affect protected or vulnerable biodiversity.
- The basic criteria are that violation of national laws and regulations are unacceptable.
- Local environmental effects – Principles include (a) soil and soil quality, that must be retained or even improved, (b) ground and surface water supply, that must not be polluted.
- Local economic effects – The production of biomass must contribute towards local prosperity.
- Social well-being – The production of biomass must not decrease the well-being of local societies.

Nevertheless, there are cases where using biofuels is a good option.

#### 1.4.1 Bio-gas

Anaerobic bacteria can convert biogenic carbon compounds to methane. This methane is identical to the methane obtained from natural gas, and can be used for the same applications.

An excellent application for the use of bio-gas is waste remediation. Waste material from the agri-food industry can be highly polluting if not treated properly, emitting noxious chemicals into water and the potent greenhouse gas methane into the atmosphere. But when used as feedstock for bio-gas production, it reduces carbon pollution by capturing methane and displacing natural gas, and also prevents water pollution (I. Venter 2014).

Bio-gas can be used to fuel Otto engines or gas turbines.

#### 1.4.2 Bio-ethanol

*(sec:BioEthanol)* Yeast and other microbes can convert sugar or starch from plants into ethanol. This technology is as old as civilization. Archaeologists may be in disagreement where beer was brewed first, but it is clear that between 3200 and 3000 BCE in ancient Sumeria the brewing of beer was an activity regulated by the government (Damerow 2012). By 2000 BCE in ancient Egypt the technology of fermentation was harnessed well enough that bakeries and breweries were co-located (*Model Bakery and Brewery from the Tomb of Meketre* 1920).

1070 Beer and wine can best be described chemically as aqueous solutions of sugars, ethanol, flavourants and colourants, with or without suspended solids. It was not until the 8th century CE that Persian and Arab scientists mastered the art of distillation and purified ethanol (Modanlou 2008), and it was not until Pasteur that yeast was seen as a living organism, and ethanol a product of its metabolism (Barnett 2000).

1075 It is unclear when ethanol was firsted as a fuel, but its flammability must have been noticed by the first distillers. By 1838 ethanol was common enough to be used in alcohol lamps as a source of heat in the chemical laboratory (Griffin 1838), and by the 1850s it was a major component of lamp oil used for illumination in the USA (Abebe 2008).

1080 Ethanol as a fuel for internal combustion engines has a clear start date: in 1826 Samuel Morley was granted a patent for an engine designed to use ethanol as fuel (Cummins 1989, p. 79).

1085 The industrial production of ethanol is a technologically advanced process. It is an active research field. A paper (Cardona and Sánchez 2007) reviewing the process technology of producing bio-ethanol written in 2007 had garnered 644 citations by January 2019.

The industrial production of ethanol consists of three steps.

1. Fermentation
2. Distillation
3. Dehydration

1090 The fermentation step is the biological process by which the yeast organism *Saccharomyces cerevisiae* convert the sugar and starch in the biological material to ethanol and carbon dioxide. This needs to be done in a sterile environment to prevent contamination by other micro-organism.

1095 The distillation step is the physical process of separating the produced alcohol from the aqueous mixture in which it is produced. This generally produces an azeotropic mixture that contains 95.6 % ethanol (Kumar, Singh, and Prasad 2010).

1100 While the processes of fermentation and distillation of ethanol for fuel is in principle the same as that of producing alcoholic beverages, the emphasis of the processes are very different. In the beverage industry the emphasis is on the development of complex flavours and a consistent, recognizable drinking experience. In the fuel industry the emphasis is on efficiency and throughput.

1105 To produce fuel-grade ethanol from the distilled azeotrope, dehydration is a necessary third step. This step can be implemented by various distillation processes that add a third compound, but in most modern plants the water is removed by adsorption onto molecular sieves (Kumar, Singh, and Prasad 2010).

### Brazilian ethanol

1110 Brazilian sugar-cane ethanol is an integral part of the country's energy network. Most Otto-engine vehicles have 'flex-fuel' engines, which can be fuelled with any blend of ethanol and petrol. The industry is regarded as sustainable ((Smeets et al. 2006).

### US maize

US maize ethanol is primarily blended with petrol to meet legislative requirements of oxygenates in fuels. The production of maize is water-intensive and the fermentation process is carbon-intensive. The US ethanol from maize has a poorer energy balance and larger carbon footprint than Brazilian ethanol from sugar cane, but has a smaller water footprint (Mekonnen et al. 2018).

Ethanol can fuel Otto engines and gas turbines.

### 1.4.3 Fischer-Tropsch fuel from biomass

(sec:FT) It is possible to heat woody biomass with steam in a low-oxygen environment to produce synthesis gas, which can be converted to a mix of fuel products in the well-known Fischer-Tropsch catalytic process. Such a fuel could be called a biofuel, but its production would be difficult to reconcile with the principles of *green chemistry*, described in Section 2.1.1.

### 1.4.4 Hydrotreated vegetable oil: “Green diesel”

(sec:GreenDiesel) The petroleum industry has developed a collection of chemical processes, such as hydrogenation, oxygenation and cracking. This collection of processes can be applied to vegetable oils to produce a fuel for diesel engines and gas turbines. This path is being followed by the aviation industry (Chiaramonti et al. 2014).

### 1.4.5 Biodiesel

Oils and fats have been used as fuels since antiquity, most obviously as a fuel for lamps: olive oil has been identified as the fuel used in lamps dating from around 600 CE (Kimpe, P. A. Jacobs, and Waelkens 2001), and at the Paris World Fair in 1900 Rudolf Diesel demonstrated an engine that ran on peanut oil (Knothe 2010).

During the energy crisis of the 1970s, the South African government looked at alternative sources of fuel. Experiments were done with sunflower oil in tractors, but there was a problem with the formation of carbon deposits around the injector nozzles. Seeing that the clogging of the injectors might be caused by the high viscosity of the sunflower oil as a fuel, the transesterification of the oil was implemented, replacing the glycerol in the sunflower oil with methanol. This created an oily liquid with a lower viscosity, which proved to be a trouble-free replacement (Van Niekerk 1980).

Such a transformed oil is termed *biodiesel*. It consists primarily of a mixture of fatty acid methyl esters, often abbreviated into the acronym FAMEs.

Compared to ethanol, biodiesel production is relatively simple, the main method sharing much with the ancient technology of making soap. This simplicity makes the production of biodiesel attractive to small and decentralized manufacturing, and consequently governments consider biodiesel production an attractive proposition: it can create job opportunities in rural population and it can create a stable market for farmers who produce vegetable oil crops.

Biodiesel is not yet carbon neutral: the methanol used to create the methyl esters is a product of the petroleum industry, and therefore biodiesel emissions contribute to global warming. But replacing fossil diesel fuel with biodiesel results in a nett reduction of carbon emissions, with the future possibility of replacing fossil-derived methanol with carbon-neutral bio-methanol (Shamsul et al. 2014).

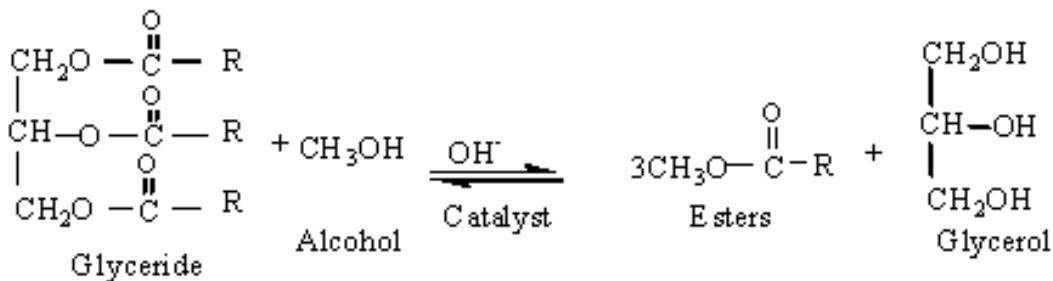


FIGURE 1.4: Transesterification

?(fig:Electron)?

## 1155 1.5 The significance of biodiesel and its quality control.

In a complex technical, economic and social environment it is likely that there will always be a need that can be best met by internal combustion engines. The decision on the type of engine and the decision on the fuel for that engine are not independent, as Cummins reminds us (Cummins 1989):

1160 “Our generation faces a similar challenge in a real liquid fuel energy shortage that will come within the lifetime now living. As we plunge into the seeking of solutions to our dilemma, we must never forget that *an engine and the fuel it consumes are inseparable partners*; the one cannot progress without the full cooperation of the other. This precept is vital to the planning of future powerplants, since an engine’s design determines 1165 its fuel and binds us to our future resource requirements.”

1170 While the variation-selection process by which real engineering progress is ‘blind’, i.e. the combination of factors that make a design successful is not known at the start of a development (Vincenti 1990), scientific knowledge provides guidance in the form of insight into processes and theoretically achievable targets for performance. Using the scientific knowledge we have today, we can risk a forecast: if society is determined that its carbon footprint should be reduced with minimal noxious pollution the application of internal combustion engines would tend towards the following:

- 1175 1. Larger: Because larger engines are more efficient than smaller engines, larger engines will be preferred. (See section 1.3.1)
2. Low-carbon: The preferred engine would be fuelled by carbon-neutral fuels. (Section 1.3.5)
3. Constant speed: Engines are most efficient when they can work at constant load. (Section 1.3.4)
- 1180 4. Efficient: Engines with high efficiency should be preferred, which implies engines with high compression or pressure ratios. (Section 1.3.2)
5. Cleanup: The exhaust should be amenable to cleanup, i.e. the exhaust gases of the fuel-engine system should not contain compounds that are incompatible with available converter or filter technologies. (Section 1.3.4)

From this it should be clear that the optimal engine of the future will not be an Otto engine, because it will always have a limited compression ratio and its catalytic exhaust cleanup will always require stoichiometric air:fuel ratios, which puts bounds to efficiency increases.

Because it is the most efficient engine, the gas turbine will play an important role. However, its exhaust temperature is very high, which begs for energy recovery to increase the efficiency of the system. Such systems are already seen in the *combined cycle gas turbine* (CCGT). But energy recovery systems are not light and small, so their best application outside aerospace (where the excess heat is utilized as thrust) appears to be stationary electricity generation.

For automotive applications the engine that will best contribute to reduce the carbon footprint of may be a large Diesel engine with exhaust cleanup to remove NO<sub>x</sub> and soot, fuelled with a carbon-neutral fuel.

In choosing between refined vegetable oil and biodiesel as a fuel, it is most likely that biodiesel will have a lower carbon footprint. Refinery operations are energy intensive, which might add to the carbon footprint of the fuel, and refineries are not usually near the point of use, so that transport will also add to its carbon footprint. The carbon emissions of biodiesel production are comparatively low. (At this point it is important to note that these observations are not definitive: carbon footprints are not predicted, they must be calculated.)

Buying large, highly efficient diesel engines with sophisticated management systems and exhaust cleanup require high capital investment. In an economically competitive environment they therefore need to bring reliable returns, which implies high availability, as measured by frequency of breakdown and length of time between maintenance stops. Such high reliability can only be achieved if the manufacturer understands the fuel-engine system well and it behaves predictably.

In a low-carbon future, therefore, well-characterized biodiesel might be expected to play a central role in non-electric ground transport and industrial application where electrification is not possible.

## 1.6 Conclusion: the role of chromatography.

The development of the fuel-engine systems depends heavily on the chemical characterization of fuels and engine emissions, and chromatography has always played a central role: some of the earliest researchers developing gas chromatography were employees of a petrochemical company (Keulemans, Kwanten, and Zaal 1955).

This thesis explores the possibility of applying comprehensive two-dimensional (supercritical fluid × gas) chromatography to the chemical analysis of biodiesel for the purpose of characterization and quality control.



$PV \neq nRT$

The ideal gas law does not apply to supercritical fluids.

# 2

## Introduction: Carbon dioxide and chromatography

1225

### ⟨Chapter2⟩ 2.1 The chemical industry

Industrialized societies depend on chemicals. (In this discussion I define chemicals as pure substances that are produced by industry for industry.) Chemicals might be used in the processing of products, or blended with other chemicals in formulations that might be sold to users as products. In a familiar example, sugar is a chemical produced by the sugar industry from sugar cane or sugar beet. It is a pure substance (sucrose) that is mostly used in the industry as an ingredient for processed food. Other uses of sugar include coatings for medication, and feedstock for engineered micro-organisms that produce pharmaceuticals. (Sugar is a rare example of a chemical that is also sold to consumers.)

The chemical industry produces a huge variety of products, from compounds as simple as hydrochloric acid to compound as sophisticated as cyanocobalamin (Figure 2.1). All of these chemicals help produce the products indispensable to industrialized societies. Nevertheless, the chemical industry is not held in high regard by people outside the industry. A part of this negative perception comes from the chemical industry's reputation for fatal accidents and pollution (Gumm 2015).

The list of incidents is long and examples easily spring to mind: In 1984 a leak at a chemical plant in Bhopal, India, caused the death of thousands of people and the injury of thousands more (R. Varma and D. R. Varma 2005). Stratospheric ozone is struggling to recover from depletion caused by the reckless emissions of chlorofluorocarbons (Ball et al. 2018). Plastics microparticles are now found everywhere in the oceans (Woodall et al. 2014), and the pesticide DDT is found in the breast milk of Inuit mothers (Gibson et al. 2016).

The chemical industry is also a prodigious producer of greenhouse gases. Apart from the carbon dioxide emitted by the production of the energy that power chemical processes, some chemical processes emit carbon dioxide as a waste product. Most notable of these is the reduction of atmospheric nitrogen as the first step in the production of nitrogen fertilizers. Of all the greenhouse gases monitored by the IPCC, only carbon dioxide, methane and nitrous oxide are found in nature: the others are exclusively products of the chemical industry (IPCC 2014).

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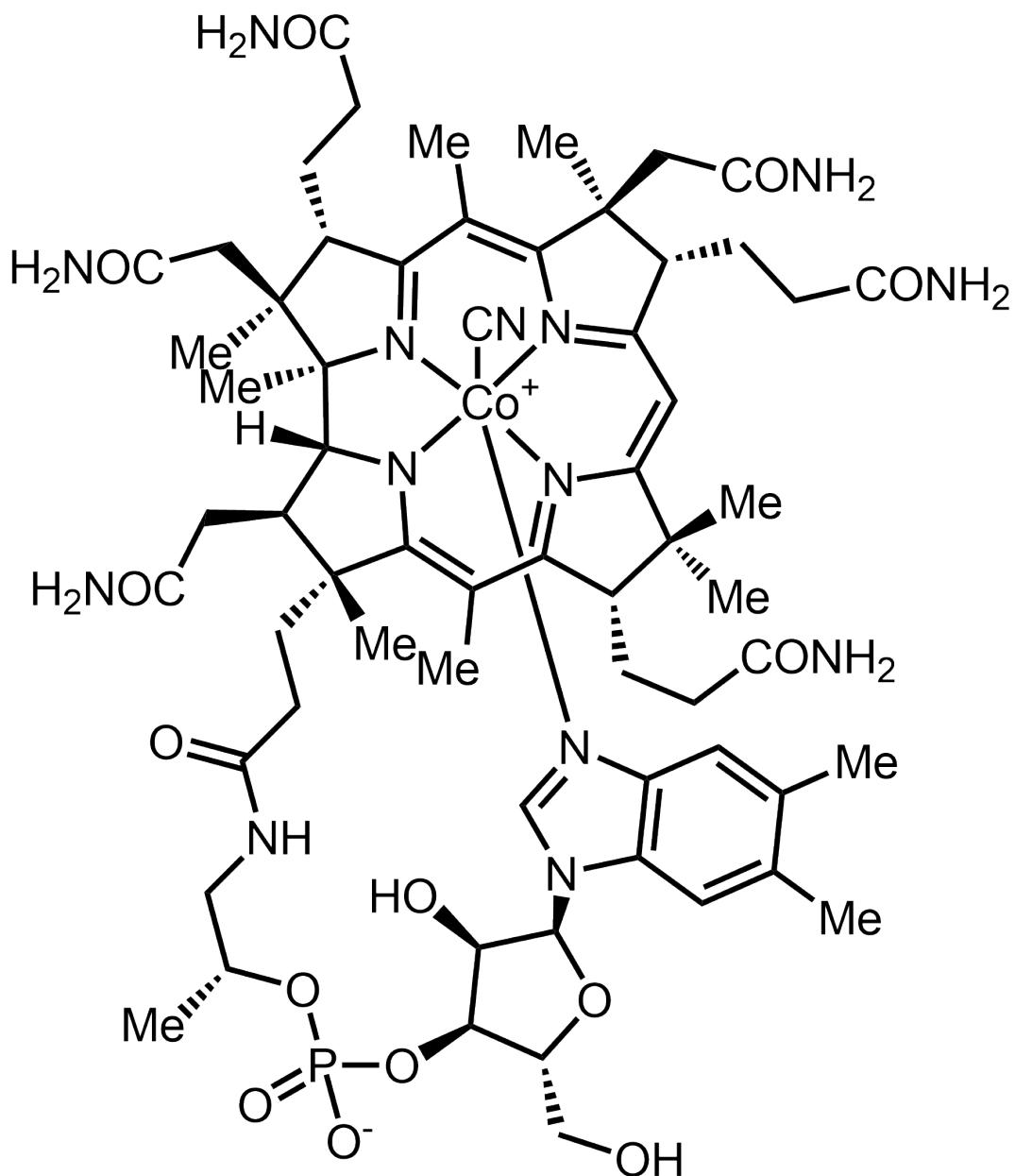


FIGURE 2.1: The chemical structure of cyanocobalamin, a form of vitamin B12. This compound is produced on the tonne scale by the chemical industry.

`(fig:vitb12)`

No chemical has ever jumped out of a lab, multiplied uncontrollably, and spread into the environment to poison or pollute: they have all been introduced into the environment by human ignorance, negligence, or recklessness. All chemicals behave well when used in properly controlled environments, but human actions can let them escape and damage and pollute. But while we try to solve the intractable problem of human behaviour, we as chemists cannot just lay blame: we must pay attention to the intrinsic safety of chemicals and chemical processes.

### 2.1.1 “Green chemistry”

The date of the birth environmental movement is conventionally set to 1962, when the biologist Rachel Carson published the book *Silent Spring*, which pointed out the destruction of nature by the unrestricted use of pesticides, and the dangers of overuse (Carson 2003). This was a direct imputation of the chemical industry, because the pesticide products contained many chemicals.

Chemists are human, and the realization uncontrolled that chemicals can have detrimental effects lead at least some chemists to reflect on their own work. This has given rise to the concept of *green chemistry*. Although the term has no rigorous definition or quantitative measure (Linhorst 2010), a set of 12 principles or guidelines are proposed:

- ?<(itm:waste)?  
1. It is better to prevent waste than to treat or clean up waste after it is formed.
- ?<(itm:incorporation)?  
1275 2. Synthetic methods should be designed to maximize the incorporation of all materials used in the process into the final product.
- ?<(itm:toxicity)?  
3. Wherever practicable, synthetic methodologies should be designed to use and generate substances that possess little or no toxicity to human health and the environment.
- ?<(itm:efficacy)?  
1280 4. Chemical products should be designed to preserve efficacy of function while reducing toxicity.
- <(itm:aux)  
5. The use of auxiliary substances (e.g. solvents, separation agents, etc.) should be made unnecessary wherever possible and innocuous when used.
- ?<(itm:energy)?  
1285 6. Energy requirements should be recognized for their environmental and economic impacts and should be minimized. Synthetic methods should be conducted at ambient temperature and pressure.
- <(itm:renewable)  
7. A raw material or feedstock should be renewable rather than depleting whenever technically and economically practicable.
- m:derivatization)?  
1290 8. Unnecessary derivatization (blocking group, protection/deprotection, temporary modification of physical/chemical processes) should be avoided whenever possible.
- ?<(itm:catalysts)?  
9. Catalytic reagents (as selective as possible) are superior to stoichiometric reagents.
- ?<(itm:persistence)?  
1295 10. Chemical products should be designed so that at the end of their function they do not persist in the environment and break down into innocuous degradation products.
- ?<(itm:monitoring)?  
11. Analytical methodologies need to be further developed to allow for real-time, in process monitoring and control prior to the formation of hazardous substances.

?<itm:safe>?

- 1300 12. Substances and the form of a substance used in a chemical process should be chosen so as to minimize the potential for chemical accidents, including releases, explosions, and fires.

1305 While these guidelines are clearly written with synthetic chemistry in mind, it does not mean that they do not apply to analytical chemistry. For example, Principle 7 suggests that, when possible, one should use hydrogen rather than helium as mobile phase in capillary gas chromatography: hydrogen is renewable, whereas helium is a finite resource and from time-to-time there are reports of shortages (Ko-rnblut 2019).

1310 One large area of the greening of chemistry is changing the use of solvents. Principle 5 recommends that solvents use be avoided, if possible — most solvents used in chemistry are ultimately derived from petroleum and are toxic to some degree. But solvents play a large role in many kinds of chemistry, and eliminating their use in the near future seems unlikely. Searching for and characterizing bio-derived, non-toxic, non-persistent solvents is an active field of research (Clarke et al. 2018).

1315 But there are a few solvents in already current use that are inherently “green”, such as water or ethanol.

One such naturally green solvent is carbon dioxide.

### 2.1.2 Carbon dioxide as a green chemical

Carbon dioxide as a chemical is used in industry in a few key areas.

- 1320 • Carbon dioxide is often used in firefighting, in the form of portable fire extinguishers or room flooding systems. In this last use it is displacing the ozone-depleting halomethane (Halon).
- 1325 • When liquid water is supersaturated with carbon dioxide, the gas desolvates slowly in the form of streams of tiny bubbles. This phenomenon makes beverages prepared from water supersaturated with carbon dioxide (or *carbonated water*) interesting to drink, and a large, international beverage industry is based on carbonated water.
- 1330 • Carbon dioxide has a freezing point of -77 °C, and the solid can be conveniently obtained by evaporating liquid carbon dioxide at atmospheric pressure. The evaporating liquid rapidly cools the stream of carbon dioxide, lowering the temperature of the stream to below the freezing point, and the gas crystallizes into the solid. The resulting ‘snow’ can be compressed into blocks, which only slowly sublimes into gaseous carbon dioxide, keeping the temperature at the freezing point. Packing frozen food products together with this ‘dry ice’ allows for it to be transported cold.
- 1335 • Pellets of dry ice can be entrained in a jet of air, and used to abrade surfaces for cleaning (Spur, Uhlmann, and Elbing 1999). This use of carbon dioxide can displace toxic solvents or abate the noxious dust produced by blasting operations.
- 1340 • Carbon dioxide is a ‘natural refrigerant’ (Pearson 2005), and can be used to displace hydrofluorocarbon refrigerants, which are potent, long-lived greenhouse gases.

- 1345 • Carbon dioxide can be used as a preservative and anti-oxidant in packaged food. If air in a packaged food is removed by purging the headspace with carbon dioxide, the growth of microbes can be discouraged, extending the shelf life of the product (Jacobsen and Bertelsen 2002).
- Carbon dioxide can be used to extract compounds from natural products.

Of these uses, extractions are economically the most important.

## 2.2 Extractions using carbon dioxide

### 2.2.1 Commercial extractions

1350 There are several commercial processes that use carbon dioxide to extract valuable products from plant material.

#### Plant oils

Vegetable oils are obtained from various crops, and can be extracted from the plant material by pressing, heating or extraction. High-pressure carbon dioxide has been  
1355 used to extract vegetable oils from various plants, although it seems that this process has only found niche applications so far (Eisenmenger et al. 2006; Grażyna and Anna 2018).

#### Hops

Hops is an essential component in the brewing of beer. It imparts a desired bitter  
1360 flavour, stabilizes the beer during storage, and assists with foam formation (Schönberger and Kostelecky 2011). Hops is a seasonal crop with a limited growing range, but the demand for beer is not limited to certain areas or seasons. The creation of hops extract makes it possible for brewers to benefit from hops without owning a hops plantation or storing and transporting dried hops over long distances. All  
1365 hops extracts produced today are extracted by carbon dioxide (Hunt et al. 2010).

#### Coffee

Coffee is an international industry, with coffee drunk in many cultures and in many forms. One of the attractions of coffee is the effects of the psychoactive substance found in the coffee bean, *caffeine*. Caffeine is a mild stimulant and promotes wakefulness. A small proportion of coffee drinkers enjoy drinking coffee, but prefer to avoid the stimulant effect, which might induce insomnia or irritability. For these  
1370 coffee drinkers the market supplies *decaffeinated coffee*.

Given the large amount of coffee traded (an estimated 167.47 million bags of coffee in the 2018-2019 coffee year (*Coffee Market Report December 2018 2018*))<sup>1</sup>, if only  
1375 a small percentage of coffee needs to be decaffeinated, it will be a large amount of coffee to process, and industrial processes will be necessary to supply the demand.

Decaffeination of coffee is achieved by selectively extracting the caffeine from green (*i.e.* unroasted) coffee beans using carbon dioxide. This is the largest use of carbon dioxide for extraction (Ramalakshmi and Raghavan 1999). The extracted  
1380 caffeine is sold for use in medication and 'energy' drinks.

<sup>1</sup>The factoid that "coffee is the second-most traded commodity after oil" has been proven to be untrue (Greenberg 2017).

## 2.2.2 Analytical Extractions

Extraction is not only an industrial process, but is part and parcel of analytical chemistry. The first extractions using carbon dioxide was not aimed at developing an industrial operation, but to develop methods for analytical chemistry. This method is usually called SFE, for supercritical fluid extraction.

## 2.2.3 Why carbon dioxide?

But what makes carbon dioxide a suitable solvent for extraction?

There are two aspects to this question. The first is about the *greenness* of carbon dioxide. It is non-toxic, non-persistent, non-flammable, non-corrosive, inexpensive, commercially available, and a waste product. (It goes without saying that this carbon dioxide is sourced from a carbon-neutral source, perhaps the brewing industry.)

The second aspect of the desirability of carbon dioxide lies in its physical properties and the conditions under which we use it.

Chemists will intuitively understand that gaseous carbon dioxide has no solvating properties, and that liquid carbon dioxide should not behave much differently than any other solvent. Both these statements are true under 'normal' circumstances.

But consider the case of an isobaric cooling of a volume of gas. The gas-liquid transition takes place because energy is removed from the system. At some point the kinetic energy of some of the molecules becomes less than the energy of the intermolecular forces, and the molecules prefer to clump together, *i.e.* it condenses. The remaining gas molecules receive the excess energy, and therefore stay in the gas state, until more energy is removed, leading to more of the gas condensing. During this process the temperature remains constant, and this temperature is known as the boiling point.

Now consider a solute (solid or liquid) in the same volume of gas being cooled. In this case, as the gas cools, the gas-solute intermolecular forces can become stronger than the gas-gas intermolecular forces at a temperature which is higher than the boiling point. In such a case the gas molecules will clump around solute molecules, while the kinetic energy of the gas molecules are still too high to allow condensation. Now the gas has obtained solvating properties and the solute will become truly dissolved in the gas.

The same effect can be imagined to happen during the isothermal compression of a gas.

If there are more than one solute in the volume of gas, some might dissolve in the gas, while others one might not. This means that the solvating gas can be *selective*. It can also be seen that the solvating power of the gas will depend on the temperature and the pressure of the gas. This means that the solvent becomes *tunable*.

While the dense gas has solvating properties, it still has the physical properties of a gas:

**1420 Diffusivity** The solvating gas maintains its low diffusion coefficients, which means that it can easily diffuse into porous material, and that solutes will rapidly diffuse through it.

**Surface tension** The solvating gas has a low surface tension, which means that it will readily 'wet' surfaces and penetrate porous material.

**1425 Viscosity** The solvating gas has a low viscosity, which means that it takes little energy to pump it.

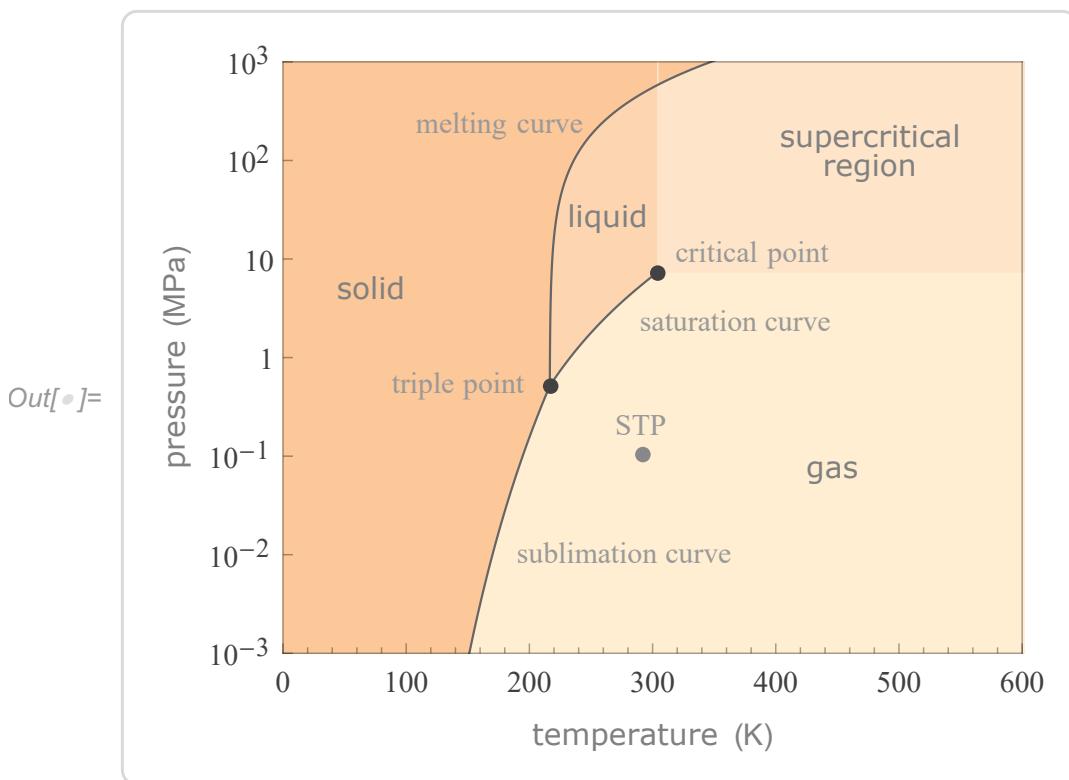


FIGURE 2.2: The phase diagram of carbon dioxide

&lt;fig:co2phase&gt;

For historical reasons such solvating gases are known as a 'supercritical fluids', because they were first obtained by heating a liquid at high pressure (Berche, Henkel, and Kenna 2009), so that the temperature and pressure of the substance is higher than its *critical point*.

<sup>1430</sup> The critical pressure of a substance is the pressure above which it is impossible to create a gas-liquid phase transition by isobaric cooling, and the critical temperature is the temperature above which it is impossible to create a gas-liquid phase transition by isothermal compression. When the gas is at its critical temperature and critical pressure, it is at its critical point. The critical point is very

<sup>1435</sup> different from the *triple point*: there is no equilibrium involved (See figure 2.2). The terms 'supercritical fluid' and 'dense gas' are synonyms (Randall 1982) — the adjective 'dense' in 'dense gas' of course implies that the gas behaviour is far from that of an ideal gas.

There are no discontinuities in binary diffusivity when temperature is changed from above to below the critical temperature (Lauer, McManigill, and Board 1983). Hence extractions are often done under conditions which are not technically 'supercritical' but still yields its benefits.

<sup>1440</sup> The critical pressure of carbon dioxide is 304.12 K (31.10 °C), and the critical pressure is 7.39 MPa (72.9 atm). This temperature and pressure are easy to achieve in the laboratory with standard chromatographic instrumentation, or in an industrial plant with standard process engineering technologies.

Carbon dioxide is gaseous at ambient conditions. This means that once it has been used in its role as extractant, and it is exposed to the atmosphere, it will rapidly evaporate, without needing added heat, and leaving no residues.

1450 Any pure compound will have a supercritical point. Compounds with technologically feasible supercritical points and useful chemical properties include ammonia, methanol, CFCs/Freon, hydrocarbons (propane, butane), water and sulfur hexafluoride. All of these lack green attributes: hydrocarbons pollute, the CFCs deplete ozone, sulfur hexafluoride is a potent greenhouse gas, and methanol and water are liquid at ambient conditions.

1455 For these reasons the term supercritical fluid extraction is practically synonymous with extraction using high-pressure carbon dioxide.

It is also possible to use supercritical fluids as reactants. This topic falls outside the scope of this discussion.

#### 1460 Modifiers

(sec:modifiers) The carbon dioxide molecule has zero dipole moment and the bulk fluid a low dielectric constant, so supercritical carbon dioxide is expected to be a non-polar solvent. (Although in practice the solvent behaviour is more complex, partly explained by the high quadrupole moment of the carbon dioxide molecule (Raveendran, Ikushima, and Wallen 2005).)

1465 While the solvating power of a supercritical fluid is certainly ‘tunable’ by adjusting its pressure and/or temperature, the range in solubility might be quite limited in practice. Just as with other solvents, it is possible to add a co-solvent or *modifier* to the supercritical carbon dioxide. This makes it possible to increase the solubility of polar compounds in the supercritical fluid. Methanol, ethanol, formic acid, and water are examples of suitable green modifiers for carbon dioxide (Herrero et al. 1470 2010).

1475 When modifiers are used the carbon-dioxide, modifier and solute forms a system with four degrees of freedom (modifier percentage, solute concentration, pressure, and temperature), which can become difficult to model. While this is a challenge for process engineers who need to design efficient industrial systems, analytical chemists can afford to be pragmatic and use heuristics to find suitable conditions (Wells, Zhou, and Parcher 2003).

#### Practical extractions

1480 Figure 2.3 shows a schematic diagram of a system set up for supercritical fluid extractions.

Carbon dioxide (a) is readily available from suppliers of industrial gases, and high-purity grades are available. Chromatography-grade solvents are usually used as modifiers (d). High-pressure pumps (HPLC type) are used to compress the carbon dioxide (b) and the modifier (c), which are mixed together at the appropriate ratio. The mixture gets pumped into the (optionally heated (f)) extraction cell (e), which contains the material that needs to be extracted. Having extracted the extract from the material, the supercritical fluid passes through a pressure-control mechanism (g). This allows the pressure of the supercritical fluid to drop to ambient, turning it into a low-density non-solvating gas. The extract becomes desolvated, and precipitates in the collection vessel (h). The operation of the system might be either static or dynamic: in static operation the supercritical fluid is pumped into the system, the flow is stopped, and the matrix/fluid mixture is given time to approach equilibrium. Then the fluid is expelled and the extract collected. In dynamic operation the supercritical fluid is pumped through the extraction cell and the extract collected continuously.

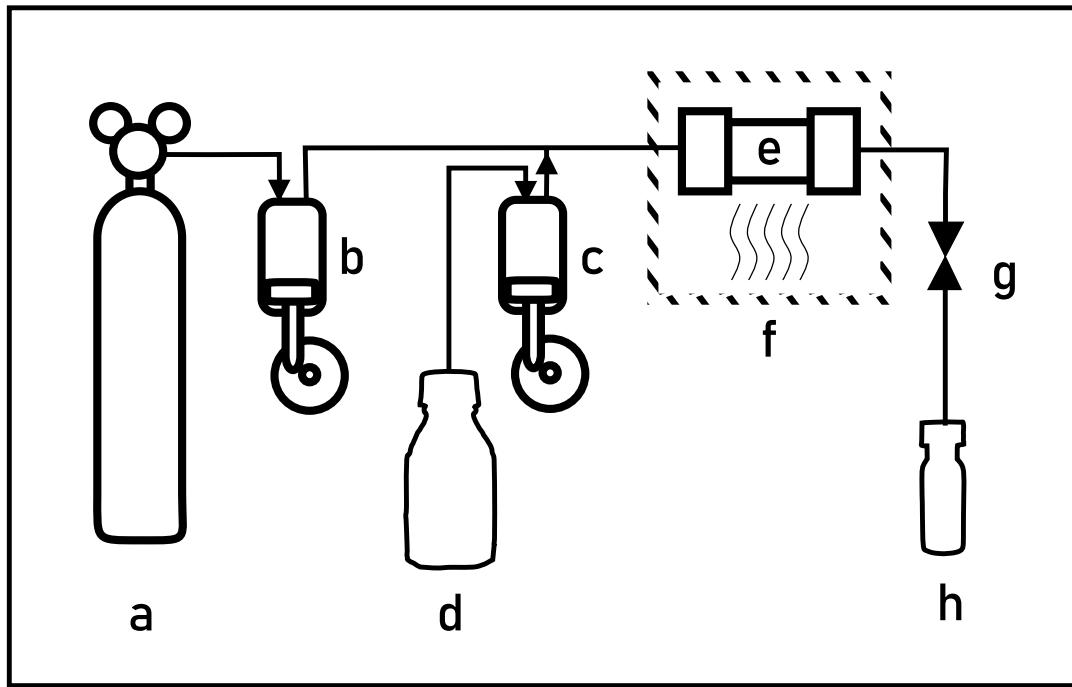


FIGURE 2.3: A diagram of an SFE system. (a) CO<sub>2</sub> supply (b) High-pressure SF pump (c) High-pressure modifier pump (d) Modifier reservoir (e) Extraction cell (f) Pressure control (g) Collection vessel

(fig:sfediaagram)

## 2.3 Supercritical Fluid Chromatography

An analyte will extract out of a matrix into a given solvent with a certain efficiency and at a certain rate. While this is important while finding an optimum extraction method, otherwise its relevance is limited.

However, *different* analytes will extract out of a matrix with *different* efficiencies and at *different* rates. In a 1906 paper the Russian botanist Tsvet applied this observation to the dynamic extraction of a bed of calcium carbonate that had mixture of plant pigments applied at the inlet end, using petrochemical solvents. The different extraction efficiencies and rates of adsorption and desorption on the calcium carbonate surface lead to the *separation* of the compounds in the mixture. Tsvet called this method of separation *chromatography* (Ettre and Sakodynskii 1993a; Ettre and Sakodynskii 1993b). With time this method became generalized, and today chromatography is a major, established scientific field with many ramifications and a myriad of applications.

Because of the different technologies used in its applications, chromatography is conventionally classed by the state of its mobile phase as either *gas chromatography* (GC) or *liquid chromatography* (LC). However, as we have seen, solvating gases can also extract analytes from solid stationary phases, and hence the term *supercritical fluid chromatography* (SFC) was created for these kinds of separations.

Supercritical fluid chromatography as practised today bears a great resemblance to *high performance liquid chromatography* (HPLC). The main reason for this is that there is a large overlap between the technology used for HPLC and the technology needed for SFC. In particular, both use high-pressure pumps and columns packed

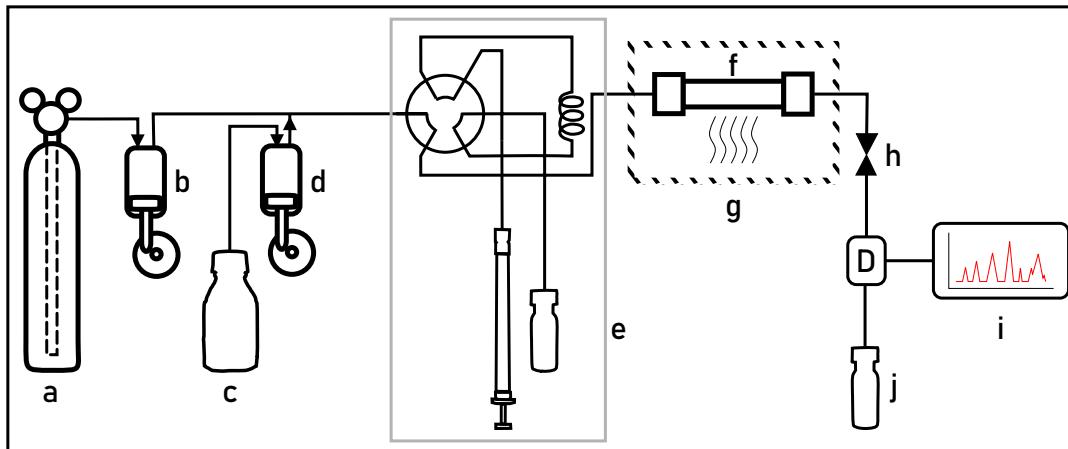


FIGURE 2.4: A diagram of an SFC system. (a)  $\text{CO}_2$  supply (b) High-pressure SF pump (c) Modifier reservoir (d) High-pressure modifier pump (e) Injection system (f) Column (g) Optional heating (h) Back-pressure control (D) Detector (i) Data system (j) Optional fraction collection.

?(fig:sfcdiagram)?

1520 with particles of very small diameter, and use optical detectors. The same instrument manufacturers who supply HPLC instrumentation also supply SFC instrumentation.

### 2.3.1 SFC and FID

1525 But SFC was not always a technique based on HPLC technology. In the 1980s SFC was practised using open tubular columns and flame detectors, so the instrument designs looked more like GC instruments than HPLC instruments, and SFC was sold as a replacement for GC (Poole 2003).

During this era the detectors used for SFC were the flame-based detectors used for CG, in particular the flame ionization detector (FID).

1530 The *flame ionization detector* was invented near-simultaneously in South Africa and New Zealand (Ettre 2008). The core of the system is a flame of hydrogen gas burning in pure air. The measured signal is the conductivity of the flame plasma, which is measured by applying a -100 V potential difference between electrodes at the tip and the base of the flame. There are very few free ions in the hydrogen flame, 1535 so the conductivity is normally low. But organic compounds introduced into the flame creates a number of free ions, which increases the conductivity of the flame gases. The change in conductivity is measured by measuring the current between the two electrodes, using an electrometer.

1540 As a first approximation the signal produced by the FID is proportional to the number of carbon atoms in the analyte. This is rather surprising, until the mechanism of its working is elucidated. At high temperature in the hydrogen-rich core of the flame, all hydrocarbon atoms are reduced to methane (Holm and Madsen 1996). Once it gets into contact with the hot oxygen in the outer layers of the flame there is a chemi-ionization reaction. The electric field acting on the ions creates a flow of ions, which can be measured as an electric current.

1545 The FID's shining strength as a detector is its tremendous linear dynamic range of  $10^7$ . Combining the linear dynamic range with its carbon sensitivity makes it an excellent detector for quantifying organic compounds.

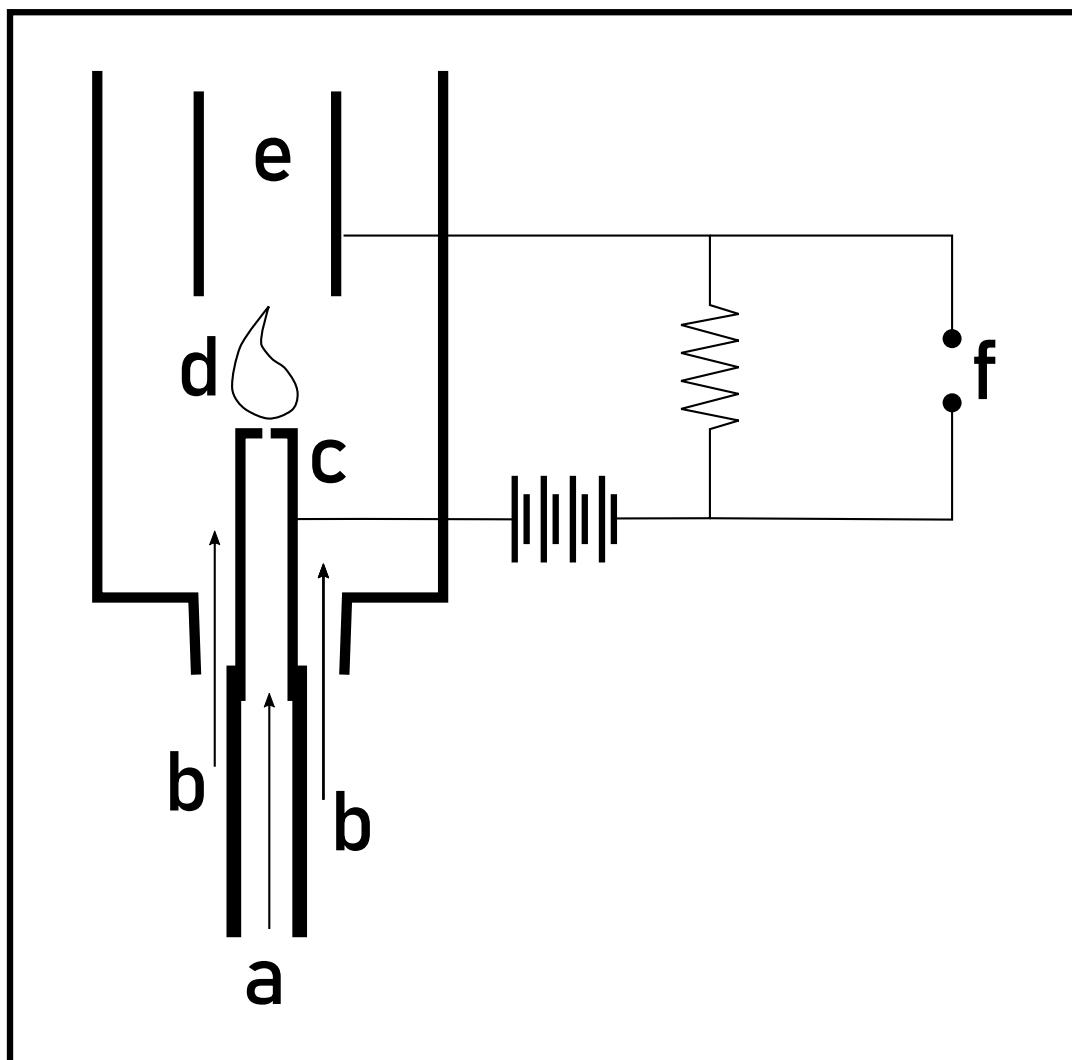


FIGURE 2.5: A schematic diagram of the flame ionization detector. (a) Mixture of column eluate and hydrogen gas (b) Clean air (c) Metallic flame tip electrode (d) Collector electrode (e) Conductivity signal

?<fig:fiddiagram>?

During period of history that SFC was seen as a variant of GC, the FID was the detector of choice. But when SFC started looking like HPLC, and the selectivity of the chromatography started being manipulated by adding modifiers (see Section 2.2.3), the FID lost its utility. The quantity of modifier added to the carbon dioxide mobile phase swamps the FID detector, making it useless as a detector. In contrast, if the chosen modifiers are transparent at the relevant wavelengths, optical detectors are more useful in SFC using modified carbon dioxide as a mobile phase.

### 2.3.2 SFC $\times$ GC

(sec: SFC $\times$ GC)

In chromatography, the analyte in the eluate can be detected to yield information about the sample, or fractions of eluate can be collected for further purification or characterization, or both. If fractions are collected, there is nothing that prevents the chromatographer from subjecting a collected fraction to another chromatographic separation.

For example, a synthetic chemist might use column chromatography to separate their target compound from side-reaction compounds, collecting fractions of the eluent. If the compounds are colourless, the chemist might use *thin-layer chromatography* (TLC) to examine the fractions for presence and purity of the relevant compounds.

If the product and the side-product were co-eluting it would not help to use the same chromatographic system (*i.e.* combination of stationary phase and mobile phase) for the TLC examination, because the co-elution would not become apparent. If, however, the chemist changed the mobile phase, or switched to a different stationary phase for the TLC examination, then it is quite likely that the two compounds would be separated.

For example, if a synthesis involving sugars were being attempted, two sugar enantiomers would likely co-elute on a silica clean-up column. Investigating the sugar fraction using a TLC plate with a chiral stationary phase would reveal the presence of stereoisomers.

The second, chiral separation (on TLC) is said to be *orthogonal* to the first (packed column) separation.

Such a separation is an example of a two-dimensional (2D) separation. In analytical chromatography, 2D separations are powerful tools.

The simplest 2D chromatography is called *heart-cutting*. In such a case a fraction of interest is collected from the first dimension ( $^1\text{D}$ ) separation, and subjected to a second dimension ( $^2\text{D}$ ) separation. For example, a fraction collected from an HPLC separation could be injected into a gas chromatograph. The separation on the GC dimension would obviously be very different from the separation on the HPLC dimension, and therefore we can say that the orthogonality is high.

Heart-cutting is a useful way to get detailed information from a sample, but it is technically demanding and labour-intensive. It is usually employed for challenging separations in well-understood samples, because peaks of interest in the first dimension must be captured for injection on the second dimension.

If, however, one collects the eluate in equal fractions and do identical separations on each of the fractions, then one enters the domain of **comprehensively coupled chromatography**. The comprehensive coupling of orthogonal chromatographies does not demand a previous understanding of the sample, and doing exactly the same separation for each fraction allows the process to be automated.

A 2D separation can be called comprehensive if the following three criteria are met (Giddings 1987):

1. Every part of the sample is separated by two distinct chromatographic processes.
2. Equal percentages of all sample fractions are separated by the second process.
- 1600 3. Compounds resolved by the first dimension separation remains resolved.

To effectively distinguish between heart-cut and comprehensive 2D couplings, a terse nomenclature was developed. Simple coupling between systems is designated by a hyphen ("‐"), for example HPLC-GC. This corresponds to the usual notation of a coupled detector coupled to a chromatograph, for example GC-FID or GC-MS. Comprehensive coupling is denoted by a multiplex sign ("×"), for example GC×GC. If it is not clear from the context, the detector can be specified using a hyphen, as in GC×GC-MS.

Of the comprehensive coupled chromatography techniques, GC×GC is the most mature, with a selection of powerful instruments on the market.

1610 This thesis further explores the implementation of SFC×GC, as first developed by Venter and Rohwer (Andre Venter and Egmont R. Rohwer 2004; Andre Venter, Makgwane, and Egmont R. Rohwer 2006).

In this implementation of SFC×GC, fractions of the eluate of the SFC is transferred to the GC. The mobile phase is changed from carbon dioxide to hydrogen 1615 in the modulating interface, and the GC dimension is a conventional open-tubular capillary separation with FID detection.

Any volatile modifiers or additives added to the SFC mobile phase will be separated from the analytes by the GC dimension. This makes it possible to use the FID as a detector in an SFC×GC chromatograph where the SFC is based on HPCL 1620 technology.

The <sup>2</sup>D separations are *fast* chromatographic separations, *i.e.* separations optimized for speed, sacrificing resolution. This is achieved using fast temperature programming of the GC separation.

The high orthogonality of the SFC and the GC separations enables novel comprehensive 2D chromatography, which we applied to the chemical analysis of biodiesel. 1625



*The nice thing about standards is that you have so many to choose from.*

Andrew S. Tanenbaum

# 3

## Biodiesel, technical standards and chemical analysis

### ⟨Chapter3⟩ 3.1 Introduction to standards

(Sec: Intro)  
1630 The fact that vegetable oil can be used to fuel Diesel engines has been known since the earliest days. Rudolf Diesel himself had exhibited an engine at the Paris Exhibition in 1998 (Knothe 2010) that ran on peanut oil. But the development of the petroleum industry late in the 19th century ensured an ample supply of fossil fuel for these engines.

1635 The development of the diesel engine happened in parallel with its fuel, for, as Cummins said "...we must never forget that *an engine and the fuel it consumes are inseparable partners*; the one cannot progress without the full cooperation of the other" (Cummins 1989). The first invention of an engine presupposes a supply of fuel, but the variation-selection process that searches for lower costs and higher efficiency  
1640 then opens up the quest for more fuels. If a possible fuel is found, meticulous engine builders would then have to test each new fuel in their engines and approve of it. But fuel suppliers would like to see their fuels used in as many engines as possible. This convergence of interests gives rise to the establishment of *technical standards*. Technical standards allow engine builders to develop engines that will run on any  
1645 fuel with certain agreed-upon standard qualities, and fuel producers can produce fuels knowing that they will work on any engine designed for that fuel.

A technical standard or just *standard* is a "document, established by consensus and approved by a recognized body, that provides, for common and repeated use, rules, guidelines or characteristics for activities or their results, aimed at the achievement  
1650 of the optimum degree of order in a given context" (Hatto 2010). Standards are used in many aspects of industrialized societies, from implantable medical devices (ISO 2019) to coffins and caskets (SABS 1993). Standards are often associated with products, but also apply to procedures (ISO 2015) and systems (ISO 2017).

Standards are not mandatory, nor do they provide the 'best' way of doing something.  
1655 The great strength of standards is their reliability. "Standards exist principally to provide a reliable basis on which common expectations can be shared regarding specific characteristics of a product, service or process" (BSI 2016).

TABLE 3.1: A few well-known standards organizations

Standards Organizations	Abbreviation	Name	Country of origin
SABS		South African Bureau of Standards	South Africa
ISO		International Organization for Standardization	International
CEN		European Committee for Standardization	Europe
ASTM		ASTM International	USA
BSI		British Standards Institution	UK
IEC		International Electrotechnical Commission	International
DIN		German Institute for Standardization	Germany
ANSI		American National Standards Institute	USA
UL		Underwriter's Laboratory	USA
ITU		International Telecommunication Union	International

1660 Standards are published by *standards organizations*, which might be national or international in character, or might be established to serve a certain industry. Standards organizations are often known by their abbreviations, and a few of them are listed in Table 3.1. The authors of these standards documents are usually **technical committees**, comprised of individuals from a wide variety of stakeholder organizations, who work towards consensus.

1665 Standards have a unique publication system. They are not published by publishing houses, but by the standards organizations themselves, who sell the documents directly to end users. Each standard is usually also better known by its number than by its title. For example, if I were to mention the document entitled "Quality management systems — Requirements" few people would know that I'm talking about the well-known quality system standard usually known as ISO 9001.

1670 The South African national standards body is the South African Bureau of Standards (SABS), which was established by an act of parliament, Standards Act, 1945 (Act No. 24 of 1945)), as amended by the Standards Act, Act No. 8 of 2008 (South African Government 2008). The SABS issues South African National Standards.

1675 Standards organizations not only write standards, they might also **adopt** them. Adoption happens when a suitable standard has already been issued by another standards organization. Standards very often refer to other standards, and standards are often based on published research. While standards are not mandatory, some legislation might refer to standards.

1680 The desire for engine designers for access to a reliable fuel and for fuel suppliers to have the largest possible market lead them to cooperate in the development of standards for fuels. In South Africa the relevant standard for petroleum-based diesel fuel is SANS 342, and in the USA the equivalent is ASTM D975. According to the Petroleum Products Act of 1977, as amended, diesel fuel sold to an end-consumer must conform to SANS 342

1685 Biodiesel is chemically very different from **petrodiesel** (diesel derived from fossil sources), and therefore the technical standards of biodiesel need to be different from the standards for petrodiesel.

## 3.2 SANS 1935: An overview

The current South African standard applicable to biodiesel is “South African National Standard 1935 Automotive biodiesel — Fatty Acid Methyl Esters (FAME) for diesel engines — Requirements and test methods” (SABS 2008). This is an adoption of the European Committee for Standardization’s (CEN) standard EN 14214 (CEN 2008). In the USA is the equivalent standard is ASTM D 7651, which is largely similar but of different heritage.

SANS 1935 consists of 18 pages. The first two pages are unnumbered: for the purposes of this discussion they will be numbered in small Roman numerals.

p(i) The first page is a title page, following the usual format for SABS standards. The top line of the page contains the ISBN (978-0-626-26349-2), and in large type the standard number (SANS 1935:2011). Then follows in capital letters “South African National Standards”, and below that the title “Automotive biodiesel — Fatty Acid Methyl Esters (FAME) for diesel engines — Requirements and test methods”. At the bottom edge of the page we find the SABS logo and some contact information.

p(ii) The second page is an informational page. It starts with a table of changes, which was still empty at the time of writing. Then follows a foreword, in which the technical committee who approved the standard is acknowledged (National Committee SABS SC 1018A). It also gives the date of publication (December 2011) and states that it supersedes SANS 1935:2004. Then there is a very significant line, which states that the standard is referenced in the Petroleum Products Act (South African Government 1977).

p1 Contains the Table of Contents

p2 Is left blank

p3 Paragraph 1: This paragraph describes the scope of the standard, which is that it applies to biodiesel as an automotive fuel.

p3 Paragraph 2: This paragraph lists all the normative standards required to comply with SANS 1935. Thirty-five standards are listed.

p4 Continues the list of normative references

p5 Paragraph 3: This paragraph contains a list of definitions. The most important one is this: “biodiesel [is a] fuel comprised of methyl esters of long chain fatty acids derived from vegetable oils.” This is a very specific description of biodiesel. It excludes animal fats as a source of fatty acids, and it excludes ethyl esters. But a note reads “Consideration for the inclusion of ethyl esters, animal fats and C8 – C12 carbon chains should be taken later.” The significance of ethyl esters is that methyl esters are not usually “carbon neutral”, that is, the methanol used in the transesterification reaction is usually obtained from the petrochemical industry, whereas ethanol from fermentation (See Section 1.4.2) could be carbon neutral. The definition also rules out hydrotreated vegetable oil (see Section 1.4.4) or biomass-derived Fischer-Tropsch diesel (see Section 1.4.3).

p5 Paragraph 4: This paragraph lists requirements

- <sup>1730</sup> **p5** Paragraph 4.1 discusses general requirements. According to these requirements biodiesel is a homogeneous liquid, free of adulterants or contaminants, to which additives might be added. It provides details regarding testing for oxidative stability and cold-flow properties
- <sup>1735</sup> **p6** Paragraph 4.2: This paragraph is about physical and chemical properties and states that biodiesel shall comply to the requirements of Table 1
- <sup>1740</sup> **p7** Paragraph 4.3: This paragraph concerns methods of test. It states that test methods shall be one of the test methods listed in Table 1.
- p7-p8** These pages contain Table 1.
- p9** Paragraph 5 describes a few simple rules for packing and marking biodiesel.
- <sup>1745</sup> **p10-p11** Annex A describes a method for the calculation of iodine value from chromatographic data. This calculation might be used instead of a direct measurement of the iodine value.
- p12** Annex B prescribes how samples for testing must be taken.
- p13** Annex C gives a list of values for calculating precision.
- <sup>1750</sup> **p14** Annex D provides an approved method for correcting density measurements. The prescribed tests require density to be measured at 15 °C, which might be inconvenient. Instead, a different test may be made at a more convenient temperature, and a correction applied.
- p15** Annex E is informative and recommends the implementation of quality management systems. It is followed by a bibliography.
- <sup>1755</sup> **p16** The final page contains information about the SABS and its services.

### 3.3 SANS 1935: Properties, requirements and methods.

On studying SANS 1935, it quickly becomes clear that the core of the document is Table 1. This table has three columns. The first column lists a **property**, the second column specifies a numerical value the property has to conform to, the **requirement**, and the third column prescribes the **test method** that must be used to obtain the value.

In the following discussion each of the requirements will be discussed, in order of increasing relevance to this thesis, grouped by method of determination (See Figure 3.1).

### <sup>1765</sup> 3.4 SANS 1935: Physical properties

#### 3.4.1 Density

A diesel engine can only deliver work proportional to the heat of combustion of the fuel, and for a given chemical composition the energy will be proportional to the

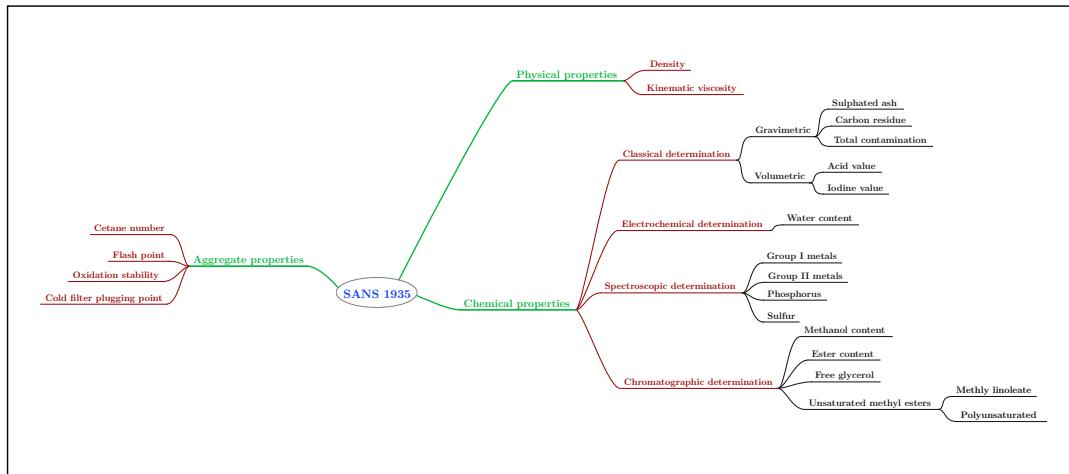


FIGURE 3.1: A mind map showing how the properties of biodiesel assessed by SANS 1935 is grouped for the purpose of discussion in this chapter.

*(fig:MindMap)*

mass of the fuel. But diesel engines inject measured volumes of fuel, not measured masses. (Fuel is also sold by volume, not mass.) Therefore, an engine designer must specify the density of the fuel to specify the power output of the engine.

According to SANS 1935 the density of biodiesel is required to be  $860 \text{ kg m}^{-3}$  to  $900 \text{ kg m}^{-3}$  at a temperature of  $15^\circ\text{C}$ . The testing methods that can be used are described in ISO 3675 and ISO 12185.

ISO 12185 prescribes the measurement of density by using an electronic instrument known as the oscillating tube density meter. This measures the density of a liquid by measuring the frequency of a freely-oscillating tube filled with the liquid under test. The frequency depends on the mass of the filled tube, and therefore on the density of the liquid. These devices are easy to use and very accurate. The temperature of the liquid is controlled electronically.

ISO 3675 prescribes the use of a hydrometer, which is an instrument that measures density by measuring the buoyant force on a floating indicator. According to the law of Archimedes, the buoyant force on a body immersed in a liquid is equal to the weight of the displaced liquid. If the liquid is denser, the force is greater. Therefore, in a denser liquid a floating indicator will float with more of the indicator above the surface of the liquid. Hydrometer technology is mature, and the devices are simple and robust. If ISO 3675 is used at a temperature other than the specified one, a temperature correction is applied, as described in Annex D.

### 3.4.2 Kinematic viscosity

The viscosity of a fluid is a measure of its resistance to flow when a force is applied to it. Kinematic viscosity is the resistance to flow of a liquid when the force of gravity is applied to it. This flow of course depends on the density of the liquid, so that kinematic viscosity is determined by measuring the liquid's viscosity and dividing it by its density. It is important that a diesel fuel have the right viscosity, because the fuel must be finely divided for rapid combustion. The size of the droplets of fuel in the spray produced by the engine's injectors is strongly influenced by the fuel's viscosity.

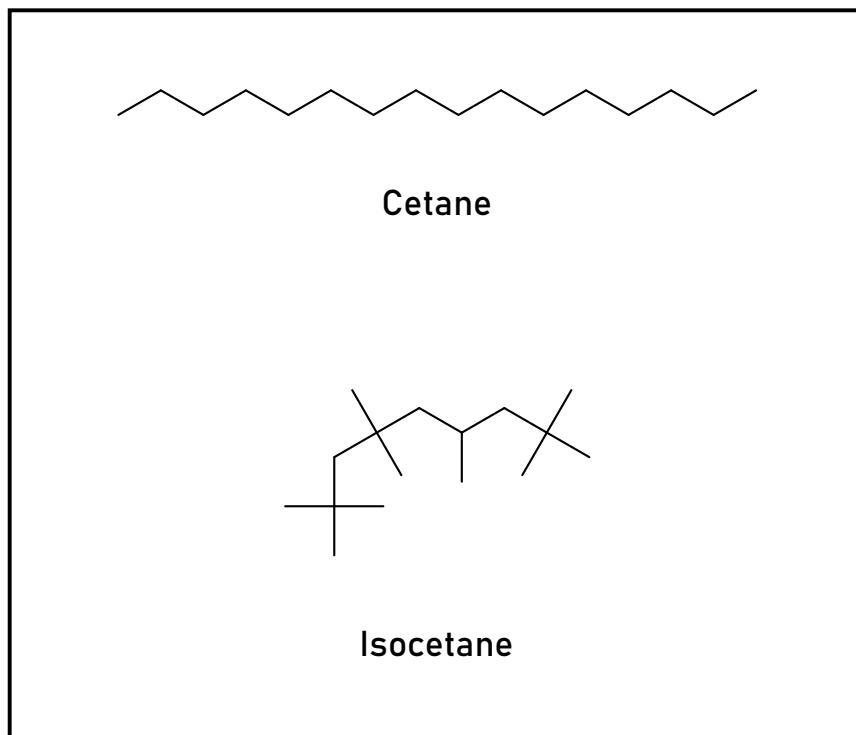


FIGURE 3.2: The chemical structures of the C16 hydrocarbon isomers cetane and isocetane.

`(fig:Cetane)`

SANS 1935 requires that kinematic viscosity be  $3.5 \text{ mm}^2 \text{ s}^{-1}$  to  $5.0 \text{ mm}^2 \text{ s}^{-1}$ . The measurement method is specified by ISO 3104, which uses a capillary viscometer: the time taken for a fixed volume of liquid to flow through a capillary. This time is then multiplied by an instrument-specific factor to yield the kinematic viscosity.

### 3.5 SANS 1935: Aggregate properties: Specialized instrumentation

Some of the requirements specified in SANS 1935 are not values that have direct correspondence to physical or chemical quantities usually used in science, but are measures that have proved useful in engineering practice.

#### 3.5.1 Cetane number

For a diesel engine to operate according to design the fuel must combust in a reliable manner. The cetane number is a number that indicates the ease of ignition of a fuel in a diesel engine. Cetane is a synonym for hexadecane, and is a liquid compound that ignites easily when injected into a diesel engine at high compression and temperature, without a spark. In contrast, its highly branched isomer 2,2,4,4,6,8,8-heptamethylnonane (also called isocetane, see Figure 3.2) ignites less easily. The intuitively-understood "ease of combustion" can be quantified as the **ignition delay**, the time between the moment the fuel is injected into the engine and the moment the ignition starts. A mixture of the two compounds will have an ignition delay somewhere between that of the two compounds. A higher number indicates easier ignition.

SANS 1935 prescribes the test method of ISO 5165. The cetane number is obtained by measuring the ignition delay in a highly specialized test engine. The engine is equipped with the necessary instrumentation to measure ignition delay. It is a single-piston, four-stroke engine that has a combustion chamber of variable volume, which gives it a variable compression ratio. While running on the fuel under test, the compression ratio is changed until a prescribed ignition delay is achieved. Then the cetane/isocetane mixture that will give the same ignition delay at that compression ratio is found. If pure cetane will give the prescribed ignition delay, then the cetane number 100 is assigned. If pure isocetane gives the prescribed ignition delay, then the cetane number 0 is assigned. If it is found that a 50:50 cetane:isocetane mixture produces the prescribed ignition delay, then the cetane number of 50 is assigned.

Cetane number gives limited insight into the chemical properties of the fuel, or indeed, engine performance. But it is a trusted measure of the suitability of a fuel for use in a diesel engine and is therefore included in the standard. Biodiesel from common feedstocks usually easily meets the cetane specifications.

### 3.5.2 Flash point

Fuels are, naturally, flammable, and different fuels are flammable to different degrees. A fuel's **flash point** can be used to quantify its degree of flammability. The flash point is the temperature at which a fuel's vapour at atmospheric pressure will ignite when it comes in contact with a flame.

The measurement of flash point is not primarily concerned with the performance of the fuel inside the engine, but is important to know because it determines how the fuel can be safely be handled during transport and storage (Worldwide Fuel Charter Committee 2009).

SANS 1935 specifies that either ISO 2719 (Procedure A) or ISO 3679 be used for determining the flash point. It seems that this aspect of SANS 1935 is out of date, because ISO 2719:2002 has been withdrawn and revised by ISO 2719:2016, which adds a Procedure C, specifically for FAME. ISO 2719 specifies the use of a Pensky-Martin closed-cup test, an automated device in which a temperature ramp heats an enclosed amount of liquid, while an ignition source is periodically introduced. The temperature at which the vapour ignites is the flash point. ISO 3679's test procedure works on a similar principle, but the device uses a smaller volume of liquid, and the liquid and its vapour are considered to be in thermal equilibrium.

### 3.5.3 Oxidation stability

(sec:Rancimat) Fossil fuels for diesel engines are largely composed of alkanes, which are chemically very inert. This means that they can be stored for long periods without significant degradation. Biodiesel, in contrast, is by definition (SABS 2008, Paragraph 4.1.1) mostly composed of fatty acids, which are less inert than alkanes. This instability has been known since antiquity as fatty food going **rancid**. In particular, double bonds in the fatty acids are liable to reactions with atmospheric oxygen (Velasco, Dobarganes, and Márquez-Ruiz 2010). These are free radical reactions, illustrated in Figure 3.3. The hydroperoxides that form are highly reactive and can react with other fuel compounds to form polymers, soaps and acids.

The oxidative stability is measured by the Rancimat, an automated instrument. In operation it bubbles a stream of hot air through a sample of biodiesel. This air then passes through a conductivity measuring cell filled with deionized water. The

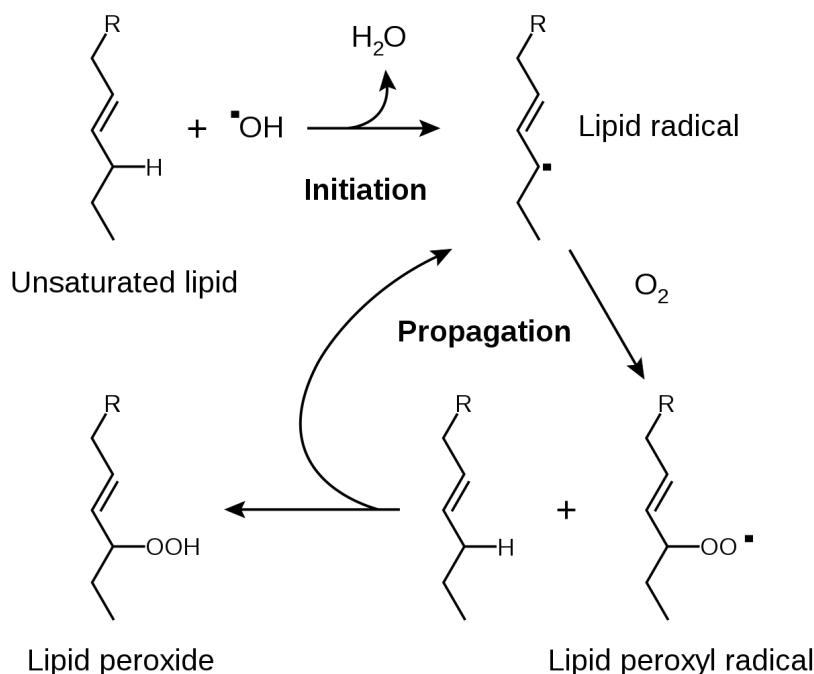


FIGURE 3.3: The peroxidation of lipids. By Tim Vickers, after (Young and McEneny 2001). Public Domain, <https://commons.wikimedia.org/w/index.php?curid=1728531>

(fig:RancidRadical)

1865 conductivity of the water increases when volatile acids dissolve in it. At the beginning of the test period very few acids are swept into the water, so the conductivity remains low. As the sample oxidizes and produces acids, the conductivity of the water slowly increases. But the oxidation reaction is **autocatalytic**, which means that the products of the oxidation reaction accelerates the oxidation. The result of the autocatalysis is that there is a sudden increase in conductivity on the conductivity curve. The time taken until this point is reached is called the **induction period**.

1870 SANS 1935 requires a minimum induction period of 6 minutes at 110 °C. Paragraph 4.1.2 permits the addition of antioxidants, compounds that prevent oxidation.

### 3.5.4 Cold filter plugging point

1875 The flow of a liquid depends on its temperature. Most obviously, its viscosity depends on the temperature. But a liquid can freeze, which will also influence its flow. The operating temperature of a diesel engine is so high that the temperature of the fuel in the engine will be high enough to guarantee adequate flow for injection, but the biodiesel needs to be pumped from the fuel tank, through a fuel filter, to the engine. If crystals of fuel form at low ambient temperatures, these crystals might plug the pores in the filter, which could lead to fuel starvation in the engine. It is therefore important that the biodiesel must be able to be pumped through the engine's fuel filter at all expected temperatures.

1885 For this purpose SANS 1935 specifies a **cold filter plugging point**, measured according to SANS 50116. It is the lowest temperature at which a given volume of biodiesel still passes through a standardized filtration device in a specified time when cooled under certain conditions.

### 3.5.5 Copper strip corrosion

Many fuel storage and transfer systems parts and engines parts are made of metals, and under adverse conditions those parts are susceptible to corrosion. Corrosion accelerates in chemical environments that include, *inter alia*, acids, water and oxygen, but corrosion is a complex process, so it is very difficult to predict which combination of factors will result in unacceptable corrosion.

SANS 1935 prescribes ISO 2160 as the test for corrosiveness. In this test a strip of pure copper metal is polished and then immersed in a sample of the biodiesel at 50 °C for 3 h. The degree of corrosion is judged by visually comparing it to a standard card and then assigning it to a corrosion class.

## 3.6 SANS 1935: Chemical properties: Classical determination

### 3.6.1 Sulfated ash

Ash is the solids remaining after the complete combustion of a fuel. This is measure of the solids that will remain after use. Ash might originate from suspended solids, soluble metallic soaps, and residual catalyst that form refractory oxides during combustion. In addition to wear and deposits in the fuel system associated with ash, it also impairs modern diesel particulate filters.

SANS 1935 specifies ISO 3987 for the determination of sulfated ash. For the determination of ash it would not be practical to actually combust the sample and weigh the residue, because the ash might be carried away, either as volatile species or as a finely divided aerosol. Therefore the sample is oxidized using sulphuric acid, and then heated to drive off the sulphuric acid. What remains is metal oxides, which can be weighed.

### 3.6.2 Carbon residue

When a mixture of organic compounds is heated under conditions of low oxygen, it can form coke. Coke is a material that is practically pure carbon. Pure, solid carbon will combust, but the reaction is kinetically limited, so it will only combust slowly under conditions of high temperature and a large excess of oxygen. If a fuel tends to form coke inside an engine, it can cause problems in operation. Vegetable oil as diesel fuel, for example, tend to form coke on the injectors (van der Walt and Hugo 1982). This coking can cause problems with injection, which would affect engine performance. Carbon residue is not a scientific measure: from long experience it has been found to correlate with coking tendencies of oils in the petroleum industry, and so found its way into the biodiesel standard.

The test method prescribed for testing for carbon residue is ISO 10370. This involves heating a sample of biodiesel to high temperature in a crucible in air, using standardized apparatus. Most of the sample burns off, leaving a residue. The mass of the residue is determined and reported.

### 3.6.3 Total contamination

Ideally, biodiesel should be a homogeneous liquid. Some undissolved material can be tolerated, but too much can plug filters. The amount of undissolved solids is termed **total contamination**.

The test of total contamination specified by SANS 1935 is contained in SANS 52662, which is synonymous with EN 12662. In this test total contamination is determined by obtaining the mass of material retained on a glass-fibre filter after passage of the biodiesel sample.

<sup>1935</sup> **3.6.4 Acid value**

The end products of oxidative degradation of biodiesel include free organic acids, so the acidity of biodiesel is a good indicator of its quality. Measuring the acidity gives an indication of how much the biodiesel has already oxidized, whereas the oxidation stability (measured by Rancimat, see Section [3.5.3](#)) indicates how well the biodiesel will withstand oxidation on storage.

<sup>1940</sup> SANS 1935 specifies the test described in SANS 54104, which is equivalent to EN 14104. This test is a titration with alcoholic KOH of a sample of biodiesel dissolved in a mixture of solvents. A glass pH electrode connected to an electronic pH meter is used to follow the titration.

<sup>1945</sup> **3.6.5 Iodine value**

As discussed above (see Section [3.5.3](#)) the oxidation of biodiesel is a major quality concern. The tendency of biodiesel to oxidize correlates with the number of double bonds in the fatty acids, or their **degree of unsaturation**.

<sup>1950</sup> The degree of unsaturation of oils and fats have long been measured by the **iodine value**, dating from 1884 (Knothe [2007](#)). Halogens will rapidly add to double bonds, so that when a mixture of fatty acids is treated with a known excess of iodinating reagent, the remaining iodine can be titrated to determine the amount of iodine absorbed by the fatty acids. The iodine value is the mass of iodine absorbed by 100 mass units of fat or oil.

<sup>1955</sup> The relevance of including iodine value in biodiesel standards has been questioned (Knothe [2002](#)), because all the information regarding unsaturation of the fatty acids in biodiesel is contained in chromatographic data which need to be obtained in other requirements. SANS 1935 seems to acknowledge this, because Annex A allows that the iodine value can be calculated from chromatographic results (see Section [3.9.5](#)).

<sup>1960</sup> Because determining iodine value is a mature technology and therefore relatively simple, it is tempting to think of it as a test that might be useful to small biodiesel producers. Unfortunately the iodine value is determined by the feedstock, so that for a feedstock from a certain vegetable oil crop there is unlikely to be any significant variation in iodine value, no matter the production process. Iodine value might be useful as a simple indicator of feedstock quality when the biodiesel is produced from waste vegetable oil, which might contain a variety of oils from different origins.

<sup>1965</sup> The prescribed method for the requirement is SANS 54111 (or, equivalently, EN 14111). In this method a known excess of Wijs's reagent (iodine chloride in acetic acid) is added to a weighed sample. The reaction mixture is then treated with potassium iodide, which converts the excess  $\text{ICl}$  to  $\text{I}_2$ , which is then titrated with potassium thiosulfate. The titration is followed potentiometrically and the equivalence point determined from the titration curve.

## 3.7 SANS 1935: Chemical properties: Electrochemical determination

1975

### 3.7.1 Water content.

The compounds that comprise biodiesel are much more polar than those of petrodiesel. Therefore, much more water can dissolve in biodiesel than can dissolve in petrodiesel. This water has several deleterious effects on the quality of biodiesel. It encourages the growth of micro-organisms, allows hydrolysis, and accelerates corrosion.

The level of water specified in the requirements of SANS 1935 is lower than the solubility of water, and therefore refers to dissolved water. Free, visible water is excluded by paragraph 4.1.4.

The method specified by SANS 1935 for the determination of water in biodiesel is ISO 12937. This standard prescribes the well-known Karl Fischer titration used for the determination of water in solvents. This is a **coulometric** titration, which means that electricity is used as titrant. The titration curve is a plot of oxidation potential of a platinum electrode against the amount of charge (current integrated over time).

## 3.8 SANS 1935: Chemical properties: Spectroscopic determination

### 3.8.1 Group I metals

The most common catalysts in biodiesel production are sodium or potassium hydroxides ( $\text{NaOH}$  and  $\text{KOH}$ ) or alkoxides ( $\text{CH}_3\text{ONa}$  and  $\text{CH}_3\text{OK}$ ). These catalysts are polar and will dissolve in the glycerol byproduct of biodiesel production. But some may remain in the biodiesel itself, and needs to be removed. This cleanup can be done by washing with water, adsorbent columns, or selective membranes (Atadashi et al. 2011).

SANS 54108 is the method specified for the determination of sodium, and SANS 54109 the method specified for potassium. Both are flame atomic absorption spectroscopy methods: a sample of the biodiesel is aspirated into a gas flame, where the sodium and potassium are atomized. The atoms will absorb light at certain wavelengths, and measuring the amount of light absorbed will give a measure of the amount. Alternatively, EN 14538 may be used to determine sodium and potassium simultaneously with calcium and magnesium (see Section 3.8.2).

### 3.8.2 Group II metals

Fatty acids neutralized by alkali and alkaline earth metal hydroxides form **soaps**. If these metals are present in the feedstock, as catalyst, or in washing water they can form soaps with the fatty acids in the biodiesel. These soaps can form deposits in engines that can affect operation. For example, deposits of calcium soaps have been reported to cause injectors to stick (Pischinger, Schlag, and Mittelbach 2000).

SANS 1935 prescribes EN 14538 as the method for determining Group II metals. This is an optical emission spectroscopy method: a sample of the biodiesel is diluted in kerosene, and injected into a inductively coupled argon plasma. The emission intensities at certain wavelengths are compared to the emission intensities of solvents containing known concentrations of the metals.

### 3.8.3 Phosphorus

Phosphorus expected in biodiesel should not affect a diesel engine's performance, but it can have a detrimental effect on the exhaust treatment system by forming ash that can clog filters and reactive species that can reduce catalyst effectiveness. Trace levels of phosphorus will be expected in biodiesel, in the form of natural phospholipids. Normal biodiesel feedstock and production methods should yield acceptable phosphorus concentration, but inorganic phosphorus might be present in biodiesel produced from used cooking oil.

SANS 54107 is the prescribed method for determining phosphorus in biodiesel. A sample of biodiesel is dissolved in xylene, and the solution introduced in aerosol form into an inductively coupled argon plasma. The high temperature of the plasma causes phosphorus atoms and/or ions to emit radiation. This emission is measured at a certain wavelength, and compared to emissions from solutions with known concentrations.

### 3.8.4 Sulfur

The amount of sulfur in biodiesel is limited not because it affects the fuel's performance, but because the fuel must be compatible with emission control systems and must not emit more sulfur than petrodiesel. Most biodiesel feedstocks are naturally low in sulfur and are therefore unlikely to exceed the limits.

SANS 1935 offers two alternative tests for sulfur. ISO 20846 is a UV fluorescence method, while ISO 20884 is an X-ray fluorescence methods. The quantum-mechanical mechanism is the same for both methods: a chemical species absorbs energy from a photon which puts it in an activated state. The species then returns to a state of lower energy, emitting a photon of different energy. In the case of ISO 20846 the species is gaseous SO<sub>2</sub> (obtained by combusting the sample), and the activating photons are from the ultraviolet part of the electromagnetic spectrum. In the case of ISO 20884 the chemical species are the bound form of the sulfur as found in the biodiesel, and the activating photons are from the X-ray region of the electromagnetic spectrum.

## 3.9 SANS 1935: Chemical properties: Chromatographic determination

*(sec:ChromDet)*

### 3.9.1 Methanol content

Methanol in biodiesel increases its flash point, and it is an indicator of poor production process control.

SANS 1935 requires a maximum mass fraction of 0.2 % of methanol in biodiesel. The prescribed test method is contained in SANS 54110, and involves heating a sealed vial partly filled with biodiesel to 80 °C. A portion of the headspace vapour is taken and injected into a gas chromatograph. The amount of methanol is quantified by comparing the methanol peak to either an internal or external standard.

### 3.9.2 Ester content

*(sec:EsterContent)*

As prescribed in Paragraph 3 of SANS 1935, biodiesel must consist of fatty acid methyl esters. The first line of Table 1 quantifies this requirement as a minimum of 96.5 % mass fraction.

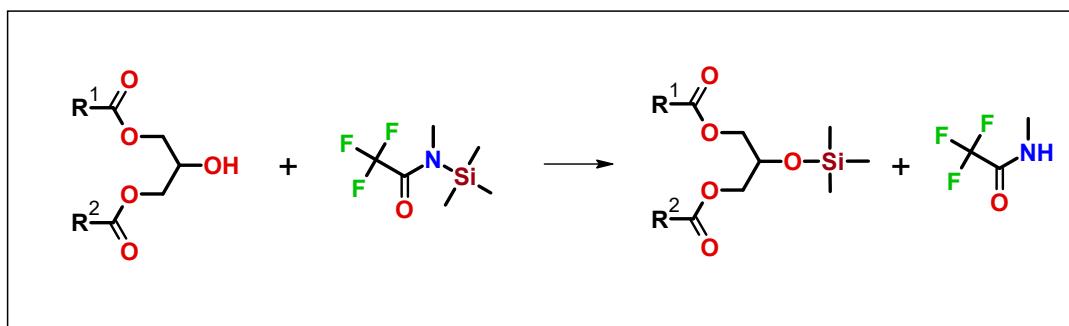


FIGURE 3.4: The derivitization of a diglyceride with N-methyl-N-(trimethylsilyl)-trifluoroacetamide (MSTFA). During GC the inert trimethylsilyl group does not interact with the stationary support, and the volatile trifluoroacetamide elutes with the solvent front.

*(fig:MSTFA)*

The specified method is SANS 54103. This document refers to ISO 5508:1990, which has been withdrawn and superseded by ISO 12966-4:2015. ISO 5508 and ISO 12966 describe gas chromatographic methods for the determination of FAMEs. ISO 5508 is obsolete: it gives conditions for packed GC columns and thermal conductivity detectors, two technologies which are now rarely found in the chromatography laboratory. Both methods, however, require polar stationary phases. The quantity of esters is determined by integrating all the peaks in a certain retention time window and comparing it to the peak area of an internal standard.

### 3.9.3 Glyceride content

*(sec:Glycerides)* Glycerol (propane-1,2,3-triol) is the “backbone” of the oil molecules that constitute the feedstock for biodiesel production. Each hydroxyl group can form an ester bond with a fatty acid, and when all three have a fatty acid bound to it is an ‘oil molecule’. IUPAC recommends that such a molecule be called a tri-O-acylglycerol (Nič et al. 2009), but by long-established custom they are called triglycerides. The conversion of a triglyceride to FAMEs is a stepwise process, with one ester bond at a time being methanolized. This means that during the reaction process, there will also be di- and monoglycerides (di- and mono-O-acylglycerols) in the reaction mixture. If the reaction is not well controlled then these glycerides will be found in the final product.

The presence of glycerides is of course an indicator of an incomplete transesterification reaction, but it has further negative effects. In particular, during cold weather, or in petroleum-blended biodiesel, some dissolved impurities might precipitate, in particular the monoglycerides (Dunn 2009; Plata, Gauthier-Maradei, and Kafarov 2015). This precipitate might block filters or otherwise interfere with engine performance.

SANS 1935 specifies that the mono-, di- and triglyceride content of biodiesel must be determined with a procedure compliant with SANS 54105 (or, equivalently, EN 14105). In this method, the biodiesel sample is treated with MSTFA (2,2,2-Trifluoro-N-methyl-N-(trimethylsilyl)acetamide) before separation by GC.

MSTFA is a **derivatization reagent**: It reacts with the hydroxyl groups in the glycerides to form trimethylsilyl derivatives, as shown in Figure 3.4. The molecule is then much more inert and will not interact with the stationary phase support, yielding peaks with better shapes.

2095

The trifluoroacetamide group is a good leaving group, and upon nucleophilic attack by the hydroxyl oxygen its bond with the trimethylsilyl (TMS) group is cleaved, leaving the TMS bound to the oxygen. The labile hydrogen atom is now replaced by the TMS group, which is inert and will not interact with polar entities in the stationary phase or column, which will lead to improved peak shapes.

### 3.9.4 Free Glycerol

2100

Free glycerol is one of the products of the transesterification of plant oils to produce biodiesel. It is a polar compound, which naturally separates from non-polar biodiesel, and any excess remaining dissolved in the biodiesel is removed during the washing step. Free glycerol contributes to injector coking.

Inappropriate processing may leave excess free glycerol in the biodiesel, and therefore determining free glycerol is an important quality-control step.

2105

Free glycerol can be determined by the same chromatographic procedure prescribed for the determination of the other glycerides in SANS 54105 (see Section 3.9.3, but SANS 1935 also offers the option of SANS 54016. This standard uses a liquid-liquid extraction of biodiesel with a mixture of ethanol, water, and hexane. The free glycerol transfers quantitatively to the bottom layer, which is then analysed 2110 with a gas chromatographic method. The benefit of this method is that the resulting chromatogram has only one peak, that of glycerol, so quantification is straightforward.

2115

The column specified in SANS 54106 method is a PLOT column, *i.e.* a Porous Layer Open Tubular column, which is a capillary lined on the inside with a layer of particles coated with the stationary phase, in this case a polar polyethylene glycol.

### 3.9.5 Polyunsaturated methyl esters

&lt;sec:ChromDetUnsat&gt;

The degree of unsaturation of the constituent fatty acids is the main determinant of the oxidative stability of biodiesel. SANS 1935 limits the amount of highly unsaturated FAMEs by two lines in Table 1. In particular, the amount of methyl linoleate 2120 is limited to less than 12 % mass fraction, and the amount of polyunsaturated fatty acid methyl esters with four or more double bonds is limited to less than 1 % mass fraction.

#### Linolenic acid methyl ester

2125

The prescribed method for the determination of methyl linoleate (C18:3) is SANS 54103, the same method as prescribed for total esters (see Section 3.9.2).

#### Highly unsaturated fatty acid methyl esters

2130

Polyunsaturated fatty acids (PUFAs) with more than four double bonds are not commonly found in plant oils, but are found in algae. Fish that feed on algae accumulate these oils, so that they are also found in fish oils, but fish oil is not a sustainable resource and should not be used for fuel (Kitessa et al. 2014). PUFAs are also highly oxidatively unstable, and therefore undesirable in biodiesel.

The method prescribed for the determination of PUFA FAMEs is EN 15779. This method uses a capillary column with a polyethylene glycol stationary phase.

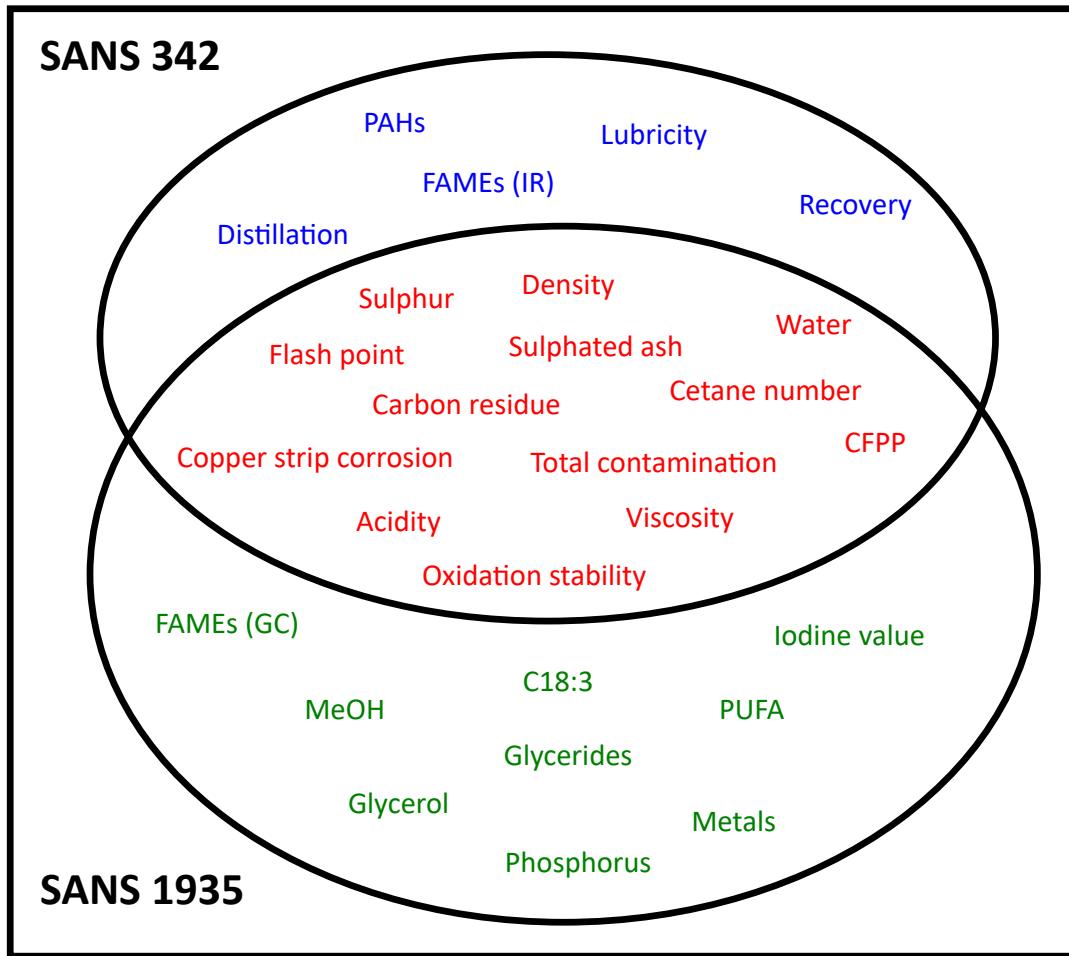


FIGURE 3.5: The overlap of requirements of SANS 1935 for biodiesel and SANS 342 for petrodiesel.

{fig:Venn}

### 3.10 Comparison of SANS 1935 and SANS 342

<sup>2135</sup> SANS 1935 sets the standard for neat biodiesel or biodiesel blendstock in South Africa, and SANS 342 sets the standard for petrodiesel. As can be expected, there is considerable overlap of the requirements of petrodiesel and biodiesel, summarized in Figure 3.5.

<sup>2140</sup> There are 5 requirements for petrodiesel in SANS 342 that have no equivalents in SANS 1935. The first two, labelled “Distillation” and “Recovery” are borrowed from petroleum engineering tests, and are intended to ensure an appropriate volatility range. These two tests fall outside the scope of this discussion.

<sup>2145</sup> The amount of **polycyclic aromatic hydrocarbons** (PAHs) in petrodiesel is limited by SANS 342 to help control their levels in the environment. PAHs have many sources, including uncombusted petrodiesel.

<sup>2150</sup> The Petroleum Products Act allows the blending of up to 5 % of biodiesel with petrodiesel without the need to label it explicitly as a biodiesel blend. The method prescribed by SANS 342 for determining the FAMEs in petrodiesel is EN 14078. This is an infrared (IR) spectrophotometry method: the amount of infrared radiation absorbed by a sample of diesel is compared to the radiation absorbed by solutions

containing known amounts of FAMEs. This method is based on the strong absorption of IR the carbonyl group at  $1745\text{ cm}^{-1}$ ; of course every fatty acid contains a carbonyl group. Petrodiesel exhibits low absorbance at this wavelength, because petrodiesel contains only trace levels of compounds with carbonyl groups. While 2155 this spectroscopic test method can be made very simple, its accuracy depends on certain assumptions about the sample, which might not be valid. (Pinho et al. 2014).

The pumps in modern diesel engines generate pressures of up to 2500 bar, which imply that the mechanical parts exert large forces on each other. Metal parts that transmit large forces need to be lubricated to limit wear to acceptable rates. Some of 2160 these pump parts are immersed in the fuel, and can therefore not be conventionally lubricated: the circulating fuel will rapidly dissolve and sweep away any grease or lubricating oil. Therefore, the lubrication depends on the lubricating properties of the fuel itself. **Lubricity** is a measure of the lubricating properties of a fuel. It is an aggregate property, and it is quantified by the **wear scar** caused when two metal test 2165 parts of specified shape are rubbed against each other with a specified force for a specified duration, while being lubricated by the liquid under test. Such tests are done using special-purpose test machines.

SANS 1935 contains no lubricity requirement, because it has been found that biodiesel always provides adequate lubricity. Also, biodiesel can be used as a lubricity improver in ultra-low-sulphur petrodiesel that might not meet the requirement. This lubricity is not provided by the FAMEs, but by the polar impurities in the biodiesel, such as free fatty acids and monoglycerides (Knothe and Steidley 2005).

### 3.11 Conclusion: room for innovation.

2175 The diversity of chromatographic biodiesel quality control methods implies that the compliance with SANS 1935 can be expensive. To comply, a biodiesel producer will have to submit samples of biodiesel to at least four different kinds of chromatographic analysis:

- EN 14105 for free and total glycerol
- EN 14103 for FAME and methyl linoleate content
- EN 14110 for residual methanol
- EN 15779 for highly unsaturated FAMEs
- (Optionally) EN 14106 for free glycerol

This might prove costly, and there has been at least one innovation to reduce the number of instruments required (McCurry and Norman 2009).

2185 As was emphasized in the introduction to this chapter (see Section 3.1), standards are not 'the best' way of determining a certain desirable property of biodiesel, they are merely trusted methods. As improved methods are developed and become trustworthy, they can be adopted as alternatives. As far as biodiesel is concerned, it seems that there is ample room for chromatographers to innovate and develop improved methods and instrumentation. This was the motivation for selecting biodiesel as a practical challenge for developing SFC $\times$ GC.

$pV \neq nRT$

The universal gas law does not apply to supercritical fluids

# 4

## Instrumentation: Supercritical Fluid Chromatography

(Chapt~~2~~<sup>2200</sup> 54) In this thesis the development of a comprehensively coupled (supercritical fluid  $\times$  gas) chromatograph and its application to the analysis of biodiesel is discussed. The discussion on the experimental work divides naturally into two parts: this chapter discusses the supercritical fluid chromatography (SFC) and the next chapter discusses the gas chromatography (GC).

### 4.1 SFC

As discussed in Chapter 2, an SFC chromatograph consists of a supply of mobile phase, a pump, a pressure control system, a modifier control system, a column, a pressure relief system and a detector. These subsystems are discussed below.

### 4.2 Mobile phase

As mobile phase we used carbon dioxide. The benefits of carbon dioxide as an solvent and mobile phase is discussed in Chapter 2. We used 99.995 % pure carbon dioxide supplied by Air Products. Our colleagues in industry also use food grade or technical grade carbon dioxide and they have not reported any significant impurities.

As modifier we used LiChrosolv®methanol from Merck. This is an 'HPLC-grade' solvent, of which the purification is optimized for the removal of UV-absorbing impurities. Using this grade of solvents is important when using optical absorbance or fluorescence detectors, because lowering levels of impurities improve limits of detection. It is quite likely that a less expensive grade of modifier would also be suitable for our SFC, because we do not use an optical detector.

### 4.3 Pump

(sec:C02Pump) As pump we used the robust, reliable Varian 8500 HPLC pump because it was available. This pump had its control electronics removed, and it was controlled from a

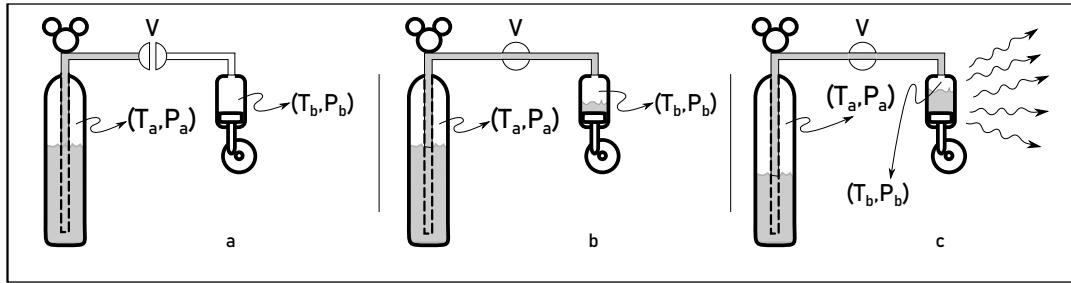


FIGURE 4.1: A schematic diagram of the process of the filling of the carbon dioxide pump. (a) The pressure in the reservoir  $P_a$  is much higher than the pressure in the receptacle  $P_b$ . The valve is closed, so there is no flow. (b) The valve is open and some carbon dioxide has flowed from the reservoir to the receptacle. But  $P_b = P_a$ , so there is no flow. (c) When the temperature in the receptacle  $T_b < T_a$ , then  $P_b < P_a$ , following Gay-Lussac's law. Flow will continue until all the vapour in the receptacle has condensed.

(fig:co2fill)

personal computer by software written for the purpose. The Varian 8500 pump is driven by a stepper motor. This kind of motor turns in discrete steps, instead of at a constant rate. It is driven by pulses of electric current, rather than a continuous current, and therefore the speed of the motor can be controlled by varying the rate of the pulses. This makes it relatively simple to control the speed of the motor from a computer. The Varian 8500 pump has a built-in pressure transducer that provides an electronic signal proportional to the pressure at the pump outlet.

The Varian 8500 pump is a single-piston pump with a 250 ml capacity that needs to be refilled between strokes. Refilling means that one has to stop chromatography, which makes it important to fill the pump to full capacity, so that chromatographic runs are not interrupted. This means that the pump needs to be filled with carbon dioxide in the liquid phase rather than the vapour phase. Compressing either would create the appropriate phase for doing chromatography, but compressing a vapour would leave one with a much smaller volume of high-pressure carbon dioxide than compressing a liquid. Filling a pump with liquid carbon dioxide is more difficult than one might imagine.

Imagine a reservoir of liquid carbon dioxide, equipped with a dip tube, connected to an empty receptacle via a valve (V). (See Figure 4.1) One can assume that the receptacle is empty and contains only carbon dioxide vapour at atmospheric pressure, say 1 bar. The vapour pressure of the vapour above the liquid in the reservoir is about 55.3 bar (5.6 MPa). When the valve (V) is opened the high-pressure vapour in the reservoir will expel the liquid carbon dioxide through the dip tube and through the valve, into the receptacle. In the low-pressure environment of the receptacle the carbon dioxide will boil. Soon there will be some liquid carbon dioxide in the receptacle, with the rest of the receptacle volume filled with gaseous carbon dioxide. When the system comes to equilibrium the pressure in the receptacle ( $P_b$ ) will equal the pressure in the reservoir ( $P_a$ ), and there will be no flow of liquid carbon dioxide. One can attempt to now increase the flow by increasing the volume of the receptacle (for example by withdrawing a pump piston), and hence decreasing the vapour pressure there. But any flow from the reservoir will lead to expansion of the headspace of the reservoir. This will lead to cooling of the vapour, and therefore lower pressure and therefore lower flow. The final result of this process is that the receptacle is never filled to capacity with liquid.

The only way to restore the flow from the reservoir to the receptacle is to create a pressure difference  $P_a - P_b$ . While it is possible to create an overpressure in the reservoir by adding a headspace gas, it is technically challenging and expensive.  
2255 It is simpler to use the Gay-Lussac gas law. According to this law the ratio of the pressure to the temperature of a gas is constant so that  $\frac{P_b}{T_b} = k$ . This means that a decrease in temperature will lead to a decrease in pressure. The way to fill the reservoir to capacity is to ensure that the temperature of headspace vapour  $T_a$  is higher than the temperature of the headspace vapour  $T_b$ . Because safety regulations  
2260 prohibit the heating of a cylinder of pressurized gas, the way to ensure  $T_a < T_b$  is to cool the receptacle. This cools the headspace vapour, and following the Gay-Lussac law the pressure  $P_b$  decreases, the difference  $P_a - P_b$  increases and the liquid flows until the receptacle is filled to capacity.

In the case of the Varian 8500 pump the cooling of the pump was achieved by  
2265 wrapping a coil of copper tubing around the cylinder, and pumping a chilled heat-exchange fluid through it. A chiller with a mechanically cooled tank with a 20 l capacity was filled with a solution of 5.0 l of diethyl glycerol in about 10 l of water. This mixture has a freezing point of  $-15^\circ\text{C}$ , which can be cooled by the chiller without freezing. (If the coolant freezes a layer of ice forms on the cooling plate of the chiller,  
2270 which isolates the remaining liquid from the cooling plate and limits the minimum temperature of the coolant.) An inexpensive submersible water pump (designed for decorative water fountains) was used to pump the coolant through the circuit. This pump can deliver 800 l of water per hour at a head of 1.2 m.

For some experiments we also used a SFT-10 pump from Supercritical Fluid Technologies (Newark, Delaware). This is a purpose-built two-piston pump with sapphire valve seats and a Peltier-cooled head. This is a much better technology than the HPLC pumps. It takes up less space and does not need refilling, since it feeds directly from the cylinder. This pump had its own microprocessor controllers on board, and flow and pressure could simply be commanded from the PC through a  
2280 USB cable.

### 4.3.1 Pressure control system

The aim of chromatography is to separate chromatograms by a certain distance. But the days of separating coloured compounds in glass packed columns are long past and the direct measurement of distances are now relegated to thin-layer chromatography.  
2285 Instead, we have to make do with proxies for distance such as **retention volumes** or **retention times**.

Retention times are particularly convenient today, because they can easily be measured by computer systems. Retention times, however, depend on a known flow rate of the mobile phase through the column. The flow rate need not be constant,  
2290 although for the sake of simplicity a constant flow rate is preferred. Ideally, the flow rate should also be adjustable. The need for an adjustable, constant flow rate can be met by using a control system. Control systems are well understood by engineers, among whom it is a major field of study (Koenig 2009).

Figure 4.2 shows a diagram of a simple process control system. Some aspect of  
2295 the process under control is measured, which yields the **process variable** (PV). The PV is compared to the **set value** (SV), and the **error** ( $e$ ) is obtained by finding the difference. The error is provided to the controller, which calculates the **manipulated variable** (MV). The MV is used to drive the final control element, which adjusts the process with the aim of producing a smaller error.

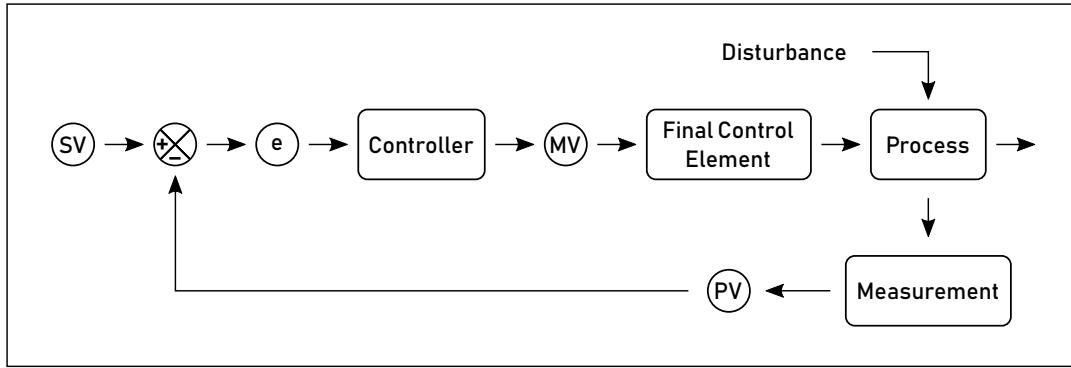


FIGURE 4.2: Schematic diagram of a closed-loop control system. (SV) Set Value, (PV) Process Variable, (e) Error, (MV) Manipulated Variable

`(fig:processcontrol)`

2300 Chromatographic systems conventionally work under pressure control or flow control regimes. Either is suitable: constant pressure systems usually yield constant flow if other parameters are held constant. Because flow measurement is more complex than pressure measurement, pressure control is the older, simpler and less expensive control method. Also, in SFC the solvent strength of the mobile phase is  
2305 pressure/density dependent. We therefore selected pressure control as our control regime.

In our system the pressure measured by the on-pump sensor was the process variable (PV). This was compared to the set value (SV) set on the computer console. The controller was a software algorithm, implemented as a virtual instrument in the  
2310 programming environment LabVIEW. The controller computed an output value for the manipulated variable (MV), which was the rate at which pulses were sent to the pump, in hertz. The digital input/output (DIO) interface of a National Instruments PCI-6014 multifunction data acquisition board created those pulses and sent them to the pump's electronic interface.

2315 The software used a PID (Proportional-Integral-Derivative) controller module, with the Integral and Derivative contributions disabled. The simpler proportional control was found totally adequate for our purpose, and it can be improved in future by finding an appropriate tuning method and applying it with the integral and/or derivative activated.

### 2320 4.3.2 Modifier Control

The modifier needs to be present in the mobile phase at a known and controlled concentration. There are various SFC-modifier mobile phase supply units on the market. They tend to be expensive and complex, because they require the use of two controlled, high-pressure pumps. Given that our needs were rather modest,  
2325 we elected to use a simpler system. Instead of accurately pumping the supercritical fluid and the modifier, we only pump and control the supercritical carbon dioxide and add measured volumes of modifier.

The modifier control unit consisted of a six-port valve, a fixed-volume measuring loop, and a mixing chamber. The unit operates by filling the measuring loop with  
2330 modifier, and then switching the measuring loop into the flowing mobile phase. The modifier is washed into the mixing chamber, where it is intimately mixed with and dissolved in the supercritical carbon dioxide. The concentration of modifier in the mobile phase is determined by the switching rate of the sampling valve. Given that

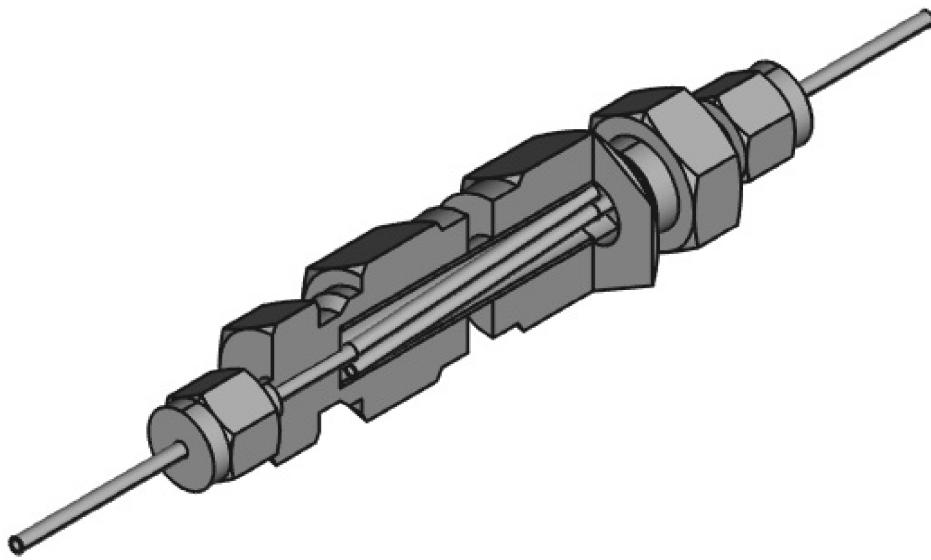


FIGURE 4.3: A cutaway diagram of the mixing chamber design. The chamber is designed for the mixing of the modifier and the supercritical carbon dioxide.

fig:mixingchamber)

about 3 volumes of supercritical carbon dioxide is needed to wash the modifier out of the loop, the highest concentration of modifier that can be added in this manner is about 25 %.

Figure 4.3 shows the chosen design of the mixing chamber. The inlet and outlet pipes of the mixing chamber extend far into the chamber, to break up plug flow and encourage rapid mixing.

The modifier injection valve position was commanded from the controlling PC by electronic voltage pulses.

## 4.4 Sample injection

The sample inlet was a two-position rotary valve with an internal sampling volume (Figure 4.4). In one position the sampling volume is filled with a syringe, and when the valve was commanded to inject the sampling volume is switched into the carrier mobile stream. The sample volume was 0.5  $\mu\text{l}$ . The injection valve position was commanded from the controlling PC by electronic voltage pulses.

The reader might be more familiar with the six-port valve as an injection system. The benefit of the internal groove injection valve is that it allows a much smaller volume to be injected. Using this small volume makes it possible to inject samples without dilution, which simplifies sample preparation. A drawback of this design is that it is not possible to vary the volume of sample injected.

### 4.4.1 Column

The separation power of a chromatographic system can be modelled by the equation

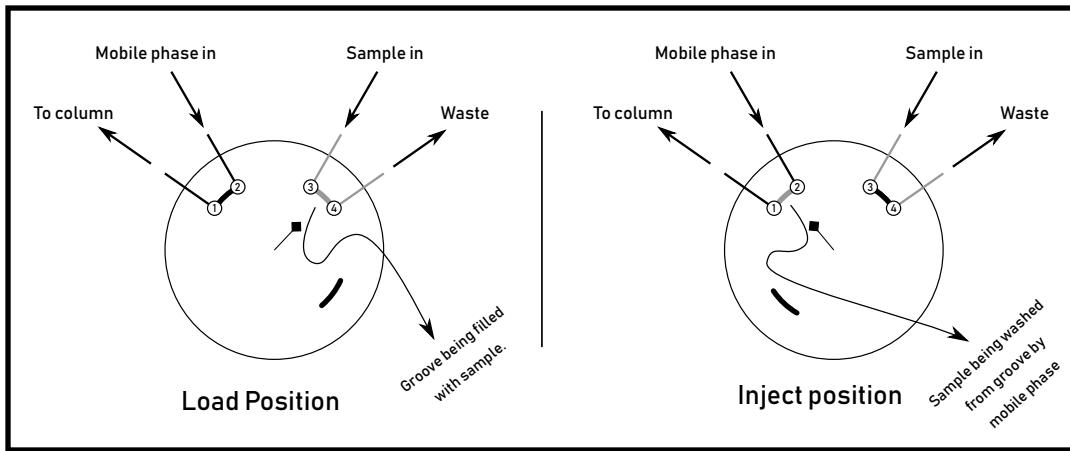


FIGURE 4.4: A schematic diagram of the sampling valve.

&lt;fig:samplingvalve&gt;

$$R_s = \left( \frac{\sqrt{N}}{4} \right) \left( \frac{\alpha - 1}{\alpha} \right) \left( \frac{k'}{1 + k'} \right)$$

were  $R_s$  is the **peak resolution**,  $N$  is the **number of plates**,  $\alpha$  is the **selectivity**, and  $k'$  is the **retention factor**.

The number of plates  $N$  can be increased simply by making the column longer, at the cost of increasing the time of the run. The maximum length of the column is determined by the pressure drop available that will still yield adequate flow. In capillary gas chromatography the openness of the column and the low viscosity of the gas-phase mobile phase routinely allows columns 100 mlong at inlet pressures of a few atmospheres. In HPLC the high viscosity of the mobile phase and the narrow, tortuous pathways between the small particles of the packing material means that the columns are typically 100 mm to 200 mm long, requiring hundreds of atmospheres of inlet pressure for adequate flow.

The low viscosity of an SFC mobile phase allows operation of much longer HPLC-type packed columns. This provides a much higher number of plates in the chromatographic system than HPLC. The SFC column we used in the SFCxGC system was a set of five HPLC columns (150 mm × 4.6 mm, 3 µm particles) (Restek, Pinnacle DB Silica) connected in series.

The stationary phase in this column is 'bare silica', which is usually considered 'polar'. That means that the packing of the column consists of porous silica particles with no organic phase covering its surface, and hence one that will interact strongly with polar molecules. In contrast, the ubiquitous 'C18' stationary phase used in reverse-phase HPLC is 'non-polar'. The particles of such a stationary phase are coated with octadecyl chains bonded to its surface, and hence will tend to interact strongly with non-polar molecules. The base material of the particles is usually silica, but that is only because manufacturing uniform particles of a given size from silica is a mature technology and not because silica is important to the separation mechanism. (In fact, some care has to be taken to deactivate the silica so that it does not contribute to the retention. Such a mixed retention mechanism can lead to **peak tailing**.)

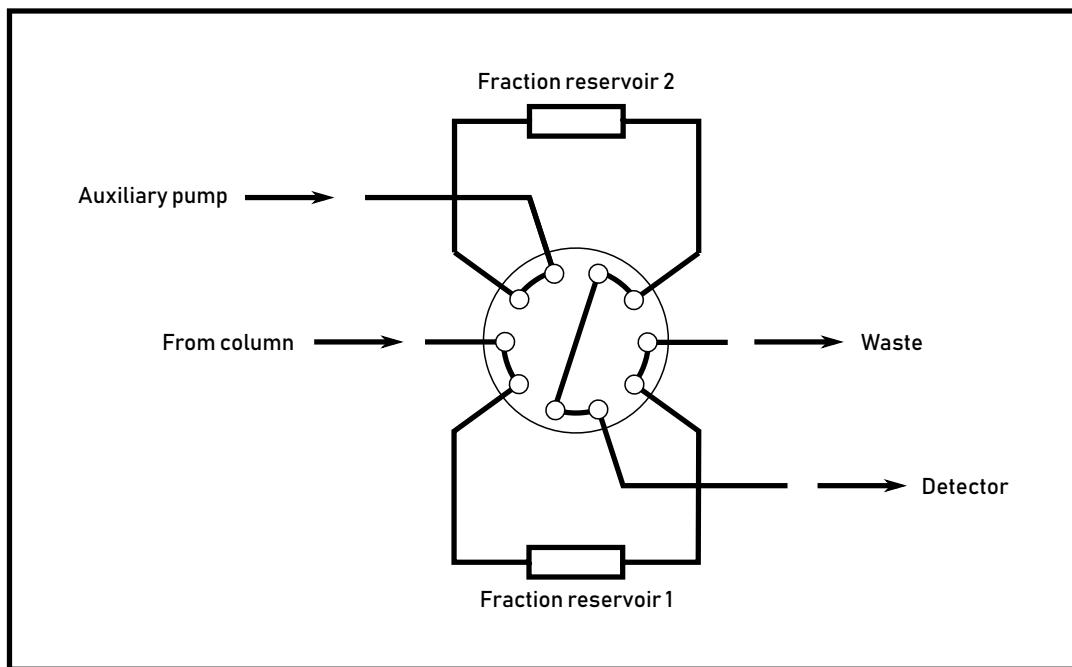


FIGURE 4.5: A schematic diagram of a valve configured to allow continuous-flow fraction collection.

ig:continuousflow)

#### 4.4.2 Stopped-flow

(sec:stopflow) Among analytical chemists it is the convention that chromatographic runs are done without interruptions. In our SFC $\times$ GC chromatograph we collect fractions of SFC eluate and separate them by GC. It is possible to collect, store and inject fractions of SFC eluate while maintaining continuous flow, for example using a dual storage loop system like the one depicted in Figure 4.5. But such systems require careful selection of volumes, fraction collection reservoirs with matching volumes, and an auxiliary pump.

Instead of such a system, we opted for a more versatile and simpler stopped-flow system. In this system a fraction of the SFC eluate is collected, and then the flow through the column is stopped. While the flow is stopped the collected fraction is subjected to separation by GC. The flow was stopped by a six-port rotary valve using three of the ports, directing flow either to a blocked-off port or to the depressurizer, in effect making it an on-off controller.

In packed columns, "eddy diffusion is always the main cause for peak broadening at any temperature and at high velocities" (Gritti and Guiochon 2006). During the period that the flow is stopped only longitudinal diffusion contributes to peak broadening, which is small compared to eddy diffusion in dense mobile phases. Using the simple stopped-flow technique of sample collection will therefore not have a major effect on the resolution of the  $^1\text{D}$  chromatography.

## 4.5 Pressure Relief

No matter the kind of SFC system one uses, at some point the eluate needs to be depressurized. This can be before or after the detector. If depressurization happens after the detector, then the details of the mechanism doesn't matter much, because

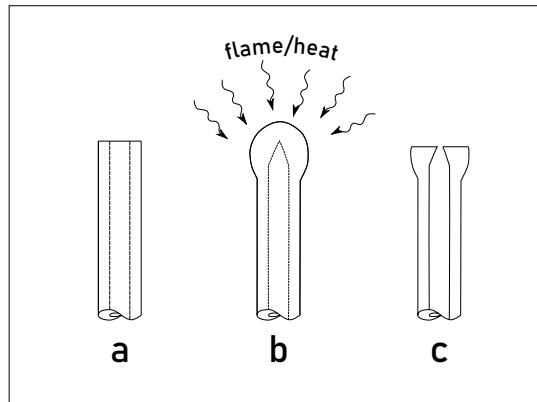


FIGURE 4.6: The steps of making a Guthrie restrictor. (a) Cut length of quartz capillary (b) Heat with flame to soften glass and create internal cone (c) Grind down end to expose orifice of the appropriate size.

*(fig:restrictor)*

the information has been obtained and the eluate can be discarded. If, as in our case, the depressurized eluate needs to still pass the detector, it is important to not lose the resolution achieved by the column. This means that the design of the depressurizer requires some care.

The main concern in depressurizer design is premature desolvation of analytes. This causes **discrimination**, which is the differential treatment of substances where identical treatments are required. In particular, compounds of higher molecular weight tend to desolve first from the supercritical fluid, and might precipitate in the wrong place if the depressurizer is not designed to prevent it.

Depressurization can be done either statically or dynamically. In a dynamic system there is an active element that controls the flow of the eluate in such a way as to maintain the pressure upstream of the active element of the controller. Such a device is often called a **back pressure regulator**, and is usually an electromechanical device with digital control. They tend to be complex and expensive.

In static depressurization there is no active pressure control. A simple **restrictor** is used to limit the flow between the high pressure of the SFC and the low-pressure outlet. Textbooks often discuss different restrictor designs. (The book by Luque de Castro, Valcárcel, and Tena 1994, provides an example.). Our preferred depressurizer was the ‘integral’ or Guthrie design (Guthrie and Schwartz 1986). Figure 4.6 shows the steps in manufacturing the Guthrie restrictor.

We found this restrictor robust and fairly simple to manufacture. It was possible to adjust the flow to a given flow rate. This design of restrictor should eliminate discrimination, because the decompression takes place in a very short space.

However, the Guthrie design proved prone to blockage. These blockages could not be eliminated by incorporating a  $0.5\text{ }\mu\text{m}$  filter before the restrictor, which prompted an investigation into the cause of blockages. To rule out the possibility that it was particles that caused the Guthrie restrictor to become blocked, we first had to determine the diameter of the orifice. This proved to be harder than expected: optical microscopy was not able to give a simple, unambiguous measure of the orifice diameter. Scanning electron microscopy (SEM) showed that the orifice diameter of a Guthrie restrictor is about  $10\text{ }\mu\text{m}$ , as shown in Figure 4.7. This makes it very unlikely that particles with an origin in the SFC system blocked the restrictor: even the  $3\text{ }\mu\text{m}$  particles from the column packing material should not block this restrictor.

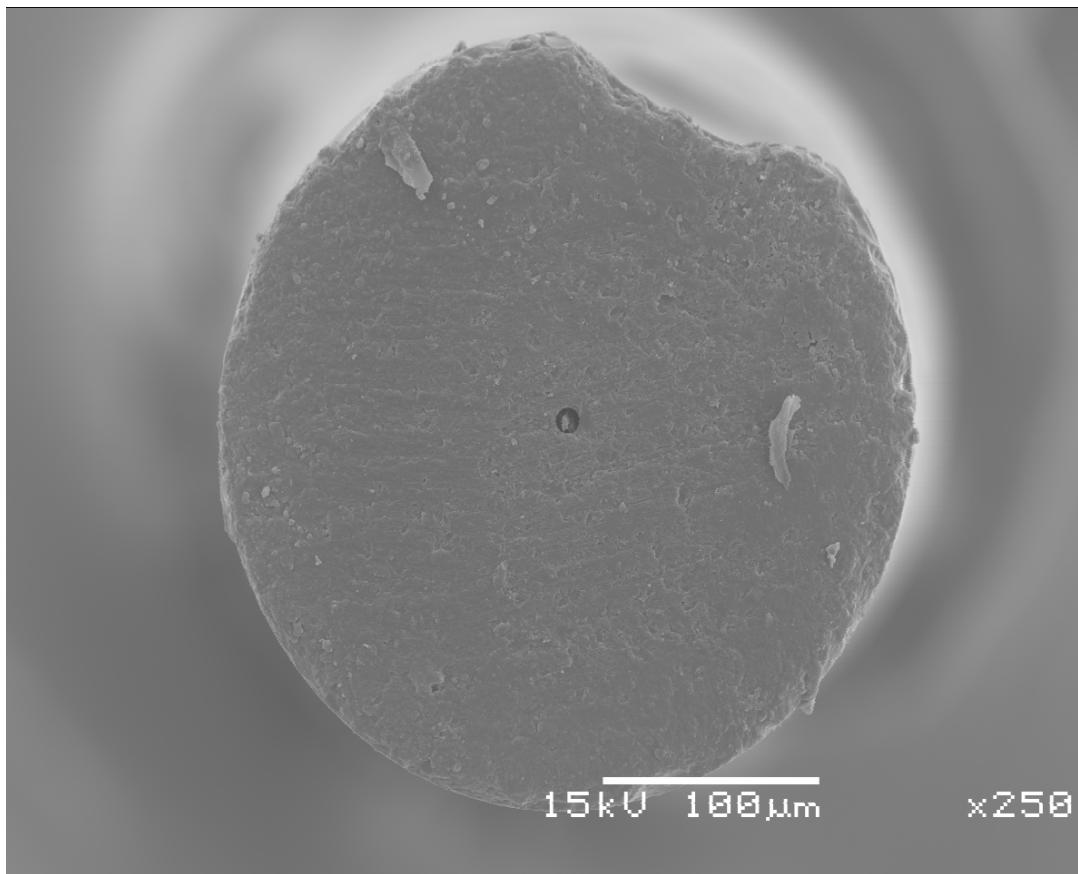


FIGURE 4.7: An electron micrograph of a restrictor tip, showing the size of the orifice.

restrictor orifice)

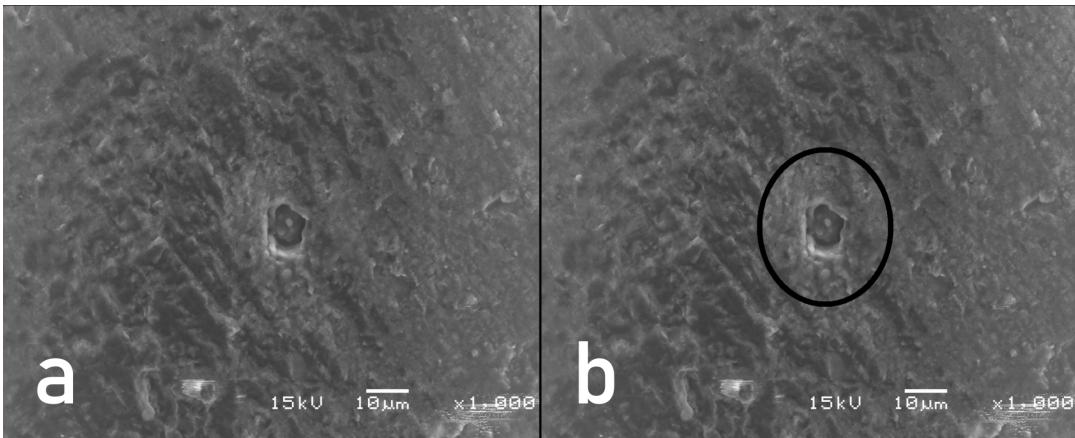


FIGURE 4.8: An electron micrograph of a blocked restrictor orifice. (a) Original image (b) Image with black circle showing estimated original orifice location

`fig:restrictorblockage`

2440 Experience had taught us that the restrictor did not block if the flow was continuous. It was only when the pressure in the restrictor cycled that blockages occurred. To eliminate the possibility that it was material from the stop valve (see Section 4.4.2) that caused the blockages, we cycled the pressure by switching the pump on and off, allowing the pressure to bleed off through the restrictor. This way of cycling the pressure did not prevent blockages.  
 2445

2450 Examining the blocked restrictors with an electron microscope revealed that the restrictors became blocked by a soft material. Backscatter SEM mode allows the energy of X-rays to be measured by energy-dispersive spectrometry, which yields information about elemental composition. While much care must be taken before this information can be used for quantitation, it revealed that the deposited material contains significant quantities of carbon and oxygen, and possibly some chlorine. This points to the probability that the material blocking the column is organic in nature, and possibly polymeric. See Figure 4.8.

2455 We could not determine the origin of the material, but it seems likely to be the column, because cycling the pressure without a column did not cause a blockage. We eliminated the possibility of blockage by sample material by using a brand new column. We suspect that the blocking material might be remnants of surfactants used in the synthesis of the silica gel stationary phase. These surfactants are leached out of the column packing by the highly diffusive supercritical fluid mobile phase, and  
 2460 they then precipitate in the restrictor when the pressure drops and the compounds desolvate. Repetitive pressure cycles causes a build-up of material which gradually blocks the orifice.

2465 The smallness of the orifice contributes to the plugging problem. With a diameter of 50 μm, a solid sphere that fits in this capillary will have a volume of 0.065 nl, or 65 pl. This means that nanogram quantities of material can easily block the restrictor.

2470 Being satisfied that an unfortunate combination of restrictor design and column packing was the likely cause of the blockages, we chose a different restrictor design. The choice was a simple linear restrictor and we trusted that heating the end of the restrictor would prevent discrimination. The final linear restrictor was 800 mm long and had an internal diameter of 0.050 mm.

## 4.6 Detector

When SFC was first developed, a capillary column was the usual column, and an FID was the usual detector. In current use, however, packed columns used with mobile phase modifiers predominate and flame detectors have lost their place: the high concentration of modifiers in the mobile phase would swamp the signal from analytes or saturate the detector. Therefore modern SFC uses predominantly UV/Vis optical detectors.

In the supercritical fluid chromatograph described in this chapter there was no dedicated detector: the role of the detector was taken by a gas chromatograph.

This gas chromatograph collected fractions from the SFC, and separated them in fast chromatographic runs with an FID detector, yielding comprehensive 2D chromatograms. This chromatograph-as-a-detector is described in Chapter 5.

2475

2480



*There are no significant technical limitations to column temperature programming in the order of a few hundred degrees per minute and equally rapid cool-down rates.*

Wolfgang Bertsch, 1997

# 5

## Instrumentation: Fast temperature programmed gas chromatography

2485

*(Chapter 5) The topic of this thesis is the development of a comprehensively coupled (supercritical fluid  $\times$  gas) chromatograph and its application to the analysis of biodiesel. The discussion on the experimental work divides naturally into two parts: the topic of the previous chapter is supercritical fluid chromatography (SFC) and of this chapter it is gas chromatography (GC).*

2490

### 5.1 Speed of analysis

The coupling between two chromatographs can be called **comprehensive** if it meets the following criteria (as discussed in Section 2.3.2):

- 2495 1. Every part of the sample is separated by two distinct chromatographic processes.
- 2. Equal percentages of all sample fractions are separated by the second process.
- 3. Compounds resolved by the first dimension separation remains resolved.

In principle, such a coupling could be implemented by manually or mechanically collecting equal-sized fractions from a first  $^1\text{D}$  chromatograph, and then injecting a portion of each fraction into a different ( $^2\text{D}$ ) chromatograph. In practice, such an approach would be slow, labour-intensive, expensive, and error-prone. But reliable devices that can repeatedly collect and re-inject fractions of eluate were invented. These devices became known as **modulators** and made comprehensively coupled chromatography practical. In continuous flow coupling, such as that found in GC $\times$ GC, the fraction collection and  $^2\text{D}$  injection happens during the uninterrupted  $^1\text{D}$  separation. In stopped-flow coupling, as used in the SFC $\times$ GC chromatography described here, the  $^1\text{D}$  SFC separation is stopped once a fraction has been collected, the fraction is separated on the  $^2\text{D}$  column, and then the SFC separation is restarted and the next fraction collected.

2500

2510 The rate at which fractions of the  $^1\text{D}$  separation are collected is determined by the peak width on the  $^1\text{D}$  chromatogram. To get adequate peak detection in the  $^1\text{D}$

chromatogram, at least 3 fractions of each  $^1\text{D}$  peak should be collected (Murphy, Schure, and Foley 1998), but more is better. For examples, if the peaks are 3 s wide the collection rate should not be less than 1 Hz. The rate at which fractions of the  $^1\text{D}$  separation is collected is known as the **modulation rate**, and its inverse is the **modulation period** ( $P_M$ ). In continuous-flow coupling (e.g. as found in GC $\times$ GC), the  $^2\text{D}$  separation should be completed in a time less than the modulation period. Evidently, this means that the  $^2\text{D}$  separation should be faster than the  $^1\text{D}$  separation.

In stopped-flow coupling, the duration of the  $^2\text{D}$  separation is largely decoupled from the peak width of the  $^1\text{D}$  separation and, by implication, the  $^1\text{D}$  modulation period. In terms of the third criterion of comprehensive coupling listed above, as long as the resolution obtained in the  $^1\text{D}$  separation is not lost during the stopped-flow period, the duration of the  $^2\text{D}$  separation can be as long as it needs to be. In SFC the longitudinal diffusion is low, so that in principle the  $^2\text{D}$  GC runs in SFC $\times$ GC need not be remarkably fast. In practice, however, **sample throughput** of a chromatographic method needs to be acceptable. One of the major determinants of sample throughput is **run time**, the time it takes from when the sample is injected until the final compound has eluted. The judgement of what is acceptable depends strongly on context: for finding the proverbial biomarker for Alzheimer's disease very long run times might be acceptable, but for routine analysis of a commodity biodiesel acceptable run times would probably be short. Therefore, when developing a practical SFC $\times$ GC chromatograph the aim should be to make run times as short as possible.

The time it takes for a stopped-flow SFC $\times$ GC run can be calculated from

$$t_T = t_{1D} \left(1 + \frac{t_{stopped}}{P_M}\right)$$

where  $t_T$  is the total time,  $t_{1D}$  is the run time the unmodulated  $^1\text{D}$  run would take,  $P_M$  is the **modulation period**, and  $t_{stopped}$  is the duration of the SFC stopped-flow state, during which the  $^1\text{D}$  run is completed. With continuous-flow coupling (as in GC $\times$ GC)  $t_{stopped} = 0$ , so that  $t_T = t_{1D}$ .

Examining the expression shows that we can decrease  $t_T$  by increasing  $P_M$ , but there is limit to this: if the modulation period becomes too long, the separation obtained in the  $^1\text{D}$  separation will be lost. So, the only way to decrease the total run times is to reduce  $t_{stopped}$ , the GC run time. For example, if a typical SFC run takes 20 min, and fractions are collected for every 5 s of SFC run, it means there will be  $20 \times 60 / 5 = 240$  fractions collected. Each of these fractions must be injected into a GC chromatograph. If each GC run took 1 min, the SFC $\times$ GC run would last 240 min = 4 h. This is a very long run time, and in the context of commodity biodiesel analysis the sample throughput of such a chromatograph would probably be judged unacceptable.

Reducing the run time of the  $^2\text{D}$  GC chromatography in SFC $\times$ GC means applying the concepts of **fast chromatography**. Fast chromatography is the discipline that studies the theories, experiments, and technologies that shorten chromatographic run times.

## 5.2 Fast gas chromatography theory

Improving the speed of chromatographic separations is an aspect of **optimization**. Optimization can be defined as “the action of making the best or most effective use of a situation or resource” (Oxford University Press 2019). “Making the most effective

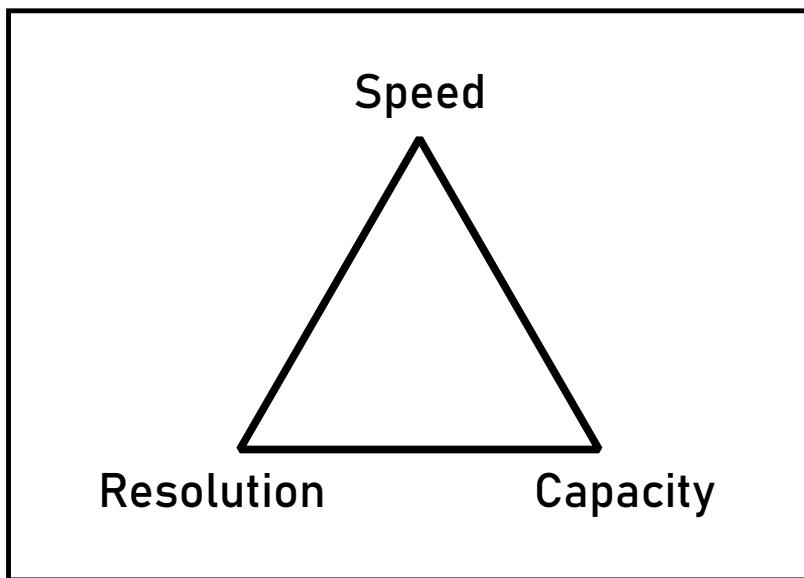


FIGURE 5.1: The chromatographer's trilemma.

(fig:trilemma)

use" means that optimization has a **goal**, which can be characterized by a numerical **criterion** while subject to some **constraints**. The relative and variable amount of "resources" used are **optimizing parameters**.

2560 There are two kinds of optimization in chromatography: selectivity and kinetic. The difference between them is not the goals of the optimization, but how the optimizing parameters are selected. **Selectivity optimization** is the process of selecting the best combination of stationary phase and analyte. The optimization parameters are the different stationary phases and derivitization methods available. **Kinetic 2565 optimization** is the process of selecting the best combination of carrier gas, temperature, flow rate, column length, column diameter, and stationary phase thickness. This differentiation between selectivity optimization and kinetic optimization is not made on theoretical grounds, but practical reasons: selectivity optimization often involves the discrete and expensive parameter of selecting a column, whereas the 2570 parameters involved in kinetic optimization can often be adjusted continuously and inexpensively. Selectivity optimization is guided by chemical knowledge; kinetic optimization is guided by **chromatographic rate theory**.

Every chromatographer involved in a kinetic optimization has met the chromatographer's trilemma (See Figure 5.1): Any chromatographic method that involves decisions about speed, resolution, and capacity can maximize only one at a time<sup>1</sup>. The fastest chromatography will have low capacity and low resolution, the chromatography giving the highest resolution will be slow with low capacity, and the chromatography of a sample with a high concentration of analyte will be slow and have low resolution (Klee and L. M. Blumberg 2002). One way the chromatographer 2580 can ease the grip of the trilemma is to decide what factor is least important for the current problem, pick a value for that is likely to be satisfactory for most conditions, keep it fixed, and then find an acceptable compromise between the other two. In the development of the SFCxGC we decided to work with a fixed sample capacity: we chose a capillary column with a 0.25 mm internal diameter and a 0.25  $\mu\text{m}$  stationary

<sup>1</sup>This trilemma applies only to capillary chromatography. When using packed columns, the capacity can be readily increased by using a column with a larger diameter and a larger amount of packing.

2585 phase. With the sample capacity held constant, it is now possible to optimize the system so that it gives adequate resolution in the shortest possible time.

2590 The **efficiency** of a column is given by the relationship  $E = \frac{t_R}{\sigma}$ , where  $t_R$  is the **retention time** of a peak, and  $\sigma$  is its standard deviation, related to the **peak width** (Leonid M. Blumberg 2018). The efficiency is a measure of how much the peak spread during its transport through the column relative to distance it travelled, using time as a proxy for distance. Efficiency is related to the more traditional but less intuitive **number of plates**  $N$  through the expression  $N = E^2$ . The length of the column ( $L$ ), divided by the square of the efficiency gives the **plate height** ( $H$ ). Plate height is a measure of how much each unit of length of the column contributes to the separation: a lower  $H$  means that the separation is more effective. The value of  $H$  in chromatographic theory is that equations can be developed that express  $H$  as functions of the adjustable parameters of kinetic optimization. Such functions help predict suitable chromatographic conditions and set limits to the expected performance of the separation.

2600 When doing chromatography, any optimization goal must contain resolution as part of the criteria; after all, the purpose of the activity is the separation of chemical compounds. Traditionally, the first optimization taught is the optimization of  $H$  as criterion, with the rate at which the carrier gas passes through the column as the variable parameter. The rate can be characterized by either the average gas velocity ( $\bar{u}$ ) or the gas flow rate ( $F$ ). Traditionally  $\bar{u}$  have been used, because it can be conveniently calculated from simple measurements:  $\bar{u} = L/t_m$ , where  $t_m$  is the **void time**, which is easily determined by measuring the retention time of an unretained compound. In modern gas chromatographs  $F$  can be continuously controlled by electronic pneumatic control (EPC) systems, and has become a more convenient parameter.

2615 Blumberg has developed the theory of optimized flow rate (Leonid M. Blumberg 1999). He found that the function that describes plate height in terms of flow rate can be simplified to two cases: a high-pressure drop and a low-pressure drop. (A high pressure drop is defined as the case when the difference in pressure between the inlet and outlet is much greater than the outlet pressure,  $\Delta p \gg p_o$ ). When the pressure drop is low

$$H = \frac{b}{F} + (c_1 + c_2)F$$

and when the pressure drop is high

$$H = \frac{9}{8} \left( \frac{b}{F} + c_1 F \right) + C_2 \sqrt{F}$$

where  $c_1$ ,  $c_2$  and  $C_2$  are constants independent of  $F$ .

2620 When the layer of stationary phase is thin, also known as the **thin-film case**, then  $c_2 = 0$  and  $C_2 = 0$ , and

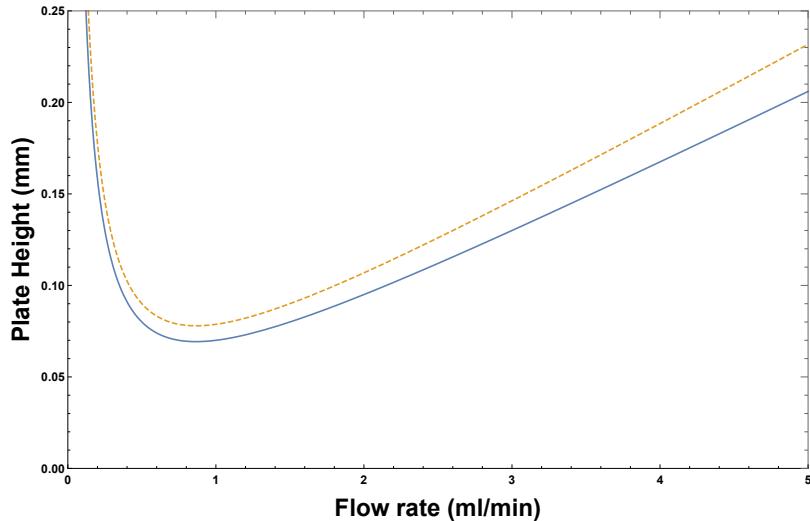
$$H = \frac{b}{F} + c_1 F$$

for the low pressure drop case and

$$H = \frac{9}{8} \left( \frac{b}{F} + c_1 F \right)$$

for the high pressure drop case. The low pressure drop case will be immediately recognized as the **van Deemter equation**, and the high-pressure drop case differs from it only by a factor of 0.125. (See Figure 5.2)

2630 It is clear that the function has a minimum, and the flow rate at this minimum  $H$  can be called  $F_{Hmin}$ . This flow is called **efficiency-optimized flow**, EOF. A remarkable feature of this theoretical result is that it is independent of column length (Leonid M. Blumberg 1999). In contrast, the **efficiency optimized velocity** (EOV) depends on the column length when the pressure drop is high (Leonid M. Blumberg

FIGURE 5.2: A plot of  $H$  vs.  $F$ .

&lt;fig:VanDeemter&gt;

1997). Because  $E = \sqrt{\frac{L}{H}}$  it means that any necessary resolution can be obtained  
 2635 by making the column as long as necessary. This theoretical result is supported by experiment: the world record for the longest capillary column is 1300 m. "This column produced excellent resolution of a petroleum sample, but took 3 hours to elute methane and 3 days to elute the last peak" (Ferguson 2013). Reducing the column length while keeping EOF will reduce the column efficiency, although the plate height will remain constant.  
 2640

In this study the aim is to reduce the time it takes to complete a run, and therefore the question is not what the best possible efficiency is, but what the shortest time is in which a specified efficiency can be maintained. The remarkably simple theoretical result is that under **speed-optimized flow** (SOF),  $F = \sqrt{2}F_{Hmin}$  (Leonid M. Blumberg 1997). Because this flow is higher than the EOF,  $H$  is larger, but this can be compensated for by increasing  $N$ , which is implemented by using a longer column. For a given solute, column diameter, stationary phase and carrier gas there is no flow that will yield a faster separation without reducing the efficiency of the separation.

The remaining parameters that can be manipulated to kinetically optimize for speed are column diameter and carrier gas diffusion. The theoretical details will not be discussed here, but it will be clear that optimizing for separation speed tends towards the selection of narrow-bore capillary columns and using hydrogen as a carrier gas. The theoretical study of optimization gives great insight into which parameters to choose, but the actual calculation can be tedious, so for practical purposes the use of online method translation calculators is recommended (Restek Corporation 2014). The improvement possible in separation speed by using method translation calculations is remarkable, but still falls short of the speed that is needed for comprehensively coupled chromatography.

SOF improves the speed of analysis by finding the combination of parameters  $F$  and  $L$  that will give the fastest separation under the constraint of constant  $H$ . If the constraint on  $H$  is relaxed, a faster separation can be achieved. Assuming that it is not possible to optimize the selectivity further (for example because the stationary phase cannot be changed), there are only two ways to relax the plate height constraint: **selective detection** and **sample cleanup**. In selective detection a detector is

chosen that only responds to the analytes, and not to any interfering compounds. For example, an **electron capture detector** (ECD) responds strongly to halides, so a GC-ECD chromatogram of a soil extract will be much simpler than the equivalent GC-FID chromatogram, which means that the GC-ECD analysis of chlorinated pesticides in soil can be much faster than the equivalent GC-FID analysis. GC-MS using **single ion monitoring** (SIM) also enables faster chromatography by simplifying the chromatogram. In sample cleanup the efficiency constraint can be relaxed because the sample becomes chemically simpler. For example, a sample might be passed through a **solid phase extraction** (SPE) cartridge before it is injected into the chromatograph. Interfering compounds are retained in the cartridge, which simplifies the sample. A simpler sample requires lower efficiency, so the SOF can be higher.

The high speed chromatography that is used in GC<sub>x</sub>GC is possible because any fraction collected from the <sup>1</sup>D separation is considerably simpler than the original sample, and therefore the efficiency constraint is relaxed, which allows a much higher SOF and therefore a very fast <sup>2</sup>D separation. Even at SOF, there might still be excess efficiency, which could be traded for higher speed. Since  $E^2 = \frac{H}{L}$ , this higher speed can be obtained by increasing the flow beyond SOF (decreasing H), or by using a shorter column (decreasing L). Optimization with the goal of the best speed/efficiency ratio using column length and flow as optimizing parameters, theory and experiment shows that flow nearer the SOF combined with shorter columns give better performance (Klee and L. M. Blumberg 2002; Reed 1999) than flow much faster than SOF combined with longer columns.

### 5.3 The General Elution Problem.

The theory developed above is for kinetic optimization: it assumes a single **retention factor**,  $k'$ . This means that the optimization is focused on the kinetic conditions required to optimally separate two compounds eluting after each other, which have similar retention behaviour. If a sample were to contain only these two compounds, then all would be well. But if another pair of compounds, with a mutually similar retention behaviour but retention behaviour very different from the first pair were to be separated, the optimum conditions will be very different. So not all the compounds in complex mixture can be optimally separated with a single set of conditions. This is known as the **general elution problem** (Skoog, Holler, and Crouch 2007, p. 779). The general elution problem makes speed and resolution mutually exclusive: a fast chromatogram will not meet satisfy resolution constraints, and a fully resolved chromatogram will not satisfy speed constraints.

The solution to the general elution problem is to adjust the conditions during the run so that the retention factors  $k'$  stay constant during the run. This strategy is implemented in LC by **gradient elution**, and in GC by **temperature programming**. In isothermal GC less volatile compounds exhibit higher  $k'$  than more volatile compounds. If kinetic conditions are optimized for the more volatile compounds, then they will elute optimally, but the heavier compounds will elute with with excess resolution. If — after the elution of the more volatile compounds — the temperature of the column is increased, then  $k'$  of the less volatile compounds will increase, so that the selectivity conditions will now match the optimized kinetic conditions, and their elution will be optimal. By careful programming the temperature can be continuously controlled so that the selectivity conditions always match the kinetic conditions, and all the compounds elute under optimum conditions. The technology for

temperature programming in GC is mature, and every column manufacturer specifies temperature limits for isothermal and ramped temperature programs for every stationary phase.

In GC $\times$ GC the general elution problem appears in the form of **wrap-around**, defined as “the occurrence of second dimension peaks in subsequent modulation sequences, caused by second-dimension retention times that exceed the modulation period of a comprehensive two-dimensional system” (Mariott, Schoenmakers, and Wu 2012). In GC $\times$ GC the  $^2\text{D}$  separation is isothermal, but because GC separations are predominantly determined by volatility, in the  $^2\text{D}$  separation the retention factors have similar values and wrap-around is just a nuisance. In SFC $\times$ GC, where the group-type separation on the  $^1\text{D}$  column yields fractions that contain compounds with a wide range of volatilities (A. Venter, E. R. Rohwer, and Laubscher 1999b), isothermal GC will have unacceptable wrap-around. Therefore, temperature programming is necessary for successful SFC $\times$ GC.

Guided by the theory described above, the fast GC used as the  $^2\text{D}$  separation in the SFC $\times$ GC chromatograph was designed to use a short GC column (1 m long) with flow higher than SOF that follows a correspondingly fast temperature program.

## 5.4 Temperature ramp rates

(sec:RampRates) Early experimenters realized the importance of temperature control in GC, and used vapour baths (Desty and Whyman 1957) or oil baths (Eggertsen, Knight, and Groenings 1956) to control temperature in their experiments, but they quickly realized that leaks or cracks could allow oil into the column. Oil that entered the column would then contaminate the stationary phase, rendering it useless. If air is used to transfer heat to the column, the gas pressure inside the column is higher than the atmospheric air pressure outside the column, so leaks let carrier gas escape but does not allow contaminants to enter. Therefore, in the modern gas chromatograph the convention is to control the temperature of GC columns by keeping them immersed in an air bath with a precisely controlled temperature.

A temperature program is traditionally described using a series of **ramps** and **holds**. A ramp is a period of time during which the temperature changes, almost always increasing. The **ramp rate** is the rate at which the temperature changes. A hold is a period of time during which the temperature is kept constant.

When developing a chromatographic method, at what rate must the temperature of a column increase so that the selectivity conditions match the kinetic conditions? Blumberg and Klee (L. M. Blumberg and Klee 2000) recommend that a good initial temperature ramp rate is  $10\text{ }^\circ\text{C}$  per void time ( $t_m$ ). For long columns the void times are long, and the ramp rates can be low. As an illustration, Figure 5.3 shows the temperature ramp rates of a state-of-the-art chromatograph. By contrast, for short, narrow-bore columns, the ramp rate needs to be thousands of  $^\circ\text{C min}^{-1}$ .

The low ramp rate of conventional air baths is caused by three factors:

- The low heat capacity of air
- The poor thermal conductivity of air
- The mass of the oven that needs to be heated.

There is very little that can be done about these matters. It would be possible to switch oven gas to hydrogen for higher thermal conductivity, and construct a low-mass oven using, say, advanced resins for construction, but this would probably

**Table 1. Typical 7890B GC Oven Ramp Rates**

Temperature range (°C)	120 V Oven* rates (°C/min)	Fast ramp rates** (°C/min) Dual-Channel	Single-Channel***
50 to 70	75	120	120
70 to 115	45	95	120
115 to 175	40	65	110
175 to 300	30	45	80
300 to 450	20	35	65

\* Results obtained with line voltage maintained at 120V

\*\* Fast ramp rates require power > 200 volts at > 15 Amps.

\*\*\*Requires G2646-60500 oven insert accessory.

FIGURE 5.3: The temperature-rate table from the Agilent 7890B chromatograph data sheet (Agilent Technologies 2019).

(fig:RampRate7890B)

involve significant safety and cost issues. What is required is a complete redesign of  
2760 the heating principle.

#### 5.4.1 Resistive heating

Fortunately, the heating rate problem has a technologically simple solution, **resistive heating**<sup>2</sup>. When a constant electric field is applied to a metal, the free electrons in the metal will be accelerated by the electric field. But the electrons are within the crystal lattice of the metal, where their mean free path is very short. The electrons will therefore collide with atoms in the crystal lattice, scattering inelastically. The energy lost in the inelastic collisions will increase the vibrational frequency of atoms, and this energy will appear as heat.  
2765

The number of electrons and their average (drift) speed in combination is described by a measure called the **current** ( $I$ ), and the electric field is can be described by the applied **voltage** ( $V$ ). The current  $I$  is proportional to the applied voltage, and the ratio  $\frac{V}{I}$  defines a proportionality constant  $R$ , called the resistance, which is a function of the conductor's material and dimensions. The total power dissipated to heat ( $P$ ) is given by  $P = IV$  or, equivalently,  $P = I^2R$  or  $P = \frac{V^2}{R}$ .  
2770

Applying a voltage  $V$  to a metal element close to a chromatographic column will heat the metal, and therefore the column in contact with it. The rate at which the piece of metal heats up depends on the power dissipated, the mass of the metal, and its heat capacity. If the volume of the metal is small enough, and the current high enough, the temperature of the metal element (and with it the temperature of the column) will increase at a rate high enough to be useful in chromatography. By suitable manipulation of  $V$ , then, the temperature of the column can then be changed at any desirable rate.  
2775

This technology has been reviewed (Wang, Tolley, and Lee 2012; M. R. Jacobs, Hilder, and Shellie 2013; Miranda 2010), and a few technologies for resistive heating of capillary columns have emerged:  
2780

- Direct heating of a metal column or a column coated with a metal layer
- Collinear heating
- Coaxial heating

<sup>2</sup>Alternative methods of heating by electromagnetic fields are **inductive heating** and **dielectric heating**.

The first SFC $\times$ GC work done (Andre Venter and Egmont R. Rohwer 2004) used a directly heated metal column. While this approach proved the concept, experience showed two shortcomings. Firstly, metallic columns are usually designed with specific high-temperature applications in mind. This means that metal columns are not available with all the stationary phases available in fused-silica columns. Secondly, it was harder than expected to control the temperature accurately. The temperature was determined by a thermocouple glued to the column, which was sensitive to local variations.

An example of collinear resistive heating is Agilent™'s 'low thermal mass' column, which includes collinear heating wire and a collinear sensing element bundled with a short fused silica column. This approach requires that the collinear heating element be wrapped in close contact with the column.

For the work presented in this thesis, a coaxial heater was used. This heater is in the form of a thin-walled stainless-steel tube that carries the electric current. It was made of a 940 mm length of SAE 304 stainless steel with an outside diameter of 1.06 mm and an inside diameter of 0.80 mm. The column was threaded inside the stainless steel tube, which put it in close contact with the heater, giving reliable heat transfer. The coaxial heater can be coiled to fit inside a conventional GC oven.

This coaxial design has the advantage that the column can be changed without changing the resistive heater. This means that there is no need to re-calibrate the heater when columns are changed.

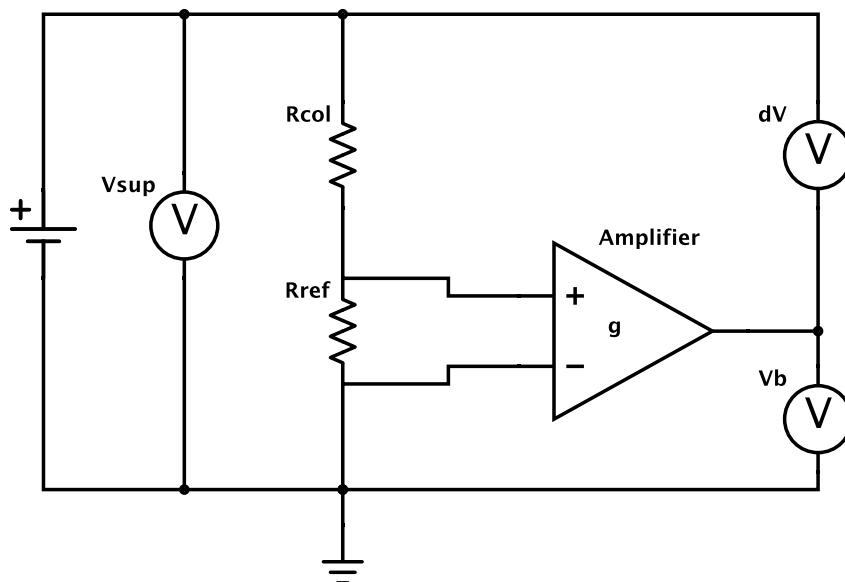


FIGURE 5.4: Electric circuit diagram of the coaxial heater.

`{fig:HeaterDiagram}`

The electrical resistance of a conductor is determined by its dimensions, the material it is made of, and the temperature of that material. In a conductor of given dimensions, therefore, a knowledge of the resistance implies a knowledge of the temperature. By following changes in the resistance, one can determine changes in the temperature, and by comparing resistance at certain temperatures with known temperatures, one can get a calibrated temperature from a given resistance.

High-powered resistive heating implies low resistance. Measuring the *absolute value* of low resistance is technologically challenging (Dyos 2012). But a simple electronic circuit can be used to *compare* the resistance of the coaxial heater with that of

a reference resistor (Figure 5.4). The circuit is supplied by a voltage  $V_{sup}$ , in general an unknown value.  $V_{col}$  and  $V_{ref}$  represents the voltage drop over the respective resistors. Because the current  $I$  through the circuit is the same for both  $R_{col}$  and  $R_{ref}$ , it is true that  $\frac{R_{col}}{V_{col}} = \frac{R_{ref}}{V_{ref}}$ , and therefore

$$R_{col} = R_{ref} \frac{V_{col}}{V_{ref}}$$

The voltage drop across  $R_{ref}$  is too small to be directly digitized, therefore it is amplified by an amplifier with gain  $g$ , so that  $V_b = gV_{ref}$ .  $V_b$  is measured, as is  $dV$ , the potential difference between the supply and the amplifier output. We can show that the coaxial heater resistance  $R_{col}$  is a function of the voltage ratio  $\frac{dV}{V_b}$ .

First,

$$V_{sup} = V_{col} + V_{ref}$$

but also

$$V_{sup} = dV + V_b$$

Therefore,

$$V_{col} + V_{ref} = dV + V_b$$

But  $V_b = gV_{ref}$ , so

$$V_{col} + V_{ref} = dV + gV_{ref}$$

$$V_{col} = dV + gV_{ref} - V_{ref}$$

$$V_{col} = dV + V_{ref}(g - 1)$$

$$\frac{V_{col}}{gV_{ref}} = dV/gV_{ref} + V_{ref}(g - 1)/gV_{ref}$$

$$\frac{V_{col}}{V_{ref}} = gdV/gV_{ref} + gV_{ref}(g - 1)/gV_{ref}$$

$$\frac{V_{col}}{V_{ref}} = g \frac{dV}{V_b} + (g - 1)$$

This proves that  $\frac{V_{col}}{V_{ref}}$  is a linear function of  $dV/V_b$ . A quick check for correctness of the expression: for a unity-gain amplifier  $g = 1$ , and  $\frac{V_{col}}{V_{ref}} = dV/V_b$ .

Therefore

$$R_{col} = R_{ref} \left( \frac{gdV}{V_b} + (g - 1) \right)$$

In practice the gain  $g$  might not be completely constant, but show a slight dependence on  $dV$ , so a general expression might be

$$R_{col} = f \left( \frac{dV}{V_b} \right)$$

## 5.5 Calibration

The assumption is that the temperature is a function of the resistance of the coaxial heater, or  $T = g(R_{col})$ . Because  $R_{col} = f(\frac{dV}{V_b})$ , we can see that  $T = g(f(\frac{dV}{V_b}))$ , or, because a function  $g$  of a function  $f$  is a function  $h$  ( $f \circ g = h$ ),  $T = h(\frac{dV}{V_b})$ . Through a calibration procedure  $h$  can be approximated by a polynomial or lookup table.

### 5.5.1 Temperature uniformity

*(sec:Uniformity)* It is highly desirable that the coaxial heater should give uniform heating, but there is no obvious guarantee an electrically heated tube will heat uniformly. We did not analyse the problem theoretically, but the following factors will play a role:

#### Resistivity's dependence on temperature

The resistance of a metal increases with temperature. This means that if one section of the coaxial heater gets hotter than the rest, its resistance will increase. If the current were to remain constant, more power would be dissipated in this section ( $P = I^2R$ ). If more power is dissipated, the temperature will increase, leading to a higher resistivity, leading to higher power dissipation, leading to a higher temperature, in a runaway cycle. In contrast, if the coaxial heater was made of a material of which the resistivity decreases with an increase in temperature, the temperature along the length of the heater would be stabilized by negative feedback. In the case of the supply voltage being held constant a higher resistance in one section would mean a lower current overall, but still a higher power dissipation in the section with higher temperature.

#### Thermal conduction

Each section of the coaxial heater is in thermal contact with its neighbours. If it were to get hotter, the heat will flow from the hotter section to the cooler neighbouring sections. This will tend to even out any temperature differences.

#### Radiation

An object radiates heat, which can be approximated by the Stefan-Boltzmann law:

$$P = A\epsilon\sigma T^4 \quad (5.1) \quad \text{eq:1}$$

where  $A$  is the surface area of the object,  $\epsilon$  is the **emissivity** of the surface,  $\sigma$  is the Stefan-Boltzmann constant, and  $T$  is the absolute temperature of the object.

If a section of the coaxial heater were to become hotter than the other sections, it would therefore radiate more heat. The increased power dissipation by the higher resistance will therefore be partially offset by a higher radiation. This will tend to moderate temperature differences.

#### Convection

When a part of the coaxial heater is hot, it will heat the air around it, through radiation and conduction. This air might then move, through buoyancy or any other force, past another part of the coaxial heater. This second part could then be heated up by the transported air. For example, a vertically mounted coaxial heater can be

2875 expected to develop a temperature gradient from bottom to top, as natural convection lets hot air transfer heat from the lower end of the heater to the upper end. In this way temperature gradients could be established and maintained.

### Examining thermal uniformity by imaging

2880 The effect of the combination of resistivity's temperature dependence, thermal conduction, radiation, and convection on the temperature uniformity of the coaxial heater could be mathematically or numerically modelled, but such an endeavour would fall outside the scope of this project. Experience had not led us to expect any significant temperature non-uniformity, but we welcomed the opportunity to get empirical confirmation.

2885 **Thermal imaging** is the process by which infrared radiation from objects can be captured in a photographic process. At near-ambient temperatures objects emit copious amounts of infrared radiation, and higher the temperature, the more is emitted. (See Equation 5.1.) Cameras designed using specialized optics and sensors allow the capture of that radiation as an image of a scene that shows objects based on their 2890 surface temperature. The technological capability of thermal imaging has improved markedly over the past years.

Using a FLIR™ T660 thermal imaging camera we obtained a video of the coaxial heater executing three consecutive temperature ramps. This camera uses a  $640 \times 480$  focal plane uncooled bolometer array as a detector, which is sensitive to radiation in the range  $7.5 \mu\text{m}$  to  $14 \mu\text{m}$ . Figure 5.5 shows the setup used to record the thermal video. To get reliable temperature information from a thermal image requires calibration or a precise knowledge about the material's **emissivity**. In the absence of such calibration, the obtained temperature measurement is more useful if seen as relative rather than absolute.

2900 The video showed that the three consecutive temperature runs heated the coaxial heater identically. Figure 5.6 is a frame from the video, analysed to give estimates of the temperatures on spots on the coaxial heater. The maximum temperature difference between any two points is  $17.5^\circ\text{C}$ , and there are no marked gradients.

2905 This examination of the uniformity of the coaxial heater helped to dispel fears that unexpected temperature gradients would interfere with the fast gas chromatography.

### 5.5.2 Temperature calibration

When doing temperature-programmed gas chromatography it is desirable to have an absolute measurement of the column temperatures. This makes it possible to 2910 translate and compare methods between instruments. It was therefore necessary to calibrate the temperature of the coaxial heater. Calibration is the comparison of measurement values from a device under test with those of a calibration standard of known accuracy. In the case of the temperature calibration of the coaxial heater, it means comparing the results of the temperature as measured by the coaxial heater with a standard temperature.

The problems of measuring a temperature inside a tube with a bore of  $0.8 \text{ mm}$  are not trivial, and to solve them we examined the available technologies to find the optimum solution.

The following temperature measurement technologies exist:

- 2920 • Liquid-in-glass thermometers



FIGURE 5.5: This photograph shows the setup used to record the thermal video.

ThermalImageSetup>

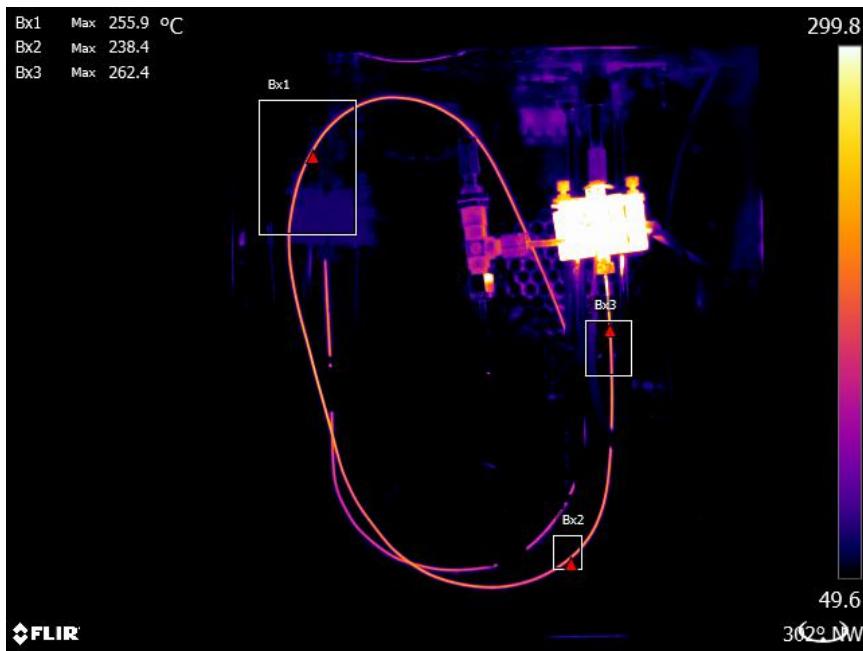


FIGURE 5.6: A thermograph of the coaxial heater at high temperature. It shows that there are no runaway hot spots.

<fig:ThermalImage>

- Sealed liquid or gas sensing instruments and bimetallic sensors.
  - Electrical resistance temperature measurement using metallic sensors
  - Thermistors and semiconductors
  - Thermoelectric temperature measurement
- 2925     • Disappearing filament optical pyrometer
- Photoelectric optical pyrometers
  - Total radiation pyrometers.

Liquid-in-glass thermometry would not be applicable because of the size of the devices, and because they don't give a desirable electrical signal. Sealed liquid or gas sensing elements are also too bulky, and like bimetallic sensors are best used where only an on/of electric signal is required.

In recent years the technology for measuring temperature by radiant energy methods have improved markedly and has become affordable, in the form of thermal cameras. However, at lower temperatures the accuracy of the recorded temperature depends heavily on the emissivity of the measured material. Thermal imaging will also only measure the outside wall surface temperature of the heater, and not the temperature on the inside of the coaxial heater. So, while thermal imaging settled questions about heater uniformity (Section 5.5.1), it was not considered ready to serve as a calibration standard.

2940 This leaves us with resistance temperature measurement with metallic sensors, thermistors and semiconductors, and thermoelectric temperature measurement.

Electrical resistance measurement using metallic sensors might have been feasible, if sensing elements of the appropriate dimensions were commercially available. A further difficulty with this method of temperature measurement is that long, thin 2945 conductors would be needed to connect the sensing element to the electronics. These conductors would add to the resistance measured by the sensing element, requiring complex correction or multi-wire measuring methods.

Thermistor and semiconductor devices could be made small enough, but all 2950 commercially available packages are too large. Besides, the temperature range used in GC ( $-50^{\circ}\text{C}$  to  $400^{\circ}\text{C}$ ) does not fall inside the operating temperature range specified by manufacturers of semiconductor devices ( $-55^{\circ}\text{C}$  to  $125^{\circ}\text{C}$  for military applications) (Department of Defence 1996).

Thermoelectric temperature measurement then remains, and was used to calibrate the coaxial heater. In particular, the **Seebeck effect** was exploited, which is the 2955 observation that a temperature gradient imposed on a metallic conductor will generate an electrical potential along the gradient. Therefore, a circuit of two dissimilar metallic conductors will generate a voltage when there is a temperature gradient along the conductors. The voltage generated is a function of the temperature difference between the hot junction of the two metals and their cold junctions at the 2960 voltmeter input. Such a pair of dissimilar conductors used to generate a voltage is known as a **thermocouple**, and this technology is mature and widely used in industry. Standard thermocouples are constructed from well-characterized alloys that generate predictable voltages for given junction temperatures. Pairs of thermocouple 2965 standardized alloys are known as 'types', of which the general-purpose Type K suited our purpose best. Thermocouple wire can be purchased in a range of gauges, down to  $25\ \mu\text{m}$  in diameter, and the signal processing for thermocouple signals have

been standardized. Thermocouple junctions can be made by welding, crimping, soft soldering, hard soldering, bolting, or simply twisting the wires together (McGee 1988). Because the temperature range to be measured was (-50 °C to 400 °C) soft soldering is not a viable choice, because soft solders have melting points around 200 °C. The possibility of corrosion and mechanical vibration suggest that twisting the wires together will not form a reliable joint, and of course there are no sub-millimetre bolts on the market.

This leaves welding, crimping and hard soldering as methods for making thermocouple junctions. Tools for crimping hair-fine wire are rare, and it is likely that a practical crimped connection will have a diameter many times the diameter of the wire, possibly making it too large to fit.

Hard soldering is usually done with a high-temperature flame, and on contact the flame will rapidly burn the fine wires. The temperature required for hard soldering is still lower than the melting point of the wires, so that hard soldering is not excluded, but we did not have the knowledge or the technology to solve the associated problems.

Welding was found to be an accessible technology for forming small, reliable joints in fine thermocouple wire.

### 5.5.3 Thermocouple Welding

Welding is the process of joining two metal parts by melting a portion of each part, allowing the molten metals first to mix, and then to solidify. This creates a permanent joint between the two metals. Welding is widely practised as an industrial process in applications ranging from shipbuilding to microelectronics. Previous work in our laboratories used capacitive discharge spot welding to melt the spot where two wires crossed. These thermocouples were found to be too fragile for the intended application, so a method was developed to yield a more robust thermocouple.

In the laboratory electricity is the most convenient source of heat. A 24 V direct current, adjustable bench power supply was used to supply the current. The two wires of the thermocouple were twisted together, and the twisted pair was connected to one pole of the power supply. A carbon electrode was connected to the other pole. The carbon electrode was carefully brought closer to the thermocouple pair until a spark jumped across the air gap. When the spark turned into an arc the heat of the arc melted the end of the twisted pair. The molten metal would then contract into a spherical globule, which grew as the arc added more heat. As more of the metal of the wire melted, the globule would move away from the carbon electrode, until the gap became too large to sustain the arc. The current would then stop, leaving a spherical welded bead at the end of the wires. (See Figure 5.7.) The process could be repeated as often as necessary to obtain a bead of the desired size.

It is worth noting that it is necessary to form an arc: if the carbon electrode happens to touch the wire so that a current flows directly from the wire to the carbon electrode the wire rapidly heats up over its length and melts. It is also interesting that it seems necessary to have a roughly broken carbon electrode surface: a polished surface would not generate an arc, or even make electrical contact with the wires. This might be because the graphite used for the electrode was formulated for use in pencils.

The apparatus used to do the welding is depicted in Figure 5.8, and Figure 5.9 shows the image seen through the microscope when the welding is done.

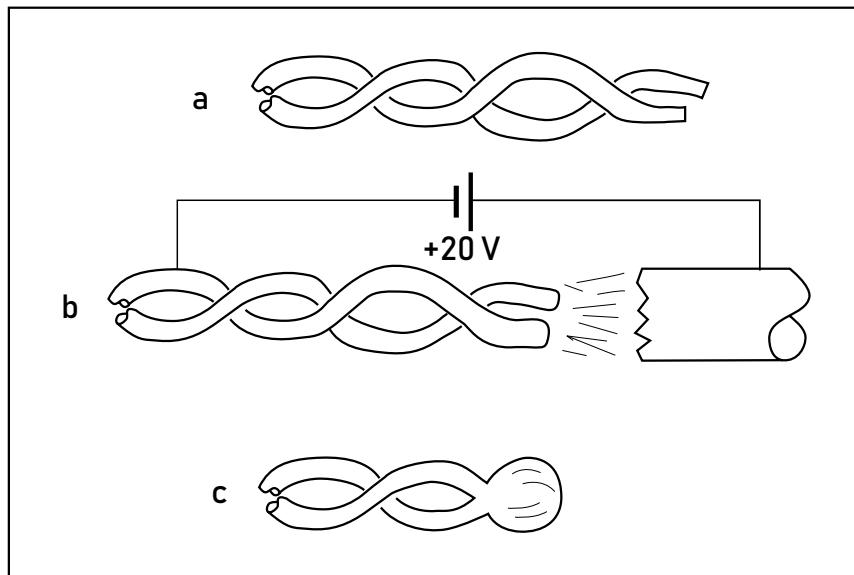


FIGURE 5.7: The process of welding fine wires to make thermocouples  
 (a) Wires twisted together (b) Electric arc heating up the wires (c)  
 Wires welded with a well-formed bead.

*(fig:WeldingSteps)*

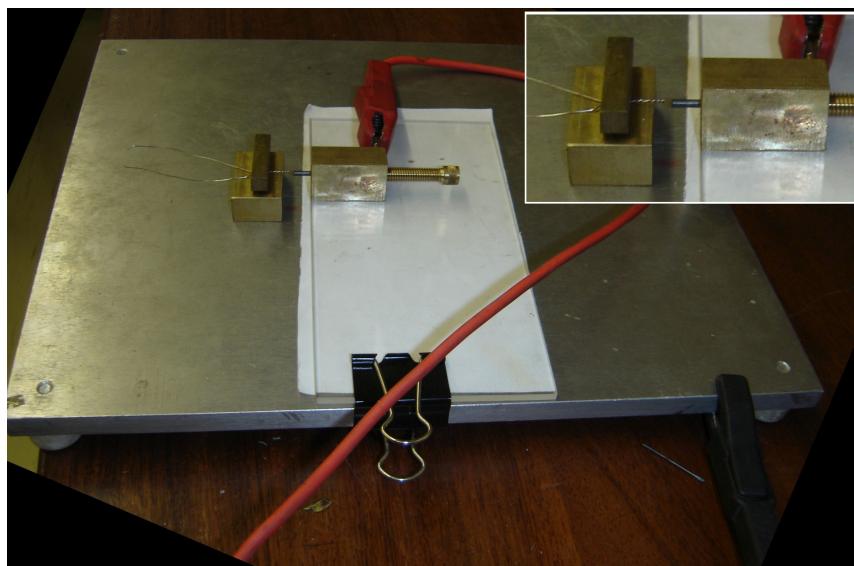


FIGURE 5.8: A view of the fine-wire thermocouple welder, as set up on a microscope base plate. The wire shown in the photograph is much thicker than that actually used. It is shown clamped between the clamping bar and the clamping weight. A thin sheet of acrylic serves to isolate the positive electrode from the negative base. The carbon electrode can be advanced towards the thermocouple twist using the screw. The black clamp at the bottom right-hand corner attached to the base plate and the red clamp attached to the screw housing provide a potential difference of approximately 20 V between the carbon electrode and the thermocouple.

*(fig:FineWireWelder)*

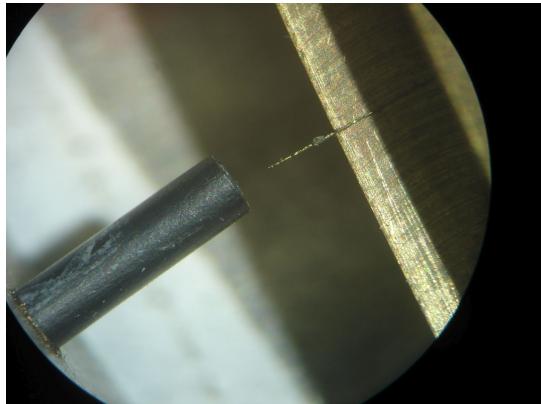


FIGURE 5.9: A microphoto of a twisted wire ready to be welded. The black carbon electrode is 2 mm in diameter.

(fig:TCWeldMicro)

#### 5.5.4 Thermocouple probe construction

- 3015 To measure the temperature inside the coaxial heater required the thermocouple to be inserted into the coaxial heater. For this a fused silica capillary with an inside diameter of 0.25 mm and an outside diameter of 0.4 mm was used to construct a probe. (These capillaries were readily available, in the form of discarded chromatographic columns.) A narrower capillary was used to draw the wires through the probe capillary, in a process described in Figure 5.10. The wires could then be welded to form a thermocouple, and then drawn to the desired position in the probe capillary.
- 3020

#### 5.5.5 Thermocouple interfacing

- 3025 The Type K thermocouple has a sensitivity of approximately  $41 \mu\text{V } ^\circ\text{C}^{-1}$ . This means the voltage generated at the temperature range of interest is too small to be conveniently digitized, and so needs to be amplified. Because the Type K thermocouple is so commonly used in industry, amplifiers have been developed specifically for thermocouple signal conditioning. We chose the AD595 integrated circuit amplifier. The component's data sheet explains its application concisely: "The AD595 is a complete instrumentation amplifier and thermocouple cold junction compensator on a monolithic chip. It combines an ice point reference with a pre-calibrated amplifier to produce a high level ( $10 \text{ mV}/^\circ\text{C}$ ) output directly from a thermocouple signal."
- 3030 (Analog Devices 1999) The output signal of the AD595 can be directly digitized for computer recording.

#### 5.5.6 Calibration procedure

- 3035 The International Vocabulary of Metrology (*International vocabulary of metrology – Basic and general concepts and associated terms (VIM) 2019*) define **calibration** as "[an] operation that, under specified conditions, in a first step, establishes a relation between the **quantity values** with **measurement uncertainties** provided by **measurement standards** and corresponding **indications** with associated measurement uncertainties and, in a second step, uses this information to establish a relation for obtaining a **measurement result** from an indication."
- 3040

In the first step of calibration, the **quantity values** used in the calibration were the temperature value of the thermocouple probe as provided by the thermocouple voltage, the AD595 amplifier, the digitization and the subsequent calculations according

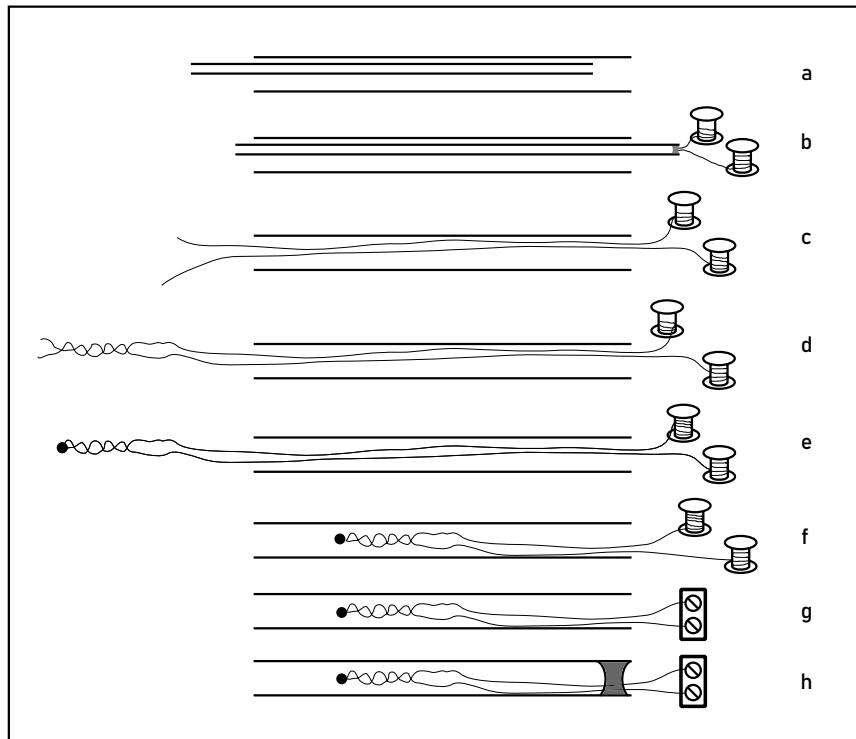


FIGURE 5.10: (a) A narrow capillary is threaded inside a wider one  
(b) The ends of a pair of thermocouple wires are fitted inside the end of the narrow capillary and anchored with cyanoacrylate adhesive.  
(c) The wires are pulled through the wider capillary using the narrow capillary.  
(d) The ends of the wires are twisted together, creating a mutual mechanical anchor.  
(e) The ends of the wires are welded together.  
(f) The wires are pulled back into the thick capillary, locating the junction at the desired position in the capillary.  
(g) The wires are trimmed and connected to a terminal block.  
(h) A drop of cyanoacrylate adhesive is used to anchor the wires permanently in the capillary.

g:FineWireThermocouple)

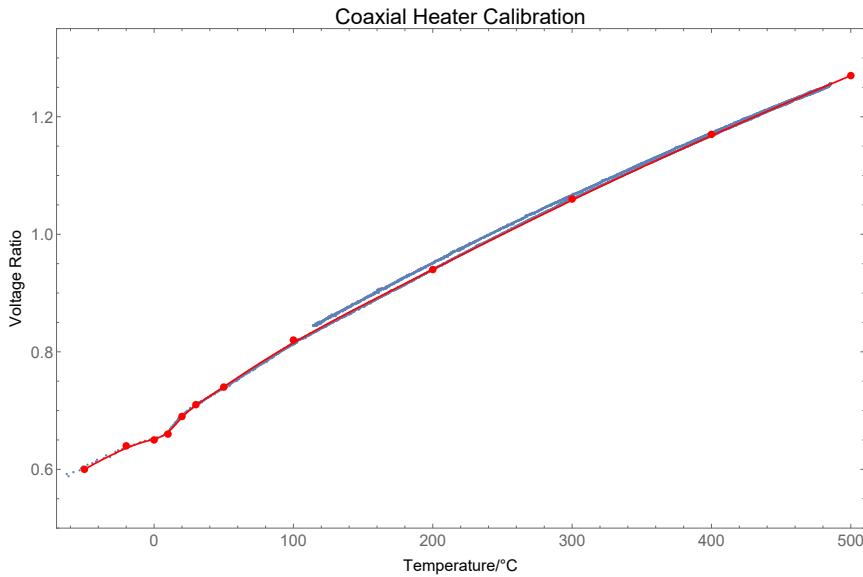


FIGURE 5.11: Calibration curve for the coaxial heater.

:CalibrationCurve>

3045 to the AD595 data sheet. The **measurement standards** were the known responses of the thermocouple (Ripple 1995), the amplifier and the digitization system. The **indication** was the voltage ratio  $\frac{dV}{V_b}$ . As this was a first attempt a calibration, **measurement uncertainties** were assumed to be negligible.

3050 In preparation for calibration the coaxial heater was installed in the oven of the Varian 3300 GC as it would be when in use. The detector was removed, and the thermocouple probe was threaded through the detector stem, through the heated T-piece block, and into the coaxial heater until the thermocouple junction was about half-way between the inlet end and the detector end of the coaxial heater. When the system was set up as it would be during use, the coaxial heater was cooled down, 3055 and a power ramp applied. The temperature of the thermocouple was recorded, together with the voltages  $dV$  and  $V_b$ . A curve could then be plotted of thermocouple temperature  $T_{TC}$  against the voltage ratio  $\frac{dV}{V_b}$  (Figure 5.11). This revealed the shape of the function  $T(\frac{dV}{V_b})$ , and completed the first step of the calibration.

3060 The second step of a calibration is to establish a **measurement result** from an **indication**. It would be traditional to fit a mathematical function such as a polynomial to the curve, but a numerical method was chosen instead. We fitted a B-spline to the curve, and extracted coordinates on the curve from the B-spline (See Figure 5.12). An interpolation function was then used to obtain the measurement result  $T$  from the indication  $dV/V_b$ .

3065 The calibration could be checked by setting the current through the coaxial heater to a minimum, at which not enough heat is generated to affect its temperature. Then the oven of the Varian 3300 GC could be set to a range of different temperatures. Once equilibrium was reached the reported temperature of the coaxial heater and the air bath temperature could be compared and the calibration adjusted.

3070 The technique of constructing a long, thin thermocouple probe allowed the calibration of the coaxial heater, which makes it possible to translate chromatographic methods. The probe has also proven useful in other applications, for example providing overheating in a GC-MS transfer line.

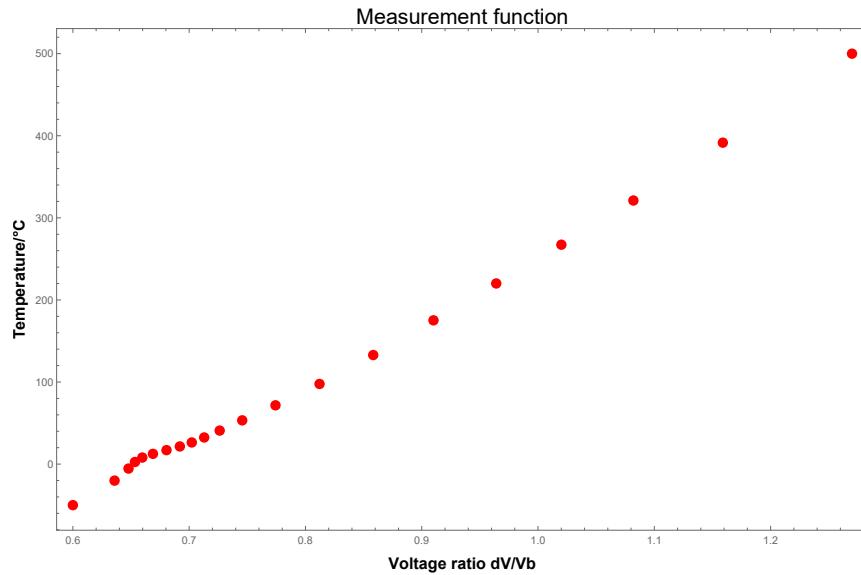


FIGURE 5.12: Measurement curve for the coaxial heater.

&lt;fig:MeasurementCurve&gt;

### 5.5.7 Cold spots

<sec:ColdSpots> For the fastest temperature programming with resistive heating the heating element should be as light as possible and carry the largest necessary current. The current doing the heating must be carried to the coaxial heater using a feed conductor. To prevent the feed conductor from heating up it must have a low resistance, and this low resistance is achieved by making the conductor as 'thick' or as 'heavy' as necessary, meaning that it is constructed of a material with a high mass per unit length.

3080 Good electrical metallic conductors are invariably also good thermal conductors, and therefore, the area around the junction of the feed conductor to the coaxial heater will always have a lower temperature than the nominal temperature of the heater. In capillary GC this is undesirable: a cold spot in a column can wreak havoc with retention times and peak shapes.

3085 Conversely, if an attempt is made to reduce the contact area between the feed conductor and the thin material of the coaxial heater, a hot spot might develop, which could burn a hole in the coaxial heater tube or damage the column.

The electrical connection between the feed conductor and the coaxial heater was 3090 therefore designed in the form of an externally heated block. This block was kept at a higher temperature than the highest expected temperature of the chromatographic temperature program. This prevented the formation of cold spots in the coaxial heater, which might lead to cold spots in the column, while also providing a large contact area so that hot spots do not develop. Each end of the coaxial heater was fitted into one of two heated blocks, where it was brazed in place. Each block had a brass tail, to which an electric feed-wire was soft-soldered. Each block was heated by four 100W Hotset™ electrical cartridge heaters, with dimensions of 6.5 mm × 40 mm. The cartridge heaters were switched on and off by solid state relays controlled from the computer. The temperature of the block was monitored through a thermocouple mounted in a blind hole in the block and the amount of power to the heaters was controlled by **pulse width modulation** (PWM) implemented in software.

### 5.5.8 Cold column

*(sec:ColdColumn)* In a cold GC stationary phase, the retention factor  $k'$  of a particular compound can be very high. This means that the analytes migrate slowly relative to the mobile phase. The lower the temperature, the higher  $k'$  becomes, so that for very low temperatures the migration of the analyte becomes negligible. In effect, the analytes are 'trapped'. This trapping, also called **cryo-trapping** or **cryo-focusing** is useful in various aspects of gas chromatography, such as two-stage thermal desorption or thermal modulation in GC $\times$ GC.

In the SFC $\times$ GC instrument described here, cryo-trapping was used as the second stage of a two-stage modulator. (The stop valve described in Section 4.4.2 represents the first stage.) The column was cooled down to very low temperatures, which trapped any analytes eluting from the first dimension in a narrow band on the GC column. Once the required amount of fraction had been collected, the eluate flow from the first dimension was stopped by closing the stop valve. Then the temperature ramp of the fast GC would start. As the coaxial heater warmed up the column the values of  $k'$  would decrease, and the narrow analyte band would start migrating.

The first SFC $\times$ GC chromatograph cooled the column by using the oven cryo-cooling capability of the Varian 3300 GC (Andre Venter and Egmont R. Rohwer 2004; Andre Venter 2003). The purpose is to cool the GC column to sub-ambient temperatures, normally needed when analysing volatile compounds. In such cases the  $k'$  values at or near room temperature are too low to provide adequate retention, and the cryo-cooling function permits temperature programs to start at sub-ambient temperatures.

The Varian 3300 cryo-cools its oven by injecting liquid carbon dioxide into the oven's air circulation fan. The evaporating liquid carbon dioxide absorbs energy from the air, which lowers the temperature of the air in the oven. A control system controls the amount of carbon dioxide admitted and the amount of heat added through the oven heaters, thereby keeping the oven at the required temperature.

The cryo-cooling function can cool down the oven to cryo-trap analytes, but there are two reasons why it is not suitable for practical trapping in SFC $\times$ GC. The first reason is the quantity of coolant required: doing SFC $\times$ GC runs revealed that about 15 kg of carbon dioxide was consumed per run. A standard cylinder of carbon dioxide contains 33 kg, which implies that a new cylinder would be required every two runs. Such a rate of use is much too high for the intended application of the instrument. The second reason using is that it is much too slow. The time spent on cooling the oven and the column is time that cannot be spent doing separations, and cooling a conventional GC oven takes a lot of time: the Varian 3300's cryo-cooling function took 30 s to cool the column down to a low starting temperature. A commercial forced-convection system ("GC Chaser" supplied by Zip Scientific) improves the cool down time of an Agilent 6890 GC oven, taking 7 minutes instead of 16, cooling down the oven from 350 °C to 30 °C. Cooling the column in an air bath has the same drawbacks of low conductivity and low heat capacity as air-bath heating has (see Section 5.4).

A system was therefore developed that injected liquid carbon dioxide into one end of the space between the column and its coaxial heater, with the other end open to the atmosphere. When the valve opens the space rapidly fills with liquid carbon dioxide, while the pressure drops from 55 atm in the cylinder to 1 atm at the outlet. The liquid boils, absorbing large quantities of heat from the column and coaxial heater in the process, so that their temperatures decrease rapidly. This system solves the speed and coolant consumption problems: because the coolant is in direct

contact with the parts that need to be cooled, the cooling is rapid, and because the coolant is applied where it is needed, only a small quantity is required.

3155 The carbon dioxide for cooling the coaxial heater was introduced through the same heated block that provided the electrical connection. (See Section 5.5.7.) A T-piece design allowed the liquid carbon dioxide to be admitted to the end of the coaxial heater, which was brazed to the block. A micro-union brazed to the block sealed the column's exit port, and the liquid carbon dioxide entered along the side  
3160 of the T (Figure D.1 and Figure D.2). A metering valve allowed the flow rate of the coolant to be adjusted, and a computer-controlled solenoid valve switched the flow on or off.

### Cryogen supply

3165 The carbon dioxide for cooling was supplied by Afrox, in high-pressure cylinders each containing 33kg of technical grade carbon dioxide. Each cylinder was internally equipped with a **dip tube**, a tube that extends from the valve at the top of the cylinder to the bottom of the cylinder. This ensures that when the valve is opened, liquid carbon dioxide is delivered.

3170 Experience taught that for repeatable cooling, the source of liquid carbon dioxide had to be near the solenoid valve. If this was not the case, when the valve was opened initially only carbon dioxide gas would be admitted, followed by a mixture of gas and liquid, and only finally liquid. (This is similar to the common experience of opening a water tap after a municipal water supply interruption: a lot of gurgling and spitting before a reliable stream of water flows from the tap.) Such unreliable  
3175 coolant flow gives unreliable cooling. To solve this problem we installed a reservoir for liquid carbon dioxide on top of the GC. The problem of filling a receptacle with liquid carbon dioxide was described in Section 4.3, so the final design of the reservoir took the form of a coil of copper tube immersed in a circulating coolant (Figure D.3). Mounting the reservoir above the cut-off valve allows the liquid to collect at the  
3180 bottom and allow gas to collect at the top, so that when the valve opens the flow into the coaxial heater contains only liquid.

### 5.5.9 Column mounting

3185 The T-piece blocks described in Section 5.5.7 and Section 5.5.8 above and also depicted in Figure D.2) acted as electrical connections for the coaxial heater and as injection point for coolant. The block is heavy compared to the coaxial heater and column, and also has to absorb the forces of the coolant inlet tube and the electrical connections. The column runs from the heated inlet/detector to a heated T-piece block, and in between it should not be exposed to any low-temperature cold spots, therefore the gap between the T-piece block and the inlet/detector should be quite  
3190 small, but the gap cannot be zero, because electrical isolation needs to be maintained. Also, fused-silica capillary columns are fragile and misalignment causes them to break. Therefore, mechanically stiff and accurate mounting was needed for the T-piece blocks, to allow the precise but adjustable alignment of the T-piece blocks with the inlet/detector.

3195 The final design for mounting the T-piece blocks was a pair of parallel rails. These rails were held in place in the Varian 3300 oven by friction, so that they could be adjusted and removed as necessary, yet were stiff enough to transfer the necessary forces without deflecting. The pointed ends of the rails pressed against a solid aluminium plate used as the floor of the oven, and at the top adjustable points pressed

3200 against pressure plates which pressed against the roof of the oven. Figure D.4 shows a technical drawing of the rails as designed.

The T-piece blocks were the electrical connections for the resistive coaxial heater, which meant they needed to be electrically isolated, but they were also heated, which meant that the insulation had to be resistant to heat. A commercially available material that met these requirements was found in the form of **silicon mica**, a composite material of mica and a silicone resin. This material has a continuous operating temperature of at least 500 °C, making it ideally suited to GC applications. The silicon mica is also machinable and can easily be shaped to the required design.

The T-piece blocks were mounted on a pair of cars riding on the round-bar rails. 3210 Each car was designed as a sandwich of plates of stainless steel and silicon mica around a pair of brass bushes. Once assembled, the cars offered a set of studs on to which the user could fit and bolt down the T-piece blocks. The positions of the cars were determined by a locking collar on one of the rails. Figure D.5 shows technical drawings of the cars that explains the design.

### 3215 5.5.10 Heating control

The amount of electrical power supplied to the coaxial heater was controlled by a bank of six PNP 2N2955 transistors, connected in parallel to distribute their heat dissipation. The final control signal was a voltage set either by a potentiometer from the front panel, or by the computer. An operational amplifier adjusted the current 3220 through the coaxial heater circuit so that a portion of the voltage applied to the coaxial heater was equal to the set-point voltage. By varying the set-point voltage the current through the coaxial heater can be controlled to provide any desired amount of heat ( $P = I^2R$ ).

### Temperature monitoring

3225 Independent of the amount of power dissipated in the coaxial heater, the current through the reference resistor was compared to the current through the column. This ratio corresponds to the resistance of the coaxial heater. This resistance is a function of the temperature of the heater. Through the calibration procedure described above the temperature of the coaxial heater can be calculated. The computer can do this 3230 fast enough so that a temperature measurement is available to continuously feed to a control system.

### PID tuning

A proportional-integral-derivative (PID) controller was used to calculate the amount of power necessary to keep the temperature of the coaxial heater as close to the 3235 temperature set point as possible. The temperature set-point, in turn, was given by the desired chromatographic temperature ramp.

The process of determining the best calculation by the PID is called **tuning**, and usually consists of determining the optimum values of a few parameters. Tuning PID controllers is a complex sub-discipline of process engineering, and outside the 3240 scope of this project, but for practical purposes we used a privately published step-by-step tuning method (Peacock 2008) which is a version of the Cohen-Coon tuning method.

Figure 5.13 shows the effect of an improved loop tuning.

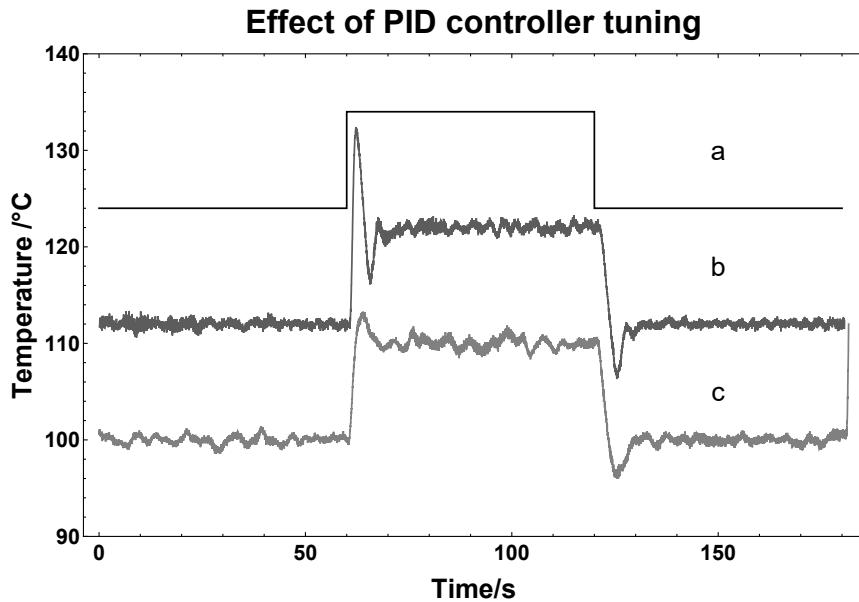


FIGURE 5.13: An illustration of effective PID controller tuning. The trace (a) represents the set point change over time and includes a step change. (b) Before tuning the temperature overshoots and then undershoots the step change in the set point. (c) After the controller has been tuned the overshoot in response to a step change in the set point is minimized.

`{fig:LoopTuning}`

A properly tuned heater helps to improve the repeatability of the chromatography through reliable temperature programs and prevents damage to the column due 3245 overheating during set point overshoots.

### Heating rates

As discussed in Section 5.4, for fast temperature-programmed gas chromatography the ramp rate needs to be in the order of thousands of degrees Celsius per minute.

3250 Figure 5.14 contains heating ramps executed by the coaxial heater. The heating rate was  $4000\text{ }^{\circ}\text{C min}^{-1}$ , and there are no significant differences between the 9 consecutive ramps.

### Cooling rates

The project did not demand a precise knowledge of, or control over, the cooling rate 3255 of the coaxial heater. The only requirement was that cooling should be as fast as possible. The cooling rate could be adjusted through the metering valve, and an optimum cooling rate of  $5100\text{ }^{\circ}\text{C} / \text{min}$  was achieved with a carbon dioxide flow rate of around 30 g/min. This allowed the column to be cooled down from  $350\text{ }^{\circ}\text{C}$  to  $-20\text{ }^{\circ}\text{C}$  in about 2 seconds. At this rate the portion of the chromatographic cycle that 3260 is not used for separation is dominated by fraction collection, and further cooling rate increases will not significantly reduce sample throughput.

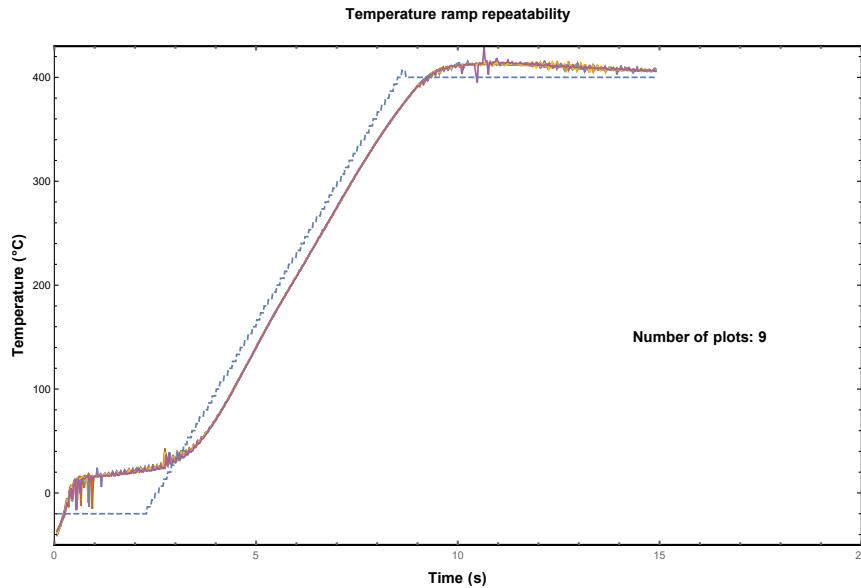


FIGURE 5.14: A graph of 9 identical, consecutive temperature ramps overlaid. The heating rate is  $4000\text{ C}^{\circ}/\text{min}$ . The temperature follows the set point faithfully up to  $350\text{ }^{\circ}\text{C}$ .

(fig:4000K/min)

## 5.6 Detector

The detector used in this fast GC was an unmodified Varian<sup>TM</sup> 3300 flame ionization detector (FID). The detector bias voltage was supplied by the original electronics, but a stand-alone high-speed electrometer (V.G. Micromass Ltd, Model M406-H) captured the signal, which was then conditioned by a bench-top amplifier (V.G. Micromass Ltd, Model M406) before it was sent to the computer. This electrometer and amplifier were fast enough to detect and amplify the signals generated by the fast GC.

## 5.7 Data acquisition and control software

The whole SFC $\times$ GC instrument was controlled from a single PC, running a single program. The program was written in LabVIEW 7.1<sup>TM</sup> (National Instruments). This software was designed to interact very easily with the National Instruments PCI-6014 multifunction data acquisition board.

LabVIEW is a **visual programming language**, so called because programs are created by manipulating icons and wires on a screen, instead of typing text. This visual aspect of it makes it very easy to develop user interfaces as **virtual instruments** and get quick results. Figure 5.15 shows the interface of the program used to control the SFC $\times$ GC instrument.

## 5.8 Data structure

In GC $\times$ GC, 2D data is recorded as a continuous FID output stream — as if it is a 1D GC chromatogram — and later converted into a 2D chromatogram using knowledge of the modulation period. For two reasons we could not use this approach. First, in

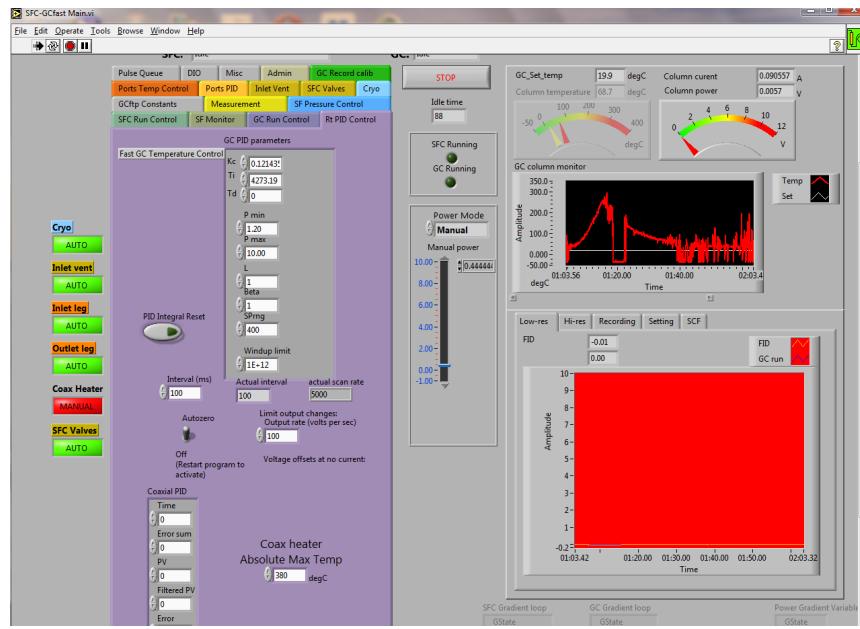


FIGURE 5.15: A screenshot of the LabVIEW virtual instrument used to control the SFC $\times$ GC instrument.

`(fig:SFCGCFastVI)`

our instrument the first (SFC) dimension runs in a stop-flow mode making continuous data recording inappropriate. Second, the duration of the cooling cycle varies, which would introduce unacceptable variation in  $^2$ D retention times.

We therefore constructed 2D chromatograms by recording a GC run for each SFC fraction injected. The  $^1$ D retention times were recorded as the start times of each individual GC run. The  $^2$ D retention times were measured from the time the GC fast temperature program started.

## 5.9 Data visualization

For data visualization we used the technical computing system Wolfram Mathematica 11.3<sup>TM</sup>. Mathematica is an extremely powerful and broad system, and we found it useful for its data manipulation and visualization tools.

The collected data could be handled in different ways. One way was to split it up into separate GC runs. Each individual GC run and its associated data could then be examined using the `Manipulate[]` function (Figure 5.16).

Then the data could be re-arranged into a list of three-element lists, with  $^1$ D retention time,  $^2$ D retention time, and detector signal as the elements of the inner lists. The Mathematica functions `List3DPlot[]` and `ContourPlot[]` could then be used to plot 3D chromatograms (Figure 5.17) or contour plots (Figure 5.18) respectively.

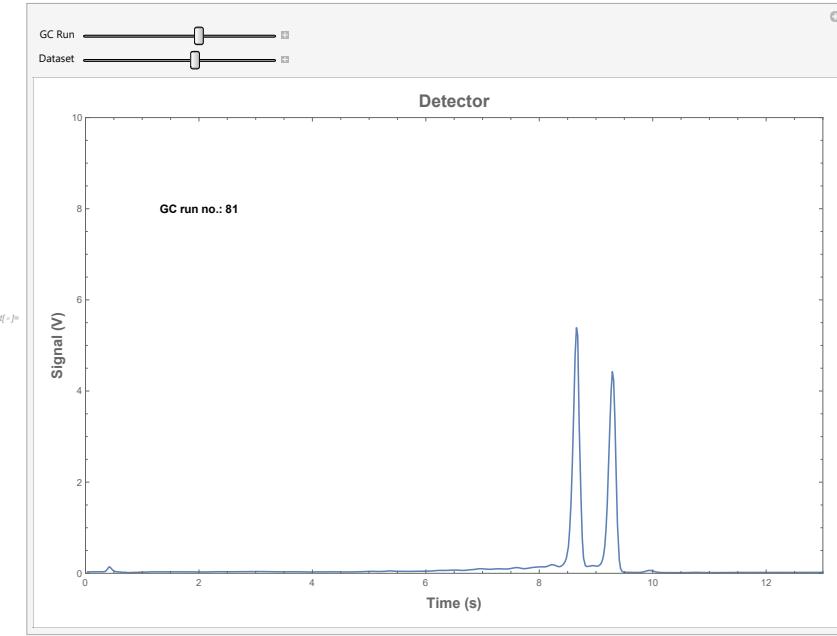


FIGURE 5.16: A single fast GC chromatogram, in a Mathematica `Manipulate[]` environment. The sliders can be used to select the which GC run to view, and which data of that run.

`{fig:SingleGC}`

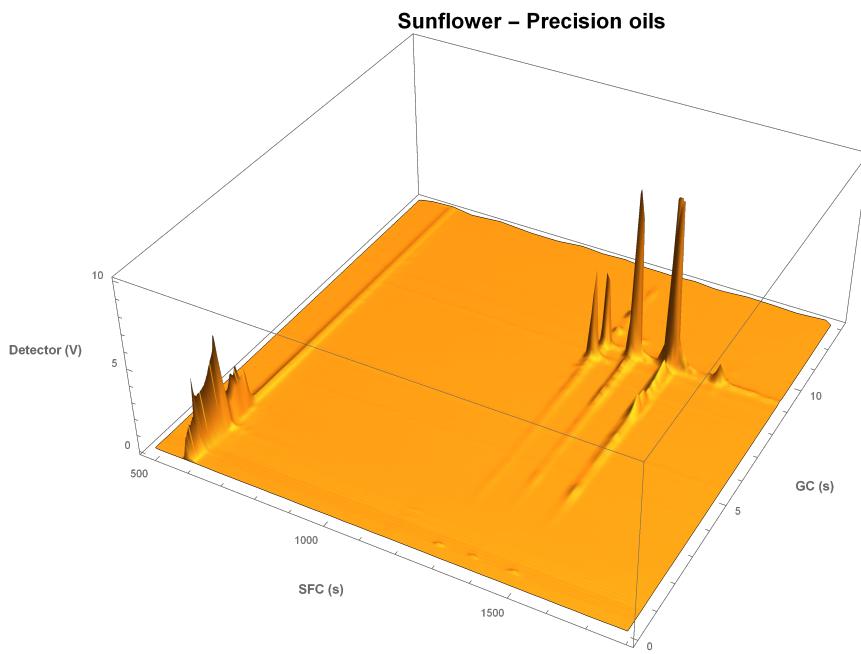


FIGURE 5.17: A 2D SFC $\times$ GC chromatogram.

`{fig:2DChromatogram}`

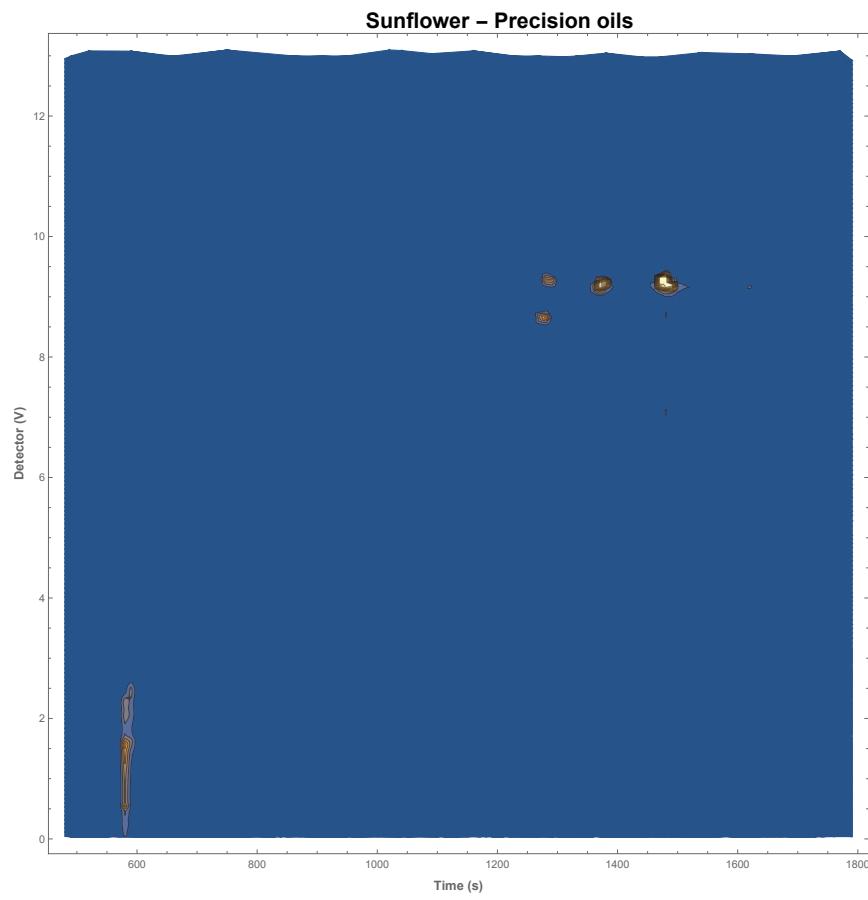


FIGURE 5.18: An SFC $\times$ GC contour plot.

`{fig:Contourplot}`

I'm really into coconut oil for everything. I cook it, eat it, put it in my hair, and use it as body lotion. I put it on my face, too - day cream, night cream, whatever. I love the smell. It reminds me of the beach. I'm not particular on what brand as long as it's organic.

Juliette Lewis

# 6

## Investigating biodiesel feedstock using SFC $\times$ GC

<Chapter 6>

### 3305 6.1 Introduction

Biodiesel consists mostly of a mixture of methyl esters of fatty acids obtained from plant oils. To comply with the relevant technical standard, SANS 1935, (SABS 2008) (see Chapter 3), it must consist of a mass fraction of 96.5 % or more of fatty acid methyl esters (FAME), but not more than 12 % linolenic acid methyl ester, and not

3310 more than 1 % of FAMEs with more than 4 double bonds. The prescribed methods for determining the quantities of these compounds are chromatographic, but the method for determining the *total amount* of FAMEs is a completely different method from that for determining the amount of *unsaturated* FAMEs. Both these methods generate complex chromatograms that need highly skilled and experienced chromatographers to interpret. There is no doubt that the use of artificial intelligence and other technological innovations for interpreting chromatograms will grow, but the paradox of automation<sup>1</sup> predicts that as biodiesel production grows, feedstocks proliferate, and complexity increase, the analytical chemist will need better tools, methods, and instruments to understand the problems that arise when automation fails.

3315 Comprehensively coupled chromatography offers a way to better exploit chemistry for improved analytical separations, adding another tool to the chemical analyst's toolbox. It does this in three ways: the first is by increasing the peak capacity of the system, the second is by improved sensitivity, and the third is by generating patterns in the data.

### 3325 6.2 SFC of FAMEs

The power of comprehensive chromatography is unlocked by orthogonality. Orthogonality is the difference in separation mechanism between the two dimensions (Marriott, Schoenmakers, and Wu 2012). When FAMES are chromatographically separated by SFC in using neat carbon dioxide as a mobile phase and unmodified silica as a stationary phase, then the separation is according to the number of double

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3330 <sup>1</sup>Automation helps you least when you need it most (Strauch 2018; Bainbridge 1983).

bonds, independent of chain length (Robertson et al. 1991; R. M. Smith and Cocks 1994; R. M. Smith, Hyytiänen, et al. 2001). This stands in strong contrast to the separation of FAMEs by GC, where the major separation is according to volatility, which can be adjusted, but not overridden, by changing the polarity (or other chemical aspect) of the stationary phase. (See Figure 6.1.)

Silver ions are often used in stationary phases to separate unsaturated compounds, and this includes stationary phases for SFC (Sandra et al. 2002; Potgieter et al. 2013). The retention mechanism is quite complex (Nikolova-Damyanova 2019), but it offers a powerful technique for the elucidation of lipid structures, as reviewed as early as 1966 (L. J. Morris 1966). Nevertheless, in this chapter the use of stationary phases modified with silver ions was neither necessary nor attempted.

The utterly different retention behaviours of FAMEs on silica with a carbon dioxide mobile phase and on GC offers high orthogonality, which promises to make comprehensive coupling worthwhile.

### 6.3 Performance of the coaxial heater.

Practical SFC $\times$ GC depends on reliably repeating fast temperature programs on a capillary gas chromatography column. The effect of these conditions on the column will be discussed in the following sections.

#### 6.3.1 Column lifetime

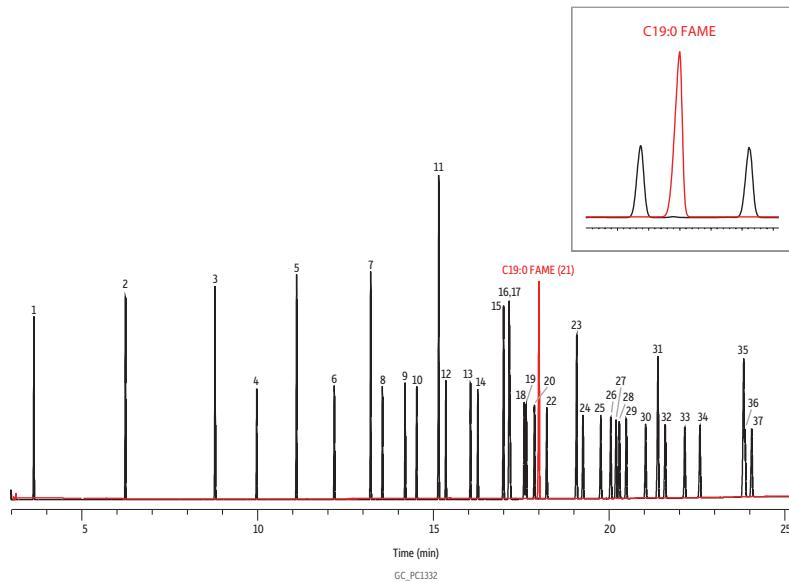
As discussed in Section 3.9, FAMEs can be separated by gas chromatography using relatively non-polar columns with high-temperature programs. For this purpose column manufacturers supply special 'high temperature' columns, which address two aspects of column lifetime: mechanical degradation and stationary phase degradation. The SFC $\times$ GC chromatograph did not use such columns, and the following discussion explains why their use could be avoided.

##### Mechanical degradation

Fused silica has a high **tensile strength**, but also a low **fracture toughness**. This means that it is strong enough to resist the forces involved in its use as a column, but very liable to fracture if it gets damaged. Very small flaws, less than a micron in depth, will cause fused silica (or any other glass) to fracture under loads far under what would be expected from its tensile strength. Such flaws can be caused by mechanical scratches or reaction with atmospheric water. To prevent this the outside of the fused silica capillary is coated with a protective layer of **polyimide resin**. The strength of the a capillary column therefore depends on the integrity of the polyimide coating. Like all resins and polymers the polyimide coating degrade faster at higher temperatures, so the longer the column spends at high temperature the sooner the polyimide will degrade to the point where it fails to protect the fused silica, and the shorter the overall mechanical lifetime of the column will be.

Traditionally, chromatographers are advised to use temperature programs that minimize the time spent at high temperature, which will contribute to longer column lifetimes. The usual way to do it is to use a temperature program with a maximum temperature no higher than necessary, but the same cumulative time at high temperature can be obtained by using higher temperatures but for shorter periods. Fast temperature programs expose the columns to high temperatures for only very brief periods.

## Food Industry FAME on FAMEWAX by EN14103 (2011)



Peaks	tx (min)	Conc. (mg/mL)	Structural Nomenclature
1. Methyl capronate	3.629	1.2	C6:0
2. Methyl caprylate	6.237	1.2	C8:0
3. Methyl caprate	8.787	1.2	C10:0
4. Methyl undecanoate	9.971	0.6	C11:0
5. Methyl laurate	11.105	1.2	C12:0
6. Methyl tridecanoate	12.179	0.6	C13:0
7. Methyl myristate	13.215	1.2	C14:0
8. Methyl myristoleate	13.549	0.6	C14:1 (cis-9)
9. Methyl pentadecanoate	14.196	0.6	C15:0
10. Methyl pentadecenoate	14.524	0.6	C15:1 (cis-10)
11. Methyl palmitate	15.152	1.8	C16:0
12. Methyl palmitoleate	15.355	0.6	C16:1 (cis-9)
13. Methyl margarate	16.052	0.6	C17:0
14. Methyl heptadecenoate	16.261	0.6	C17:1 (cis-10)
15. Methyl stearate	16.995	1.2	C18:0
16. Methyl oleate	17.156	1.2	C18:1 (cis-9)
17. Methyl elaidate	17.168	1.2	C18:1 (trans-9)
18. Methyl linoleate	17.583	0.6	C18:2 (all-cis-9,12)
19. Methyl linolelinate	17.641	0.6	C18:2 (all-trans-9,12)
20. Methyl γ-linolenate	17.874	0.6	C18:3 (all-cis-6,9,12)
21. Methyl nonadecanoate	18.052	2.0	C19:0
22. Methyl α-linolenate	18.223	0.6	C18:3 (all-cis-9,12,15)
23. Methyl arachidonate	19.075	1.2	C20:0
24. Methyl (2)-11-eicosenoate	19.255	0.6	C20:1 (cis-11)
25. Methyl 11,14,17-eicosatrienoate	19.763	0.6	C20:2 (all-cis-11,14,17)
26. Methyl eicos-8,11,14-trienoate	20.046	0.6	C20:3 (all-cis-8,11,14)
27. Methyl hexadecanoate	20.197	0.6	C16:0
28. Methyl arachidonate	20.290	0.6	C20:4 (all-cis-5,8,11,14)
29. Methyl 11,14,17-eicosatrienoate	20.488	0.6	C20:3 (all-cis-11,14,17)
30. Methyl 5,8,11,14,17-eicosapentaenoate	21.036	0.6	C20:5 (all-cis-5,8,11,14,17)
31. Methyl behenate	21.39	1.2	C22:0
32. Methyl erucate	21.595	0.6	C22:1 (cis-13)
33. Methyl decosadenoate	22.150	0.6	C22:2 (all-cis-13,16)
34. Methyl tricosanoate	22.584	0.6	C23:0
35. Methyl lignocerate	23.826	1.2	C24:0
36. Methyl docosahexaenoate	23.863	0.6	C22:6 (all-cis-4,7,10,13,16,19)
37. Methyl nervonate	24.055	0.6	C24:1 (cis-15)

**Column Sample** FAMEWAX, 30 m, 0.25 mm ID, 0.25  $\mu$ m (cat.# 12497)  
Food industry FAME mix (cat.# 35077)  
Methyl nonadecanoate (cat.# 35055)  
Standard cat.# 35055 was dissolved in toluene.

**Diluent:** 1  $\mu$ L split (split ratio 100:1)

**Injection** Inj. Vol.: 24.0  $\mu$ L

**Liner:** Topaz 4.0 mm ID Precision inlet liner w/wool (cat.# 23305)

**Oven** Inj. Temp.: 240 °C

**Oven Temp.:** 60 °C (hold 2 min) to 200 °C at 10 °C/min to 240 °C at 5 °C/min (hold 7 min)

**Carrier Gas** He, constant flow

**Flow Rate:** 1.7 mL/min

**Detector** FID @ 250 °C

**Instrument** Agilent 7890B GC

**Notes** This chromatogram is an overlay of two injections: food industry FAME standard (black) and C19:0 methyl ester in toluene (red). An excellent separation of C19:0 (used in EN14103 as an internal standard) and the most prevalent FAMEs found in biodiesel blends was achieved. Note that C4:0 from the food industry FAME standard elutes in the solvent front.



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FIGURE 6.1: This figure shows that when separating FAMEs by GC, using specialized column, separation is primarily by fatty acid chain length, and secondarily by number of double bonds. Reproduced with permission of Restek Corporation.

3380 (Column mechanical lifetime can also be improved by using metal capillaries. Metals are not afflicted by brittle fracture the way fused silica is, but they don't offer the chemical inertness of fused silica. To exploit the fracture toughness of metal, technologies have been developed to deactivate metal surfaces (E. R. Rohwer, Pre-torius, and Hulse 1986; D. A. Smith 2002), which makes metal columns a modern possibility. Column manufacturers offer metal columns specifically designed for biodiesel analysis.)

3385 Another durability benefit of the stainless steel coaxial heater is that the outer surface of the capillary column is protected from mechanical damage by its encasement in the stainless steel tube of the coaxial heater. During development of the coaxial heater there were occasions where the column was accidentally overheated by poorly-controlled resistive heating. Even though the polyimide coating had been completely charred away and the capillary was left fragile, the column was intact and still provided separations.

3390 Using carbon dioxide as a between-program coolant also protects the polyimide coating from oxidative damage by purging the column environment of oxygen.

### Stationary phase degradation

3395 The stationary phases inside the GC columns face degradation similar to that of the polyimide. Column manufacturers minimize stationary phase degradation by using improved technologies like cross-linked resins, or polymers with backbones resistant to certain degradation mechanisms (Day et al. 2003). Using fast temperature programs help extend stationary phase lifetime by minimizing the cumulative time the column spends at high temperature.

### Column bleed

3400 Short columns and fast temperature programs have another beneficial side effect: reduced **column bleed**. Any resin- or polymer-based stationary phase degrades over time, and this degradation is faster at higher temperatures (McNair, Miller, and Snow 2019, p. 66). It is observed as a rise in baseline during the high temperature part of a temperature program. Various stationary phase technologies can be employed to reduce column bleed, but because the amount of column bleed depends on the amount of stationary phase it is clear that a 1 m column will have 1/30th of the column bleed of a 30 m column with the same stationary phase. Despite using temperature programs that went up to the temperature limit recommended by the column manufacturer, column bleed was not a significant feature in the obtained chromatograms.

#### 6.3.2 Thermal shock

3415 An aspect of fast temperature programming with such a temperature range as used in this study ( $-30^{\circ}\text{C}$  to  $400^{\circ}\text{C}$ ) is **thermal shock**. Thermal shock occurs when an object is subjected to a rapid change in temperature. Under such non-equilibrium conditions different parts of the object will have different temperatures. The material that the object is made of has a certain **coefficient of thermal expansion**, which means that the relative sizes of the parts of the object at different temperatures will change at different rates. This will cause stress between the parts, and if the stress exceeds the strength of the material, the material will fracture. Fused silica has a low

TABLE 6.1: A summary of retention time repeatability of alkanes separated on the fast temperature programmed chromatograph.

Compound	n	t <sub>r</sub> (s)	S.D. of t <sub>r</sub> (s)	R.S.D. of t <sub>r</sub> (%)
Dodecane	73	5.07	0.023	0.46
Hexadecane	73	6.58	0.052	0.78

<sup>3420</sup> coefficient of expansion<sup>2</sup>, and therefore experiences low thermal shock. The amount of material in a capillary is also relatively small, which keeps temperature gradients small and therefore stresses low. We didn't observe any column failures that could be ascribed to thermal shock.

### 6.3.3 Thermal fatigue

<sup>3425</sup> Another cause of material failure due to temperature is **thermal fatigue**. "Thermal fatigue (TF) is the gradual deterioration and eventual cracking of a material by alternate heating and cooling during which free thermal expansion is partially or completely constrained" (Rao and Raj 2001). The expansion of the portion of the fused silica column subjected to alternate heating and cooling is not constrained, <sup>3430</sup> and the part of the fused silica column that is constrained (in the sealing ferrules) is not subjected to alternate heating and cooling (it is held at constant temperature in the heated T-piece block), therefore we do not expect thermal fatigue. We did not explicitly test for the occurrence of thermal fatigue, but we did not encounter column failure that could be attributed to thermal fatigue.

### <sup>3435</sup> 6.3.4 Corrosion

One failure of the coaxial heater could be attributed to corrosion. The joint between the coaxial heater and the heated T-piece block is brazed, and there was a failure of the thin-walled stainless steel tube near that joint, in the portion heated during the brazing operation. Visual inspection seemed to indicate thinning caused by corrosion. <sup>3440</sup> An acid flux was used during the brazing, which might have contributed to the corrosion, especially in the presence of water condensed from the atmosphere during the cooling of the column.

### 6.3.5 GC retention time precision

During the development of the resistive heater it was shown that the measured temperature ramps are highly repeatable. But the retention times of compounds being separated by GC are highly sensitive to temperature, so it was necessary to obtain a measure of the retention time repeatability.

To estimate the retention time variance, two alkanes (dodecane and hexadecane) were added to the SFC mobile phase at low concentration. These compounds are not retained on the silica at all, and are therefore present in all the GC runs at the same concentration and all the peaks should be identical. Variance in peak retention, -area, and -height therefore indicate the repeatability of the fast GC system. Variability of retention time is summarized in Table 6.1.

<sup>2</sup>( $0.55 \times 10^{-6} \text{ K}^{-1}$ , compared to the  $9.0 \times 10^{-6} \text{ K}^{-1}$  of soda lime glass or the  $4.0 \times 10^{-6} \text{ K}^{-1}$  of borosilicate glass

TABLE 6.2: Oils used for FAME analysis

&lt;tab:OilSamples&gt;

Oil	Brand	Species
Canola	Spar	<i>Brassica napus</i>
Sunflower oil	Pick n Pay	<i>Helianthus annuus</i>
Coconut oil	Lemcke	<i>Cocos nucifera</i>
Flax seed oil	Lemcke	<i>Linum usitatissimum</i>

The peak widths were about 500 ms, so the 20 ms standard deviation for hexadecane means that the variation in retention time is only about 10 % of the peak width.  
 3455 The relative standard deviations (RSD) of retention times were similar to those obtained in GC $\times$ GC (Shellie, Xie, and Marriott 2002), and much smaller than those of our original “proof of concept” instrument (Andre Venter 2003; Andre Venter and Egmont R. Rohwer 2004)

## 3460 6.4 Study of potential biodiesel feedstock by SFC $\times$ GC

Biodiesel can be produced from any combination of a variety of vegetable oils<sup>3</sup>. Indeed, to ensure conformance to standards it might be necessary to produce biodiesel from blends or combinations of oils. But as an introduction it is instructive to start by analysing the FAMEs obtained from neat oils.

### 3465 6.4.1 Samples

Various samples of edible vegetable oil were obtained from supermarkets. (See Table 6.2)

### 6.4.2 Sample preparation

Fatty acids in oils might be either free or bound to glycerol. To quantitatively convert them to FAME therefore requires that the bound fatty acid be transesterified, and the free fatty acids esterified. Both transesterification and esterification can be achieved by acid catalyst, but this reaction is quite slow. Basic catalysts can rapidly transesterify acyl glycerols, but will not esterify free fatty acids.  
 3470

The method used involved first treating the oil sample with sodium hydroxide dissolved in dry methanol. The methanol acts as a solvent but also provides an excess of methanol so that the transesterification reaction is driven to completion. On completion of the reaction an excess of acid is added, which neutralizes the base and esterifies any free fatty acids. Then an organic solvent and water are added, which provides two phases: the non-polar FAMEs dissolve in the organic layer, and the polar glycerol and salts dissolve in the water layer.  
 3475  
 3480

### Method

This method is based on an official method (AOCS 2017), modified in two respects. First, the acidic catalyst boron trifluoride is replaced by sulphuric acid, and second, hexane is used instead of heptane.

<sup>3</sup>At this time SANS 1935 excludes animal fats, but they might be included in future. All of the following discussion also applies to animal fats.

- 3485 1. Transfer 4 drops of melted sample to a 20 ml glass stoppered test tube. Add a few boiling stones
2. Add 2 ml NaOH/methanol solution (0.5N) and boil for 11 min under reflux.
3. Add 2 ml 10% H<sub>2</sub>SO<sub>4</sub>/methanol solution via the condenser and boil for 2 min
4. Add 2 ml hexane via the condenser and boil for 1 min.
- 3490 5. Remove the test tube from the heat source and leave to cool.
6. Add 4 ml saturated NaCl solution and mix gently.
7. Allow the phases to separate. Transfer the hexane layer to a vial and use for injection.

### 6.4.3 SFC

3495 A 0.5  $\mu$ l volume of the hexane layer was injected into the neat carbon dioxide mobile phase, as described in Section 4.4. The column was five HPLC columns (150 mm  $\times$  4.6 mm, 3  $\mu$ m particles) (Restek, Pinnacle DB Silica) connected in series.

### 6.4.4 Modulation

3500 The modulation period was 10 s, during which the fractions of the SFC eluate were transferred to the GC inlet via a linear restrictor. In the hot (350 °C) inlet the fractions evaporated instantly and were swept onto the cold (-20 °C) column where they were trapped in the stationary phase. 3 s after the collection period ended the vent valve opened for 1 s to vent excess pressure from the GC inlet.

### 6.4.5 GC

3505 The fast temperature program ramped the GC column temperature from -20 °C to 350 °C in 10 s (2200 °C s<sup>-1</sup>), maintaining 350 °C for 2 s. Then the cooling system activated and cooled the column to -20 °C or below, ready to trap the next SFC fraction. The FID detector was kept at 350 °C.

3510 In this way a series of GC chromatograms of SFC fractions were recorded, which were assembled into a 2D chromatogram. Figure 6.2 shows a 2D chromatogram of a sample of FAME prepared from canola oil. This 2D chromatogram consists of 132 fast GC chromatograms collected in approximately 90 minutes.

### 6.4.6 Results and discussion

3515 The chromatogram shown in Figure 6.2 is very simple to interpret. In the SFC dimension, separation is by number of double bonds. This means that FAMEs without double bonds elute first in this dimension, followed by FAMEs with one, two and three double bonds respectively. There is good resolution between the peaks in the SFC dimension.

3520 The volatility of a FAME depends primarily on the length of the carbon atom chain, therefore the separation in the GC dimension is by chain length. Because these FAMEs are of biological origin, we expect their chains to have an even number of carbon atoms. From the literature we know that oleic acid (C18:1) is the most common fatty acid in canola oil (Joint FAO/WHO Codex Alimentarius Commission

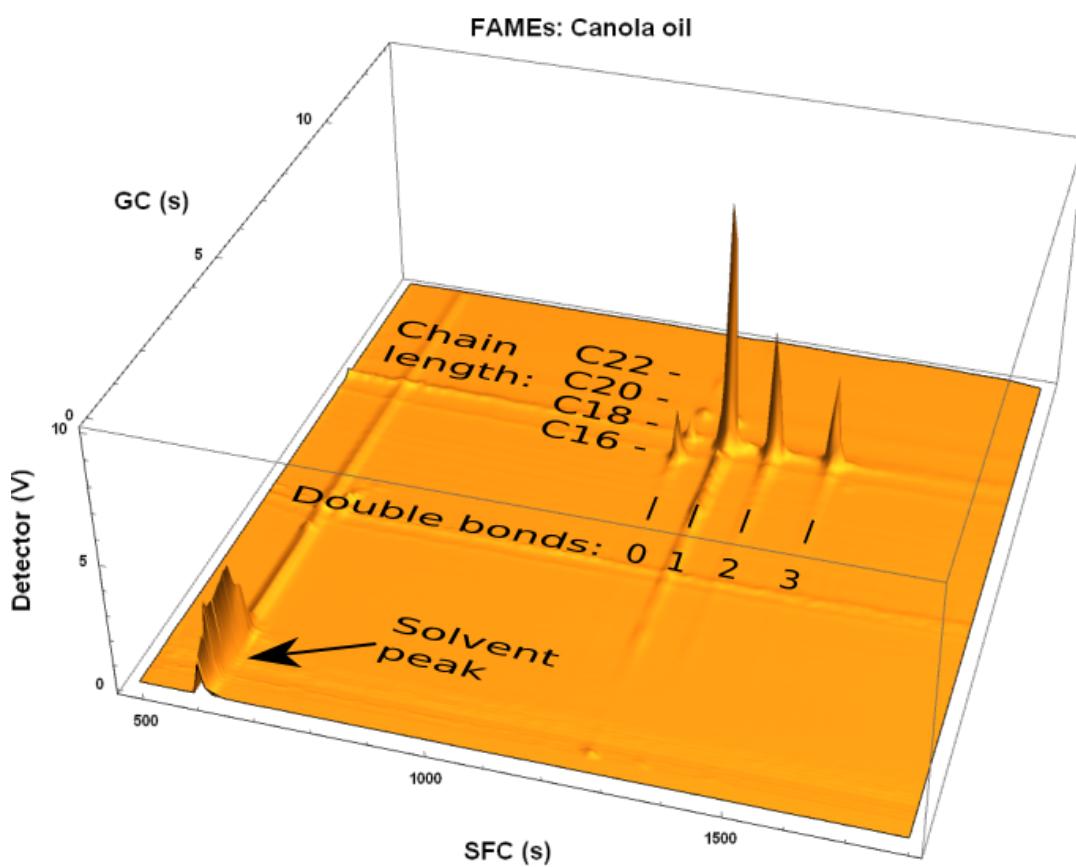


FIGURE 6.2: A 2D chromatogram of FAMEs derived from canola oil.  
It is clear that the oil consists mostly of unsaturated fatty acids.

{fig:2DCanola}

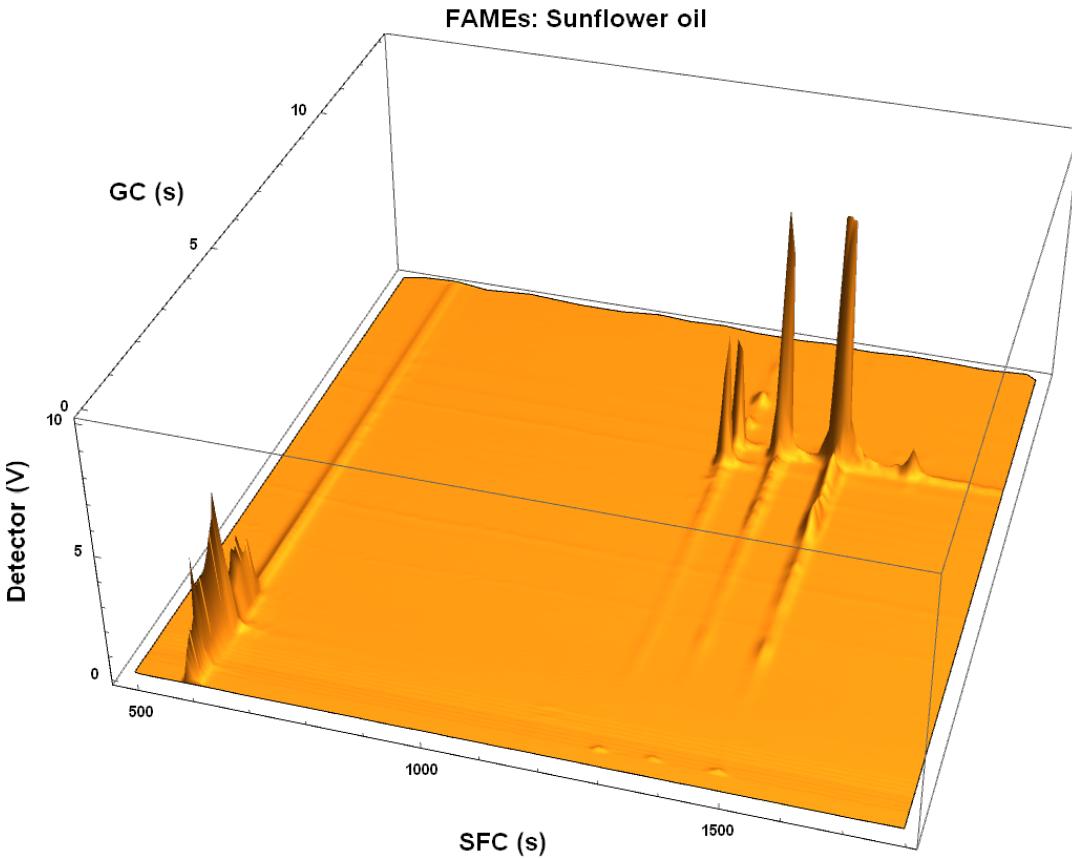


FIGURE 6.3: A 2D chromatogram of FAMEs derived from sunflower oil. It is clear that the oil consists mostly of unsaturated fatty acids, with a high proportion of linoleic acid (C18:2)

(fig:2DSunflower)

2019), and that therefore the major peak in the chromatogram is a C18 FAME. This allows us to easily identify the peaks for C16, C18, C20 and C22 in the GC dimension.

The information provided by this chromatogram confirm that canola oil is a viable biodiesel feedstock: The unsaturated compounds have mostly a single double bond, which should make it oxidatively stable, and the quantities of unsaturated FAMEs are low, which means that it will most likely have suitable cold-flow properties, because (cis) unsaturated FAMEs have higher freezing points than saturated FAMEs.

#### 6.4.7 Sunflower oil

Figure 6.3 shows an SFC $\times$ GC chromatogram of FAMEs obtained from sunflower oil. The biggest peak in the chromatogram corresponds to linoleic acid methyl ester (C18:2). This agrees with the literature, which indicates that the main, traditional cultivar of sunflower produced in South Africa is of the high-linoleic type, with a fatty acid profile of 69 % linoleic acid, 20 % oleic acid, and 11 % saturated fats (Joint FAO/WHO Codex Alimentarius Commission 2019).

This sunflower oil chromatogram was compared to a fatty acid profile obtained by an accredited oil analysis laboratory<sup>4</sup>. (See Table E.1.) The fatty acid profile data

<sup>4</sup>Precision Oil Laboratories, info@precisionoils.co.za, +27 15 307 7208, SANAS Testing Laboratory T0802

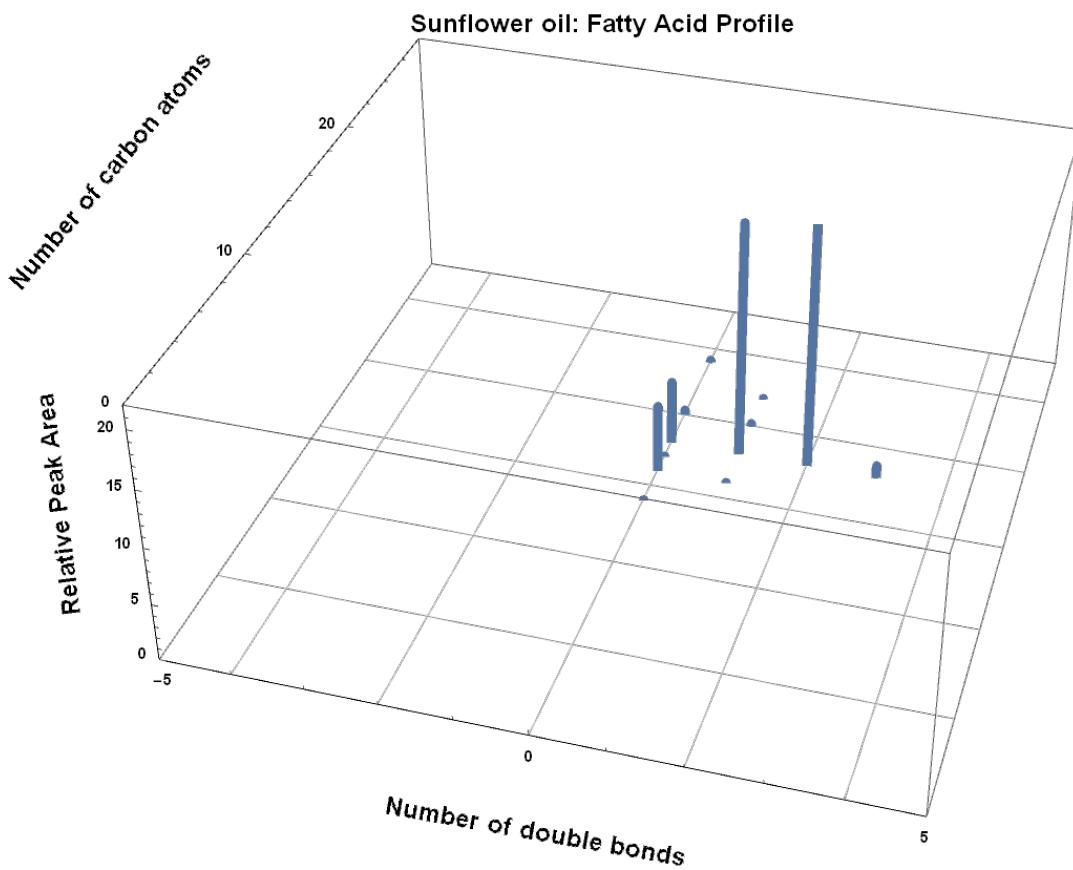


FIGURE 6.4: The fatty acid profile of a sunflower oil, when plotted on a suitably-scaled bar chart, shows how well the SFC $\times$ GC chromatogram of the same oil (Figure 6.3) represents the same data without any processing.

ig:2DSunflowerBarChart)

was plotted on a 3D bar chart (See Figure 6.4). The axes of the bar chart were chosen to plot the fatty acid profile information in a similar space and scale as the chromatogram (Figure 6.3). It can be seen that the bar chart shows the same information as the chromatogram. The information in the bar chart has been obtained from a 3545 1D chromatogram by integrating peaks, calculating amounts of FAMEs from peak areas, and then plotting those amounts in the bar chart, whereas the SFC $\times$ GC chromatogram is a plot of FID response and involves no quantification.

This chromatogram suggests that sunflower oil is a suitable feedstock for producing biodiesel: the relatively high amount of unsaturated FAs would seem to imply suitable cold flow properties. The relatively high degree of unsaturation does 3550 suggest caution regarding oxidative stability.

#### 6.4.8 Coconut oil

Figure 6.5 shows the SFC $\times$ GC chromatogram of FAMES from coconut oil (*Cocos nucifera L.*). It is clear that the dominant FAMES in coconut oil are saturated.

The high saturation of coconut oil makes it oxidatively stable, but the high melting point of the saturated FAMES also means that it might not meet cold flow requirements. 3555

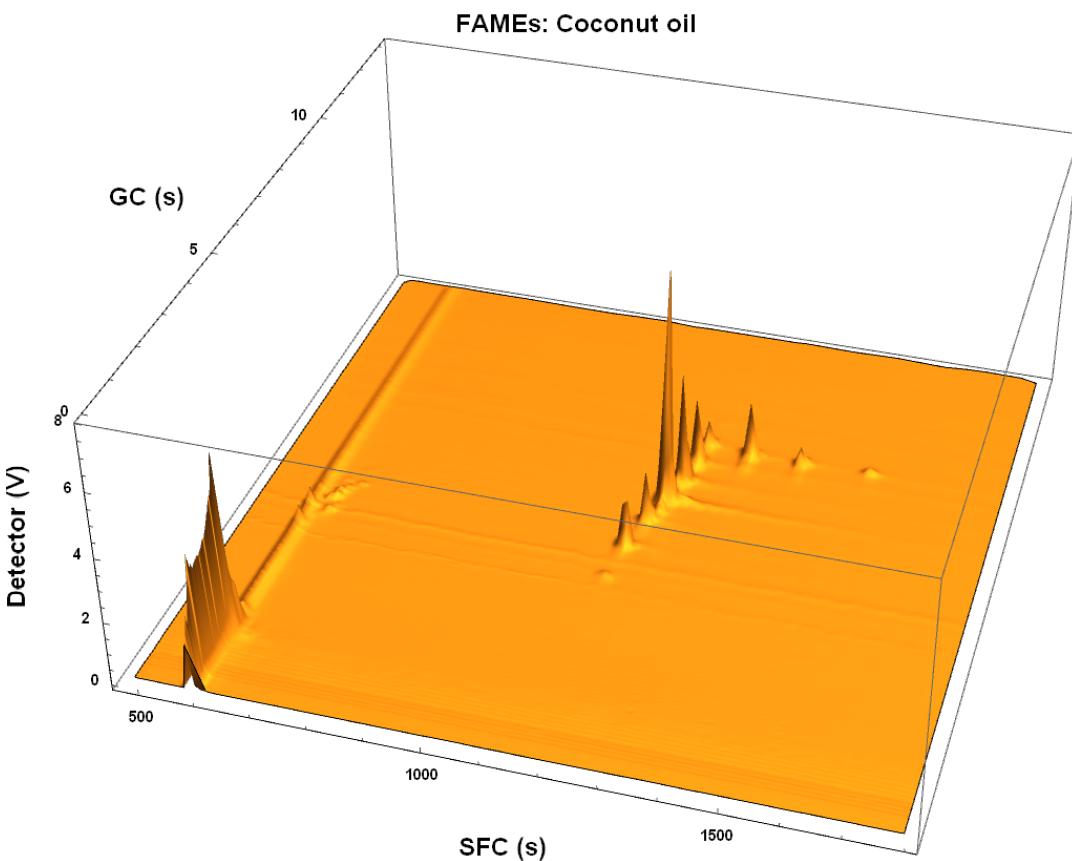


FIGURE 6.5: A 2D chromatogram of FAMEs derived from coconut oil.  
It is clear that the oil consists mostly of saturated fatty acids.

(fig:2DCoconut)

The high oxidative stability of coconut oil makes it resistant to going rancid, which means it has been used as a natural ingredient for cosmetics since antiquity (Berdick 1972).

3560

#### 6.4.9 Flax seed oil

Figure 6.6 shows an SFC $\times$ GC chromatogram of FAMEs obtained from flax seed oil. This oil is obtained from the seeds of the annual plant *Linum usitatissimum*, and was sold as an edible oil.

3565

The chromatogram shows that the major fatty acid component of flax seed oil is linolenic acid (C18:3). This will immediately inform a producer that flax seed oil would not be a suitable feedstock for biodiesel: SANS 1935 limits the linolenic acid fraction to less than 12 % (See section 3.9.5).

3570

Flax seed oil only recently became considered a food or a food supplement: the 1911 edition of the Encyclopaedia Britannica only mentions it as a food by saying "The oil is to some extent used as food in Russia and in parts of Poland and Hungary" (*Linseed 1911*), and remarks that the oil is edible when it is cold pressed. This is not surprising, because the oil goes rancid very quickly, leaving it unpalatable. Historically the more common name for the oil has been **linseed oil** and it was mainly used for industrial purposes, particularly in paint and varnish formulations. As the encyclopedia puts it: "Commercial linseed oil has a peculiar, rather disagreeable sharp taste and smell." The oil used in artists' oil paints is most commonly refined

3575

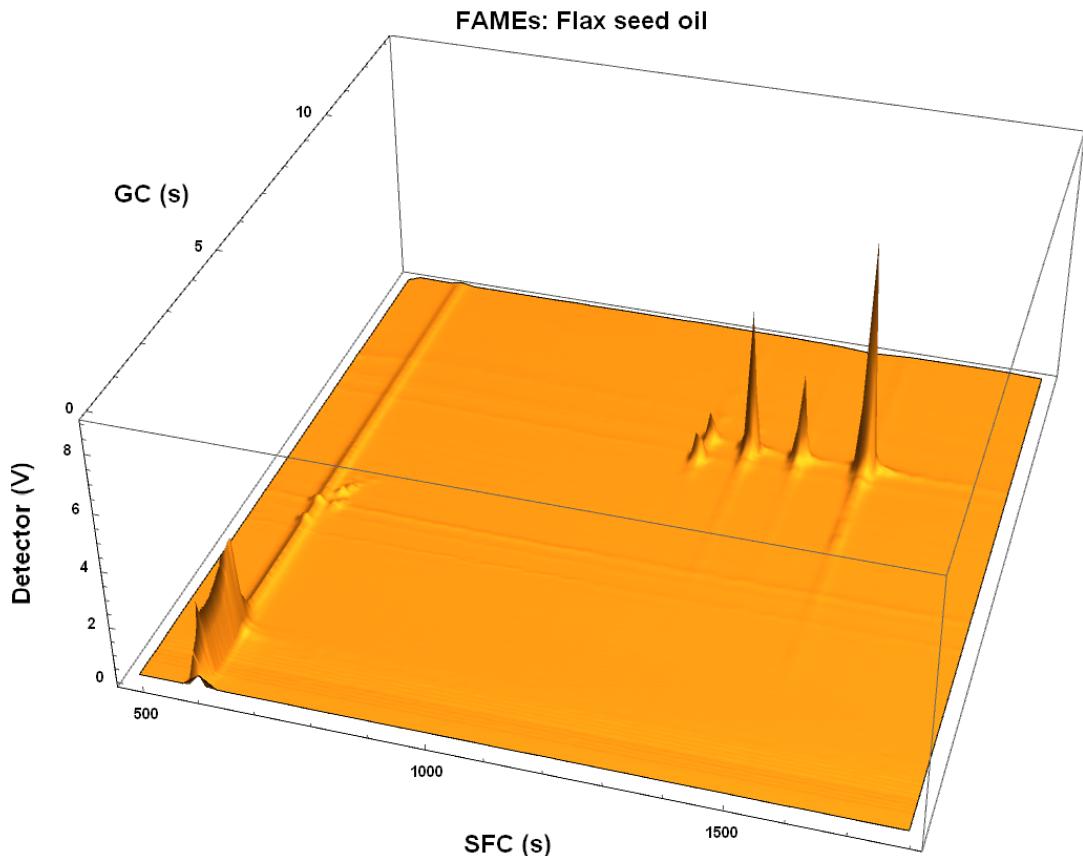


FIGURE 6.6: A 2D chromatogram of FAMEs derived from flax seed oil. It is clear that the oil consists mostly of highly unsaturated fatty acids.

`(fig:2DFlax)`

and prepared linseed oil. This oil is classed as a **drying oil**, because a layer of this oil exposed to air becomes solid, or "dry". This is the result of the high degree of unsaturation of the oil, which rapidly causes oxygen-mediated polymerization. On exposure to air the oil gradually turns more viscous and eventually solid, making it a suitable base for paint, particularly if the oil has been pre-treated. In biodiesel such polymerization could cause clogging or the formation of objectionable sludge.

## 6.5 The study of biodiesel blends by SFC $\times$ GC

### 3585 6.5.1 Introduction

Biodiesel can be blended in all proportions with diesel obtained from petroleum (**petrodiesel**). The volatility of biodiesel is slightly lower than that of petrodiesel, but in 1D gas chromatography the peaks of biodiesel compounds with higher volatility overlap significantly with the peaks of petrodiesel compounds of lower volatility. 3590 The high complexity of the petrodiesel therefore interferes with the determination of the biodiesel compounds. One example of an attempt to overcome this problem is the use of a highly polar ionic liquid column, which increases the retention of the FAMEs in the blend relative to the retention of the petrodiesel hydrocarbons (Ragonese et al. 2009).

3595 On a polar stationary phase like bare silica with a non-polar mobile phase like carbon dioxide, the hydrocarbons from petrodiesel are not expected to be significantly retained. But, as shown in the previous section, under such conditions FAMEs are retained and separated. When we inject a blend of biodiesel and petrodiesel on SFC $\times$ GC, we therefore expect the biodiesel portion to be completely 3600 separated from the petrodiesel portion.

### 6.5.2 Sample

A B50 biodiesel blend was prepared by mixing equal volumes of biodiesel and petrodiesel in the laboratory. The biodiesel sample was donated by a commercial testing laboratory and had been produced from waste cooking oil by an anonymous 3605 producer. The petrodiesel sample (Shell Extra Diesel 500 ppm) was obtained from a consumer filling station. The injection system described in Section 4.4 was used to inject a 0.5  $\mu$ l volume of this undiluted B50 blend.

### 6.5.3 SFC

The SFC used neat carbon dioxide at 200 bar inlet pressure as mobile phase. The column 3610 consisted of five HPLC bare silica columns (150 mm  $\times$  4.6 mm, 3  $\mu$ m particles) (Restek, Pinnacle DB Silica) connected in series.

### 6.5.4 Modulation

The SFC eluate fractions of 2 s were collected on the GC column cooled to a temperature of  $-20^{\circ}\text{C}$ . The inlet vent was held closed for 4 s, and then opened for 1 s to 3615 release excess pressure.

### 6.5.5 GC

The column used in the fast gas chromatograph was an OV-5 column, which has a stationary phase comprised of 5 % diphenyl, 95 % dimethylpolysiloxane. It was 1 m

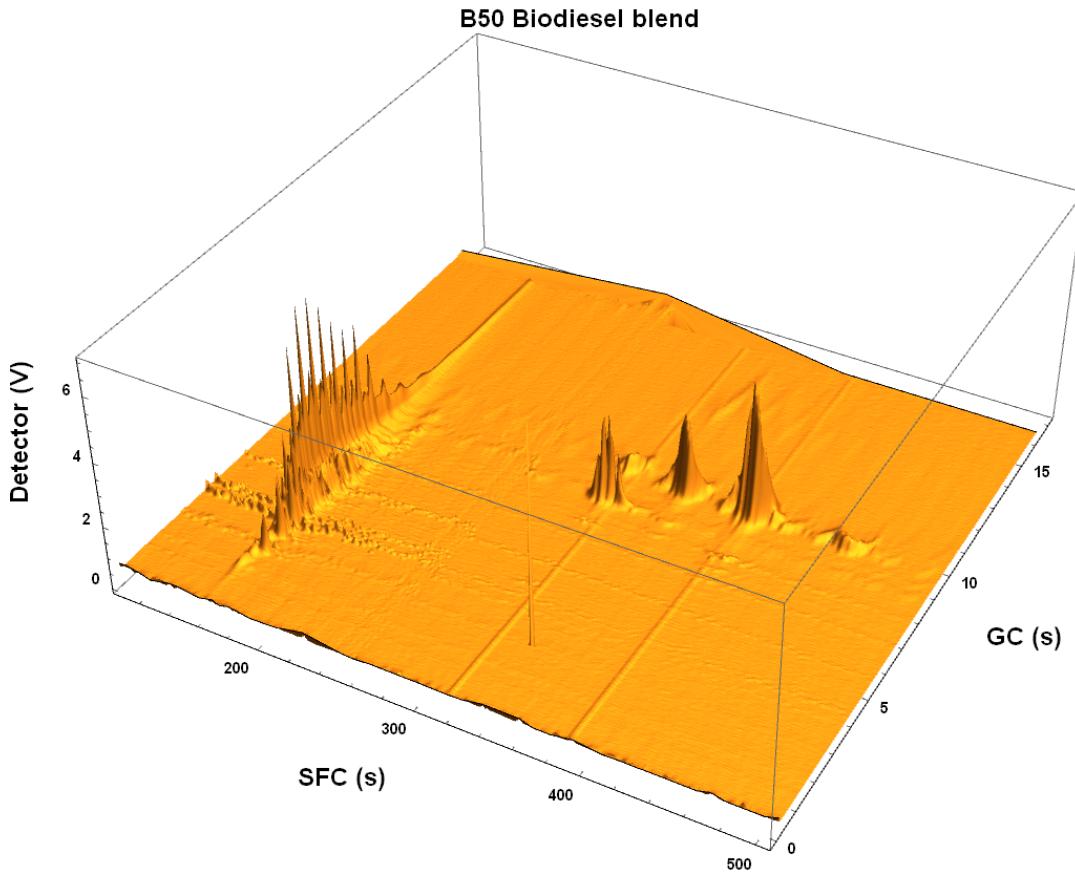


FIGURE 6.7: The separation of the biodiesel fraction and the petroleum fraction of a biodiesel blend. SFC: Neat CO<sub>2</sub> at 200 bar on Restek Pinnacle DB Silica column, 5 × 150 mm × 4.6 mm, 3  $\mu$ m particles. GC: Hydrogen carrier on OV-5 1 m × 0.25 mm × 0.25  $\mu$ m, temperature program  $-20^{\circ}\text{C}$  to  $250^{\circ}\text{C}$  at  $1800^{\circ}\text{C min}^{-1}$  hold for 6 s. GC runs: 371.

(fig:B50)

long, with an internal diameter of 0.25 mm. The thickness of the stationary phase was 0.25  $\mu$ m.

The temperature was ramped from  $-20^{\circ}\text{C}$  at a rate of  $1800^{\circ}\text{C min}^{-1}$  to  $250^{\circ}\text{C}$ , where it was maintained for 6 s. After the temperature program had ended the column was cooled to  $-20^{\circ}\text{C}$  again and the next fraction was collected.

The detector was an FID at  $250^{\circ}\text{C}$ . The FID response was recorded, and so a total of 371 fast GC chromatograms were collected to compile the 2D chromatogram.

### 6.5.6 Results and discussion

(sec:B50Discuss) The chromatogram of the biodiesel/petrodiesel blend (Figure 6.7) shows that the blend was separated into two distinct groups of compounds. At a <sup>1</sup>D (SFC) retention time of about 160 s the hydrocarbons from the petrodiesel part of the blend elutes. To those familiar with petrochemical chromatography the characteristic unresolved complex mixture ("hydrocarbon hump") topped by alkane peaks will be instantly recognizable as a representing a petroleum product. (See Figure 6.8 for an example.)

Between the <sup>1</sup>D retention times of 300 s and 480 s the peaks of the FAMEs are seen, clearly separated by number of double bonds in the SFC dimension, and by

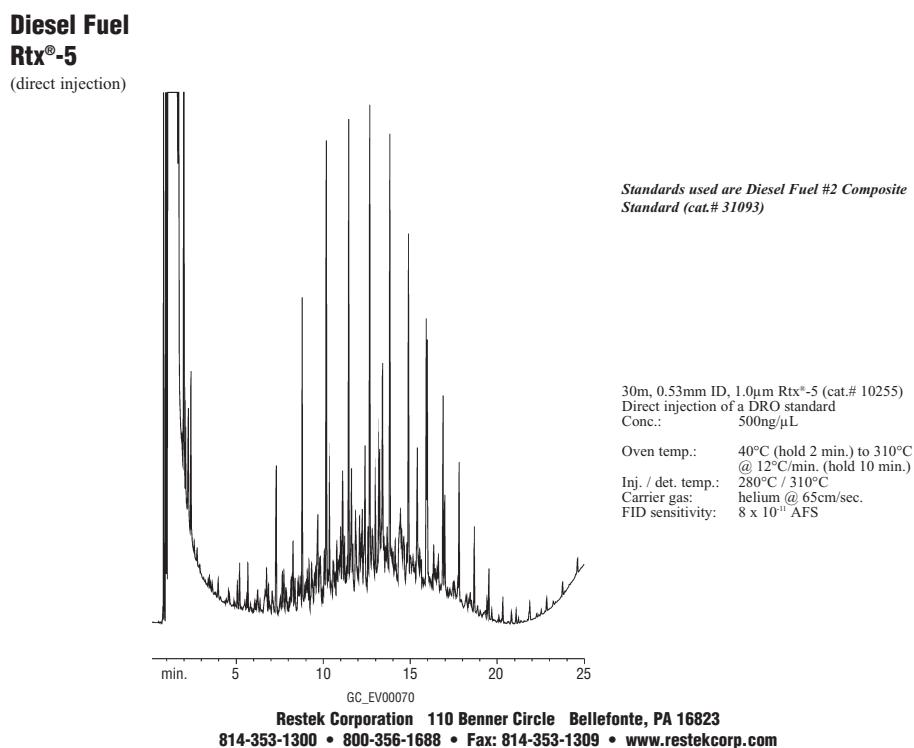


FIGURE 6.8: An example of a GC chromatogram obtained from a petroleum-based diesel fuel sample, showing the unresolved complex mixture topped by alkane peaks.

{fig:HCHump}

3635 volatility in the GC dimension. When compared to Figure 6.3, one can surmise that the feedstock for the biodiesel was waste sunflower oil.

It is worth pointing out that the hydrocarbons from the petrodiesel fraction of the blend were also separated on the SFC dimension (A. Venter, E. R. Rohwer, and Laubscher 1999a). This is not surprising: the standard ASTM D 5186 provides 3640 a method for the determination of aromatic compounds in diesel by SFC (ASTM 2019). The application of this method separates the unretained non-aromatics from the monoaromatics and from the polynuclear aromatics. The SFC×GC method provides additional information about the volatility of those aromatics which could help to quantify or identify them, but that falls outside the scope of this discussion.

3645 This separation shows that SFC×GC can be used to investigate the biodiesel component of petrodiesel/biodiesel blends without sample pre-treatment or specialized columns. The intrinsic orthogonality provides easy identification of components and the excess separation space allows for the addition of suitable internal standards for reliable qualitative and quantitative analysis with the robust FID with its predictable 3650 response. Expensive mass spectrometric detection is unnecessary.

Being able to determine the biodiesel content of biodiesel blends may prove useful in at least two scenarios:

The first is in monitoring blending. Biodiesel, essentially a mixture of esters, is quite polar compared to petrodiesel, essentially a mixture of hydrocarbons. This 3655 means that blending might not be a matter of just pumping the relevant volumes of the respective fuels into a tank and relying on diffusion to complete the mixing. Being able to determine the amount of biodiesel in different samples of a blend will provide assurance that the blend is homogenous and should perform according to expectations. While such a test is perhaps better performed using a spectroscopic 3660 method, SFC×GC can provide information for calibration, and will of course be invaluable during trouble shooting.

The second application of SFC×GC is in regulating biodiesel content. In some political environments fuels might be taxed according to their biologically-derived content, for example to promote agriculture or to meet carbon emissions targets. 3665 Sometimes biological content is not encouraged by taxation, but simply mandated. Such incentives and mandates are of course liable to corruption, and therefore the ability to monitor the blending of the fuels is needed. The most reliable way to differentiate between organic matter derived from biological sources and organic matter derived from fossil sources is a radiochemical method, where the content of radioactive  $^{14}\text{C}$  is determined. (Organic material from fossil sources contains no  $^{14}\text{C}$ .) The technical standard ASTM D6866 provides an approved method. But radiochemical methods require specialized equipment and trained staff, whereas fuel laboratories more often have experience with chromatographic techniques and might find 3670 SFC×GC a useful technique to provide evidence that a diesel fuel blend contains the stated amount of biodiesel.

## 6.6 Conclusion

Comprehensively coupled chromatography offers better analytical separations in three ways: improved peak capacity, improved sensitivity, and more structured data. The examples presented here do not pretend to prove improved peak capacity: 3680 biodiesel is not normally considered a chromatographically-challenging complex sample. Neither is improved sensitivity relevant: determining biodiesel components is not a trace analysis problem for which high sensitivity is necessary. What these

examples do show is that the high orthogonality between SFC and GC provides data that contains a clear and powerful pattern that greatly simplifies interpretation.

3685 In the example of FAMEs derived from various vegetable oils the pattern shows how the degree of unsaturation provides separation on the first dimension, and the volatility provides separation on the second dimension. These separations, combined in one 2D chromatogram, provides intuitive understanding of the chemical composition of the oils which facilitates prediction about the contribution these oils  
3690 will make to the final biodiesel product.

In the example of examining the FAME content of a biodiesel blend, the pattern shows that SFC readily separates the petroleum-derived hydrocarbons from the FAMEs of biological origin. The biodiesel content can be determined by focusing data analysis efforts on the relevant parts of the chromatogram.

3695 SFC $\times$ GC can assist with proving compliance: it provides a chromatographic method that combines the determination of total FAMEs and unsaturated FAMEs in a single method, and might provide a way to quantify the biodiesel component in blends with petrodiesel, and cost-effective, accurate and reliable quantitation of the biodiesel component in blends with petrodiesel can be expected.

3700 The provided examples also prove that it is possible to design a system that solves the problems of doing practical, rapidly-repeated fast temperature programmed GC, ramping temperatures from below ambient to the limits of the column in hundreds of consecutive, identical runs. This high performance is possible without resorting to advanced or expensive technology.



*It is a bad plan that admits of no modification.*

Publilius Syrus

# 7

3705

## Application: Modified SFC

### ⟨Chapter7⟩ 7.1 Introduction

The previous chapter described SFC $\times$ GC runs that used neat carbon dioxide as mobile phase for the SFC separation. Carbon dioxide as a mobile phase is compatible

3710 with a flame ionization (FID) detector: under the conditions found in the hydrogen-rich portion of the flame the carbon dioxide cannot be reduced to methane, the first step in the process that produces the ions that gives the FID its name. But most modern SFC development involves the use of modifiers and additives<sup>1</sup>. These modifiers are usually organic compounds, and in the FID flame they produce ions. When the 3715 modifier-containing eluate of an SFC column passes through the detector, the quantity of the modifiers will be much higher than quantity of the analytes. If the detector is an FID, the modifiers will produce many more ions than the analyte will produce, and therefore the the modifier signal will swamp the analyte signal. Mixing organic modifiers and additives with the carbon dioxide mobile phase therefore renders SFC 3720 incompatible with flame ionization as a detection method. As a result, most modern SFC chromatographs use optical detectors, most commonly UV spectrophotometry adopted from HPLC.

While the use of optical detectors has not hampered the development of SFC, they do have a few shortcomings, which might be imposing artificial limits on the 3725 exploitation of the versatility of carbon dioxide as a mobile phase:

- While SFC is compatible with optical detection (carbon dioxide does not absorb visible or ultraviolet radiation), there is a wide range of potential modifiers that are not compatible with spectrophotometric detection because they absorb strongly in the UV, but might have chemical properties useful for manipulating SFC separations.
- Optical detectors can only detect analytes with chromophores, and the intensity of the signal depends strongly on the species.

<sup>1</sup>Modifiers are added to the carbon dioxide to manipulate the solubility of the analytes in the mobile phase, in exactly the same way that mixtures of solvents are used in, say, gradient elution in HPLC. Additives are added in small amounts, and manipulate the interaction of the analyte with the stationary phase.

- 3735 • Optical detectors have a limited dynamic range, which complicates quantification.
- Expensive 'HPLC-grade' solvents do not provide better chromatography than other grades of solvents: the processes used to purify them are just optimized to reduce the amount of UV-absorbing impurities that can interfere with analysis.

3740 In comparison, the FID can detect practically any organic compound over a wide concentration range, giving an integrated-over-time signal proportional to the mass of the compound.

If only there were a way to separate the analytes modifiers from the modifiers before detection ...

3745 When GC is comprehensively coupled to SFC, the modulator will focus each fraction of SFC eluate, and then the GC will separate all the compounds in the fraction, be they analytes, modifiers, or additives. If the modifiers are more volatile than the analytes then the modifiers will elute before the analytes, and the analyte can be detected or quantified using an FID while the signal from the modifier is ignored.

3750 Of course it is quite likely that the peak from the modifiers will **overload** the GC. Both the detector and the column may be overloaded. If the *detector* is overloaded, then it just means that the detector no longer responds linearly to the amount of material eluting from the column — this might include responding with its maximum output, leaving flat-topped peaks. If the *column* is overloaded, then the capacity of the stationary phase is exceeded and the peaks will **front**: they will be asymmetrical, 3755 with the side of the peak towards earlier elution times having a slope smaller than the slope of non-overloaded peaks, and the side of the peak towards later elution times having a larger slope.

3760 The requirement that the modifiers added to the SFC mobile phase be volatile (so that they can elute before the analyte) is not too onerous a restriction. Higher volatility correlates with lower diffusivity, so that the preferred high-diffusivity modifiers for SFC would also tend to be volatile enough to elute early on GC.

A practical SFC $\times$ GC chromatograph opens up the field for the use of carbon dioxide mixed with modifier as a mobile phase for analysing complex mixtures found in biodiesel production, biodiesel quality control, and compliance monitoring.

## 3765 7.2 SFC $\times$ GC using modified carbon dioxide.

In this section a SFC $\times$ GC chromatogram is presented that shows that methanol added as a modifier to the SFC mobile phase does not interfere with the fast GC any more than a sample's solvent interferes in the usual 1D GC. The 2D separation space of SFC $\times$ GC therefore gains in control over elution in the SFC dimension at the 3770 cost of losing the capability to detect compounds more volatile than methanol on the GC dimension.

### 7.2.1 Sample

3775 The sample was a 1:1 blend of petrodiesel and biodiesel, prepared in the laboratory. The biodiesel sample was donated by a commercial testing laboratory, and had been produced from palm oil by an anonymous producer. The petrodiesel (Shell Extra Diesel 500 ppm) was obtained from a commercial filling station.

### 7.2.2 SFC

The SFC used carbon dioxide at 200 bar inlet pressure and room temperature as mobile phase. The flow rate is of carbon dioxide at atmospheric pressure was reportedly 3780 175 ml min<sup>-1</sup>. The carbon dioxide mobile phase was modified with 5 % mass fraction of HPLC-grade methanol (Merck LiChrosolv). The column consisted of five HPLC bare silica columns (150 mm  $\times$  4.6 mm, 3  $\mu$ m particles) (Restek, Pinnacle DB Silica) connected in series.

### 7.2.3 Modulation

3785 A fraction was collected for each 5 s SFC flow time. After the flow from the SFC was stopped, 3 s was allowed for the last vapour in the inlet to be swept into the cold ( $-20^{\circ}\text{C}$ ) column for trapping, and then the inlet vent valve was opened for 1 s to relieve excess inlet pressure.

### 7.2.4 GC

3790 The GC column was a Restek Rxi-5Sil MS 0.250 mm  $\times$  0.25  $\mu\text{m}$   $\times$  1 m.

For each fast GC program the temperature was ramped from  $-20^{\circ}\text{C}$  to  $320^{\circ}\text{C}$  in 10.3 s (33  $^{\circ}\text{C s}^{-1}$ ) and kept there for 10 s before cooling the column down to  $-20^{\circ}\text{C}$  to prepare for trapping the next fraction.

A total of 233 fast chromatograms were recorded and combined into a 2D chromatogram.

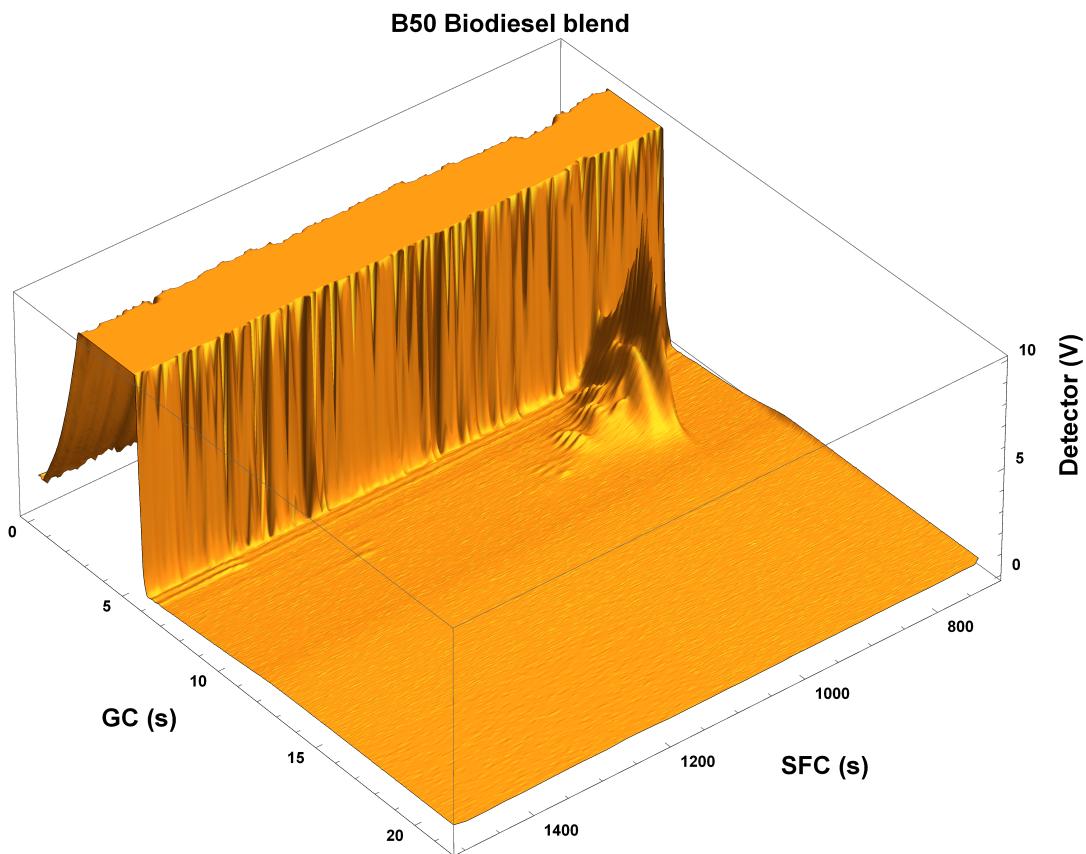
### 7.2.5 Results and discussion

The chromatogram obtained for this run is shown in Figure 7.1. Note the unusual orientation of the chromatogram, chosen to best show the data: In the SFC dimension the later elution times are closer to the reader.

3800 The intention of this chromatographic run was to explore the effect of the methanol modifier on the retention behaviour of the biodiesel fraction of the B50 blend. However, the run time was not long enough for the FAMEs to elute and the only peaks on the chromatogram are the group of hydrocarbons from the petrodiesel. Nevertheless, the chromatogram shows the possibilities that arise when a modifier is added to the SFC mobile phase.

3805 The 'wall' that appears on the chromatogram is the tailing edge of the peak of the modifier added to the SFC. These peaks are present in all the fractions at equal concentration. The flat top of the peak is caused by **detector overload**. It should be emphasized that this overload is not caused by saturation of the FID response, but by limits imposed by the chosen signal amplifier. The FID has a dynamic range of  $10^6$  and could comfortably accommodate the modifier peak, but the amplifier gain was chosen to best match the analyte signal with the input range of the 16-bit analog-to-digital converter.

3815 The 'corrugations' on the wall can be ascribed to two sources. First is variations in retention time of the modifier peak, which can be caused by variations in GC gas flow and variations in the fast temperature programs. Second, variations in the amount of modifier collected by the modulator will also cause slight changes in the apparent position of the peak. This second reason is probably the major contributor for this run, caused by variations in the timed collection period introduced by the 3820 imprecision of the electromechanical valve actuator and by variations in the SFC flow.



---

FIGURE 7.1: The modifier used in the SFC dimension elutes as a solvent front on the GC dimension, and does not otherwise affect the separation.

{fig:Modifier}

The chromatogram shows the hydrocarbon “hump” in the GC dimension and the separation of the aromatics in the SFC dimension, as discussed in Section 6.5.6.  
3825 It can be seen that the ‘solvent peak’ of the modifier interferes with the more volatile of the hydrocarbons, but the rest of the separation space is available for separation.

### 7.2.6 Conclusion

Adding modifiers to SFC expands the versatility of carbon dioxide as a mobile phase, but in 1D SFC the use of modifiers precludes the use of the universal flame ionization detector (FID), because the modifier signal will swamp the analyte signal.

3830 Adding GC-FID as a <sup>2</sup>D separation allows the use of any volatile modifier to manipulate retention in the SFC dimension, including modifiers that might not be UV-transparent. Modifiers present in the fraction collected by the modulator elute like the customary solvent peak in the GC dimension, which means that the signal from the modifier does not interfere with the signal of the analyte.

3835 The example shown above shows that SFC $\times$ GC will reliably separate any modifiers used in the SFC dimension from the analytes, allowing the FID to be used as a detector for gradient-elution SFC separations.



*"Begin at the beginning," the King said, very gravely,  
"and go on till you come to the end: then stop".*

Lewis Carroll

# 8

## Conclusion

`\todopage{1}` To do (1) To do (2)  
<sup>3840</sup>  
`\tchapage{82}`

### 8.1 Introduction

The world needs to change. If not you, who? If not now, when?

### 8.2 Synopsis

#### 8.2.1 Chapter 1: Introduction: Climate, fuel, and biodiesel

- <sup>3845</sup> The molecular basis of sustainability (Anastas and Zimmerman 2016) demands that chemists discuss sustainability in terms of compounds. Chapter 1 opens with the idea that energy is an essential component of industrialized societies, but that choosing of fossil fuels as our source of energy causes pollution, which threatens to nullify the benefits they bring. It describes the process which causes carbon dioxide to be emitted in large quantities, and how its interaction with planetary radiation makes it a pollutant. The concept of "carbon footprint" is introduced, which allows the comparison of activities in terms of their carbon pollution, enabling decision makers to select the least polluting option. The discussion then focuses on internal-combustion engines, which is a major source of carbon and noxious pollution, and it <sup>3850</sup> is shown that higher-efficiency engines have lower carbon footprints. A discussion of ways to reduce the carbon footprint of internal combustion engines shows that, in the cases where they cannot be replaced by electric motors, a reduction of carbon footprint can be obtained by preferring large, high-performance Diesel engines fuelled by biodiesel. The discussion concludes with the idea that the success of such <sup>3855</sup> engines will demand high quality biodiesel, and that chromatography will play a central role in ensuring that quality.

#### 8.2.2 Chapter 2: Introduction: Carbon dioxide and chromatography

Chapter 2 starts with a discussion of the chemical industry and the need to move towards "green" chemistry. It introduces carbon dioxide as a renewable resource,

3865 discusses its various uses in industry, and then focuses on its application in extraction and chromatography. It describes how carbon dioxide becomes a solvent at high pressures and densities, and then introduces supercritical fluid chromatography (SFC). Fractions of eluate from SFC can be analysed by gas chromatography, and if a suitable set of criteria is met, then the combination is called SFC $\times$ GC.

### 3870 **8.2.3 Chapter 3: Biodiesel standards**

As discussed in Chapter 1, a reliable high-performance engine requires a reliable fuel. Chapter 3 discusses the concepts of technical standards, which establishes requirements with which fuels must comply. The discussion then focuses on the technical standard SANS 1935, which lists the requirements that biodiesel offered for sale in South Africa must comply with, concluding with the role chromatography plays in

### **8.2.4 Chapter 4: Instrumentation: Supercritical Fluid Chromatography**

3880 Chapter 4 explains the experimental equipment used for chromatography using high-pressure carbon dioxide as a mobile phase. It starts with describing the mobile phase, how it is stored and pumped, how modifier is added, and how the sample is injected. It describes problems with designing the restrictor that maintains the pressure, and concludes with remarks about using a gas chromatograph as a detector for SFC.

### **8.2.5 Chapter 5: Instrumentation: Fast GC**

3885 Chapter 5 opens with a discussion on the time aspect of comprehensive SFC $\times$ GC chromatography, and shows that for practical analysis the GC dimension must be *fast*. The theory of fast GC is discussed, which leads to the need for fast temperature programming. The design of a resistively heated coaxial heater is described, including its calibration and control. The discussion then covers the need for a cold column, and the design of coaxial cooling using boiling liquid carbon dioxide is described. Next, the discussion covers the design of hardware to mount the coaxial heater of the fast GC in a conventional GC oven. The chapter concludes with a description of the detector and the data flow from signal to final chromatogram.

### **8.2.6 Chapter 6: Investigating biodiesel feedstock using SFC $\times$ GC**

3895 Chapter 6 discusses the application of the developed SFC $\times$ GC instrument to the study of fatty acid profiles of various potential biodiesel feedstocks, and shows that SFC $\times$ GC can be used to separate biodiesel from petrodiesel in biodiesel blends.

### **8.2.7 Chapter 7: Application: Modified SFC**

3900 Chapter 7 demonstrates that the use of SFC with modifiers does not preclude the use of the flame ionization detector when GC is used as a second dimension.

## **8.3 Contributions of this study**

The research presented in this thesis contributes to the field of chromatography in the following aspects:

### 8.3.1 Coaxial heater cooling

<sup>3905</sup> The biggest new contribution is proving the concept of active cooling of short chromatography columns using precisely-applied carbon dioxide. To the best of our knowledge there is no other fast temperature-programmed chromatograph that has sub-ambient ramp start temperatures or a faster repetition rate.

### 8.3.2 Coaxial heater reliability

<sup>3910</sup> The success of the SFC $\times$ GC is undoubtedly attributable to the high repeatability of the fast temperature programming. The

### 8.3.3 Length of SFC column

## 8.4 Voids in this study

### 8.4.1 Lack of optimization

<sup>3915</sup> 8.4.2 Luck with voltage ranges

## 8.5 Design

Although this was a scientific project, I learned a lot about design. If I'd start the project again

## 8.6 Disasters

<sup>3920</sup> 8.6.1 About the isolation amplifier

### 8.6.2 Thinking that we'd elute alcohols

We could separate paraffins, olefins and aromatics, using the long column.

### 8.6.3 Variable modulation period

<sup>3925</sup> It is possible to vary the modulation period during an SFC $\times$ GC run. This would allow one to speed up analysis, by collecting fewer fractions where less information would be expected.

## 8.7 Limitations of this SFC $\times$ GC design

### 8.7.1 Cooling control

The fast GC temperature was poorly controlled in the  $-20^{\circ}\text{C}$  to  $50^{\circ}\text{C}$  region of the <sup>3930</sup> temperature ramps, because the coaxial heater was not isolated. It was mounted in a GC oven, together with the heated T-piece blocks. The heat leaking from these uninsulated blocks accumulated in the oven, until the oven reached a temperature of about  $50^{\circ}\text{C}$ . Once the coolant flow stopped the temperature of the coaxial heater would immediate rise to the temperature of the oven.

<sup>3935</sup> This was not a problem in the chosen biodiesel application, but better control over the beginning of the temperature ramp would be necessary

## 8.8 Suggested design improvements

### 8.8.1 Four-wire resistance measurement

The resistive heater had a design that depended on measuring the resistance of the  
3940 thin-walled stainless steel tube. The resistance was measured by comparing the potential difference between the ends of the heater with the potential difference over a reference resistor carrying an identical current. The circuit that carried the current also measured the potential difference. Because the current was high, the circuit had to be constructed in such a way that it no other significant resistance in it, which  
3945 meant that care had to be taken to used heavy-guage cable and only use soldered joints.

These troubles could have been avoided if the current circuit and the measuring circuit were separated, using 'four wire' resistance measurement. In such a design the potential difference is measure using a circuit that connects directly to the voltmeter. The voltmeter has a high input impedance, which means that the current that the circuit carries is low, and stray resistances will play much less of a role. Also, this would have made it possible to use bolted joints on the current-carrying circuit, which could have simplified operations.  
3950

### 8.8.2 Legs heating and cooling integrated in detector and inlet

### 8.8.3 Siltek-treated coaxial heater

(`todo` page: 3) To do (3)

**To do...**

`<todobl1:1>`

- 1 (p. 117): Move `zlabelchap:pdfstartpage` to appropriate chapter. Remove before publication

`<todobl1:2>`

- 2 (p. 117): Remove `usepackage{stampinclud}` before publication.

`<todobl1:3>`

- 3 (p. 120): Discuss material that passes through column during trapping.





## Frequently Asked Questions

### ?⟨AppendixA⟩? A.1 How do I change the colors of links?

3965 The color of links can be changed to your liking using:

`\hypersetup{urlcolor=red}, or  
\hypersetup{citecolor=green}, or  
\hypersetup{allcolor=blue}.`

If you want to completely hide the links, you can use:

3970 `\hypersetup{allcolors=.}, or even better:  
\hypersetup{hidelinks}.`

If you want to have obvious links in the PDF but not the printed text, use:

`\hypersetup{colorlinks=false}.`



# B

## Glossary

3975

?⟨AppendixB⟩? A list of technical terms that might be unfamiliar to the reader.

**Azeotropic mixture** A mixture of liquids in which the vapour phase and the liquid phase have the same composition. The mixture cannot be further purified by simple distillation.

<sup>3980</sup> **Payload** The portion of the weight of a vehicle's load that is in excess of what it needs to complete a journey. This



# C

## Thermocouple Production

?<sup>(Appendix C)</sup>? How the thermocouple probe was made:

3985

- 1.
2. A piece of 0.25 mm polyimide-coated fused-silica capillary of about 500 mm length was cut and mounted with sticky tape on a wooden metre stick.
3. A longer length of 0.1 mm fused-silica capillary was threaded through the 0.25 mm capillary.

3990

4. The end of one of the thermocouple wires was inserted into the end of the 0.1 mm capillary. A drop of cyanoacrylate adhesive was touched to the end of the capillary. Capillary action drew the liquid adhesive into the capillary and fixed the wire in place.

3995

5. The wire was drawn carefully into the capillary by pulling on the 0.1 mm capillary.
6. Once the end of the wire protruded through the end of the 0.25 mm capillary the end of the 0.1 mm capillary was cut off.

4000

7. The wire was anchored at one end with adhesive tape, pulled tight, and anchored at the other end.

8. The procedure was repeated for the other wire.

9. The two thermocouple wires (Goodfellow) was clamped in a twisting bar. The twister bar has a square profile, 8 mm on a side.

4005

10. The wires were flamed with a cigar lighter until they were red a dull red hot. (At any higher temperature the wires would melt.) This chars the polyimide coating.

11. The flamed portion of the wires were lightly sanded with 1200 grit water paper. A pair of small pliers had its beak lined with the abrasive, and lightly stroked up and down the wire to remove the char.

12. A 6 mm tube was inserted between the wires to serve as spacer. and moved until about 10mm away from the twister bar.  
4010
13. The wires were twisted by turning the twister bar until the twister portion was about 5mm long.
14. The spools were rewound to retract the wire, until the start of the twist rested on the clamping bar.
15. The clamping weight was lowered onto the clamping bar, keeping the pair of wires in place.  
4015
16. A small pair of scissors was used to snip off the end of the
17. The welding electrode was brought into position. This was a carbon rod in the form of a pencil lead (Schwann Stabilo), 2 mm in diameter, mounted on a screwing connector. The welding circuit consisted of a bench power supply set to approximately 20V. connected. The negative terminal was clamped to the aluminium base plate of the microscope, on which rested the brass clamp bar. The positive terminal was clamped to the carbon electrode. A voltage of approximately 23 V was applied.  
4020
18. It was discovered that the carbon electrode should not have a polished end, but a roughly broken end.  
4025
19. The electrode was moved closer to the clamped twisted wire.
20. At the right point a spark would jump from the carbon to the wire, melting the end of the wire. The molten wire would draw into a globule on the end of the wire, withdrawing from the electrode and so breaking the spark, ending the heating.  
4030
21. If the wire would actually touch the electrode the wire would heat up red hot and melt off, usually destroying the twist and requiring making a new twist.
22. If all went well, there would be a hemispherical weld at the end of the twist where the two wires would be joined.  
4035
23. The thermocouple wires was withdrawn into the capillary until the end just protruded, kept in place with a pair of rubber-tipped self-closing tweezers.
24. If two sets of thermocouples were needed, the procedure would be repeated for another pair of wires.
25. The wires would be pulled back, one pair at a time.  
4040
26. The other end of the capillary was taped to the connector pad.
27. The wires was flamed and scraped to remove the polyimide isolation, and screwed down on a screw connector block.
28. The resistance between the protruding end of the thermocouple and the connector block was measured to ensure electrical connection. The resistance for the Chromel is  $1440 \Omega m^{-1}$ , and for the Alumel  $600 \Omega m^{-1}$   
4045
29. A drop of cyanoacrylate adhesive was put on the end of the capillary to anchor the wires and to prevent high-pressure gas from blowing the wires out of the capillary.

# D

## Engineering Drawings

4050

?{AppendixD}?

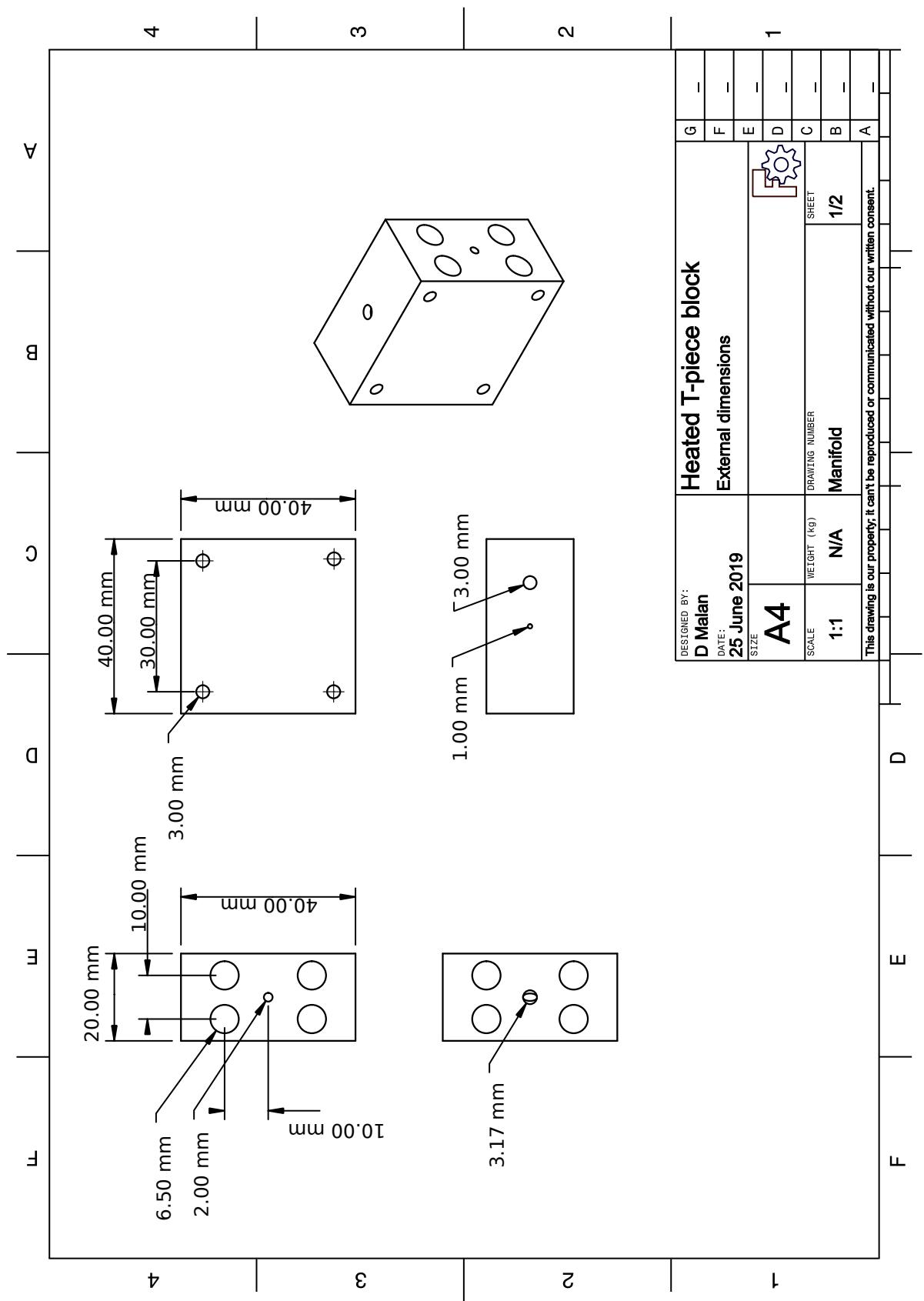


FIGURE D.1: A technical drawing showing the dimensions of the heated T-piece blocks.

`<fig:ManifoldDims>`

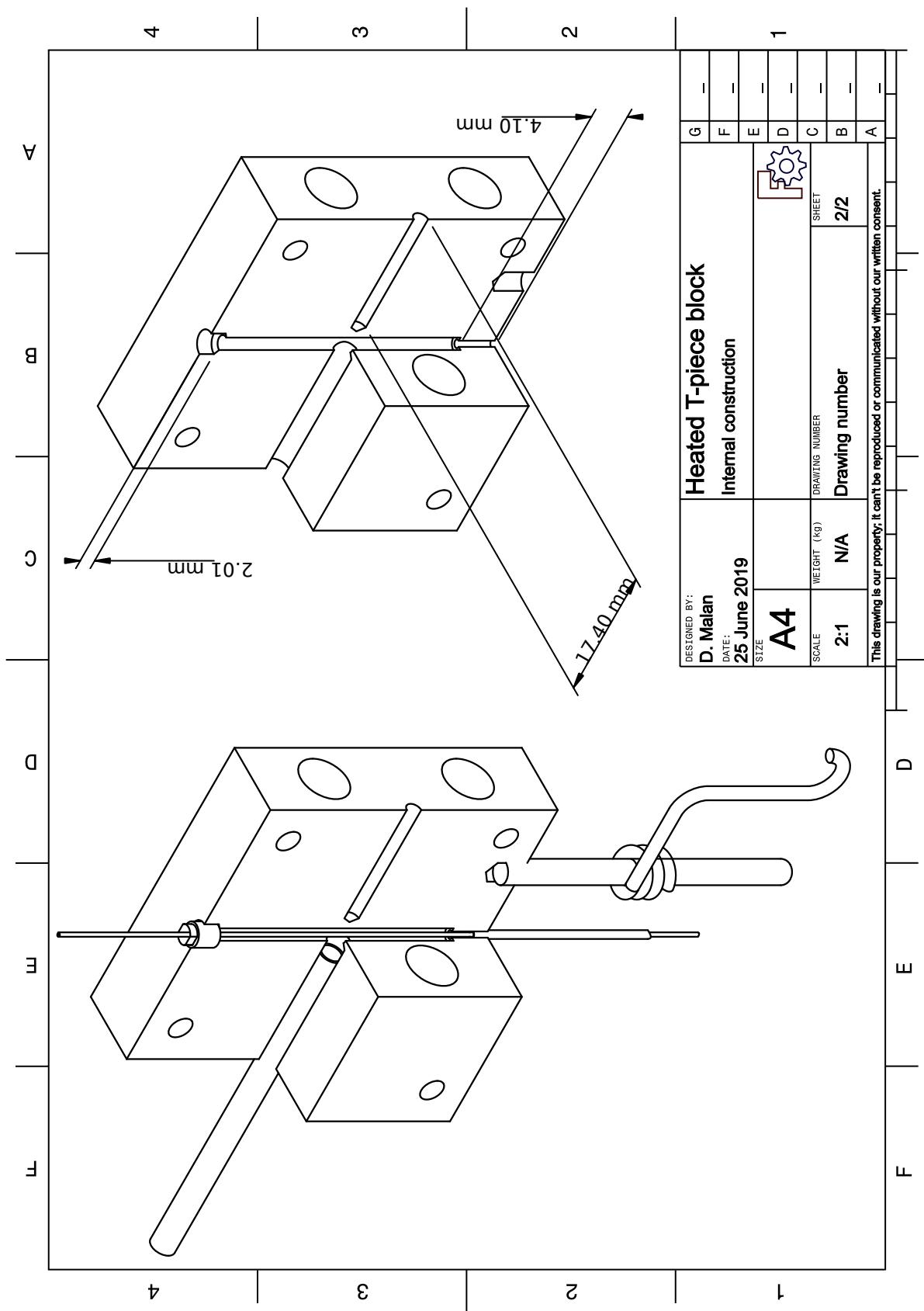


FIGURE D.2: A technical drawings of the internal construction and assembly of the heated T-piece blocks.

*(fig:ManifoldAssy)*

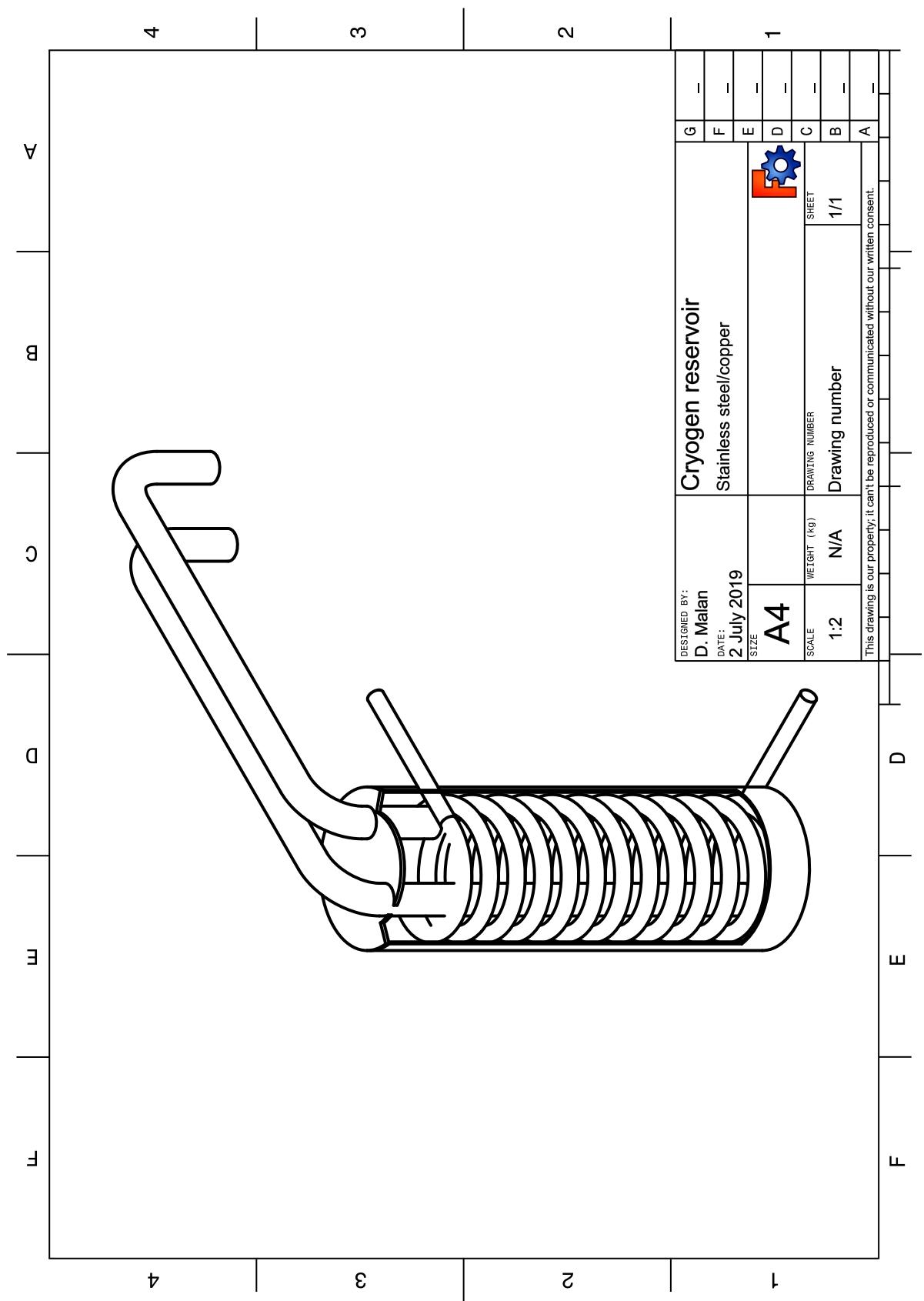


FIGURE D.3: Cut-away technical drawing of coolant reservoir.

&lt;fig:CryogenReservoir&gt;

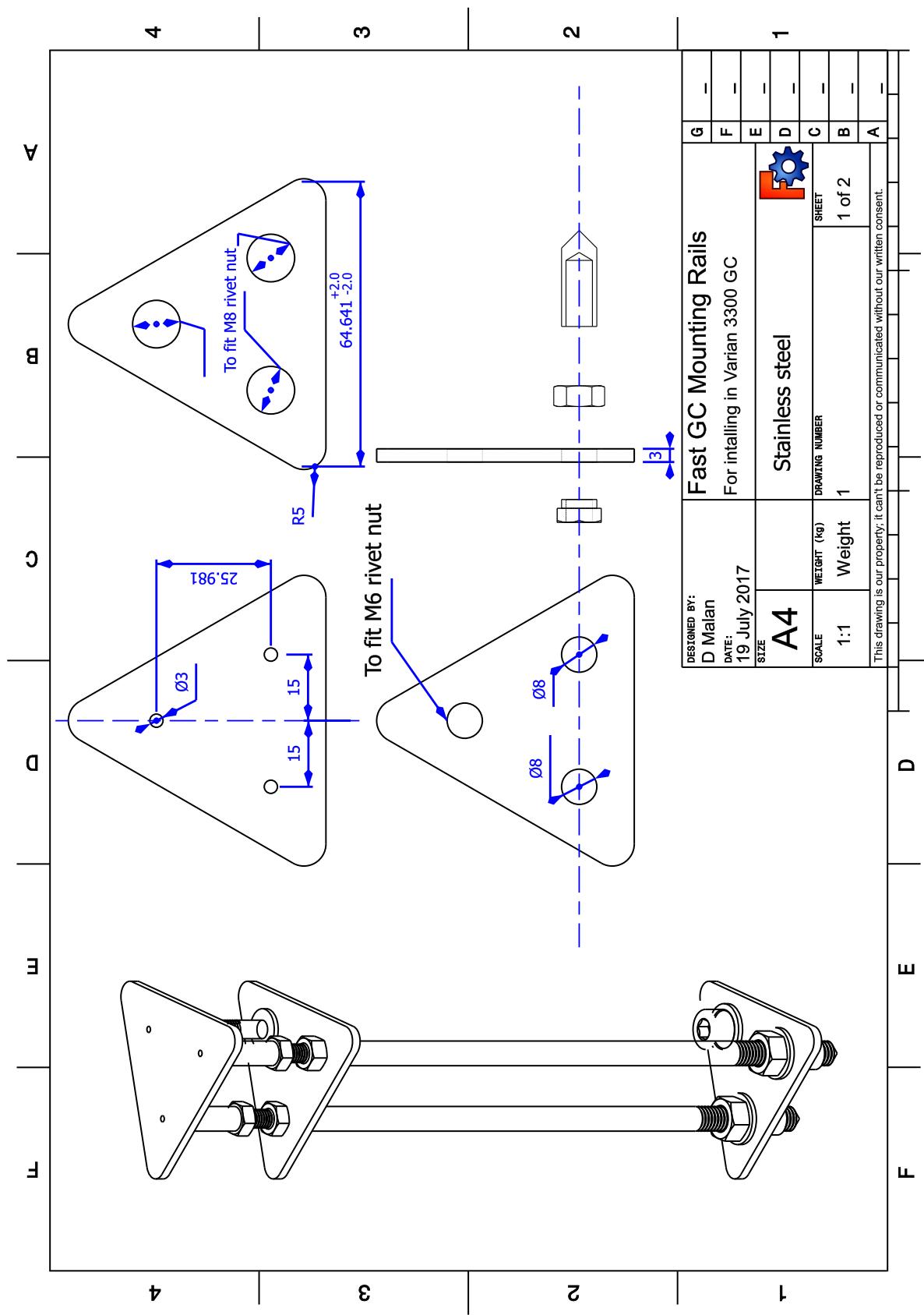


FIGURE D.4: A technical drawing of the rails carrying the T-piece mounting block.  
 (fig:RailsDrawing)

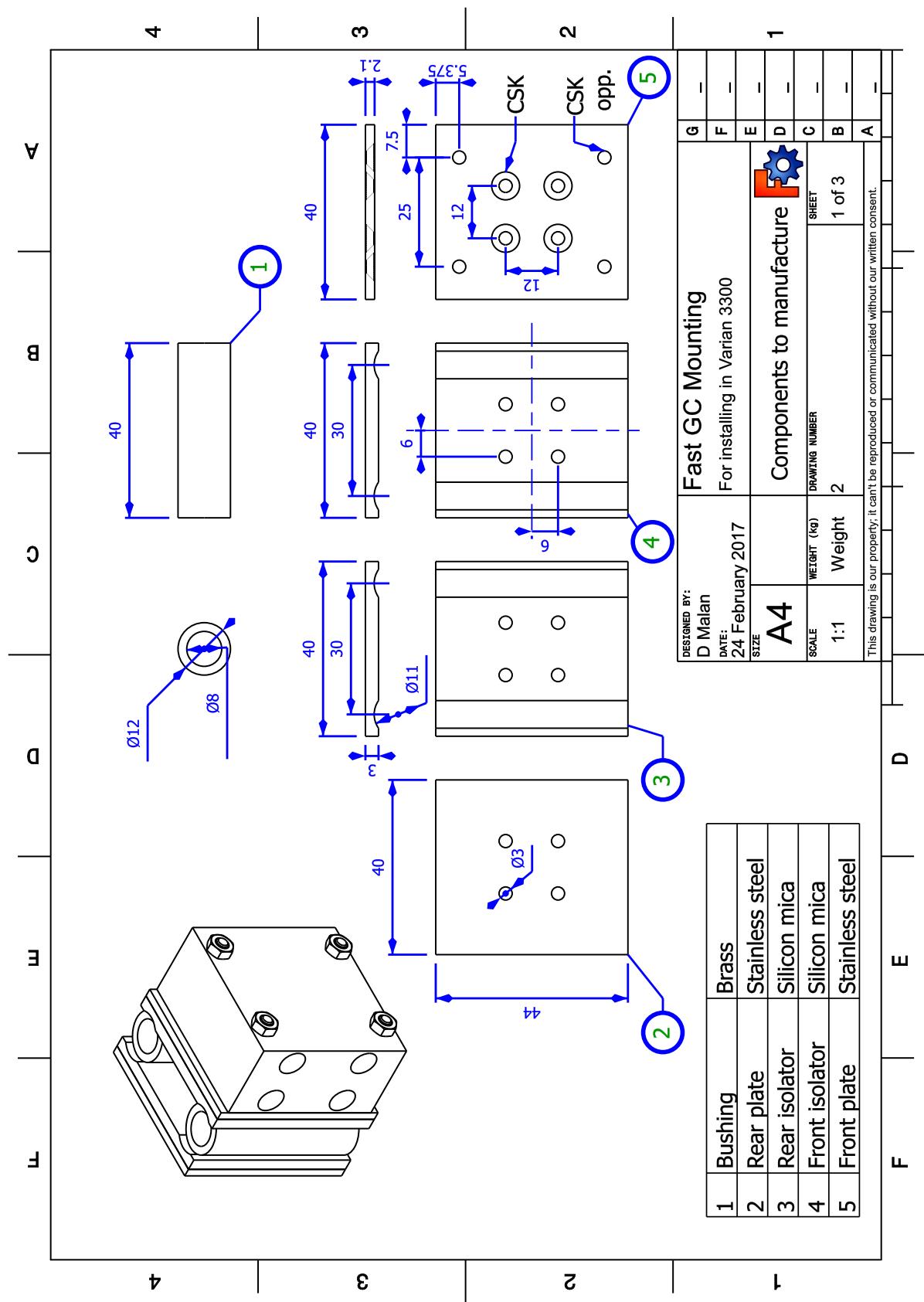


FIGURE D.5: A technical drawing of the T-piece block mounting, showing parts and dimensions.

{fig:CarsDrawing1}

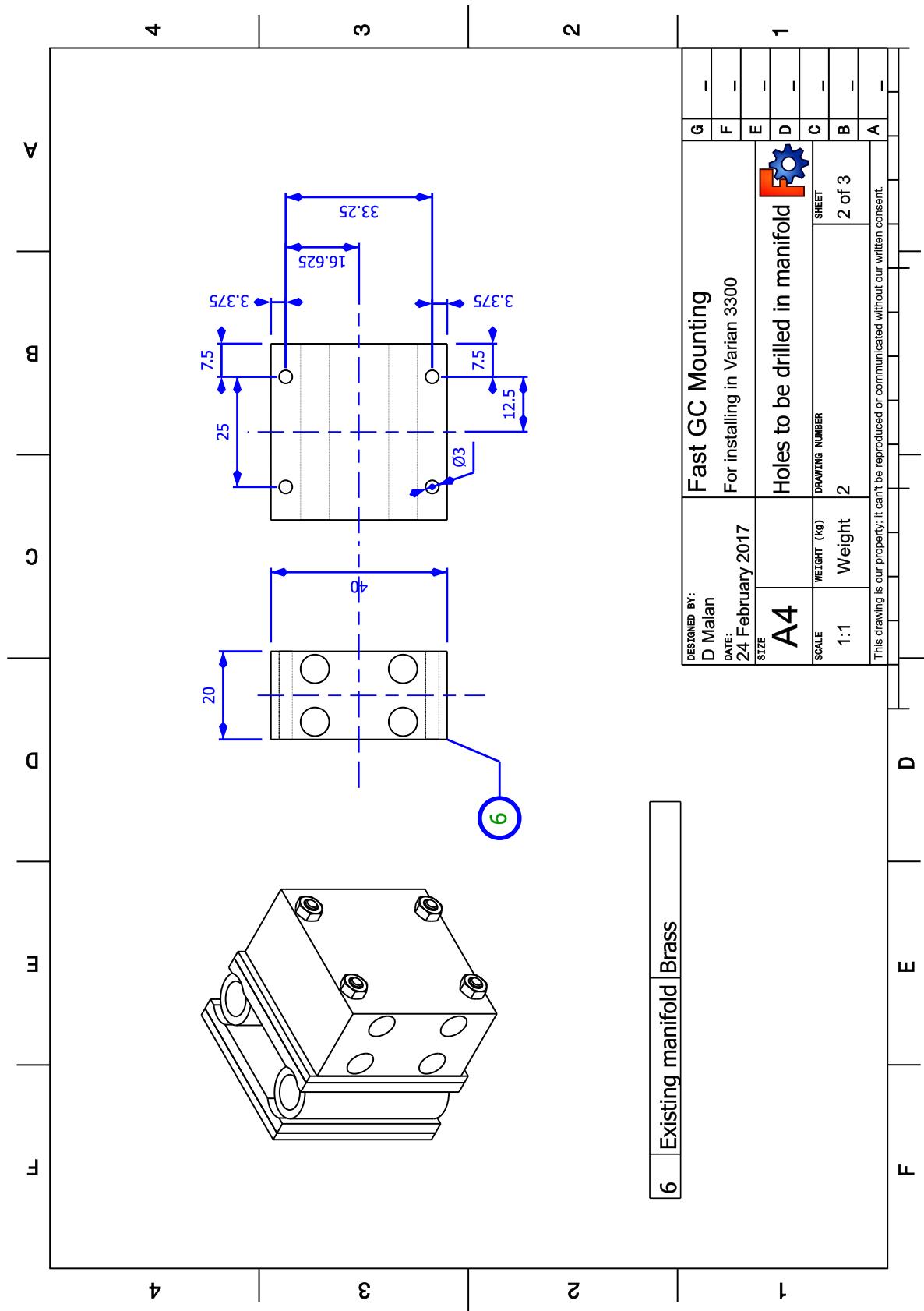


FIGURE D.5: (continued) A technical drawing of the T-piece block mounting, showing final assembly and dimensions.

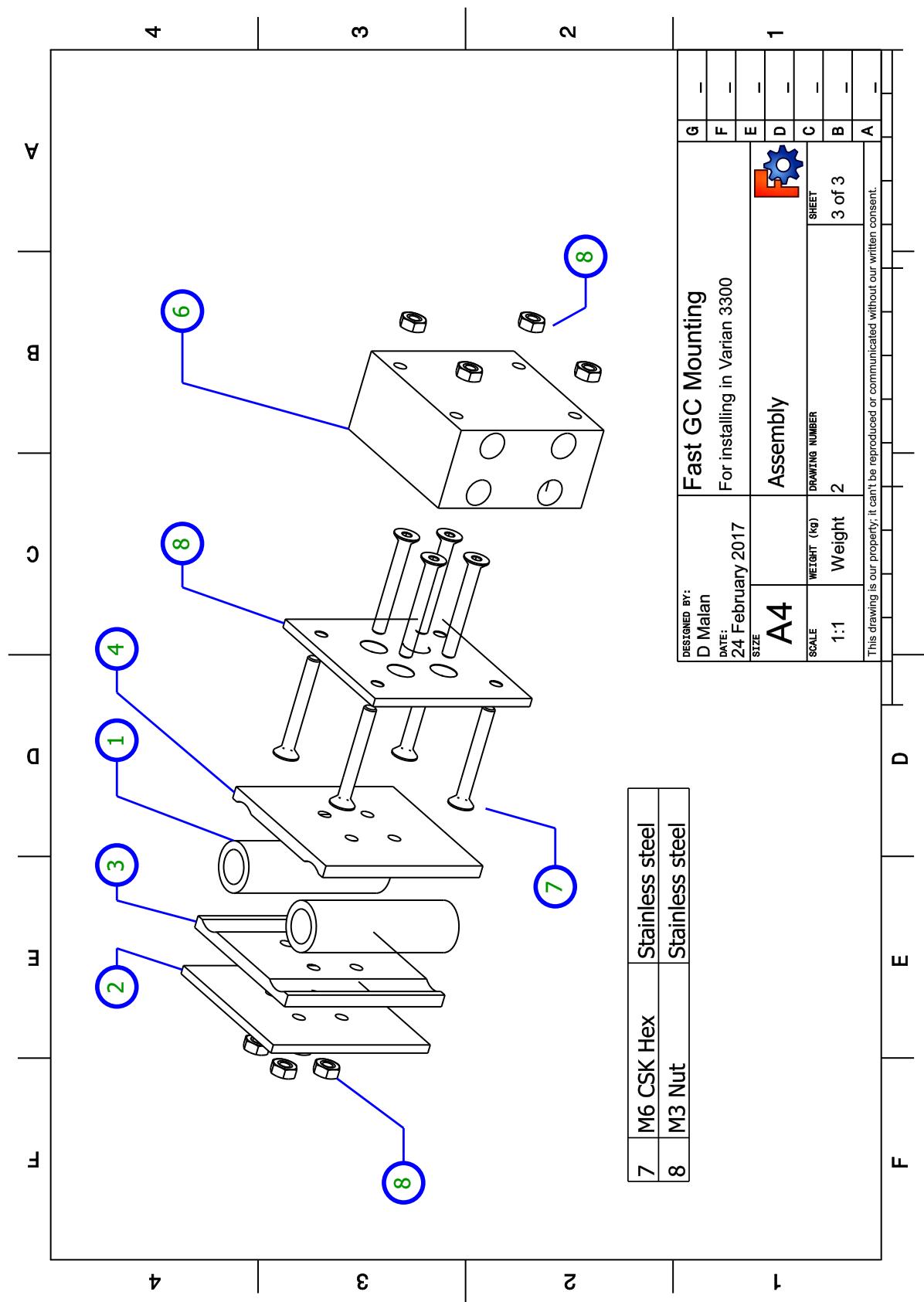


FIGURE D.5: (continued) A technical drawing of the T-piece block mounting, showing assembly method.

# E

## Fatty acid profiles

?{AppendixE}?

TABLE E.1: The fatty acid profile of sunflower oil sample

SunflowerPrecisionOils)

Fatty acid	Fraction
C14:0	0.05
C16:0	6.14
C16:1	0.04
C17:0	0.05
C17:1	0
C18:0	5.77
C18:1 t	0
C18:1 c	21.21
C18:2 t	0
C18:2 c	63.91
C18:3n6	0.11
C18:3n3	1.04
C20:0	0.47
C20:1	0.24
C20:2	0
C21:0	0
C22:0	0.77
C22:1	0
C24:0	0.19
C24:1	0

TABLE E.2: Fatty acid profile of coconut oil according to the FAO  
(Joint FAO/WHO Codex Alimentarius Commission 2019)

?<tab:CoconutFAO>?

Fatty acid	Lower	(Upper)
C6:0	ND	0.7
C8:0	4.6	10
C10:0	5	8
C12:0	45.1	53.2
C14:0	16.8	21
C16:0	7.5	10.2
C16:1	ND	
C17:0	ND	
C17:1	ND	
C18:0	2	4
C18:1	5	10
C18:2	1	2.5
C18:3	ND	0.2
C20:0	ND	0.2
C20:1	ND	0.2
C20:2	ND	
C22:0	ND	
C22:1	ND	
C22:2	ND	
C24:0	ND	
C24:1	ND	



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