

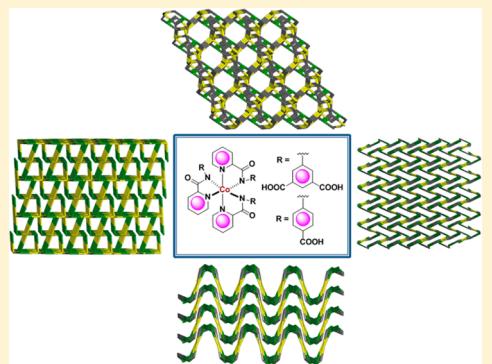
Three-Dimensional Heterometallic Coordination Networks: Syntheses, Crystal Structures, Topologies, and Heterogeneous Catalysis

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Supporting Information

ABSTRACT: This work presents the synthesis of $\{\text{Co}^{3+}\text{-Zn}^{2+}\}$ and $\{\text{Co}^{3+}\text{-Cd}^{2+}\}$ heterometallic coordination networks. These networks are originated from two unique Co^{3+} -based metalloligands containing appended arylcarboxylic acid groups at the strategically placed positions. Such appended arylcarboxylate groups coordinate the secondary metal ions, Zn^{2+} and Cd^{2+} , to afford distinct three-dimensional networks. All four networks display orderly arrangement of secondary metal ions and unique network topologies including an unprecedented one. These networks have been shown to act as the heterogeneous and reusable catalysts for the Knoevenagel condensation reactions and cyanation reactions of assorted aldehydes. Cyanation reactions nicely demonstrate the substrate size-exclusion catalysis.



INTRODUCTION

Coordination networks are important due to their well-defined architecture^{1–10} and applications in sorption,^{11,12} separation,^{13,14} sensing,^{15–17} ion-exchange and transport,^{18,19} proton conduction,^{20–22} optics,^{23–27} magnetism,^{28,29} devices,^{30,31} and catalysis.^{32–36} Due to their ordered architecture, porosity, and stability, such materials have shown significant results as heterogeneous catalysts for assorted organic transformations.^{32–40} It is important to note that such catalytic properties are largely feasible due to the crystalline nature and subsequent orderly arrangement of catalytic sites and the materials' ability to support facile diffusion of substrate and/or reagent.^{32–40} A large number of homometallic coordination networks have been synthesized and used for various applications including heterogeneous catalysis. On the other hand, heterometallic coordination networks (HCNs) have not been explored in such great deal due to the synthetic challenge of placing two different metals in close proximity. HCNs typically display higher thermal stability while also offering the possibility to explore the influence of primary metal on the properties of the secondary metal ion. In this context, a well-defined metalloligand,¹ already containing a primary metal ion, is most suited to react with a suitable secondary metal ion to produce HCNs.

Our research group has developed several metalloligands appended with assorted functional groups that afford systematically designed ordered architectures.^{41–54} Depending on the number, position, and orientation of the appended functional groups being offered from a metalloligand, we have been able to showcase the synthesis of hydrogen bonded assemblies,^{41,42} trimetallic complexes,^{43–46} as well as ordered 2D and 3D architectures.^{47–54} These examples illustrate the importance of

metalloligands for the construction of HCNs compared to homometallic networks typically obtained from the conventional reaction between a ligand and a metal ion. In addition, judicious selection of the Lewis acidic and/or redox-sensitive secondary metals in such designed materials has been utilized by displaying homogeneous as well as heterogeneous catalysis.^{43–54} Herein, we report the syntheses, crystal structures, topological analyses, and catalytic properties of two $\{\text{Co}^{3+}\text{-Zn}^{2+}\}$ and two $\{\text{Co}^{3+}\text{-Cd}^{2+}\}$ HCNs using two different Co^{3+} -based metalloligands offering appended aryl carboxylic acid groups. These HCNs display interesting network topologies including one unprecedented topology. We further illustrate the applications of such HCNs in heterogeneous catalysis for the Knoevenagel condensation and cyanation reactions.

EXPERIMENTAL SECTION

Materials and Methods. Reagents of analytical grade were procured from Sigma-Aldrich, Alfa-Aesar, and Spectrochem and were used without further purification. The solvents were dried and/or purified using standard literature methods.⁵⁵ The metalloligands $[\text{Co}(\text{L}^{3,5\text{-COOH}})_3]$ (1) (where $\text{HL}^{3,5\text{-COOH}}$ = 5-(picolinamido)-isophthalic acid) and $[\text{Co}(\text{L}^{4\text{-COOH}})_3]$ (2) (where $\text{HL}^{4\text{-COOH}}$ = 4-[(pyridine-2-carbonyl)-amino]benzoic acid) were synthesized as per our recent reports.^{42,56}

Synthesis of HCNs. $[(1)\text{Zn}_3(\text{H}_2\text{O})_5\cdot 9\text{H}_2\text{O}\cdot \text{C}_3\text{H}_8\text{O}]_n$ (3). Compound 3 was synthesized by treating a solution of $\text{Zn}(\text{OAc})_2\cdot 2\text{H}_2\text{O}$ (0.53 g, 0.24 mmol) in CH_3OH (2 mL) with a solution of

Received: May 20, 2015

Revised: July 4, 2015

Published: July 9, 2015

Table 1. Crystallographic Data Collection and Structure Refinement Parameters for Networks 3–6

	Network 3	Network 4	Network 5	Network 6
Molecular formula	C ₄₅ H ₅₇ CoN ₆ O ₃₀ Zn ₃	C ₈₄ H ₉₀ Co ₂ N ₁₂ O ₅₂ Cd ₃	C ₇₈ H ₇₄ Co ₂ N ₁₂ Na ₂ O ₃₁ Zn ₂	C ₃₉ H ₅₈ CoN ₆ NaO ₂₆ Cd
Fw	1388.78	2510.42	1943.86	1186.97
T (K)	293(2)	293(2)	293(2)	293(2)
Crystal system	Monoclinic	Monoclinic	Monoclinic	Monoclinic
Space group	P2 ₁ /c	P2 ₁ /c	C2/c	C2/c
a	15.0290(4)	13.562(5)	17.317(5)	21.277(5)
b	19.2206(6)	17.989(5)	25.397(5)	24.814(5)
c	21.3862(11)	19.981(5)	21.540(5)	20.774(5)
α	90	90	90	90
β	118.363(3)	93.910(5S)	94.431(5)	113.662(5)
γ	90	90	90	90
V (Å ³)	5436.1(4)	4863(3)	9445(4)	10046(4)
Z	4	2	4	8
d (g cm ⁻³)	1.697	1.714	1.367	1.570
F (000)	2792	2496	3936	4752
Goodness of fit (<i>F</i> ²)	1.067	1.084	1.019	1.036
R ₁ , wR ₂ [<i>I</i> > 2 (<i>I</i>)]	0.0417, 0.1117	0.0614, 0.1627	0.0797, 0.2294	0.0515, 0.1506
R ₁ , wR ₂ [all data] ^a	0.0462, 0.1143	0.0690, 0.1676	0.0890, 0.2389	0.0608, 0.1576
Largest diff. peak and hole (e Å ⁻³)	0.82 and -0.85	1.06 and -2.08	1.38 and -0.94	1.34 and -0.60

$$^a R_1 = \|F_o\| - \|F_c\| / \|F_o\|; wR_2 = \{\sum [w(\|F_o\|^2 - \|Fc\|^2)^2] / \sum wF_o^4\}^{1/2}$$

metalloligand **1** (0.1 g, 0.11 mmol) in CH₃OH (5 mL). The reaction mixture was stirred at room temperature for 2 h. Subsequently, a yellow colored precipitate was collected, washed several times with methanol, and dried under vacuum at ambient temperature. The crude product thus obtained was redissolved in water and layered with isopropyl alcohol at room temperature. After a period of three to four days, red colored crystals of network **3**, suitable for diffraction studies, were obtained. Yield: 0.13 g (85%). Anal. Calcd for C₄₂H₃₁CoN₆O₂₀Zn₃·10H₂O·C₃H₈O: C, 40.19; H, 3.67; N, 6.25; Co, 4.38; Zn, 14.58%; Found: C, 40.44; H, 4.09; N, 6.38; Co, 4.44; Zn, 15.02%. FTIR spectrum (attenuated total reflectance (ATR), selected peaks): 3423 (OH), 1605 (COO), 1585 (C=O) cm⁻¹. Absorption spectrum (solid state; $\lambda_{\text{max}}/\text{nm}$): 520, 350.

{(1)Cd₄(H₂O)₄·18H₂O}_n (**4**). Compound **4** was synthesized in a similar manner with an identical scale as that of network **3**; however, using Cd(OAc)₂·2H₂O instead of Zn(OAc)₂·2H₂O. Red-colored crystals of network **4** were obtained after a period of 5–6 d. Yield: 0.12 g (86%). Anal. Calcd for C₈₄H₅₄Co₂N₁₂O₃₄Cd₄·18H₂O: C, 37.83; H, 3.40; N, 6.30; Co, 4.42; Cd, 16.86%; Found: C, 37.45; H, 3.42; N, 6.04; Co, 4.66; Cd, 16.55%. FTIR spectrum (ATR, selected peaks): 3422 (OH), 1561 (COO, C=O) cm⁻¹. Absorption spectrum (solid state; $\lambda_{\text{max}}/\text{nm}$): 522, 345.

{(2)Zn₂}·{Na₂(H₂O)₁₃}·7H₂O_n (**5**). Compound **5** was synthesized by treating a solution of Zn(OAc)₂·2H₂O (0.62 g, 0.28 mmol) dissolved in CH₃OH (2 mL) with a solution of metalloligand **2** (0.1 g, 0.13 mmol) in CH₃OH (5 mL). The reaction mixture was stirred at room temperature for 1 h. A yellow-colored product was collected, washed several times with methanol, and dried under vacuum at ambient temperature. The crude product thus obtained was redissolved in water and layered with *tert*-butyl alcohol at room temperature. After a period of three to four days, red-colored crystals of network **5**, suitable for diffraction studies, were obtained. Yield: 0.095 g (83%). Anal. Calcd for C₇₈H₇₄Co₂N₁₂O₃₁Zn₂Na₂·7H₂O: C, 44.69; H, 4.23; N, 8.02; Co, 5.62; Zn, 6.24%; Found: C, 44.78; H, 4.53; N, 8.61; Co, 5.87; Zn, 6.01%. FTIR spectrum (ATR, selected peaks): 3423 (OH), 1605 (COO), 1585 (C=O) cm⁻¹. Absorption spectrum (solid state; $\lambda_{\text{max}}/\text{nm}$): 540, 410.

{(2)Cd(H₂O)}·{Na(H₂O)₅}·11H₂O_n (**6**). Compound **6** was prepared in a similar fashion with an identical scale as that of network **5**; however, Cd(OAc)₂·2H₂O was used in place of Zn(OAc)₂·2H₂O. Reddish-brown-colored crystals of network **6** were obtained after a period of 4–5 d. Yield: 0.13 g (87%). Anal. Calcd for C₃₉H₃₆CoN₆O₁₅CdNa·11H₂O: C, 38.36; H, 4.79; N, 6.88; Co, 4.83;

Cd, 9.20%; Found: C, 38.06; H, 5.14; N, 6.74; Co, 5.07; Cd, 9.46%. FTIR spectrum (ATR, selected peaks): 3422 (OH), 1561 (COO, C=O) cm⁻¹. Absorption spectrum (solid state; $\lambda_{\text{max}}/\text{nm}$): 520, 350.

General Procedure for the Knoevenagel Condensation. The respective aldehyde and malononitrile were taken in 5 mL ethanol and the solid catalyst (1 mol %, **3**, **4**, **5** or **6**; based on the repeating unit of the network) was added and the reaction mixture was stirred at 60 °C for 30 min. The progress of the reaction was monitored by TLC using 5% EtOAc/hexanes. After completion of the reaction, solvent was removed under reduced pressure and the residue was triturated with ethyl acetate. The catalyst thus separated was collected by filtration, and the filtrate was concentrated to afford the crude organic product. This product was purified by the column chromatography on 100–200 mesh silica using 1:1 EtOAc/hexanes as the eluent. The products were identified, quantified, and analyzed by GC and GC–MS techniques as well as ¹H NMR and ¹³C NMR spectra whenever required.

General Procedure for the Cyanation Reaction. In a typical cyanation reaction, 2 mol % catalyst (**3**, **4**, **5**, or **6**) was added to a mixture of aldehyde (1 mmol) and (CH₃)₃SiCN (5 mmol). The resulting mixture was stirred at room temperature and the progress of the reaction was monitored by TLC. After completion of the reaction, 2 mL EtOAc was added to the reaction mixture and the solid catalyst was filtered off. The filtrate was concentrated under the reduced pressure. The crude product was purified by flash column chromatography on silica gel using 10% ethyl acetate/hexanes mixture as the eluent. The products were isolated, quantified, and analyzed by the GC and GC–MS techniques and ¹H NMR and ¹³C NMR spectral techniques.

Physical Methods. The FTIR spectra were recorded with the PerkinElmer Spectrum-Two spectrometer with Zn–Se ATR. The NMR spectral measurements were carried out with a Jeol 400 MHz spectrometer with TMS as the internal standard. The absorption spectra were recorded either with PerkinElmer Lambda 25 or Lambda 35 spectrophotometers. The elemental analysis data were obtained with an Elementar Analysen Systeme GmbH Vario EL-III instrument. The metal contents (Co, Zn, and/or Cd) were measured with the help of GBC Avanta atomic absorption spectrometer using standard calibration plots. Gas chromatographic analyses were done with the PerkinElmer Clarus-580 instrument having Elite-5 column. GC–MS studies were performed with a Shimadzu instrument (QP 2010) with an RTX-SSIL-MS column. Thermal gravimetric analysis (TGA) was carried out with DTG 60 Shimadzu at 5 °C min⁻¹ heating rate under the nitrogen atmosphere. The X-ray powder diffraction (XRPD)

studies were performed either with an X'Pert Pro from Panalytical or a Bruker AXS D8 Discover instrument ($\text{Cu K}\alpha$ radiation, $\lambda = 1.54184 \text{\AA}$). The samples were ground and subjected to the range of $\theta = 5\text{--}50^\circ$ with a scan rate of $1\text{--}0.72^\circ/\text{min}$ at room temperature.

Single Crystal Structure Determination. Single crystals suitable for the X-ray diffraction studies were grown by the layering of alcohol (isopropyl alcohol or *tert*-butyl alcohol) into an aqueous solution of the respective network. The intensity data were collected at 293 K with an Oxford XCalibur CCD diffractometer equipped with graphite monochromatic $\text{Mo K}\alpha$ radiation ($\lambda = 0.71073 \text{\AA}$).⁵⁷ An empirical absorption correction was applied using spherical harmonics implemented in the SCALE3 ABSPACK scaling algorithm.⁵⁷ The structures were solved by the direct methods using SIR-92⁵⁸ and refined by the full-matrix least-squares refinement techniques on F^2 using SHELXL-97⁵⁹ in the WinGX module.⁶⁰ The hydrogen atoms were fixed at the calculated positions with isotropic thermal parameters, whereas all non-hydrogen atoms were refined anisotropically. In all cases, hydrogen atoms of the coordinated and lattice water molecules could not be located from the difference Fourier map; however, their numbers have been included in the empirical formulas. In the case of HCN 5, data were corrected for the disordered water molecules by using the SQUEEZE procedure implemented in PLATON⁶¹ which considerably improved the structure convergence. For HCN 6, oxygen atoms O10w, O13w, O14w, and O16w of lattice water molecules were refined isotropically. Details of the crystallographic data collection and structural solution parameters are provided in Table 1.

■ RESULTS AND DISCUSSION

Synthesis and Characterization. The metalloligands $[\text{Co}(\text{L}^{3,5-\text{COOH}})_3]$ (1) and $[\text{Co}(\text{L}^{4-\text{COOH}})_3]$ (2) contain

Scheme 1. Synthesis of Heterometallic Coordination Networks 3–6 Using Metalloligands 1 and 2

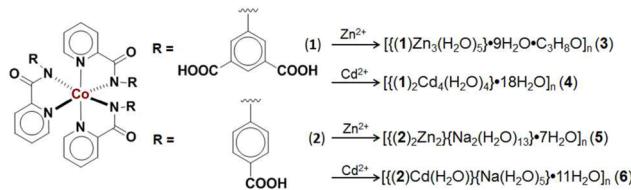


Table 2. Selected Bond Distances and Ranges for Networks 3–6

bond distance	3	4	5	6
Co(1)–N(1)	1.933(3)	1.923(4)	1.933(5)	1.923(3)
Co(1)–N(2)	1.929(3)	1.945(4)	1.958(5)	1.940(4)
Co(1)–N(3)	1.943(3)	1.954(4)	1.953(5)	1.944(3)
Co(1)–N(4)	1.953(3)	1.958(4)	1.957(4)	1.956(4)
Co(1)–N(5)	1.944(3)	1.943(4)	1.922(5)	1.964(4)
Co(1)–N(6)	1.936(3)	1.972(4)	1.930(5)	1.953(4)
M ^a –O _{carboxylate}	1.941(3)– 2.269(2)	2.173–2.573	1.951(5)– 2.428(9)	2.262(6)– 2.617(5)
M ^a –O _{water}	2.017(3)– 2.180(3)	2.220–2.450	---	2.266(5)

^aWhere M = Zn²⁺ (for networks 3 and 5) and Cd²⁺ (for networks 4 and 6).

appended but uncoordinated arylcarboxylic acid groups. The metalloligands 1 and 2, respectively, offer six and three

appended arylcarboxylic acid groups at the strategically placed positions. Such metalloligands upon reaction with secondary Zn(II) and Cd(II) ions provide the desired HCNs 3–6 (Scheme 1). In each case, reaction with secondary metal ion resulted in a distinct color change followed by the precipitation of the product. The crude product was recrystallized from water assisted by layering either with isopropyl alcohol or *tert*-butyl alcohol. All four HCNs exhibit strong bands between 1605 and 1560 cm⁻¹ for the $\nu_{\text{carboxylate}}$ stretches and broad features between 3400 and 3430 cm⁻¹ due to the presence of coordinated as well as lattice water molecules (Figures S1 and S2, Supporting Information, SI).

HCNs 3–6 were also characterized by their diffuse-reflectance absorption spectra (Figures S3 and S4, SI). HCNs 3 and 4 display nearly identical spectra with absorption bands at ca. 520 and ca. 350 nm. On the other hand, HCNs 5 and 6 exhibit spectral features at 540 and 410 nm and 520 and 350 nm, respectively. The low-energy bands are tentatively assigned to the d–d transition based on the Co(III) core of the metalloligand, whereas the high-energy features are most probably charge-transfer in nature.

The XRPD patterns of the freshly prepared samples of HCNs 3–6 are in good agreement with those simulated from the single-crystal structural data. Such a fact indicates phase purity of the bulk crystalline sample (Figures S5–S8, SI).

All four HCNs display a very similar thermal decomposition profile (Figures S9 and S10, SI). The first weight change, observed between 25 and 150 °C, corresponds to the loss of lattice as well as coordinated water molecules. This weight loss is followed by the decomposition of the networks that, respectively, starts at ca. 300 °C for HCNs 3 and 4 and at ca. 350 °C for HCNs 5 and 6.

Crystal Structures. All four HCNs were characterized by the single crystal X-ray diffraction studies to understand the network generation via the mediation of arylcarboxylate-appended metalloligands. For all four networks, Table 2, and Tables S1 and S2 (Supporting Information), respectively, provide the selected bond distances and detailed bonding parameters.

In all four HCNs, the central Co³⁺ ion is coordinated by three bidentate ligands. Importantly, bidentate ligands are arranged meridionally around the Co³⁺ ion. Such a *meridional* geometry around the Co³⁺ ion, less symmetrical than that of conventional *facial*, results in unique orientation of the appended arylcarboxylate groups. Such an orientation of arylcarboxylate groups generates interesting networks after coordination with the secondary metal ions. The bidentate ligands ligate the Co³⁺ ion via five-membered chelate rings between N_{amide} and N_{pyridyl} atoms. The Co–N_{amide} bond distances (avg. 1.941–1.949 Å) are comparable to that of Co–N_{pyridine} ones (avg. 1.933–1.940 Å). All diagonal angles are close to linear; ranging from 173.1° to 176.7°, again suggesting a symmetrical arrangement of donors around the Co³⁺ ion. As a result of symmetrical bonding, the geometry around the Co(III) ion is nearly octahedral. Notably, in all four networks; the role of Co³⁺-based metalloligand is limited to structural anchoring and directing the appended arylcarboxylate groups to limited orientations. While HCNs 3 and 4 originate from metalloligand 1 offering two carboxylate groups at 3- and 5-positions of the arene ring, networks 5 and 6 are constituted of metalloligand 2 with only one carboxylate group at the 4-position of the arene ring.

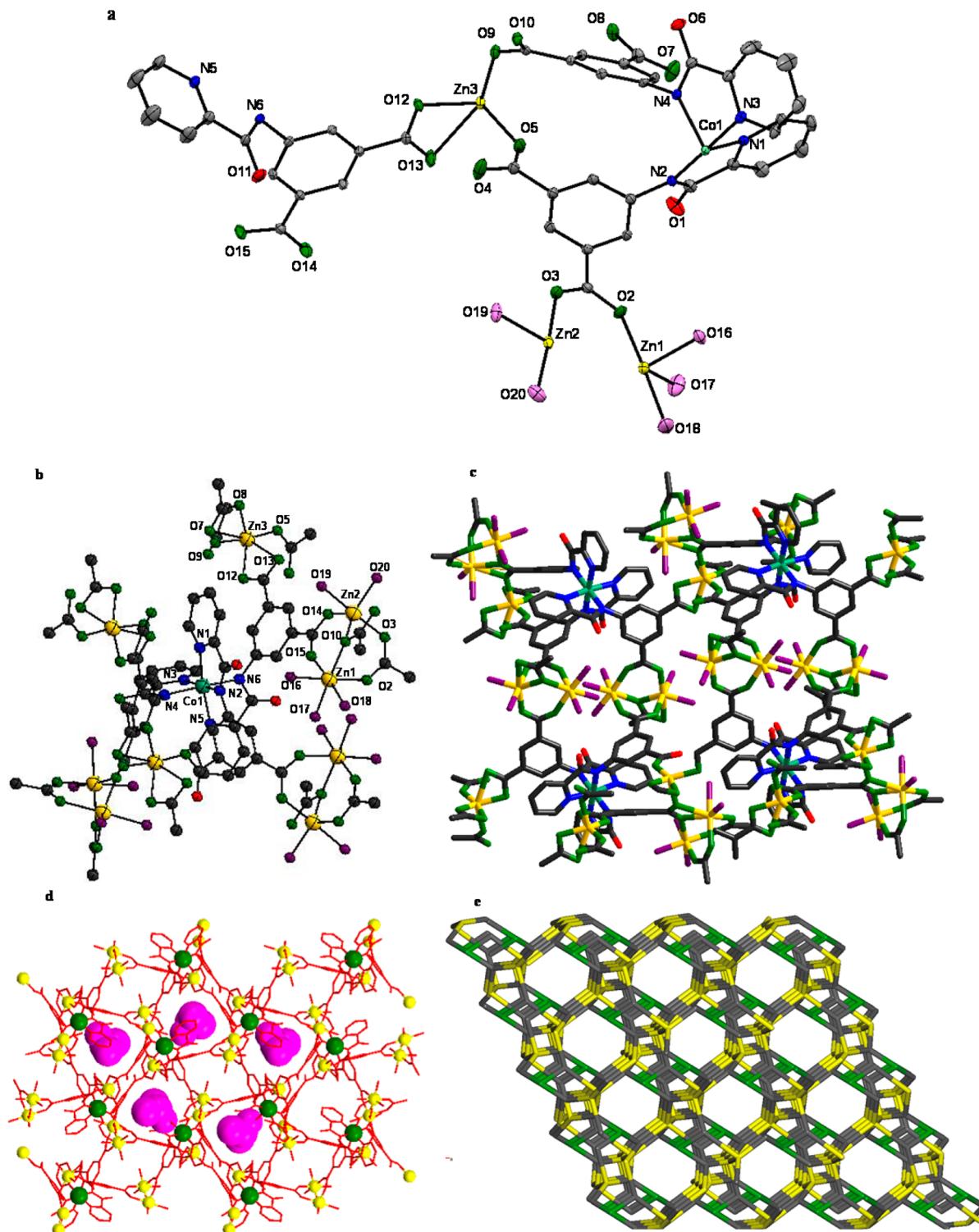


Figure 1. (a) Asymmetric unit of network 3; thermal ellipsoids are drawn at 30% probability level, whereas hydrogen atoms, isopropyl alcohol, and water molecules are omitted for clarity. (b) Ball and stick representation of a selected part of the crystal structure of network 3. Color code: sea green, Co; gold, Zn; blue, N; green, O_{carboxylate}; violet, O_{water}; red, O_{amide}; gray, C. (c) Extended structure of network 3. (d) Extended structure of network 3 along the *a* axis showing the presence of isopropyl alcohol (space-filling mode; purple color) within the network lattice. (e) Topological representation of network 3. Color code: green, Co; yellow, Zn; gray, contribution from the ligand.

Both HCNs 3 and 4 crystallized in the monoclinic cell with $P2_1/c$ space group. The unit cell of 3 contains one metalloligand, three Zn^{2+} ions, five coordinated water molecules, one lattice isopropyl alcohol, and nine lattice water molecules. The structure of network 3 shows that the

Co^{3+} -based metalloligands are connected through the secondary Zn^{2+} ions (Figure 1). Every metalloligand offers three appended arylcarboxylate groups, therefore a total of six carboxylate groups. Such arylcarboxylate groups coordinate separate $Zn(II)$ ions leading to a 3D architecture. The six-

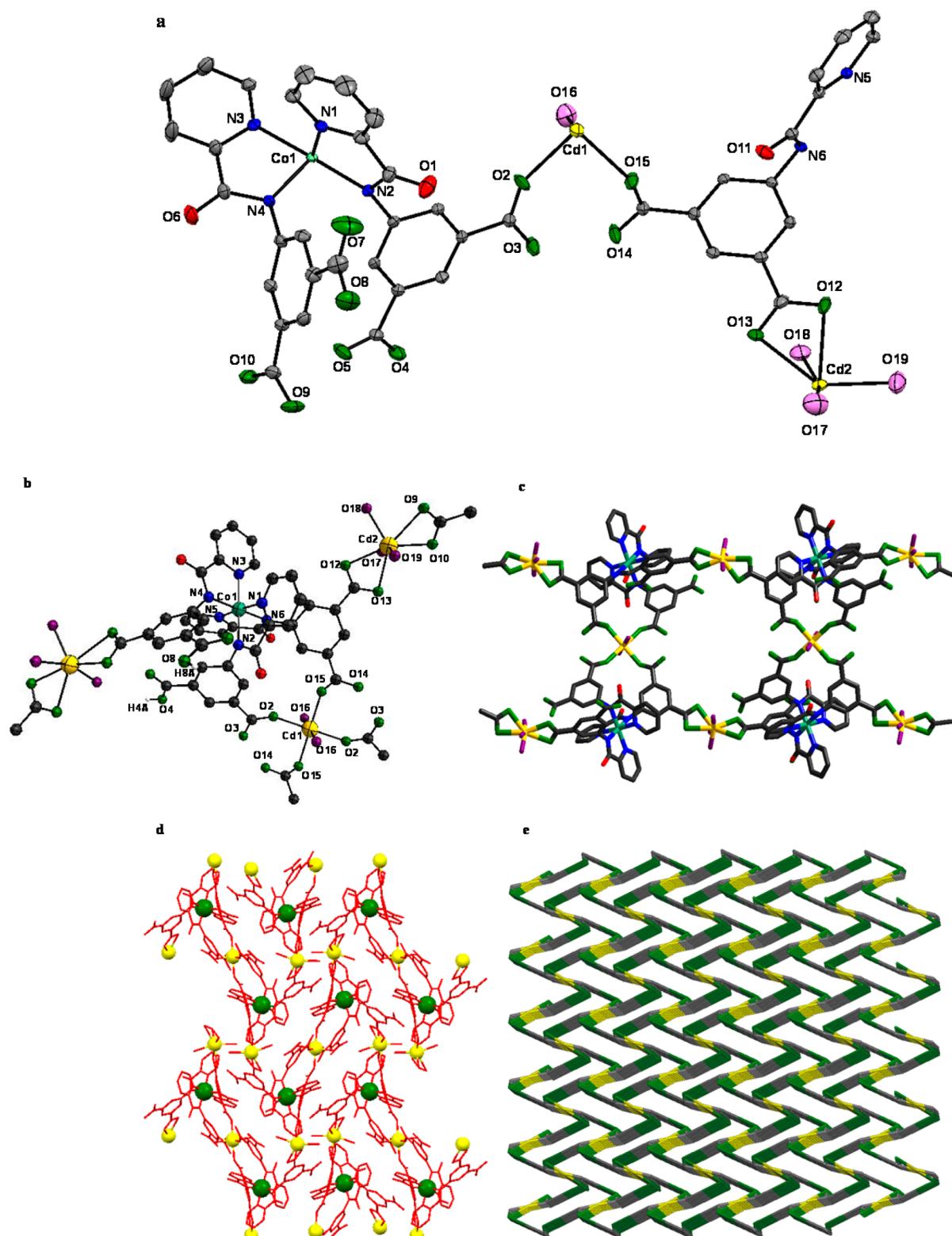


Figure 2. (a) Asymmetric unit of network 4; thermal ellipsoids are drawn at 30% probability level, whereas hydrogen atoms and water molecules are omitted for clarity. (b) Ball and stick representation of a selected part of the crystal structure of network 4. Color code: sea green, Co; gold, Cd; blue, N; green, O_{carboxylate}; violet, O_{water}; red, O_{amide}; gray, C. (c) Extended structure of network 4. (d) Extended structure of network 4 in a view along the a axis. (e) Topological representation of network 4. Color code: green, Co; yellow, Cd; gray, contribution from the ligand.

coordinated Zn1 ion is ligated to three O_{carboxylate} atoms from three different metalligands and three water molecules. The five-coordinated Zn2 ion ligates three O_{carboxylate} atoms from

three different metalligands and two water molecules. The five-coordinated Zn3 ion is connected to five O_{carboxylate} atoms in which four donors are from two metalligands while the

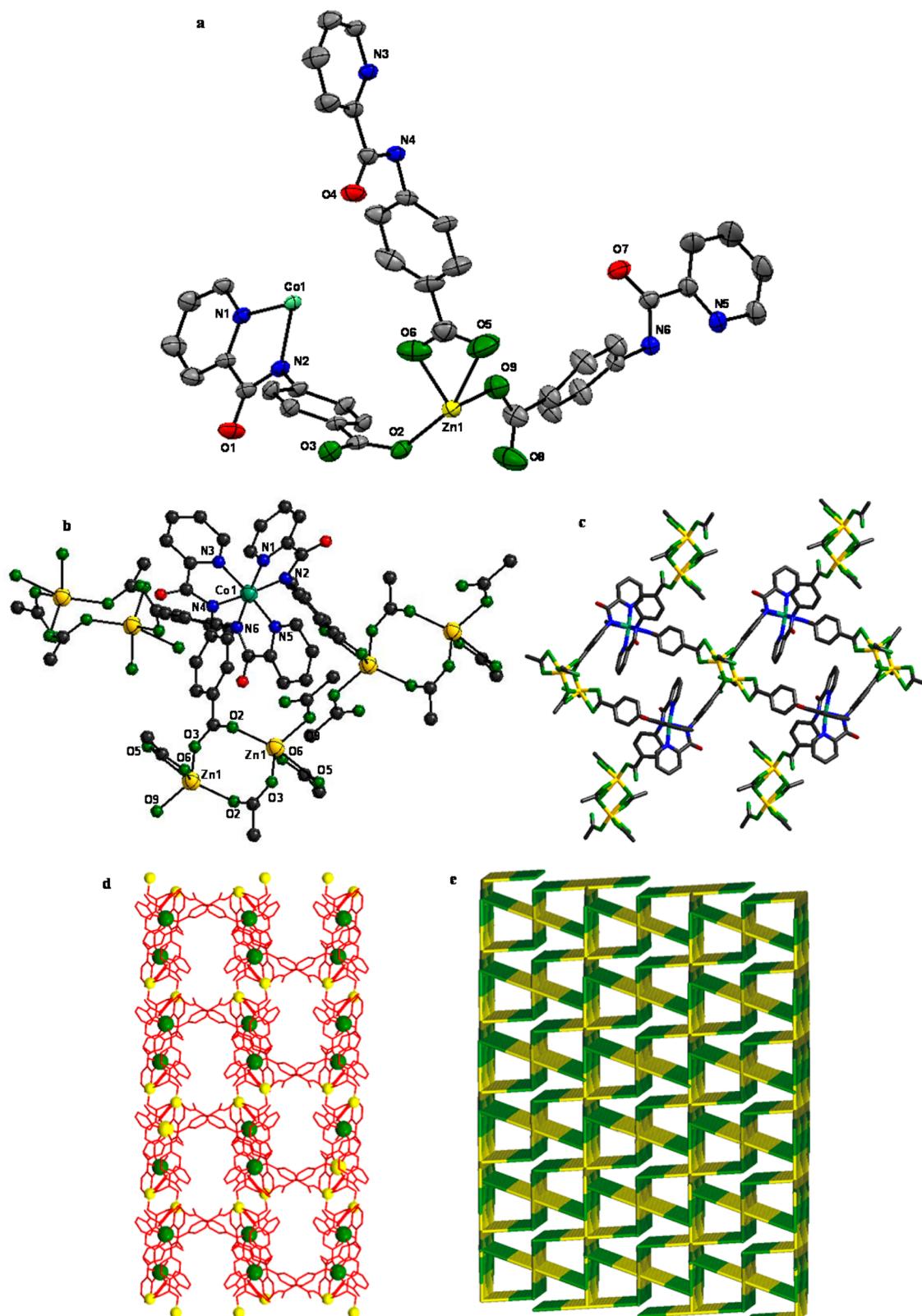


Figure 3. (a) Asymmetric unit of network 5: thermal ellipsoids are drawn at 30% probability level, whereas hydrogen atoms, Na^+ ion, and water molecules are omitted for clarity. (b) Ball and stick representation of a selected part of the crystal structure of network 5. Color code: sea green, Co; gold, Zn; blue, N; green, O_{carboxylate}; violet, O_{water}; red, O_{amide}; gray, C. (c) Extended structure of network 5. (d) Extended structure of network 5 in a view along the *a* axis. (e) Topological representation of network 5. Color code: green, Co; yellow, Zn.

fifth one is from a separate metalloligand. Notably, guest isopropyl alcohol is trapped within the crystal lattice with the help of hydrogen bonds (H-bonds) involving Zn-bound water

molecules as well as lattice water molecules. Consequently, isopropyl alcohol occupies the channels created in the packing of HCN 3.

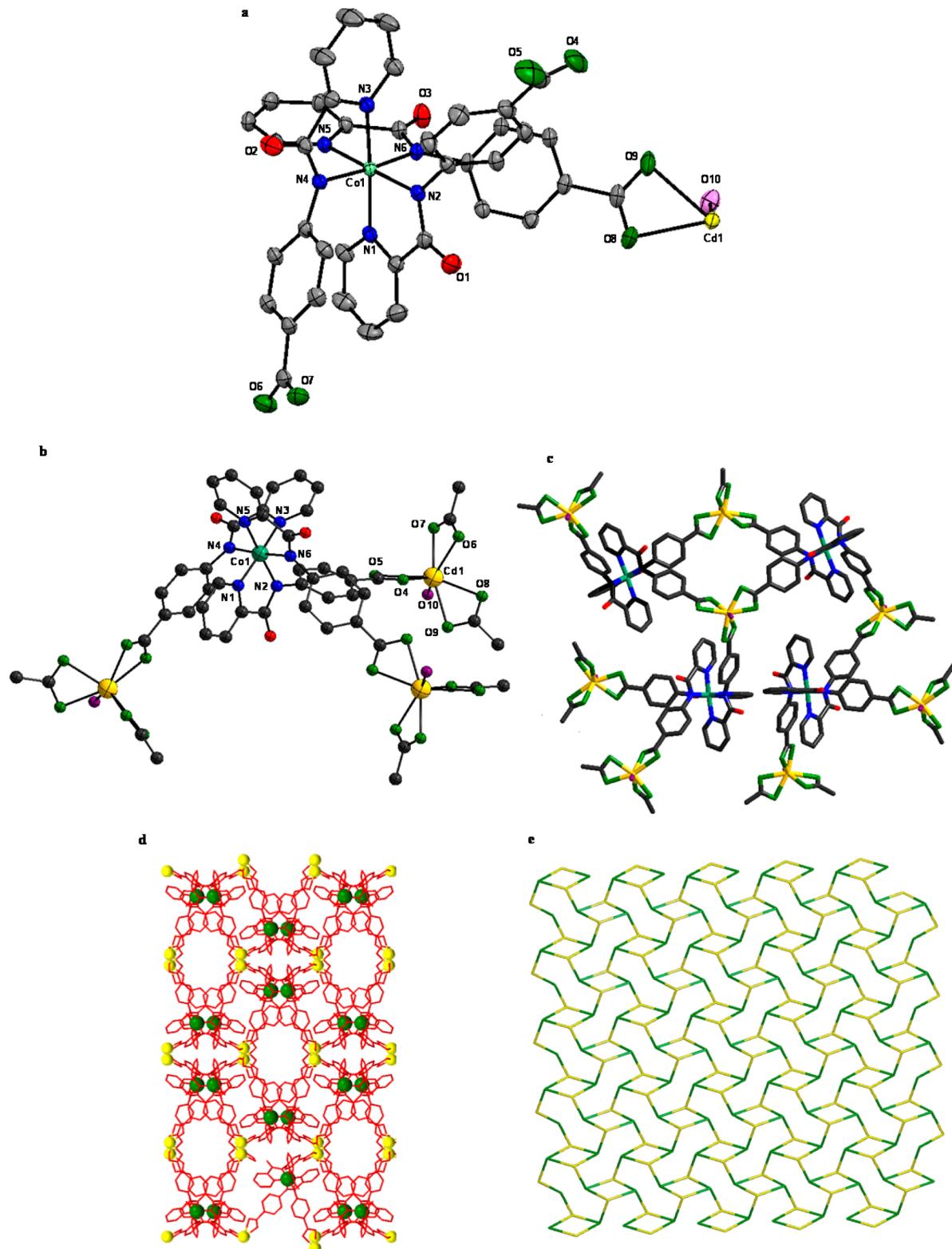
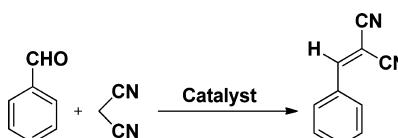


Figure 4. (a) Asymmetric unit of network 6: thermal ellipsoids are drawn at 30% probability level, whereas hydrogen atoms, Na^+ ion, and water molecules are omitted for clarity. (b) Ball and stick representation of a selected part of the crystal structure of network 6. Color code: sea green, Co; gold, Cd; blue, N; green, $\text{O}_{\text{carboxylate}}$; violet, O_{water} ; red, O_{amide} ; gray, C. (c) Packing diagram of network 6. (d) Extended structure of network 6 in a view along the *c* axis. (e) Topological representation of network 6. Color code: green, Co; yellow, Cd.

The unit cell of network 4 contains 2 metalloligands, 4 Cd^{2+} ions, and 4 coordinated and 18 uncoordinated water molecules. The molecular structure of 4 exhibits a 3D network where Co^{3+} -based metalloligands are connected through the secon-

dary Cd^{2+} ions (Figure 2). As noted for 3, every metalloligand offers six carboxylate groups. Interestingly, four appended groups out of six coordinate to two Cd^{2+} ions in the carboxylate form whereas the remaining two groups exist as carboxylic acid

Table 3. Control Experiments for the Knoevenagel Condensation of Benzaldehyde with Malononitrile Using Network 3



entry	catalyst [mol %]	solvent	time [min]	temp. [°C]	yield ^a [%]
1	-	Ethanol	60	25	0
2	3 [1]	Ethanol	30	25	70
3	3 [1]	Ethanol	30	40, 50, 60	80, 100, 100
4	3 [1]	Ethanol	15, 30	50	85, 100
5	3 [0.5, 1, 2]	Ethanol	30	50	75, 100, 100
6	3 [1]	MeOH, CH ₃ CN, EtOAc, Toluene	30	50	85, 82, 78, 80

^aYield was calculated using the gas chromatograph.

and take part in intermolecular hydrogen bonding (H-bonding) with the lattice water molecules. The six-coordinated Cd1 ion is ligated to four O_{carboxylate} atoms, two each from one metalloligand, whereas the rest of the axial sites are occupied by the water molecules. The seven-coordinated Cd2 ion is ligated to four O_{carboxylate} atoms, two each from every metalloligand, and three water molecules.

Both HCNs **5** and **6**, originated from metalloligand **2** appended with *para*-arylcarboxylate groups, crystallize in the monoclinic cell with C2/c space group. The higher symmetry associated with networks **5** and **6**, when compared to **3** and **4**, immediately suggests the difference of placing fewer arylcarboxylate groups on a metalloligand in the former. The unit cell of HCN **5** contains 2 metalloligands, 2 Zn²⁺ ions, 2 Na⁺ ions, and 13 coordinated water molecules. Any metal-

ligand offers three *para*-arylcarboxylate groups and two such groups out of three coordinate one Zn²⁺ ion, whereas the remaining one bridges two Zn²⁺ ions. A repetition of such a bonding pattern generates a 3D network (Figure 3). The geometry around the five-coordinated Zn(II) ion is severely distorted with τ_5 parameter of ca. 0.6.⁶² It may be noted that τ_5 has ideal values of 0 and 1 for perfect square pyramidal and trigonal bipyramidal geometries, respectively.⁶² The structure displays the presence of a Na₂(OH₂)₁₃ cluster in which two sodium ions are bridged by one water molecule while coordinated to the rest of terminal water molecules. Such a cluster is held in place by an array of H-bonds involving O_{amide} and lattice water molecules.

The unit cell of network **6** contains 1 metalloligand, 1 Cd²⁺ ion, 1 Na⁺ ion, and 6 coordinated and 11 lattice water molecules. The molecular structure of **6** shows that the metalloligands are connected through the secondary Cd²⁺ ions resulting in a 3D network (Figure 4). Every seven-coordinated Cd(II) ion receives three pairs of bidentate arylcarboxylate coordinations from three different metalloligands while the seventh coordination site is occupied by a water molecule. Interestingly, crystal structure also shows the presence of a Na(OH₂)₅ cluster which is held in place by a series of H-bonds including O_{amide} and lattice water molecules.

Solvent-accessible voids (SAVs) of different volumes were found in all crystal structures. Networks **3** and **4** displayed SAVs of volume 1575.1 Å³ and 551.3 Å³ which correspond to 29% (5436.2 Å³) and 11.3% (4863.0 Å³) of their unit cell volumes, respectively. On the other hand, SAVs of 1819.4 Å³ and 2543.3 Å³ corresponding to 19.3% (9445 Å³) and 25.3% (10046.3 Å³) of unit cell volume were noted for networks **5** and **6**, respectively.

Topological Analyses. The topology for all four HCNs was studied with the help of TOPOS 4.0.^{63,64} Although all four networks are similar in terms of involving Co³⁺ based

Table 4. Knoevenagel Condensation of Assorted Aldehydes Using Networks 3–6

Entry ^a	R	Product	Yield [%] ^b			
			Network 3	Network 4	Network 5	Network 6
1	C ₆ H ₅		>99	>99	>99	>99
2	4-NO ₂ C ₆ H ₄		>99	>99	>99	>99
3	4-ClC ₆ H ₄		>99	>99	>99	>99
4	3-MeOC ₆ H ₄		99	>99	>99	99
5	Naphthalen-2-yl		99	>99	>99	99
6	Furan-2-yl		99	>99	>99	99
7	Thiophen-2-yl		99	>99	>99	99

^aConditions: Catalyst: 1 mol %; Temperature: 50 °C; Time: 30 min. ^bYield was calculated using the gas chromatograph.

Table 5. Cyanation Reactions of Assorted Aldehydes with Trimethylsilylcyanide Using HCNs 3–6

Entry ^a	Substrate	Molecular drawing of the substrate	Molecular dimension of substrate (l × w × h)	Yield [%] ^b			
				3	4	5	6
1			4.2 × 6.03 × 1.4	90	92	98	>99
2			4.22 × 6.98 × 1.4	95	96	99	>99
3			4.21 × 6.68 × 2.0	78	85	99	99
4			4.20 × 8.47 × 2.06	80	87	94	94
5			4.85 × 7.77 × 2.10	78	80	90	92
6			6.03 × 6.64 × 1.4	40	50	80	88
7			5.66 × 7.21 × 1.4	36	45	78	85
8			6.22 × 9.28 × 1.4	38	45	40	55

^aConditions: Catalyst: 1 mol %; Temperature: 50 °C; Time: 30 min. ^bYield was calculated using the gas chromatograph.

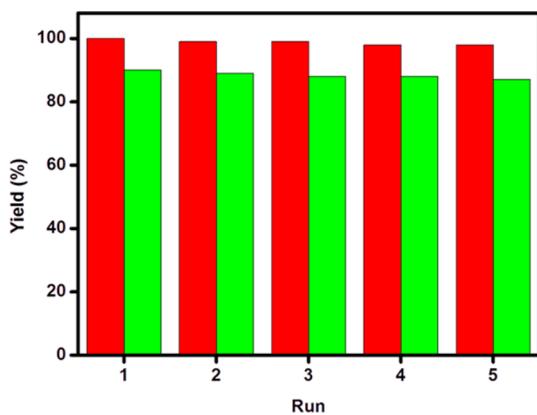


Figure 5. Recycling of network 3 up to five cycles in the Knoevenagel condensation reactions (red bars) and the cyanation reactions (green bars).

metalloligands and secondary metal ions, their topologies are considerably distinct from each other. It is noteworthy that each network offers secondary building units (SBUs), constituted of Zn²⁺ or Cd²⁺ ions that are coordinated to a

certain number of arylcarboxylate moieties.^{54,65–67} These SBUs result in a number of nodes leading to different types of topologies for each coordination network.^{65–72} In each case, such SBUs expand through Co³⁺ metalloligand units which themselves act as the nodes.⁵⁴

Both HCNs 5 and 6 contain similar 3-connected Co³⁺ metalloligand nodes with point symbols {4².6} and {4.8²}, respectively.^{65–67} Further, the arylcarboxylate groups are connected to the secondary Zn²⁺ or Cd²⁺ ions to generate Zn₂(COO)₆ (in 5) or Cd(COO)₃ (in 6) SBUs that serve as the 6-connected and 3-connected cluster nodes, respectively. The TOPOS analysis provides flu type topology with point symbol {4².6}₂{4⁴.6².8⁷.10²} and fes type topology with point symbol 4.8² for 5 and 6, respectively (Figures 3e and 4e).

In HCNs 3 and 4, every HL_{3,5-COOH} ligand offers two carboxylate groups, resulting in a total of 6 arylcarboxylate moieties that can coordinate to the secondary Zn²⁺ or Cd²⁺ metal ions to form different types of SBUs.^{65–67} Besides this, both HCNs offer Co³⁺ metalloligand based nodes which are 3-connected for both HCN 3 and 4 with point symbols {5.8²} and {3.10²}, respectively. Interestingly, in the case of HCN 3, a new type of topology has been observed where arylcarboxylate groups bind with Zn²⁺ ions to give rise to Zn(COO)₄ and

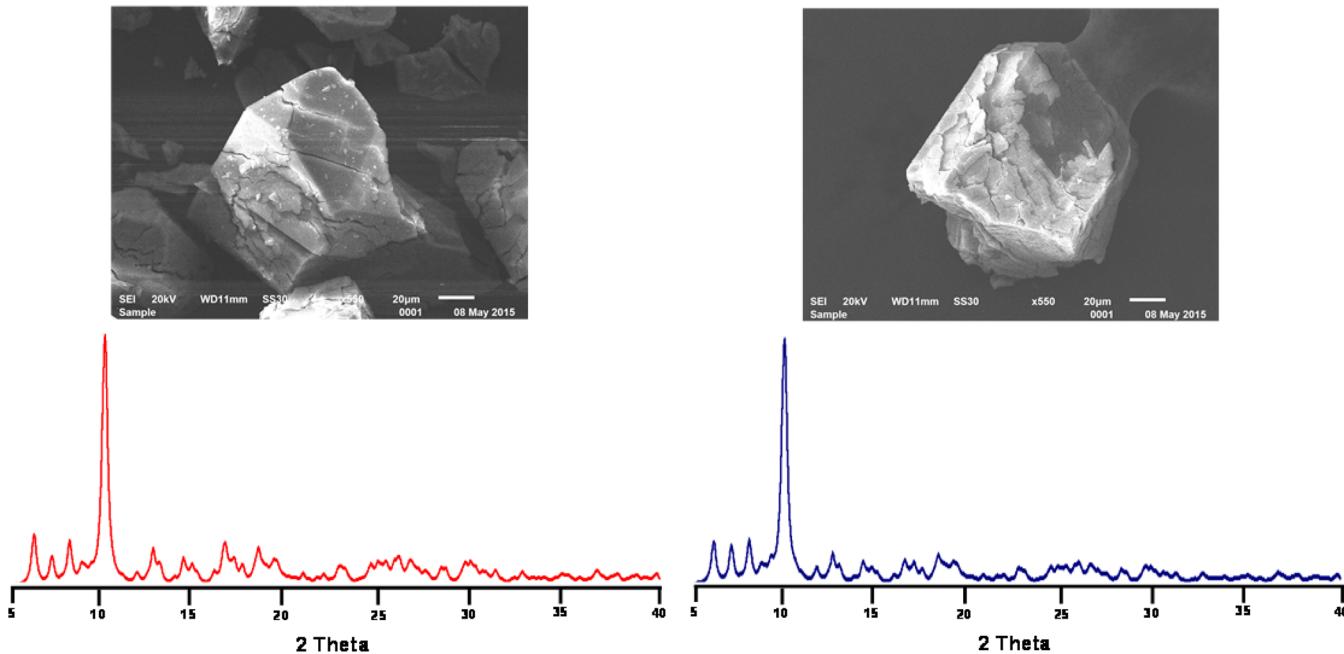


Figure 6. PXRD patterns along with the SEM images of network 3: before (left panel) and after the Knoevenagel condensation reaction of benzaldehyde with malononitrile (right).

$\text{Zn}_3(\text{COO})_3$ SBUs with 4-connected nodes. Such SBUs are connected through the Co^{3+} based metalloligand nodes generating a network with the point symbol $\{5.6.7.8^2.9\}\{5.6.7\}\{5.8^2\}_2\{5.8^3.9^2\}\{8^3\}$ (Figure 1e). In the case of 4, aryldicarboxylate groups give rise to $\text{Cd}(\text{COO})_2$ and $\text{Cd}(\text{COO})_4$ SBUs. In $\text{Cd}(\text{COO})_2$, Cd2 is bonded to two arylcarboxylate groups in chelating mode resulting in a rod-like connection. Since this Cd atom (Cd2) does not act as a distinct node, it will not contribute toward the overall topology of the network. Hence, there are only 3- and 4-connected nodes as a result of Co^{3+} based metalloligands and Cd1 centers, respectively. The TOPOS analysis for HCN 4 generates the point symbol $\{3.10^2\}_4\{3^2.10^2.11^2\}$ (Figure 2e).

Application of HCNs in Heterogeneous Catalysis. An important structural feature of the four HCNs is the utilization of two different types of metalloligands in the construction of networks. Notably, metalloligand 1, offering six arylcarboxylate groups, affords quite densely packed HCNs 3 and 4. On the other hand, metalloligand 2, offering only three arylcarboxylate groups, generates HCNs 5 and 6 with loosely packed architectures. Such distinct architectures suggest that while HCNs 5 and 6 may allow facile diffusion of substrates and reagents due to such a fact; the other two HCNs may only interact with the substrates and/or reagents on the surface. Therefore, the four HCNs provide a good opportunity to compare their role in the catalysis in order to understand the implication of design concept. Further, structural investigations of all four HCNs illustrate the presence of Lewis acidic metal ions, $\text{Zn}(\text{II})$ and $\text{Cd}(\text{II})$, arranged in an orderly manner. Such secondary metal ions are coordinated by the arylcarboxylate groups in addition to a certain number(s) of labile water molecules. Importantly, such labile water molecules could be potentially displaced by the suitable substrates (discussed below). Furthermore, the 3D nature of all HCNs, their stability and robustness, and insolubility in common organic solvents provide an option to carry out the catalytic reactions heterogeneously. We selected Knoevenagel condensation

(KC) and cyanation reaction (CR) of aldehydes to illustrate the importance and uniqueness of the four HCNs.

Exchange Studies. One of the important parameters in heterogeneous catalysis is to validate that the substrates are able to replace loosely coordinated solvent molecule(s) on a Lewis acidic metal ion. We selected HCN 3 to provide such evidence. In order to confirm whether the coordinated and lattice water could be exchanged, 3 was heated under vacuum to remove most of the coordinated and lattice water molecules. Such a dehydrated sample was then saturated in a controlled environment of D_2O vapors. The FTIR spectrum of D_2O -exchanged sample displayed shifted $\nu_{\text{O}-\text{D}}$ stretches at ca. 2500 cm^{-1} suggesting replacement of the coordinated and lattice water molecules (Figure S11, SI). Notably, other spectral features did not show significant change confirming the framework stability. Further, HCN 3 was impregnated with substrate by suspending its crystals in a CH_2Cl_2 solution of benzaldehyde and analyzing with FTIR spectra, powder diffraction, and scanning electron micrograph (SEM) images. The FTIR spectrum exhibited red-shifted $\nu_{\text{C=O}}$ stretches (at 1680 cm^{-1}) for the impregnated $\text{C}_6\text{H}_5\text{CHO}$ when compared to neat sample (at 1700 cm^{-1}) (Figure S12, SI). This fact points out that benzaldehyde has been potentially activated owing to its coordination to the Lewis acidic metal ion. On the other hand, powder XRD pattern displayed broad and moderately shifted features indicative of exchange (Figure S13, left panels; SI). Importantly, SEM images display that the crystals retain their shape and crystallinity albeit with minor scars on the surface (Figure S13, right panels; SI). Collectively, these studies show that the reagents (D_2O and $\text{C}_6\text{H}_5\text{CHO}$) have the potential to diffuse through the porous network and replace the coordinated water molecule(s).

Knoevenagel Condensation. Knoevenagel condensation^{73,74} (KC) is a well-known synthetic method for the formation of C–C bonds needed for creating assorted organic synthons and significant classes of compounds. KCs can be catalyzed by a variety of chemicals and reagents whereas several

Lewis acidic metals have also been successfully utilized.^{75–85} Lewis acidic metals based heterogeneous catalysts stand out due to their potential reusability, stability, and ease of product separation.^{49,51} However, functional group tolerance and substrate scope has been one of the limiting factors with heterogeneous catalysts, and therefore next-generation catalysts are required.^{32–40} In this context, Lewis acidic metal-based coordination networks have displayed noteworthy results by several groups including us.^{49,51,75–85} The catalytic activities of HCNs 3–6 was investigated by employing several aromatic and heterocyclic aldehydes using malononitrile as an active methylene compound. In order to optimize the reaction conditions, HCN 3 was used as a representative catalyst for the KCs of benzaldehyde with malononitrile. As can be seen from Table 3, catalyst 3 is essential for the KC, as no product was observed without it (entry 1). Entries 2 and 3 depict the importance of mild elevated temperature of 50 °C for the quantitative conversion. It is further noted that quantitative product formation requires 30 min (entry 4) and only 1 mol % catalyst loading (entry 5). Entry 6 suggests that EtOH is the best solvent for KCs as other solvents resulted in comparatively poor performance. With these optimized reaction conditions, all four HCNs offered KCs of assorted aldehydic substrates with malononitrile (Table 4). Importantly, only 1 mol % of the four HCNs resulted in nearly quantitative conversion for most of the substrates. Further, substrates carrying either electron-withdrawing or electron-donating groups (entries 2–4, Table 2); naphthal group (entry 5); and heterocyclic rings (entries 6 and 7) were equally effective. These examples adequately illustrate the importance of the present HCNs in carrying out the Lewis acid metal assisted KCs heterogeneously and suggested their potential catalytic roles in other similar reactions.

Cyanation Reactions. CRs of carbonyl-based substrates provide a convenient route to cyanohydrins which are significant intermediates for a variety of fine chemicals and organic synthons.^{86–89} Extensive research has been directed to develop improved Lewis acid based catalysts for the reaction of cyanide ion with carbonyl-based substrates.^{90–93} In this context, heterogeneous catalysts incorporating Lewis acidic sites are advantageous for the ease in catalyst separation and product isolation.^{32–40,49,51} The successful performance of the present HCNs in KCs advocated their potential role in CRs as well. As displayed in Table 5, all four HCNs were able to catalyze CRs of assorted aldehydes using trimethylsilylcyanide as the cyanide source. It is important to mention that CRs reactions were performed without any requirement of solvent taking advantage of the liquid nature of most of the reagents.⁹⁴ Benzaldehyde as well as several functionalized analogues provided the corresponding cyanohydrins in good to excellent yield.

The CRs also provided the option to understand the size exclusion catalysis as several related aldehydes with varying molecular dimensions are available. We, therefore, performed CRs with assorted aldehydic substrates differing in molecular dimensions as mentioned in Table 5. Importantly, benzaldehyde, 4-nitrobenzaldehyde, 4-chlorobenzaldehyde, 4-methoxybenzaldehyde, and 3-methoxybenzaldehyde offered very good to nearly quantitative yield of the respective product with all four HCNs. Such a fact suggests that these substrates have comparable uncluttered access to the catalytic secondary metal centers in all four HCNs. However, 1-naphthaldehyde and 2-naphthaldehyde as well as 9-anthrinaldehyde provided comparatively lower yield of the product especially with HCNs 3 and 4.

We reason that, as the molecular size of the substrate increases, more and more substrates are excluded from approaching the catalytic metal ions present within the interior of the materials. Therefore, the catalytic sites available on the surface actually carry out the CR as evidenced by the lower product yield. A notable difference is clearly observable between the HCNs produced by metalloligand 1 when compared to the ones generated by 2. As already explained, both HCNs 5 and 6 offer much larger pores and therefore fewer bulkier substrates were excluded, whereas densely packed HCNs 3 and 4 excluded maximum numbers of substrates. Such a notable difference in catalysis is suggestive of the architectural distinctness between two categories of HCNs.

Recyclability Experiments. As all four HCNs carried out heterogeneous KCs and CRs quite efficiently, we became interested in testing the recyclability of these networks. Importantly, addition of EtOAc to the reaction mixture resulted in clean separation of HCNs. Such networks were filtered, dried in air, and reused for the subsequent catalysis without any requirement of special treatment. We tested the recovered HCNs for the KCs as well as CRs for five consecutive cycles and did not notice any significant decrease in the catalytic performance (Figure 5). All four recovered networks were characterized by the powder XRD measurements and gratifyingly showed identical reflections to that of as-synthesized samples, thus confirming that the structural integrity remains intact during the catalysis (Figures 6 and S14–S16, SI). The FTIR spectrum of recovered network 3 as a representative example was also in excellent match to that of pristine sample (Figure S17, SI). In order to understand the probable morphological changes to the HCNs during the course of the catalytic reaction, recovered network 3 was evaluated by the SE micrographs (Figure 6). Although SEM studies did not show significant changes to the surface morphology, there was evidence of minor scars and cracks. These studies further establish the reusable nature of the present HCNs in Lewis acid assisted catalytic reactions.

CONCLUSION

We have utilized two Co³⁺-based metalloligands with appended arylcarboxylic acid groups at the strategically placed positions for the construction of {Co³⁺-Zn²⁺} and {Co³⁺-Cd²⁺} heterometallic coordination networks. In the resultant 3D networks, secondary Zn²⁺ and Cd²⁺ ions receive coordinations from the appended arylcarboxylate groups of the metalloligands. These 3D networks exhibit interesting network topologies including an unprecedented one. All networks acted as the heterogeneous and reusable catalysts for the Knoevenagel condensation reactions and cyanation reactions of assorted aldehydes. These examples adequately illustrate metalloligand strategy for the development of promising heterometallic networks.

ASSOCIATED CONTENT

S Supporting Information

Figures for the FTIR spectra, diffuse reflectance absorption spectra, TGA, Powder XRD, and SEM images; and tables for the bond distances and bond angles. The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acs.cgd.5b00697.

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The manuscript was written through contributions of all authors.

Notes

The authors declare no competing financial interest.

ACKNOWLEDGMENTS

R.G. acknowledges the financial support from the Science & Engineering Research Board (SERB), New Delhi. Authors sincerely thank CIF-USIC of the University of Delhi for the instrumental facilities. S.S. acknowledges UGC for the SRF fellowship.

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