Additions and Corrections

2008, Volume 42, Pages 1315–1323

Purnendu K. Dasgupta,* Yining Liu, and Jason V. Dyke: Iodine Nutrition: Iodine Content of Iodized Salt in the United States

The abstract of this article contains the following statement regarding freshly opened iodized salt containers: "Forty-seven of 88 samples fell below the USFDA recommended I content while 6 exceeded it." However, the left column of page 1318 contains the following conflicting statement at the bottom: "Some 46 samples (52%) fell below the USFDA suggested iodization range (which is admittedly higher than that in most other countries) while 6 of 88 samples (7%) exceeded the recommended range."

The statement on page 1318 is in error. The corrected statement should read as follows: "Some 47 samples (53%) fell below the USFDA suggested iodization range..." In principle, the correct information was available in Figure 3. We apologize for this inadvertent error and thank the research staff of *Body and Soul* magazine for pointing this out to the authors.

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2008, Volume 42, Pages 4133–4139

Troy M. Benn* and Paul Westerhoff: Nanoparticle Silver Released into Water from Commercially Available Sock Fabrics

1. Page 4137, Figure 5. The Freundlich isotherm data describing the adsorption of silver to wastewater treatment biomass was erroneously calculated. The correct data are presented here in Figure 5.

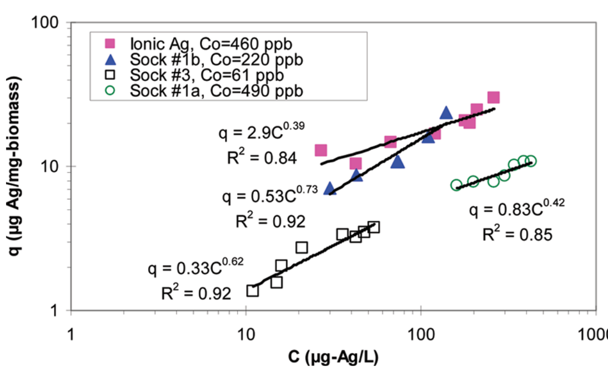

FIGURE 5. Batch adsorption isotherms for the wash solutions of three socks (#1b, #3, #1a) and an ionic silver solution (Ionic Ag). Initial silver concentrations varied from 61 to 490 ppb, and pH values ranged from 5.8 to 7.4.

2. Pages 4136−4137. The values for the Freundlich adsorption capacity and intensity parameters, K and $1/n$, were mistakenly reported. The correct values of K ranged from 0.33 and 2.9 (μg-Ag/g-biomass)(L/μg-Ag)$^{1/n}$. The correct average value for $1/n$ is 0.54 (unitless).

3. Page 4138, Figure 6. The silver concentration in the effluent of a WWTP would exceed the EPA salt and fresh water criteria maximum concentrations (CMCs) at lower influent concentrations than initially reported.

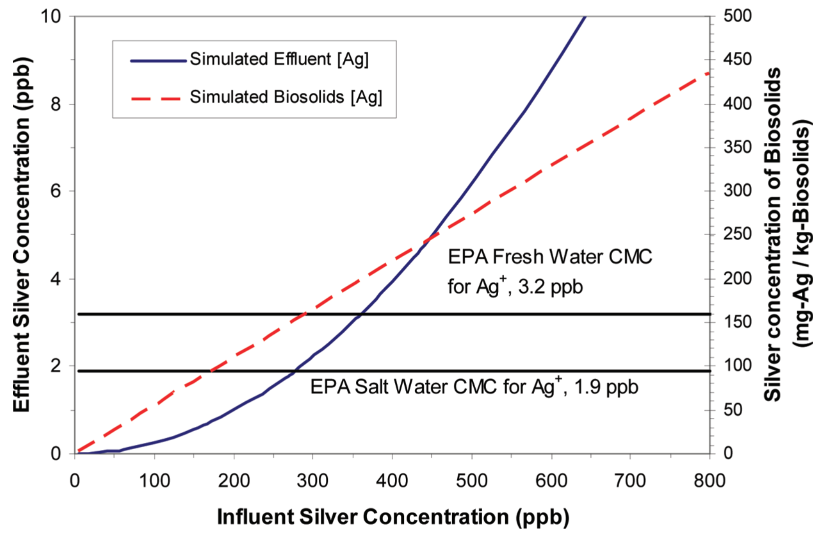

FIGURE 6. Model results illustrating the removal of influent silver for a typical WWTP (model parameters Freundlich K and $1/n = 1.0$, 0.5; $\tau = 0.5$ d; $\theta = 5$ d; $X = 2000$ mg/L). The silver concentration in the treated effluent would exceed the EPA salt and fresh water criteria maximum concentrations (CMCs) at influent concentrations of 280 and 360 ppb, respectively. The concentration of silver in the waste activated sludge flow is represented by the dashed line and the secondary *y*-axis.

4. Page 4138. The revised WWTP model for silver removal leads to the following corrections:

- Using a common municipal WWTP influent silver concentration of 5 μg/L, the model results in an effluent silver concentration of 0.001 μg/L, and the wasted biosolids silver concentration would be 2.8 mg-Ag/kg-biosolids.
- WWTP effluent would exceed the EPA CMC for salt and fresh water at influent silver concentrations of about 280 and 360 ppb, respectively. These influent concentrations are 2 orders of magnitude higher than those commonly observed for municipal WWTPs.
- WWTP effluent would not exceed the 100 ppb secondary drinking water standard for silver until the influent concentration reached approximately 2100 ppb (not shown graphically).
- By controlling mixed liquor suspended solids (MLSS) from 2000 to 4000 mg/L and the θ:τ ratio from 5 to 20, a WWTP could treat a range of influent silver concentrations between 140 and 1110 μg/L to less than 1.9 μg/L.

Supporting Information Available

Table S1 showing WWTP maximum allowable influent silver concentrations, C_0 (ppb), that would result in an effluent silver concentration lower than 1.9 ppb as a function of operational parameters MLSS and θ:τ. This material is available free of charge via the Internet at http://pubs.acs.org.

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