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Addition/Correction

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Correction to "Activity Coefficients at Infinite Dilution for Hydrocarbons in Fatty Alcohols Determined by Gas–Liquid Chromatography"

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In the original document (*J. Chem. Eng. Data*, **2011**, *56* (4), 850–858) on pages 851 and 852, the values tabulated in Tables 1 and 2 correspond to natural logarithm of the activity coefficients at infinite dilution ($\ln \gamma^{\infty}$) and not the values of γ^{∞}. In consequence Tables 3 and 4 and Figures 2 to 18 given in the manuscript are not correct. The resulting correct tables and figures are reported below. As a consequence some sentences are also corrected in the text.

The sentence of page 855 which says in part, "The γ^{∞} values are relatively slight (< 1) in the two solvents"; is not correct. The correct sentence is "The γ^{∞} values are larger than unity (> 1) in the two solvents".

The correct sentence (in second paragraph) of page 857 is "The values of *s* show their high polarity" instead of the incorrect sentence "The values of *s* show their weak polarity."

Table 3. Values of $\Delta H^{E,\infty}$, ΔH^{dis}, and ΔH^{vap} as Determined from Equations 8 to 10 in the Two Solvents

	octadecanol			eicosanol			literature
	$\Delta H^{E,\infty}$	ΔH^{dis}	ΔH^{vap}	$\Delta H^{E,\infty}$	ΔH^{dis}	ΔH^{vap}	ΔH_{vap}
solute	kJ·mol^{-1}	kJ·mol^{-1}	kJ·mol^{-1}	kJ·mol^{-1}	kJ·mol^{-1}	kJ·mol^{-1}	kJ·mol^{-1}
hexane	1.259	−29.127	30.387	0.926	−28.409	29.335	28.84[a]
ethanol	0.590	−41.000	41.590	1.985	−37.744	39.728	39.33[b]
tetrachloride	0.049	−30.352	30.400	0.982	−29.163	30.144	33.47[b]
butyl acetate	0.810	−39.935	40.744	1.643	−39.003	40.647	43.89[c]
toluene	0.258	−35.501	35.760	11.342	−23.919	35.262	33.43[a]
diethyl ether	0.323	−25.007	25.330	0.528	−24.435	24.963	27.37[c]

[a]Reference 20. [b]Reference 32. [c]Reference 33.

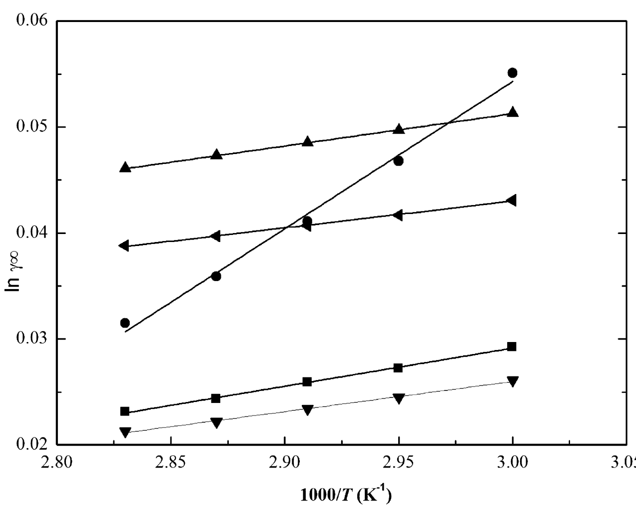

Figure 2. Experimental activity coefficients at infinite dilution for the alkanes and olefins in the solvent Octadecanol as a function of temperature; ■, pentane; ●, hexane; ▲, octane; ▼, pentene; ◀, heptene; ——, linear regression.

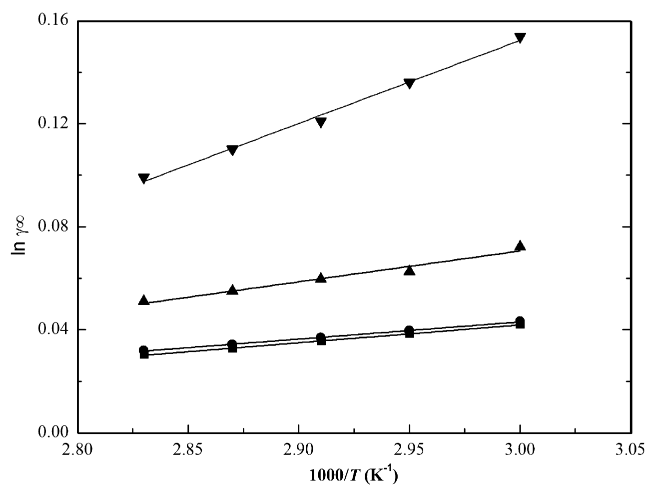

Figure 3. Experimental activity coefficients at infinite dilution for the alcohols in the solvent octadecanol as a function of temperature; ■, methanol; ●, ethanol; ▲, propanol; ▼, butanol; ——, linear regression.

The correct sentence (in the last paragraph) of page 857 is "The Q_{12}^{∞} values are higher for octadecanol (Figure 18)..." instead of the incorrect sentence "Comparing the Q_{12}^{∞} values of octadecanol with eicosanol (Figure 18), the Q_{12}^{∞} values of octadecanol are almost twenty times higher."

In the caption of Figure 1, it should be "eicosanol " instead of "eicoanol".

The authors apologize for the mistakes.

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Table 4. Abraham Parameters for Solvents Studied at Different Temperatures[a]

solvent	T/K	c	r	s	a	b	l	R^2
octadecanol	333.15	0.159	−0.751	1.361	2.699	−0.737	0.699	0.992
	338.15	0.146	−0.741	1.339	2.591	−0.736	0.679	0.992
	343.15	0.132	−0.728	1.316	2.484	−0.735	0.660	0.993
	348.15	0.120	−0.716	1.293	2.381	−0.734	0.642	0.993
	353.15	0.107	−0.703	1.271	2.280	−0.732	0.624	0.993
eicosanol	338.15	0.531	−1.373	2.035	2.535	−1.282	0.499	0.989
	343.15	0.482	−1.282	1.951	2.420	−1.215	0.492	0.986
	348.15	0.440	−1.207	1.873	2.318	−1.163	0.483	0.984
	353.15	0.402	−1.143	1.805	2.224	−1.120	0.475	0.981

[a] R^2 is the correlation coefficient.

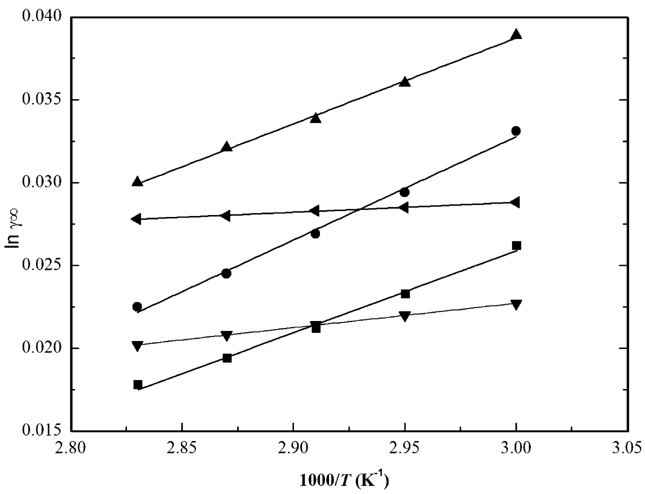

Figure 4. Experimental activity coefficients at infinite dilution for the chloroalkanes in the solvent octadecanol as a function of temperature; ■, dichloromethane; ●, chloropropane; ▲, chlorobutane; ▼, chloroform; ♦, tetrachloromethane; ——, linear regression.

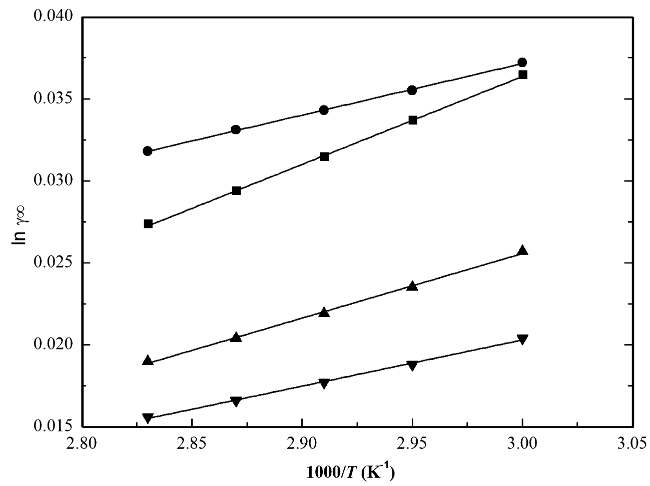

Figure 6. Experimental activity coefficients at infinite dilution for the cycloalkane, aromatic and ethers in the solvent octadecanol as a function of temperature; ■, cyclohexane; ●, toluene; ▲, diethyl ether; ▼, THF; ——, linear regression.

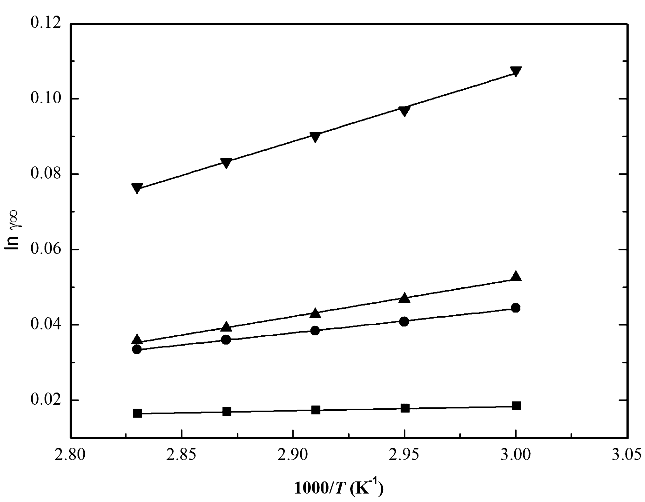

Figure 5. Experimental activity coefficients at infinite dilution for the ketone and acetates in the solvent octadecanol as a function of temperature; ■, propanone; ●, ethyl acetate; ▲, butyl acetate; ▼, acetonitrile; ——, linear regression.

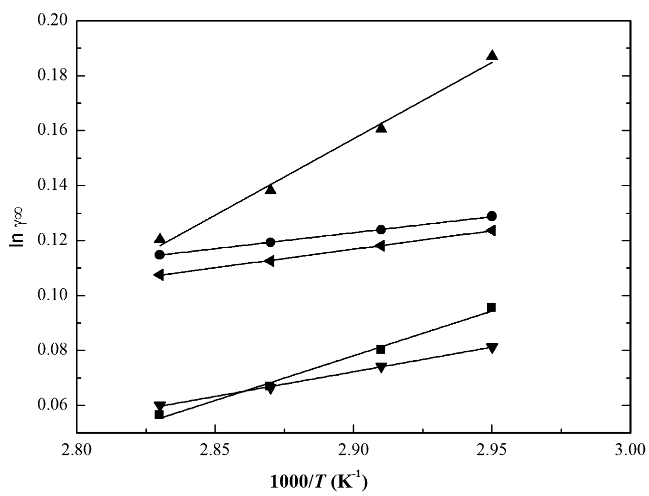

Figure 7. Experimental activity coefficients at infinite dilution for the alkanes and olefins in the solvent eicosanol as a function of temperature; ■, pentane; ●, hexane; ▲, octane; ▼, pentene; ◀, heptene; ——, linear regression.

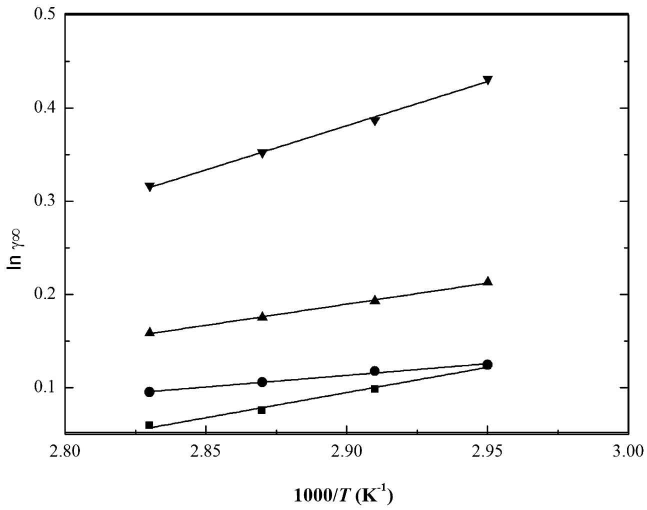

Figure 8. Experimental activity coefficients at infinite dilution for the alcohols in the solvent eicosanol as a function of temperature; ■, methanol; ●, ethanol; ▲, propanol; ▼, butanol; ——, linear regression.

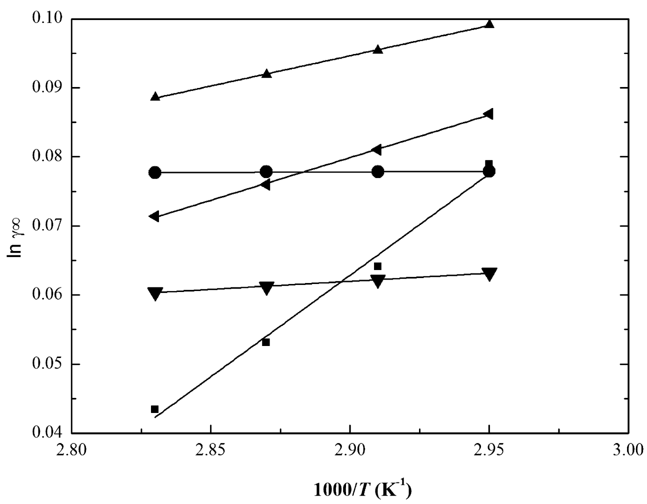

Figure 9. Experimental activity coefficients at infinite dilution for the chloroalkanes in the solvent eicosanol as a function of temperature; ■, dichloromethane; ●, chloropropane; ▲, chlorobutane; ▼, chloroform; ◀, tetrachloromethane; ——, linear regression.

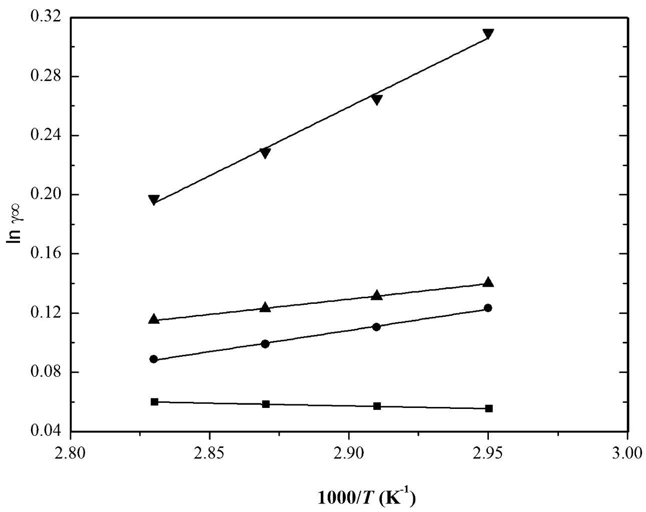

Figure 10. Experimental activity coefficients at infinite dilution for the ketone and acetates in the solvent eicosanol as a function of temperature; ■, propanone; ●, ethyl acetate; ▲, butyl acetate; ▼, acetonitrile; ——, linear regression.

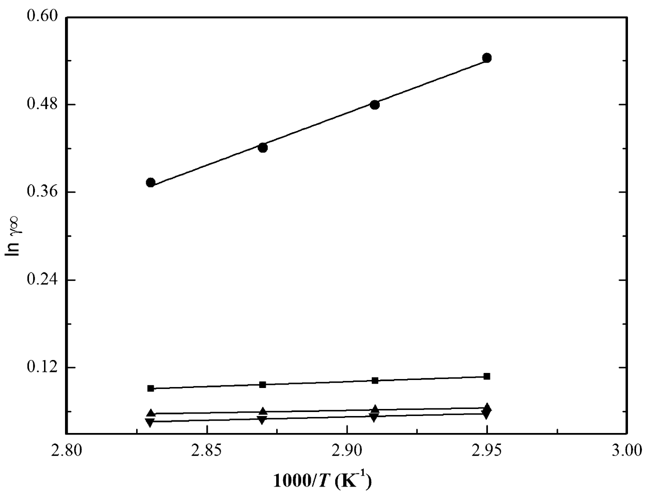

Figure 11. Experimental activity coefficients at infinite dilution for the cycloalkane, aromatic and ethers in the solvent eicosanol as a function of temperature; ■, cyclohexane; ●, toluene; ▲, diethyl ether; ▼, THF; ——, linear regression.

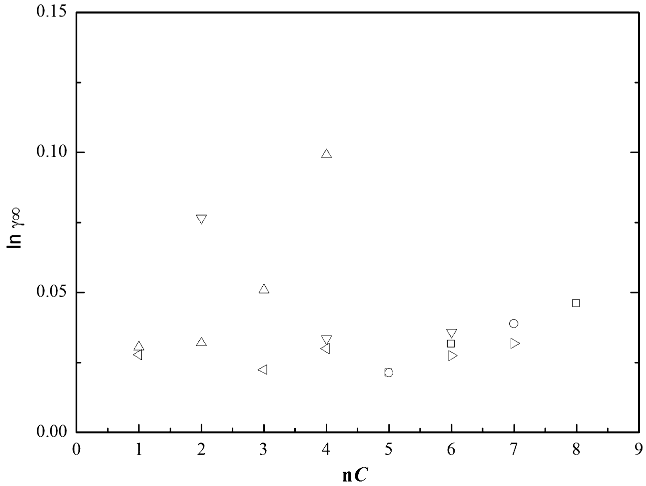

Figure 12. Variation of experimental activity coefficients for the solutes in the solvent octadecanol according to the carbon's atom numbers at T = 353.15 K. □, alkanes; ○, olefins; △, alcohols; ▽, acetate; ◁, chloroalkanes; ▷, aromatics and cycloalkane.

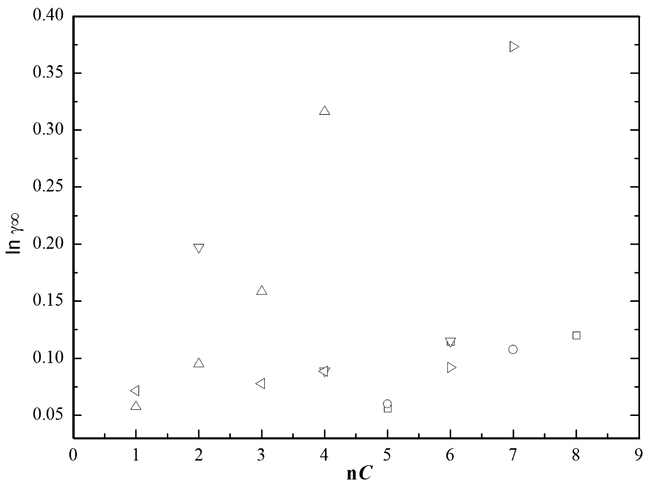

Figure 13. Variation of experimental activity coefficients for the solutes in the solvent eicosanol according to the carbon's atom numbers at T = 353.15 K. □, alkanes; ○, olefins; △, alcohols; ▽, acetate; ◁, chloroalkanes; ▷, aromatics and cycloalkane.

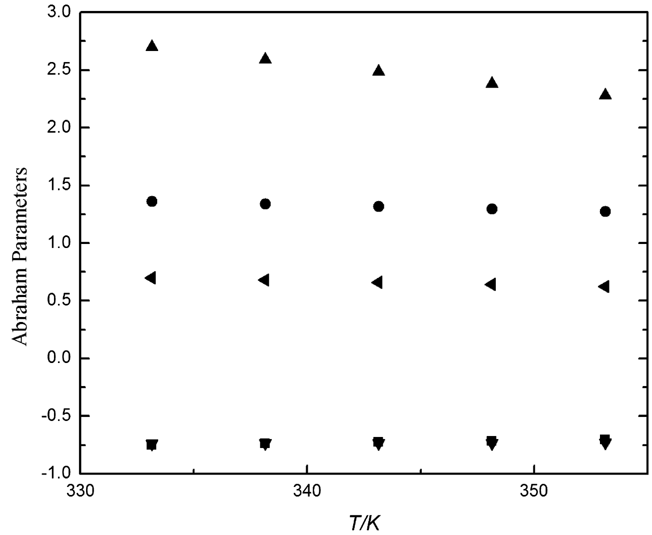

Figure 14. Evolution of Abraham parameters with temperature for octadecanol. ■, r; ●, s; ▲, a; ▼, b; ◀, l.

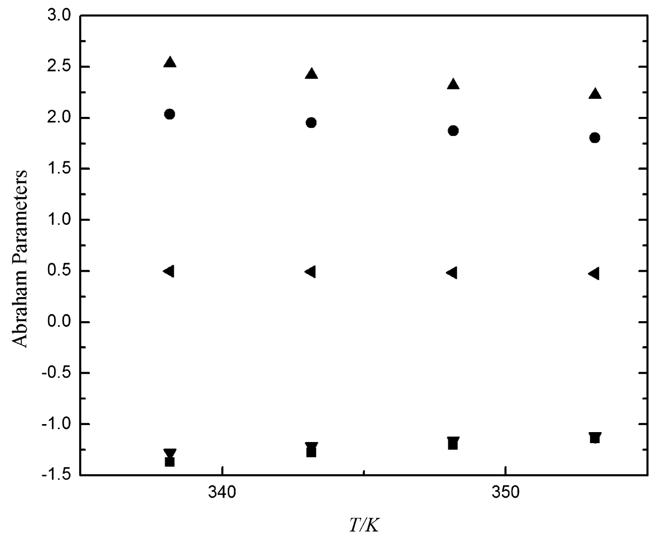

Figure 15. Evolution of Abraham parameters with temperature for eicosanol. ■, r; ●, s; ▲, a; ▼, b; ◀, l.

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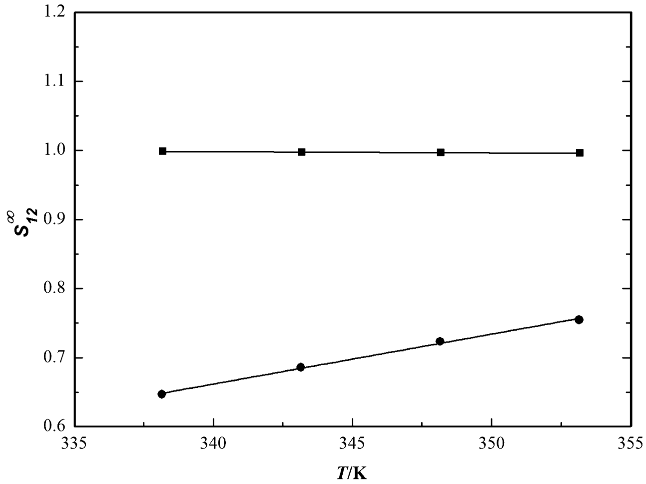

Figure 16. Selectivity at infinite dilution for two solvents at different temperatures regarding the separation of cyclohexane from toluene: ■, in octadecanol; ●, in eicosanol; ——, linear regression.

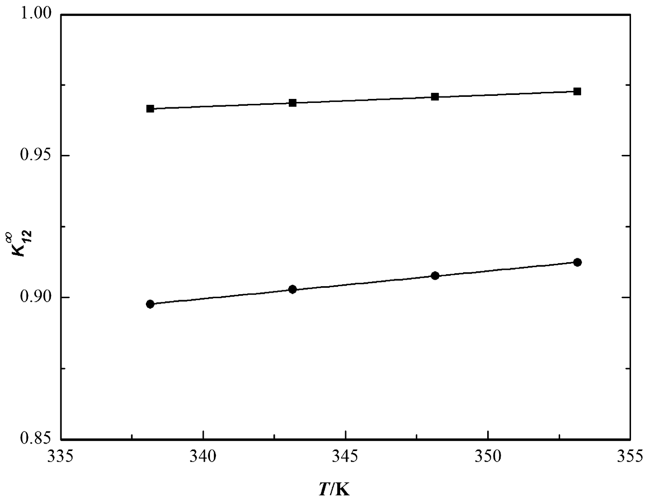

Figure 17. Capacity at infinite dilution for two solvents at different temperatures regarding the separation of cyclohexane from toluene: ■, in octadecanol; ●, in eicosanol; ——, linear regression.

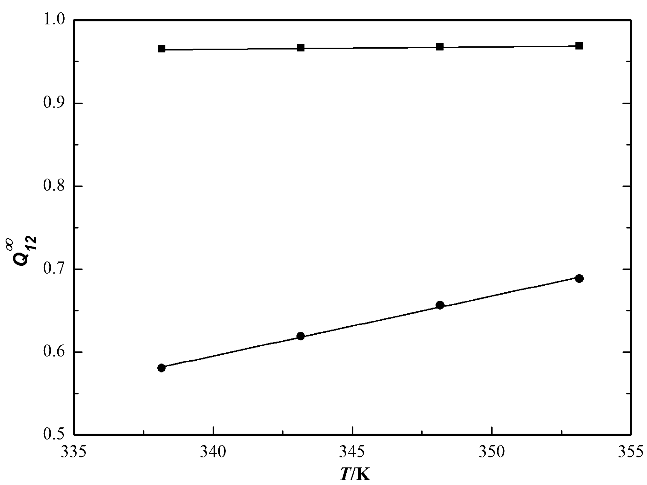

Figure 18. Quantity at infinite dilution for two solvents at different temperatures regarding the separation of cyclohexane from toluene: ■, in octadecanol; ●, in eicosanol; ——, linear regression.