

**Viscosity and Density of Carbon Dioxide + 2,6,10,15,19,23-Hexamethyltetracosane (Squalane).** Fausto Ciotta, Geoffrey Maitland, Matthieu Smietana, J. P. Martin Trusler,\* and Velisa Vesovic, *J. Chem. Eng. Data* **2009**, 54, 2436–2443.

Tables 5 and 8 contained several errors. The corrected tables are given below.

**Table 5. Coefficients of Equations 1 to 3 for Viscosity and Statistical Parameters**

coefficient	$x_1 = 0$	$x_1 = 0.4233$	$x_1 = 0.6039$	$x_1 = 0.7882$
$\ln[A_\eta/(\text{mPa}\cdot\text{s})]$	$-2.62 \cdot 10^0$	$-3.28 \cdot 10^0$	$-2.48 \cdot 10^0$	$-4.42 \cdot 10^0$
$B_\eta/\text{K}$	$7.50 \cdot 10^2$	$9.32 \cdot 10^2$	$5.62 \cdot 10^2$	$1.29 \cdot 10^3$
$C_\eta/\text{K}$	$-1.72 \cdot 10^2$	$-1.28 \cdot 10^2$	$-1.60 \cdot 10^2$	$-4.11 \cdot 10^1$
$d_0$	$-5.94 \cdot 10^{-1}$	$2.54 \cdot 10^1$	$4.99 \cdot 10^0$	$1.37 \cdot 10^1$
$d_1$	$-1.39 \cdot 10^3$	$-2.12 \cdot 10^4$	$-4.96 \cdot 10^3$	$-1.11 \cdot 10^4$
$d_2$	$1.20 \cdot 10^6$	$4.77 \cdot 10^6$	$1.47 \cdot 10^6$	$2.51 \cdot 10^6$
$e_0$	$7.07 \cdot 10^2$	$3.36 \cdot 10^3$	$1.71 \cdot 10^3$	$3.15 \cdot 10^3$
$e_1$	$-8.25 \cdot 10^{-1}$	$-1.47 \cdot 10^1$	$-6.75 \cdot 10^0$	$-1.39 \cdot 10^1$
$e_2$	$-8.70 \cdot 10^{-4}$	$1.68 \cdot 10^{-2}$	$6.95 \cdot 10^{-3}$	$1.58 \cdot 10^{-2}$
$10^2 \Delta_{\text{AAD},\eta}$	0.80	0.43	0.77	1.41
$10^2 \Delta_{\text{MAD},\eta}$	1.97	1.31	3.70	4.08
$10^2 \Delta_{\text{Bias},\eta}$	0.13	0.10	0.12	-0.06

**Table 8. Deviation of Literature Density Data for Squalane from Equation 7**

source	$10^2 \Delta_{\text{AAD},\rho}$	$10^2 \Delta_{\text{bias},\rho}$	$10^2 \Delta_{\text{MAD},\rho}$
Tomida et al. <sup>10</sup>	0.33	0.29	0.71
Pensado et al. <sup>15</sup>	0.12	0.12	0.32

JE1008153

10.1021/je1008153

Published on Web 08/18/2010

**Liquid–Liquid Coexistence Curves of {x 1-Butyl-3-methylimidazolium Tetrafluoroborate + (1 - x) 1,3-Propanediol} and {x 1-Butyl-3-methylimidazolium Tetrafluoroborate + (1 - x) 1,4-Butanediol}.** Hekun Lv, Yimin Guo, Xueqin An, and Weiguo Shen,\* *J. Chem. Eng. Data* **2010**, 55, 2484–2488.

In lines 22 and 23 of Table 5 of the original paper, the refractive index data for  $x = 0.686$  were mistyped. The corrected Table 5 is given below:

**Table 5. Refractive Indexes  $n$  at Wavelength  $\lambda = 632.8$  nm for {x [C<sub>4</sub>mim][BF<sub>4</sub>] + (1 - x) 1,4-Butanediol} at Various Compositions and Temperatures**

$T/\text{K}$	$n$	$T/\text{K}$	$n$	$T/\text{K}$	$n$
$x = 0.094$					
306.251	1.4367	308.240	1.4361	310.169	1.4355
307.203	1.4364	309.231	1.4358	312.271	1.4349
314.325	1.4343				
$x = 0.190$					
309.545	1.4322	312.778	1.4312	316.135	1.4302
311.248	1.4317	314.291	1.4308	317.296	1.4299
$x = 0.286$					
309.877	1.4289	312.995	1.4279	315.792	1.4271
311.224	1.4285	314.394	1.4275	317.348	1.4266
$x = 0.381$					
309.264	1.4264	312.315	1.4255	315.655	1.4246
310.639	1.4260	313.793	1.4251	317.383	1.4241
$x = 0.484$					
306.699	1.4247	309.725	1.4238	312.876	1.4229
308.217	1.4243	311.331	1.4234	314.312	1.4225
$x = 0.579$					
304.220	1.4237	308.275	1.4225	312.195	1.4214
306.241	1.4231	310.223	1.4220	314.294	1.4208
$x = 0.686$					
303.255	1.4223	307.292	1.4212	311.217	1.4202
305.261	1.4218	309.202	1.4207	313.283	1.4196
$x = 0.786$					
301.241	1.4216	304.235	1.4208	309.241	1.4195
303.162	1.4211	307.212	1.4200	311.854	1.4188
$x = 0.894$					
296.244	1.4221	300.129	1.4211	304.141	1.4200
298.195	1.4216	302.155	1.4206	306.686	1.4194

The coefficient of the  $x^4$  term on the right side of eq 4 was mistyped. The correct equation is:

$$n(T^0 = 298.15 \text{ K}, x) = 1.4436 - 0.0526x + 0.0654x^2 - 0.0885x^3 + 0.0888x^4 - 0.0368x^5 \quad (4)$$

The authors apologize for these mistakes.

JE100709E

10.1021/je100709e

Published on Web 07/19/2010