ADDITIONS AND CORRECTIONS

Evaluation of the Nonrandom Hydrogen Bonding (NRHB) Theory and the Simplified Perturbed-Chain-Statistical Associating Fluid Theory (sPC-SAFT). 2. Liquid−Liquid Equilibria and Prediction of Monomer Fraction in Hydrogen Bonding Systems. Ioannis Tsivintzelis, Andreas Grenner, Ioannis G. Economou,* and Georgios M. Kontogeorgis

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Page 5657. Equation 17 and the points that represent experimental data in Figure 8a, as previously published, are incorrect.

Considering that ρ in eq 15 of the original paper is the molar density, eq 17 should be

$$K = v^* \exp\left(\frac{-\Delta G^H_{AB}}{RT}\right) \tag{17}$$

Figure 8a should illustrate experimental data, as well as the model predictions, for the fraction of non-hydrogen-bonded molecules in pure methanol. However, Luck[1] reported experimental data for the fraction of the "free −OH" groups, which, according to von Solms et al.,[2] was identified as the fraction of the unbonded hydrogen atoms of the hydroxyl groups. Details for the transformation of the experimental data in fractions of unbonded molecules (monomers) can be found in the literature.[2,3] The revised Figure 8a, which illustrates the correct experimental values for the monomer fraction in pure methanol, is

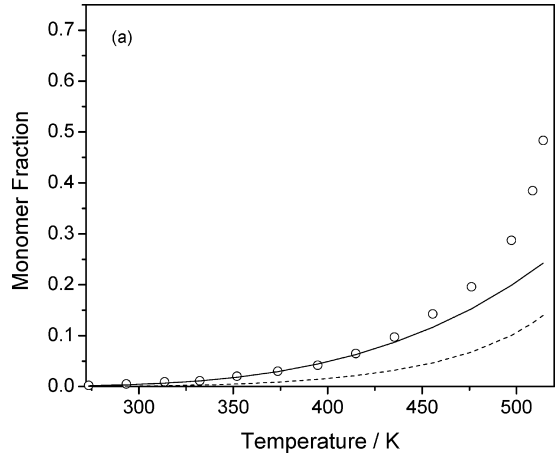

Figure 8. (a) Fraction of non-hydrogen-bonded molecules in pure methanol. Experimental data (points), NRHB (solid lines), and sPC-SAFT predictions (dashed lines).

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Literature Cited

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(3) von Solms, N.; Jensen, L.; Kofod, J. L.; Michelsen, M. L.; Kontogeorgis, G. M. Measurement and modelling of hydrogen bonding in 1-alkanol + *n*-alkane binary mixtures. *Fluid Phase Equilib.* **2007**, *261*, 272.

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