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ADDITIONS AND CORRECTIONS

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Andrew W. Yau and Huw O. Pritchard*: Toward a Unified Master-Equation Theory of Thermal Decomposition Reactions. Analytic Solution for Diatomic Dissociation. Pages 134–149.

Editors Note: We reprint herewith Figures 3 and 7 of this article which was published in the January 11, 1979 issue. These figures as originally printed in this issue lacked important detail. Sometimes we receive poor figures or drawings from our authors; in this case Professor Pritchard supplied us with perfectly good figures, and errors were made in the printing process. We apologize to Professor Pritchard and to our readers for this unhappy result which we correct herewith.—Bryce Crawford, Jr.

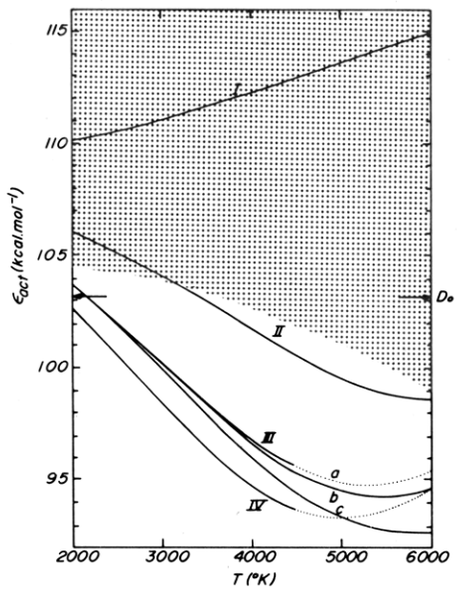

Figure 3. Arrhenius activation energy ϵ_{act} as a function of temperature for various models of the dissociation of H_2 in Ar. $\epsilon_{act} = d(\ln k_{diss})/d(1/kT)$ where the dissociation rate constant is calculated as follows: (I) equilibrium, SSH model, without rotation ($k_{eq,0}$), i.e., curve III of Figure 2; (II) nonequilibrium, SSH model, without rotation ($k_{diss,0}$) i.e., curve IV of Figure 2; (III) rotationally averaged nonequilibrium SSH model A, assuming (a) equilibrium in J, (b) slight disequilibrium in J (cutoff at $J \geq 30$), (c) severe disequilibrium in J (cutoff at $J \geq 20$); (IV) rotationally averaged nonequilibrium SSH model B, assuming equilibrium in J, as in IIIa above.

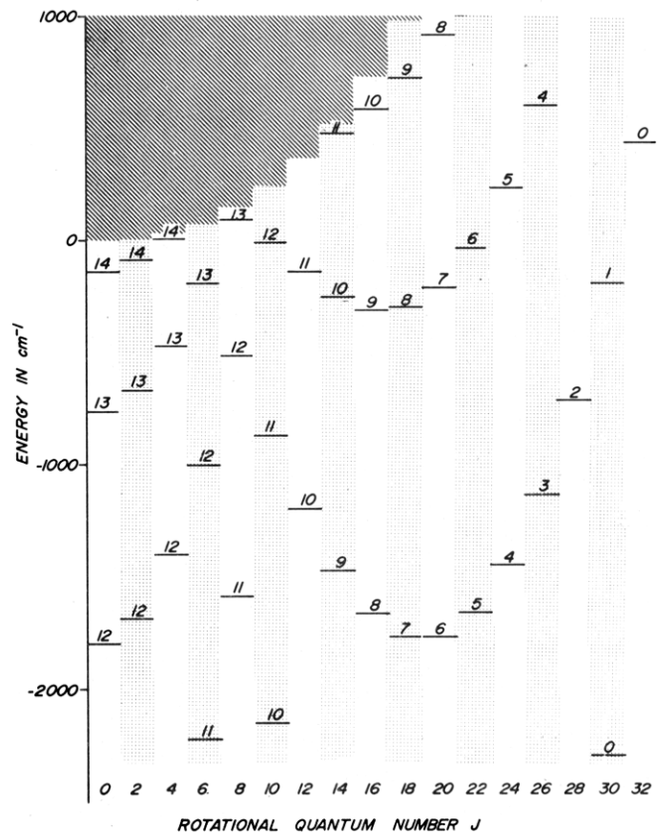

Figure 7. Energy-level diagram for para-H_2 in the ground electronic state. The vibrational quantum number v is given for each vibrational level. The energy-level diagram would be appropriate for calculations in the temperature range of 50–100 K.