

**Date: 4-06-2021**

**Class: 10<sup>th</sup> Genesis**

**Subject: Science**

**Test code: SEP06(21041306)**

**Chemistry**

***M. Marks: 20***

1. (a) What happens when a piece of iron metal is placed in copper sulphate solution? Name the type of reaction involved. (2 marks)  
(b) write balanced chemical equation with state symbols for the following reaction:  
Barium chloride solution reacts with sodium sulphate solution to give insoluble barium sulphate and a solution of sodium chloride
2. What is meant by (a) displacement reaction, and (b) double displacement reaction? Explain with the help of one example each. (2 marks)
3. What happens when a zinc strip is dipped into a copper sulphate solution?  
(a) Write the equation for the reaction that takes place. (2 marks)  
(b) Name the type of reaction involved.
4. What is a decomposition reaction? Give an example of a decomposition reaction. Describe an activity to illustrate such a reaction by heating. (3 marks)
5. (a) Explain the term “Corrosion” with an example. Write a chemical equation to show the process of corrosion of iron. (3 marks)  
(b) What special name is given to the corrosion of iron?  
(c) What type of chemical reaction is involved in the corrosion of iron?
6. (a) Explain the term “rancidity”. What damage is caused by rancidity? (3 marks)  
(b) What type of chemical reaction is responsible for causing rancidity?  
(c) State and explain the various methods for preventing or retarding rancidity of food.
7. (a) What happens when an aqueous solution of sodium sulphate reacts with an aqueous solution of barium chloride? (5 marks)  
(b) Write the balanced chemical equation for the reaction which takes place.  
(c) State the physical conditions of reactants in which the reaction will not take place.  
(d) Name the type of chemical reaction which occurs.  
(e) Give one example of another reaction which is of the same type as the above reaction.

