Determination of Adsorption Isotherms of Hydrogen on Titanium in Sulfuric Acid Solution Using the Phase-Shift Method and Correlation Constants. Jin Y. Chun and Jang H. Chun,* *J. Chem. Eng. Data* **2009,** *54,* 1236–1243.

In our previously published paper, the reversible hydrogen electrode (RHE) scale should be corrected as the standard hydrogen electrode (SHE) scale. The authors regret that the error has been published. The correction does not affect the results and discussion in the paper.

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Isobaric Vapor−Liquid Equilibria for (Acetic Acid + Cyclohexane) and (Cyclohexane + Acetylacetone) at a Pressure of 101.3 kPa and for (Acetic Acid + Acetylacetone) at a Pressure of 60.0 kPa. Zhidong Yang, Jiawen Zhu, Bin Wu,* Kui Chen, and Xiaohe Ye, *J. Chem. Eng. Data* **2010,** *55,* 1527–1531.

In the second and third column of Table 1 of the above paper, the experimental and literature values for the density of cyclohexane at a temperature of 298.15 K are reported. Because of an error in the conversion of temperature values in degrees Celsius to Kelvin, these data are in fact for cyclohexane density at 293.15 K. The correct data for cyclohexane density at 298.15 K are shown.

The authors apologize for this mistake.

Table 1

	ρ/g·cm^{-3} (298.15 K)		n_D (293.15 K)		T_b/K (1 atm)	
compound	exptl	lit.	exptl	lit.	exptl	lit.
acetylacetone	0.9717	0.9721	1.4489	1.4494	413.60	413.55
cyclohexane	0.7734	0.7739	1.4259	1.4264	353.90	353.87
acetic acid	1.0469	1.0446	1.3715	1.3720	391.44	391.50

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Heat Capacities of Binary Mixtures of Acetic Acid with Acetic Anhydride and Methenamine at Different Temperatures. Li-Yan Dai,* Qian Li, Ming Lei, and Ying-Qi Chen, *J. Chem. Eng. Data* **2010,** *55,* 1704–1707.

In Table 5 of the original paper, the heat capacity value of the binary mixture of methenamine (1) with acetic acid (2) at x_1 = 0.2221, T = 328.15 K is reported. Because of a transcription error, this piece of data is wrong. The correction data are shown.

Table 5

T/K	C_p/J·mol^{-1}·K^{-1}	T/K	C_p/J·mol^{-1}·K^{-1}
	$x_1 = 0.0220$		
308.15	137.35	323.15	152.45
313.15	142.53	328.15	157.24
318.15	147.51	333.15	161.64
	$x_1 = 0.0454$		
308.15	144.80	323.15	160.18
313.15	150.32	328.15	165.08
318.15	155.25	333.15	169.05
	$x_1 = 0.0703$		
308.15	151.94	323.15	166.78
313.15	157.16	328.15	172.00
318.15	162.14	333.15	176.96
	$x_1 = 0.0967$		
308.15	159.76	323.15	175.70
313.15	165.58	328.15	180.88
318.15	170.38	333.15	185.28
	$x_1 = 0.1249$		
308.15	170.64	323.15	186.63
313.15	176.29	328.15	191.96
318.15	181.92	333.15	197.37
	$x_1 = 0.1551$		
308.15	180.97	323.15	198.28
313.15	187.61	328.15	204.18
318.15	193.63	333.15	209.36
	$x_1 = 0.1873$		
308.15	191.09	323.15	209.59
313.15	198.06	328.15	214.96
318.15	203.66	333.15	220.93
	$x_1 = 0.2221$		
308.15	203.69	323.15	222.76
313.15	210.89	328.15	229.35
318.15	216.63	333.15	235.17

The authors apologize for this mistake.

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