Corrections

Hypothetical Thermodynamic Properties. Subcooled Vaporization Enthalpies and Vapor Pressures of Polyaromatic Hydrocarbons. William Hanshaw, Marjorie Nutt, and James S. Chickos,* *J. Chem. Eng. Data* **2008,** *53,* 1903–1913.

Page 1912. The headings in Table 8 should be changed as given below.

Table 8. Summary of the Parameters of the Third-Order Polynomial, Equation 3, Calculated by Correlation for the PAHs of this Study

	$\frac{A}{T^3}$	$\frac{B}{T^2}$	$\frac{C}{T}$	D
biphenylene	147336207.1	−1912159.03	−119.691	5.477
diphenylmethane	115104919.3	−1585090.3	−1317.607	7.085
fluorene	281912267.6	−3039483.96	2358.69	3.348
1-methylfluorene	186870736.8	−2448715.91	737.475	4.527
4H-cyclopenta(*def*)phenanthrene	281948344	−3381686.99	3081.917	1.903
2-phenylindene	249783869.2	−3103982.53	2204.907	3.087
1-methylanthracene	253190269.5	−3124542	1885.73	3.537
2-methylanthracene	196916080.8	−2517940.73	70.957	5.244
9-methylanthracene	256952510.4	−3172745.66	1933.223	3.456
1-methylphenanthrene	181137987.8	−2275711.89	−913.277	6.312
pyrene	272098419.3	−3361297.08	2143.069	3.095
9,10-dimethylanthracene	284338787.7	−3522193.91	2591.339	2.668
2-ethylanthracene	257375467.1	−3214538.04	1815.719	3.679
2,3-benzofluorene	244120457.7	−3097765.94	780.109	4.616
1,2-benzanthracene	282250430.8	−3627106.05	1948.441	3.323
2,3-benzanthracene	290595955.4	−3674856.98	1937.572	3.354
triphenylene	268335765.5	−3471269.68	1311.11	4.051
p-terphenyl-d_{14}	250438718.5	−3269234.63	1407.876	4.207
triphenylmethane	76236039.7	−1716464.79	−2378.221	7.648
benzo[*a*]pyrene[a]	362132279	−4462435	3363.69	1.852
benzo[*e*]pyrene	361778000.9	−4467130.9	3358.284	1.935
benzo[*k*]fluoranthene	340115133.8	−4248262.11	2729.014	2.689
3,4-benzofluoroanthene	326007471.4	−4123674.37	2417.768	2.985
benz[*e*]acephenanthrylene	344544546.3	−4296000.47	2953.31	2.393
benzo[*ghi*]perylene	404342365.1	−4955693.15	3885.116	1.278
1,2:3,4 dibenzanthracene[b]	422327238.1	−5189596.52	4340.621	1.277
1,2:5,6 dibenzanthracene[b]	423491900.7	−5187595.06	4420.352	1.118
p-quaterphenyl	375623111.3	−4787612.49	2943.198	3.038
1,3,5-triphenylbenzene[b]	353704496.4	−4572620.47	2525.485	3.679

[a] Calculated as an average from the temperature dependence of retention time using the standards in run 14 and the standards reported in ref 5. [b] Calculated from the data reported in ref 5.

Also, note that benz[*e*]acephenanthrylene and 3,4-benzofluoroanthene are the same material. The vapor pressures calculated for the two using the constants given in Table 8 give an indication of the reproducibility of the results from two different sources. The results for 3,4-benzofluoroanthene are the mean of five sets of measurements: for benz[*e*]acephenanthrylene, the data are the results of a single set of measurements. The retention times for benz[*e*]acephenanthrylene were included in a previous mixture but not reported.[5] Also note that the retention times for triacontane are incorrectly inserted two rows below where they should appear in Table 8 of ref 5 and that the retention times for dotriacontane have been omitted. The retention times as they should appear in ref 5 and their temperature dependence are provided below.

Retention Times for Benz[*e*]acephenanthrylene Not Previously Reported in Reference 5

T/K	532.5	537.5	542.55	547.5	552.6	557.6	562.6
				t/min			
methylene chloride	0.262	0.261	0.261	0.260	0.262	0.265	0.250
hexacosane	3.601	3.088	2.676	2.331	2.040	1.803	1.598
octacosane	5.87	4.954	4.221	3.615	3.113	2.703	2.355
benz[*e*]acephenanthrylene	6.281	5.458	4.778	4.202	3.705	3.290	2.929
benzo[*a*]pyrene	7.615	6.595	5.757	5.043	4.436	3.923	3.541
1,3,5-triphenylbenzene	8.62	7.325	6.276	5.404	4.680	4.111	3.541
triacontane	9.657	8.022	6.735	5.680	4.830	4.111	3.541
dotriacontane	15.873	13.011	10.766	8.960	7.503	6.313	5.341

	slope		
	K	intercept	r^2
hexacosane	-9051.9 ± 83	15.80 ± 0.007	0.9996
octacosane	-9761.8 ± 71	16.614 ± 0.006	0.9997
triacontane	-10448 ± 67	17.387 ± 0.006	0.9998
dotriacontane	-11141 ± 60	18.18 ± 0.005	0.9998
benz[*e*]acephenanthrylene	-8064.9 ± 46	13.355 ± 0.004	0.9998
benzo[*a*]pyrene	-8072.4 ± 116	13.172 ± 0.01	0.9990
1,3,5-triphenylbenzene	-9212.2 ± 70	15.182 ± 0.006	0.9997

We would like to thank Dr. Rob Chirico for bringing this to our attention.

JE800660Y

10.1021/je800660y
Published on Web 10/08/2008

Solubility of Tridodecylamine in Supercritical Carbon Dioxide. Mohammad Kaboudvand and Hassan S. Ghaziaskar,* *J. Chem. Eng. Data* **2008,** *53,* 1841–1845.

Page 1844. Figure 4 should be replaced by:

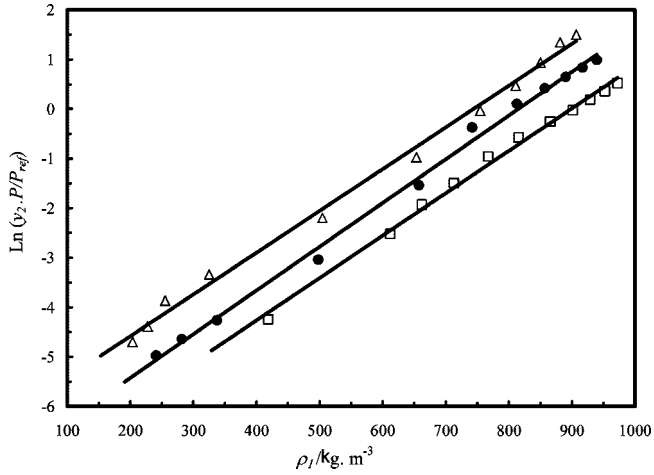

Figure 4. Correlation plots of $\ln(y_2 \cdot P/P_{ref})$ versus ρ_1 for tridodecylamine in the pressure range of (8.0 to 40.0) MPa: □, $T = 308.1$ K; ●, $T = 318.1$ K; △, $T = 328.1$ K using the Bartle equation. The linear regression lines and the experimental points are shown on the graph.

Reference 24 should be replaced by

Literature Cited

(24) Bartle, K. D.; Clifford, A. A.; Jafar, S. A. Solubility of solids and liquids of low volatility in supercritical CO_2. *J. Phys. Chem. Ref. Data* **1991**, *20*, 713–757.

JE8007233

10.1021/je8007233
Published on Web 10/21/2008